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1. Introduction

Over last two decades, there has been a great development in the field of localized surface plasmon resonance (LSPR). LSPR can be defined as nanoparticles (NPs) much smaller than the wavelength of the incident light in size, exhibiting collective oscillation of free carriers when the frequency of electromagnetic wave matches the frequency of free carriers oscillating

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against the restoring force of oppositely-charged carriers.¹ Noble metal NPs such as Au and Ag are the most systematically exploited plasmonic materials and mainly exhibit their LSPRs in the visible region.² Recently, it has been disclosed that LSPR emerges not only in metals but also in heavily doped semiconductors with appreciable free carrier density, including aliovalent doped/self-doped metal oxides, metal chalcogenides, metal nitrides and many others.³ For example, tin doped indium oxide (ITO),⁴ aluminum doped zinc oxide (AZO),⁵ niobium doped titanium oxide (NTO),⁶ oxygen-deficient molybdenum oxide (MoO_{3-x})^{7,8} and tungsten oxide (WO_{3-x}),⁹ titanium nitride (Ti_xN)¹⁰ as well as nonstoichiometric copper chalcogenide (Cu_{2-x}E, E = S, Se and Te)¹¹⁻¹⁸ NPs exhibit intense LSPR. In particular, copper vacancy (hole) doped Cu_{2-x}S NPs have attracted numerous attention



since the pioneering work by Zhao *et al.* in 2009, where the oscillation of holes generated in valence band results in strong absorption and scattering of incident light.¹⁹ Because of lower order of magnitude of carrier density ($N_h \approx 10^{20}\text{--}10^{22} \text{ cm}^{-3}$) with respect to noble metals ($N_e \approx 10^{23} \text{ cm}^{-3}$), Cu_{2-x}S NPs support LSPR responses in near-infrared (NIR) and mid-infrared (MIR) regions.^{20–23} Unlike noble metals with restricted carrier density, once the size, morphology and the surrounding medium is fixed, the LSPR response is locked and cannot be flexibly adjusted. In Cu_{2-x}S NPs, even a minor variation of carrier density has a strong effect on the spectral position and strength of the LSPR. This allows for a controllable LSPR tuning over a wide spectral range, either during the synthesis or post-treatment, making the Cu_{2-x}S NPs very fascinating in a manifold of research fields.^{24–30}

In addition to variable carrier density, another advantage of Cu_{2-x}S over noble metals is the existence of various crystal structures and stoichiometries, varying from the most copper-rich high/low chalcocite Cu_2S to the most copper-poor covellite CuS .³¹ Due to the absence of any free holes, the full stoichiometric Cu_2S is LSPR inactive, while the nonstoichiometric Cu_{2-x}S are all in LSPR-active state.^{11,19} The LSPRs of Cu_{2-x}S NPs are highly dependent on their crystal structures. As previously demonstrated, djurleite $\text{Cu}_{1.94}\text{S}$ and digenite $\text{Cu}_{1.8}\text{S}$ nanodisks

(NDs) show two distinct LSPR oscillation peaks assigned to out-of-plane mode and in-plane mode in the NIR and MIR regions, respectively, while roxbyite $\text{Cu}_{1.8}\text{S}$ NDs and covellite CuS NDs only exhibit one oscillation peak assigned to in-plane mode in NIR region.^{32–34} Besides the crystal structure, another crucial parameter is the morphology. Indeed, morphology dependent LSPR has been demonstrated in Cu_{2-x}S nanospheres, nanodisks and nanorods.^{32,35} Thanks to the tunable LSPR, the Cu_{2-x}S NPs have been well exploited in photothermal therapy, photocatalysis, sensing, switchable magnetic probe, two-photon upconversion and many others.^{3,24}

Herein, we present a review on the emerging Cu_{2-x}S NPs, particularly their optical responses and LSPR related applications. In the first section, we start the introduction of crystal structures and stoichiometries of Cu-S system. Following this subsection is a discussion on the hole generation in the valence band of intrinsically unstable Cu_2S . Then, we introduce an optical modeling of Cu_{2-x}S for better understanding of their LSPR and it is followed by vigorous discussion on the flexible manipulation of LSPR in Cu_{2-x}S *via* hole density tuning, crystal structure control and transformation as well as morphology control. Furthermore, we briefly discuss the relationship between LSPR and photoluminescence of Cu_{2-x}S . In the second section, we introduce recently well exploited plasmonic applications including photoacoustic imaging and photothermal therapy, photocatalytic organic synthesis, photo(electro)catalytic water splitting, two-photon upconversion as well as SERS probe endowed by the LSPR of Cu_{2-x}S . In the third section, we give a short summary of this review and we emphasize some important open questions relevant to the Cu_{2-x}S synthesis, chemistries, physics and industrial applications.

2. Optical responses

Since the optical responses of Cu_{2-x}S NPs are determined not only by their crystal structure and stoichiometry but also by the morphology. In this section, we firstly introduce the crystal structures and Cu-S system of copper sulfide, then we give a brief discussion on the generation of copper vacancies in the



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valence band of fully stoichiometric Cu_2S . After that, an optical modeling for Cu_{2-x}S to better understand the origin of the LSPR is provided and it is followed by an elaborate discussion on the flexible manipulation of LSPR in Cu_{2-x}S NPs *via* hole density tuning, crystal structure control and transformation as well as morphology control. Furthermore, the relationship between LSPR and photoluminescence of Cu_{2-x}S is discussed.

2.1 Crystal structures and stoichiometries of Cu-S system

In Cu-S system, the sulfur 3s and 3p orbitals can strongly hybridize with copper 3p and 4s orbitals, resulting in more covalent character. Moreover, the low size and electronegativity differences between sulfur and copper, the ability to form S-S bonds and the probability of Cu-Cu bonds within the structure, allow for an abundance of crystal structures and stoichiometries in copper sulfide, varying from most copper-sufficient chalcocite $\text{Cu}_{2.0-1.94}\text{S}$, djurleite $\text{Cu}_{1.97-1.94}\text{S}$, to roxbyite $\text{Cu}_{1.78-1.81}\text{S}$, digenite $\text{Cu}_{1.8}\text{S}$ and anilite $\text{Cu}_{1.75}\text{S}$, and to most copper-deficient covellite CuS (Table 1).^{3,31} The chalcocite phase exists in three different forms at different temperatures: low chalcocite phase is formed at $T > 104\text{ }^\circ\text{C}$, high chalcocite phase is at $104\text{ }^\circ\text{C} < T < 436\text{ }^\circ\text{C}$ and cubic chalcocite phase is at $T \geq 436\text{ }^\circ\text{C}$. Among these three structures, low chalcocite phase is the most stable one at room temperature.³⁶ However, at or above $105\text{ }^\circ\text{C}$, the low chalcocite phase transforms into high chalcocite phase, which is a solid-liquid hybrid phase.³⁷ Djurleite phase is similar to low chalcocite phase, because of its monoclinic structure with a slightly decreased copper content. Both the djurleite and low chalcocite phase are very common species and are usually coexisted.^{38,39} The djurleite phase was initially discovered by Djurle and later was established as the mineral djurleite in his honor.^{40,41} The unit cell content of djurleite is $\text{Cu}_{62}\text{S}_{32}$ ($\text{Cu}_{1.94}\text{S}$) with 62 copper atoms in the structure, where 52 copper atoms are in a triangular coordination with sulfur, 9 copper atoms are in a tetrahedral coordination, and the last one is in a linear coordination.³⁹ When an extra Cu atom is added to the structure, the unit cell content becomes $\text{Cu}_{63}\text{S}_{32}$ ($\text{Cu}_{1.97}\text{S}$), exhibiting the higher and lower limit of $\text{Cu}_{1.94-1.97}\text{S}$ for djurleite. Another case with lower Cu:S stoichiometric ratio (1.8:1) is digenite phase in the crystal structure. Bulk digenite $\text{Cu}_{1.8}\text{S}$ has three polymorphs depending on the temperature: high-temperature phase with face-centered cubic sulfur anion structure, which is stable at $T > 73 \pm 3\text{ }^\circ\text{C}$, low-temperature phase also has face-centered cubic symmetry of sulfur anion

Table 1 Cu-S system: mineral name, stoichiometry and their corresponding crystal structures (data extracted from ref. 3, 31 and 57)

