


REVIEW

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The role of metal substitutions in the development of Li batteries, part I: cathodes

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Metal substitutions into known structures have served a pivotal role in developing two of the three main components in state-of-the-art Li-ion batteries: the cathode and electrolyte. In this first installment of this review, we will discuss the design principles established for cathodes, focusing on the benefits and limitations of substitutions in terms of the performance metrics of highest import for high-energy applications such as electric vehicles and grid storage. We emphasize the high number of metrics that must be simultaneously optimized and review the methods best suited to determine such metrics. In particular, partial metal substitutions have proven to be important in increasing energy density, but also in improving safety and cyclability. However, for many materials, cathode particle coatings have been found to be as effective as partial metal substitution in improving the lifetime of the battery since lifetime in state-of-the-art batteries is limited by reactions between the cathode particles and the liquid electrolyte, which can be addressed through surface coating. Though, the effects of substitutions and coatings can be additive to some degree. After a detailed overview of the main classes of cathodes used commercially and in development; we find a number of potentially interesting cathodes that operate at such high potentials that the use of liquid electrolytes becomes impractical that open the door to large leaps in Li-ion battery energy density. This motivates the development of solid electrolytes, the topic of part II of this review.

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1. Motivation

Metal substitution is a common technique used to tune and enhance the properties of many classes of solid-state materials. It has a long history in the form of metal alloys where one metal is mixed with another (*i.e.* some of the metal is substituted for another) to obtain more desirable physical and chemical properties than the initial metal alone.¹ A few examples are worth considering as an introduction to the various competing roles that substitutions can play in optimizing materials properties. Titanium alloys continue to be studied actively, with metal substitutions impacting physical properties such as Young's modulus, hardness, and thermal conductivity; all of which improve the suitability for a wide array of applications in aerospace and the automotive industry.² By contrast, on-going work in aluminum alloys focuses on improving corrosion resistance.³ While the physical properties are typically bulk properties requiring quite high levels of substitution to have a significant impact, the corrosion resistance is a surface chemical property that can be greatly improved with low-level substitutions. In particular, targeted substitutions at vulnerable locations on the material's surface can play a dramatic role.³

Herein, we will review the multiple roles that metal substitutions have played in improving Li-ion battery materials. Analogous to metallic alloys, the substitutions can either play a role in improving the bulk properties (such as energy density) or the surface dominated properties (such as electrochemical stability of the electrolytes). Multiple examples of which will be discussed throughout this article.

The above examples of substitutions to make alloys are typically in the % levels. There are also numerous cases in the literature where far lower substitutions into other materials can be very significant. The most important of which is, of course, doping into crystalline semiconductors. These substitutions are routinely used to tune a material's electronic properties, producing desirable band gaps for semiconducting materials,⁴ leading to improvements in light-emitting diodes⁵ and solar cells;^{6,7} as well as improving catalytic yields in metal-organic frameworks.^{8,9} In the case of battery materials, there are certain classes of electrode materials that are electronically insulating such that metal substitutions have been utilized to improve their electronic properties. It therefore becomes apparent that methods used in a wide variety of materials science and engineering prove to be of interest in improving battery materials. In this article, we will use the term doping exclusively for cases where the substitution is at a very low level and impacting the electronic properties.

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In this work, we focus on the role of metal substitution in the development of Li-ion battery materials. Specifically, in this first installment of this 2 part review, we evaluate the role metal substitutions have played in the development of Li-ion cathode materials. Li-ion batteries (LIBs) are fundamental to our modern lives. They power our smartphones, tablets, laptops, and, increasingly, our vehicles. Their prevalence is due to their high energy density compared to other battery chemistries, allowing them to store a large amount of energy in portable applications. While Li-ion battery technology has improved significantly over the past several decades, there is still room, and need, for improvement. Electric vehicles (EVs) still lag behind internal combustion engine (ICE) vehicles in terms of range.¹⁰ In order for EVs to reach the range of ICE vehicles and for Li-ion batteries to find new uses, such as in other modes of transport, we must continue to improve Li-ion battery energy density, power, and lifetime. Much of the focus on battery improvement lies on the cathode material. It is the limiting factor of LIB capacity, accommodating fewer Li^+ ions per unit weight than the anode, thus defining the maximum capacity of the battery. Presently, layered transition metal oxide cathodes have the largest market share of cathode materials. The most popular of which are $\text{LiNi}_x\text{Mn}_y\text{Co}_z\text{O}_2$ (NMC) and LiCoO_2 (LCO).¹¹ These materials have high energy densities in excess of 600 W h kg^{-1} ,¹² as well as a respectable lifetimes, but contain Co which is expensive and leaves an undesirable environmental impact on extraction and disposal.^{13,14} LCO was the first of the layered cathode materials to be commercialized,¹⁵ but by substituting Ni and Mn for Co, its performance was improved while reducing the cobalt content, leading to the superior NMC.^{16,17} In a similar vein, $\text{LiNi}_x\text{Co}_y\text{Al}_z\text{O}_2$ (NCA) was developed through metal substitution into layered transition metal oxide materials and also holds an appreciable market share.¹¹ These two examples of massively successful commercialization of layered oxides with metal substitutions has served as motivation for a great number of further studies looking at the impact of substitutions into many classes of cathodes for Li batteries. This idea of metal substitution to improve cathode structural and electrochemical properties has continued with small amounts of metals being substituted into these layered transition metal oxides to continue to improve their overall performance.

Outside of layered transition metal oxides, metal substitution has been applied to other cathode materials such as olivine phosphate cathode materials have an appreciable spot in the market, namely LiFePO_4 , which has a lower capacity than the layered transition metal oxides, but is safer and has a longer lifetime. Substitution of iron for other metals has led to the development of “high voltage” olivine cathodes such as LiCoPO_4 and LiNiPO_4 , which operate near 5 V vs. the 4 V of NMC and NCA, which could bring about a leap in LIB energy density. In addition to olivine cathode materials, spinel and disorder rock salts show promise in the high voltage space. Currently, these high voltage cathodes are restricted by electrolyte development, as traditional carbonate electrolytes break down when operating at these high potentials. Though, through

partial metal substitution, these high-voltage cathodes are becoming more compatible with electrolyte and through the continued improvement of both the cathode and electrolyte, high voltage, high energy Li-ion batteries may be realized.

Part I of this review will provide an overview of cathode improvements resulting from metal substitutions (both complete and partial), as well as provide an overview of the methods that should be used for evaluating these materials in a manner that provides accurate screening of battery performance. Specifically, partial metal substitutions have been used extensively to enhance the following properties in Li-ion cathodes:

- (1) Energy density (bulk property, % level substitutions required to have an impact)
- (2) Material stability (surface property, preferential substitutions at the surface of particles can be impactful; also bulk property, cation migration and phase transformations can be suppressed)
- (3) Rate performance (both bulk and surface property, often dominated by particle morphology).

Given the fundamental differences between these 3 properties, it is not surprising that successful substitution techniques differ for each key property, and the challenge is to optimize all 3 simultaneously. The methods used to reliably evaluate each of these properties also differ and will be reviewed in the next section, prior to a detailed discussion of the role of metal substitutions into improving each class of cathode in detail.

2. Cathode material metrics

There are a number of performance metrics that must be optimized simultaneously in designing advanced cathodes for Li-ion batteries. These include: (1) a high energy density, (2) a long cycle life, (3) a high energy conversion efficiency, (4) a good rate performance, (5) a high safety during battery operation, and (6) a low cost. These metrics are often shown in radar plots similar to that shown in Fig. 1. Thus, any study looking at improving one metric with the use of a substitution must also be mindful of any unintended detrimental impacts on the other metrics. This, in a nutshell, is what makes going beyond the state-of-the-art so difficult in Li-ion batteries. This section focusses on describing the methods used to extract the important metrics to assess battery performance, and in particular focusses on what conditions should and should not be used in order to ensure that conclusions hold up under real conditions in commercialized cells. Here, we assume a target application of electric vehicles as this is currently driving battery research and development worldwide. Electric vehicles being the objective has consequences for what cycling rates are appropriate, and it should be noted that for niche applications (e.g. power tools where faster rates are needed) different conditions (e.g. faster cycling) become appropriate. For the purpose of this review, we seek to understand how substitutions improve performance for conditions comparable to those of interest for electric vehicles, *i.e.* charging in 15–60 min,



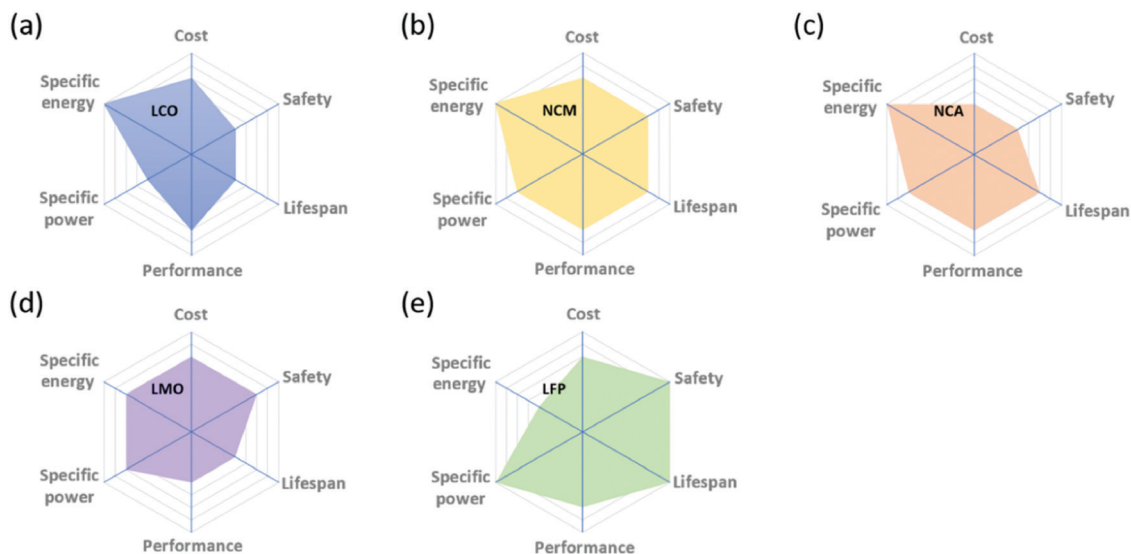


Fig. 1 Radar plots demonstrating the relative qualities of different classes of cathode materials: LiCoO_2 (LCO), $\text{Li}[\text{Ni},\text{Co},\text{Mn}]\text{O}_2$ (NCM), $\text{Li}[\text{Ni},\text{Co},\text{Al}]\text{O}_2$ (NCA), LiMn_2O_4 (LMO), and LiFePO_4 (LFP). Reproduced with permission from ref. 18. Copyright (2020 and 2019, respectively) Elsevier.

a slow discharge on the order of 10 h, and usage on the scale of 10 years.

2.1 Standard electrochemical testing

Electrochemical techniques are at the heart of cathode material evaluation. Typically, the electrochemical cycling is performed galvanostatically, *i.e.* using constant current. Though other methods are sometimes used such as constant current-constant voltage (CCCV), where the cell is first charged at a constant current until a threshold potential is reached, at which point that potential is maintained for a specified interval to simulate how batteries are charged in practice. For the sake of cathode evaluation, purely galvanostatic charging is the accepted technique in academia.

The rate at which cells are charged is often reported as a C-rate where a current of 1C is the current at which the cell will be charged to its expected capacity in one hour (and, *e.g.* C/5 would require 5 h). The expected capacity is determined by multiplying the specific capacity (mA h g^{-1}) by the mass of active material in the prepared electrode. This is trivial for well-characterized materials, as their specific capacities are known under a variety of conditions. However, for novel materials whose capacities are not known, C-rate may still be used, using an estimate of the specific capacity to determine the C-rate. It is of the utmost importance that the capacity used for the determination of the C-rate be reported, as the observed capacity of the cell is highly dependent on experimental conditions and cannot be assumed from data shown elsewhere in the article. Often, the capacity of a first charge cycle will be used as the expected capacity of a cell and the C-rate determined from this value. Measuring charging current in this unit compensates for variation in cathode capacity, making method standardization easier, as well as easing interpretation for the reader. Alternatively, current densities may be reported instead of C-rate. The current in mA g^{-1} is typically used to denote the

current per gram of active material and gives the same information as the C-rate. The two can be converted by dividing the current density by the specific capacity of the material. The current may also be reported in units of mA cm^{-2} . This, however, is the most difficult for the reader to digest and compare with other studies, as the cathode loading must be reported to allow the reader to calculate the C-rate. All electrochemical metrics will have a different value for different rates making it challenging to compare results from two different articles, more details follow on how this was handled in this review.

Although galvanostatic cycling is the typical protocol for battery performance determination, researchers also often use the more traditional technique of cyclic voltammetry. In cyclic voltammetry, an alternating linear sweep potential is applied to the cell and the current response is measured, indicating at which potential redox processes are occurring. It may be due to intercalation or side reactions. Different materials have signature CV curves, with peaks typically corresponding to TM redox or phase transitions (Fig. 2). CV experiments should be conducted over the same, or a slightly larger, voltage window as cycling tests. dQ/dV plots yield similar information to CV (CV can be converted to dQ/dV using the sweep rate) but are obtained from galvanostatic cycling. This has the inherent advantage of being able to obtain dQ/dV plots during traditional cycling experiments. Additionally, unlike CV, dQ/dV peaks do not deform due to changes in resistance of the cell. Instead, they shift laterally, making them easier to interpret (Fig. 2). It is important to note that while these methods can be used to evaluate cell degradation, they alone do not give direct insight into the mechanism behind cell degradation. With novel materials, it is best to use these methods in conjunction with other *in situ* and *ex situ* techniques to determine what CV and dQ/dV peaks correspond to physically.

In the next sections, we focus on how key battery performance metrics should be extracted from electrochemical



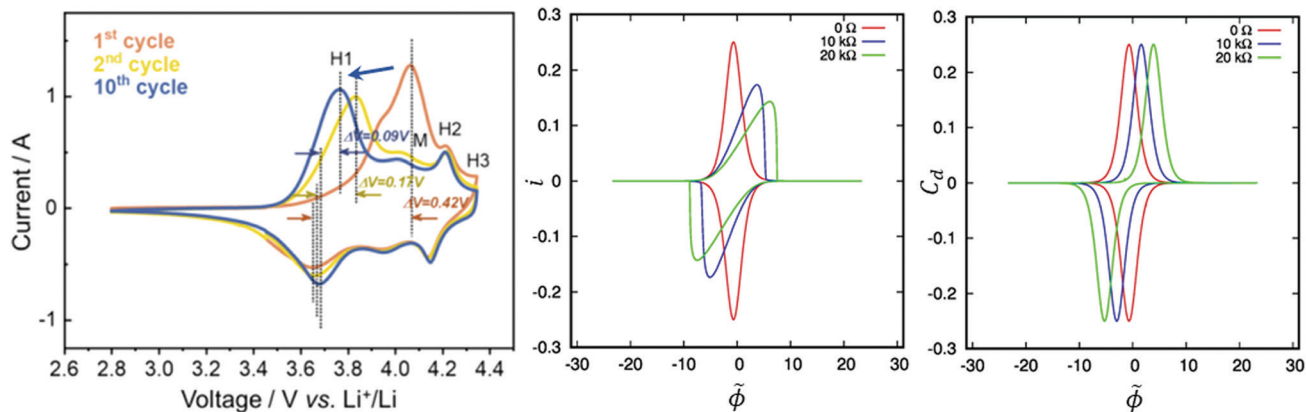


Fig. 2 CV during cycling of NMC 811 with H1 phase transition onset potential shift denoted with the blue error (left), influence of bulk resistance on CV (center) and dQ/dV (right). Reproduced with permission from ref. 19 and 20. Copyright (2020 and 2019) Elsevier.

testing and describe some of the pitfalls that can lead to results not holding up in batteries under typical conditions for commercialized products.

2.2 Extracting energy density

Specific capacity may be converted to another important battery metric, energy density, by multiplying the specific capacity of the material by the average voltage of the cell during discharge. The accepted method of representing the charge/discharge of a cell is using a charge/discharge curve, also called a voltage or potential curve as shown in Fig. 3. It is a plot of the potential against the capacity or the stoichiometry of the intercalant. The charge and discharge capacities may be interpreted from the length of the charge and discharge curves, respectively and the shape of the curves provide information about Li^+ intercalation in the material. These curves are sensitive to material composition,^{21,22} charging rate,²³ temperature,²⁴ and the state of health of the material.²⁵ Charge/discharge plots are useful in identifying how the charge and discharge of the material changes under different charging conditions and over repeated charges (Fig. 3). An important point to note is that in the case of a two-electrode system, which is typical for most studies, the curve is affected by both the cathode and anode of the cell. It is common to use a Li metal anode to maintain excess Li and

sufficient capacity to not limit the capacity over many cycles if Li is consumed. However, a solid electrolyte interphase composed of electrolyte decomposition products forms on the anode affecting the kinetics of charge transfer which makes voltage measurements unreliable at current densities above 1 mA cm^{-2} .²⁶ A three-electrode cell may be used to obtain the curve uniquely representative of the cathode, typically using a Li pseudo-reference electrode.²⁷ Three-electrode techniques are not the norm in galvanostatic experiments evaluating cathode materials and will not be further discussed in this review.

2.3 Extracting lifetime with long-term cycling

The lifetime of a cell can be evaluated through repeated charging and discharging of the cell and observing the capacity degradation and voltage decay. Specifically, the discharge capacity, or reversible capacity, is of interest as that is the origin of the accessible energy of the cell. Thus, the charge capacity is often omitted when presenting long-term cycling data (Fig. 3). A cycle life of a cell does not have a universal definition. Commonly, it is considered to be when the cell capacity declines to 80% of the original reversible capacity.³⁰ The phenomenon of the capacity decreasing over multiple cycles is termed “capacity fade”. The lifetime of a cell may also be characterized as the point where capacity fade accelerates sharply before reaching 80%

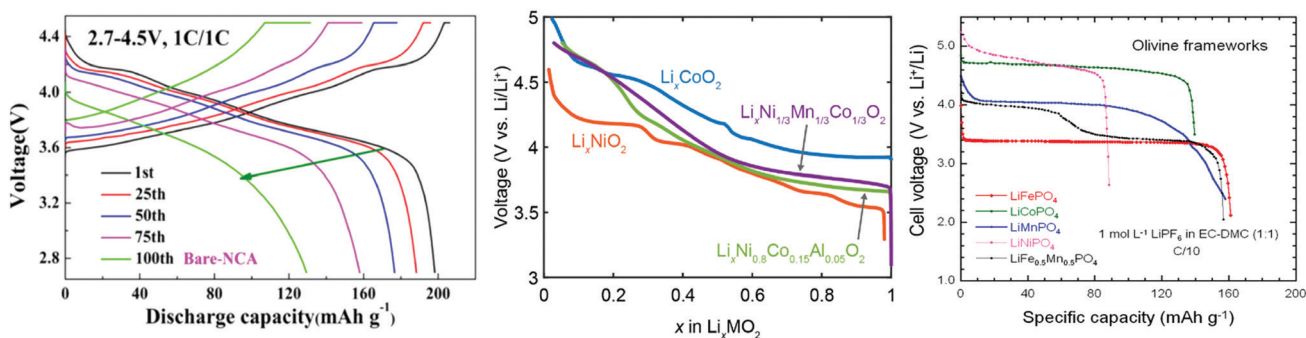


Fig. 3 Charge/discharge curves of $\text{LiNi}_{0.85}\text{Co}_{0.10}\text{Al}_{0.05}\text{O}_2$ (left). Discharge curves of various layered oxides with respect to Li composition instead of capacity (center) and for olivine cathode materials. (right) Figures reproduced with permission from ref. 21, 28 and 29. Copyright (2017) Wiley, (2019) American Chemical Society, (2017) Multidisciplinary Digital Publishing Institute.

