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Selective deposition of metal oxide nanoflakes on graphene electrodes to obtain high-performance asymmetric micro-supercapacitors†

Zhenyuan Xia,^a Viktoriia Mishukova,^c Szymon Sollami Delekta,^c Jinhua Sun,^a Jaime S. Sanchez,^a Jiantong Li^c and Vincenzo Palermo^{a,b}

To meet the charging market demands of portable microelectronics, there has been a growing interest in high performance and low-cost microscale energy storage devices with excellent flexibility and cycling durability. Herein, interdigitated all-solid-state flexible asymmetric micro-supercapacitors (A-MSCs) were fabricated by a facile pulse current deposition (PCD) approach. Mesoporous Fe₂O₃ and MnO₂ nanoflakes were functionally coated by electrodeposition on inkjet-printed graphene patterns as negative and positive electrodes, respectively. Our PCD approach shows significantly improved adhesion of nanostructured metal oxide with crack-free and homogeneous features, as compared with other reported electrodeposition approaches. The as-fabricated Fe₂O₃/MnO₂ A-MSCs deliver a high volumetric capacitance of 110.6 F cm⁻³ at 5 μA cm⁻² with a broad operation potential range of 1.6 V in neutral LiCl/PVA solid electrolyte. Furthermore, our A-MSCs devices show a long cycle life with a high capacitance retention of 95.7% after 10 000 cycles at 100 μA cm⁻². Considering its low cost and potential scalability to industrial levels, our PCD technique could be an efficient approach for the fabrication of high-performance MSC devices in the future.

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Introduction

Nowadays, electrical energy storage technology has revolutionized the whole world and changed our daily life with numerous portable electronic devices such as smartphones, digital watches, wireless earbuds and other wearable devices.^{1,2} Due to the importance of energy storage devices, the 2019 Nobel Prize in Chemistry was awarded for the pioneering work on the rechargeable Li-ion battery. Meanwhile, supercapacitors, as another leading electrochemical energy storage technology, have also attracted intensive interest in both academia and industry in the last decade.³ In particular, planar micro-supercapacitors (MSCs) possessing a thin-layer structure, high power density, rapid charging/discharging capabilities and long cycle life are considered to be promising microscale power sources for on-chip electronic systems.^{4–6} Carbon-based

MSCs are excellent candidates for microscale power sources compared with conventional stacked-geometry supercapacitors, as they possess integrated thin-film electrodes with short ionic diffusion distances and separator-free structures. Besides this, the fabrication of MSCs with a wide working voltage window is a promising approach for achieving both high energy density and power density.⁷

Asymmetric supercapacitors (ASCs), consisting of two different electroactive materials as positive and negative electrodes, could make full use of two different potential windows in the same cell and significantly increase the overall energy density of devices. To date, transition metal oxides have been intensively studied in ASCs since they could offer high specific capacitance by taking advantage of ultrafast surface redox reactions of transition metal cations.^{7,8} According to the working potential windows of various transition metal oxides in aqueous electrolytes, most of the oxides are used as positive electrode materials with a stable potential range of around 0–1 V (vs. SCE), such as pseudocapacitive oxides MnO₂, ZnO and battery-type oxides Co₃O₄, V₂O₅, SnO₂, NiO, etc.^{9–15} On the other hand, there are few studies reporting negative electrode materials with a working potential range below 0 V (vs. SCE), which includes Fe₂O₃, MoO₃, and metal nitrides (VN), etc.^{16–19} Among these oxides, MnO₂ and Fe₂O₃ as positive and negative electrode materials, respectively, are the most promising candi-

^aDepartment of Industrial and Materials Science, Chalmers University of Technology, Hörsalsvägen 7B, 41258 Göteborg, Sweden. E-mail: zhenyuan@chalmers.se

^bIstituto per la Sintesi Organica e la Fotoreattività, CNR, via Gobetti 101, 40129 Bologna, Italy

^cSchool of Electrical Engineering and Computer Science, KTH Royal Institute of Technology, Electrum 229, 16440 Kista, Sweden. E-mail: jiantong@kth.se

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dates for ASC devices because of their high theoretical capacitance, low cost, abundance and environmental friendliness. Moreover, the well-separated potential windows of MnO_2 (−0.1 to 0.8 V vs. SCE) and Fe_2O_3 (−0.9 to 0 V vs. SCE) could significantly increase the operating voltage of ASCs in aqueous electrolytes.

In order to assemble planar ASCs for on-chip integration, various fabrication strategies, such as photolithography, laser scribing, screen/inkjet printing, and electrodeposition have been developed to pattern carbon and other electro-active materials into interdigital electrodes to form asymmetric microsupercapacitors (A-MSCs).^{20–27} Among these approaches, printing is a cost-effective, cheap, fast and scalable approach for facile fabrication of micropattern electrodes. Conductive materials (e.g., active carbon, carbon nanotubes, nanofibers, and graphene) can also be printed to give interdigitated electrodes.²³ In comparison with screen printing, inkjet printing is a flexible and non-contact method that allows precise control of high-resolution patterning without the need for prepatterned masks. However, direct inkjet printing has some limitations: unlike screen printing, which could easily pattern carbon materials and pseudo-capacitive metal oxide mixtures with sufficient thickness, inkjet printing is not suitable for thick patterns. More importantly, micro-printing asymmetric electrodes composed of two different materials would require two printing nozzles and a precise alignment of the two printing patterns, to avoid shortcuts or defects.^{28,29}

Unlike the inkjet printing technique, electrodeposition has its unique advantages in thin film fabrication of transition metal oxide materials. Electrodeposition is a mature and scalable technique for nanocrystalline metal plating or metal oxide functional coating in industry. It allows electrosynthesis of metal oxides from solution of metal salts and the subsequent deposition onto any complex-shape substrates through a one-step process.^{21,30–32} In addition, by tuning the electrochemical conditions (e.g. pulse or direct current output, deposition time, electrolytes), we can spatially control the loading amount and surface morphology of metal oxide on selected electrodes.^{33–40} Although electrodeposition has been intensively employed in the fabrication of nanostructured metal oxide films for bulk supercapacitors,^{41,42} there are still some challenges in this thin film formation process. For example, conventional anodic or cathodic direct current deposition of metal oxides will result in internal stresses on the substrates with the formation of delamination layers after coating, which is unfavorable for stable MSC electrodes. Besides, metal oxide materials usually suffer from poor electronic and ionic conductivity, affecting the rate capability and cycling stability of the ASC devices.