Mineral name	Stoichiometry	S packing	Crystal structure
Low chalcocite (α)	$\text{Cu}_{2.0-1.99}\text{S}$	hcp	Monoclinic
High chalcocite (β)	$\text{Cu}_{2.0-1.94}\text{S}$	hcp	Hexagonal
Cubic chalcocite (γ)	Cu_2S	fcc	Cubic
Djurleite	$\text{Cu}_{1.97-1.94}\text{S}$	hcp	Monoclinic
Digenite	$\text{Cu}_{1.8}\text{S}$	fcc	Cubic
Roxbyite	$\text{Cu}_{1.81}\text{S}$ ($\text{Cu}_{58}\text{S}_{32}$)	hcp	Triclinic
Anilite	$\text{Cu}_{1.75}\text{S}$ (Cu_7S_4)	fcc	Orthorhombic
Covellite	CuS	hcp	Hexagonal

structure but different lattice constants and metastable rhombohedral phase obtained by cooling from $73 \pm 3\text{ }^\circ\text{C}$ resembles high chalcocite, with one copper atom located at four equivalent positions in each sulfur tetrahedron.^{42,43} These three structures are close related to each other, and therefore, phase identification of $\text{Cu}_{1.8}\text{S}$ NPs remains a challenge.⁴⁴ Roxbyite phase $\text{Cu}_{1.81}\text{S}$ ($\text{Cu}_{58}\text{S}_{32}$) is just as complex in its entirety as the structures of low chalcocite and djurleite phase.^{45,46} It has been described that, it possesses a triclinic unit cell with 90 atoms, constructing a distorted hexagonal close packing of 32 sulfur atoms with 58 individual copper atoms occupying triangular coordination with sulfur. Other layers sandwiched between the close packed sulfur layers composed of double or split layers of Cu atoms. Some of which have two-fold linear coordination, but mostly they have three-fold triangular and four-fold tetrahedral coordination.⁴⁶ The roxbyite phase is unusual in bulk forms but has been readily synthesized in NPs.^{23,47,48} Anilite $\text{Cu}_{1.75}\text{S}$ (Cu_7S_4) phase has an orthorhombic crystal system, where the sulfur atoms adopt approximately face-centered cubic arrangement and the copper atoms occupy the interstices.⁴⁹ In the orthorhombic unit cell of 28 copper atoms, 20 of which are in triangular coordination and 8 are in tetrahedral coordination with sulfur.⁵⁰ Moreover, according to the theoretical study based on DFT theory, anilite phase was shown to be the most stable structure among the known Cu_{2-x}S phases.⁵¹ The most Cu-deficient covellite CuS phase has a unique layered structure in spite of simplest stoichiometry, where its copper atoms locate in two different environments: planar trigonal CuS_3 layer are sandwiched by two tetrahedral CuS_4 layers along the *c*-axis, with the triple layers connected to adjacent triple layers through disulfide (S-S) covalent bonds.⁵² Another description of covellite CuS is that it is composed of alternating CuS_3 - CuS_3 - CuS_3 layers and S-S layers along the *c*-axis.⁵³ It is worth noting that CuS forms this unique structure, rather than another simpler available structures, is due to more energetically favorable covalently bonding between sulfur atoms, than bonding to copper atoms.⁵⁴ The unique crystallographic feature in covellite CuS phase allows for an intercalation and a reaction of alkali metal cations such as Li^+ between the S-Cu-S layers, endowing CuS NPs as a promising electrode material for Li-ion battery.⁵⁵ The above discussed crystal structures of Cu_{2-x}S can be catalogued by either asymmetric hcp (hexagonal close packing) or fcc (face-centered cubic) of sulfur anions, with copper cations occupying interstitial spaces in the sulfur lattice.⁵⁶ The crystal structures of important copper sulfide phases are shown in Fig. 1a-h. However, The Cu-S crystal systems mentioned here would not be complete without the other less discussed crystal structures such as geerite $\text{Cu}_{1.6}\text{S}$, spionkopite $\text{Cu}_{1.39}\text{S}$, yarrowite $\text{Cu}_{1.12}\text{S}$ and villamanitite CuS_2 , which are possible in the bulk forms but are somewhat unusual in NPs.⁵⁷ In contrast to electron-based noble metals, LSPR responses in Cu_{2-x}S originate from collective oscillation of free holes in valence band. Therefore, the fully stoichiometric Cu_2S is LSPR inactive due to the absence of any free holes, while the nonstoichiometric Cu_{2-x}S are all in LSPR-active state. Furthermore, the LSPR responses in Cu_{2-x}S are highly sensitive to both of the crystal structure and the Cu/S stoichiometry. Among the LSPR-active Cu_{2-x}S , it is worth noting that, the aforementioned



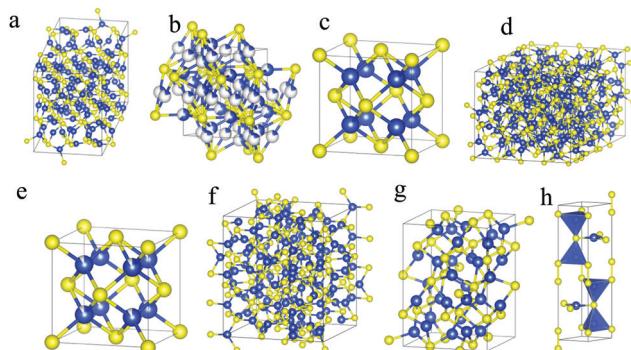


Fig. 1 Crystal structures for (a) low chalcocite, (b) high chalcocite, (c) cubic chalcocite, (d) djurleite, (e) digenite, (f) roxbyite (g) anilite, and (h) covellite phases. Blue and yellow spheres represent copper and sulfur atoms, respectively. Schematics were generated using VESTA.

covellite CuS has been demonstrated to be intrinsically metallic rather than being a heavily doped semiconductor. The S–S covalent bonding and S hypervalency (S_2^{2-} and S^{2-}) generates two free holes which are almost equally distributed among the six layers per unit cell (Cu_6S_6), leading to metallic character with hole density up to $\sim 10^{22} \text{ cm}^{-3}$.^{58,59}

2.2 Generation of holes in Cu_2S valence band

In copper sulfides, the top of valence band is essentially contributed from the sulfur 3p orbitals, and the bottom of the conduction band is mainly contributed from the copper 4s orbitals.²⁴ It is assumed that each copper atom provides one 4s electron to bonding, while each sulfur atom provides six electrons. In the fully stoichiometric Cu_2S , the valence band is completely filled and the materials would perform as intrinsic semiconductor. The copper cations in the lattice have a valence of +1, and the sulfur anions have a valence of -2. Even though the low chalcocite phase is the most stable one among the three chalcocite phases, it is prone to generate Cu vacancies. When the Cu_2S NPs are exposed to oxidative environment, electrons are extracted along with Cu^+ in order to keep electrostatic neutrality, and therefore copper vacancies (holes) are generated in the top of the valence band (Fig. 2a).⁶⁰ This hole generation mainly affects the valence of sulfur anion framework since the valence band has a major contribution from sulfur 3p orbitals, whereas the valence of copper cation inside the NP lattice keeps close to +1.²⁴ For example, $Cu^+ + (\text{copper } 4s)e^- + 1/2O_2 = CuO$, i.e., $Cu^+ + 1/2O_2 = CuO - (\text{copper } 4s)e^-$, indicates a lack of one bonding electron in sulfur 3p orbital. Since the sulfur 3p orbitals contribute to the valence band, an extraction of Cu^+ results in a lack of one bonding electron (h^+) in sulfur 3p orbital, that is, produce one hole in valence band. In order to better understand the optical properties of $Cu_{2-x}S$, it is beneficial to recall some basic aspects of the band structure. The optical bandgap for hole doped $Cu_{2-x}S$ is defined as the energy difference between the bottom of the conduction band and the highest occupied valence band. The optical bandgap is hole concentration dependent and it is larger than that of undoped Cu_2S . This optical bandgap variation by hole doping in valence

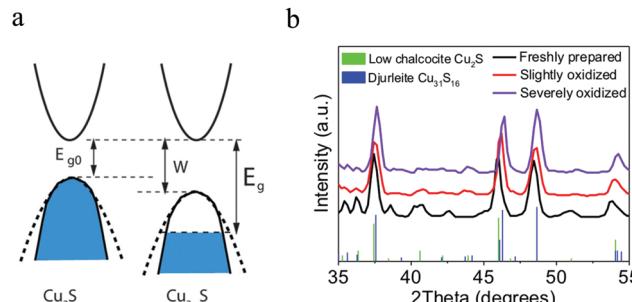


Fig. 2 (a) Band structures of Cu_2S and $Cu_{2-x}S$. In $Cu_{2-x}S$, both of the intrinsic (E_{g0}) and optical bandgaps (E_g) increased with hole generation. (Reproduced with permission from ref. 60, copyright 2014 Royal Society of Chemistry.) (b) Crystal structure transformation from low chalcocite Cu_2S to djurleite $Cu_{31}S_{16}$ ($Cu_{1.94}S$) triggered by air oxidation. In spite of very similar XRD patterns, the shift of XRD peaks to higher diffraction angles indicates the removal of copper atoms and leads to the crystal structure transformation. (Reproduced with permission from ref. 62, copyright 2018 American Chemical Society.)

