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Chemical vapor deposition and temperature-dependent Raman characterization of two-dimensional vanadium ditelluride

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Recently, ultrathin two-dimensional (2D) metallic vanadium dichalcogenides have attracted widespread attention because of the charge density wave (CDW) phase transition and possible ferromagnetism. Herein, we report the synthesis and temperature-dependent Raman characterization of the 2D vanadium ditelluride (VTe₂). The synthesis is done by atmospheric pressure chemical vapor deposition (APCVD) using vanadium chloride (VCl₃) precursor on fluorophlogopite mica, sapphire, and h-BN substrates. A large area of the thin film with thickness ~10 nm is grown on the hexagonal boron nitride (h-BN) substrate. Temperature-dependent Raman characterization of VTe₂ is conducted from room temperature to 513 K. Remarkable changes of Raman modes at around 413 K are observed, indicating the structural phase transition.

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1. Introduction

Metallic transition metal dichalcogenides (TMDs) with fascinating properties in their bulk counterpart such as charge density waves (CDW),^{1–7} superconductivity,^{8–10} and magnetism,^{11–13} have attracted huge interest over several decades. Charge density wave (CDW) is an incident originating from the interaction of electron–phonon at one and two dimensions, in which the phase transition is generally associated with an obvious change of the resistivity and magnetic susceptibility.^{14–17} CDW has drawn special attention because of its competition with superconductivity at low temperatures.¹⁸ Dissimilar from superconductivity states, CDW states can remain at room temperature and above,¹⁷ which extends the probability for phase transition devices being handled at room temperature. In recent years, a wide range of potential applications such as an oscillator,¹⁹ nonvolatile memory,²⁰ and high sensitive room-temperature bolometer²¹ have been reported based on CDW phase transition.

Vanadium ditelluride (VTe₂) is interesting as of now rarely investigated 2D CDW materials.^{22–24} VTe₂ is a member of group VB TMDs with one layer of V atoms sandwiched between two layers of tellurium atoms, stacked together by weak van der

Waals forces. VTe₂ shows exceptional magnetic characteristics at the 2D limit and CDW property.^{25,26} Earlier study of VTe₂ exhibits the structural phase transition at 482 K from the high-temperature 1T phase to the low-temperature 1T' phase and the CDW order of 3 × 1 × 3 in the 1T' phase.^{27,28} Room temperature ferromagnetism and two possible CDW transitions have been reported in a few-layer 1T-VTe₂ at 135 and 240 K.^{25,26} Ohtani *et al.* reported the CDW transition temperature for VTe₂ is about 474 K by electrical and magnetic measurements.²⁹ Won *et al.* reported the CDW phase transition of VTe₂ single crystal at 420 K synthesized by flux method.³⁰ However, the periods of CDW orders, electronic structure, and CDW mechanisms of VTe₂ remain unclear.

Despite several investigations on vanadium ditelluride (VTe₂) due to the unique structural phase transition property, the details synthesis process is still lacking. Typically 2D VTe₂ materials are obtained *via* chemical vapor transport (CVT) and molecular beam epitaxy (MBE) method. CVT usually needs high pressures and long reaction times. MBE is an effective way to prepare high-quality 2D materials. But, the lack of cost and time-efficiency restricts the scalable application of this technique. Accordingly, developing facile, convenient, and batch producible synthetic routes for VTe₂ material is highly necessary. In recent years, to synthesis 2D materials, chemical vapor deposition (CVD) is considered one of the main processing methods. High crystalline quality and impurity-free samples with the potential for scaling up are the fundamental advantages of the CVD method. Here, we report the facile synthesis of 2D VTe₂ single crystals on fluorophlogopite mica and sapphire substrates and large area VTe₂ thin film on the h-BN substrate by APCVD method. The synthesized VTe₂ nanosheets are characterized by optical microscope (OM), scanning electron

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microscopy (SEM), atomic force microscopy (AFM), energy-dispersive spectroscopy (EDS), X-ray photoelectron spectroscopy (XPS), and Raman spectroscopy. Moreover, the structural phase transition is verified in the synthesized VTe_2 nanosheets by the temperature dependent Raman characterization.

2. Experimental

2.1 Materials

VCl_3 (99.99%) and Te (99.999%) were purchased from Alfa Aesar. All the starting materials were analytical grade and were used without further purification.