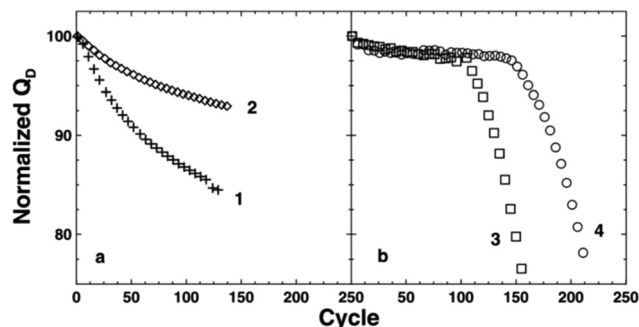


Fig. 4 Charge vs. cycle number plots showing gradual capacity loss (left) and rollover failure (right). Reproduced with permission from ref. 31. Copyright (2013) IOP Publishing.³¹

capacity, referred to as a “knee” in the capacity vs. cycle number plot, or “rollover failure” (Fig. 4). Dahn showed that rollover arises from poor Li diffusion kinetics arising from the formation of an impeding layer on the negative electrode, comprised of electrolyte degradation products.³¹

The rate used to evaluate cycle life varies and this can have dramatic impact on the results obtained. C/3, C/2, or 1C are popular rates, though higher rates may be used to demonstrate high-rate capabilities of cathode materials, and lower rates may also be used, especially to compare materials that show very poor longevity. It is very important to recognize that the lifetime performance of Li-ion batteries is not only a function of the

number of cycles but also of time. Thus, a cell that maintains a high capacity over many cycles at 1C will typically lose much more capacity at C/10 due to the fact that there is more time for negative side reactions to occur, depending on the voltage to which it is charged.³² This approach to inflate the lifetime metrics of a material is unfortunately commonplace in studies examining nanostructured materials where high rates are achievable. One must consider the total measurement time and then extrapolate to reasonable battery lifetimes in order to obtain useful metrics. That said, at higher rates, structural degradation may be so severe that capacity will degrade much faster at 2C and above compared to 1C, depending on the rate capability of the cathode material.

Regardless of the chosen rate for long-term cycling, the first one to three cycles are typically performed at a low rate of C/10 or C/20. This is to ensure proper formation of the solid electrolyte interphase (SEI), which serves to protect the electrodes from reactions with the electrolyte that negatively affect performance. The appropriate potential window in which the cell is cycled may vary depending on the cathode material being studied.

Given all the different experimental parameters affecting the lifetime metric for the same material, in the context of this article, we must compare the substituted and unsubstituted from the same article in order to get any usable information. A good example of this is shown in Fig. 5a, c, and d where it is clear that lifetime metrics depend strongly on the cycling rate;

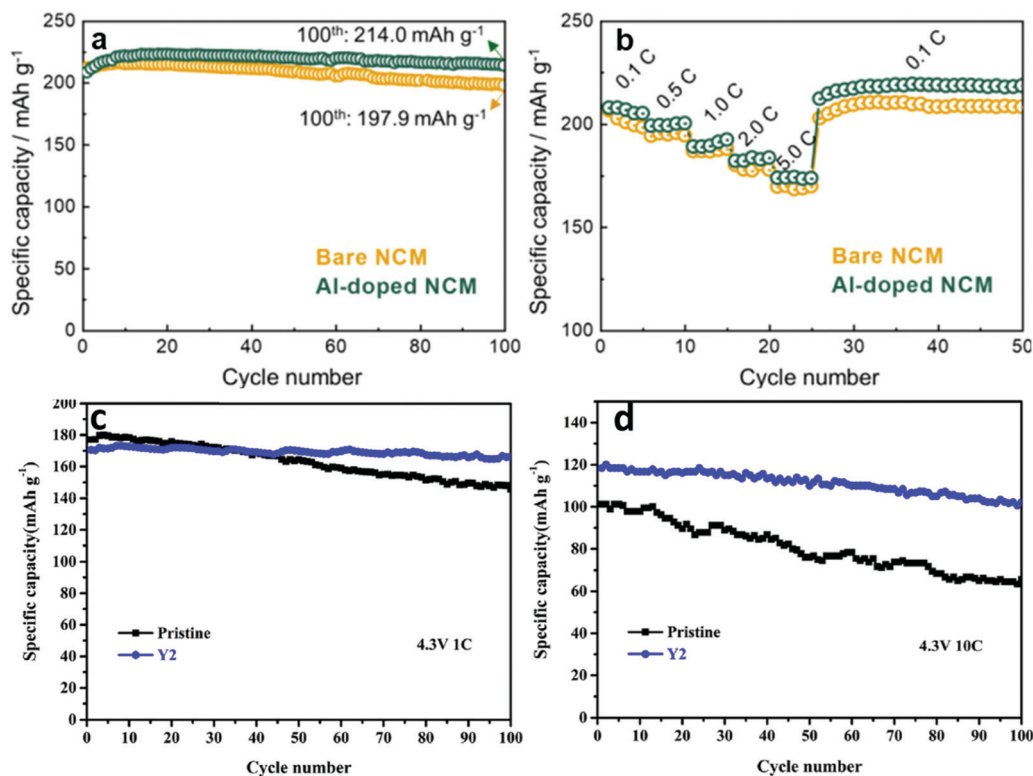


Fig. 5 (a) Long term cycling of NMC and Al-substituted NMC at a rate of 0.1C (a). (b) Rate performance and recovery of NMC and Al-substituted NMC.¹⁹ (c) Long term cycling performance of NMC and Y-substituted NMC at 1C (c) and 10C (d). Figures reproduced with permission from ref. 19 and 33. Copyright (2020 and 2019) Elsevier.



and so we can only determine the impact of a substitution by comparing to the unsubstituted from the same article.

2.4 Extracting lifetime with coulombic efficiency

Another important metric during the charge/discharge of the cell that can, in many cases, be an indication of lifetime is the ratio of the discharge capacity and the previous charge capacity. This is referred to as the coulombic efficiency (CE) and represents the charge efficiency of the cell: how much of the lithium put into the anode during charge is given back during discharge. When a battery is charged, not all the lithium goes into intercalation into the electrodes, there are many side reactions that can occur, driven by the electric potential put across the cell, including the formation of the SEI. This lithium is often lost for subsequent cycling such that it leads to capacity fade and eventual cell death. Thus, CEs are typically less than 1 (1 would represent a battery that lasts forever). The CE typically begins smaller and increases approaching 1, as the initial formation of the SEI causes a large discrepancy, which decreases as the SEI matures.³⁴ Typical values of the CE are >99% after initial SEI formation and state-of-the-art Li-ion batteries show CEs >99.9% and these are the batteries that can last up to 10 years in commercial settings. Thus, it is completely uninformative to present the CE on a scale of 0–100%, but nonetheless this practice is common and leaves the reader unable to evaluate the CE properly. It is also important that if CE is to be taken as a useful metric of battery health, it must be measured with a sufficient accuracy ($\sim \pm 0.01\%$) to have any predictive power. In such cases, CE values have correlated extremely well with long-term cycling^{31,35–37} and have been used predictively to accelerate the design of advanced batteries. However, when CE values are reported that do not meet the above standards, there is no correlation to lifetime and so no CE values will be discussed in this review article unless they meet the standard of usefulness (0.01% precision).

2.5 Rate performance

The rate performance of a cell is tested similarly to lifetime, but with a varying charge rate. The cell is first cycled multiple times at a low rate (C/10 or C/20) to ensure good SEI formation and to serve as a capacity baseline. Then, the rate is increased incrementally with several cycles occurring at each rate (Fig. 4b). Depending on the material, the rate may be increased as high as 10C or 20C, *i.e.* a full charge in 6 or 3 minutes. After the highest rate, the cell is then cycled several times at the first rate and the capacity compared to the first cycles to determine if performance was affected by the high-rate exposures. The current target application of many battery researchers are electric vehicles and the key target for rates is to be able to charge in 15 minutes (4C charging). As such, rate tests provide an important metric for advanced batteries. For example, in Fig. 5b, we see 33% capacity loss when cycled at 5C compared to C/10. However, it should also be mentioned that in state-of-the-art batteries, rate performance is limited by the anode³⁸ and as such this is not so essential in the current review article dealing with cathodes. The exception is where substitutions

have such a detrimental impact on rate performance such that the cathode in fact becomes limiting.

2.6 Voltage fade

As mentioned previously, voltage fade during long-term cycling also decreases the energy density of a battery. It is important to note that many battery management systems rely on voltage as an indicator of state-of-charge,³⁹ and this relation cannot be used if the voltage of the battery gradually decreases over time and this makes battery management far more challenging. Battery health monitoring and prediction is an active area of development and is steadily improving making voltage fade less of an issue with regard to battery state of health estimation.⁴⁰ Nonetheless, it remains an important metric to track given that cathode materials experience structural transformations as they age and, in some materials, the accumulated changes cause the average voltage during the discharge of the cell to decrease significantly. This causes a reduction in the available energy of the cell. While not a large problem with current market-leading battery chemistries, it is a large factor hindering next-generation materials from being realized. It can be evaluated simultaneously with capacity fade by taking the average voltage of each discharge cycle and monitoring the decay. The mechanisms behind both capacity and voltage fade are discussed in more detail within the material sections.

2.7 Safety

The most important aspect of safety for Li-ion batteries is preventing thermal runaway.⁴¹ Thermal runaway occurs when the battery reaches a temperature where exothermic reactions occur at such rates that the resulting heating of the battery is sufficient to accelerate the rate of these exothermic reactions to such an extent that the electrolyte ignites causing the cell to rupture violently.⁴² The cathode material plays a role in the self-heating of the battery as well as the onset of thermal runaway by reacting with the electrolyte. The most common techniques of evaluating the onset temperatures of exothermic reactions are differential scanning calorimetry (DSC) and accelerating rate calorimetry (ARC).^{43–46} DSC consists of measuring the difference in energy it takes to heat a reference and the sample as a function of temperature.⁴⁷ Accelerating rate calorimeters confine the sample to an adiabatic environment where energy loss is minimal and measures the rate at which the temperature increases as a function of temperature to monitor self-heating.⁴⁸ ARC is a more precise technique than DSC to determine the onset temperature for exothermic degradation of battery components leading to thermal runaway.^{48,49} Either technique must be performed on the fully charged cathode (this is the most reactive state of the cathode) in the presence of electrolyte in order to correlate to thermal runaway in a full battery. Fig. 6 shows an example of ARC and DSC on various cathode materials.

3. Advanced characterization to determine mechanisms

The previous section outlined the techniques used to screen for important battery metrics. These are essential in evaluating the



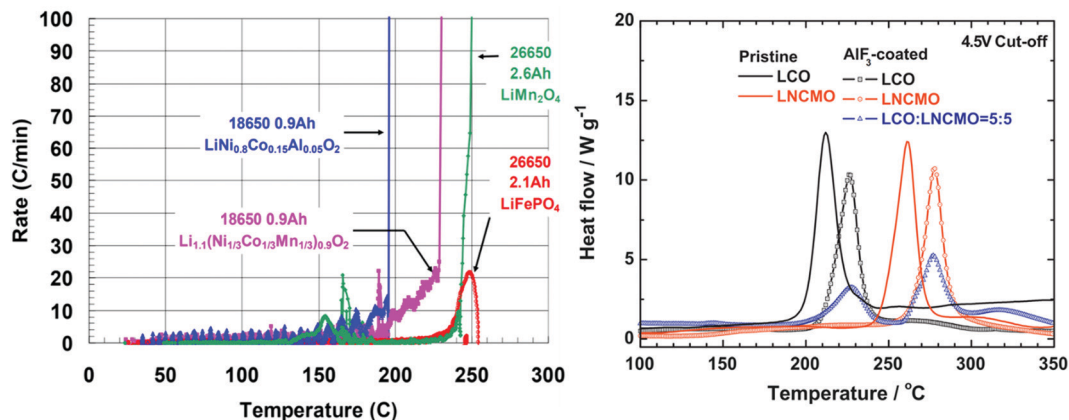


Fig. 6 ARC of various batteries using different cathode materials in EC:PC:DMC electrolyte and graphite anodes (left) and DSC of various cathode materials (right). Figures reproduced with the permission of ref. 50 and 51. Copyright (2012) IOP Publishing and (2011) Elsevier.

usefulness of substitutions into the cathodes. However, a great variety of techniques are also used extensively in order to further understand the mechanisms involved in either improving or deteriorating battery performance with substitutions. A few of the key techniques that will be discussed throughout this review are described in this section.

3.1 Electrochemical impedance spectroscopy

Aging of a cell is typically accompanied by impedance growth due to a thickening SEI and material degradation, specifically at the surface.⁵² An increase in cell impedance can negatively impact the performance of the cell, lowering efficiency and reducing rate capability, and may be indicative of reduced safety due to increased risk of thermal runaway.⁵³ The method of choice to evaluate cell impedance is electrochemical impedance spectroscopy (EIS). EIS consists of perturbing the battery with a low amplitude alternating current pulse and monitoring the voltage response. This is repeated over a wide range of frequencies (Hz–MHz) and the impedance at each perturbation is determined.⁵⁴ The data are typically represented using a Nyquist plot wherein the impedance is plotted parametrically in the plane consisting of the imaginary impedance on the y-axis and the real impedance on the x-axis. For Li-ion batteries, the plot typically takes the shape of successive semicircles. The utility of EIS relies on modelling the obtained response with a simple circuit. With a circuit model, a resistor and capacitor in parallel respond with a semicircle on the Nyquist plot and placing a resistor in series with these parallel components results in a shift in the semi circle's position on the real axis. A typical response (Fig. 7), can be fitted with multiple of these parallel resistors and capacitors in series. The values of the fitted circuit can be assigned to different components of the battery. This is where care needs to be taken, as the model is not prescriptive: it does not assign meaning to the obtained values, that is up to the researcher. The accepted interpretation of the model is that there is a “bulk” impedance between electrodes where the electrolyte and separator are, R_0 in Fig. 7, as well as a interfacial impedances such as between cathode material and electrolyte which is caused by the surface

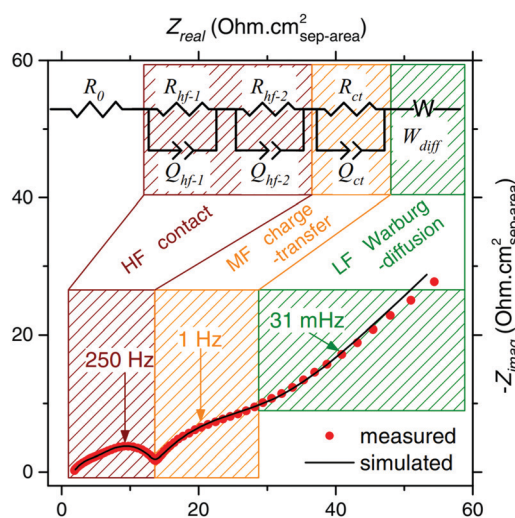


Fig. 7 Example Nyquist plot of a LIB with interpretation. Contact resistance (red), charge-transfer resistance (yellow), Warburg diffusion (green). Figure reproduced with the permission of ref. 55. Copyright (2016) IOP Publishing.

structure of the cathode and the SEI. This is typically the area of interest when EIS is used, as it is desirable for a cell to keep a consistent impedance as it ages and a large increase in impedance represents a poor state of health for the battery.

3.2 Scanning electron microscopy

The performance of a cathode material is closely tied to its structure. Therefore, it is critical to understand and evaluate changes to the structure during and after cycling to identify changes that occur that affect the performance of the cell. These changes can occur on multiple length scales, from the size of cathode particles (typically tens of microns) to the displacement of individual atoms within the crystal lattice. There are many microscopy, diffractometry, and absorbance techniques to characterize the structure of cathode materials across these scales; the most popular of which will be discussed here, starting with scanning electron microscopy (SEM).

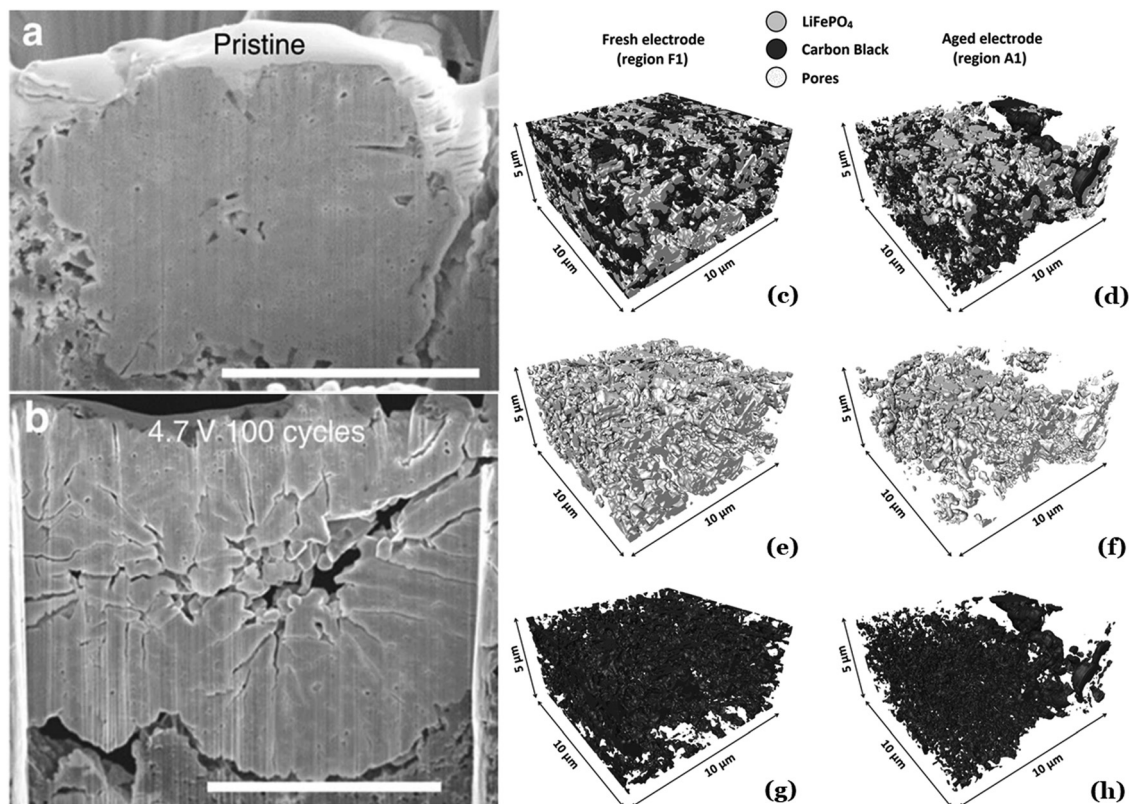


Fig. 8 Example of cracks formed in NMC cathode material after cycling. 3D reconstruction of fresh and aged LFP electrode. Figures reproduced with permission from ref. 61 and 68. Copyright (2016) Elsevier and (2017) Nature.

