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Visible-light-promoted cross-coupling reaction of hypervalent bis-catecholato silicon compounds with selenosulfonates or thiosulfonates†

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Under mild metal-free conditions, we successfully developed a visible-light-promoted free radical cross coupling of hypervalent bis-catecholato silicon compounds with selenosulfonates or thiosulfonates and efficiently constructed unsymmetrical 1°-alkyl-alkyl selenide compounds which are difficult to prepare. This protocol uses mild reaction conditions and easily available substrates, has a wide range substrate scope, high yields, and simple operation and no need of expensive photosensitive catalysts. It provides a green and simple method for the preparation of unsymmetrical alkyl-alkyl selenide compounds. At the same time, this strategy is also applicable for the efficient construction of aryl-alkyl selenides, heteroayl-alkyl selenides and aryl-alkyl sulfides.

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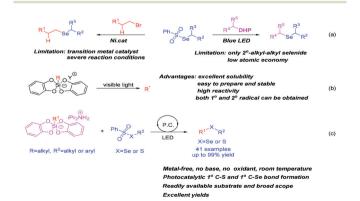
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Introduction

Selenium is an essential trace element for the human body, and it mainly exists in internal organs and blood in the form of selenoamino acids and peptides in organisms. Selenium is the basic component of some enzyme active parts, and affects the activity and function of enzymes. In the human body, it participates in the formation of thyroxine deiodinase, glutathione peroxidase and other proteins in the form of selenocysteine.2 In recent years, many commercial selenium supplement products have appeared on the market. Relevant studies have shown that adequate selenium supplementation is beneficial to health. In addition, organic selenium compounds are widely used in organic chemistry,3 selenium-containing biologically active molecules,⁴ pesticides⁵ materials. It can be seen that the synthesis of organic selenium compounds has important practical significance. In the past few years, selenosulfonates have been widely used in the construction of organoselenide compounds as an important class of reagents and intermediates. Common strategies include transition metal-catalyzed C-Se bond construction⁷ (Scheme 1a, left), C-Se bond construction under metal-free catalysis⁸ and C-Se bond construction under photocatalysis⁹ (Scheme 1a, right). However, these strategies need to use expensive transition metal catalysts or additives. On the other

hand, low atom economy and confined secondary alkyl-alkyl selenides scope limit their further applications. Therefore, it is of great significance to develop a simple, mild and efficient method to prepare organoselenide compounds with wide substrate scope.

In recent years, visible light-mediated photoredox catalysis has received more and more attention. Due to its special reaction mode, many transformations that could not be achieved through the transition metal catalytic process before are possible. However, finding free radical precursors with sufficient stability and functional group compatibility is a big challenge. For instance, organoborates (especially alkyl trifluoroborate), ¹⁰ 4-alkyl-1,4-dihydropyridines (DHPs)¹¹ have been widely applied



Scheme 1 (a) Ni-catalyzed C-Se bond construction and C-Se bond construction under photocatalysis. (b) Hypervalent silicon compounds as a source of C-radicals. (c) Visible-light-promoted free radical cross coupling of hypervalent bis-catecholato silicon compounds with selenosulfonates or thiosulfonates.

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in the synthesis of various compounds. However, the main drawbacks are the formation of polluting boron fluoride side products and the continuing difficulty in generating primary radicals under photocatalytic conditions. Therefore, it is particularly important to find a free radical precursor that is more environmentally friendly and has high substrate universality. In recent years, hypervalent silicon compounds¹² have been widely used in various free radical reaction strategies as a kind of highly active free radical precursor. Kumada et al. 13 first reported the free radical reaction of organic pentafluorosilicate under copper catalysis conditions, however its poor solubility severely limited its application in organic synthesis. The hypervalent silicon derivative of bis-catechol reported by Frye's group has received widespread attention due to its more soluble properties (Scheme 1b).14 These ideal radical precursors can generate various functionalized alkyl radicals especially the primary alkyl radicals smoothly under lowenergy visible light using ready-made photocatalysts, which has attracted wide attention from chemists. Prof. Molander's group¹⁵ and other groups¹⁶ have done a lot of excellent work in this area. Herein, we developed a free radical cross-coupling reaction between bis-catechol hypervalent silicon derivatives and alkylselenosulfonates under visible light mediation, and successfully constructed primary unsymmetrical alkyl-alkyl selenide compounds. This strategy is also applicable to the construction of aryl-alkyl selenides, heteroayl-alkyl selenides and aryl-alkyl sulfides (Scheme 1c). More importantly, we used this strategy to successfully construct the selenocarbamate compound 3r. This shows that the strategy has broad application prospects in the biological field.