2.2 CVD growth process of VTe_2

A quartz tube with a diameter of 1 inch is used to synthesis the VTe_2 single crystal. The length of the furnace is about 46 cm. A mixture of H_2/Ar was used as the carrier gas. Two quartz boat with a volume of about $8 \text{ cm} \times 1.1 \text{ cm} \times 1.2 \text{ cm}$ (each boat) containing VCl_3 and Te powders were placed upstream of the furnace at 300°C and 450°C respectively. Sapphire or freshly cleaved (001) facet of mica or h-BN/ SiO_2/Si substrates were placed in the downstream area of the furnace. The distance between the precursors and the substrates ranges from 5 cm to 21 cm. To purge the reaction chamber, 500 sccm Ar flow was introduced for 30 minutes before the heating process and then followed by 100 sccm Ar and 20 sccm H_2 for growth. With a ramp rate of $20^\circ\text{C min}^{-1}$, the furnace was heated to the growing temperature 800°C . The growth time was 15 minutes. The furnace was naturally cooled down.

2.3 Characterization of VTe_2

The VTe_2 crystals were characterized using optical microscopy (Olympus BX51M), SEM (Hitachi S-4800; acceleration voltage of 5.0 kV) equipped with an energy-dispersive spectrometer, AFM (DI 3100 multimode microscope), and XPS (ESCALAB250Xi). Raman spectra were measured on a home-built vacuum, variable temperature, low-wavenumber Raman system with 532 nm excitation. A NA = 0.82 low-temperature objective (LT-APO/VIS/

0.82, attocube systems AG, Germany) was used for laser focusing and signal collection. Below $1 \text{ mW } \mu\text{m}^{-2}$ the laser power was kept. Amplified-spontaneous emission (ASE) filters, beam splitter, and notch filters (Ondax Inc., USA) were used to achieve low-wavenumber detection.

2.4 Device fabrication and electrical measurement

A copper TEM grid was used as a shadow mask for metal evaporation. As contact electrodes, 15 nm Ti and 50 nm Au were used. Devices were put in the air and measured by a Keithley 4200-SCS semiconductor device analyzer.

3. Results and discussion

The synthesis procedure of VTe_2 flakes with the APCVD method is schematically illustrated in Fig. 1a. Briefly, VCl_3 and Te powders were used as the precursors. Fluorophlogopite mica/sapphire/h-BN substrates were placed in different positions in the downstream part of the furnace and reaction was carried out usually at 800°C . Fig. 1b shows the crystal structure of the 1T phase VTe_2 . In the space group, $P\bar{3}m1$ (164) with lattice constants of $a = b = 3.636 \text{ \AA}$ and $c = 6.51 \text{ \AA}$ VTe_2 is crystallized

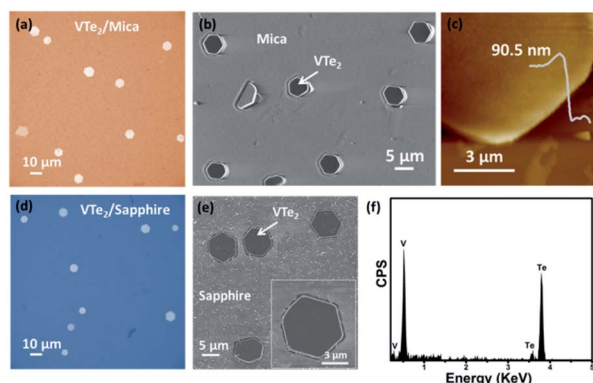


Fig. 2 (a) Optical microscopy (b) scanning electron microscopy (SEM) and (c) atomic force microscopy (AFM) images of as-synthesized VTe_2 single crystals on fluorophlogopite mica substrates. (d and e) Optical and SEM images of the as-synthesized VTe_2 single crystals on sapphire substrates respectively. (f) Elemental analysis of VTe_2 crystal on sapphire substrate by energy-dispersive spectroscopy (EDS).