These challenges can be overcome by the introduction of new electrochemical deposition techniques and the combination of advanced current collectors with metal oxides. Herein, we used pulse current deposition (PCD) as an alternative cost-effective and simple technique^{43,44} for the fabrication of A-MSCs based on pseudo-active metal oxide (MO_x) nanomaterials and graphene-based current collectors. Generally,

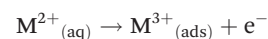
electrochemically exfoliated graphene (EG) was inkjet printed on a Kapton film²³ to form interdigitated patterns, which served as reliable conductive deposition substrates, thanks to the high electrical conductivity, excellent electrochemical stability, large specific surface area and low cost of graphene. In our exfoliation process, the total cost including those of starting material graphite foil (ca. 0.5 € per g), electrolyte 0.1 M $(\text{NH}_4)_2\text{SO}_4$ (ca. 1 € per 100 mL), and power consumption (ca. 1 € per kW h) is less than 5 € per g on the lab scale. Then, MnO_2 and Fe_2O_3 were selected as positive and negative electrode materials, respectively, with well-separated potential windows and matchable capacitances. Crack-free and mesoporous MnO_2 and Fe_2O_3 thin films were fabricated by selective PCD coating on the EG layer and subsequent annealing treatment. To the best of our knowledge, there is no previous research using pulse electrodeposition in micro-scale supercapacitor applications. Our results indicated that the PCD method could significantly improve the coating uniformity, boosting the cycling stability of our prepared A-MSC devices, as compared with the conventional direct current deposition approach. The current technique described here is promising for the rapid fabrication of high-performance A-MSCs with high versatility of pseudo-capacitive metal oxide materials.

Results and discussion

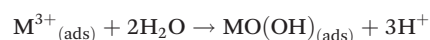
Fig. 1 shows the fabrication process of $\text{Fe}_2\text{O}_3/\text{MnO}_2$ -based A-MSCs. Generally, it was divided into the following steps. First, conductive EG inks were printed on a Kapton substrate to obtain planar interdigitated micropatterns. After annealing treatment of the EG layer, either MnO_x or FeOOH nanoflakes were electrodeposited on the EG pattern with a two-electrode configuration (EG as the working electrode and a Pt wire as the counter electrode). Then, the obtained samples were annealed at 300 °C to convert the deposited precursors into either Fe_2O_3 or MnO_2 . For the single metal oxide based MSCs, 1 M LiCl was used as the aqueous electrolyte for either Fe_2O_3 or MnO_2 based electrodes. For A-MSCs, solid electrolyte of LiCl/PVA gels was used to fabricate the all-solid-state $\text{Fe}_2\text{O}_3/\text{MnO}_2$ A-MSCs (named Fe–Mn).

The electrosynthesis and deposition process of MO_x by the anodic approach can be described by the following equations:⁴⁵

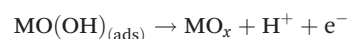
Step 1:



Step 2:



Step 3:



The proposed mechanism for anodic deposition of manganese oxide involves three steps: (1) oxidation of cationic Mn^{2+} precursors in the solution near the anode electrodes; (2) and the subsequent hydrolysis of Mn^{3+} to form the $\text{MnO}(\text{OH})$ inter-



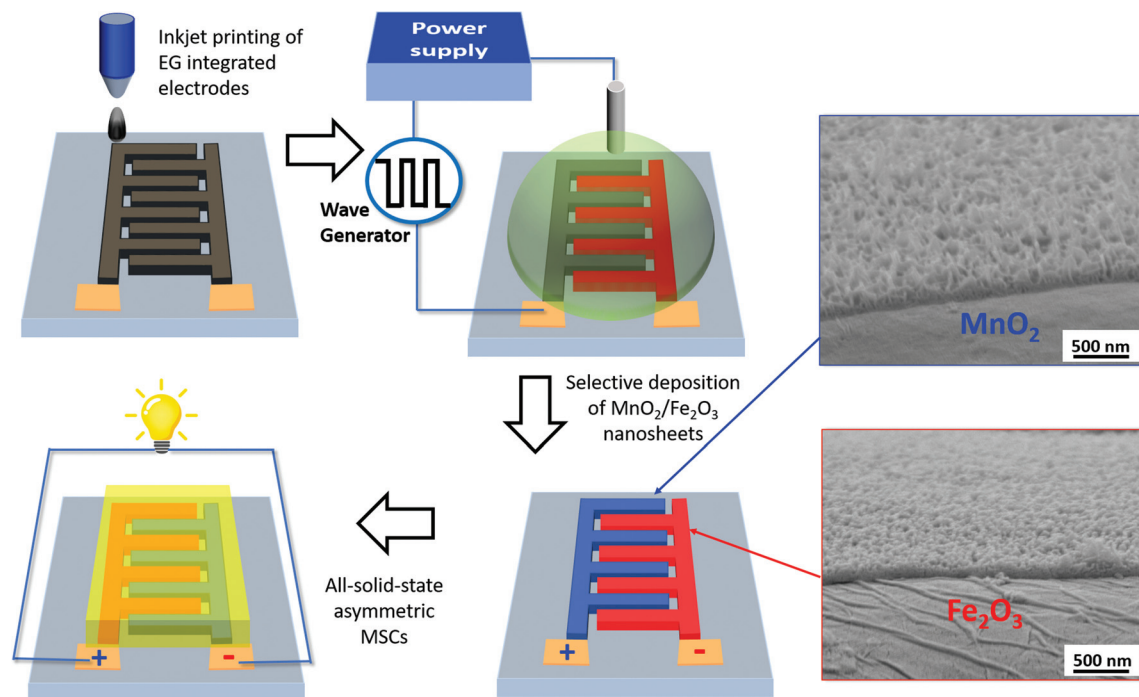


Fig. 1 Schematic illustration of the electrochemical deposition process for Fe_2O_3 and MnO_2 based inkjet-printed graphene A-MSCs (inset, cross-section SEM images of Fe_2O_3 and MnO_2 nanoflakes).

mediate which precipitates on the anode electrode; (3) in the last step, the adsorbed manganese hydroxide is further oxidized to MnO_2 with the release of a proton. Manganese acetate (MnAc) was selected as the precursor material for electrodeposition of MnO_2 , because acetate counterions could reduce the potential at which film deposition occurs and results in a high deposition rate.^{46,47} As for iron oxide, the deposition mechanism is similar to that of manganese oxide, starting from the oxidation of Fe^{2+} cations to Fe^{3+} and precipitation of the obtained iron hydroxide on the anode surface, but without the last oxidation step.¹⁷ Ammonium iron(II) sulfate was used as the common source material for Fe_2O_3 anodic deposition. In both cases, the mesopores might be formed from the scaffold-like void spaces that originate between the aggregating metal hydroxide nanoparticles.