band is called the Burstein–Moss effect, which has been well exploited in copper chalcogenide NPs.^{11,20} In addition to the Burstein–Moss effect in Cu_2S , another important point to address is the intrinsic instability, which is derived from its thermodynamic propensity toward a copper deficiency because of the low chemical potential of Cu^0 .⁶¹ Previous experimental investigations have found that Cu_2S is unstable at ambient conditions and rapidly transforms into hole doped $Cu_{2-x}S$ phases by air oxidation, most probably djurleite $Cu_{1.94}S$, which has been thermodynamically stable under ambient conditions (Fig. 2b).⁶² Owing to the lower crystallographic symmetry and thus lower free energy than low chalcocite Cu_2S , djurleite is more stable in bulk and in NPs.^{19,63,64}

2.3 Optical modeling

The origins of LSPR in hole-doped $Cu_{2-x}S$ NPs remain poorly understood. In order to better understand their LSPR, it is useful to introduce some general concepts with regard to optical responses in noble metals. We select the most prominent theory Drude model to reproduce and predict some of the optical responses of $Cu_{2-x}S$, and it is described by the following complex dielectric function:

$$\epsilon(\omega) = \epsilon_1(\omega) + i\epsilon_2(\omega) \quad (1)$$

where

$$\epsilon_1 = \epsilon_\infty - \omega_p^2/(\omega^2 + \Gamma^2) \quad (2)$$

and

$$\epsilon_2 = \omega_p^2 \Gamma / [\omega(\omega^2 + \Gamma^2)] \quad (3)$$

are the real and imaginary part of the dielectric function, respectively. The density of free holes in $Cu_{2-x}S$ should be large enough to achieve a negative value of ϵ_1 to support LSPR. In addition, the smaller value of ϵ_2 is demanded for better plasmonic performances or applications.^{60,65} The free carrier damping constant Γ depicts the damping caused by the scattering



with electrons, phonons, lattice defects or impurities during collective oscillations.²⁴ ε_∞ is the high frequency dielectric constant and ω is the frequency of incident light. The relationship between bulk plasma frequency ω_p and free carrier density N is shown in the following equation:

$$\omega_p^2 = (Ne^2)/(\varepsilon_0 m^*) \quad (4)$$

where m^* is the effective mass of free carrier, ε_0 is the vacuum dielectric permittivity and e is elementary charge.²⁴ As shown in eqn (4), the plasma frequency ω_p directly correlates with the carrier density N , providing a tool to estimate the carrier density of a specific sample. This indicates that an increase in carrier density gives rise to an increase of the ω_p and a blue-shift of the LSPR peak. In the case of Cu_{2-x}S , the dominate morphology is the ND. In the electro-static limit, where the NP diameter is much smaller than the wavelength ($d \ll \lambda$), the dipole polarizability (α) of a ND can be approximated employing basic scattering theory for an oblate spheroid with semi-axis $a < b = c$.³² To calculate the hole density N_h of Cu_{2-x}S NDs, the α can be expressed as:

$$\alpha_j = [3\varepsilon_0 V(\varepsilon_j - \varepsilon_m)]/[3\varepsilon_m + 3L_j(\varepsilon_j - \varepsilon_m)] \quad (5)$$

where $j = 1, 2$, or 3 , V is spheroid volume, L_j is a morphology-dependent constant calculated from the ratio of axes of the oblate spheroid with $L_1 + L_2 + L_3 = 1$ and $L_2 = L_3$. In a resonance condition, the denominator in eqn (5) must be zero:

$$3\varepsilon_m + 3L_j(\varepsilon_j - \varepsilon_m) = 0, \quad \varepsilon_{1,j} = \{ -[(1 - L_j)]/L_j \} \varepsilon_m \quad (6)$$

where $\varepsilon_{1,j}$ is the real part of dielectric function (ε_j) of the Cu_{2-x}S , ε_m is dielectric constant of the surrounding medium. In eqn (2), the term Γ can be calculated from absorption spectrum shown in the following equation:

$$\Gamma = c(\lambda_{1/2})/(\lambda_{\max}^2) \quad (7)$$

where c is the speed of light, $\lambda_{1/2}$ is the full width at half-maximum (FWHM) of the LSPR peak and λ_{\max} is the absorption maximum wavelength. It is worth noting that, because of the non-parabolicity of the valence band in Cu_{2-x}S , the effective mass of hole is highly dependent on the hole density, suggesting that one cannot extract the hole density directly from eqn (4). Occasionally, the hole density can be estimated applying eqn (4) to the LSPR of Cu_{2-x}S NDs with several different crystal structures, taking the different effective hole masses of the various Cu_{2-x}S structures into consideration.^{43,66,67} Furthermore, Kriegel *et al.* reported that the Drude model cannot sufficiently describe the experimental absorption spectrum of Cu_{2-x}S NPs directed by copper vacancies. The holes in the valence band of Cu_{2-x}S results in a deviation from the purely metallic behavior, meaning that such holes cannot be treated as fully free particles and that the effects of hole localization are important.³⁵ However, as aforementioned, the covellite CuS with highest hole density represents a special case. In an exhaustive work, Xie *et al.* investigated CuS NDs with different sizes using the discrete dipole approximation method performed within the framework of Drude–Sommerfeld model to demonstrate their morphology and size dependent LSPR.⁶⁸ The experimental spectra of several CuS NDs of various

sizes were satisfactorily reproduced by extracting the Drude parameters, indicating that the LSPR responses of CuS are consistent with the predictions by the Drude model and the holes in this crystal structure can be treated as fully free. Such treatment is likely owing to the fact that the carriers in CuS derive from the inherent metallic-like character resulting from the significant density of lattice-constitutional valence band free holes, which is quite different from the other copper sulfide structures.

2.4 Flexible LSPR manipulation of Cu_{2-x}S NPs

Except for Cu_2S , all the copper-deficient Cu_{2-x}S phases are in LSPR-active state. Besides the well-known size effect, the LSPR of Cu_{2-x}S NPs can be flexibly manipulated by several means including hole density tuning, structural control and transformation as well as morphology control according to the above optical modeling. Particularly, hole density tuning is realized through chemical redox reactions, crystal structure transformation is triggered by removal/insertion of Cu^+ and by sulfur sources, and morphology control is directed by metal ions and copper/sulfur precursor reactivity.

2.4.1 Hole density tuning *via* chemical redox reactions.

According to eqn (2)–(4) in Section 2.3, it is believed that even a minor variation of hole density in Cu_{2-x}S NPs can result in strong effect on spectral position and intensity of the LSPR. Therefore, it allows for a controllable LSPR tuning over a wide spectral range through chemical redox reactions by oxidizing/reducing agents. Oxidizing agent such as air (O_2), iodine (I_2), alkylamines as well as cerium(IV) ammonium nitrate $[(\text{NH}_4)_2\text{Ce}(\text{NO}_3)_6]$ can induce a generation of copper vacancies in the valence band of Cu_{2-x}S , leading to an increase of hole density.^{20,61,69,70} As a consequence of oxidation, evolved LSPR band, gradually blue-shifted LSPR peak, increscent peak intensity and decrescent peak width were observed (Fig. 3a). In contrast, reducing agent such as alkylthiols, 3-mercaptopropionic acid (MPA), sodium biphenyl as well as diisobutylaluminium hydride (DIBAH) can induce a filling of copper vacancies, leading to a decrease of hole density.^{11,61,62,69} As a consequence of reduction, suppressed LSPR band, gradually red-shifted LSPR peak, decrescent peak intensity and increscent peak width were observed (Fig. 3b). The presence of oxidants such as oxygen favors the extraction of Cu^+ from the $\text{Cu}_{1.8}\text{S}$ ND lattice through the formation of CuO (hole generation), while the electron donors such as MPA favors electron injection into the ND lattice and thus the filling of holes (Fig. 3c). It is worth noting that, upon oxidation and reduction, the accompanied optical bandgap also blue-shifts and red-shifts, known as Burstein–Moss effect, which is expected from a change of hole density within the ND. The estimated hole density varied from $N_{h1} = 4.47 \times 10^{21} \text{ cm}^{-3}$ to $N_{h2} = 5.88 \times 10^{21} \text{ cm}^{-3}$ upon oxidation, while the hole density decreased to $N_{h3} = 5.20 \times 10^{21} \text{ cm}^{-3}$ after MPA reduction. In spite of the variation of LSPR and optical bandgap, the crystal structure of roxbyite is remained.