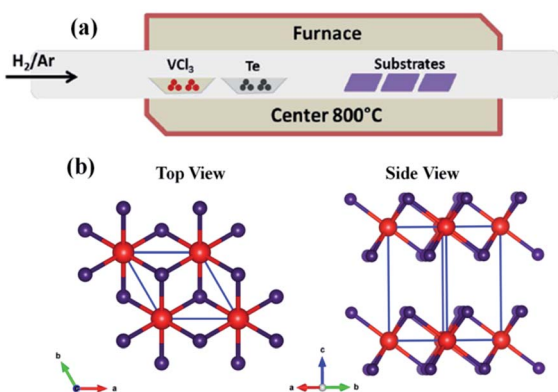


Fig. 1 (a) Schematic of synthesis setup and (b) the crystal structure of 1T- VTe_2 . Red and violet balls are indicating vanadium(V), and tellurium (Te) atoms respectively.

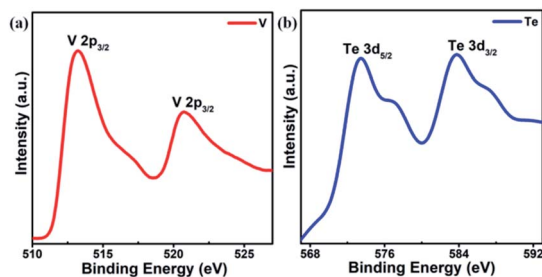


Fig. 3 XPS spectra of 1T- VTe_2 . (a) The $\text{V } 2p_{3/2}$ and $2p_{1/2}$ peaks are seen to position at 513.1 eV and 520.7 eV respectively. (b) The $\text{Te } 3d_{5/2}$ and $3d_{3/2}$ peaks are at 573.3 eV and 583.7 eV respectively.



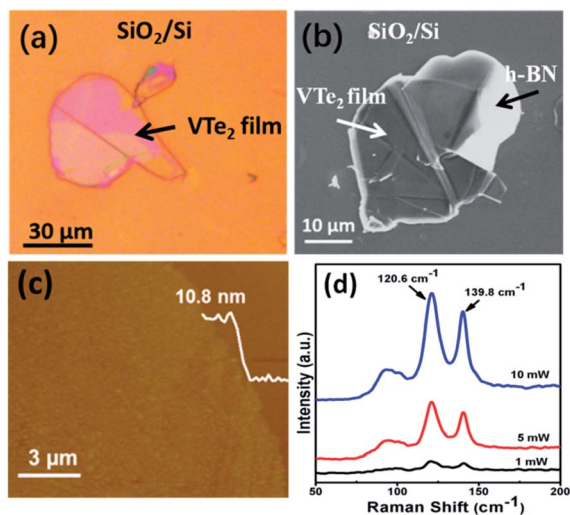


Fig. 4 (a) Optical microscopy, (b) scanning electron microscopy (SEM), (c) atomic force microscopy (AFM) images and (d) room-temperature Raman spectra of as synthesized VTe_2 thin film on the h-BN substrate.

in a trigonal layered structure.²⁵ In octahedral coordination, every monolayer is composed of a layer of hexagonally organized vanadium ions, sandwiched between two layers of tellurium ions. The layers are held jointly by way of weak van der Waals forces.

The morphology of CVD grown VTe_2 crystals on mica and sapphire is shown in Fig. 2a–e. We were able to synthesize VTe_2 single crystals on mica and sapphire substrates at the suitable growth parameters (see details in the Experimental section). It is noticeable that the hexagon is the dominant shape of VTe_2 crystals (Fig. 2a and d) implying the single-crystalline nature of the samples. Atomic force microscopy (AFM) was used to measure the thickness of as-synthesized VTe_2 crystal. Fig. 2c shows the thickness of VTe_2 on the fluorophlogopite mica substrate was 90.5 nm. Energy-dispersive spectroscopy (EDS) was employed to identify the chemical composition of the synthesized VTe_2 crystals on the sapphire substrate. Both V and Te elements are visualized by the EDS profile (Fig. 2f).

