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Organic cation (DMPI) intercalated graphite anode for high voltage next generation dual-ion batteries†

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Developing advanced energy storage systems to address the intermittency of renewable energy sources is crucial for meeting ever-increasing energy demands. Among post lithium-ion battery systems, dual-ion batteries (DIBs) have shown bright prospects in developing low-cost and safe batteries with good electrochemical performance. Herein we have modelled, for the first time, organic cation intercalated graphite systems. Imidazolium based ionic liquid 2,3-dimethyl-1-propyl imidazolium chloride (DMPI-Cl) with the AlCl_3 salt has been implemented as an electrolyte. Using first principles calculations, we have performed DMPI cation intercalation into the graphite anode based on different plausible staging mechanisms. The intercalation energy characteristics indicate favourable intercalation of DMPI into graphite following the staging mechanism, which has been further confirmed by a simulated X-ray diffraction study. A higher cell voltage (3.7–4.6 V range) comparable to lithium-ion batteries along with a maximum capacity of 62 mA h g^{-1} has been achieved. Charge transfer analysis presents $+0.87 |e|$ charge transfer from DMPI to graphite, indicating DMPI cation intercalation into graphite during the charging process. Moreover, the metallic character of the DMPI cation intercalated graphite system and diffusion barrier as low as 0.2 eV suggest a constant electronic conductivity and better rate performance, respectively. Furthermore, we have explained the reason behind the inapplicability of the 1-ethyl-3-methyl imidazolium (EMI) cation as an organic cation for dual-graphite batteries and hence highlight the need to explore alternative ionic liquids. These results provide clear understanding of DMPI cation intercalation into graphite anodes and could be helpful in fabricating dual graphite electrode-based DIBs with better electrochemical performance compared to conventional DIBs using metal anodes.

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1. Introduction

Increasing utilization of renewable energy sources, like wind, hydroelectric and solar power, is essential today for a sustainable economy. Due to the volatility, randomness, and intermittency in such energy sources, they cannot be directly deployed for uninterrupted power supply. As a result, development of highly efficient energy storage and conversion systems is required to assist with the widespread usage of portable electronic devices in daily life.¹ Various types of rechargeable metal-ion batteries have

been developed in the recent past such as Li-ion,^{2–4} Na-ion,^{5–7} Mg-ion,^{8,9} Zn-ion^{10,11} and Al-ion^{12,13} batteries. Among them, lithium-ion batteries (LIBs) have profoundly dominated the energy storage market due to their high output voltage and gravimetric capacity.^{2,4} However, using expensive transition metal oxides as cathode materials and less abundant Li metal as the anode may not be sustainable for large-scale application of LIBs in future.^{14,15} Moreover, the safety issues and thermal runaway events are major concerns that have not been completely resolved.¹⁶ Altogether, the high cost and the safety issues related with LIBs make them a less suitable choice for green energy storage. Hence, new materials like graphite are being explored to develop better performance low-cost batteries. Owing to its redox-amphoteric nature and layered structure, graphite can be reduced as well as oxidized by electrochemical reactions with proper cation/anion uptake and release. Thus, dual graphite batteries can be conceptualized, also called dual-ion batteries (DIBs), where both a graphite anode and cathode are involved in intercalation/deintercalation of cations and anions, respectively,

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† Electronic supplementary information (ESI) available: Relative energies of the different binding sites, details of molecular dynamics analysis at different temperatures, comparative study of stage-1 and stage-2 with varying DMPI ions, pictorial representation of all four stages, NBO charge analysis of DMPI and EMI cations, XRD pattern of the DMPI intercalated graphite anode, stability study of different stages of the AlCl_3 intercalated graphite cathode, and NEB diffusion barriers. See DOI: 10.1039/d1ma00351h



during the charge/discharge cycle of a battery.^{17–20} In 1989, McCullough *et al.* first reported a DIB using graphite as both the cathode and anode along with the nonaqueous electrolyte of ClO_4^- .²¹ Graphite-based DIBs are considered as promising alternative rechargeable batteries because they have a high working voltage (4.5 V), better safety and lower cost compared to the conventional LIBs.^{22–25} In DIBs, the working principle is different from the rocking chair mechanism as in LIBs. Here, the cations and anions are both reversibly intercalated in the graphite anode and cathode simultaneously in the charging process, while the ions are deintercalated from graphite back to the electrolyte during the discharging process.^{26,27} Commonly used DIBs have lithium salts (like LiPF_6 and LiTFSI) with organic solvents (ethylene carbonate and ethyl methyl carbonate) as an electrolyte, which faces several fatal problems like (i) electrolyte decomposition at a high working voltage (>4.5 V),²⁸ (ii) graphite exfoliation due to solvent co-intercalation along with ion intercalation,²⁸ and (iii) active battery material loss due to solid–electrolyte interphase (SEI) formation.²⁹ Hence, compatible electrolytes and electrodes are required to reduce the exfoliation of graphite and replace lithium metal. Carlin *et al.* have presented room temperature ionic liquids (RTILs) as a better alternative to replace conventional volatile organic solvents.³⁰

Ionic liquids (ILs) are low melting point salts and have emerged as an important part in electrochemistry to design new classes of electrolytes. ILs show several encouraging properties for electrochemical applications such as low vapour pressure, a broad electrochemical window, and higher ionic conductivity.^{31–33} A special feature of using ILs as electrolytes is that the electrochemical reaction is mainly driven due to their high oxidative and reductive stability and thus they possess a large electrochemical stability window.^{34–36} Several recent studies have considered DIBs involving IL electrolytes and revealed that the organic cations of ILs such as the 1-ethyl-3-methyl imidazolium cation (EMI^+) and *N*-butyl-*N*-methyl pyrrolidinium (BMP^+) can be intercalated into the graphite anode.^{37–39} The broad electrochemical stability window of ILs allows for a high cut-off charging voltage (4.0–5.0 V).^{37–39} Carlin and co-workers have also studied ILs composed of cations like 1-ethyl-3-methyl imidazolium (EMI^+) and 2,3-dimethyl-1-propyl imidazolium (DMPI^+) and anions such as CF_3SO_3^- , AlCl_4^- , $\text{C}_6\text{H}_5\text{CO}_2^-$, PF_6^- and BF_4^- as electrolytes in dual ion batteries.^{25,30} Our previous reports have shown that the AlCl_4^- anion along with different IL based organic cations (imidazolium and pyrrolidinium) can be used as electrolytes in Al DIBs.^{33,40} Using organic cations as charge carriers has the added advantage of preventing unwanted co-intercalation of solvents into the graphite anode (as in the case of LiPF_6) as well as diminished chances of solvent decomposition at high charging voltage.

Consequently, apart from the intercalation mechanism, there is a lack of clear molecular level understanding of the intercalation capacity and intercalated species for organic cation intercalation in graphite-like anodes. In this study, we have computationally modelled, for the first time, an organic cation intercalated graphite anode system successfully. We have chosen imidazolium-based IL 2,3-dimethyl-1-propyl

imidazolium chloride (DMPI-Cl) with AlCl_3 at a 1 : 1 molar ratio as the electrolyte. We have investigated the intercalation mechanism by considering the staging manner of DMPI cation intercalation into the graphite anode. First principles calculations are performed for a systematic study of the structure, stability, electronic properties, and theoretical capacity along with the average voltage of the DMPI cation intercalated graphite electrode. Furthermore, we have also investigated the diffusion pathways of DMPI cations in the graphite system. Along with this, we have also explained why EMI^+ cannot be used as an organic cation for dual-graphite battery research, although EMI based IL electrolytes are extensively used in battery technology.^{12,41,42} On the basis of the obtained theoretical insights, we believe that organic cation intercalated graphitic anodes could be utilised in DIBs to achieve better electrochemical features compared to DIBs using metal anodes, and our study would motivate further developments in cheaper dual graphite battery technology.