2.4.2 Crystal structure control and transformation.

In addition to hole density, LSPR manipulation can be realized by crystal structure control and transformation. The capability

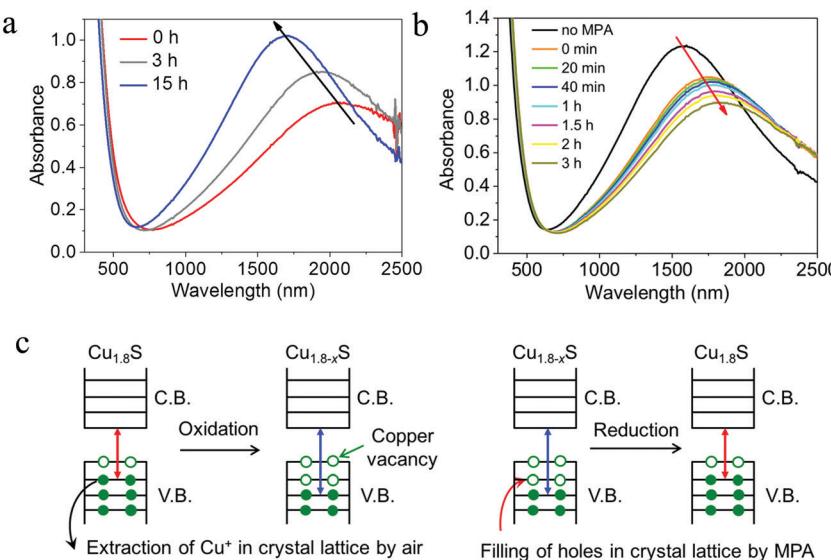


Fig. 3 Chemical redox reactions of roxbyite $\text{Cu}_{1.8}\text{S}$ NDs. Time evolution of UV-vis-NIR absorption spectra of $\text{Cu}_{1.8}\text{S}$ NDs under (a) air oxidation and (b) MPA reduction. (c) Scheme of oxidation induced copper vacancy generation and reduction induced hole filling. Note: $\text{Cu}_{1.8-x}\text{S}$ suggests the copper-poor roxbyite phase. Chemical redox reactions change the hole density or the stoichiometry of roxbyite copper sulfide, but do not change the crystal structure. (Reproduced with permission from ref. 69, copyright 2020 Springer.)

to access different crystal structures from a given NP holds great potential for fundamental studies and diverse applications. The crystal structure control has been realized by different Cu/S feed ratios and sulfur sources.^{21–23} For crystal structure transformation, started from covellite CuS , *via* either Cu^+ insertion or alkylthiol thermal-treatment, a series of Cu_{2-x}S phases can be achieved.^{71,72} Actually, the most important factor among Cu_{2-x}S with different crystal structures on the LSPR is the hole density,³³ because the well-studied crystal structures such as low/high chalcocite, djurleite, roxbyite and covellite are catalogued by asymmetric hcp of sulfur anions, with copper cations taking up interstitial spaces between the sulfur lattices. As aforementioned, according to eqn (2)–(4) in Section 2.3, crystal structure control and transformation will result in a variation of hole density, leading to much stronger effect on the LSPR peak position and absorption strength. In the case of crystal structure transformation, upon an addition of Cu(i) complex, terakis (acetonitrile) copper(i) hexafluorophosphate ($[\text{Cu}-(\text{CH}_3\text{CN})_4]\text{PF}_6$), covellite $\text{Cu}_{1.1}\text{S}$ NDs were successively enriched with copper and converted into $\text{Cu}_{1.3}\text{S}$, $\text{Cu}_{1.5}\text{S}$ and Cu_2S , respectively (Fig. 4a). $\text{Cu}_{1.3}\text{S}$ and $\text{Cu}_{1.5}\text{S}$ could be assigned to metastable “copper-rich covellite” phase, evidenced by a slight shift of XRD peaks of covellite to lower diffraction angles. The Cu_2S could not be assigned considering just a single phase, and might be a mixture of anilite $\text{Cu}_{1.75}\text{S}$ and low chalcocite Cu_2S . Combined X-ray photoelectron spectroscopy (XPS) and electron paramagnetic resonance (EPR) data indicate that, during this transformation the valence of copper cation keeps close to +1, compensated by the variation of sulfur anion valence from -1 in CuS to -2 in Cu_2S . Along with structural change, a suppression and even a quenching of LSPR band of covellite $\text{Cu}_{1.1}\text{S}$ was also observed (Fig. 4b). This tendency of LSPR variation is similar with that by chemical reduction, but has much stronger effect. Accordingly, crystal structure transformation with large tuning of hole density,

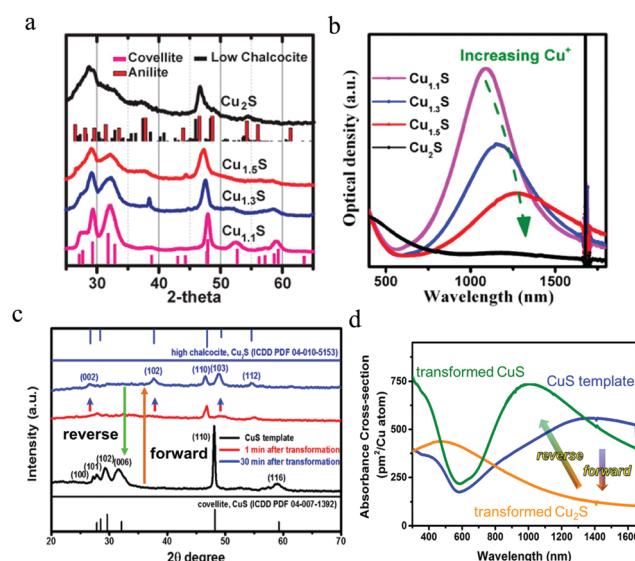


Fig. 4 Crystal structure transformation of covellite copper sulfide NDs. (a) Cu^+ complex reduction induced and (c) 1-DDT thermal-reduction induced structural change and (b and d) corresponding UV-vis-NIR absorption spectra evolution. In (c), the transformation from Cu_2S to CuS was triggered by sulfur-organic solutions including sulfur-oleic acid, sulfur-oleylamine and sulfur-octadecene. (a) (Reproduced with permission from ref. 71, copyright 2013 American Chemical Society.) and (b) (Reproduced with permission from ref. 72, copyright 2017 American Chemical Society.)

provides a favorable tool to manipulate LSPR response for special applications. Furthermore, not only the Cu(i) complex, but also alkylthiol such as 1-dodecanethiol (1-DDT) could induce structural transformation of covellite CuS to chalcocite Cu_2S . Upon thermal-treatment, 1-DDT molecules broke disulfide bonds in CuS and removed sulfur from the lattice, which drives an expansion of lattice in high chalcocite Cu_2S along the *c*-axis



direction (Fig. 4c). Subsequently, the transformed Cu_2S was recovered to CuS by sulfur-organic solutions. Interestingly, a morphology as well as a particle size variation is accompanied with the forward and reverse structural change. Excluding the effect of morphology and size variation, the LSPR response is almost quenched by phase transformation from CuS to Cu_2S , and partially recovered by the reverse phase transformation from Cu_2S to CuS (Fig. 4d).

2.4.3 Morphology control. In metals, plasmon oscillation mode and thus the number of LSPR peak is mainly determined by NP morphology. Control of NP morphology provides a powerful tool to tune their LSPR responses. For example, anisotropic silver nanorods (NRs) and triangular nanoplates possess multiple LSPR peaks, while Ag nanospheres (NSS) have only LSPR peak.^{73–75} Nonetheless, in the case of Cu_{2-x}S NPs, the effect of morphology on the LSPR is not clearly understood and it seems that the morphology has little influence on their plasmon oscillation mode. Djurleite $\text{Cu}_{1.94}\text{S}$ tetradekahedral NPs prepared by Sn^{4+} directed seeding growth of NDs only exhibit one LSPR peak in spite of anisotropic morphology. Heating these NPs in the presence of oleylamine results in pseudomorphic structural conversion to roxybite $\text{Cu}_{1.8}\text{S}$, which merely increases the plasmon oscillation strength rather than the oscillation mode (Fig. 5a).⁷⁶ Moreover, djurleite $\text{Cu}_{1.94}\text{S}$ NRs were prepared by selective utilization of copper precursors. In contrast to silver NRs with two oscillation modes, *i.e.*, transversal mode (T-mode) and longitudinal mode (L-mode), the $\text{Cu}_{1.94}\text{S}$ NRs only show one LSPR peak and an oxidation of the NRs by Ce^{4+} ammonium nitrate could not induce the anisotropic oscillation mode (Fig. 5b). Furthermore, not only the NRs, but also the hexagonal NDs as well as prisms does not show any morphology related plasmon oscillation mode.⁷⁷ As aforementioned, according to eqn (2) and (4)–(7) in Section 2.3, the hole density for the singular LSPR peak in the ND can be determined. In the case of Cu_{2-x}S NR with semi-axis $a > b = c$, the determination of hole density is similar with that of Cu_{2-x}S ND. L_j for rod is also a morphology-dependent constant calculated from the ratio of axes of the prolate spheroid with $L_1 + L_2 + L_3 = 1$ and $L_2 = L_3$. The value of L_1 is dependent on the ratio of (b/a) and if $L_1 = 0$ and $L_2 = L_3 = 0.500$, the morphology is rod.³² However, different from above reports, contradicting reports on the effect of morphology on the