Fig. 3 shows the X-ray photoelectron spectroscopy (XPS) spectra of 1T- VTe_2 single crystal grown on fluorophlogopite mica. The binding energies of V $2p_{3/2}$ and V $2p_{1/2}$ are 513.1 eV and 520.7 eV respectively are near to those in 1T- VS_2 and 1T- VSe_2 ^{31–34} and lower than that in VO_2 .³⁵ The binding energies of Te $3d_{5/2}$ and Te $3d_{3/2}$ are 573.3 eV and 583.7 eV respectively comparable to the oxidation state of Te. In VTe_2 , the binding energy difference (ΔE) between V $2p_{3/2}$ and Te $3d_{5/2}$ is 60.2 eV, which is near to that of V and Te primary substance (60.7 eV), demonstrating that the V–Te bonds in VTe_2 are weak because of the small electronegativity difference between V and Te.

Moreover, we have investigated the growth of 2D VTe_2 on h-BN/ SiO_2 / Si substrate (Fig. 4a–c). It has been seen that large area VTe_2 thin film (size ~ 30 – $50 \mu\text{m}$) with thickness $\sim 10.8 \text{ nm}$ were grown on h-BN substrate applying the same growth condition

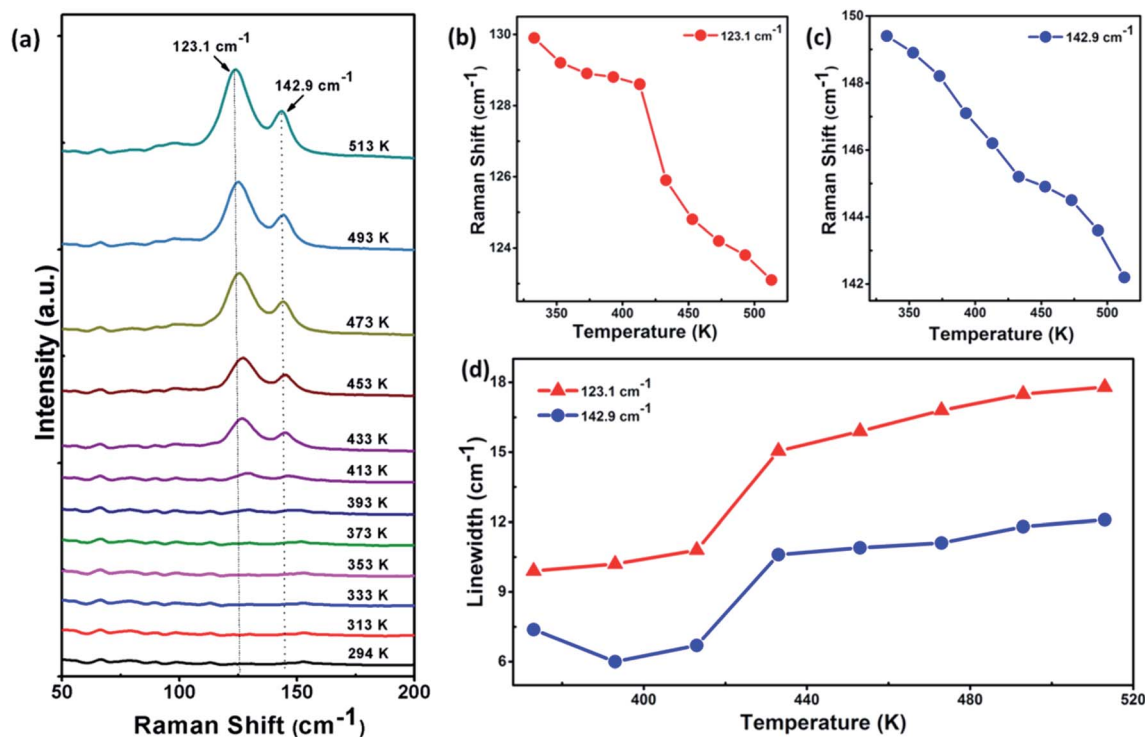


Fig. 5 (a) Temperature-dependent Raman spectra of VTe_2 single crystal at different temperature between 513 and 294 K. (b and c) Temperature dependent Raman shift of the Raman modes 123.1, 142.9 cm^{-1} . (d) Changes in linewidth of the Raman modes 123.1, 142.9 cm^{-1} as a function of temperature.