SEM is useful for characterizing both as-synthesized cathodes and evaluating particle degradation after cycling. Particle morphology directly affects the performance of cathode materials. The size and shape of particles can significantly impact both energy density and rate performance. SEM can be used to qualitatively evaluate particle morphology and validate proper synthesis. It is possible to use SEM images to quantitatively determine particle size with the aid of image analysis software, but it is standard to use laser particle size distribution instruments to obtain particle size distributions instead. SEM is particularly well-suited to comparing pristine and cycled material to qualitatively compare the degree of particle degradation through cracking (Fig. 8a and b).^{56–59} This is of particular use in understanding the long-term cycling performance, particularly when surface coatings or core-shell structures are used in an attempt to prevent particle fracture. Using focused ion beam SEM tomography, 3D reconstructions of the cathode can be made before and after cycling to obtain a more complete image of the structural evolution of the material (Fig. 8c–h).^{60–63} SEM is also commonly paired with energy-dispersive spectroscopy (EDS) to map the distribution of consistent metals across the particle. This can be used to verify the uniform distribution of a coating or can validate the synthesis of uniformly substituted materials or core-shell and gradient materials by imaging the cross-section of the particle obtained *via* focused ion beam.^{64–67}

3.3 X-ray and neutron diffraction

To determine the bulk crystal structure of a cathode material, X-ray diffraction spectroscopy (XRD) is the most popular method. XRD spectra can be refined using Rietveld refinement to obtain structural parameters, crystallite size, as well as the amount of impurity phases.^{69–75} XRD spectra are also used to evaluate certain structural defects such as Li/transition metal mixing. XRD can also be used *in situ* to investigate changes in cathode structure during cycling to identify when structural transformations occur as illustrated in Fig. 9.^{76–81} This can be accomplished using cell materials with an X-ray transparent window. This allows the determination of structural changes during battery operation and is essential in identifying the mechanisms taking place during cathode use. Substitutions have been shown to suppress certain transformations during cycling^{82–84} (e.g. Mg substitution into NMC 811 in Fig. 9).⁸⁵

Like XRD, neutron diffraction can be used to investigate the structure of cathode materials. It operates on a similar principle to XRD, only using neutrons instead of X-rays. The main advantage of neutron diffraction is that it is more sensitive to lighter elements (Li and O is of particular interest in the present context) and can distinguish between elements with similar atomic numbers.⁸⁶ It has been used to study the distribution of transition metals in transition metal oxides. For example, in $\text{LiNi}_{1/3}\text{Mn}_{1/3}\text{Co}_{1/3}\text{O}_2$, neutron diffraction was used to determine that there is non-random ordering of Ni,



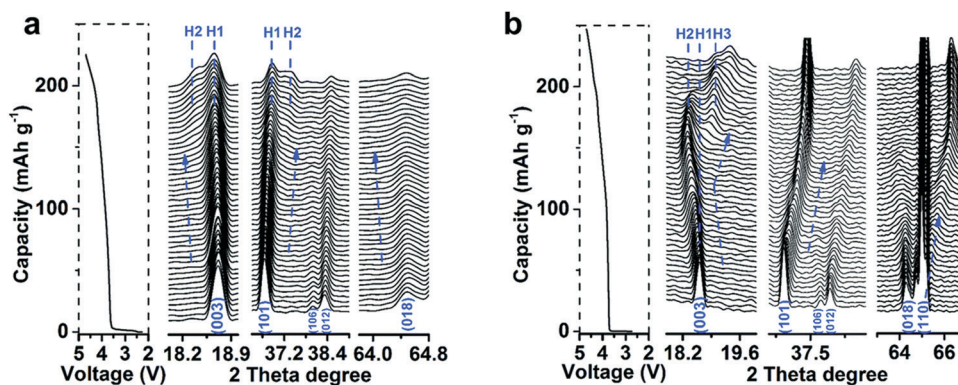


Fig. 9 *In situ* XRD of Mg-substituted (a) and pristine (b) NMC 811 showing suppression of transition to H3 shown via the (003) peak shift and splitting as well as the reduced shifting of the 101 peak. Figure reproduced with the permission of ref. 85. Copyright (2019) Royal Society of Chemistry.

Mn, and Co within transition metal layers.⁸⁷ It has also been used to determine the degree of Li/Ni mixing through Rietveld refinement (Fig. 10).⁸⁸ Recently, *in situ* neutron powder diffraction has also been performed on materials in the pouch cell format: Liu investigated irreversible structural changes in $\text{Li}[\text{Li}_{0.2}\text{Ni}_{0.18}\text{Mn}_{0.53}\text{Co}_{0.1}]\text{O}_2$ ⁸⁹ and Goonetilleke used it to compare the structural stability of various NMC cathode compositions, finding NMC 442 to be more stable than NMC 111 and NMC 811, with the smallest change in the *c* lattice parameter when charged.⁹⁰

3.4 Transmission electron microscopy

To investigate local crystalline structures, transmission electron microscopy (TEM) may be used. It is particularly useful for evaluating structural changes at the surface of cathode particles as well as changes occurring on the atomic scale. For instance, one can observe the transition from layered to spinel to

disordered rock salt occurring at the surface of NCA and Ni-rich NMC particles as they are cycled (Fig. 11).^{91–94} The technique has also been used to evaluate SEI growth.⁹⁵ TEM in conjunction with energy-dispersive X-ray spectroscopy (EDX) or electron energy loss spectroscopy (EELS) is also used to evaluate changes in metal compositions of cathode materials at the surface.^{96,97} TEM is standard in determining the thickness and structure of surface coatings^{98–102} and can be used to observe stacking faults in the material or other crystal defects.^{103–106}

An example of TEM being used to explore nanoscale changes in the cathodes is the formation of O–O dimers in Li-rich Ir oxides.¹⁰⁷ Although this technique can be used to observe the atomic scale, it should be noted that the materials are very sensitive to beam damage and that great care must be taken during TEM, including using low energies and aberration correction.¹⁰⁸

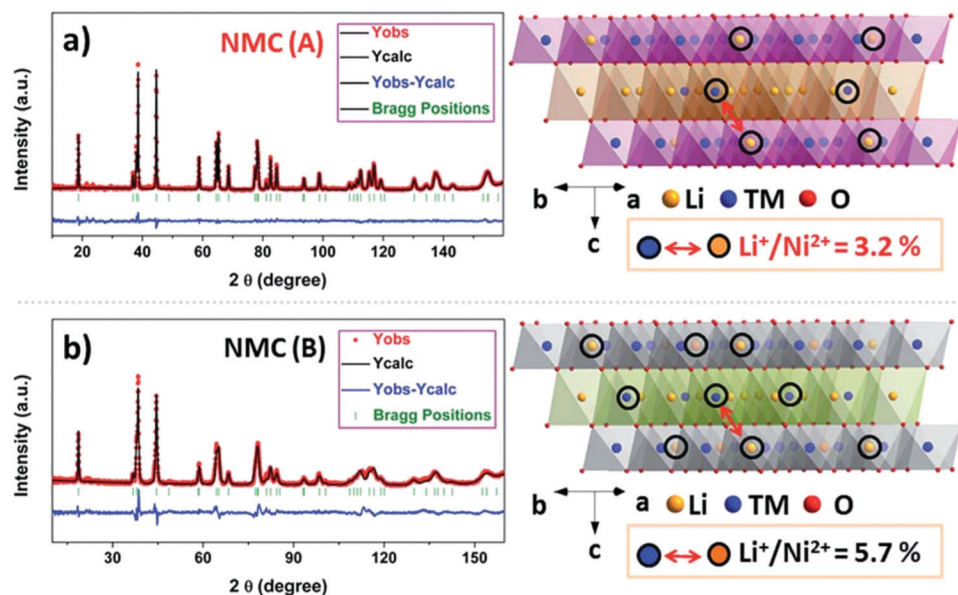


Fig. 10 Neutron diffraction refinement of NMC to determine extent of Li/Ni mixing. Reproduced with the permission of ref. 88. Copyright (2017) Royal Society of Chemistry.



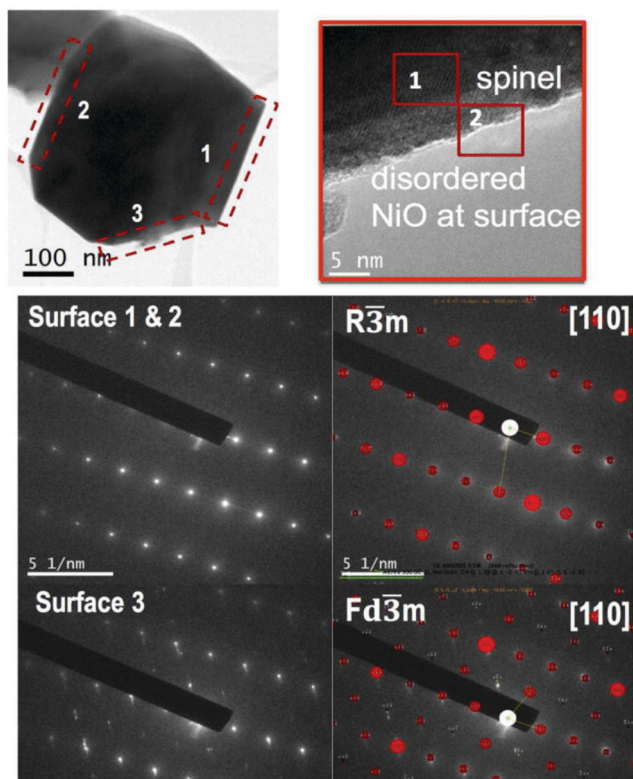


Fig. 11 TEM images demonstrating decomposition of NCA particle surface. Reproduced with permission from ref. 91. Copyright (2016) AIP Publishing.⁹¹

3.5 Spectroscopies

While the above techniques are important for understanding physical structure and structural changes during the cycling of cathode materials, they tell little of the changes in oxidation state of metals during charge/discharge. X-ray absorbance techniques can be used to identify the electrochemical activity of constituent metals in the electrode. In particular, X-ray absorption near-edge structure (XANES) can be used *in situ* to determine the average oxidation state of metals in the cathode material during charge and discharge, elucidating which

metals are oxidized at different states of charge and how it may change over multiple cycles.^{109–111} This can be achieved by fitting the edge of the spectrum. For example, the oxidation state of Ni has been investigated during the cycling of $\text{Li}_4\text{NiTeO}_6$ cathode materials as demonstrated in Fig. 12.¹¹²

While XANES can be used to study the bulk oxidation state of metals in a material, X-ray photoelectron spectroscopy (XPS) may be used to investigate the oxidation state of metals at the surface of the cathode material.^{113–115} For example, in the same study by Ting,¹¹² XPS was used to determine the oxidation state of Ni at the surface of the model cathode material, thus allowing for a comparison of the behaviour of bulk and surface metals, a powerful combination of techniques to probe changes in the surface of cathode materials. XPS may also be used to investigate the chemistry of the SEI^{116–118} and explore the reactivity and mechanism of action of various surface coatings.^{119–121}

4. Impact of substitutions into LiMO_2 compositions

Having reviewed the key methodologies used in studies looking at the impact of substitutions into Li cathodes, the rest of this review will examine in detail the impact of substitutions into particular classes of cathodes. We begin with this section looking at layered oxides.

4.1 Unsubstituted LiMO_2

The predominant class of Li-ion cathodes for high energy batteries is layered transition metal oxides. The history of these cathodes goes back to 1980 with the discovery of LiCoO_2 (LCO) as a cathode material by Goodenough¹²² leading to the creation of the first modern lithium-ion battery by Yoshino in the 1980s,¹¹ consisting of a LiCoO_2 (LCO) cathode and carbonaceous anode.¹²³ This cell was commercialized by Sony and was the first in a plethora of LIB chemistries to come in the past four decades. To this day, the most-used cathode materials are isostructural with LCO.

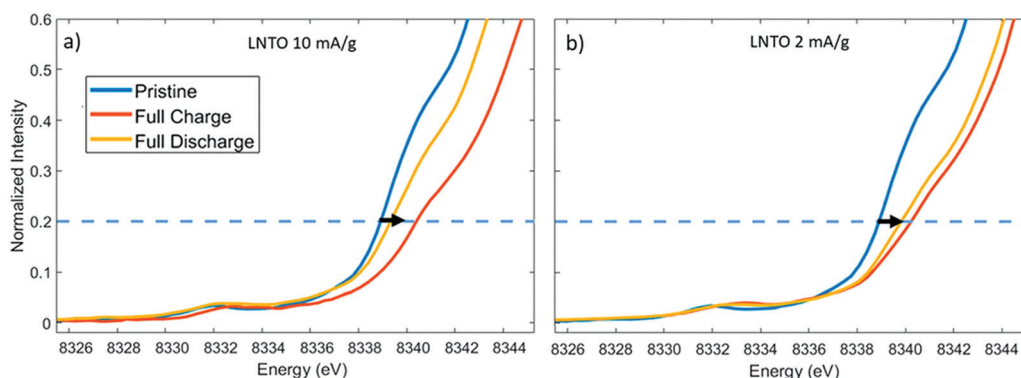


Fig. 12 Relevant portion of the Ni K-edge for determining oxidation state from XANES spectra of $\text{Li}_4\text{NiTeO}_6$.¹¹² The edge position can be used, along with references, to calculate the average oxidation state of Ni. Figure reproduced with permission from ref. 112. Copyright (2020) American Chemical Society.



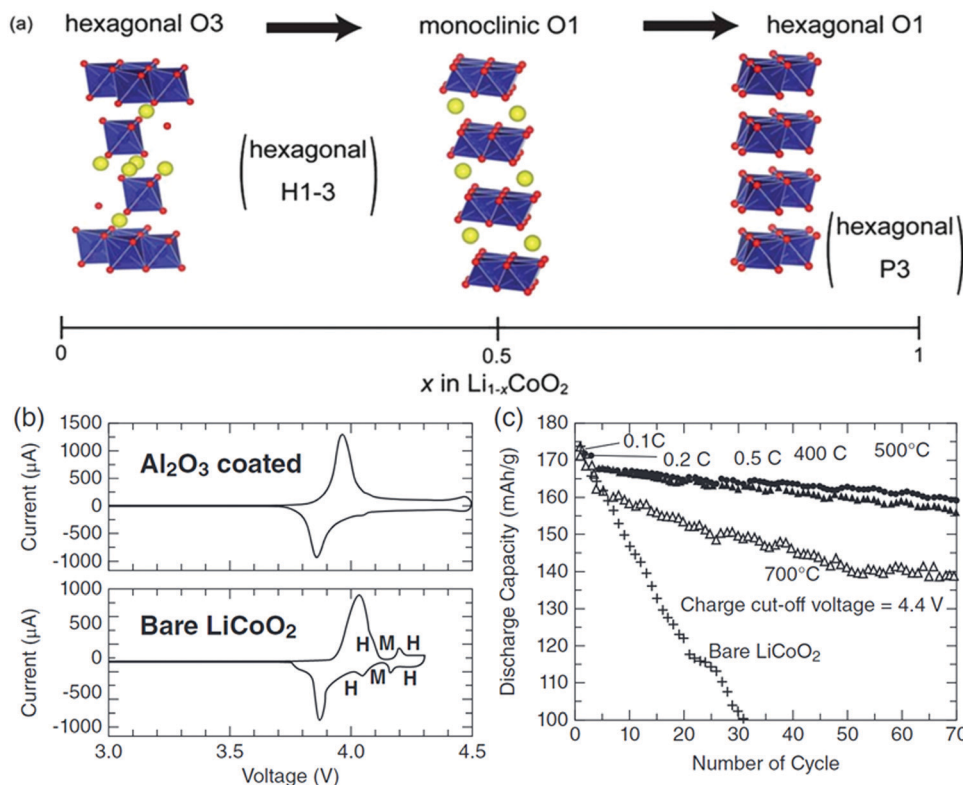


Fig. 13 O3 and monoclinic and hexagonal O1 structures of LCO (a). Cyclic voltammograms (b) of bare and coated LCO. The phase transformations yielding the O1 structure (labelled H \rightarrow M \rightarrow H) are suppressed by the coating, yielding better long-term cycling (c). Figures reproduced with permission from ref. 126 and 127. Copyright (2012) Royal Society of Chemistry and copyright (2017) Wiley.^{126,127}

LCO consists of planes of Li atoms sandwiched between extensive slabs of CoO_2 , forming a rocksalt structure that can be described in the $R\bar{3}m$ space group. The structure is shown in Fig. 13 and is called the O3 structure in the battery literature as Li takes octahedral sites and 3 CoO_2 slabs must be stacked in the vertical direction before repetition. The cobalt atoms occupy octahedral sites within CoO_2 slabs and Li are found in octahedral sites between the slabs. This material reversibly intercalates Li through these 2D sheets. The removal of lithium during charge is, in principle, compensated by Co^{3+} converting to Co^{4+} during charge. The theoretical capacity of LCO assuming complete removal of lithium is 273 mA h g^{-1} .¹²⁴ However, in practice the capacity is substantially less ($\sim 140 \text{ mA h g}^{-1}$), as only half of the Li can be accessed before the structure converts from the O3 stacking of the pristine material to the O1 stacking of CoO_2 (see Fig. 13a and 15c).¹²⁵ The shearing required to transform from the O3 to the O1 structure results in very poor long-term cycling as shown in Fig. 13c. This rapid fade has been mitigated dramatically by either substitution (to be discussed below) or by coatings such as Al_2O_3 (Fig. 13b and c). The battery literature contains many such cases where either coatings or substitutions may improve certain battery metrics, in such cases it is important to look at all metrics to see which solution is optimal.

There are a number of other materials that can be made with this same structure by completely replacing Co with another

transition metal. Two materials, in particular, were widely studied to replace LCO: LiNiO_2 (LNO) and LiMnO_2 , both with theoretical capacities very close to LCO.¹²⁸ LNO can be synthesized quite readily, though it is difficult to synthesize with perfect stoichiometry, as some Ni^{3+} is reduced to Ni^{2+} and mixes into the lithium layer due to similar atomic radii of Ni^{2+} and Li^+ , resulting in reduced capacity and structural stability.^{129–132} The material shows good cycling on the first cycle (Fig. 15d). However, it suffers from structural instability during electrochemical cycling, notably due to the M \rightarrow H2 transition (Fig. 15f). Nickel can migrate to vacant Li^+ sites during charging and this migration occurs to such an extent that the material converts to a disordered rocksalt structure and this results in the blocking of Li^+ intercalation pathways. This leads to an increased impedance as well as active material loss and a dramatically shortened cycle life. Overall, this leads to a cathode material with a high initial capacity, but very poor cycle life as shown in Fig. 15e. Numerous substitutions into LNO have been attempted to prevent these transformations as discussed in the next section.

Another metal that has been extensively studied as a replacement for Co is Mn. The use of Mn is particularly attractive due to the fact that it is less costly than Ni or Co and is more environmentally benign. LiMnO_2 has, in fact, been prepared in an orthorhombic layered structure (Fig. 14). Although it has a high theoretical capacity of 285 mA h g^{-1} , it is easily converted

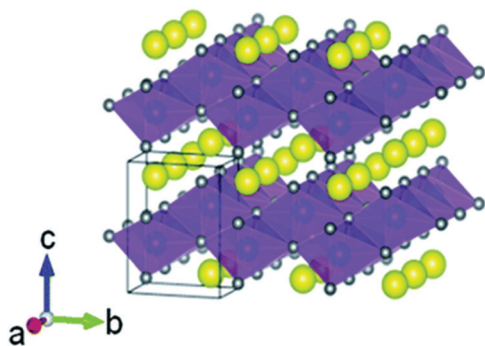


Fig. 14 Layered structure of LiMnO_2 . Figure reproduced with permission from ref. 133. Copyright (2018) Royal Society of Chemistry.

to a spinel-like LiMn_2O_4 phase upon Li^+ deintercalation and has proven difficult to synthesize, often containing Li_2MnO_3 and LiMn_2O_4 phases which increase the rate of capacity fading. Although the conversion to LiMn_2O_4 actually increases the capacity of the material (Fig. 15h), the average voltage is so

low that it is not competitive. Additionally, LiMnO_2 is susceptible to Mn dissolution in the electrolyte, degrading the structure at the surface and shortening the cycle life. Even with these deficiencies, it has reached practical capacities of $\sim 160 \text{ mA h g}^{-1}$ and has seen some commercial use in primary batteries (non-rechargeable), though its lower energy density and poor lifetime compared to other layered oxides limit its application. Fig. 15g–i show the electrochemistry of LMO. The differential capacity plot in Fig. 15i illustrates the structural transformation from the first few cycles onward with the evolution of the peaks at 3.7–4.2 V.