Unlike the conventional direct current (galvanostatic) deposition (DCD), in this work, we used pulse current deposition for the efficient coating of MO_x functional layers. Pulse deposition has been widely used in the industrial plating of nanocrystalline metals to improve deposits for several decades, while it is scarcely studied in the field of nanostructured metal oxides. Thus, the two electrodeposition techniques for optimizing the coating qualities in energy storage devices are worth comparing. The deposition conditions for DCD and PCD methods are similar, in both cases, the applied current density for the target electrodes was a fixed value of $47 \mu\text{A cm}^{-2}$ ($5 \mu\text{A}$ per 0.105 cm^{-2} for the single electrode). The deposition time for MnO_2 in the DCD mode was 5 min, 10 min, 20 min, and 40 min; whereas the deposition time for Fe_2O_3 in

the DCD mode was 10 min, 20 min, 40 min, and 80 min, separately. For pulse current deposition (PCD), a square-wave pulse signal was generated with a period of 5 ms and a duty ratio of 0.5 (5 ms 'on time' and 5 ms 'off time' for one duty cycle, as shown in Fig. S1†). Accordingly, the efficient deposition time for both MnO_2 and Fe_2O_3 in the PCD mode was the same as those of the samples obtained in the DCD mode, while the total deposition time was double due to the extra 'off time'.

Fig. 2 and S2, 3† show SEM images of MnO_2 and Fe_2O_3 prepared by the DCD or PCD method at different coating times with post-treatment annealing. As shown in Fig. 2a and S2c,† MnO_2 deposited by the galvanostatic approach exhibited a considerable number of cracks after coating for 20 minutes. Spalling and warping of the coated layers were clearly observed on a scale of hundreds of micrometers. The density of cracks and defects was even higher for a longer deposition time ($t = 40 \text{ min}$ Fig. S2d†), and the cracked films tended to peel off from the substrate, probably due to shrinkage stress of the thick layer.⁴⁰ Similar issues of cracking and delamination were observed on the Fe_2O_3 film deposited by the galvanostatic method after 40 minutes of coating time (Fig. 2d and S3a–d†), with smaller warping cracks of around 2–5 micrometer long. In contrast, both MnO_2 and Fe_2O_3 films prepared by PCD, using the same current density, revealed a uniform coating layer well adhering to the EG sublayer even after 40 minutes of deposition for MnO_2 and 80 minutes of deposition for Fe_2O_3 (Fig. 2e, f and S3e–h†). The better film quality obtained by PCD with respect to the DCD current could be attributed to more efficient transport of ions achievable in the former tech-



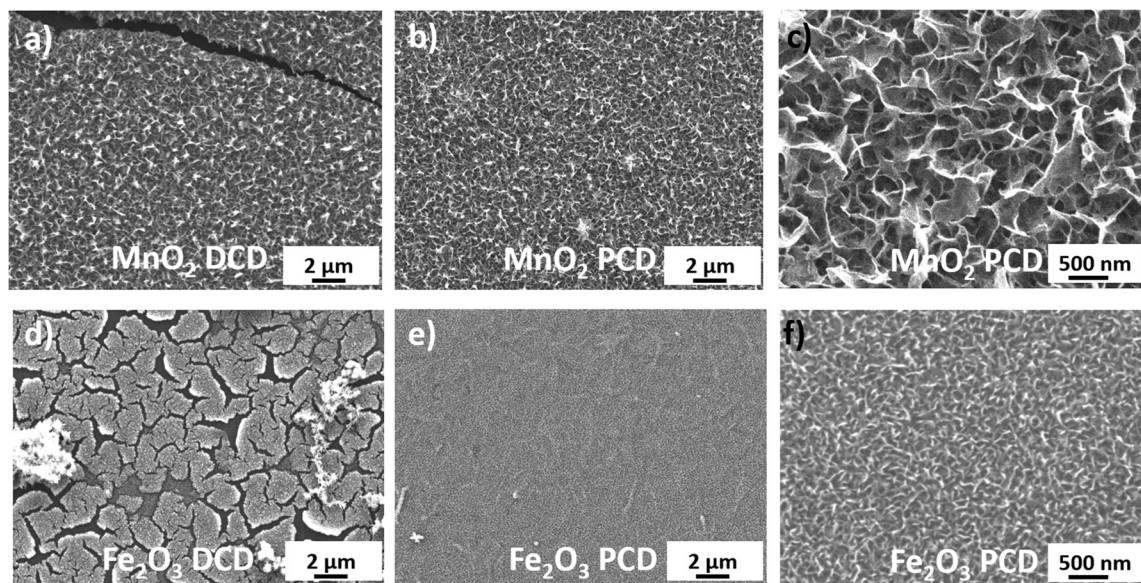


Fig. 2 SEM images of the MnO_2 thin film coated on the EG pattern by (a) direct current deposition (DCD mode, $t = 20$ min) and (b and c) pulse current deposition methods (PCD mode, $t = 40$ min); SEM images of the Fe_2O_3 thin film coated on the EG pattern by (d) direct current deposition (DCD mode, $t = 40$ min) and (e and f) pulse current deposition methods (PCD mode, $t = 80$ min).

nique: the “on” time in PCD can generate a high instantaneous positive potential at the anode, and the huge overvoltage changes with ultra-short current pulses provide high localization of the Faradaic process with a small interelectrode gap on the anode side;^{48,49} meanwhile, the “off” time during the periodic pulse deposition gives manganese or iron ions sufficient replenishment time to diffuse from the bulk solution to the EG surface.^{9,50}