2. Computational details

First principles calculations have been performed using the Vienna Ab initio Simulation Package (VASP).^{43,44} All geometry optimizations and respective calculations have been performed using the generalized gradient approximation of the Perdew–Burke–Ernzerhof (GGA-PBE) functional.⁴⁵ The projected augmented wave (PAW) method is used for treating the interaction between the ion core and valence electrons.⁴⁶ An energy cut off of 470 eV has been implemented. During structural optimization, the Brillouin zone has been sampled with a Γ centered k -point grid of $11 \times 11 \times 5$ for the unit cell and $2 \times 2 \times 1$ for the considered supercells. All structures have been optimized with an energy criterion of 10^{-5} eV \AA^{-1} and force criterion of ≤ 0.01 eV \AA^{-1} for all the atoms to obtain full relaxation between the atomic and lattice positions. The DFT-D3 approach has also been considered for van der Waals corrections in our calculations.⁴⁷ We have modelled four different stages for both the DMPI cation and AlCl_4^- anion intercalation process, where a $6 \times 6 \times 2$ supercell of graphite containing 288 carbon atoms for the stage-1, stage-2 and stage-4 systems and a $6 \times 6 \times 3$ supercell containing 432 carbon atoms for the stage-3 calculation have been chosen with the same concentration of both cations and anions. Density of states (DOS) calculations have been performed for a $4 \times 4 \times 1$ supercell with a Γ centered k -point mesh of $9 \times 9 \times 1$. Bader charge analysis has been performed using the Henkelman program to determine the quantitative charge transfer upon intercalation of the DMPI cation into graphite layers.^{48–50} *Ab Initio* Molecular Dynamics (AIMD) simulations have been performed with the *NVT* ensemble for a broad temperature range of 300–600 K for a 5 picosecond (ps) timescale. The Nosé–Hoover thermostat⁵¹ with a Nosé mass parameter of 0.01 has been implemented to control the temperature fluctuations. The activation barriers for the DMPI cation diffusion pathways have been calculated using the climbing image nudged elastic band method (CI-NEB).⁵² The minimum energy paths (MEPs) are



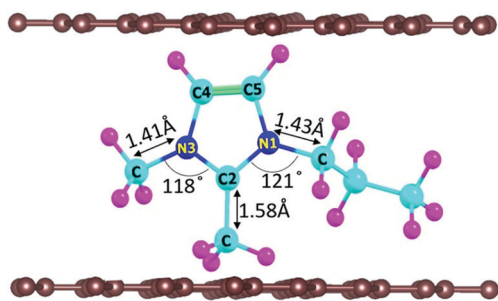
initialized by considering six image structures between the fully optimized initial and final geometries with an energy convergence criterion of 10^{-3} eV. During the charging process, the net electrons obtained from the external circuit (with a very small time lag) induce the intercalation of DMPI cations, which subsequently transfer their charge, thereby reducing graphite. The reverse phenomenon is occurring for the AlCl_4 anion transferring charge to consequently oxidise the graphite cathode. This time lag is very short compared to the time required to reach thermodynamic equilibrium of the DMPI-graphite system. Hence, in our DFT study we have considered a neutral system of both DMPI-graphite and AlCl_4 -graphite as we carry out our calculations in equilibrium conditions. Such consideration of neutral systems has yielded electrochemical performance results comparable to experimental reports in the past.^{27,54} Zero-point energy (ZPE) corrections have also been included for the diffusion barrier calculation using the formula $\text{ZPE} = \sum_i \frac{h\nu_i}{2}$, where h is Planck's constant and ν_i is the vibrational frequency. The ZPE is calculated by considering the degrees of freedom of intercalated DMPI into the graphite system.

3. Results and discussion

3.1. Structural changes and stable binding sites

In DIBs, the cations (DMPI in our case, Scheme 1) and anions are reversibly intercalated/deintercalated from the cathode and anode in the charging/discharging process, respectively. Hence, it is important to study the structural changes upon intercalation. We have chosen two possibilities of DMPI cation intercalation in the graphite layer, where one is oriented perpendicular to the graphite layers and the other is parallel to the graphite layers (Fig. 1a and b).

Owing to the strong van der Waals interaction between the imidazolium ring and graphite layers, the parallel orientation of the DMPI cation into the graphite layer is found to be stable. The perpendicular orientation shifts to parallel upon relaxation with the dispersion correction. However, without the vdW correction the perpendicular orientation is retained, which proves that the extensive interactions between π -stacked graphite and the parallel orientation (imidazole ring) of DMPI are responsible for this configuration being more stable.



Scheme 1 Considered intercalant cation, 2,3-dimethyl-1-propyl imidazolium (DMPI), with labelled atoms. Here, the blue, cyan, magenta, and brown colours indicate N, C, H, and graphite layer carbon atoms, respectively.

We have further investigated the structural distortions such as changes in bond lengths and bond angles obtained due to the anisotropic van der Waals interaction between layered graphite and the DMPI cation. Upon intercalation of DMPI, the $\angle \text{C2-N3-C}$ bond angle decreases from 124° to 118° , and $\angle \text{C2-N1-C}$ decreases from 126° to 121° . Similarly, bond length changes are also observed in intercalated DMPI, where the bond lengths of N1-C, N3-C and C2-C change to 1.43, 1.41 and 1.58 Å from 1.47, 1.34 and 1.48 Å in free DMPI, respectively. The alkyl group variance between DMPI and EMI is that extra methyl and propyl groups are present in the C2 and N1 positions for DMPI, respectively, compared to EMI. In our earlier report we have studied the effect of alkyl group variation in the N1 position of the imidazolium ring.³³ But a fundamental understanding is required on the effect of addition of a methyl group in the C2 position of the imidazolium ring. So, with respect to the C2 position of DMPI, we have arranged four possible sites of binding in graphite layers, (i) S1 (Top), (ii) S2 (Bridge 1), (iii) S3 (Bridge 2) and (iv) S4 (Hollow), as shown in Fig. 1c-f. In the S1 (Top) site, C2 of DMPI occupies the top position of a C atom of graphite. In the S2 (Bridge 1) site, C2 occupies the bridging position between two non-bonded carbon atoms, whereas in the S3 site (Bridge 2) it occupies the bridging position of the C-C bond. Similarly, for the S4 (Hollow) site C2 occupies the centre of hexagons (C_6) of the graphite layer. Our relative energy calculations in Table S1 (ESI[†]) show that the hollow site with respect to the C2 position is most stable for DMPI cation binding into the graphite layers. Particularly, the hollow and top site have a very small (0.009 eV) relative energy difference. Hence, there is a high possibility of the equilibrium existence of both hollow and top sites. We have carried out AIMD simulations to verify the thermal stability of the DMPI cation intercalated graphite system as well as the equilibrium existence of both hollow and top binding sites. Upon AIMD simulation at 300 K for 5 ps, we have observed that, while the AB stacking of graphite layers is unaltered, the intercalated DMPI cation shifts from the most stable site (Hollow) to the second most stable (Top) site of binding as shown in Fig. S1 (ESI[†]). Owing to the very small energy difference (~ 0.009 eV), these structures can change the adsorption sites at the expense of thermal energy ($3/2 k_B T = 0.03$ eV, where k_B and T are Boltzmann's constant and temperature, respectively). As the calculated relative energies are less than the DFT intrinsic error, therefore all binding sites are equally accessible and DMPI can keep on changing binding sites. This has also been observed during the AIMD simulation as shown in Fig. S2 (ESI[†]). For sites B1 and B2, the relative energies are 0.04 and 0.05 eV, respectively, which are higher than the thermal energy (0.03 eV). Upon further simulation of DMPI bonded at the most stable hollow site of graphene at 300–600 K, the system was found to be stable (Fig. S3, ESI[†]). The bond lengths and bond angles of the DMPI cation do not change significantly, although the position and orientation of DMPI do fluctuate a little. Hence, from the simulation study we could confirm that these two sites (Hollow and Top) are equally stable, and we have considered the hollow site for our further studies.