Table 1 Comparison of the reported structural phase transition temperature of VTe₂

Synthesis method	Characterization technique	Phase transition temperature	Reference
CVD	Raman spectroscopy	413 K	This work
MBE	Raman spectroscopy and electrical measurements	135 K and 240 K	26
CVT	Electrical and magnetic measurements	473 K	29
Flux method	Electrical measurement	420 K	30
CVT	Electrical and magnetic measurements	475 K	52

used on sapphire and fluorophlogopite mica. Fig. 4d shows the room temperature Raman spectra of VTe₂ film on the h-BN substrate at different laser intensities. At the laser intensity of 10 mW, three main peaks are observable at 92.6, 120.6, and 139.8 cm⁻¹. The Raman spectrum of VTe₂ crystal is consistent with the previous reports.²⁶ According to earlier room temperature Raman measurements, the observable Raman peaks at 120.6 and 139.8 cm⁻¹ are in plane E_g and out of plane A_{1g} modes respectively.^{25,26} The weak Raman peak at 92.6 cm⁻¹ could be related to the E_g phonons possibly coming from symmetry points.^{36,37}

In recent years to explore the CDW phase transitions in TMDs, different research tools have been employed, such as scanning tunneling microscopy (STM),^{31,38–40} transmission electron microscopy (TEM),^{41,42} and angle-resolved photoemission spectroscopy (ARPES),^{31,40,43} and Raman spectroscopy. Raman spectroscopy is a non-destructive, high sensitive micro-area analysis method. It has been adopted as a very reliable technique to investigate the vibrational properties of 2D TMDs materials. Over the past few years, temperature-dependent Raman spectra of 1T-VS₂,³² 1T-VSe₂,^{32,44–46} 1T-TaS₂,²¹ 1T-TaSe₂,⁴⁷ 2H-TaSe₂,^{48,49} 1T-TiSe₂,⁵⁰ and 2H-NbSe₂⁵¹ have been reported. Fig. 5a shows the temperature-dependent Raman

spectra of VTe₂ single crystal (thickness of ~90 nm on sapphire) in the temperature range between 513 and 294 K. There are two Raman peaks observed at around 123.1, 142.9 cm⁻¹ corresponding to the E_g and A_{1g} vibration modes, respectively. It is noticeable that the overall Raman signal of VTe₂ crystal is very weak from room temperature to 393 K, but become strong and sharp when the temperature reached at 413 K. This phenomenon was reproducible to several run and samples. The frequencies of the observed Raman modes are plotted as a function of temperature (Fig. 5b and c). The energy changes of the Raman modes 123.1, 142.9 cm⁻¹ are more explicit with reducing the temperature. It has seen that the temperature dependence of these Raman modes' frequencies is almost the same and exhibits remarkable frequency changes at about 413 K, indicating that VTe₂ undergoes a structural phase transition. This is also confirmed by the changes in the linewidth of 123.1, 142.9 cm⁻¹ Raman modes. In Fig. 5d, the changes in the linewidth of these Raman modes are plotted as a function of temperature and shows a monotonous line expanding from the temperature at 413 K. This result is consistent with the observed CDW phase transition at 420 K by Won *et al.*³⁰ but different from the observed structural phase transition at 474 K and 482 K.^{26,29} Table 1 summarizes the contrast of phase transition temperatures of VTe₂ growth by different methods and characterized using different techniques.

Furthermore, the electrical measurement was conducted for as-synthesized VTe₂ single crystal (thickness ~90 nm, sample length ≈ 13.63 μm, and width ≈ 6.14 μm) on a mica substrate. Fig. 6 shows a linear I_{ds}-V_{ds} curve with a resistance of ≈ 185 Ω and estimated conductivity is ≈ 1.3 × 10⁵ S m⁻¹ at room temperature.

4. Conclusions

In conclusion, we have successfully synthesized 2D VTe₂ single crystals on mica and sapphire substrates and large area VTe₂ film on the h-BN substrate *via* the APCVD method. Temperature-dependent Raman shift, the linewidth of the 123.1, 142.9 cm⁻¹ Raman modes of VTe₂ crystal show the structural phase transition at around 413 K. Our investigation will provide the fundamental spectroscopy for 2D VTe₂ crystal.

Conflicts of interest

There are no conflicts to declare.