4.2 NMC: Ni, Mn co-substitution into LiCoO_2

With the above limitations of LCO, LNO, and LMO, it is not surprising that significant efforts have been invested in looking for solid solutions with 2 or more of Ni, Mn and Co in the LCO structure. In the early 1990s, Delmas¹⁴¹ suggested a solid solution of $\text{LiNi}_{1-y}\text{Co}_y\text{O}_2$ which allowed up to 0.6 Li to be reversibly deintercalated, increased from the 0.5 of LCO, and with optimization could reach 0.75 reversibly.¹⁴² The material

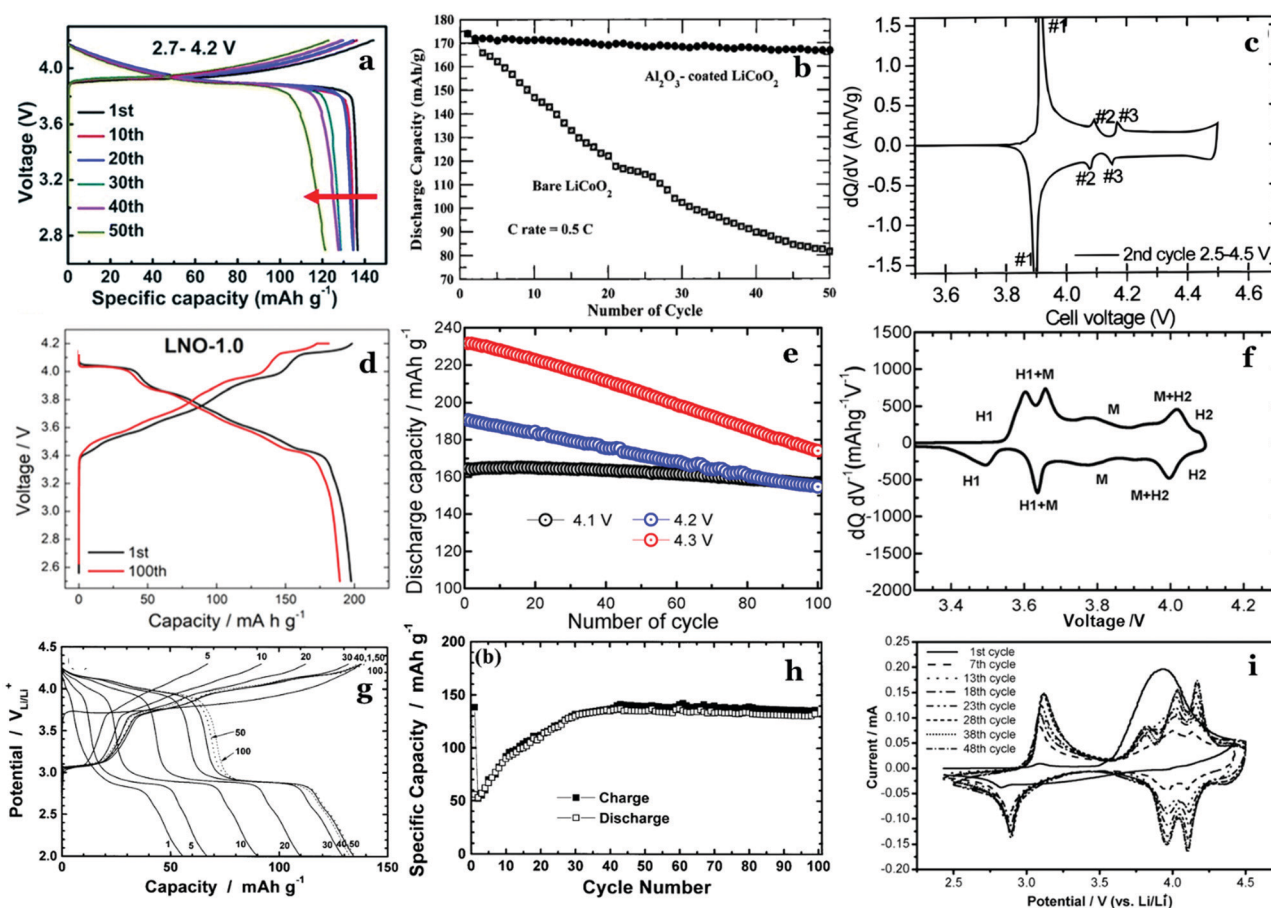


Fig. 15 (a) Charge/discharge curves for LCO cycling at a rate of 0.2C.¹³⁴ (b) First 50 cycles of bare and coated LCO cycled at 0.5C.¹³⁵ (c) dQ/dV plot of LCO second cycle illustrating structural transitions.¹³⁶ (d) First and 100th cycle (at C/3) of LNO prepared *via* optimized synthesis in pressurized atmosphere.¹³⁷ (e) Long term cycling of LNO at various upper cutoff potentials all at a rate of 0.5C.¹³⁸ (f) dQ/dV of LNO first cycle, denoting hexagonal (H) and monoclinic (M) phases and transitions.¹³⁸ (g) Charge/discharge curves of *o*- LiMnO_2 cycled at C/3.¹³⁹ (h) Long term cycling of *o*- LiMnO_2 at a rate of C/3.¹³⁹ (i) dQ/dV of *o*- LiMnO_2 over multiple cycles.¹⁴⁰ Figures reproduced with permission from ref. 134–140. Copyright (2019) Royal Society of Chemistry, copyright (2000) American Chemical Society, copyright (2000) IOP Publishing, copyright (2020) American Chemical Society, copyright (2017) American Chemical Society, copyright (2002) Elsevier, copyright (2007) Elsevier.^{134–140}

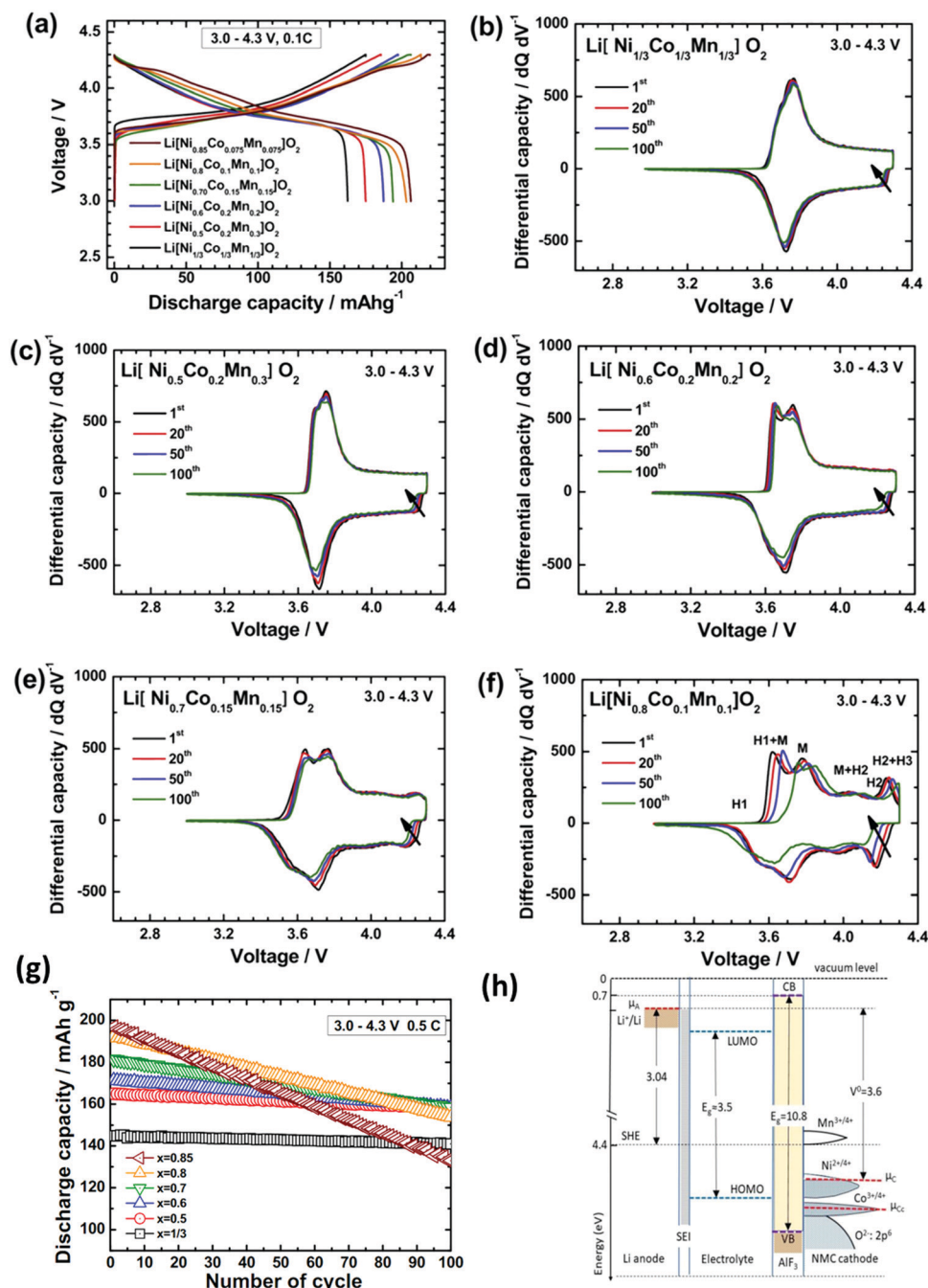


Fig. 16 (a) Charge/discharge curves for a variety of NMC compositions. (b)–(f) Differential capacity plots of various NMC compositions with increasing Ni content. (g) Long-term cycling of various NMC compositions.¹⁴⁴ (h) Energy level diagram of NMC (with AlF₃ coating) relative to Li/Li⁺.¹⁴⁵ Figures reproduced with permission from ref. 144 and 145. Copyright (2013) Elsevier and copyright (2019) AIMS Press.^{144,145}

had the same structure as LCO, but with some Co replaced by Ni. It was found that the phase transitions of LNO were suppressed at 20% atom cobalt at potentials less than 4.08, only showing a single phase transition while LNO undergoes two (H1 → M → H2).¹⁴³ However, the LNi_{0.8}Co_{0.2}O₂ had worse cycling performance. The Dahn and Ohzuku groups simultaneously explored solid solutions of the form Li[Ni_xMn_xCo_{1-2x}]O₂ wherein the square brackets indicate the transition metal layer. In this material, Co takes the 3+ state while Ni and Mn take the 2+ and

4+ states, respectively.^{16,17} These showed very good electrochemistry, comparable to LCO without phase transitions at high voltage as shown in Fig. 16(b–f). Fig. 16 shows the electrochemical properties of a wide range of NMC compositions.

Although this early work on NMC compositions gave rise to commercialized materials, given that nickel can take either the 2+ or 3+ oxidation states during synthesis, the chemistry of the NMCs is far richer than a single composition line. If one allows for Li non-stoichiometry (more to come on this in the Li-rich



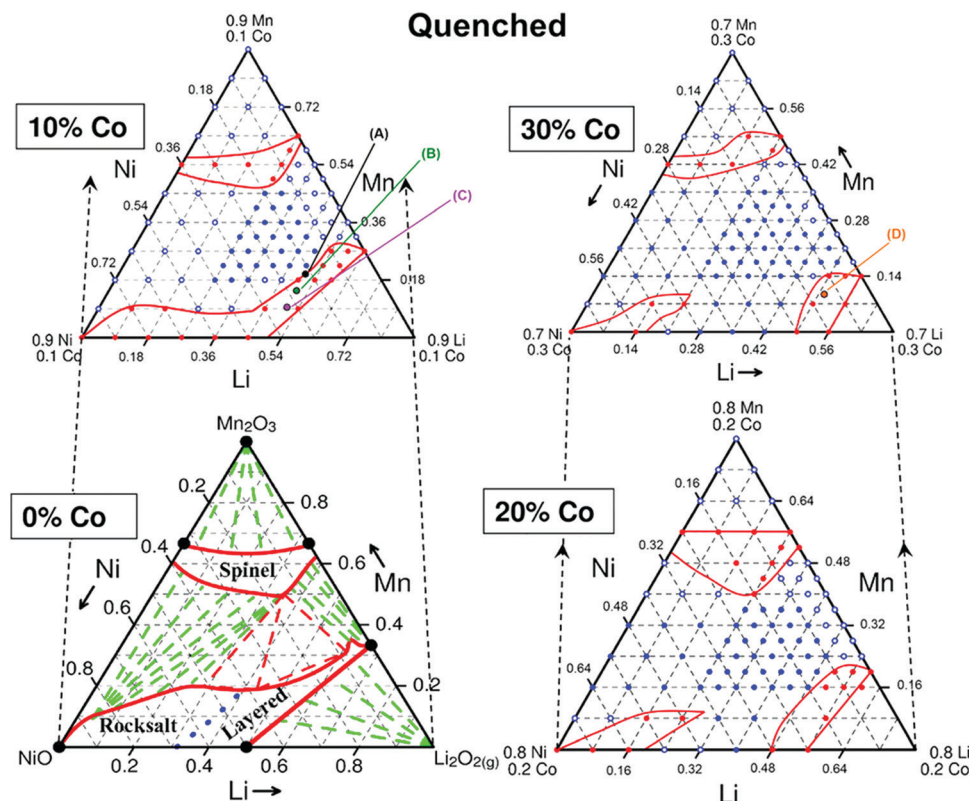


Fig. 17 Slices through the Li–Co–Mn–Ni–O pseudo-quaternary pyramid. The purple region in each slice represents the layered oxides of interest as Li-ion cathodes. Reproduced with permission from ref. 147. Copyright (2015) American Chemical Society.

oxide section below) then the solid solution region is a 3-D volume within the Li–Co–Mn–Ni–O pseudo-quaternary pyramid as illustrated in Fig. 17. In the combinatorial work that gave rise to the phase diagrams, it was found that the materials transform during cooling from high temperature such that the quenched samples may sometimes take different phases that the more slowly cooled samples.^{146–148} Recently, in addition to structural characterization, there have been reports of combinatorial electrochemical characterization in the layered metal oxide cathode space paving the way to rapid complete cathode screening, allowing orders of magnitude more data to be obtained in the same amount of time as traditional methods.¹⁴⁹

Thus, a great deal of work has gone into examining NMC compositions of the form $\text{Li}[\text{Ni}_x\text{Mn}_y\text{Co}_{1-x-y}]\text{O}_2$, which are often referred to by their TM compositions, for example NMC 532 refers to $\text{Li}[\text{Ni}_{0.5}\text{Mn}_{0.3}\text{Co}_{0.2}]\text{O}_2$, some of which are summarized in Table 1. Over a wide composition of x and y (particularly with low y to avoid high Mn compositions), NMC maintains the layered structure of LCO with the Co layers becoming transition metal layers with a distribution of all three TMs. Co is present in the 3+ oxidation state, while Ni is in the 2+ or 3+ oxidation state, depending on the relative amount of Mn, which is predominantly found in the 4+ oxidation state. If the ratio of Ni to Mn is 1 : 1, Ni is in the 2+ state. If Ni is in excess, the excess Ni is in the 3+ state.¹⁵⁰ There is, however, evidence to suggest that this is not always the case, as NMC 442 has been shown to have 20% of the Mn content in the 3+ oxidation state and 20% of the Ni in the 3+

Table 1 Summary of NMC compositions and typical capacities when charged to 4.3 V

Composition	Abbreviation	$Q_{\text{Theoretical}}$ (mA h g^{-1})	Q_{Actual} (mA h g^{-1})	Ref.
$\text{LiNi}_{1/3}\text{Mn}_{1/3}\text{Co}_{1/3}\text{O}_2$	NMC 111	278	145–175	166–168
$\text{LiNi}_{0.4}\text{Mn}_{0.4}\text{Co}_{0.2}\text{O}_2$	NMC 442	279	150–175	166 and 168–171
$\text{LiNi}_{0.5}\text{Mn}_{0.3}\text{Co}_{0.2}\text{O}_2$	NMC 532	278	150–170	172–174
$\text{LiNi}_{0.6}\text{Mn}_{0.2}\text{Co}_{0.2}\text{O}_2$	NMC 622	277	145–180	175–177
$\text{LiNi}_{0.8}\text{Mn}_{0.1}\text{Co}_{0.1}\text{O}_2$	NMC 811	276	170–200	164 and 178–180

oxidation state to compensate.¹⁵¹ During cycling, the dominant redox-active metal is Ni, undergoing redox between $\text{Ni}^{2+}/\text{Ni}^{4+}$, while it has been suggested that $\text{Co}^{3+}/\text{Co}^{4+}$ redox can be accessed once two thirds of Li are depleted in NMC111.^{152,153} However, an irreversible change from O3 to O1 stacking occurs before 0.75 Li can be extracted.¹⁵⁴ In practice NMC has a capacity of around 150–190 mA h g^{-1} when cycled to 4.3 V,¹⁵⁵ depending on the composition, with a longer cycle life than both LNO and LMO attributed to an increased structural stability. NMC is also a safer cathode material than LCO, with a higher onset temperature for thermal runaway (Fig. 18).

While NMC is one of the most performant cathodes to date, like state-of-the-art-LCO, the cathode particles can crack due to the change in volume from intercalation/deintercalation and phase transformations, resulting in a larger surface area to



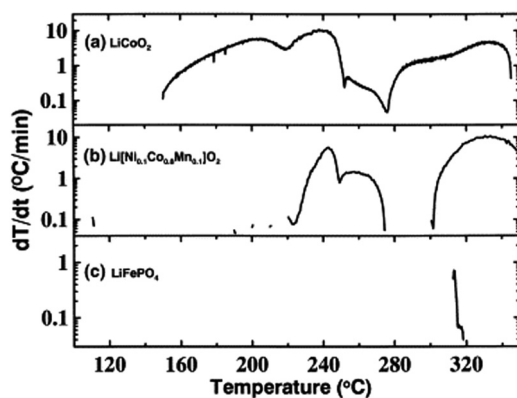


Fig. 18 ARC of LCO, NMC 118, and LFP showing the later onset of thermal runaway in both NMC and LFP compared to LCO. Figure reproduced with permission from ref. 159. Copyright (2004) Elsevier.

react with electrolyte and poorer overall performance.^{68,156–158} Thus, the stability could be improved through substitutions and coating, which will be discussed in the next section.

Another issue addressed by the wide composition window of NMCs is related to an on-going effort to minimize the cobalt content. Cobalt is the most expensive element in the material and its sourcing is ethically questionable due to unsafe mining practices, contaminating local communities, and in some cases utilizing child labour.^{160–162} As a result, contemporary cathodes have been moving toward Ni-rich compositions. The most studied in order of increasing Ni content are $\text{LiNi}_{0.4}\text{Mn}_{0.4}\text{Co}_{0.2}\text{O}_2$ (NMC 442), then $\text{LiNi}_{0.5}\text{Mn}_{0.3}\text{Co}_{0.2}\text{O}_2$ (NMC 532), $\text{LiNi}_{0.6}\text{Mn}_{0.2}\text{Co}_{0.2}\text{O}_2$ (NMC 622), and $\text{LiNi}_{0.8}\text{Mn}_{0.1}\text{Co}_{0.1}\text{O}_2$ (NMC 811). Specific discharge capacity tends to increase with increasing Ni content, but comes at the cost of decreased structural stability and cycle life.¹⁴⁴ Particularly problematic forms of structural instability are cation mixing and surface reconstruction due to metal loss through reactions with the electrolyte. In cation mixing, Ni migrates to Li sites, blocking Li diffusion. This phenomenon can be identified using a few indicators in XRD spectra: the c/a ratio, the intensity ratio of the 003 and 104 peaks.^{163,164} Additionally, Ni-rich cathodes suffer from poor thermal stability due to the reactivity Ni^{4+} with the electrolyte, decreasing cycle life, poor rate performance, and decreased safety. A review of these issues can be found here.¹⁶⁵ Several methods have been explored to resolve the issue of structural instability and surface reactivity, most notably: coatings, metal substitutions, and core-shell structures (considered to be non-uniform substitutions within the context of this review article). Coating consists of applying a layer of a different material to the surface of cathode particles to prevent surface reactions, structural degradation, and cracking. Metal substitutions involve making a change to the TM composition through the entire material. Core-shell structures are a hybrid of these two approaches: the core of the particle is composed of one layered oxide, while the shell of the particle (usually significantly thicker than a traditional coating), is composed of a different, more structurally resilient, cathode material. The two materials must be close enough in structure to maintain the particle's integrity during

battery operation. Herein, we focus primarily on metal substitutions but will also briefly explore other methods for the sake of comparison in order to evaluate whether the substitutions are warranted or if, for example, a simple coating yields a better battery material.