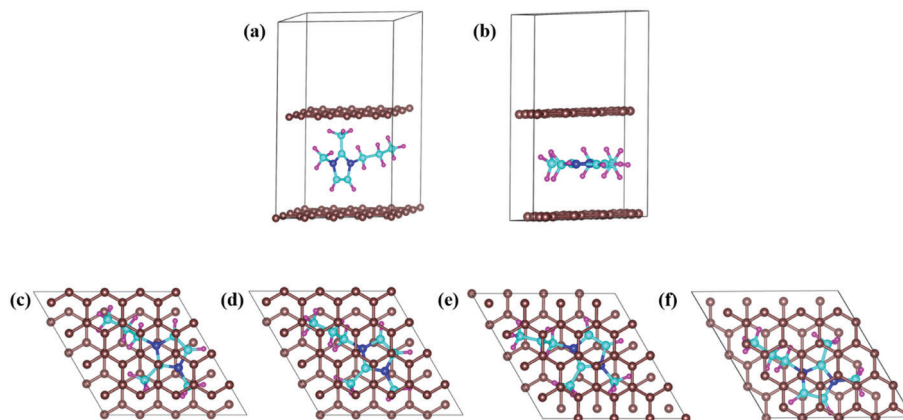


Fig. 1 DMPI intercalated structures: (a) perpendicular orientation, and (b) parallel orientation. Optimized structures of the DMPI intercalated system; (c) S1 (Top), (d) S2 (Bridge 1), (e) S3 (Bridge 2), and (f) S4 (Hollow). Here, the brown, blue, cyan, and magenta colours represent the graphite layer, and N, C, and H of the DMPI cation, respectively.

3.2. Model system

In DIBs, graphite intercalation compounds (GICs) are formed between the graphite layer and intercalant species.⁵³ The formation of GICs follows a systematic staging mechanism of intercalant insertion based on various competing interactions among the intercalate and intercalant. The van der Waals interaction between the parallelly stacked graphite layers needs to be overcome by the incoming DMPI cations. Again, the intercalated DMPI cations may experience intermolecular repulsive forces among themselves. Hence, the intercalation phenomena can proceed by maintaining a balance among these competing forces.^{54–57} Consequently, different stages of intercalation can be considered to occur during the charging/discharging process, which has been observed experimentally,^{55,58} as well as theoretically.^{54,56} The staging features can be represented by the stage index, which stands for the number of unintercalated graphene layers present between the intercalated layers. The stage-*n* system refers to the “*n*” number of graphene sheets or “*n* – 1” empty host galleries present between two intercalated layers. For instance, stage-1 GICs correspond to intercalation happening after every graphene layer, while in stage-2 intercalation will happen after a gap of two empty layers. In this manner, different stages are expected in DIBs during intercalation

of DMPI cations into graphite. Hence, we have chosen the four most favourable intercalation stages for the same DMPI cation concentration. For that we have modelled a $6 \times 6 \times 2$ supercell of graphite containing 288 carbon atoms for the stage-1, stage-2 and stage-4 systems and their corresponding relative energies have been compared as shown in Fig. 2. The stage-2 configuration is observed to be the most stable for DMPI cation intercalation, followed by stage-4 and stage-1 for the same concentration of DMPI cations. The energy required to overcome the repulsive forces between intercalant species is less than the energy required to overcome the van der Waals forces between graphite layers required for gallery height opening for intercalation. Hence, DMPI intercalation favours stage-2 rather than stage-1. Stage-4 is less stable than stage-2, which may be due to the van der Waals repulsion among the long carbon chains of four DMPI cations intercalated into a single graphene layer rather than two DMPI intercalated in two separate layers as in stage-2. Upon increasing the number of DMPI cations intercalating to 8, the relative energy difference between stage-1 and stage-2 intercalation decreases (Fig. S4, ESI[†]), thus indicating that stage-1 would be more stable compared to the other stages for higher concentrations of the DMPI cation. Owing to the different supercell size of stage-3, the result is not comparable to the other stages. However, the DMPI

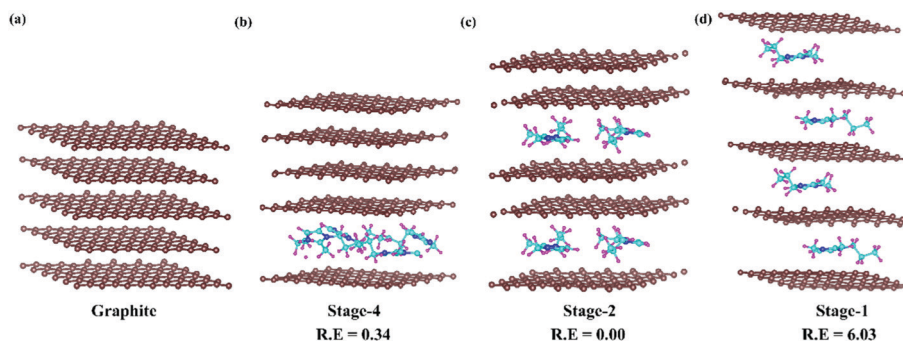


Fig. 2 Systematic representation of (a) pristine graphite and the staging mechanism of DMPI cation intercalation, (b) stage-4, (c) stage-2 and (d) stage-1. R. E. is in eV units. Here, the brown, blue, cyan and magenta colours represent the graphite layer, and N, C, and H of the DMPI cation, respectively.



intercalation mechanism is followed during charging, and the same deintercalation mechanism is followed at the time of the discharging process. Overall, from the theoretical stability calculation, we could describe the formation of different DMPI cation intercalated stages.

Furthermore, we have considered the maximum possible intercalation in our supercells for each stage. A $6 \times 6 \times 3$ supercell containing 432 carbon atoms has been chosen for the stage-3 calculation. In the case of stage-1, a 288 carbon containing graphite system can intercalate a maximum of 16 DMPI cations to give a stable GIC, which can be represented by the formula unit $[C_{288}(\text{DMPI})_{16}]$. Other lower stoichiometry of DMPI has been chosen such as 4, 8 and 12 per 288 carbon atoms of graphite, maintaining the criteria of stage-1 characteristics. Thus, the generalised formula unit would be $[C_{288}(\text{DMPI})_n]$ where the n values are 4, 8, 12, and 16 for stage-1. Similarly, for the other stages formula units have been considered such as $n = 2, 4, 6,$ and 8 for stage-2; and $n = 1, 2, 3,$ and 4 for stage-4. For stage-3, the formula unit would be $[C_{432}(\text{DMPI})_n]$ where $n = 2, 4, 6,$ and 8 DMPI. All optimised structures have been provided in the ESI† (Fig. S5–S8). Overall, from the different considered stoichiometries, we have tried to get an idea about the maximum storage capacity of these systems and compared our results with experimental reports to investigate more about the staging mechanism followed by DMPI cation intercalation in graphite, Fig. 3.²⁵

3.3. Binding energy

The binding energy is a good criterion to investigate the stability of intercalated graphite systems. We have calculated the binding energy for various numbers of DMPI intercalation into graphite following different staging mechanisms, using the following equation

$$E_{\text{Binding}} = \frac{E_{[(\text{DMPI})_x C_m]} - E_{C_m} - x E_{\text{DMPI}}}{x} \quad (1)$$

where x is the number of DMPI cations, and $E_{[(\text{DMPI})_x C_m]}$, E_{C_m} and E_{DMPI} are the total energies of the DMPI intercalated graphite system, the bulk graphite system and a single DMPI cation, respectively. Hence, a more negative binding energy value

Table 1 Binding energy per DMPI cation (eV) and interlayer distance (Å) for all the stages with different concentrations