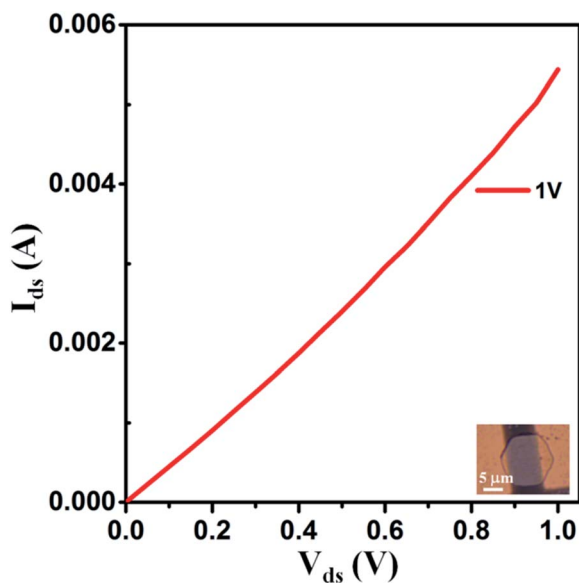


Fig. 6 Current (I_{ds}) versus source-drain voltage (V_{ds}) of as synthesized VTe₂ single crystal. Sample length ≈ 13.63 μm and width ≈ 6.14 μm.