4.3 NCA: Co, Al co-substitution into LiNiO_2

The other mixed metal layered oxide of note is $\text{LiNi}_{1-x-y}\text{Co}_x\text{Al}_y\text{O}_2$ (NCA). The rationale of this material is similar to that of high Ni NMCs: Ni acts as the main redox-active component, while Co serves to enhance structural order and Al increases structural stability, though Co can contribute electrochemically at high potentials. The most common composition of NCA is $\text{LiNi}_{0.8}\text{Co}_{0.15}\text{Al}_{0.05}\text{O}_2$.¹⁵⁵ NCA materials have a high theoretical capacity of 279 mA h g^{-1} and in practice achieves $180\text{--}200 \text{ mA h g}^{-1}$.¹⁸¹ It also has improved rate capability over NMC, giving it a higher specific power making it more suitable for the booming electric vehicle industry. As with NMC, LCO, and LNO, NCA also has a layered rock salt structure with alternating transition metal oxide and lithium layers with a mix of metals in each TM layer. NCA also suffers from similar issues to Ni-rich NMC materials. Li/Ni mixing is still an issue, causing structural degradation and impeding Li^+ intercalation paths. $\text{LiNi}_{0.8}\text{Co}_{0.15}\text{Al}_{0.05}\text{O}_2$ has been shown to have the best electrochemical performance out of a range of compositions,¹⁸² balancing a high capacity with decent cycling life. There is also evidence that Aluminate impurities can form at concentrations above 5%.^{182,183} Additionally, NCA has an improved thermal stability when compared to aluminum-free $\text{LiNi}_{0.8}\text{Co}_{0.2}\text{O}_2$, improving safety and high-temperature performance. Like other Ni-based cathode materials, Ni^{4+} is highly unstable at high temperature, reacting with HF in electrolyte and causing material degradation, namely at the surface.¹⁸⁴ Ultimately, NCA suffers from worse cyclability compared to non-nickel-rich NMC cathodes. Like NMC, coatings and substitutions have been studied to improve the structural stability and electrochemical performance, these will be discussed in the next section.

4.4 Substitutions into NMC and NCA

One appeal of studying the impact of substitutions on the layered oxides is that metal substituents can be easily incorporated into existing synthesis methods thus ensuring easy scale-up. The most commercially-relevant synthesis route for layered TM oxides is a solution-based co-precipitation synthesis.^{185,186} This method takes place in two stages. First, a precursor material is prepared by dissolving the desired metals in the form of sulfates, nitrates or acetates into a pH-controlled solution at the desired stoichiometric ratios. Ammonium is then added to the solution to act as a chelating agent, and NaOH is added to cause co-precipitation. With careful control of the reaction conditions (temperature, pH, stir rate, flow rates of reactants), the morphology can be tuned to give large spherical particles as shown Fig. 19.¹⁸⁷ The second phase is to add a lithium source such as LiOH is then added to the dried and rinsed precursor ($\text{M}(\text{OH})_2$ where M is the mixture of TMs) and heated to high temperature $800\text{--}1050 \text{ }^\circ\text{C}$ are commonly



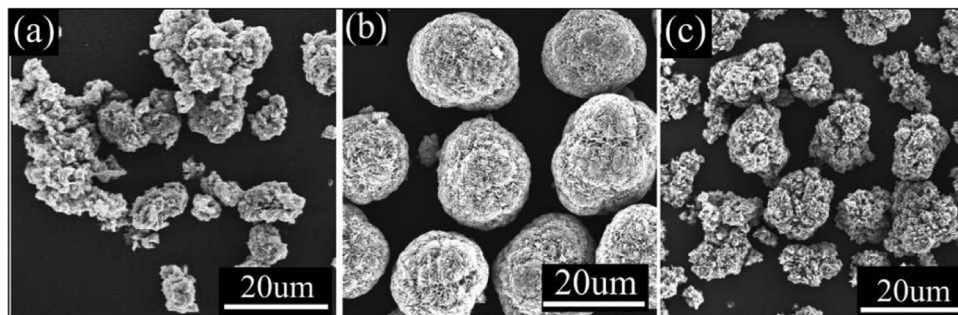


Fig. 19 Optimized synthesis of layered oxide cathode material. Figure reproduced with permission from ref. 187. Copyright (2015) Elsevier.

used to obtain a lithiated layered cathode material. Successful syntheses have been reported by either adding the metal substituent at the precursor or at the lithiation phase. Although these are in principle equivalent, one would expect a more uniform substitution when done at the precursor stage while substitutions at the lithiation stage might yield a structure closer to a core-shell (*i.e.* more of the substituent would be found at the surface of particle unless aggressive sintering conditions are used). Alternatively, the material can be produced through a purely solid-state reaction wherein metal oxides or hydroxides (including Li) are mixed in the desired stoichiometric ratio using a mortar and pestle or milling before being heated to the desired temperature to yield the final product.^{188,189} However, this tends to give a less homogeneous product and requires much longer sintering times in order to reach the equilibrium layered phase.¹⁹⁰ Given that the industrial norm is to use co-precipitation, this is the synthesis route of choice for studies of substitutions into NMC or NCA. After synthesis, it is important to determine the composition of the material. While in the case of doping into semiconductors, the substitution levels are so low that the level of substitution can only be confirmed by the impact on the properties, this is not the case in the vast majority of substitution studies into battery materials. Specifically, the %-level substitutions can be readily verified using inductively coupled plasma coupled with optical emission spectroscopy, atomic emission spectroscopy, or mass spectroscopy (ICP-OES, ICP-AES, or ICP-MS).^{191–193} Though ICP techniques do not distinguish between coatings, bulk substitutions, and additional material phases, it is useful for verifying the presence of the substituted metal and, more importantly, the Li content of the synthesized material.¹⁹⁴ To verify whether the metal substitution was successful in penetrating the material, electron microscopy and X-ray techniques such as SEM-EDX and XRD may be used to determine where the substituent ended up within the material (see Section 3 for more detail).¹⁹⁵

Many metals have been substituted into layered oxide materials to increase their performance and longevity, including Al, Ti, Mg, Zr, Nb, Y, La (Fig. 20). The effects of the substitutions are manifold and a single substituent may result in multiple effects making it difficult to fully evaluate its impact. Substitutions have been shown to prevent cation mixing, suppress phase transitions, and prevent reaction with electrolyte. For layered

metal oxide cathode materials, the two most well-studied substitutions are Al and Ti. Both have been shown to increase lifetime, but at the cost of capacity, as both substituents are electrochemically inactive. These will be explored in more significant detail due to the volume of experimental and theoretical work.

Fig. 20 shows a summary of the impact of substituent metals and select surface coatings on discharge capacity and percent capacity loss per cycle on various NMC compositions. Fig. 21 shows the same information, but for NCA cathodes. In all cases, one finds a wide variety of performance for unsubstituted NMC/NCA, which makes comparison between studies impossible. Thus, for Fig. 20 and 21, we report the performance for the unsubstituted material for that study in blue and the optimum substituted performance in red. For long-term cycling, the minimum number of cycles was 50 and maximum 250. It is clear that many substituents and coatings decrease capacity fade while having only a small impact on capacity, and sometimes a beneficial one. It is also clear that there is significant variation in pristine cathode materials making direct comparisons between different articles difficult, though looking at the relative picture of the performance of pristine materials *vs.* their substituted counterparts useful conclusions can be made. It is also worth noting that in many cases, a series of substituent concentrations were presented in the papers from which these data were extracted; only the best performing composition is reported here.

The variance in material performance is exemplified by the inconsistent behaviour of Al-substituted NMC. In most reports of Al-substituted NMC, the initial capacity is decreased,^{222,223} though there are some experiments that show the contrary using a similar composition and synthesis method. An increased capacity has also been reported in NMC 111 by Hashem *et al.*²²⁴ who suggest that the increased capacity is due to reduced anti-site defects and the reduced cyclability is due to strain placed on the lattice through Al substitution. It is also possible that the discrepancy may be the result of synthesis conditions, as the material by Hashem *et al.* was synthesized by the sol-gel method while the others cited in this work used the co-precipitation technique. In Ni-rich NMC, Al³⁺ has been shown to increase the retained capacity by nearly 20% while only reducing the initial capacity by less than 1.2%.¹⁹⁸ This is attributed to a relief in lattice strain caused by Co³⁺ as it is substituted by Al³⁺ introducing stronger



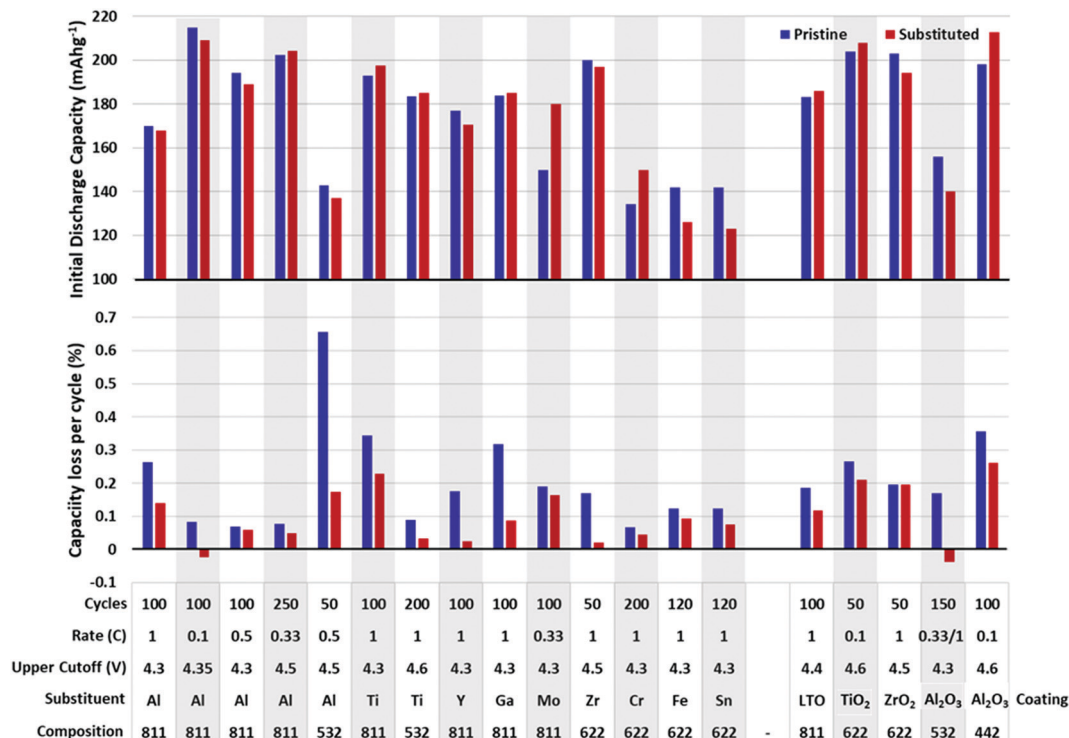


Fig. 20 Initial reversible capacity (top) and capacity lost per cycle (bottom) of pristine and substituted or coated NMC materials. Data extracted from ref. 19, 33, 178, 188 and 196–208. Below each plot is listed the cycling rate, upper cutoff voltage, substituent/coating, composition of unsubstituted. The gap in the x-axis separates substitutions from coatings.

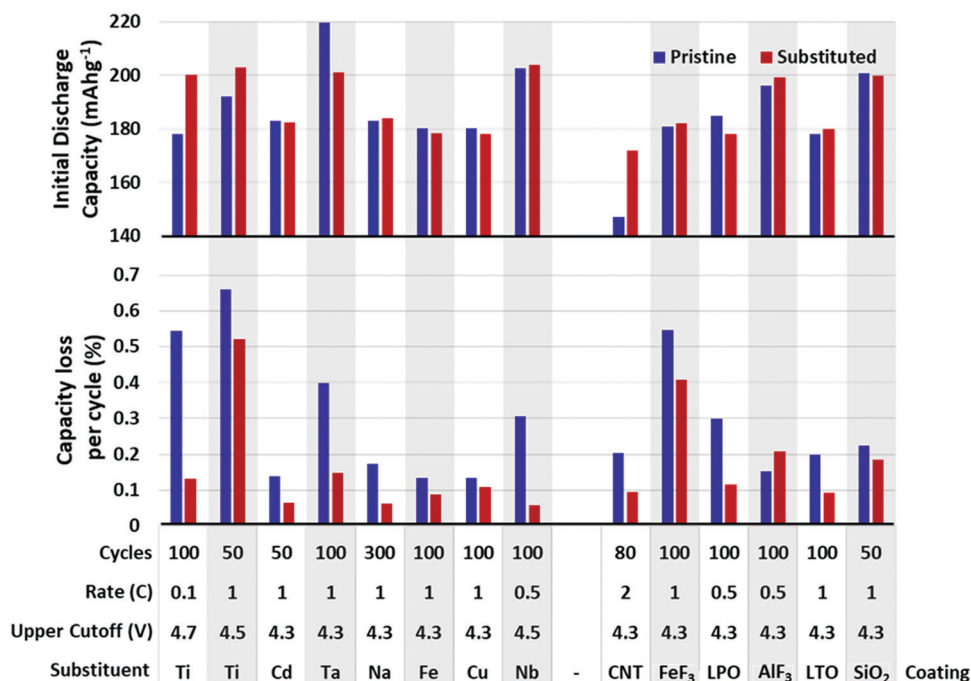


Fig. 21 Initial reversible capacity (top) and capacity lost per cycle (bottom) of pristine and substituted or coated NCA materials. Data extracted from ref. 209–221.

Al–O bond that reduces the amount of cracking and electrolyte penetration. This is supported by calculations on the NMC

532 composition, which point to the strong Al–O ionic-covalent bonding stabilize the cathode structure.²²⁵ The decreased



capacity fade observed by Zhao¹⁹⁸ was also attributed to the formation of Al–O compounds on the surface of the NMC particles, acting as a weak coating that can scavenge HF and prevent structurally compromising reactions between the cathode and electrolyte.¹⁹⁷ The protective effects at the surface can be improved through gradient substitution, where the more susceptible surface of the material has a higher Al content to protect from material degradation, while the less vulnerable centre maintains its pristine, or near-pristine composition, gaining the improved capacity of the unsubstituted material without compromising cycle life.²²⁶ Al³⁺ substitution also increases thermal stability and safety of Ni-rich NMC materials, offsetting the onset of exothermic reactions and making their onset more gradual.^{197,226} Aluminum substitution has also been shown to decrease Li/Ni mixing in NMC 662 when substituted for Mn and in 811 which likely contributes to the increased stability seen in these compositions when substituted.^{226,227}

Titanium also proves to be an effective substitution for layered metal oxides. Like Al, it has been shown to increase capacity retention significantly, while actually increasing reversible capacity.^{188,199,228} Though, there have been capacity decreases reported in NMC 811.²²⁹ When Ti⁴⁺ is substitutes Co³⁺ the charge is compensated *via* Mn⁴⁺ reduction to Mn³⁺. While Mn³⁺ is Jahn–Teller active, the structural stability gained through Ti substitution appears to outweigh any negative effects of Mn³⁺.²²⁸ The increase in cycle life in Ti-substituted NMC is broadly attributed to stronger Ti–O bonds improving the structural integrity of the material and preventing particle cracking and is supported by DFT calculations.^{229,230} In Ni-rich NMC, Ti⁴⁺ substituent increases the reversibility of the H2–H3 phase transition and reduces lower polarization during cycling, which would also improve electrochemical performance.²³¹ Ti has also been shown to be an effective substitution in NCA. In contrast with NMC, Ti⁴⁺ has been shown to reside in octahedral Li⁺ sites. This expands the interlayer distance, allowing for easier Li⁺ movement as well as reducing the severity of Li/Ni mixing.^{210,232} Liu *et al.* have found the concentration of 1% atom to be optimal. Ti may also act as “pillar” ions, supporting the structure through repeated lithiation/delithiation and preventing collapse.²³³

Zirconium has also been explored as a substituent in both NMC and NCA.²⁰² Similar to Ti, Zr has a larger radius than Ni, Mn, and Co in their regular oxidation state in NMC, causing a volume expansion of the lattice and interlayer spacing. Zr has also been found occupying Li⁺ sites, increasing the lattice spacing on the *c* axis and serving similarly to Ti in NCA as a “pillar” ion, stabilizing the structure during lithiation/delithiation and discouraging Li/Ni mixing. Zr has been found to perform similarly in terms of cycling performance and capacity loss as Ti substitution.²³⁴ Zr and F co-substitution was also shown to increase capacity retention by over 15% over 200 cycles in NCA.²³⁵

The poor performance of Fe substitution in both NMC and NCA is attributed to a hindrance of Li diffusion. Though it has also been shown that the incorporation of Fe³⁺ increases the

thermal stability of NMC 622.²⁰⁴ Fe was also shown to decrease Li/Ni disorder in NCA along with Cu substitution.²¹⁴ The electrochemical performance of some other substituents are shown in Fig. 20 and 21.

While many metals have been used as substituents in layered transition metal oxides, there is still much to explore. Multi-metal substitutions are studied sparingly and open an enormous space of possible combinations to be optimized. Several metals have not been used as substituents in the literature, specifically many of the lanthanides and heavy d-block elements, and the concentration and method of substitution has not been optimized for many materials. There is also the issue of ensuring that scale-up is feasible for the substituted layered oxide materials, as many studies use solid-state syntheses that are not employed at the industrial scale. Thus, there is work to be done adapting metal-substituted cathode synthesis with scalability in mind.

4.5 Substitutions vs. coatings

Coatings for layered metal oxide materials come in a wide variety. There are metal oxides,^{172,236,237} polyanionic lithium compounds,^{238,239} fluorides,^{240,241} and more that have been used with some success. Coatings serve to stabilize the surface of the cathode particle without inhibiting Li⁺ diffusion. There are several methods of applying coatings to cathode materials such as co-precipitation, sol–gel, and various deposition techniques. Wet chemical techniques are desirable, as they are economical compared to other methods while being easier to integrate into industrial cathode syntheses. In co-precipitation synthesis, the metal oxide precursor is synthesized as described previously. Then the synthesized cathode and soluble form of the material to be coated (*e.g.* the acetate, nitrate, or sulphate of the metal to be coated as a metal oxide) are stirred in a reactor, dried, and annealed at 300–500 °C.^{242–244} This method also works with metal phosphate and fluoride coatings by adding the metal phosphate for phosphate coatings, or the metal nitrate and ammonium fluoride for fluoride coatings, in solution with the prepared pristine cathode material before drying and annealing.^{245,246} Note the reduced temperature in the annealing process compared to the calcination temperature of 800–1050 °C used for the layered oxide synthesis. If the temperature used for sintering the coating is too elevated, substitutions into the layered oxide may occur. In such cases, non-uniform coatings will occur, potentially leaving some surface area of the cathode material exposed, thereby reducing the effectiveness of the coating.²⁴⁷ Sol–gel techniques suffer a similar deficiency. Alternatively, deposition techniques such as atomic layer deposition and sputtering can be used to achieve a more uniform coating. However, due to their cost, it is unlikely these methods will be used in full-scale cathode production.

Alumina is perhaps the most studied coating, as discussed previously, it was used to stabilize LCO and has been used to stabilize NMC.^{208,248–253} It has been shown to prevent TM reduction at the surface, oxygen loss, and even prevents phase transformations.^{208,254,255} It has also been shown to scavenge



HF, forming protective Al-F compounds, resulting in a healthier SEI layer.²⁵⁶ Titanium coatings have been shown to be similarly effective.^{206,257–259} See Guan for a more rigorous review of cathode coatings.¹⁸

Alternatively, performance can be improved using a core-shell structure, or a core-gradient-shell structure.^{260–267} The idea behind a core-shell particle is to have high energy density cathode material in the core, surrounded by a more stable, yet still active, cathode material to protect the core from degradation (Fig. 22a–c). Fig. 22d demonstrates how effective this approach can be in prolonging cycle life, while only minimally reducing the specific capacity of the overall cathode. This technique can also improve the safety of Li-ion batteries by improving the thermal stability of the cathode overall. For example, a core-shell gradient was able to suppress the cathode self-heating onset temperature by over 40 °C.²⁶⁸ Being an active cathode material itself, the shell is ionically conductive toward Li⁺ and does not significantly dampen rate performance.^{266,267} One shortcoming of core-shell particle design is the susceptibility to internal separation of the core and shell.²⁶⁰ Upon repeated expansion/contraction of the materials, strain is placed on the interface, eventually causing a separation of the two materials. This separation allows electrolyte to access the

more-susceptible core and react with its surface, causing material degradation and metal dissolution. This problem has been partially addressed by using gradient compositions to smooth the transition between core and shell. This has reduced the propensity for internal separation of core and shell to occur, but cracking of the particle remains an issue. Fig. 23 shows the effectiveness of several core-shell compositions.