Results and discussion

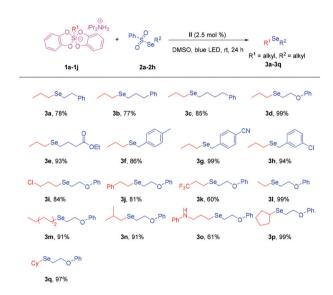
Initially, we tried the reaction of propylsilicate 1a, Se-(2-phenoxyethyl) benzenesulfonoselenoate 2d, and organometallic photocatalyst Ru(bpy)₃(PF₆)₂ in (dry) DMF at room temperature for 14 h under a N₂ atmosphere. Gratifyingly, the reaction proceeded smoothly to give the target product (2-phenoxyethyl) (propyl)selane 3d. With this promising result in hand, we tried to optimize the reaction conditions. First we investigated the ratio effects of 1a and 2a. It was found that when the ratio of 1a to 2a is 1.67:1.0, 3d could be obtained in 64% GC-yield (Table 1, entries 1-4). Then, we tried the reactions in several different dry solvents such as DMA, MeCN, and DMSO. It was found that when anhydrous DMSO was used, the target product 3d could be obtained in 73% GC-yield (Table 1, entries 5-7). We also investigated other different photosensitive catalysts. To our delight, 3d was obtained in 85% GC-yield when the reaction was carried out utilizing acriflavin II as the catalyst (Table 1, entries 8-12). It should be noted that the reaction could not proceed under an air atmosphere. (Table 1, entry 13). 3d was obtained in 99% isolated yield when the reaction time was prolonged to 24 h (Table 1, entry 14).

With the optimized reaction conditions in hand, we set about assessing the generality of this protocol. First, we exam-

Table 1 Optimization of the reaction conditions for the preparation of unsymmetrical alkyl-alkyl selenide compoundsa,k

Sio Ph Se OPh cat. (2.5 mol %) Se OPh Solvent, Blue LED, rt, 14 h						
1a		2d N ₂			3d	
Entry	1a (equiv.)	2a (equiv.) Cat	Solvent	GC-yield ^b	
1	1.00	2.0	1	DMF	45%	
2	1.50	1.0	1	DMF	56%	
3	1.67	1.0	1	DMF	64%	
4	2.50	1.0	1	DMF	58%	
5	2.00	1.0	1	DMA	29%	
6	2.00	1.0	1	DMSO	73%	
7	2.00	1.0	1	MeCN	48%	
8	2.00	1.0	2	DMSO	85%	
9	2.00	1.0	3	DMSO	75%	
10	2.00	1.0	4	DMSO	56%	
11	2.00	1.0	5	DMSO	69%	
12	2.00	1.0	6	DMSO	81%	
13 ^c	2.00	1.0	2	DMSO	N.R	
14^d	2.00	1.0	2	DMSO	99% ^e	

^a Reaction conditions: propylsilicate 1a, Se-(2-phenoxyethyl) benzenesulfonoselenoate 2d; photosensitive catalyst (2.5 mol%), solvent(dry) (1 ml), room temperature, 14 h. ^b Yields were determined by GC-MS analysis using biphenyl as an internal standard, GC-yield. c Under an air atmosphere. d 24 h. e Isolated yield.



Scheme 2 Substrate scope for unsymmetrical alkyl-alkyl selenide compounds. Reaction conditions: substituted silicate 1 (0.34 mmol), substituted selenosulfonates 2 (0.2 mmol), II (2.5 mol%), (dry) DMSO (1 mL), room temperature, blue LED, N₂, 24 h. Isolated yields.