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References

- 1 J. Wilson, F. Di Salvo and S. Mahajan, *Adv. Phys.*, 1975, **24**, 117.
- 2 T. Straub, T. Finteis, R. Claessen, P. Steiner, S. Hufner, P. Blaha, C. S. Oglesby and E. Bucher, *Phys. Rev. Lett.*, 1999, **82**, 4504.
- 3 F. Weber, S. Rosenkranz, J. P. Castellan, R. Osborn, R. Hott, R. Heid, K. P. Bohnen, T. Egami, A. H. Said and D. Reznik, *Phys. Rev. Lett.*, 2011, **107**, 107403.
- 4 M. Eichberger, H. Schafer, M. Krumova, M. Beyer, J. Demsar, H. Berger, G. Moriena, G. Sciaini and R. J. D. Miller, *Nature*, 2010, **468**, 799.
- 5 C. Dang, M. Guan, S. Hussain, W. Wen, Y. Zhu, L. Jiao, S. Meng and L. Xie, *Nano Lett.*, 2020, **20**(9), 6725–6731.
- 6 W. Wen, C. Dang and L. Xie, *Chinese Phys. B*, 2019, **28**(5), 058504.
- 7 W. Wen, Y. Zhu, C. Dang, W. Chen and L. Xie, *Nano Lett.*, 2019, **19**(2), 1805–1813.
- 8 T. Yokoya, T. Kiss, A. Chainani, S. Shin, M. Nohara and H. Takagi, *Science*, 2001, **294**, 2518–2520.
- 9 B. Sipos, A. F. Kusmartseva, A. Akrap, H. Berger, L. Forro and E. Tutis, *Nat. Mater.*, 2008, **7**, 960.
- 10 R. Ang, Y. Tanaka, E. Ieki, K. Nakayama, T. Sato, L. J. Li, W. J. Lu, Y. P. Sun and T. Takahashi, *Phys. Rev. Lett.*, 2012, **109**, 176403.
- 11 K. Xu, P. Chen, X. Li, C. Wu, Y. Guo, J. Zhao, X. Wu and Y. Xie, *Angew. Chem., Int. Ed.*, 2013, **52**, 10477.
- 12 X. Zhu, Y. Guo, H. Cheng, J. Dai, X. An, J. Zhao, K. Tian, S. Wei, X. C. Zeng, C. Wu and Y. Xie, *Nat. Commun.*, 2016, **7**, 11210.
- 13 H. L. Zhuang and R. G. Hennig, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2016, **93**, 054429.
- 14 M. Dressel, *Naturwissenschaften*, 2007, **94**, 527.
- 15 P. Monceau, in *Physics and Chemistry of Low-Dimensional Inorganic Conductors*, ed. C. Schlenker, J. Dumas, M. Greenblatt and S. van Smaalen, Plenum Press Div Plenum Publishing Corp, New York, 1996, vol. 354, p. 371.
- 16 K. Rossnagel, *J. Phys.: Condens. Matter*, 2011, **23**, 213001.
- 17 J. A. Wilson, F. J. Di Salvo and S. Mahajan, *Adv. Phys.*, 2001, **50**, 1171.
- 18 A. M. Gabovich, A. I. Voitenko and M. Ausloos, *Phys. Rep.*, 2002, **367**, 583.
- 19 G. Liu, B. Debnath, T. R. Pope, T. T. Salguero, R. K. Lake and A. A. Balandin, *Nat. Nanotechnol.*, 2016, **11**, 845–850.
- 20 I. Vaskivskiy, I. A. Mihailovic, S. Brazovskii, J. Gospodaric, T. Mertelj, D. Svetin, P. Sutar and D. Mihailovic, *Nat. Commun.*, 2016, **7**, 11442.
- 21 X. Wang, H. Liu, J. Wu, J. Lin, W. He, H. Wang, X. Shi, K. Suenaga and L. Xie, *Adv. Mater.*, 2018, **30**, 1800074.
- 22 M. Chhowalla, H. S. Shin, G. Eda, L.-J. Li, K. P. Loh and H. Zhang, *Nat. Chem.*, 2013, **5**, 263.
- 23 Y. Ma, Y. Dai, M. Guo, C. Niu, Y. Zhu and B. Huang, *ACS Nano*, 2012, **6**, 1695.
- 24 K. Xu, P. Chen, X. Li, C. Wu, Y. Guo, J. Zhao, X. Wu and Y. Xie, *Angew. Chem., Int. Ed.*, 2013, **52**, 10477.
- 25 J. Li, B. Zhao, P. Chen, R. Wu, B. Li, Q. Xia, G. Guo, J. Luo, K. Zang, Z. Zhang, H. Ma, G. Sun, X. Duan and X. Duan, *Adv. Mater.*, 2018, **30**, 1801043.
- 26 X. Ma, T. Dai, S. Dang, S. Kang, X. Chen, W. Zhou, G. Wang, H. Li, P. Hu, Z. He, Y. Sun, D. Li, F. Yu, X. Zhou, H. Chen, X. Chen, S. Wu and S. Li, *ACS Appl. Mater. Interfaces*, 2019, **11**, 10729.
- 27 K. Bronsema, G. Bus and G. Wiegers, *J. Solid State Chem.*, 1984, **53**, 415.
- 28 A. Nakamura, T. Shimojima, M. Matsuura, Y. Chiashi, M. Kamitani, H. Sakai, S. Ishiwata, H. Li, A. Oshiyama and K. Ishizaka, *Appl. Phys. Express*, 2018, **11**, 092601.
- 29 T. Ohtani, S. Onoue and M. Nakahira, *Mater. Res. Bull.*, 1984, **19**, 1367–1375.
- 30 <http://meetings.aps.org/link/BAPS.2018.MAR.C18.5>.
- 31 M. Bonilla, S. Kolekar, Y. Ma, H. C. Diaz, V. Kalappattil, R. Das, T. Eggers, H. R. Gutierrez, M. H. Phan and M. Batzill, *Nat. Nanotechnol.*, 2018, **13**, 289–293.
- 32 M. Hossain, J. Wu, W. Wen, H. Liu, X. Wang and L. Xie, *Adv. Mater. Interfaces*, 2018, 1800528.
- 33 X. Chia, A. Ambrosi, P. Lazar, Z. Sofer and M. Pumera, *J. Mater. Chem. A*, 2016, **4**, 14241–14253.
- 34 J. Zhang, C. Zhang, Z. Wang, J. Zhu, Z. Wen, X. Zhao, X. Zhang, J. Xu and Z. Lu, *Small*, 2018, **14**, 1703098.
- 35 D. Malarde, I. D. Johnson, I. J. Godfrey, M. J. Powell, G. Cibir, R. Quesada-Cabrera, J. A. Darr, C. J. Carmalt, G. Sankar, I. P. Parkin and R. G. Palgrave, *J. Mater. Chem. C*, 2018, **6**, 11731–11739.
- 36 R. Samnakay, D. Wickramaratne, T. R. Pope, R. K. Lake, T. T. Salguero and A. A. Balandin, *Nano Lett.*, 2015, **15**, 2965–2973.
- 37 H. Wang, Y. Chen, M. Duchamp, Q. Zeng, X. Wang, S. H. Tsang, H. Li, L. Jing, T. Yu, E. H. T. Teo and Z. Liu, *Adv. Mater.*, 2018, **30**, 1704382.
- 38 C. J. Arguello, E. P. Rosenthal, E. F. Andrade, W. Jin, P. C. Yeh, N. Zaki, S. Jia, R. J. Cava, R. M. Fernandes, A. J. Millis, T. Valla, R. M. Osgood and A. N. Pasupathy, *Phys. Rev. Lett.*, 2015, **114**, 037001.
- 39 J. Feng, A. Tan, S. Wagner, J. Liu, Z. Mao, X. Ke and P. Zhang, *Appl. Phys. Lett.*, 2016, **109**, 021901.
- 40 H. Ryu, Y. Chen, H. Kim, H. Z. Tsai, S. Tang, J. Jiang, F. Liou, S. Kahn, C. Jia, A. A. Omrani, J. H. Shim, Z. Hussain, Z. X. Shen, K. Kim, B. I. Min, C. Hwang, M. F. Crommie and S. K. Mo, *Nano Lett.*, 2018, **18**, 689–694.
- 41 J. Shi, X. Chen, L. Zhao, Y. Gong, M. Hong, Y. Huan, Z. Zhang, P. Yang, Y. Li, Q. Zhang, Q. Zhang, L. Gu, H. Chen, J. Wang, S. Deng, N. Xu and Y. Zhang, *Adv. Mater.*, 2018, 1804616.
- 42 W. T. Adam, R. Hovden, D. Wang, Y. D. Kim, J. Okamoto, K. A. Spoth, Y. Liu, W. Lu, Y. Sun, J. C. Hone,