Fig. 1 and 2c, f show the typical surface morphology of MnO_2 and Fe_2O_3 after 10 and 40 minutes of PCD, respectively. The two kinds of MO_x nanoflakes were well distributed on the microelectrode surface with mesoporous structures. The average pore size was on the order of 100 nm from MnO_2 nanoflakes (Fig. 2c) with the film thickness on the order of 200 nm, as estimated from SEM cross-section images (Fig. S4b†). The MnO_2 layer thickness was tunable according to the deposition time: it increased from 150 nm to 300 nm by increasing deposition time from 5 to 20 minutes (Fig. S4a–c†). In the case of Fe_2O_3 , a more compact porous structure with closely packed small nanoflakes (average pore size of 50 nm) and average film thickness of around 150 nm was observed ($t = 40$ min, Fig. 2f and S4e†). The layer thickness of Fe_2O_3 increased from 90 nm to 230 nm by increasing deposition time from 20 to 80 minutes (Fig. S4d–f†). These mesoporous structures from pulse current deposited MnO_2 and Fe_2O_3 created high accessible surface area, which can provide effective ion transport pathways and consequently enhance the overall electrochemical performance of the final device. The XRD patterns of our MO_x samples (Fig. S5†) consist two broad peaks at 22.4° and 26.7° , which primarily come from the (110) and (210) phase⁵¹ of the polyimide substrate and the (002) phase of the graphene layer.⁵² Low intense signals at around 28.7° for

MnO_2 and 35.6° for Fe_2O_3 also reveal the presence of very small oxide crystallites, possibly corresponding to pyrolusite $\beta\text{-MnO}_2$ (110) and maghemite $\gamma\text{-Fe}_2\text{O}_3$ (311).^{53,54}

We used Raman spectroscopy to compare bare EG, $\text{MnO}_2\text{-EG}$ and $\text{Fe}_2\text{O}_3\text{-EG}$ (Fig. 3a). The Raman spectra showed two distinct peaks at 1352 cm^{-1} and 1580 cm^{-1} for all three samples, attributed to the D peak (intervalley double resonance, usually from defect-related scattering) and the G peak (in-plane E_{2g} optical phonon mode) from the defective EG layer.^{55,56} Another weak peak at $\sim 2700\text{ cm}^{-1}$ was due to the 2D peak of graphene (the second order of the D peak). The Raman spectrum at low wavenumbers revealed the scattering peaks of either MnO_2 or Fe_2O_3 nanoflakes. The dominant peaks of Fe_2O_3 were at around 350 (T_{2g} mode), 511 (E_g mode), and 700 cm^{-1} (A_{1g} modes), ascribed to the maghemite phase of $\gamma\text{-Fe}_2\text{O}_3$.⁵⁷ The Raman spectrum of the MnO_2 film showed one main peak at 643 cm^{-1} , which could be identified as Mn–O stretching vibrations from the tetragonal $\beta\text{-MnO}_2$ structure.⁵⁸

The chemical composition of bare EG, $\text{MnO}_2\text{-EG}$ and $\text{Fe}_2\text{O}_3\text{-EG}$ layers was confirmed by XPS analysis, showing C 1s, Mn 2p, Fe 2p and O 1s core levels (see in Fig. 3b–d and S6†). The C 1s spectra were measured for the bare EG sample; a sharp peak was observed at 284.6 eV (Fig. 3b) due to the $\text{sp}^2\text{ C}=\text{C}$ bond and $\text{sp}^3\text{ C}-\text{C}$ bond, and other peaks at 286.2 , 288.2 , and 288.9 eV were assigned to the C–OH/C–O–C, C=O and O–C=O components, respectively. The existence of these oxygen functional groups was due to the printed EG ink, which can be attributed to the unavoidable oxidation of graphite during the anodic exfoliation process of EG sheets. Fig. 3c shows the Mn 2p spectrum from MnO_2 , which consists of two main peaks at 641.6 and 644.3 eV for Mn $2p_{3/2}$, and one peak at 653.2 eV for Mn $2p_{1/2}$. The energy separation between the Mn $2p_{3/2}$ and Mn



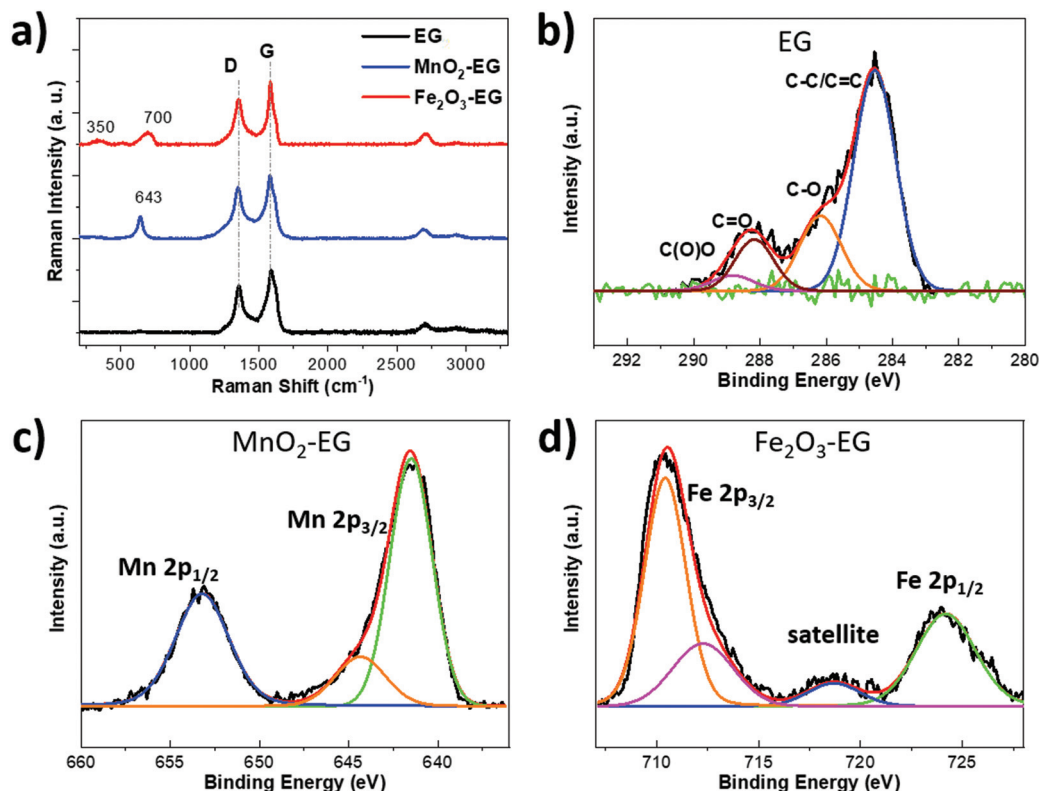


Fig. 3 (a) Raman spectrum of EG, MnO₂-EG and Fe₂O₃-EG composites, and XPS spectrum of (b) C 1s, (c) Mn 2p, and (d) Fe 2p observed in bare EG, MnO₂-EG, and Fe₂O₃-EG composites.