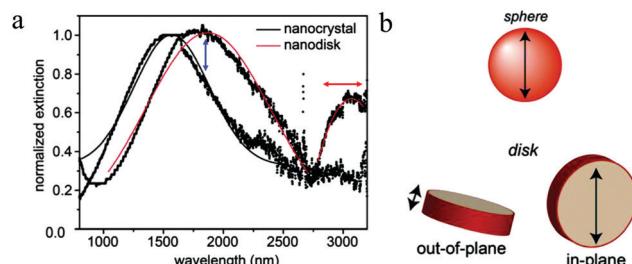


Fig. 6 (a) NIR and MIR absorption spectra and (b) scheme of LSPR polarizations for Cu_{2-x}S NSs and NDs. The LSPR energy, line shape and relative oscillation intensities of the two plasmon bands can be tuned by varying the aspect ratio or free hole density of the ND. (Reproduced with permission from ref. 32, copyright 2011 American Chemical Society.)

plasmon oscillation mode of Cu_{2-x}S NPs has been published. Hsu *et al.* prepared chalcocite/djurleite Cu_{2-x}S NSs and NDs using a solventless synthesis and the NSs show one dipolar oscillation mode, while for NDs, the oscillation splits into a higher energy (1600–1900 nm) out-of-plane mode and a low energy (3000–3100 nm) in-plane mode (Fig. 6a and b).³² In contrast, metallic covellite CuS NDs only show an in-plane oscillation mode.^{33,71} Therefore, it is highlighted to consider the crystal structure alongside NP morphology to avoid confusion in assigning LSPR peaks to corresponding oscillation modes.

2.5 Surface chemistry and LSPR in Cu_{2-x}S

As aforesaid, the free hole density and thus the LSPR in Cu_{2-x}S NPs can change spontaneously under exposure of redox environments. Controllable surface hole density in Cu_{2-x}S is essential to any application which requires constant plasmonic behavior, including solar energy conversion, sensing, or phototherapy. Nevertheless, the relationship between surface chemistry and LSPR evolution/stabilization in Cu_{2-x}S has been less investigated. Wei *et al.* reported the effect of oxygen exposure on the variation of LSPR in most stable covellite CuS NDs.⁷⁸ Oxygen molecules adsorbed on the ND surface induced a blue-shift of LSPR peak in the sample and it could be reverted by the repassivation of organic surfactant, oleylamine (Fig. 7a). This surface chemistry relevant LSPR is explained by the fact that, the adsorbed oxygen molecules acted as an electron acceptor, which withdraws electrons from the CuS ND, resulting in an increase of hole density. Furthermore, how different organic surfactants affecting the absorption spectra of CuS NDs was investigated. In spite of approximately the same refractive index, the different electron withdrawing abilities of the surfactant in terms of Lewis acids and bases is responsible for the shift of LSPRs.⁷⁸ Therefore, it is highly important to stabilize the LSPR absorption profile when making use of the Cu_{2-x}S for special applications. Manna *et al.* prepared CuS @amorphous CuPd_xS heterostructure to tune and lock the LSPR in covellite CuS NDs.⁷⁹ This core@shell structure was obtained by reacting the as-synthesized CuS NDs with Pd^{2+} in the presence of a reducing agent ascorbic acid. The reorganization of the CuS lattice caused by preferential accommodation of Pd atoms between the bilayer formed by the S–S covalent bonds, led to the formation of protective amorphous CuPd_xS layer.⁷⁹

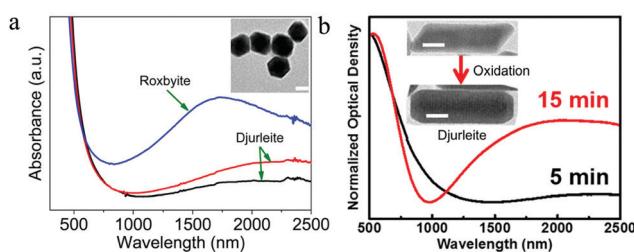


Fig. 5 UV-vis-NIR absorption spectra of (a) roxybite $\text{Cu}_{1.8}\text{S}$ and djurleite $\text{Cu}_{1.94}\text{S}$ tetradekahedral NPs, (Reproduced with permission from ref. 76, copyright 2016 American Chemical Society.) and (b) Djurleite $\text{Cu}_{1.94}\text{S}$ NRs. Scale bar in (a) and (b) is 20 nm. (Reproduced with permission from ref. 77, copyright 2017 American Chemical Society.)



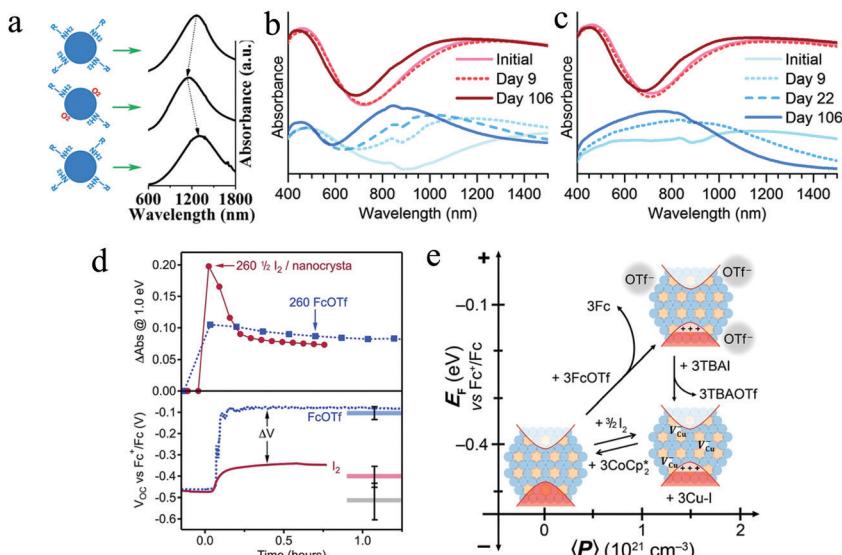


Fig. 7 (a) Scheme of the process of LSPR shifts in CuS NDs with unload and reload of organic surfactant oleylamine. (Reproduced with permission from ref. 78, copyright 2013 American Chemical Society.) Time evolution of LSPR absorption spectra of (b) monoclinic and (c) tetragonal chalcocite NPs with (red, top) and without (blue, bottom) MoS_4^{2-} -treatment. The LSPR absorption in the initial samples indicates that the chalcocite NPs have been oxidized after purification. (Reproduced with permission from ref. 81, copyright 2016 Royal Society of Chemistry.) (d) Differential LSPR absorption intensity and open-circuit potentials (V_{OC}) plotted vs. time for Cu_2S NPs oxidized with same equivalent of $1/2 \text{I}_2$ (solid) and FcOTf (dashed). (e) Summary of redox control over hole density, copper vacancy concentration, and E_F in Cu_{2-x}S NPs. (Reproduced with permission from ref. 83, copyright 2018 American Chemical Society.)

In the case of Cu_2S NPs, the loss of copper is a surface-facilitated process.¹¹ Organic surfactants such as 1-dodecanethiol, which is widely used in the synthesis of Cu_2S NPs, as well as commonly employed oleylamine cannot protect the sample from oxidation sufficiently and results in a rapid evolution of LSPR absorption bands.^{19,80} Spontaneous hole generation has limited the use of chalcocite copper sulfides as a solar light absorber in photovoltaics. In order to passivate the chalcocite Cu_{2-x}S NPs, Plass *et al.* chose tetrathiomolybdate (MoS_4^{2-}) as an effective replacement surfactant due to its strong chelation to copper cations, which could form an inert layer where oxidation could not occur and prevented Cu^+ loss from the sample.⁸¹ $(\text{NH}_4)_2\text{MoS}_4$ -treated Cu_{2-x}S NPs were stable, preventing evolution in the LSPR absorption in air (Fig. 7b and c) and other oxidative environments. Notably, the LSPR stabilization by $(\text{NH}_4)_2\text{MoS}_4$ treatment is a concentration dependent process. An increase of the $(\text{NH}_4)_2\text{MoS}_4$ concentration results in the dissolution of Cu_{2-x}S NPs and the reformation of amorphous Cu_2MoS_4 .⁸¹ In addition to the surface passivation by inorganic surfactant, the chalcocite Cu_{2-x}S can be stabilized by alloying with metal cations such as Ag^+ and Fe^{3+} .^{36,82} Another important phenomenon of surface chemistry and LSPR in Cu_{2-x}S NPs is the stabilization of excess free holes. Gamelin *et al.* applied spectroelectrochemical potentiometry as a quantitative *in situ* technique to probe the low chalcocite Cu_2S NP Fermi-level energies (E_F) during redox reactions. They investigated spectroscopically indistinguishable LSPR bands in Cu_2S NPs with and without generation of lattice copper vacancies and demonstrated that the copper vacancies are more effective in the stabilization of excess free holes than surface anions.⁸³ With the injection of same equivalent of oxidant