Stages	No. of DMPI cations	Binding energy (eV)	Interlayer distance (Å)
1	4	−2.48	6.92
	8	−3.58	7.14
	12	−3.90	7.16
	16	−3.59	7.54
2	2	−2.51	5.23
	4	−3.99	5.30
	6	−4.01	5.39
	8	−3.80	5.53
3	2	−2.45	4.64
	4	−4.04	4.61
	6	−4.26	4.71
	8	−3.76	5.03
4	1	−2.64	4.37
	2	−4.17	4.39
	3	−4.54	4.38
	4	−3.90	4.47

indicates the more feasible intercalation of the DMPI cation into the graphite system as shown in Table 1. At a low concentration of DMPI cations (4 DMPI), intercalation into graphite is less favourable for every stage with lower binding energy (~ 2.5 eV) compared to intercalation at higher concentrations. This low binding energy may be due to the activation barrier which needs to be overcome during gallery height opening against the interplanar van der Waals interactions among graphite layers. Further DMPI cation intercalation into the already opened galleries becomes more feasible in the case of each stage with higher calculated binding energies. This phenomenon is also reported in several experimental studies on DIBs, where the experimental coulombic efficiency is very poor in the first few cycles.^{25,27} Moreover, the binding energy (~ -3.5 eV) of DMPI cation intercalation is found to be higher than anion (AlCl_4^- : -1.5 eV; PF_6^- : -2.9 eV) binding in graphite, which is possibly due to the π - π interaction between the DMPI imidazolium ring and the graphite layers.^{54,55} However, upon approaching maximum DMPI cation intercalation the binding energy again decreases for each stage, which could be due to the electrostatic repulsion between neighbouring DMPI cations in the same gallery. As the binding energy value is still negative, the effective repulsive forces between cations must be less than

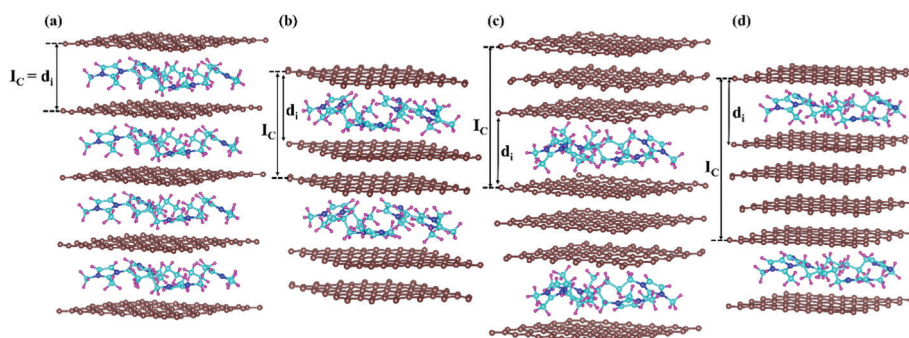


Fig. 3 Schematic representation of all optimised structures of the four different DMPI cation intercalated systems, (a) stage-1, (b) stage-2, (c) stage-3 and (d) stage-4, where l_c and d_i are the periodic repeating distance and intercalant gallery height, respectively, and the brown, blue, cyan, and magenta colours represent the graphite layer, and N, C, and H of the DMPI cation, respectively.



the attractive forces due to intercalation. In an experimental report by Lv *et al.*, a higher cut-off working voltage has been considered to maintain better coulombic efficiency, which could be due to our calculated high binding energies (2.4–4.5 eV) for DMPI cation intercalation in the graphite anode for all the possible stages.²⁵ Additionally, we have also tried to address the advantage of using DMPI rather than EMI for intercalation into the graphitic anode for dual graphite batteries. We have calculated the binding energy of EMI cation intercalated graphite layers. Our calculated low binding energy (+0.007 eV for stage-1) for EMI cation intercalation validates the infeasibility of using EMI with graphite anodes, which has been reported experimentally.⁵⁹ Hence, very poor electrochemical performance is observed which can be ascribed to the lack of alkyl substituents in the C2 position of the imidazolium ring of the EMI cation. The imidazolium cations with protons in the C2 position are less stable inside the graphite layers. This observation can be explained by applying natural bond orbital (NBO) charge calculations,⁶⁰ where the C2 carbon of DMPI has almost double the charge (0.493 |e|) of the C2 carbon (0.288 |e|) of EMI (Fig. S9, ESI†). So, the resultant interaction of DMPI–graphite is more compared to EMI–graphite. This is also reflected in the experimental XRD studies, where any characteristic peak is absent for the EMI–graphite system.⁵⁹ Overall, from the binding energy calculations, we can conclude that the DMPI cation has fair binding strength, which is essential in DIBs, and graphite can be used as a potential anode.

3.4. Staging mechanism and XRD analysis

To determine the staging mechanism actually occurring during DMPI intercalation into graphite, we have simulated XRD patterns for our optimised intercalated systems. We have compared our simulated XRD patterns with the experimental results.²⁵ We notice that the graphite structure gets deformed upon intercalation of the DMPI cations along with an increase in the interlayer spacing. As the gallery height (3.35 Å) of graphite is less than the DMPI cation size (4.6 Å),⁶¹ the graphite interlayer spacing increases in the very first intercalation step. After the intercalation of DMPI cations, the average gallery height expansion in graphite is found to be 115%, which is comparable to reports of various anion intercalation in graphite.^{54,55} The average interlayer distance can be calculated by the formula,

$$\text{Average interlayer distance} = \frac{\text{total height of the stage} - n \text{ having } n - 1 \text{ galleries or } n \text{ graphite layers}}{n - 1 \text{ galleries or } n \text{ graphite layers}}$$

The average interlayer distance for all the stages has been included in Table 1. The distance between two layers of DMPI intercalated graphite *i.e.*, the intercalant gallery height (d_i) remains similar (~7.1–7.6 Å) for all the stages, which is comparable with the experimental report (7.16 Å) for DMPI intercalation.⁵⁹ Our simulated XRD pattern of the DMPI intercalated system for all the stages ($n = 1$ –4), as shown in Fig. 4, helps to understand the structural changes in the graphite system. The XRD patterns suggest the structural changes in graphite upon intercalation with characteristic shifting of the

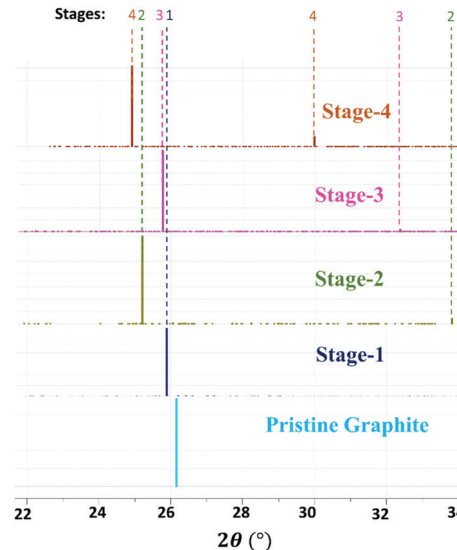


Fig. 4 Schematic diagram of the simulated XRD pattern of pristine graphite and different stages of DMPI intercalated systems where 2θ ranges from 22 to 34°.