5. Li-Rich layered oxide cathodes

5.1 Ordered Li-rich oxides

At the same time as NMC solid solutions were being discovered and studied for commercialization, another class of layered oxide cathodes was discovered. In 2001, the Dahn group looked at Li[Ni_xLi_{1/3–2x/3}Mn_{2/3–x/3}]O₂ which take the same layered structure as LCO, NMC and NCA, but with some excess lithium present on the transition metal layer.^{277–279} Fig. 24 shows the electrochemistry for these materials. Up to 4.4 V, the cycling looks very much like any other layered material. However, above 4.5 V, the material transforms during first charge and the result is a far larger reversible capacity well beyond that expected based on Ni redox alone (e.g. for $x = 1/4$, Ni

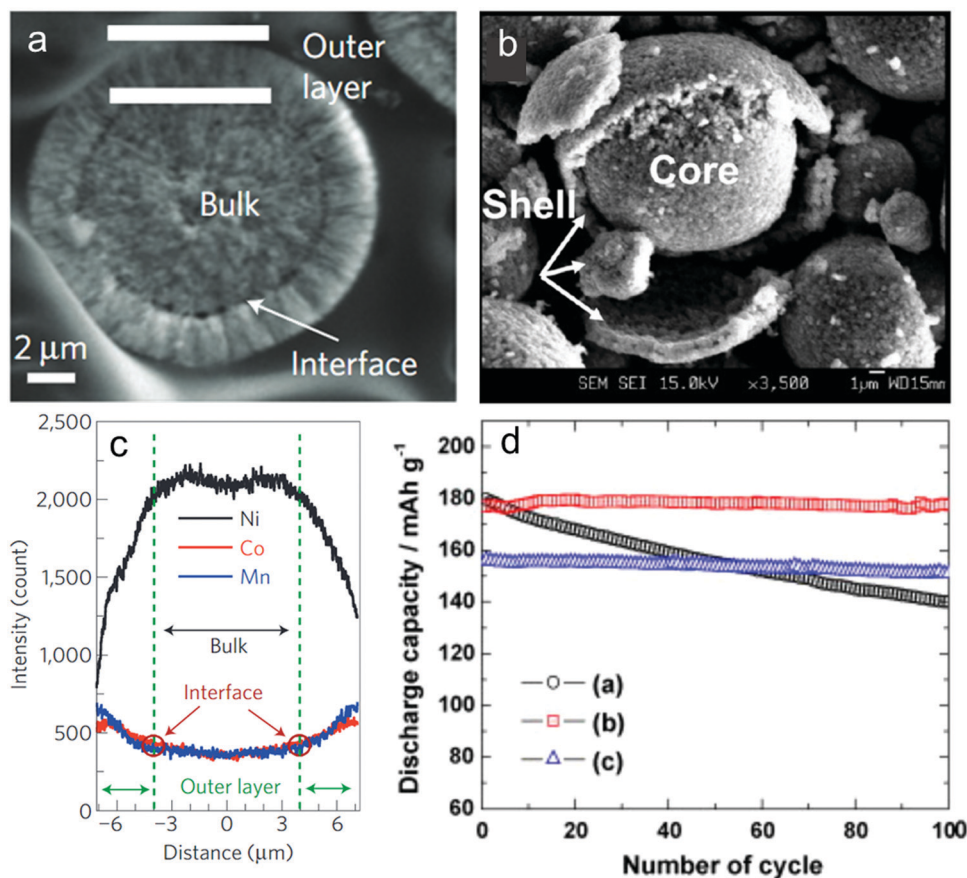


Fig. 22 (a) Pristine core-shell particle. (b) Particle with core-shell separation. (c) Metal concentration distribution in core-shell particle. (d) Long-term cycling of (a) NMC 111 core material, (b) LiNi_{0.5}Mn_{0.5}O₂ shell material, (c) Li[(NiCo_{1/3}Mn_{1/3})_{0.8}(Ni_{1/2}Mn_{1/2})_{0.2}]O₂ core-shell material. Figures reproduced with permission from ref. 260, 269 and 270. Copyright (2014) American Chemical Society, copyright (2009) Nature, copyright (2010) Elsevier.



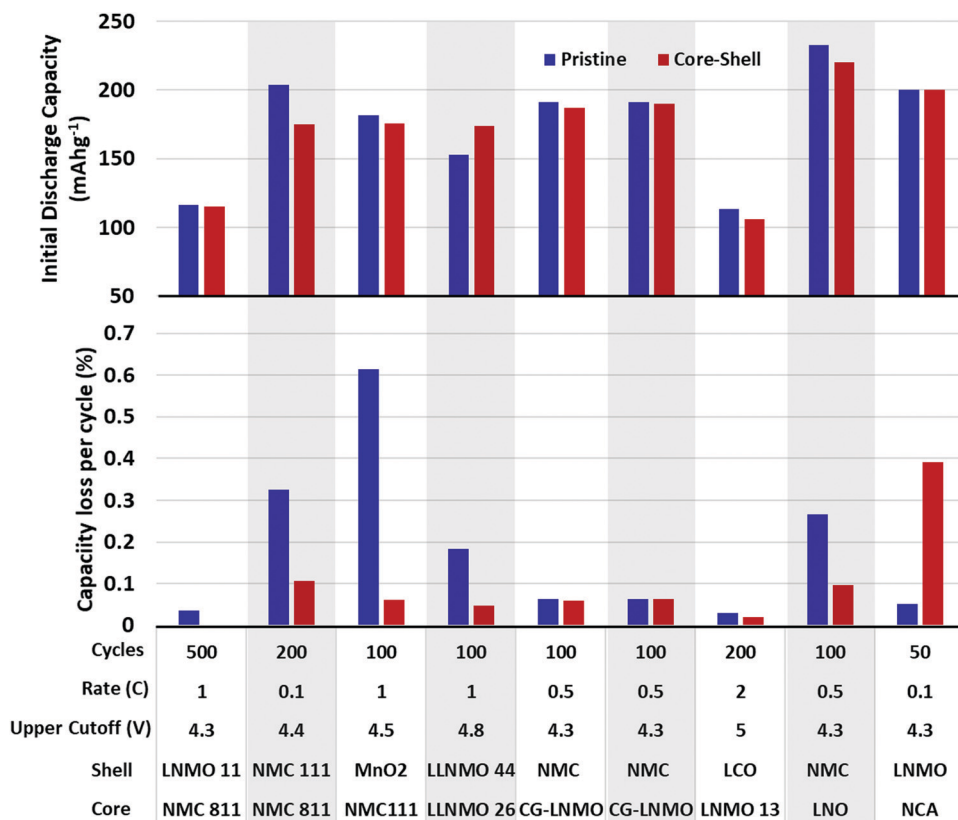


Fig. 23 Initial reversible capacity (top) and capacity lost per cycle (bottom) of core-shell materials. Data extracted from ref. 261, 262 and 271–276.

redox accounts for the first 120 mA h g⁻¹ then the remaining 180 mA h g⁻¹ on charge is unaccounted for).

This peculiar class of material has now been studied extensively both in terms of designing a better commercial cathode but also from the fundamental point of view to understand the mechanisms involved. We will review both here. This new class of cathode materials are the Li-rich cathode materials, where more than one Li atom is present per atom of transition metal in the layered structure such that the stoichiometry is of the form Li_x[Li_{1-x}M_{1-x}]O₂. The materials of highest interest for commercialization have Ni, Mn, and Co as the M atoms such that they are referred to as Li-rich NMC. As shown previously in Fig. 17, there is a wide range of compositions in the layered region in the Li–Ni–Mn–Co–O system, and a high portion of these are Li-rich. A few key compositions have been studied extensively including: Li_{1.2}Mn_{0.6}Ni_{0.2}O₂ ($x = 0.2$ in the Li[Ni_xLi_{1/3-2x/3}Mn_{2/3-x/3}]O₂ notation),^{280–283} and Li_{1.2}Mn_{0.54}Co_{0.13}Ni_{0.13}O₂.^{284–288} Each of these materials show the same key features: (1) a large reversible capacity beyond that expected from metal redox, (2) an irreversible transformation on first charge, and (3) a gradual decrease in average voltage with extended cycling (voltage fade). The Li-rich oxides therefore show a significant increase in capacity over all commercialized cathodes, reaching greater than 250 mA h g⁻¹.²⁸⁹ However, this class of material is plagued by voltage fade, poor rate capability, and a gradual transformation towards a spinel-like material particularly in the Li-rich materials that have a high Mn

content.^{25,95,290,291} Nonetheless, Fig. 24 makes it clear that composition has a great impact on the cycling performance and so continued study in varied compositions, including substitutions into Li-rich NMCs, are certainly warranted.

5.2 Mechanisms during battery operation of Li-rich oxides

The complex electrochemistry taking place in the Li-rich oxides has taken over 15 years to understand, with some clarity still lacking. It is impossible to discuss the role of substitutions into these materials without first discussing the mechanisms involved during battery use. As mentioned above, the simple model of TM redox compensating all Li removal that worked so well for 30 years in layered oxides fails entirely when it comes to Li-rich oxides. Ten years after their discoveries, two groups simultaneously identified that much of the reversible redox that is generated after the first charge comes from reversible oxidation of oxygen.^{55,285–287,292–295} There is still debate on whether or not the oxidation of oxygen results in O–O dimers in the Li-rich NMC,²⁹⁶ but there is no longer any debate as to whether or not reversible oxidation of oxygen occurs. This accounts for a great deal of the extra capacity seen, and a great hunt has been underway to find materials that engage oxygen redox while minimizing oxygen gas loss. Oxygen loss in Li-rich oxides certainly occurs,^{280,297} but only from the surface of particles²⁸⁶ and this results some activation of Mn (*i.e.* it can be reduced from 4+ to 3+ during the next discharge) but it also



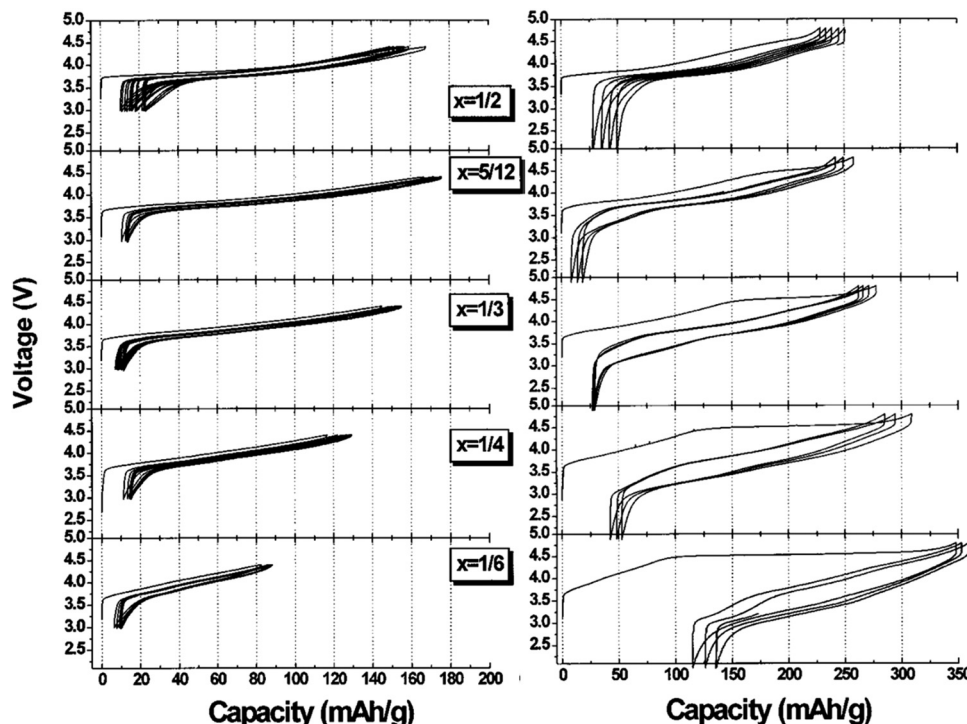


Fig. 24 Voltage curves for $\text{Li}[\text{Ni}_x\text{Li}_{1/3-2x/3}\text{Mn}_{2/3-x/3}]\text{O}_2$ up to 4.4 V (left) and up to 4.8 V (right). Whereas all samples look similar to traditional layered oxides like LCO up to 4.3 V, they all undergo an irreversible change upon cycling above 4.5 V and this leads to a large reversible capacity as high as 230 mA h g^{-1} . Figure reproduced with permission from ref. 279. Copyright (2002), IOP Publishing.²⁷⁹

gives rise to some irreversible capacity if there is insufficient Mn to compensate.

However, none of these mechanisms explain voltage fade. Studies on model systems (Li–Ru–Ti–O) showed that during the first charge, there is a massive migration of TMs from the TM layers to the vacant Li layers.^{292,298} This occurs reversibly, with a return to the highly ordered layered structure at the end of discharge. Mobile atoms like Ti do this readily, but gradually with continued cycling, the Ti was found to be trapped in tetrahedral sites near the Li layer. This trapping of Ti simultaneously gives voltage fade and starts the transformation towards the spinel-like material since in Li-rich NMCs. Thus, it appears that the voltage fade comes from an imperfect TM migration during cycling.²⁹⁹ The working principle has been that, in the Li-rich NMC materials, the severe voltage fade is attributed to the formation of spinel-like domains where Mn gets trapped in tetrahedral sites. Additionally, surface stability remains an issue due to Mn dissolution through electrolyte loss, causing structural disorder at the surface, increasing impedance, lowering capacity and contributing to voltage fade. Thus, improving the structural stability and interfacial stability of these materials is of the utmost importance. Several coatings and substitutions have been applied to achieve these goals and will be reviewed in the next section.

5.3 Substitutions into Li-rich oxides

There has been a significant effort investigating the effects of substitutions in lithium-rich materials (Fig. 25). As mentioned

earlier, these Li-rich materials suffer from significant capacity and voltage fade, thus improving structural stability is the goal of substitutions in this case. Stemming from work done on layered oxides, Al and Ti have been substituted extensively into Li-rich materials to varying effect. In $\text{Li}_{1.5}\text{Mn}_{0.675}\text{Ni}_{0.1675}\text{Co}_{0.1675}\text{O}_2$, Al^{3+} occupies Li^+ sites, improving the kinetics of lithiation/delithiation and reducing impedance, thus increasing rate performance.³⁰⁰ In $\text{Li}[\text{Li}_{0.23}\text{Ni}_{0.15}\text{Mn}_{0.52}\text{Al}_{0.10}]\text{O}_2$ it was found that the presence of Al improved long-term cycling performance, and rate performance.³⁰¹ It has been hypothesized that Al helps prevent Mn migration causing the conversion to spinel-like phases. In $\text{Li}_{1.2}\text{Ni}_{0.16}\text{Mn}_{0.56-x}\text{Al}_x\text{Co}_{0.08}\text{O}_2$, Al^{3+} was substituted for Mn^{4+} at compositions ranging from $x = 0$ – 0.08 and mainly occupied TM layer sites while causing the formation of some Ni^{3+} to compensate for the aliovalent substitution. This reduced initial capacity, likely due to suppression of material activation, while greatly increasing capacity retention and reducing voltage fade by over 50% in the case of $x = 0.08$.³⁰²

Ti has been shown to function similarly to Al. Ti^{4+} has also been shown to occupy Li^+ sites and increase repulsion with Mn^{4+} , suppressing Mn migration while reducing oxygen loss due to stronger Ti–O bonds.³⁰³ Feng *et al.* showed through theoretical work that Ti in the Li later reduced the band gap in Li_2MnO_3 , increasing the electrical conductivity of the material and confirmed that the activation energy of Mn migration increased compared to the pristine material.³⁰⁴

Chromium has also been substituted into Li-rich oxides with some success. When Cr was substituted into TM sites, it suppressed



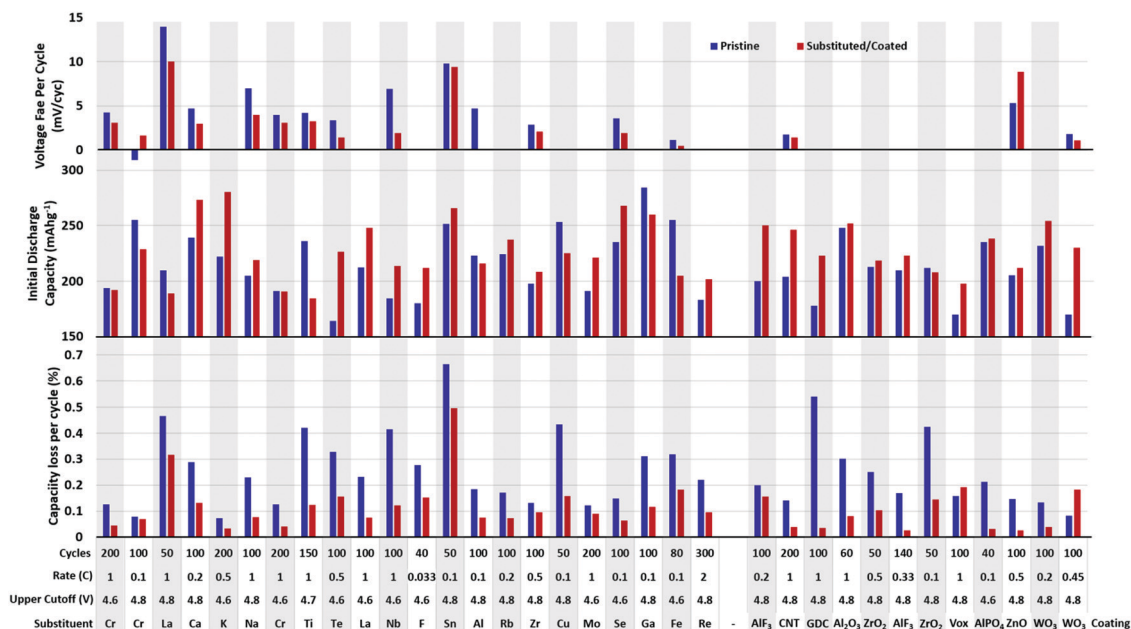


Fig. 25 Voltage loss per cycle (top), initial reversible capacity (middle), and capacity lost per cycle (bottom) of pristine and substituted or coated NCA materials. Data extracted from ref. 100, 303, 305 and 308–337.

voltage fade, but also lowered the energy density. The authors hypothesized that the stronger Cr–O bonds may hinder anionic redox. Stronger TM–O bonds may also prevent full activation of the Li-rich material, lowering capacity.³⁰⁵

Magnesium substitution can behave similarly to Ti substitution, replacing Li and acting as a pillar, while also opening Li-ion channels, improving structural stability and rate capability.³⁰⁶ Mg²⁺ substitution has also been shown to lower capacity if it replaces Mn. This capacity loss is attributed to the inactivity of Mg and potentially also due to stronger Mg–O interactions preventing anionic redox or hindering activation of Li₂MnO₃.³⁰⁷ Similarly to the stoichiometric layered metal oxides, there is still much to explore in the multi-metal substitution space with Li-rich materials. The greatest need is in preventing the voltage fade, where numerous studies have shown this can be done *via* metal substitutions in model systems (using expensive metals such as ruthenium and iridium),²⁹³ the key now is to obtain the same performance using materials that are cost-effective. Some work has been done in this regard, using selenium³⁰⁸ and iron³⁰⁹ substituents, but there is far more optimization required to see the potential benefit of metal substitution on reducing voltage decay.

5.4 Disordered rocksalt Li-rich oxides

In the Li–Ni–Mn–Co–O system (Fig. 17), there are the ordered layered materials that have been of high interest, but at lower Li content we see the rocksalt region. The rocksalt structure (or disordered rocksalt) is identical to that of the layered oxide, except every metal layer is a random distribution of Li and TMs. At low Li contents ($x < 1.2$ in Li_xM_{2–x}O₂), the Li lies in 3-D networks but there is insufficient Li to ensure percolation and as such these prove to be very poor cathode materials due to

limited Li diffusion. In 2014, the Ceder group showed that at higher Li content one can obtain a disordered rocksalt phase with percolation of Li, such that the cathode can be charged by 3-D diffusion through the disordered Li channels wherein Li may diffuse through paths without any nearest-neighbour TMs.³³⁸ This was obtained on Li–Mo–Cr–O materials that in fact crystallized in the layered structure but converted to the disordered rocksalt (DRX) structure during the first charge.