(ferrocenium triflate: FcOTf or $1/2 \text{I}_2$), the LSPR absorption of Cu_2S NPs rapidly increases and then slowly drops over several minutes to converge to an equilibrium value. The more positive equilibrium potential (V_{eq}) when NPs oxidized by FcOTf than by I_2 , meaning that the holes formed in valence-band by FcOTf oxidation are nearly 300 mV less stable than those formed by I_2 oxidation at the same hole density ($\langle P \rangle$) (Fig. 7d). This difference indicates that these two oxidants yield substantially different p-doped NPs. The Cu-I bond formation could induce and stabilize the holes in valence band and provide the necessary driving force to induce copper vacancy formation. Furthermore, the oxidized Cu_{2-x}S NPs by I_2 were reduced by the addition of decamethylcobaltocene (CoCp^*_2). The CoCp^*_2 both eliminated valence band holes and pushed the Cu^+ back into the Cu_{2-x}S NP lattice to reform the stoichiometric Cu_2S phase (Fig. 7e).⁸³

2.6 Correlation between LSPR and photoluminescence in Cu_{2-x}S

The investigation on exciton-plasmon coupling is highly interesting for the exploration of photocatalytic charge transfer, emission enhancement or suppression as well as the Fano resonances.²⁴ For Cu_{2-x}S with dually excitonic and plasmonic responses, the photoluminescence (PL) is highly correlated with its LSPR. Kriegel *et al.* monitored the PL of Cu_{2-x}S NPs as a function of hole density, which is dynamically tuned by air oxidation and DIBAH reduction (Fig. 8a and b).¹¹ As expected, a suppression and even a quenching of PL intensity was observed upon oxidation as the number of copper vacancy increased, which leads to more efficient in non-radiative recombination of excitons. The quenching of PL could be attributed to several



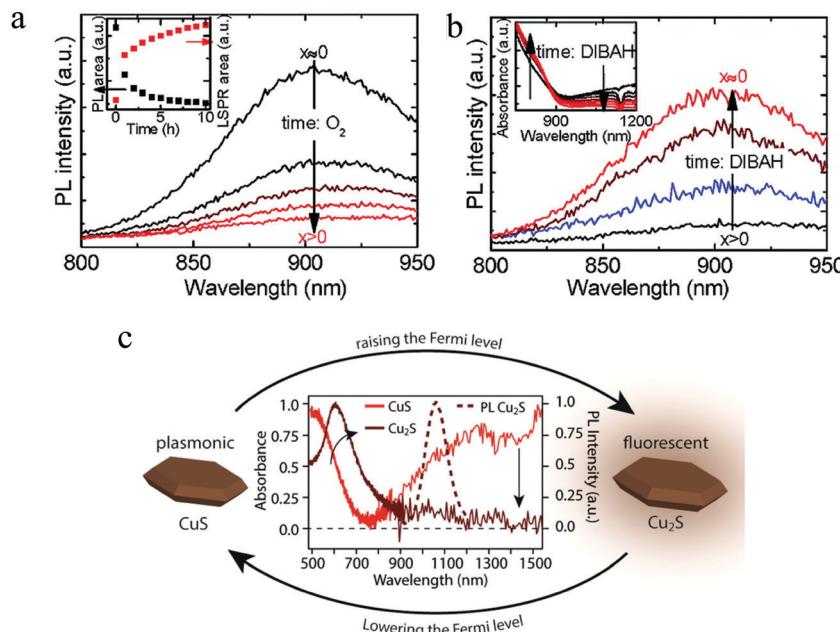


Fig. 8 Time evolution of the PL spectra of Cu_{2-x}S NPs (a) during oxidation by air and (b) subsequent reduction by DIBAH. PL intensity conversely varied with the density of copper vacancy. (Reproduced with permission from ref. 11, copyright 2012 American Chemical Society.) (c) Crystal structure transformation from plasmonic covellite CuS to fluorescent low chalcocite Cu_2S by electrochemical intercalation of Cu^+ . The PL of the Cu_2S NP film centered around 1050 nm and the quantum yield was established to be ca. 1.8%. (Reproduced with permission from ref. 85, copyright 2017 American Chemical Society.)

processes such as Auger recombination, carrier trapping as well as exciton–plasmon energy transfer.^{11,84} Conversely, PL intensity was gradually recovered upon reduction, owing to a decrease of copper vacancy. Furthermore, time-resolved PL spectroscopy indicates that the exciton lifetime is shortened with the emergence of LSPR, while it is lengthened with the LSPR bleaching. In addition, electrochemical intercalation of Cu^+ into plasmonic CuS NP film resulted in an irreversible phase conversion to the fluorescent Cu_2S (Fig. 8c).⁸⁵ As aforesaid, the as-synthesized low chalcocite Cu_2S NPs are usually unstable under ambient conditions and transform easily to djurleite phase. However, the Cu_2S films prepared by this electrochemical reduction were stable and show unique PL characteristics in spite of the uncertain origin of the light emission. Furthermore, the copper vacancy in roxbyite $\text{Cu}_{1.8}\text{S}$ dramatically declines the PL intensity. Although 1-DDT molecules could partially reduce the number of copper vacancy in roxbyite $\text{Cu}_{1.8}\text{S}$ triangular nanoplates, the residual ones result in complete quenching of the PL intensity.⁸⁶ As a consequence, the presence or absence of LSPR absorption in Cu_{2-x}S NPs should be carefully considered for particular applications.

3. Plasmonic applications

In this section, we will focus on the applications of Cu_{2-x}S NPs resulting from their LSPR responses in NIR and MIR regions. In order to comprehend the origin of plasmonic applications, it is worth recalling the sequence of events following absorption/scattering of photons by a well-known plasmonic Au NP,⁸⁷ which can be somewhat closely correlated with Cu_{2-x}S .

In particular, the potential of highly resonant absorption/scattering of NIR and MIR photons and subsequent hole–hole/hole–phonon scattering, vibrational resonance of crystal lattice and coupling to the environment as well as the near electric field enhancement characteristics of Cu_{2-x}S NPs have been well explored for a diversity of applications. (i) The absorbed NIR and MIR light is dissipated to the surrounding medium in the form of heat, making the Cu_{2-x}S NPs as local heat sources and promising candidates for biological applications such as photoacoustic imaging, photothermal therapy as well as other perfect IR absorbers. (ii) The LSPR-induced hot carriers generated *via* inelastic scattering in Cu_{2-x}S transfer to metal and semiconductor counterparts through heterostructured interface, endowing them suitable for organic synthesis, photo(electro)catalytic water splitting and many others. (iii) The amplification of local electric field intensity by LSPR enables Cu_{2-x}S NPs to enhance two-photon upconversion of luminescent materials and SERS sensitivity.

3.1 Photoacoustic imaging and photothermal therapy

After illumination, Cu_{2-x}S NPs can rapidly convert NIR/MIR light energy into heat for photoacoustic imaging (PAI) and photothermal therapy (PTT). Recently, PAI has developed rapidly as a noninvasive biomedical imaging technique that detects the thermoelastic expansion and ultrasonic signals resulted from photothermal expansion of biological tissues after light illumination. This technique is a combination of molecular specificity of optical imaging depth and spatiotemporal resolution of sonography.^{88,89} In the NIR-II region (1000–1700 nm), there is low photon absorption and scattering by biological tissues. As a result, selection of Cu_{2-x}S NPs with intense NIR-II absorption