(002) peak of graphite. To index the stage number (n) of the GICs, two characteristic peaks are observed, the $(00n + 1)$ and $(00n + 2)$ planes along the stacking direction, in the XRD analysis. d_{00n+1} represents the spacing between adjacent layers. The intercalant gallery height (d_i), gallery expansion (Δd) and periodic repeating distance (I_c) of a particular stage index ' n ' are calculated using the d -spacing values as $I_c = d_i + (n - 1) \times 3.35 = \Delta d + n \times 3.35 = (n + 1) \times d_{00n+1}$, and similarly for d_{00n+2} , $I_c = (n + 2) \times d_{00n+2}$, where $\Delta d = d_i - 3.35$ and 3.35 Å is the interlayer distance between two consecutive graphite layers.⁶² In the case of pristine graphite an intense peak is found at $2\theta = 26.5^\circ$ belonging to its (002) plane generally.¹² For our simulated XRD pattern, we can identify such a peak at 26.16° for the pristine system which shifts to 25.88° , 25.18° , 25.75° and 24.96° for stage-1, stage-2, stage-3 and stage-4 intercalation, respectively, as shown in Fig. 4. As the distortion and lateral displacement in the graphite layer stacking increase with an increase in the DMPI intercalation, other intense peaks belonging to various planes can also be identified at lower ranges of 2θ in the XRD plot as presented in Fig. S10 (ESI†). In the previous report of experimental DMPI intercalation by

Lv *et al.*, the (002) peak of the system was found to shift from $2\theta = 26.4^\circ$ for pristine graphite to $2\theta = 26.0^\circ$ for the GIC.²⁵ Consequently, from our XRD plots we can suggest that stage-1 intercalation ($2\theta = 25.88^\circ$) of the DMPI cation in graphite occurs predominantly (Fig. S4, ESI†), which also agrees with an experimental study by Sutto *et al.*⁵⁹

3.5. Electrochemical properties

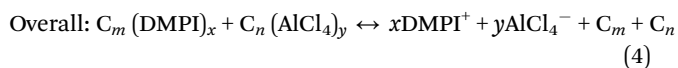
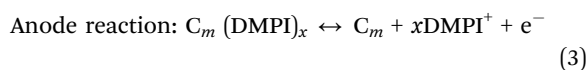
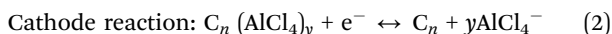
In this section, we have discussed the electrochemical performance of a DIB with dual-graphite electrodes, where



DMPI cations are paired with AlCl_4 anions. Upon charging, the DMPI cations and AlCl_4 anions are expected to undergo intercalation into the graphite anode and cathode, respectively, while during discharging both the ions are expected to diffuse back to the electrolyte. Hence, to investigate the electrochemical properties of DMPI with the graphite anode, it is necessary to also include the AlCl_4 anion intercalation into the graphite cathode.

3.5.1. AlCl_4 intercalation into the graphite cathode. The anion intercalation phenomenon for layered structures like graphite electrodes has been well studied both theoretically and experimentally.^{12,54} Our previous study has also established the exact staging mechanism pathway for AlCl_4 anion intercalation into the graphite cathode.⁵⁴ Motivated by those studies, we have modelled four stages ($n = 1-4$) for AlCl_4 anion intercalation keeping the concentration of AlCl_4 anions fixed. From our relative energy calculations, we could confirm that in the initial period of the charging process stage-4 AlCl_4 intercalation is more favourable (Fig. S11, ESI[†]). From our simulated XRD data, we could identify the (005) and (006) planes at $2\theta = 23.32^\circ$ and 28.07° with a d -spacing of 3.81 and 3.17 Å, respectively. For the experimental diffraction patterns, 2θ values are obtained at 23.7° and 28.5° with a d -spacing of 3.77 and 3.15 Å, which are comparable with our simulated patterns.¹²

3.5.2. Voltage calculations. The open circuit voltage (OCV) is one of the most important parameters in deciding on a suitable battery, which stands for the output voltage of a battery at full intercalation capacity or at full charge in other words. In the DIB system, the IL electrolytes (DMPI- AlCl_4) behave not only as a charge carrier but also as an active material towards electrochemical activity. The OCV depends upon both the cathode and anode reactions of the DIB system. The discharge voltage can be calculated from the following cathodic and anodic reactions,



where x and y are the number of DMPI cations and AlCl_4 anions, respectively. C_m and C_n are the graphite anode and graphite cathode, respectively. From the above reaction it is evident that the reaction mechanism of the DIB also includes a contribution from the electrolyte medium.

The cell voltage (V) can be calculated from the Nernst equation, $V = \frac{-\Delta G_{\text{cell}}}{zF}$, where z and F are the number of electrons transferred and the Faraday constant, respectively, while ΔG_{cell} is the change in Gibbs free energy during the chemical reactions.

$$\Delta G_{\text{cell}} = \Delta H_{\text{cell}} - T\Delta S_{\text{cell}}; \Delta G_{\text{cell}} = \Delta E_{\text{cell}} + P\Delta V_{\text{cell}} - T\Delta S_{\text{cell}}$$

As our calculations are performed at a 0 K temperature, so the changes in volume (ΔV_{cell}) and entropy (ΔS_{cell}) of the reactions

are negligible. Thus, the Gibbs free energy change only depends upon the change of internal energy (ΔE_{cell}).

Therefore, the internal energy change of the overall reaction would be,

$$\Delta E_{\text{cell}} = \{xE_{\text{DMPI}^+} + yE_{\text{AlCl}_4^-} + E_{C_m} + E_{C_n}\} - \{E_{[C_n(\text{AlCl}_4)_y]} + E_{[C_m(\text{DMPI})_x]}\} \quad (5)$$

where $E_{[C_n(\text{AlCl}_4)_y]}$, $E_{[C_m(\text{DMPI})_x]}$, $E_{\text{AlCl}_4^-}$ and E_{DMPI^+} are the total energy of the AlCl_4 intercalated graphite system and DMPI intercalated graphite system, and the AlCl_4 and DMPI fragments, respectively. E_{C_m} and E_{C_n} are the total energy of the graphite system. $E_{\text{AlCl}_4^-}$ and E_{DMPI^+} are calculated by optimizing the AlCl_4 anions and DMPI cations as a molecular species due to the nonavailability of their crystal structure.

Then the average voltage would be,

$$V = \frac{\{E_{[C_n(\text{AlCl}_4)_y]} + E_{[C_m(\text{DMPI})_x]}\} - \{xE_{\text{DMPI}^+} + yE_{\text{AlCl}_4^-} + E_{C_m} + E_{C_n}\}}{z} \quad (6)$$

To understand the exact mechanism of intercalation in a DMPI- AlCl_4 DIB, we have considered the different staging intercalation behaviour of both cations and anions at the initial and final periods of the charging process. Therefore, based on our model we have considered all combinations of staging behaviour possible upon intercalating four cations and anions into the graphite anode and cathode to replicate the initial charging environment, which is presented in Table 2. The highest calculated voltage is 5.25 V, which is obtained from the combination stage-2 DMPI-stage-4 AlCl_4 intercalated system. Experimental evidence suggests that the voltage plateaus range within 3.1–4.3 V with the cut-off charging voltage at 4.4 V²⁵ indicating that the discharging voltage would remain lower than 4.4 V. However, our observation and literature suggest that stage-4 intercalation in the case of AlCl_4 is more feasible compared to the other stages during the initial charging process. The combinations of stage-3 DMPI-stage-4 AlCl_4 and stage-4 DMPI-stage-4 AlCl_4 yield voltages of 5.23 and 5.16 V, respectively, which are also higher than the considered cut-off voltage in the experimental report. Interestingly, the combination of stage-1 DMPI-stage-4 AlCl_4 gives an output voltage of 3.74 V, which matches with the experimental voltage range 3.1–4.3 V. In addition, our calculated results and experimental evidence support that stage-1 formation of DMPI is more feasible than the other stages.^{25,59} Thus we propose that,

Table 2 Calculated theoretical voltage (V) with different stages of DMPI cations and AlCl_4 anions

AlCl_4	DMPI			
	Stage-1	Stage-2	Stage-3	Stage-4
Stage-1	1.61	3.12	3.17	3.03
Stage-2	2.10	3.61	3.66	3.52
Stage-3	3.25	4.76	4.80	4.67
Stage-4	3.74	5.25	5.23	5.16



during the initial charging process, stage-1 DMPI–stage-4 AlCl_4 intercalation occurs. However, when the charging process approaches completion, both DMPI and AlCl_4 ions are expected to follow a stage-1 type of intercalation into the graphite system. Considering the combination of stage-1 DMPI–stage-1 AlCl_4 results in an OCV of 4.6 V for maximum intercalation (16 cations and anions) at both the electrodes, which is comparable with the upper limit of the reported experimental voltage.²⁵