2p_{1/2} is 11.6 eV, in agreement with previous reports on MnO₂ materials.⁵⁹ The O 1s peak was deconvoluted in two main components: Mn–O–Mn component at 529.8 eV and Mn–OH component at 531.1 eV (Fig. S6a†). The –OH bond might come from the adsorption of O₂ and water molecules on the MnO₂ surface. Meanwhile, Fig. 3d shows the chemical state of Fe 2p, which was used to detect the presence of iron oxide. There were one main peak and a shoulder peak at 710.4 and 718.8 eV for Fe 2p_{3/2}, and one peak at 724.2 eV for Fe 2p_{1/2}. The Fe 2p_{3/2} was also associated with a shake-up satellite around 7 eV higher than the main peak, which is close to the experimental iron(III) oxide data reported in the literature.^{40,53} The O 1s peak (see in Fig. S6b†) was fitted using two main components: Fe–O–Fe component at 529.7 eV and Fe–OH component at 531.3 eV, similar to the MnO₂ O 1s result.

The detailed morphology and composition of Fe₂O₃ and MnO₂ were further examined by TEM and EDS analysis at the nanoscale level. Fig. 4a and S7a† show the representative TEM images of crumpled MnO₂ nanoflakes. High resolution TEM (HRTEM) image (Fig. 4a) revealed the relative lattice fringes with *d* spacings of 0.31 nm and 0.16 nm, which were corresponding to the (110) and (211) lattice plane of pyrolusite MnO₂. In most cases, the orientation of MnO₂ nanoflakes was random with no preferred direction, as evidenced by the broad diffraction rings from the fast Fourier transform (FFT) pattern (Fig. S7c†).¹⁰ Furthermore, the elemental composition of Mn, O and Cu was confirmed by energy dispersive X-ray spec-

troscopy analysis (EDS) of the TEM copper grid of MnO₂ flakes (Fig. S8a†). In the case of Fe₂O₃, a needle-like Fe₂O₃ nanoflake agglomerate was observed in TEM image (Fig. S7b†) with an average thickness of 5–10 nm and a length of around 100 nm, which is consistent with the SEM results. HRTEM images in Fig. 4b and S7b† show Fe₂O₃ nanoflakes randomly distributed on the TEM grid, which also explains the partial amorphous nature of the Fe₂O₃ film by XRD. The crystalline lattice of Fe₂O₃ sheets showed *d* spacings of 0.30 nm, 0.25 nm and 0.17 nm corresponding to the (220) (311) and (422) lattice planes of the maghemite structure (Fig. S7d†), which is consistent with the Raman and XRD results. The presence of Fe and O elements were confirmed by EDS analysis (Fig. S8b†).

The electrochemical behaviour of Fe₂O₃ and MnO₂ on EG MSCs were studied in a two-electrode configuration with a 1 M LiCl aqueous electrolyte. We first used cyclic voltammetry (CV) to study what happens by coating only one of the two EG electrodes with MO_x. MnO₂ was coated by 10 minutes of PCD on the positive electrode, to obtain asymmetric MnO₂/EG electrodes. On the other hand, Fe₂O₃ was coated by 40 minutes of pulse deposition on the negative electrode to obtain asymmetric EG/Fe₂O₃ electrodes. To clarify, we refer to the different electrode couples as: positive electrode material/negative electrode material, in this order. These asymmetric samples were compared with pristine, symmetric EG electrodes.

There was no significant difference in the capacitance of MnO₂/EG EG/Fe₂O₃ as compared to symmetric EG/EG devices,



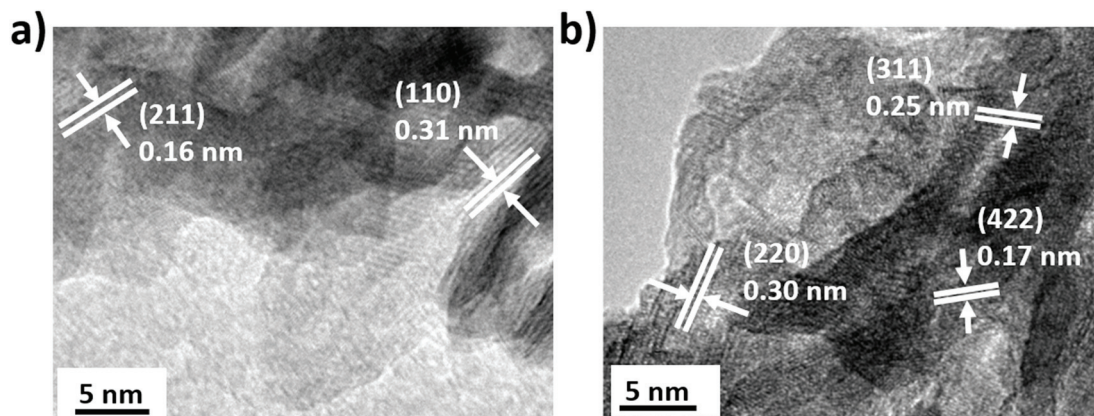


Fig. 4 High magnification TEM images of (a) MnO_2 and (b) Fe_2O_3 nanoflakes.

as shown in Fig. S9†. The areal capacitance values of $\text{EG}/\text{Fe}_2\text{O}_3$ and MnO_2/EG asymmetric MSCs were 47 and $52 \mu\text{F cm}^{-2}$ at a scan rate of 100 mV s^{-1} , these values are *ca.* 3 times higher than the capacitance of initial EG MSCs ($\sim 15 \mu\text{F cm}^{-2}$). The small capacitance value of either bare EG or MO_x/EG MSCs was attributed to the low EDLC capacitance of the bare EG electrode, and unbalanced charges between the EG and MO_x electrodes results in poor capacitor performance of the whole asymmetric device.