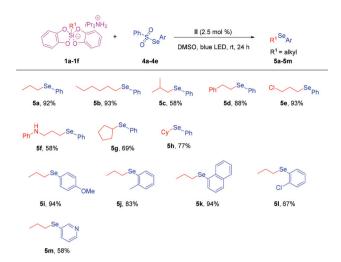
ined the reactivity of different substituted selenosulfonates (Scheme 2). The reaction of propylsilicate 1a with Se-phenethyl benzenesulfonoselenoate 2a gave the target product 3a in 78% yield. The reaction of propylsilicate 1a with Se-(3-phenylpropyl)

benzenesulfonoselenoate 2b afforded the desired product 3b in 77% yield. The reaction of propylsilicate 1a with Se-(4-phenylbutyl) benzenesulfonoselenoate 2c furnished 3c in 85% yield. To our delight, when ethyl 4-((phenylsulfonyl)selanyl) butanoate 2e was applied to the reaction, the target product 3e was obtained in 93% yield successfully, which demonstrates excellent functional group tolerance. Next, we investigated the universality of different substituted benzylselenosulfonates. It should be noted that whether it is a methyl-substituted benzylselenosulfonate 2f or an electron-withdrawing group functionalized benzylselenosulfonate 2g, or a halogen-substituted benzylselenosulfonate 2h were applied to the reactions with 1a, the desired products 3f-h could be obtained in good to excellent yields, respectively.

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Then, we further investigated the universality of different substituted alkylsilicates (Scheme 2). The reaction showed a broad functional group compatibility, in which both halogen substituted alkylsilicate, phenyl substituted alkylsilicate and trifluoromethyl substituted alkylsilicate were well tolerated. The target products 3i-k were uniformly obtained in excellent yields. It was found that the length of the chain has no effect on the yields. The target products 31-n were uniformly obtained in excellent yields too. Surprisingly, the secondary amine substituted alkylsilicate substrate was also well tolerated. The product 30 could be obtained in 61% yield. When cyclopentylsilicate and cyclohexylsilicate were applied to the reaction, the target products 3p and 3q were isolated in excellent yields, respectively.

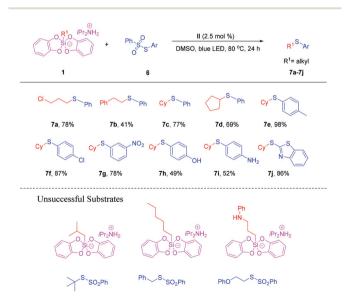
Encouraged by the above achievements, we turned attention to extending the method to arylselenosulfonates. However, under the previous conditions, the reaction of propylsilicate 1a with Se-phenyl benzenesulfonoselenoate 4a failed to afford product 5a. To our satisfaction, when the reaction temperature was increased to 80 °C, it smoothly led to the target product 5a in 92% yield. With the optimal modified reaction conditions in hand, we explored the scope of different substituted alkylsilicates (Scheme 3). Similarly, the reaction showed a broad functional group compatibility. When hexylsilicate was applied to the reaction, the target product 5b could be obtained in 93% yield. When isobutylsilicate was applied to the reaction, the target product 5c could also be isolated in 58% yield. When phenethylsilicate and chloropropylsilicate were applied to the reaction. The target products 5d and 5e could be obtained in 88% and 93% yields, respectively. The secondary amine substituted alkylsilicate substrate was well tolerated and the target product 5f could be obtained in 58% yield. Subsequently, we examined the substrate scope of arylselenosulfonates (Scheme 3). It was found that when Se-phenyl arylsulfonates bore electron donating groups (4-OMe, 3-Me), the target products 5i and 5j could be obtained in 94% and 83% yields, respectively. The reaction of 1a with Se-(naphthalen-1-yl) benzenesulfonoselenoate 4c led to the desired product 5k in 94% yield. When Se-(2-chlorophenyl) benzenesulfonoselenoate 4c and Se-(pyridin-3-yl) benzenesulfonoselenoate 4d were applied to the reaction, the target products 5l and 5m could also be obtained in 67% and 58% yields, respectively. This study pro-



Scheme 3 Substrate scope for aryl-alkyl selenides compounds. Reaction conditions: substituted silicate 1 (0.34 mmol), substituted selenosulfonates 4 (0.2 mmol), II (2.5 mol%), (dry)DMSO (1 mL), 80 °C, Blue LED, N₂, 24 h. Isolated yields.

vides a new method for the construction of heteroaryl alkyl selenide compounds.