as a contrast agent, can effectively improve the quality (signal-to-noise ratio) of PAI. Ku *et al.* synthesized 11 nm CuS NPs and used for PAI with a 1064 nm laser, well within the NIR-II region. The CuS NPs allowed clear imaging of lymph nodes located 12 mm below the skin surface. Additionally, CuS NPs within agarose gel embedded in chicken breast muscle were imaged at a depth of ~5 cm with a high in-plane resolution of ~800 μ m as well as an imaging sensitivity of ~0.7 nmol per image voxel.⁹⁰ In order to relieve the *in vivo* toxicity to the blood and normal tissues, Mou *et al.* prepared ultrasmall Cu_{2-x}S NPs (<5 nm) with intense LSPR absorption in NIR-II region and employed them as an excellent PAI nanoplatform, to identify the site, size and morphology of the mouse tumor. Moreover, the Cu_{2-x}S NPs can be completely cleaned through feces and urine, implying their high biosafety.⁹¹ In addition to PAI, PTT also is a noninvasive treatment using NIR light absorbers to generate thermal energy under laser irradiation, leading to elevated local temperatures and irreversible damage to adjacent tumor cells. Compared to traditional cancer therapies such as surgery, chemotherapy as well as radiation therapy, PTT has the merits of low cost and minimal side effects.^{92,93} Therefore, the Cu_{2-x}S NPs composed of earth abundant and nontoxic elements show strong LSPR absorption in NIR-II region, holding great promise in the area of PTT. Ding *et al.* constructed a dual plasmonic hybrid Au–Cu₉S₅ core/shell NPs that have high LSPR absorption in NIR-II window and exhibit 50% higher absorption at 1064 nm compared to the individual Cu₉S₅ NPs. The core/shell NPs showed large optical absorption cross section, high photothermal transduction efficiency at 37%, large tissue penetration depth, excellent X-ray attenuation ability as well as low cytotoxicity, which enable them for robust photothermal therapy in the NIR-II window with low NP dose as well as laser flux, making them as potential theranostic materials for X-ray CT imaging.⁹⁴ More recently, Zhu *et al.* used biocompatible Au@Cu_{2-x}S NPs for PTT and demonstrated that, under NIR laser irradiation (808 nm, 1.5 W cm⁻²), tumor volume decreased about 50% in 22 days treatment in the presence of Au@Cu_{2-x}S NPs. For control, the same tumor increased by 2–3 times without the NPs.⁹⁵ Chang *et al.* prepared Au@void CuS yolk-shell NPs (YSNPs) for “chemophototherapy”. A series of Au NPs with LSPR absorption peaks at 520, 700, 808, 860, and 980 nm were embedded in CuS hollow NPs to find the most efficient Au–CuS YSNPs according to the resonance energy transfer process. The Au_{808nm}–CuS and Au_{980nm}–CuS YSNPs were found able of inducing the largest temperature elevation and generating the most significant hydroxyl radicals under the irradiation of 808 and 980 nm laser, accordingly damaging the most severe 4T1 cell through oxidative stress mechanism.⁹⁶ In addition to the tunable Au core@fixed CuS shell dual plasmonic nanocomposites, another strategy is the fabrication of Au core@variable Cu_{2-x}S shell NPs. Sun *et al.* selectively grown copper sulfide shells such as covellite (CuS), digenite (Cu_{1.8}S), and non-stoichiometric Cu_{1.5}S on the surfaces of Au NP cores. This synthetic method provides not only fine control of the core and shell sizes but also adjustment of the hole density in the Cu_{2-x}S shell. The unique multimodal photothermal behaviors of these core@shell NPs upon selective photoexcitations at their distinct LSPR frequencies, enable the authors to quantitatively evaluate and

compare the intrinsic photothermal efficiency of multiple types of hot charge carriers, excited-state lifetimes, energy distributions as well as energy transfer pathways.⁹⁷ Nonetheless, the utilization of high power intensity above therapeutically permissible values, limits their practical applications. Additionally, the size of the core/shell NPs exceeded 20 nm in all these studies. To circumvent the above limits to a certain extent, Wang *et al.* reported the preparation of physiologically stable 6.5 nm chalcocite/djurleite Cu_{2-x}S NPs with LSPR centered at 1300 nm and analyzed their PTT efficacy (Fig. 9a and b).⁹⁸ The Cu_{2-x}S NPs demonstrated high photothermal heating and therapy efficiency under an 808 nm laser with of 0.6 W cm⁻² power intensity. Interestingly, the Cu_{2-x}S NPs not only worked as photothermal therapy agent, but also as photodynamic therapy (PDT) agent by producing high levels of reactive oxygen species, which are linked to local Cu(i) release triggered deacetylation reaction of non-fluorescent dye. The combination of PTT and PDT will evoke further investigations on the different therapeutic mechanisms of plasmonic copper chalcogenide NPs.

3.2 Photocatalytic organic synthesis

In general, there are two proposed mechanisms for plasmonic Cu_{2-x}S NPs to enhance the photocatalytic performances of organic synthesis, including local heating effect and charge transfer resulting from the hot carrier (hole) generation. In most studies, metal–semiconductor heterostructures were selected for photocatalytic tests, where noble metals such as Pt and Pd work as the catalytic center and Cu_{2-x}S serve as the LSPR component. Cui *et al.* prepared Pd/Cu₇S₄ heterostructured NPs and promoted the conversion efficiency of Suzuki coupling reaction, oxidation of benzyl alcohol as well as hydrogenation of nitrobenzene.⁹⁹ The LSPR peak of Pd/Cu₇S₄ was mainly located in the region of 1500–2000 nm (Fig. 10a). The authors showed that for all three catalytic reactions, the maximum conversion efficiency was obtained under 1500 nm laser irradiation (Fig. 10b). The intense LSPR absorption excited by 1500 nm laser leads to the formation of hot holes in the valence band of Cu₇S₄, some of which possess sufficient energy to overcome the Schottky barrier at the Cu₇S₄–Pd interface, and inject into the Pd domain. The accumulation of hot holes in Pd make it positively charged, serving as a key contribution to the enhancement of photocatalytic activity of Pd/Cu₇S₄, while the local heating effect only contributes a portion of catalytic activity. Furthermore, in recycling experiments, the Pd/Cu₇S₄ NPs were stable and could maintain their structures and photocatalytic activity after five cycles. This synthetic strategy can be easily applied to other Cu₇S₄@M (M = Pt, Ag, Au) NPs, bringing new insights for design and fabrication of highly effective photocatalysts for efficient conversion of solar energy to chemical energy.⁹⁹ In addition, Wang *et al.* exploited a new and distinctive combination method to prepare a number of Janus-like Pd-dispersed CuS (amorphous PdCuS/partially crystalline CuS) NPs with tailored compositions and proportions *via* cation exchange, and used for selective semihydrogenation of phenylacetylene.¹⁰⁰ This synthetic approach promoted not only the utilization efficiency of Pd metal atoms, but also the stability of anchored active sites. By optimizing the amount of



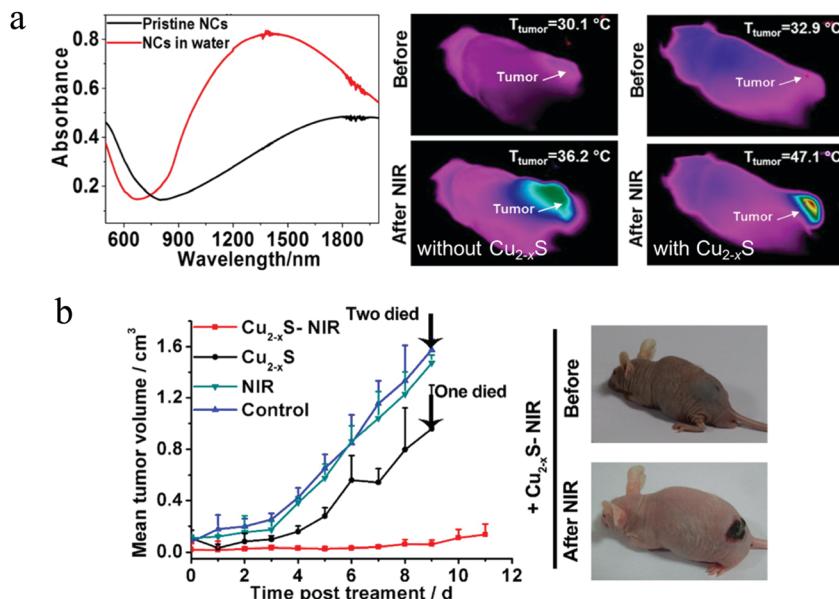


Fig. 9 (a) LSPR absorptions of the Cu_{2-x}S NPs before (black curve) and after (red curve) ligand exchange as well as thermal images of the tumor sites with and without injection of the Cu_{2-x}S NPs. After NIR laser light irradiation, the local temperature of tumor sites with Cu_{2-x}S NPs increased more rapidly than that of tumor sites without Cu_{2-x}S NPs. (b) Variation of mean tumor volume in the control group/NIR group/Cu_{2-x}S group/Cu_{2-x}S-NIR group and images of nude mice taken before and at second day with Cu_{2-x}S-NIR treatment. (Reproduced with permission from ref. 98, copyright 2015 American Chemical Society.)

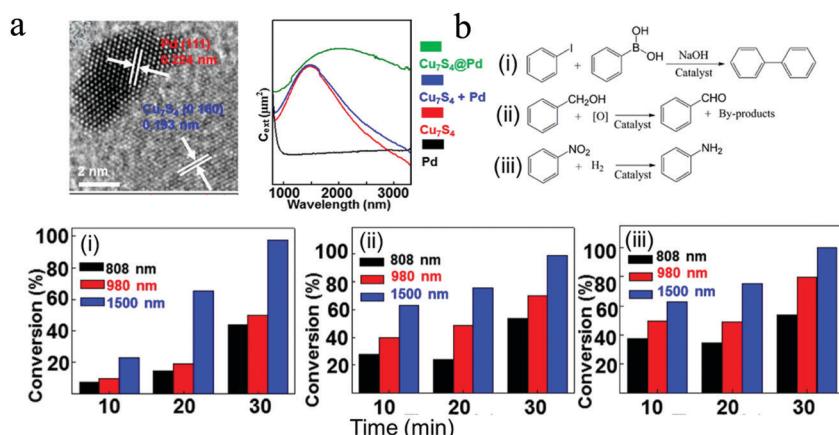


Fig. 10 Photocatalytic organic synthesis by (a) Pd/Cu₇S₄ heterostructures. (b) Conversion efficiency of (i) Suzuki coupling reaction of iodobenzene with phenylboronic acid, (ii) oxidation of benzyl alcohol, (iii) hydrogenation of nitrobenzene under 808 nm, 980 nm and 1500 nm laser irradiation (0.45 W cm⁻²) for 30 min. (Reproduced with permission from ref. 99, copyright 2015 American Chemical Society.)

dispersed Pd and deactivation effects from copper and sulfur, the PdCuS/CuS NPs exhibited high catalytic activity with a conversion of 90.3% and selectivity (towards styrene) of 91.1%. The authors believed that this dispersing a noble metal in a sulfide substrate is a feasible way to achieve effective photocatalysts for organic reactions with superior selectivity, activity and stability.