Then, we tried to study symmetric MSCs, where both electrodes were functionalized with the same material, affording either $\text{Fe}_2\text{O}_3/\text{Fe}_2\text{O}_3$ or $\text{MnO}_2/\text{MnO}_2$ electrode couples. Fe_2O_3 based symmetric MSCs were named Fe-1, Fe-2, and Fe-3, according to their pulse deposition time of 20, 40 and 80 minutes, respectively, whereas MnO_2 based symmetric MSCs were named Mn-1, Mn-2, and Mn-3, according to their pulse deposition time of 5, 10 and 20 minutes (see in Fig. 5 and S10†), respectively. We used different deposition times for Fe and Mn in order to achieve similar areal capacitance for devices for the two types of materials, as detailed below.

Fig. 5a, b and S11b† show the CV curves of Fe-2 and Mn-2 devices. Both devices exhibited quasi-rectangular CV shapes at different scan rates. A slight distortion of the rectangular shape could evidence the pseudocapacitive behaviour of MO_x . The pseudocapacitive behaviour of Fe_2O_3 could be explained by a fast and reversible faradaic reaction of the redox $\text{Fe}^{3+}/\text{Fe}^{2+}$ couple during the charge and discharge process.^{40,60} Similarly, the pseudocapacitive behaviour of MnO_2 could be described by the redox transitions of interfacial oxidation species at various Mn oxidation states (*e.g.*, $\text{Mn}^{3+}/\text{Mn}^{4+}$).⁶¹ The areal capacitances of Fe-2 and Mn-2 were 253 and $244 \mu\text{F cm}^{-2}$ at a scan rate of 100 mV s^{-1} , respectively, which is almost 15 times higher than that of the bare EG MSC device. Furthermore, the areal capacitances of MO_x based symmetric MSCs, deposited at various times, are plotted in Fig. 5c and d. In all cases, the areal capacitance increased with the PCD coating time. All symmetric MO_x -MSCs showed much higher areal capacitances as compared to that of the bare EG one. By using different deposition times, we could achieve matchable capacity values

for Fe-1/Mn-1, Fe-2/Mn-2, and Fe-3/Mn-3 devices (see in Fig. S11†). The above results help us to rationally design asymmetric MSC devices with charge balancing between the positive MnO_2 electrode and the negative Fe_2O_3 electrode.

After matching the loading amount of the positive and negative electrode, we produced asymmetric MSC device coating MnO_2 for 10 min on the positive electrode and Fe_2O_3 for 40 min on the negative electrode $\text{MnO}_2/\text{Fe}_2\text{O}_3$. This time we used the LiCl/PVA gel electrolyte instead of the liquid electrolyte, simply because the PVA based polymer electrode features a high water absorbing and holding capacity, easy handling without the risk of leakage of the liquid electrolyte, which is suitable for flexible A-MSC devices and on-chip integration.

The asymmetric $\text{MnO}_2/\text{Fe}_2\text{O}_3$ sample (Fe-Mn) was tested in an operating potential window range from 0 to 1.6 V, according to previous studies on symmetric metal oxide devices.⁶⁰ Fig. 6a shows the CV curves of Fe-Mn-A-MSC obtained at different scan rates, from 5 mV s^{-1} to 100 mV s^{-1} . A nearly rectangular CV shape at different scan rates indicates fast and reversible redox reactions of both electrodes, MnO_2 and Fe_2O_3 . The maximum areal capacitance achieved from Fe-Mn ($1534 \mu\text{F cm}^{-2}$ at 5 mV s^{-1}) is much higher than that from the bare EG device ($39 \mu\text{F cm}^{-2}$ at 5 mV s^{-1}) in the PVA/LiCl gel electrolyte by a factor of 40 (see in Fig. S12†). Our result obtained for a PCD device is also superior to that for DCD based Fe-Mn A-MSC prepared with the same deposition time (the maximum areal capacitance of DCD based device is $348 \mu\text{F cm}^{-2}$ at 5 mV s^{-1} , Fig. S13†), further confirming the advantage of pulse current deposition in high performance MSC devices. Moreover, the galvanostatic charge-discharge (GCD) measurements under different current densities from 5 to $50 \mu\text{A cm}^{-2}$ (Fig. 6b) shows a small deviation from the triangular shape and relatively symmetric shape of GCD curves, indicating good electrochemical performance of our metal oxide electrodes, in agreement with previous CV results. Fig. 6c presents the rate performance of the Fe-Mn device at different current densities. A high stack capacitance of 110.6 F cm^{-3} was achieved at a current density of $5 \mu\text{A cm}^{-2}$, which significantly exceeds that of the previously reported graphene-based devices (*e.g.*



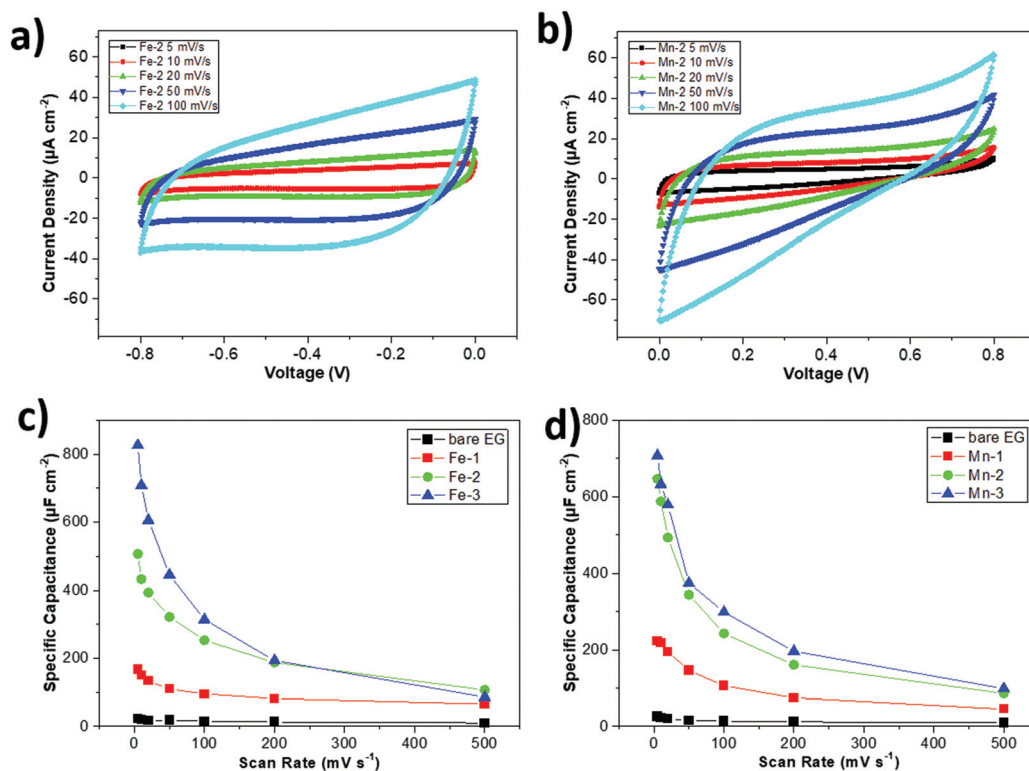


Fig. 5 Electrochemical performance of symmetric MSCs. CV curves of (a) Fe-MSC-2 and (b) Mn-MSC-2 symmetric devices at different scan rates; specific areal capacitances of (c) Fe-MSCs and (d) Mn-MSCs with different loading amounts at different scan rates.