Since that time, there has been a search for Li-rich materials that can be made directly in the DRX structure and this has proven to be a very rich chemistry allowing for vast substitutions (both cationic and anionic with F being partially substituted for O). As such, these materials are of high interest in the current review article. It should be noted that one DRX material, LiNi_{0.5}V_{0.5}O₂, shows a high capacity of 264 mA h g^{−1} despite not having excess Li,³³⁹ thus the percolation condition may not prove essential in these materials. Regardless of composition, the synthesis method of choice for DRX materials is mechanochemistry wherein ball milling is used to make the desired materials. Room-temperature ball milling prevents the layered materials from forming and also keeps the particles small, which proves necessary in order to extract the lithium. An example of a DRX formed during synthesis is Li–Mn–V–O and the electrochemistry of this material is shown in Fig. 26. The cycling curve no longer shows an irreversible transformation during the first charge, in fact the irreversible capacity is negligible.³⁴⁰ The discharge capacity is also extremely high, approaching 300 mA h g^{−1}, slightly greater than that achieved in the best Li-rich layered oxides. Despite these very promising features, there are a number of detrimental features in the electrochemical performance that are making industrial application of these materials difficult. In particular, the transport



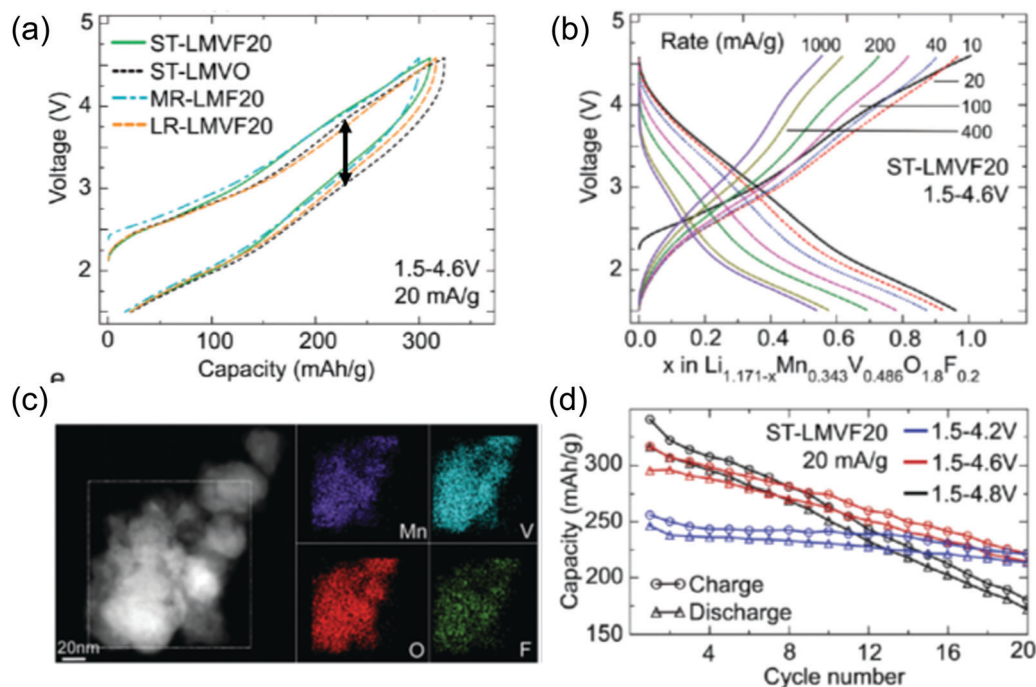


Fig. 26 Battery performance of DRX Li-Mn-V-O-F: (a) first cycle voltage curves for $\text{Li}_{1.143}\text{Mn}_{0.286}\text{V}_{0.572}\text{O}_2$ (ST-LMVO) as well as some materials with partial substitution of oxygen with fluorine, (b) the rate performance, (c) SEM with EDX map showing the small particles, and (d) extended cycling showing the rapid capacity fade. Adapted with permission from ref. 340. Copyright (2018) Royal Society of Chemistry.

properties are poor and as such the materials must be made with nanoscale size particles (Fig. 26c) in order to shorten the diffusion lengths, and they must be coated with carbon to increase the electronic conductivity. Both these approaches have proved very successful in cathodes operating well within the stability window of the electrolytes (e.g. LiFePO_4 to be discussed in the next section), but the DRX materials are typically cycled to 4.6 V or higher. The high surface area of the small particles with carbon present to catalyze reactions yields a high rate of electrolyte degradation. This in part leads to the poor long-term cycling seen in Fig. 26d. Another concern with the DRX materials is the large hysteresis seen during cycling (difference between charge and discharge voltages, shown as a black double-ended arrow in Fig. 26a). This is often on the order of 1 V; since the average voltage is on the order of 3.1 V³⁴¹ this hysteresis results in an inefficiency of about 33% in the battery simply based on voltage. Commercial Li-ion batteries are about 90% efficient and the cathode typically has a hysteresis around 1% or less (e.g. 1% in Fig. 24 up to 4.4 V). The high efficiency is one of the strong selling points of Li-ion batteries over other energy storage/conversion systems such as hydrogen fuel cells. A large hysteresis implies a much larger strain on the electric grid in order to obtain the same battery performance. It should be noted that the Li-rich oxides also have a larger hysteresis but this is considerably mitigated compared to DRX (0.25–0.4 V is seen in Fig. 24 up to 4.8 V). Also, similarly to Li-rich layered oxides, the DRX materials show voltage fade as shown in Fig. 27 where it is about 20 mV per cycle. This fade is larger than most of the Li-rich oxides discussed previously. Great care must now be taken to

track these key metrics as more studies are performed on the DRX class of cathodes.

Despite these challenges, the DRX materials are promising given their very high capacities and they also are of high interest given that they can accommodate a wide variety of substituents.³⁴¹ Recently, DRX materials were demonstrated with both 5TMs and 11TMs, implying a very rich chemistry of possible substitutions.³⁴¹ This class of materials is still quite early in its development, and it is now essential to track the key metrics limiting its use (hysteresis, voltage fade, and capacity fade) as substitutions are made. In the case of the 6 cation substitution, the impact on these metrics was mitigated, but there was a marked improvement in the rate performance suggestive that the transport properties were improved.³⁴¹ Despite this, the hysteresis remains high (0.9–1.3 V during cycling and 0.4 V during relaxation measurements, GITT). The GITT measurements show that a significant portion of the hysteresis is intrinsic to the material and cannot be eliminated by cycling more slowly. We speculate that this is due to the disordered 3D percolation paths, which one would expect to have a higher intrinsic resistance to Li transport than do the ordered 2D sheets in traditional layered oxides.

It is too early to discuss systematic studies looking at the impact of substitutions into DRX materials as there are no clear compositions that are being optimized. But there are a number of studies looking at the performance of various DRX materials and they are summarized in Table 2. A more complete review³⁴² has been published recently and clearly demonstrates that to date the emphasis has been solely on maximizing the discharge capacity and improving the lithium conductivity, with few



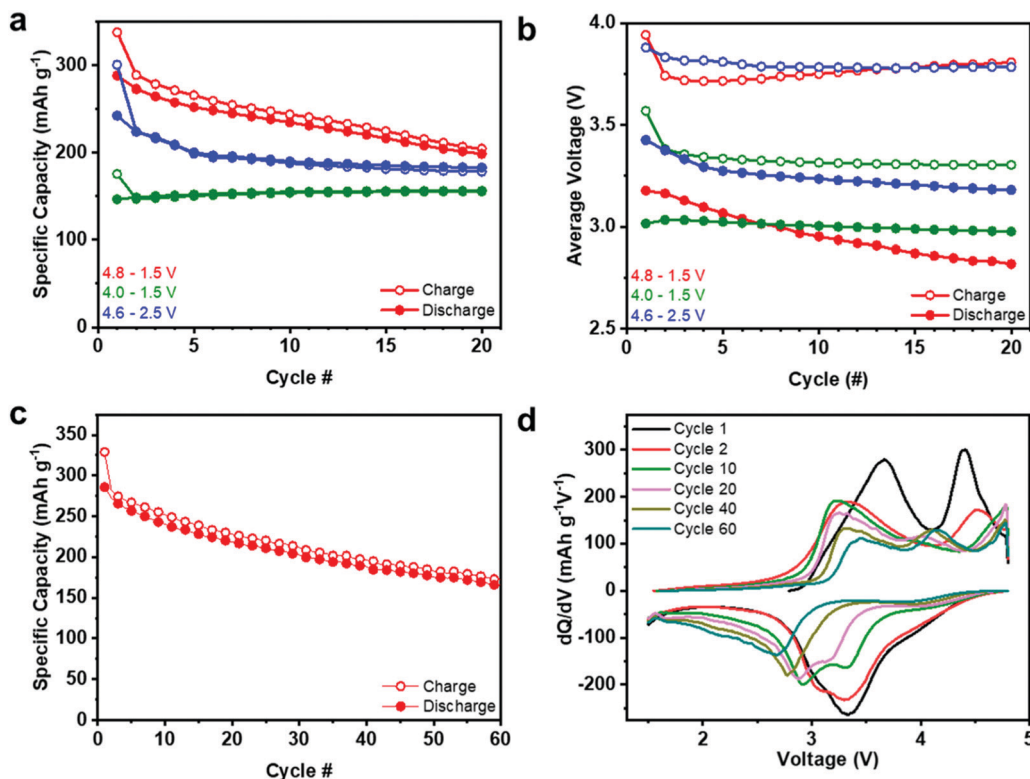


Fig. 27 Long-term cycling showing specific capacity (a), average voltage (b), specific capacity over extended cycling (c), and dQ/dV for a Li-Mn-Nb-O-F disordered rocksalt material. The voltage fade in (b) is approximately 20 mV per cycle over the full voltage window. Figure reproduced with permission from ref. 346. Copyright (2020) American Chemical Society.

Table 2 Electrochemical properties of disordered rocksalt materials: the first discharge capacity (Q_1), the first cycle irreversible capacity (ΔQ_1), the capacity after N cycles (Q_N), the voltage window used (V_{\min} to V_{\max}), the current used (I), the hysteresis (extracted at both high and low voltage as shown in Fig. 26a), and the capacity fade per cycle as a percent of the first discharge capacity ($\Delta Q/\Delta N$). When average voltage is reported, these are used for hysteresis calculation^a

Formula	Q_1 (mA h g ⁻¹)	ΔQ_1 (mA h g ⁻¹)	Q_N (mA h g ⁻¹)	N	V_{\min} (V)	V_{\max} (V)	I (mA g ⁻¹)	Hysteresis (HV-LV)	$\Delta Q/\Delta N$ (%)	Ref.
Li _{1.18} Fe _{0.34} Ti _{0.45} O ₂	223	41	142	50	1.5	5	10	1.6–0.8	0.73	347
Li _{1.17} Mn _{0.34} V _{0.49} O _{1.8} F _{0.2}	290	23	175	20	1.5	4.6	10	0.8–1.1	1.98	340
Li ₂ V _{0.5} Ti _{0.5} O ₂ F	280	–45	190	50	1.3	4.1	66	0.59 (2.31) ^a	0.64	348
Li ₂ VO ₂ F	330	–10	130	50	1.3	4.1	66	0.51 (2.53) ^a	1.21	348
Li _{1.2} Ni _{0.3} Ti _{0.3} Nb _{0.2} O ₂	221	44	180	50	1.5	4.5	40	0.4–2.2	0.37	349
Li _{1.25} Nb _{0.25} Mn _{0.5} O ₂	200	125	150	26	1.5	4.8	6.4	0.9–0.5	0.96	350
Li _{1.3} Ta _{0.3} Mn _{0.4} O ₂	248	77	85	30	1.5	4.8	10	0.9–1.5	2.19	351
Li _{1.2} Mn _{5/8} Nb _{7/40} O _{1.95} F _{0.05}	330	20	200	20	1.5	4.8	15	0.75 (3.2 V) ^a	1.97	346
LiNi _{0.5} V _{0.5} O ₂	264	10	160	50	1.3	4.5	10	0.8–1.0	0.79	339
Li ₂ Mn _{2/3} Nb _{1/3} O ₂ F	238	22	205	25	1.5	4.6	20	0.5–1.0	0.55	352
Li _{1.3} Nb _{0.3} Mn _{0.4} O ₂	300	20	180	20	1.5	4.8	25	1.5–2.5	2.00	353

^a Hysteresis is the difference between average charge and discharge voltage. In brackets is the average discharge voltage.

efforts concerned with the other metrics. The two main metrics of concern in Table 2 are the hysteresis that is high for all materials, and the poor long-term cycling. The long-term cycling is extremely poor, but this is in no small part due to the very large voltage window used, which is well beyond the stability window of the carbonate-based electrolytes and, again, this becomes exaggerated by the high surface area cathodes. As a result, it is impossible to conclude whether any improvements seen are due to substitutions in the cathodes or variants of electrolyte formulation between different articles. This provides

an opportunity to develop this literature further to see if there is real potential for commercialization. A smaller window of 3–4.6 V would be of high interest as liquid electrolytes react with oxygen gas evolved from the cathode at high voltages in both Li-air batteries and in Li-rich layered oxides, so avoiding the <3 V region may be informative and useful.^{295,343–345} We therefore consider it impossible to determine the impact of the cathode on long-term cycling when it is operating in liquid electrolytes up to voltages above 4.4 V where the electrolytes are highly reactive, especially when the materials have extremely



high surface areas due to nanometric particle sizes. This is therefore the first class of cathodes discussed herein that necessitates improved electrolytes that are stable at higher voltages than the current carbonate electrolytes, and we refer the reader to part II of this review article to discuss solid electrolytes that have the potential to make such cathodes viable.

Despite the limitation of needing a high-voltage electrolyte, continued studies on the impact of substitutions into DRX are expected to be plentiful given the wide compositional freedom this class of materials permits. Most importantly, the mitigation of the large overpotential will be essential to make these materials competitive with the state-of-the-art layered oxide cathodes. Focus should be directed to mitigating this property, with the hope that progress in the electrolytes will be made co-incidentally.

6. High voltage spinel: $\text{LiNi}_{1.5}\text{Ni}_{0.5}\text{O}_4$ spinel

The final class of oxide cathodes that will be considered here are spinel structures. Although low-voltage spinels have been commercialized, there is little progress to be made. We focus here on high-voltage spinels which continue to be quite promising, though facing the same limitations with electrolyte stabilities as other cathodes operating about 4.3 V vs. Li.

$\text{LiNi}_{0.5}\text{Mn}_{1.5}\text{O}_4$ (LNMO) is another heavily-studied high-voltage cathode material with an operational voltage of around 4.7 V. It crystallizes into two different cubic structures (Fig. 28), with the formation of the phases is highly dependent on temperature. The ordered $P4_332$ structure is formed near 700 °C and the disordered $Fd\bar{3}m$ structure is formed at higher temperatures.³⁵⁴ Generally, the disordered structure demonstrates higher stability and energy density. Like other high-voltage cathodes, the nominal voltage of LNMO is outside the window of stability of contemporary electrolytes restricting its practical application. At this voltage range, the material reacts with the electrolyte and significant Mn dissolution occurs.^{355,356} Many protective coatings have been tried to reduce the degree of reaction with the electrolyte.³⁵⁷ However, surface reactions and Mn dissolution are not the only mechanisms of failure for Mn cathodes. Mn^{3+} is Jahn–Teller active and can

lead to the transformation of tetragonal to cubic phases within the material. This increases the volume change during cycling and causes irreversible phase transitions that degrade the structural integrity of the particle.³⁵⁸ Partial metal substitution has been implemented to improve the structural stability of the bulk material.

LNMO also displays poor rate capability due to its relatively poor electronic and ionic conductivity.³⁵⁹ Research has focused on controlling particle morphology to address these issues by reducing the length of diffusion pathways. However, it remains critical that the surface be protected when synthesizing smaller particles, as the increased surface area results in more rapid Mn dissolution.

Many metals have been substituted into LNMO to stabilize the structure and improve electrochemical performance. Mg, for instance, has been shown to increase the electronic conductivity while stabilizing the structure, improving rate performance and reducing capacity fade during long-term cycling.³⁶¹ Na has also been shown to decrease particle size and cation ordering, resulting in a more conductive and more rate-performant material with a higher capacity.³⁶² Al substitution for either Ni or Mn has been shown to improve the electrochemical properties of LNMO. Substitution for Mn reduces capacity, but increase stability, while substitution for Ni has a higher capacity but also a larger capacity fade. It was shown that co-substitution was the optimal approach in the case of Al.^{363,364} With Cr substitution, LNMO can retain more than 82% capacity after 1000 cycles at 1C at an initial discharge capacity of 130 mA h g^{−1} in a traditional LiPF_6 EC:DMC electrolyte (Fig. 29).³⁶⁴ Thus, it appears that metal substitution is an effective method for improving the electrochemical properties of LNMO spinel cathode materials and may prove critical to their realization as commercial materials. Though they could still benefit from improved electrolyte stability and further structural stabilization.

7. Polyanion cathode materials

7.1 LiFePO_4

Polyanion cathodes are another broad class of intensely studied cathode materials. Of the polyanions, phosphates are the most

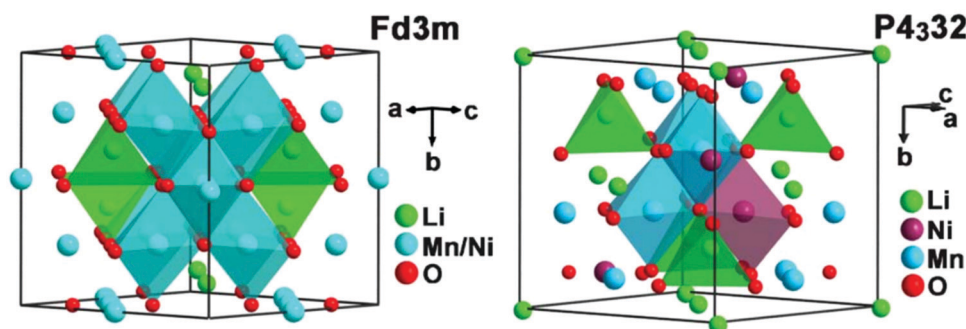


Fig. 28 Ordered (left) and disordered (right) structure of LNMO. Reproduced with permission from ref. 360. Copyright (2012) Royal Society of Chemistry.