Furthermore, we have calculated the gravimetric capacity (C) for our considered DIB system. The gravimetric capacity of both the cathode and anode can be calculated using the following equation,⁵⁴

$$C_{\text{cathode}} = C_{\text{anode}} = \frac{nxF}{M_f} \quad (7)$$

where n is the number of electrons transferred per formula unit, x is the number of ions intercalated, F is the Faraday constant, M_f is the mass of the formula unit, and C_{cathode} and C_{anode} are the capacity of the cathode and anode, respectively. The total cell capacity (C_{total}) can be expressed as a combination of the individual capacity of both the cathode and anode⁶³

$$C_{\text{total}} = \frac{1}{\frac{1}{C_{\text{anode}}} + \frac{1}{C_{\text{cathode}}} + \frac{1}{C_M}} \quad (8)$$

where $1/C_M$ is the specific mass of the other cell components such as the electrolyte, separator, and current collector. $1/C_M$ varies significantly with the cell to cell design process, so the contribution of $1/C_M$ in the total theoretical cell capacity is neglected. Therefore the theoretical cell capacity can be given by,

$$C_{\text{total}} = \frac{C_{\text{anode}} \times C_{\text{cathode}}}{C_{\text{anode}} + C_{\text{cathode}}} \quad (9)$$

Our calculated gravimetric capacity for the early stage of the charging process (four intercalated ions) is 15 mA h g^{-1} , which reaches a maximum capacity of 62 mA h g^{-1} upon full intercalation (sixteen ions). In the experimentally produced DMPI– AlCl_4 DIB, the initial gravimetric capacity was reported to be 27 mA h g^{-1} , which could further be increased up to 82 mA h g^{-1} at higher current density with subsequent loss in the coulombic efficiency of the cell.²⁵ During the initial charging process by considering four ion intercalation, as the DMPI cations follow stage-1 and AlCl_4 anions follow stage-4 intercalation, a partial gravimetric capacity of 15 mA h g^{-1} is calculated. In a fully charged system considering sixteen ion intercalation, both DMPI and AlCl_4 ions follow stage-1 intercalation, resulting in a maximum capacity of 62 mA h g^{-1} . Based on our theoretical insights and available experimental observations, we believe that graphite-like layered materials can be further explored as anode materials for intercalating large size organic cations in order to achieve higher output voltages ($\sim 4.6 \text{ V}$) compared to batteries based on intercalation of AlCl_4^- ions only ($\sim 2.0 \text{ V}$).^{12,54} Moreover, we have calculated the volume expansion for DMPI intercalation in graphite to be $\sim 115\%$. So, the expansion due to DMPI intercalation is less compared to AlCl_4 intercalation in the

graphite system (150–160%).⁵⁴ However, the reported cycle life values (1000 cycles) are less for the DMPI–graphite system, which could be due to the strong interaction between DMPI and graphite, hindering repeated intercalation/deintercalation cycles. Along with this, we have also tried to provide rationale for the experimentally observed self-discharge nature of the DMPI– AlCl_4 DIB with the help of a free energy calculation of the cell reactions. The cathodic reaction for AlCl_4 intercalation into the graphite cathode (reaction 2) is well established in DIB studies and AlCl_4 anions are not known to actively participate in the self-discharge process. The anodic reaction of DMPI intercalation into graphite is calculated to have a reaction free energy of $+0.08 \text{ eV}$, which is much lower than that of the cathodic reaction ($+5.6 \text{ eV}$). Thus, in a bias free discharging environment, DMPI deintercalation is expected to occur much spontaneously compared to AlCl_4 deintercalation. This partiality in spontaneity of the individual cell reactions may be the reason for the DMPI– AlCl_4 battery experiencing a self-discharge process, which needs to be further investigated.

3.6. Electronic properties

We have studied the electronic structure of the DMPI cation intercalated graphite system to understand the nature of the interaction between the DMPI cation and graphite host, as organic cation intercalation is less common compared to anion intercalation into the layered graphitic structure. During charging, electrons flow from the cathode to the anode through the external circuit, which allows the adsorption/intercalation of cations and anions into the anode and cathode material, respectively. Similarly, during discharge, the reverse phenomenon happens as electrons flow from the anode to the cathode, thus providing electricity. Hence, constant electronic conductivity of the electrodes during the charge/discharge cycle is essential in a battery. The electronic conductivity can be interpreted from the DOS plots on the basis of the nature of the Fermi region. Fig. 5a presents the total DOS (TDOS) and projected DOS (PDOS) for the DMPI intercalated graphite system. Graphite is known to show good in-plane electronic conductivity because of the presence of the p_z electronic states at the Fermi level and no interplane conductivity as the s , p_x and p_y electronic states are far from the Fermi level. In comparison with the DOS plot of the well-known

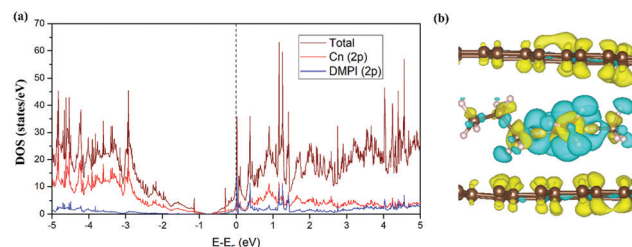


Fig. 5 (a) TDOS and PDOS of the DMPI cation intercalated graphite. The Fermi level is set to zero. (b) Isosurface ($0.0008 |e| \text{ \AA}^{-3}$) for the CDD plot of the DMPI intercalated system, where the yellow colour indicates electron density accumulation and the cyan colour indicates electron density depletion.



pristine graphite system, in our DOS plot of cation intercalated graphite, the Fermi level is shifted towards the conduction band due to the charge transfer from DMPI to graphite. Hence, the electronic conductivity is maintained as shown by the presence of electronic states at the Fermi level in Fig. 5a. The DMPI-graphite system is found to be metallic in nature due to overlap between the 2p orbital of the imidazolium ring of the cation containing C and N atoms and the 2p orbitals of graphite C atoms. The shifting of the Fermi level towards the conduction band indicates the reduction of graphite layers, which satisfies the criteria for selection of a suitable DIB anode material.

The qualitative charge transfer between the DMPI cation and graphite layers can be illustrated by plotting the charge density difference (CDD) of the systems before and after intercalation. Thus, the CDD can be calculated from the following equation,

$$\rho_{\text{CDD}} = \rho^{\text{total}} - \sum_i \rho_i^{\text{fragments}} \quad (10)$$

where ρ^{total} is the total charge density of the DMPI-graphite system and $\rho^{\text{fragments}}$ is the charge density of the individual fragments. The CDD plot has been shown in Fig. 5b, where the yellow colour represents charge density accumulation and the cyan colour represents charge density depletion. From the CDD plot, we can infer that charge from the loosely bound π -electron cloud of the DMPI imidazolium ring is transferred to the graphite layers. Furthermore, the C2 containing methyl group of DMPI also transfers a lesser amount of charge to the graphite layers. The charge transferred from the propyl chain of the DMPI cation is less because of the tightly bound σ -electrons of the C-C and C-H bonds. Overall, net charge is gained by the graphite layers and lost by the DMPI cation, thus indicating a partial ionic interaction between them. To justify this qualitative explanation, we have determined the quantitative charge transfer using Bader charge analysis. Quantitatively, 0.87 $|e|$ charge transfer is found to occur from DMPI to the graphite layer, which signifies the electrochemical reduction of graphite and DMPI being cationic, during the charging process. Overall from the DOS and CDD study, we can conclude that graphite can also be used as an anode material for DMPI cation intercalation as well as a cathode material for AlCl_4 anion intercalation, thus behaving as an amphoteric electrode material for dual graphite battery technology.^{17,19,20}

3.7. Diffusion pathway of DMPI cations

The fast charge/discharge rate of the DIB can be explained by the diffusion barriers of DMPI cations in the graphite layers. The diffusion barrier has a high impact on the battery performance, where the charging/discharging rate can be determined by the mobility of DMPI cations on the AB stacked plane of the graphite layers. Hence, trouble-free diffusion and high mobility of DMPI cations are a prerequisite for developing a dual graphite battery with a fast charge/discharge rate. To determine the diffusion barriers of the DMPI cation within the graphite lattice using the CI-NEB method, we have considered four minimum energy pathways (MEPs) of diffusion initiating from the most stable site (Hollow) of the DMPI cation to the next most stable binding