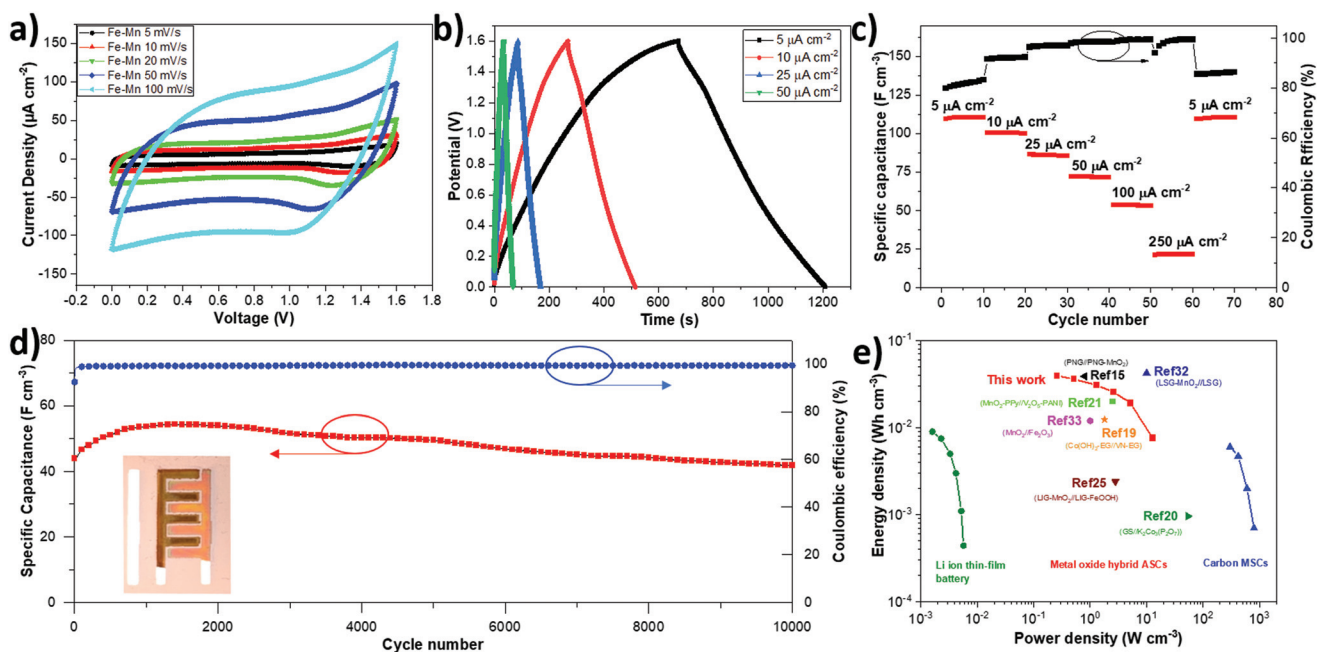


Fig. 6 Electrochemical performance of asymmetric Fe-Mn A-MSC prepared by PCD in LiCl/PVA electrodes. (a) CV curves of Fe-Mn at different scan rates; (b) charge-discharge curves of Fe-Mn at different current densities; (c) rate capability of Fe-Mn at different current densities; (d) long-term cycling stability of Fe-Mn at a current density of $100 \mu\text{A cm}^{-2}$; (e) Ragone plot of specific volumetric energy density and power density for Fe-Mn compared with other microscale energy-storage devices.



1.3 F cm⁻³ for onion-like graphene based MSCs and 3.1 F cm⁻³ for laser-written graphene oxide MSCs),^{62,63} and also superior to others state-of-the-art transition metal oxide based A-MSCs, *e.g.*, MnO₂/Fe₂O₃ A-MSCs on Cr/Ni patterns (60 F cm⁻³), PNG-MnO₂/PNG on a plasma reduced and nitrogen-doped graphene oxide substrate (38.1 F cm⁻³), and NPG-MnO₂/NPG-Ppy on a nano-porous gold substrate (45.3 F cm⁻³).^{15,32,38} To demonstrate the device stability of our asymmetric supercapacitor, we also performed long-term cycling of Fe–Mn A-MSC by GCD at 100 μA cm⁻². Interestingly, our device retained about 95.7% of the initial value with a high coulombic efficiency (99.3%) after 10 000 cycles, as shown in Fig. 6d. Furthermore, the bending test was performed under different angles with CV scan at 100 mV s⁻¹ (see in Fig. S14†). The capacitance of the Fe–Mn device maintained almost the same value after bending at *ca.* 90° and 180°, showing the potential applicability of this A-MSC as a flexible supercapacitor device.

Finally, the volumetric energy and power density of Fe–Mn were compared with the state-of-the-art performance metrics obtained from the literature (Fig. 6e). The Fe–Mn A-MSC device delivers a maximum energy density of 39 mW h cm⁻³ (corresponding power density of 259 mW cm⁻³) and a maximum power density 12.9 W cm⁻³ (corresponding energy density of 7.6 mW h cm⁻³). The energy density value is much higher than those of typical carbon-based micro-supercapacitors, and comparable with those of a Li thin-film battery (≤10 mW h cm⁻³) and other reported pseudocapacitive hybrid A-MSCs (Table S1†) such as LIG-MnO₂/LIG-FeOOH (2.4 mW h cm⁻³), MnO₂/Fe₂O₃ (12 mW h cm⁻³), Co(OH)₂-EG/VN-EG (12.4 mW h cm⁻³), and PNG/PNG-MnO₂ (38.1 mW h cm⁻³) micro-supercapacitors.^{15,19,30,38,64}

Conclusions

In summary, pulse current deposition has been demonstrated as a facile and highly efficient electrochemical approach for the fabrication of high-performance planar asymmetric MSCs. Crack-free MnO₂ and Fe₂O₃ mesoporous nanoflakes were uniformly electrodeposited on inkjet-printed graphene micropatterns as positive and negative A-MSC electrodes respectively. The fabricated A-MSCs exhibit simultaneously increased volumetric capacitance and working voltage window, thus leading to significantly higher energy density than other MSCs previously reported in the literature. The low-cost of electrodeposition techniques facilitate upscaling MO_x coating from the laboratory level to large scale industrial production. The versatility of such an electrodeposition approach paves the way towards the fabrication of other metal oxide structures, in search for synergies between metal cations and enhanced electrochemical properties.