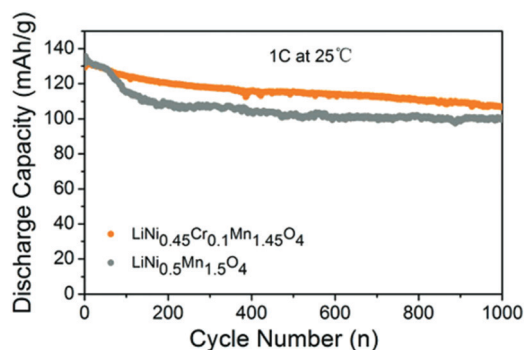


Fig. 29 Long-term cycling of Cr-substituted LNMO. Reproduced with permission from ref. 364. Copyright (2018) Wiley.

common. LiFePO_4 (LFP) was first published by Goodenough in 1997 and entered commercial use shortly after.³⁶⁵ Typically synthesized mechanochemically or through carbothermal reduction at mass-production scale, but also by co-precipitation and solid-state reactions among other techniques,³⁶⁶ LFP has an olivine structure (Fig. 30) and can be represented in the $Pnma$ space group (Fig. 30).³⁶⁷ Lithium intercalates and deintercalates through 1D channels along the c -axis. This 3D interlaced structure gives LFP a high thermal stability. It is stable up to 350 °C in N_2 atmosphere³⁶⁵ due to the strong bonding of oxygen to both Fe and P compared to layered oxide materials where the oxygen is only strongly bound to the transition metal. A high thermal stability also results in a safer battery with less risk of thermal runaway.³⁶⁸ Upon cycling, the volume change of LFP is less than 7%,³⁶⁶ thus there is low strain on the material during cycling contributing to its long cycle life. The structure of LFP, while stable, does not lend itself well to cathode performance as it has intrinsically low electrical conductivity and reduced Li^+ conductivity compared to layered cathode materials.³⁶⁹ These deficiencies combine to result

in poor rate performance. An effective strategy to solve this problem is applying conductive coatings to the active material while nanosizing the particles.³⁷⁰ While many different conductive coatings have been explored, carbon coatings have become the standard. Carbon coating has been shown to increase the conductivity from $10^{-9} \text{ S cm}^{-1}$ to $10^{-5} \text{ S cm}^{-1}$, within an order of magnitude of commercial layered oxides.³⁷¹ The poor Li transport properties of LFP have been somewhat addressed synthetically by tailoring synthesis to yield smaller particles, reducing the length of Li diffusion pathways.³⁷² The dramatic improvement in cycling due to nanosizing and carbon coating is shown in Fig. 30. Therefore, through synthetic optimization, contemporary LFP has a capacity near its theoretical capacity of 170 mA g^{-1} and can cycle for thousands of cycles at 1C while retaining over 95% capacity.³⁷³ However, LFP has a low operating voltage of 3.45 V vs. Li and thus a relatively low energy density of 590 W h kg^{-1} .³⁷⁴

7.2 Substitution into LFP: LMP (M = Mn, Co, Ni)

In addition to coating, many metals have been substituted into LFP in an attempt to increase capacity, operating voltage, and rate capability. However, there has been limited success with some claims of increased electronic conductivity, and other attributing increased conductivity to impurities formed on the surface. LFP substituted with Mg has shown a decrease in capacity, but better rate performance. However, this is likely attributable to the smaller particle size of the Mg-substituted material.³⁷⁵ Vanadium and nickel, for example, have been co-substituted into LFP and increased the capacity by up to 15% at a 1C rate and at a rate of 10C maintains 75% of the 0.1C capacity, demonstrating good rate performance.³⁷⁶ Again, the particles of the substituted material are a smaller and more uniform particle size, which may partially explain the increase in rate capability. There have been reports of metal substitution

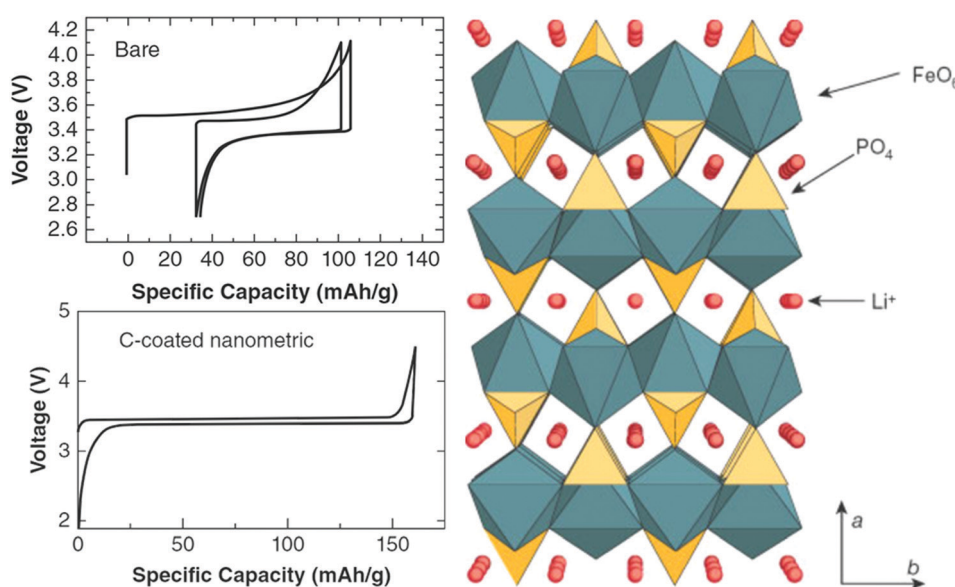


Fig. 30 Charge/discharge curves of bare and C/coated LFP (left). Olivine structure of LFP (right). Reproduced with permission from ref. 127 and 367. Copyright (2017) Wiley and copyright (2001) Nature.



increasing electronic conductivity. For instance, Mn substituted into Fe sites has been shown to decrease charge transfer resistance and increase Li diffusion. However, the substituted material showed increased capacity fade, implying a reduced structural stability.³⁷⁷ Ni substitution has been shown to achieve similar results, increasing capacity and rate capability of unsubstituted LFP, but at the cost of reduced cycle life.³⁷⁸ Improved conductivity through bulk substitution was called into question when Nazar *et al.* explained increased conductivity through the formation of connected carbon networks and conductive metal phosphides impurities at the surface of the particles.³⁷⁹ However, Molenda *et al.*³⁸⁰ observed small increases in conductivity with Al^{3+} , Zr^{4+} , and W^{6+} substitution of 25% atom and recently, a small increase in conductivity (less than one order of magnitude) was observed by gadolinium substitution into LFP.³⁸¹ Overall, there is evidence that partial metal substitutions can effectively enhance the electrochemical performance of LFP, specifically increasing the reversible capacity and rate performance, but often at the cost of reduced long-term stability and the particles must be coated with a conductive material to cycle effectively. Complete metal substitution has been shown to yield materials of high theoretical energy densities. Substituting Fe for another metal entirely, such as Mn, Co, or Ni, gives materials with much higher operating voltages (Fig. 31) with similar theoretical capacities, and thus higher theoretical energy densities. However, these materials come with their own challenges. Changing the Fe for a different metal has led to higher voltage materials that maintain the stable olivine structure. Mn, Co, and Ni phosphate cathodes have been made with varying degrees of success.

LiMnPO_4 (LMP) exhibits an operating voltage of 4.1 V, giving it a higher theoretical energy (679 W h kg^{-1}) density than LFP.³⁸³ However, the cyclability of LMP is far inferior to that of LFP due to instability caused by the presence of Mn. As LMP is delithiated, Mn^{2+} is oxidized to Mn^{3+} which is Jahn–Teller active and causes distortion in the lattice which leads to the formation of defect phases and reduced Li diffusivity. Another deficiency with

pristine LMP is electrical conductivity. LMP has an electronic conductivity significantly lower than that of LFP, making it impractical without modification.³⁸⁴ Like LFP, conductive carbon coating can improve the conductivity of LMP. However, carbon coating on LMP proves to be less stable than that on LFP. Many metals have been substituted into LMP in an attempt to improve its electrochemical properties. Cr, Fe, Mg, V and Na have all been shown to improve the reversible capacity of LMP (Fig. 34) and significantly improve structural stability. Mixed results have been found using Zn as a substituent.³⁸⁵ This is attributed to differences in synthetic conditions, which have a large influence on the effectiveness of metal substitution, with Fang and Wang³⁸⁶ reporting improved electrochemical performance using a sol-gel and solid-state methods respectfully, while Wang *et al.*³⁸⁶ reported decreased performance using a hydrothermal technique. There is also evidence that partial substitution of Na for Li can improve the charge transfer properties and lithium diffusion coefficient, likely due to the expansion of Li diffusion channels caused by the large Na ions.³⁸⁷ DFT calculations have supported transition metal substitution at the Mn site increasing electronic conductivity by narrowing the band gap and increasing structural stability through stronger metal–oxygen bonding.³⁸⁸ Recently, Mg substitution of Mn has been shown to achieve 400 cycles with 100% capacity retention at 1C, showing that LMP may have potential as a commercial high-voltage cathode (Fig. 32).³⁸⁹

LiCoPO_4 (LCP) is an attractive alternative to LFP and LMP due to its high operating voltage of 4.8 V and theoretical capacity near 167 mA h g^{-1} , yielding a high theoretical energy density of $\sim 800 \text{ W h kg}^{-1}$ due to the higher operating potential than any commercialized cathode.³⁹⁰ While the theoretical energy of these materials is high, in practice they are plagued by a lack of stability and aggressive electrolyte decomposition due to the high operating potentials. This complicates studying these materials, as there is significant electrolyte degradation upon repeated cycling.³⁹¹ Despite this severe electrolyte degradation, a significant amount of work has been conducted on these materials to optimize their capacities and better stabilize the interface such that they may be studied while high-voltage electrolyte materials are developed. Like LMP, LCP exhibits much lower electrical conductivity than LFP making and thus requires a conductive coating and reduced particle size to cycle effectively.³⁹² LTO has been shown to significantly prevent electrolyte decomposition with a capacity retention of 93% after 100 cycles when coated on unsubstituted LCP. Other than coating, metal substitutions have been shown to increase the reversible capacity of LCP, these include Fe, V, Y, and Ca.^{393–396} However, these substituents do not significantly address the cyclability issues of the material. Fe substitution has been shown to decrease antisite defects during cycling, decreasing structural instability and increasing the cycle life of the material.³⁹³ It is also clear that synthesis technique is critical in the effect of substituent metals, as Di Lecce *et al.* report a decrease in electrochemical performance and less spherical particle morphology using a higher annealing temperature.³⁹⁷ Thus far, these substitutions do not address the cyclability

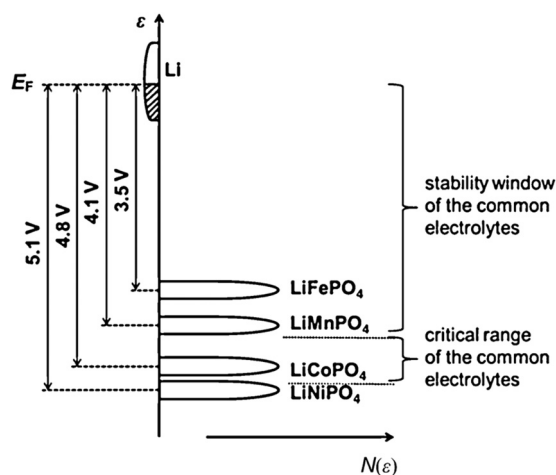


Fig. 31 Relative energies of LiMPO_4 cathode materials vs. Li/Li^+ . Figure reproduced with permission of ref. 382. Copyright (2014) Springer.



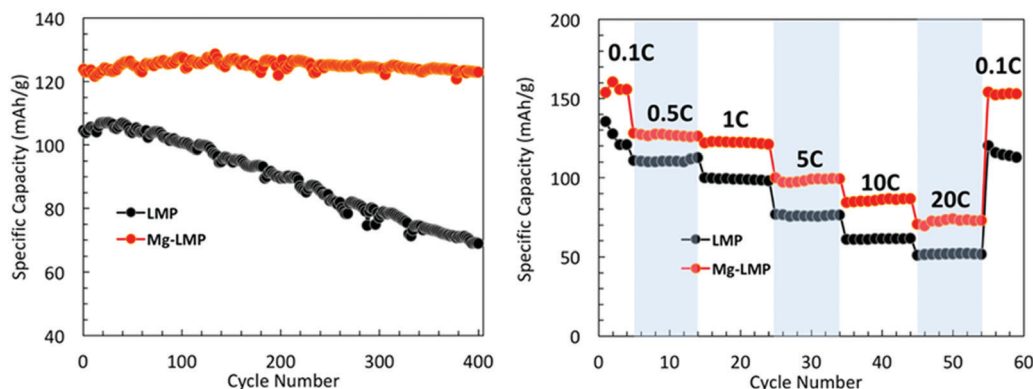


Fig. 32 Long-term cycling (left) and rate performance of Mn-substituted LMP (right). Figure reproduced with the permission of ref. 389. Copyright (2018) American Chemical Society.³⁸⁹

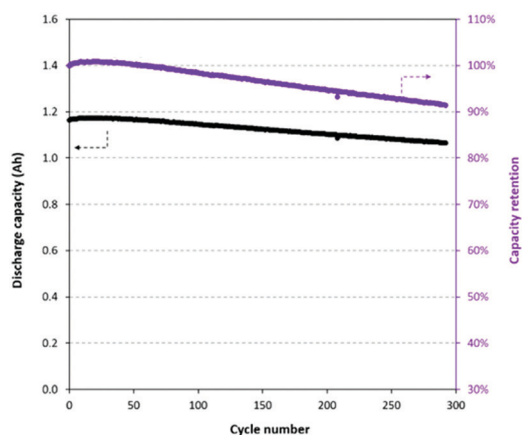


Fig. 33 Capacity retention of Fe, Cr, Si co-substituted LCP full cell in LiFSi/Py₁₃FSI electrolyte cycled at 0.2C from 3.5 to 4.9 V. Reproduced with permission from ref. 398. Copyright (2020) Multidisciplinary Digital Publishing Institute.

issue with LCP. Recently, an impressive multi-substituted material was achieved by Liu *et al.* where Fe, Cr, and Si were co-substituted into LCP. This material maintained a >90% capacity after 290 cycles at a 0.2C rate with a LiFSi/Py₁₃FSI electrolyte (Fig. 33).³⁹⁸ While the initial reversible capacity was only 120, the energy was over 570 W h kg⁻¹. Progress on LCP as a high-energy cathode is promising, but there is still much work to be done on electrolyte/cathode interface stability before it may be used commercially. Additionally, LCP has a substantial drawback that cannot be solved: the high cobalt content. Like layered oxides, reducing Co content in cathode materials is important, due to its price, environmental impact and socio-economic consequences, as discussed previously in this article.

LiNiPO₄ (LNP) is an alternative to LFP with a nominal voltage higher than that of LCP at 5.1 V. This has made the material impractical in virtually all common liquid electrolytes and has made study of the material extremely difficult due to electrolyte decomposition. It has a theoretical capacity of 170 mA h g⁻¹ and energy of over 860 W h kg⁻¹.⁴⁰⁹ Like the other phosphates, conductive surface coatings are necessary to

increase conductivity and is necessary to attempt to reduce electrolyte degradation. In addition, it is difficult to synthesize LNP without impurity phases, or significant antisite defects.³⁸² Partial metal substitution has been used to improve the quality of the synthesized material, reducing antisite defect, where Ni occupies Li sites, and increasing overall structural stability. In practice, many synthesized LNP cathodes with coating and metal substitution break 100 mA h g⁻¹. Mg²⁺ substitution has been shown to increase structural stability,⁴¹⁰ while Nd³⁺ substitution was able to achieve a capacity of 95 mA h g⁻¹ at 0.1C and retained 85% capacity after 50 cycles at 2C.⁴¹¹ Recently, Co-substituted LNP nanosheets were able to be cycled for 200 cycles at 0.1C with 98% capacity retention and an initial reversible capacity of 133 mA h g⁻¹ at that rate using an LNP/Li₄Ti₅O₁₂ full cell with solid polymer electrolyte.⁴¹² Thus, it appears through structural control and metal substitution that LNP may be realized as a high-energy cathode material if combined with a highly stable solid electrolyte. There may therefore be a future for LNO based batteries, but this relies on the development of electrolytes.

There are other high-voltage cathode compositions such as some pyrophosphates that may show promise, but have not yet been demonstrated appreciable cyclability at high voltage in full cells as LCP has. A range of pyrophosphate cathodes and other high voltage cathodes have been summarized elsewhere^{413,414}

High-voltage lithium metal phosphate cathode materials do show promise for future battery technology if their structural and electronic challenges can be overcome. Metal substitutions can certainly be used to improve material stability and also improve both the electronic and ionic conductivities. Though much work remains to be done, as many elements have not yet been attempted in published partial metal substitutions. Additionally, there appears to be benefit in multi-element substitution as shown in LCP, given the multiple properties that require improvement.³⁹⁸ Multi-element substitutions are poorly explored due to the increased dimensionality of the composition space. However, novel high-throughput experimental techniques pave the way for meaningful screening and optimization of multi-metal substitutions.¹⁴⁹



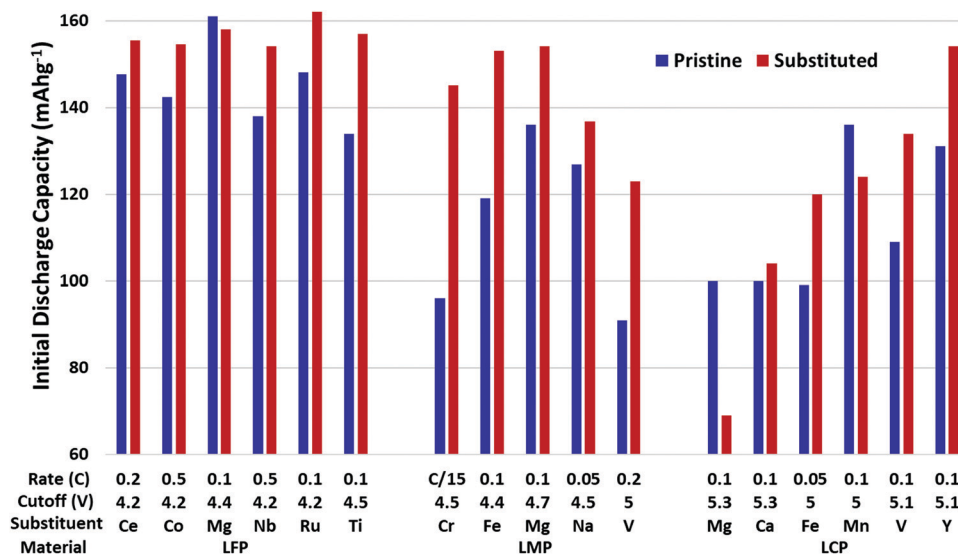


Fig. 34 Discharge capacities of substituted phosphate cathodes. Data collected from ref. 375, 387, 389, 393, 395, 396 and 399–408.

8. Conclusions and outlook

Cathodes for Li-ion batteries are diverse with many compositions having been studied under the broad classes of layered oxides, olivine polyanion, spinel, and disordered rocksalt, each with deficiencies that can be improved. Stoichiometric layered oxides stand to improve their accessible capacity while reducing the cobalt content. This can be achieved by substituting Ni for Co. However, this introduces structural instability as Li/Ni mixing, surface reactions, and structural degradation become more prevalent. Partial metal substitution can reduce the severity of these problems; either acting to stabilize the structure through stronger M–O bonds or by acting as “pillars” to prevent Mn migration and layer collapse. This has the added benefit of increasing rate capability in many cases. Compared to coatings, it seems partial metal substitution is similarly effective on average, though coatings often come at the cost of rate performance. Thus far, it appears neither single metal substitutions, nor coatings alone, are “miracle” solutions for the problems facing these materials, but they do yield significant improvements.

In Li-rich cathodes (both ordered and disordered rocksalt), structural stability is an even larger problem, plaguing the materials with a high degree of voltage fade and capacity loss. Metal substitutions have been shown to be an effective strategy to help mitigate both these issues. Like in their stoichiometric counterparts, partial metal substitution may occur on TM sites or Li sites. The former improving structural stability through stronger bonding, while the latter improves structural stability by increasing the interlayer distance and discouraging metal migration through repulsion. These electrochemically inactive metal substitutions often come with the cost of reduced capacity, a larger reduction than in stoichiometric layered materials, as they can suppress the activation of Li-rich cathode materials. Coatings are also an effective strategy to increase cycle life, but in general are not as effective at suppressing voltage fade as

metal substitution. However, substitutions and coatings are not exclusive, they can be used in conjunction to maximize performance. Further work in Li-rich oxides is sorely needed to make them viable as next-generation cathodes.

Finally, for high-voltage materials, electrolyte reactions at the material interface are a large issue, obscuring the origins of capacity during cycling. Despite this, it is clear that these materials can benefit from improved structural and interfacial stability. Partial metal substitutions are effective, especially in the case of LiCoPO_4 where Fe, Cr, Si co-substitution has been shown to yield a cathode that has the potential to make a marked leap in Li-ion energy density.

Metal substitutions undeniably have a role to play in the development of next-generation cathode materials. While many single element substitutions have been explored, the next steps are to investigate and optimize multi-element substitutions in layered and polyanion cathode materials. Currently, multi-element substitutions are poorly explored largely due to the experimental demand of adding extra dimensionality to the composition space. This limitation can be remedied, in part, using novel high-throughput experimental techniques to screen many perspective materials simultaneously. Combined with computational modelling to better direct experimental efforts, materials discovery can be significantly accelerated. Additionally, there are many elements that have not been substituted into the materials here that should be investigated for their potential in improving battery performance, such as many of the lanthanides as well as heavier d-block elements. As shown earlier in this work, there is potential for these d- and f-block elements to positively impact material performance. Thus, there is worth exploring their impact across all classes of cathode discussed in this work. Concerning high-voltage materials though, the stability window of the electrolyte becomes problematic in investigating the effects of metal substitution in detail. Thus, it is imperative that electrolytes with a larger



stability window be developed to significantly move development of high voltage cathodes forward. Solid electrolytes are promising in this respect, and these will be reviewed in part II of this review.

Conflicts of interest

There are no conflicts to declare.

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