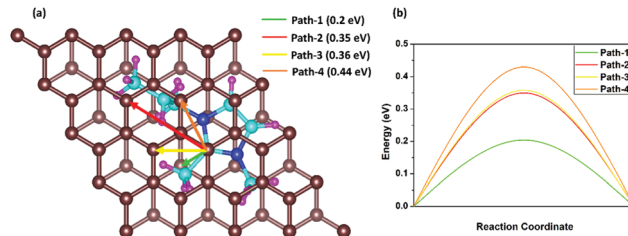


Fig. 6 Schematic representation of the four diffusion pathways, (a) nearest possible most stable hollow site in AB stacking, where the green, red, yellow and orange colour arrows indicate path-1, path-2, path-3 and path-4, respectively. (b) Energy profile diagram of the four diffusion processes, where the path-1, path-2, path-3 and path-4 energy barriers are 0.2 eV, 0.35 eV, 0.36 eV and 0.44 eV, respectively.

site (Hollow) on the AB stacked graphite in Fig. 6a.⁵² The MEPs have been shown in Fig. 6b and the structural energy profile diagram has been displayed in Fig. S12 (ESI[†]). For diffusion of DMPI cations along path-1, a minimum energy barrier of 0.2 eV is obtained. Similarly, for path-2, path-3, and path-4 the diffusion barrier is calculated to be 0.35, 0.36, and 0.44 eV, respectively.

Our calculated diffusion energy barrier values are very much comparable with the various reported alkali metal cation intercalated graphite systems such as for Li or Na (0.2–0.4 eV).^{64,65} Therefore, our diffusion studies indicate that the organic DMPI cation intercalated anode can be used for high performance DIBs with a high discharge voltage, low diffusion barrier, and high electronic conductivity.

4. Conclusion and outlook

In this work, we have carried out a systematic computational investigation of organic cation (DMPI) intercalation into the graphite anode for the first time. Based on dispersion-corrected density functional theory calculations, we have investigated the staging mechanism of DMPI cation intercalation, the charge transfer mechanism from the graphite anode, the diffusion barrier, and electrochemical properties like the voltage and specific capacity. Here, we have modelled four different stages (stage-1, stage-2, stage-3, and stage-4) of the DMPI cation intercalated graphite system to investigate the staging mechanism and electrochemical properties. To verify the stability, we have calculated the binding energy of the DMPI cation intercalated graphite system, where the DMPI cation prefers to intercalate in the parallel over the perpendicular orientation into the AB stacked graphite plane due to π - π interactions of the aromatic imidazolium cation with the graphite hexagonal rings. Further, we have validated the thermal stability of DMPI cation intercalation at the stable hollow site of AB stacked graphite using AIMD simulations. We also report the existence of a thermal equilibrium between all the possible sites of DMPI intercalation. Upon intercalation of DMPI cations, the interlayer spacing of graphite is found to increase from 3.34 Å to 7.2 Å, which facilitates trouble-free diffusion of DMPI cations through the graphite layers. Our calculated diffusion barriers support this fact, and we observe a quite small barrier of 0.2 eV for DMPI cation diffusion inside



graphite. Our calculated total energy values for different stages of DMPI cation intercalation follow the stability trend as: stage-2 < stage-4 < stage-1 during the initial charging process. However, the later stages of DMPI intercalation into graphite are expected to follow a stage-1 type intercalation mechanism as is evident by the decrease in the relative energy of stage-1 compared to the other stages with an increase in the concentration of DMPI cations. Our simulated XRD patterns for the stage-1 DMPI cation intercalation process and stage-4 AlCl_4 anion intercalation process match with the experimental XRD patterns of the electrodes during the charging process, thus validating the staging mechanism proposed by us. Our calculated average voltages for the early and later periods of intercalation (3.7 V and 4.6 V) are in good agreement with the experimental range (3.1–4.3 V). These results signify the importance of using graphite anodes for organic cation intercalation along with AlCl_4 anion intercalation in cathodes to achieve better electrochemical performance compared to DIBs using metal anodes along with AlCl_4 intercalated graphitic cathodes. The metallic character of graphite before and after DMPI cation intercalation obtained from the DOS plots confirms the constant electronic conductivity of the graphite anode. The CDD and quantitative Bader charge analysis suggest 0.87 $|e|$ charge transfer occurring from the DMPI cation imidazolium ring to the host graphite carbon atoms, thus indicating the electrochemical reduction of graphite during the charging process. Additionally, we have addressed the reason behind the self-discharge nature of DMPI- AlCl_4 graphite DIBs. Some of the shortcomings of dual-ion batteries are electrolyte decomposition and exfoliation of the electrode materials. There is minimal scope of electrolyte decomposition occurring at our calculated output voltage of 3.7 V as it is lower than the experimentally determined cut-off voltage (4.4 V). Exfoliation can result in the loss of crystallinity of the electrode. Exfoliation can occur due to (1) a lack of proper synchronization of volume expansion and contraction of the electrode upon the charging and discharging process, and (2) all ions not being fully deintercalated in the discharge process. As the volume expansion is limited for the DMPI intercalated graphite system, hence exfoliation can occur from the presence of DMPI cations inside the graphite layer in the discharge process due to the high binding energy between DMPI cations and graphite. However, experimental data showing the presence of DMPI cations inside graphite in the discharge state is not available and hence further experimental investigation is required. Overall, this article provides an in-depth understanding of the interaction between DMPI cations and graphite and helps in explaining the different staging mechanisms for both cations and anions to obtain precise electrochemical properties such as the voltage and storage capacity for DMPI- AlCl_4 graphite DIBs. We believe that these findings will motivate further exploration of various organic cations which can be suitably intercalated into anodes and coupled with AlCl_4 intercalated graphite cathodes to design high performance dual ion batteries on a par with lithium-ion battery technology.

Conflicts of interest

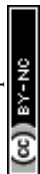
There are no conflicts to declare.

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References

- B. Dunn, H. Kamath and J. M. Tarascon, *Science*, 2011, **334**, 928–935.
- J. B. Goodenough and Y. Kim, *Chem. Mater.*, 2010, **22**, 587–603.
- J. B. Goodenough and K. S. Park, *J. Am. Chem. Soc.*, 2013, **135**, 1167–1176.
- B. L. Ellis, K. T. Lee and L. F. Nazar, *Chem. Mater.*, 2010, **22**, 691–714.
- L. P. Wang, L. Yu, M. Srinivasan, Z. J. Xu and X. Wang, *J. Mater. Chem. A*, 2015, **3**, 9353–9378.
- S.-W. Kim, D.-H. Seo, X. Ma, G. Ceder and K. Kang, *Adv. Energy Mater.*, 2012, **2**, 710–721.
- V. Palomares, P. Serras, I. Villaluenga, K. B. Hueso, J. Carretero-González and T. Rojo, *Energy Environ. Sci.*, 2012, **5**, 5884–5901.
- M. M. Huie, D. C. Bock, E. S. Takeuchi, A. C. Marschilok and K. J. Takeuchi, *Coord. Chem. Rev.*, 2015, **287**, 15–27.
- F. Liu, T. Wang, X. Liu and L. Fan, *Adv. Energy Mater.*, 2021, **11**, 2000787.
- C. Xu, B. Li, H. Du and F. Kang, *Angew. Chem., Int. Ed.*, 2012, **51**, 933–935.
- B. Lee, H. R. Lee, H. Kim, K. Y. Chung, B. W. Cho and S. H. Oh, *Chem. Commun.*, 2015, **51**, 9265–9268.
- M. C. Lin, M. Gong, B. Lu, Y. Wu, D. Y. Wang, M. Guan, M. Angell, C. Chen, J. Yang, B. J. Hwang and H. Dai, *Nature*, 2015, **520**, 325–328.
- P. Bhauriyal, G. Bhattacharyya, K. S. Rawat and B. Pathak, *J. Phys. Chem. C*, 2019, **123**, 3959–3967.
- J. M. Tarascon and M. Armand, *Nature*, 2001, **414**, 359–367.
- S. Hy, H. Liu, M. Zhang, D. Qian, B. J. Hwang and Y. S. Meng, *Energy Environ. Sci.*, 2016, **9**, 1931–1954.
- J. Wen, Y. Yu and C. Chen, *Mater. Express*, 2012, **2**, 197–212.
- F. P. McCullough, C. A. Levine and R. V. Snelgrove, *US Pat.*, 4865931A, 1989.
- T. Placke, O. Fromm, S. F. Lux, P. Bieker, S. Rothermel, H.-W. Meyer, S. Passerini and M. Winter, *J. Electrochem. Soc.*, 2012, **159**, A1755–A1765.
- I. A. Rodríguez-Pérez and X. Ji, *ACS Energy Lett.*, 2017, **2**, 1762–1770.
- J. A. Read, A. V. Cresce, M. H. Ervin and K. Xu, *Energy Environ. Sci.*, 2014, **7**, 617–620.
- P. F. McCullough, C. A. Levine and R. V. Snelgrove, *US Pat.*, 4830938A, 1989.
- X. Zhou, Q. Liu, C. Jiang, B. Ji, X. L. Ji, Y. Tang and H. M. Cheng, *Angew. Chem., Int. Ed.*, 2020, **59**, 3802–3832.