Experimental section

Electrochemical exfoliation of graphene

The EG was produced according to the procedure described in our previous publication.²³ Briefly, two pieces of graphite foil

(Alfa Aesar, 0.13 mm thick, product number: 43078) were inserted into 0.1 M (NH₄)₂SO₄ aqueous electrolyte solution as the cathode and anode (at a distance of 2 cm), respectively, and biased at a DC voltage of 10 V for exfoliation. The obtained graphite powder was collected and the rinsed powder was dispersed in DMF at a concentration of about 2 mg mL⁻¹ by ultrasonication for 10 min.

Graphene ink formulation

The centrifugation-assisted solvent exchange technique was used to formulate EG ink. First, ~40 mL EG/DMF dispersion was centrifuged at 3000 rpm for 15 min to remove the largest particles. Then the supernatant was harvested and centrifuged again at 10 500 rpm for 15 min to separate EG from DMF. After DMF was removed, 16 mL of the mixed solvent of cyclohexanone and terpineol (with a volume ratio of 3 : 1) and 320 mg of ethyl cellulose (viscosity 4 cP for 5 w/v% in 80 : 20 toluene : ethanol, Sigma-Aldrich, product number: 200646) were added to disperse the sedimented EG and sonicated for 30 min. Finally, the EEG ink was centrifuged at 2000 rpm for 3 min to further remove big particles. The final ink concentration was about 2.3 mg mL⁻¹.

Fabrication of planar EG-MSCs

EG ink was printed using a commercial piezoelectric Dimatix Materials Printer (DMP 2800, Dimatix-Fujifilm Inc.) equipped with a 10 pL cartridge (DMC-11610). First, EG inks are printed at the drop spacing of 25 μm and substrate temperature of 45 °C on different substrates to form the electrodes for MSCs. Afterward, the samples were annealed on a hot plate in air for 1–2 h at temperatures of around 380 °C for the Kapton substrate.

Selective deposition of the MnO₂ or Fe₂O₃ thin film on EG electrodes

A polydimethylsiloxane (PDMS) mask was fixed on the top of MSC along with an electrochemical setup composed of a two-electrode cell: a Pt wire as the counter electrode and the target substrate as the working electrode. For MnO₂ coating, few droplets of 10 mM Mn(COOCH₃)₂ and 40 mM NaCOOCH₃ solution in water were deposited inside the mask as the electrolyte for MnO_x deposition. For Fe₂O₃ coating, few droplets of 10 mM (NH₄)₂Fe(SO₄)₂ and 40 mM NaCOOCH₃ solution were used as the electrolyte. Galvanostatic direct current electrodeposition was performed under a positive bias to the target electrode, maintaining a current density of 47 μA cm⁻² to coat MnO_x or FeOOH films on EG-based MSC electrodes with time from 5 to 80 min. Pulse current deposition was conducted using a square-wave pulse current of 47 μA cm⁻² (5 μA current on the 0.105 cm⁻² single electrode area, as shown in Fig. S1b†) with a period of 5 ms and a duty ratio of 0.5 with the time from 5 to 80 min. A function generator (Tenma 72-14111) was used for the generation of the wave function signal, and Keithley 2450 was connected with the wave generator as the power source. After electrochemical deposition, the composites were sequentially washed with DI water and dried under a gentle flow of nitrogen gas. Then, the MSCs were annealed at 300 °C under ambient air conditions.



Electrochemical characterization

All the electrochemical properties of the fully printed MSCs were characterized using a two-electrode configuration Biologic SP-300 potentiostat (for CV profiles). The areal capacitance C_A was calculated either from the CV profiles or the GCD curves. For the CV profiles, $C_{A,CV} = \int_0^{\Delta V} (I_C - I_D) dV / (2A\nu\Delta V)$, where I_C and I_D are the charging and discharging currents, respectively, ΔV is the voltage window, A is the active area of the metal oxide film, and ν is the scan rate. The total device area was calculated according to the pattern size of active materials ($A = (2.3 \times 0.7 \times 4 + 6 \times 0.7) \times 2 = 21 \text{ mm}^2 = 0.21 \text{ cm}^2$, see in Fig. S1b†). For the GCD curves, $C_{A,GDC} = I\Delta V / A\Delta t$ where I is the discharging current and Δt is the discharging time. All the areal energy density is calculated as $E_A = 0.5C_{A,CV}\Delta V^2$, and the areal power density is $P_A = E_A/t = \nu E_A / \Delta V$ with $t = \Delta V / \nu$ being the discharging time. The volumetric energy (power) density is calculated by dividing E_A (PA) by the average MOx layer thickness (T) measured from our SEM cross-sections ($T = (151 + 197)/2 = 174 \text{ nm}$).

Materials characterization

The surface morphology of the MnO_2 or Fe_2O_3 thin film was monitored by field emission scanning electron microscopy (JEOL JSM-7800F Prime), under an accelerating voltage of 8 kV. Transmission electron microscopy (FEI Tecnai T20, 200 kV) was used to characterize the metal oxide crystalline structures. Raman spectroscopy was performed using a WITec alpha300 R, with a 532 nm laser line and a laser spot size of 300 nm. XRD measurements were performed on a Bruker D8 Advance diffractometer at 40 kV and 40 mA with Cu $K\alpha$ radiation ($\lambda = 0.1542 \text{ nm}$). The chemical configuration was measured using a VersaProbe III X-ray photoelectron spectroscope, in an ultra-high vacuum (UHV) apparatus (base pressure $1 \times 10^{-10} \text{ mbar}$) using a non-monochromatic Mg- $K\alpha$ excitation source.

Conflicts of interest

The authors declare no competing financial interest.

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