In order to investigate the applicability of this free radical cross-coupling strategy, we tried the reaction of alkylsilicates with thiosulfonates under similar reaction conditions (Scheme 4). It was found that the reactivity of thiosulfonates is lower than that of selenosulfonates. There are limited four substituted alkylsilicates, which were successfully applied to the reaction to afford the desired products. The target products 7a-d were obtained in 41%-78% yields, respectively. Next, we examined the reactivity of different substituted thiosulfonates.



Scheme 4 Substrate scope for aryl-alkyl sulfides compounds. Reaction conditions: substituted silicate 1 (0.34 mmol), substituted thiosulfonates 6 (0.2 mmol), II (2.5 mol%), (dry)DMSO (1 mL), 80 °C, blue LED, N2, 24 h. Isolated yields.

It is worth pointing out that the reaction showed a broad functional group compatibility, in which both electron-donating group Me and electron-donating withdrawing groups such as Cl and NO₂, were well tolerated. The target products 7e-7f could be observed in excellent yields. To our delight, when S-(4-hydroxyphenyl) benzenesulfonothioate and S-(4aminophenyl) benzenesulfonothioate were applied to the reaction, the target products 7h and 7i were obtained in moderate yields, respectively. The reaction of S-(benzo[d]thiazol-2yl) benzenesulfonothioate with cyclohexylsilicate led to 7j in 86% yield. Unfortunately, alkyl thiosulfonates failed to afford the desired products under the identical reaction conditions.

In order to improve the atom economy of the reaction and explore the further application of this reaction, we successfully recovered catechol with good yield (Scheme 5a). The successful recovery and reuse of catechol shows that this protocol has high atom economy. In order to explore the further application of this reaction, we tried the scale-up reaction of products 3q, 5h and 7c. To our delight, the target products 3q, 5h and 7c could be obtained in excellent yields, respectively (Scheme 5b). In addition, according to the relevant literature, 16 the treatment of compound 7c with different oxidizing agents successfully led to sulfoxide 8a and sulfone 8b in 88% and 78% yields, respectively (Scheme 5c). The above experiments show that this strategy has huge advantages in the potential industrial applications. What is exciting is that the reaction of cyclohexylsilicate with methyl (2S)-2-((tertbutoxycarbonyl)amino)-3-(phenoxyselenonothioyl)propanoate 2i under the optimized conditions successfully constructed the target product methyl (S)-2-((tert-butoxycarbonyl) amino)-3-(cyclohexylselanyl)propanoate 3r in 55% yields

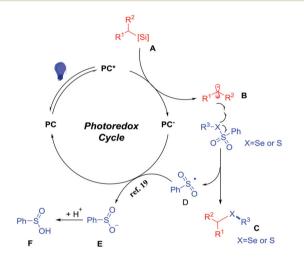
> 3q, 97% 7c, 77% PhSeSO₂Ph 5h, 1mmol 91% PhSSO₂Ph 7c, 3mmol 85% TMSCI (1.0 equiv) m-CPBA (4.0 equiv) H₂O₂

Scheme 5 Related derivative reactions. (a) Recovery and reuse of catechol. (b) The scale-up reaction of products 3q, 5h and 7c. (c) Under suitable oxidation conditions, the sulfide could be selectively oxidized to sulfoxide or sulfone. (d) Successfully constructed the target product methyl (S)-2-((tert-butoxycarbonyl)amino)-3-(cyclohexylselanyl) propanoate.

(Scheme 5d). This shows that the strategy has broad application prospects in the biological field.

Further mechanistic studies were conducted to gain deeper insight into the reaction pathway. Adding stoichiometric amounts of radical scavenger TEMPO or BHT and 1,1diphenylethylene to the reaction significantly suppressed the conversion, which revealed that a radical mechanism might be involved in this transformation. The reaction of cyclohexylsilicate with diselenide compound 9 instead of selenosulfonategave 3q in only <10% yield. This result indicates that diselenide compound 9 was not the main reaction intermediate of this reaction (Scheme 6).