- L. F. Kourkoutis, P. Kim and A. N. Pasupathy, *Proc. Natl. Acad. Sci. U. S. A.*, 2015, **112**, 15054.
- 43 L. Bawden, S. P. Cooil, F. Mazzola, J. M. Riley, L. J. Collins-McIntyre, V. Sunko, K. W. B. Hunvik, M. Leandersson, C. M. Polley, T. Balasubramanian, T. K. Kim, M. Hoesch, J. W. Wells, G. Balakrishnan, M. S. Bahramy and P. D. C. King, *Nat. Commun.*, 2016, **7**, 11711.
- 44 M. Hossain, Z. Zhao, W. Wen, X. Wang, J. Wu and L. Xie, *Crystals*, 2017, **7**, 298.
- 45 B. Dardel, M. Grioni, D. Malterre, P. Weibel, Y. Baer and F. Levy, *J. Phys.: Condens. Matter*, 1993, **5**, 6111–6120.
- 46 M. M. Ugeda, A. J. Bradley, Y. Zhang, S. Onishi, Y. Chen, W. Ruan, C. O. Aristizaba, H. Ryu, T. Mark, M. T. Edmonds, H. Z. Tsai, A. Riss, S. K. Mo, D. Lee, A. Zettl, Z. Hussain, Z. X. Shen and M. F. Crommie, *Nat. Phys.*, 2016, **12**, 92–97.
- 47 H. X. Luo, W. W. Xie, J. Tao, H. Inoue, A. Gyenis, J. W. Krizan, A. Yazdani, Y. M. Zhu and R. J. Cava, *Proc. Natl. Acad. Sci. U. S. A.*, 2015, **112**, E1174–E1180.
- 48 E. Morosan, H. W. Zandbergen, B. S. Dennis, J. W. G. Bos, Y. Onose, T. Klimczuk, A. P. Ramirez, N. P. Ong and R. J. Cava, *Nat. Phys.*, 2006, **2**, 544.
- 49 T. Shimada, F. S. Ohuchi and A. Koma, *Surf. Sci.*, 1993, **291**, 57–66.
- 50 P. Goli, J. Khan, D. Wickramaratne, R. K. Lake and A. A. Balandin, *Nano Lett.*, 2012, **12**, 5941–5945.
- 51 X. Xi, L. Zhao, Z. Wang, H. Berger, L. Forró, J. Shan and K. F. Mak, *Nat. Nanotechnol.*, 2015, **10**, 765–769.
- 52 T. Ohtani, K. Hayashi, M. Nakahira and H. Nozaki, *Solid State Commun.*, 1981, **40**, 629–631.