- 23 Y. Sui, C. Liu, R. C. Masse, Z. G. Neale, M. Atif, M. AlSalhi and G. Cao, *Energy Storage Mater.*, 2020, **25**, 1–32.
- 24 K. Beltrop, S. Beuker, A. Heckmann, M. Winter and T. Placke, *Energy Environ. Sci.*, 2017, **10**, 2090–2094.
- 25 Z. Lv, M. Han, J. Sun, L. Hou, H. Chen, Y. Li and M. C. Lin, *J. Power Sources*, 2019, **418**, 233–240.
- 26 T. Placke, A. Heckmann, R. Schmuck, P. Meister, K. Beltrop and M. Winter, *Joule*, 2018, **2**, 2528–2550.
- 27 K. V. Kravchyk, P. Bhauriyal, L. Piveteau, C. P. Guntlin, B. Pathak and M. V. Kovalenko, *Nat. Commun.*, 2018, **9**, 1–9.
- 28 W. Märkle, N. Tran, D. Goers, M. E. Spahr and P. Novák, *Carbon*, 2009, **47**, 2727–2732.
- 29 S. Das, P. Bhauriyal and B. Pathak, *J. Phys. Chem. C*, 2020, **124**, 7634–7643.
- 30 R. T. Carlin, H. C. De Long, J. Fuller and P. C. Trulove, *J. Electrochem. Soc.*, 1994, **141**, L73–L76.
- 31 M. Galiński, A. Lewandowski and I. Stepniak, *Electrochim. Acta*, 2006, **51**, 5567–5580.
- 32 E. W. Castner and J. F. Wishart, *J. Chem. Phys.*, 2010, **132**, 120901.
- 33 S. S. Manna, P. Bhauriyal and B. Pathak, *Mater. Adv.*, 2020, **1**, 1354–1363.
- 34 E. Knipping, C. Aucher, G. Guirado and L. Aubouy, *New J. Chem.*, 2018, **42**, 4693–4699.
- 35 S. Rothermel, P. Meister, G. Schmuelling, O. Fromm, H. W. Meyer, S. Nowak, M. Winter and T. Placke, *Energy Environ. Sci.*, 2014, **7**, 3412–3423.
- 36 F. Endres, S. Zein, E. Abedin, A. Y. Saad, E. M. Moustafa, N. Borissenko, W. E. Price, G. G. Wallace, D. R. Macfarlane, P. J. Newman and A. Bund, *Phys. Chem. Chem. Phys.*, 2008, **10**, 2189–2199.
- 37 A. Wang, W. Yuan, J. Fan and L. Li, *Energy Technol.*, 2018, **6**, 2172–2178.
- 38 Z. Li, J. Liu, J. Li, F. Kang and F. Gao, *Carbon*, 2018, **138**, 52–60.
- 39 J. Fan, Z. Zhang, Y. Liu, A. Wang, L. Li and W. Yuan, *Chem. Commun.*, 2017, **53**, 6891–6894.
- 40 S. Das, S. S. Manna and B. Pathak, *ACS Omega*, 2021, **6**, 1043–1053.
- 41 J. Vatsala Rani, V. Kanakaiah, T. Dadmal, M. Srinivasa Rao and S. Bhavanarushi, *J. Electrochem. Soc.*, 2013, **10**, 160.
- 42 N. Jayaprakash, S. K. Das and L. A. Archer, *Chem. Commun.*, 2011, **47**, 12610–12612.
- 43 G. Kresse and J. Furthmüller, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 44 Y. Chen, F. Peng, Y. Yan, Z. Wang, C. Sun and Y. Ma, *J. Phys. Chem. C*, 2013, **117**, 13879–13886.
- 45 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 46 J. J. Mortensen, L. B. Hansen and K. W. Jacobsen, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2004, **71**, 035109.
- 47 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 24103.
- 48 G. Henkelman, A. Arnaldsson and H. Jó Nsson, *Comput. Mater. Sci.*, 2006, **36**, 354–360.
- 49 R. F. W. Bader, *Chem. Rev.*, 1991, **91**, 893–928.
- 50 E. Sanville, S. D. Kenny, R. Smith and G. Henkelman, *J. Comput. Chem.*, 2007, **28**, 899–908.
- 51 D. J. Evans and B. L. Holian, *J. Chem. Phys.*, 1985, **83**, 4069–4074.
- 52 G. Henkelman, B. P. Uberuaga and H. Jónsson, *J. Chem. Phys.*, 2000, **113**, 9901–9904.
- 53 M. S. Dresselhaus and G. Dresselhaus, *Adv. Phys.*, 2002, **51**, 1–186.
- 54 P. Bhauriyal, A. Mahata and B. Pathak, *Phys. Chem. Chem. Phys.*, 2017, **19**, 7980–7989.
- 55 S. Kumar, P. Bhauriyal and B. Pathak, *J. Phys. Chem. C*, 2019, **123**, 23863–23871.
- 56 T. Placke, G. Schmuelling, R. Kloepsch, P. Meister, O. Fromm, P. Hilbig, H.-W. Meyer and M. Winter, *Z. Anorg. Allg. Chem.*, 2014, **640**, 1996–2006.
- 57 P. Bhauriyal, A. Mahata and B. Pathak, *J. Phys. Chem. C*, 2017, **121**, 9748–9756.
- 58 T. Ohzuku, Y. Iwakoshi and K. Sawai, *J. Electrochem. Soc.*, 1993, **140**, 2490–2498.
- 59 T. Sutto, T. Duncan and T. Wong, *Electrochim. Acta*, 2009, **54**, 5648–5655.
- 60 F. Weinhold and C. R. Landis, *Chem. Educ.: Res. Pract. Eur.*, 2001, **2**, 91–104.
- 61 S. Takahashi, N. Koura, S. Kohara, M.-L. Saboungi and L. A. Curtiss, *Plasmas Ions*, 1999, **2**, 91–105.
- 62 J. A. Read, *J. Phys. Chem. C*, 2015, **119**, 8438–8446.
- 63 U. Kasavajjula, C. Wang and A. J. Appleby, *J. Power Sources*, 2007, **163**, 1003–1039.
- 64 K. Persson, Y. Hinuma, Y. S. Meng, A. Van Der Ven and G. Ceder, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2010, **82**, 125416.
- 65 K. Nobuhara, H. Nakayama, M. Nose, S. Nakanishi and H. Iba, *J. Power Sources*, 2013, **243**, 585–587.