Based on the above experimental results and previous reports, ^{17,18} we proposed a plausible mechanism in Scheme 7. The photocatalyst assists the single electron transfer of the alkylsilicate to generate an alkyl radical B. Subsequently, alkyl radical reacts B with selenosulfonates or sulfonothioates to afford the desired products and a sulfone radical intermediate. 19 Then the sulfone radical is reduced to a sulfone anionic intermediate E. Subsequently, protonation of intermediate E affords the benzosulfinic acid.20



Scheme 6 Mechanistic studies. (a) Radical scavengers are added under standard conditions, the target product could only be obtained in trace yields. (b) Treatment of cyclohexyl silicon and diselenide 9 under standard reaction conditions could not lead the target product.

Scheme 7 A plausible mechanism.

Conclusion

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In summary, we have successfully developed a light-mediated two-component free radical cross-coupling reaction between alkylsilicate compounds and selenosulfonates under metalfree catalysis conditions. At the same time, this strategy is also applicable to the efficient construction of aryl-alkyl selenides, heteroayl-alkyl selenides and aryl-alkyl sulfides. A new method for efficiently constructing C_{sp3}-Se and C_{sp3}-S bonds without any additives was developed. Successful construction of the unsymmetrical primary alkyl-alkyl selenide compounds, which are difficult to prepare, is the highlight of this work. It provides a green and simple method for the preparation of alkyl-alkyl selenide and aryl-alkyl selenide compounds. The recovery and reuse of catechol and the scale-up experiments show the broad prospects of this strategy in industrial applications. Successful construction of the selenocarbamate compound 3r shows that the strategy has broad application prospects in the biological field.

Experimental section

General procedure

Synthesis of 3d. In a glovebox, an oven-dried screw-capped 8 mL vial equipped with a magnetic stir bar was charged with diisopropylammonium bis(catecholato)propylsilicate (129.6 mg, 0.34 mmol) and Se-(2-phenoxyethyl) benzenesulfonoselenoate 2d (68.4 mg 0.2 mmol), cat (1.4 mg, 2.5 mol%), and DMSO (1.0 mL) were added via a syringe. The reaction mixture was stirred at r.t for 24 h under a 40 W LED lamp (40 W; $\lambda = 450-460$ nm; 5 cm away; made in TanLu. Ltd; borosilicate glass). After 24 h, the crude reaction mixture was diluted with ethyl acetate (20 mL) and washed with water (20 mL × 3). The organic layer was dried over Na₂SO₄, filtered, and concentrated. The residue was purified by flash chromatography to afford the pure product 3d.

Synthesis of 5a. In a glovebox, an oven-dried screw-capped 8 mL vial equipped with a magnetic stir bar was charged with diisopropylammonium bis(catecholato)propylsilicate (129.6 mg, 0.34 mmol) and Se-phenyl benzenesulfonoselenoate 4a (59.6 mg 0.2 mmol), cat (1.4 mg, 2.5 mol%), and DMSO (1.0 mL) were added via a syringe. The reaction mixture was stirred at 80 °C for 24 h under a 40 W LED lamp (40 W; λ = 450-460 nm; 5 cm away; made in TanLu. Ltd; borosilicate glass). After 24 h, the crude reaction mixture was diluted with ethyl acetate (20 mL) and washed with water (20 mL × 3). The organic layer was dried over Na2SO4, filtered, and concentrated. The residue was purified by flash chromatography to afford the pure product 5a.

Synthesis of 7c. In a glovebox, an oven-dried screw-capped 8 mL vial equipped with a magnetic stir bar was charged with bis(catechol)diisopropyl ammonium cyclohexyl silicate 1f (129.6 mg, 0.34 mmol) and S-phenyl benzenesulfonothioate 6a (50.0 mg 0.2 mmol), cat (1.4 mg, 2.5 mol%), and DMSO (1.0 mL) were added via a syringe. The reaction mixture was

stirred at 80 °C for 24 h under a 40 W LED lamp (40 W; λ = 450-460 nm; 5 cm away; made in TanLu. Ltd; borosilicate glass). After 24 h, the crude reaction mixture was diluted with ethyl acetate (20 mL) and washed with water (20 mL × 3). The organic layer was dried over Na2SO4, filtered, and concentrated. The residue was purified by flash chromatography to afford the pure product 7c.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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