3.3 Photo(electro)catalytic water splitting

Hydrogen production through photocatalytic water splitting using solar energy is an attractive solution to the global energy and environmental issues.¹⁰¹ Initialized by Honda-fujishima effect, numerous endeavors have been made to develop efficient

photocatalysts for water splitting.¹⁰²⁻¹⁰⁴ However, the solar spectrum used by most photocatalysts is limited to the UV and visible regions. NIR light, which accounts for around 50% of the solar energy, cannot be efficiently utilized. Thanks to the intense LSPR absorption, Cu_{2-x}S NPs can efficiently use the NIR light. Nevertheless, it is suggested that the ultrafast hot carrier relaxation and annihilative exciton recombination in single Cu_{2-x}S may give rise to low energy conversion efficiency.^{11,105} In order to increase the photocatalytic activity, Lian *et al.* prepared Cu₇S₄/CdS heterostructured NPs and investigated their photocatalytic water splitting performance under NIR light (Fig. 11a and b).¹⁰⁶ Although the single Cu_{2-x}S NPs exhibit intense LSPR absorption in NIR region, they could not catalyze water splitting under NIR illumination ($\lambda > 800$ nm).



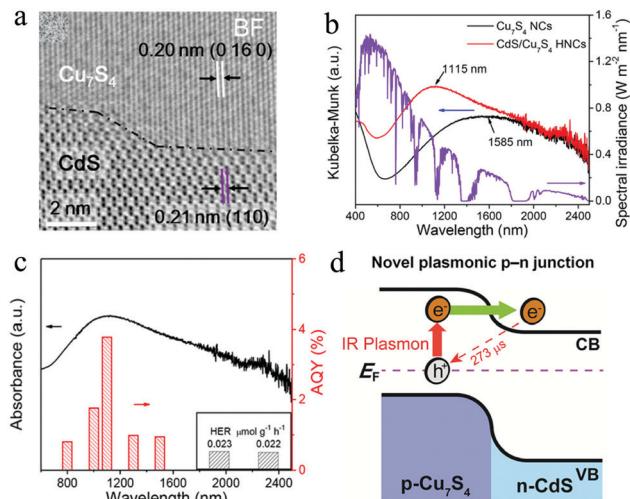


Fig. 11 (a) Photocatalytic H₂ evolution over Cu₇S₄/CdS NPs under NIR illumination ($\lambda > 800$ nm) in a Na₂S–Na₂SO₃ aqueous buffer solution (pH = 12). (b and c) Absorption spectrum and apparent quantum yield for H₂ evolution of the Cu₇S₄/CdS NPs under the monochromatic light illumination (6 mW cm⁻²) demonstrating the efficient NIR light to hydrogen conversion. (d) NIR plasmon induced hot electron generation and interfacial migration. (Reproduced with permission from ref. 106, copyright 2019 American Chemical Society.)

Furthermore, a mixture of the Cu₇S₄ and CdS NPs did not promote the H₂ evolution, demonstrating the great importance of intimate heterostructure interface between these two components. The apparent quantum yield (AQY) of Cu₇S₄/CdS NPs reached 3.8% at 1100 nm, which is much higher than other NIR active composites such as Pt-W₁₈O₄₉/g-C₃N₄ heterostructure, Pt-modified Au NRs and Au NRs/La₂Ti₂O₇ (Fig. 11c).^{107–109} The transient absorption results revealed that LSPR-induced hot electron transfer from Cu₇S₄ to CdS realizes long-lived charge separation (>273 μs), leading to high NIR light to hydrogen conversion efficiency (Fig. 11d). Furthermore, Bhavani *et al.* prepared hollow skeletal-type Cu₇S₄ superstructures integrated with CdS NRs and investigated their photocatalytic hydrogen generation *via* water splitting under solar light irradiation using lactic acid as a hole scavenger.¹¹⁰ The Cu₇S₄/CdS composite exhibited a 3 times higher H₂ generation rate than that of single CdS. Moreover, an addition of a small amount of nickel salts to the reaction solution in order to produce a Ni–Cu₇S₄/CdS composite, dramatically improved the photocatalytic H₂ generation, with an activity 20 times higher than the single CdS and 7 times higher than the Cu₇S₄/CdS. The authors suggested that the enhancement of photocatalytic activity in Ni–Cu₇S₄/CdS composite could be attributed to effective charge carrier separation by the unique properties of the skeletal-type Cu₇S₄ cages as well as superior charge transportation by the Ni salts in the reaction solution.¹¹⁰ Along similar lines, Zhang *et al.* decorated the hierarchical Cu₇S₄/ZnIn₂S₄ with black phosphorus quantum dots (BPQDs) and investigated the photocatalytic hydrogen production under the irradiation of wavelength $\lambda > 420$ nm.¹¹¹ The intimate hierarchical structure of Cu₇S₄ with ZnIn₂S₄ and carrier separation ability by BPQDs were synergistically beneficial for boosting the

photocatalytic activity of H₂ generation with a rate about 6.8 times higher than that of the pristine ZnIn₂S₄.¹¹¹ More recently, complex structure such as n-p-n type CdS@Cu_{2-x}S–MoS₂ and p-n-p type Cu_{2-x}S@CdS–Bi₂S₃ nanocomposites have been fabricated for efficient photocatalytic H₂ generation.^{112,113} Liu *et al.* successfully prepared a specific stratified CdS@Cu_{2-x}S–MoS₂ NPs with the MoS₂ nanosheets epitaxially and vertical to the surface of CdS@Cu_{2-x}S NRs by applying cation exchange and subsequent seeding growth. This stratified heterostructures exhibited high photocatalytic activity ($\lambda > 400$ nm) for H₂ generation *via* water splitting at a rate of rate of 14184.8 μmol g⁻¹ h⁻¹ in the absence of any noble metal cocatalysts, which is 12.7 and 97.2 times higher than that of CdS@Cu_{2-x}S and pure CdS and is comparable to the activity of Pt-decorated CdS@Cu_{2-x}S. This superior photocatalytic performance was ascribed to fast charge separation and migration within the framework originated from the unique heterostructure and efficient Cu^I@MoS₂ (Cu^I-intercalated within the MoS₂) catalytic and protective layers as well as the strong Cu_{2-x}S light exciton resonance absorption.¹¹² In addition, Guo *et al.* reported the design and fabrication of tandem Cu_{2-x}S@CdS–Bi₂S₃ sandwich heterojunctions with well-defined structures and interfaces *via* a continuous growth deposition method on the surface of pre-formed hollow Cu_{2-x}S octahedron (Fig. 12a and b).¹¹³ The unique sandwich-layered hollow nanostructure can provide large specific surface areas and sufficient reaction sites and promote the charge separation and migration.

Furthermore, the p-n-p heterojunction coupled with hot electron transfer by the LSPR effect of Cu_{2-x}S increased the photocatalytic H₂ generation performance (Fig. 12c). The optimized Cu_{2-x}S/CdS/Bi₂S₃ photocatalyst exhibited a remarkable H₂ generation rate of 8012 μmol h⁻¹ much higher than the other 3 test samples and the excellent stability of this ternary hollow nanostructure is beneficial to practical applications (Fig. 12d and e). Another working strategy to increase the water splitting efficiency of Cu_{2-x}S is to control the location and distribution of copper vacancies. Ren *et al.* reported the synthesis of surficially defective Cu_{1.94}S nanoflake arrays (NFAs) photocathode from Cu foil, *i.e.*, Cu–(Cu₂S@Cu_{1.94}S).¹¹⁴ The surficial copper vacancies can serve as adsorptive and catalytic sites for water splitting, and the defect free internal Cu₂S transport holes rapidly to the external circuit. Under full-spectrum (300–2500 nm, 200 mW cm⁻²) illumination, the LSPR of Cu_{2-x}S NFAs enhance light harvesting and lead to rapid generation of electron–hole pairs in the near surface layer. The electrons readily separate from the holes and immigrate rapidly to the catalytic sites, where they reduce water molecules into H₂. As a result, the presence of copper vacancy and LSPR effect in Cu_{2-x}S, improves the photoelectrocatalytic performance to an altitude of 2.67 times than that of defect free NFAs, which cannot take advantages of catalytic activity and LSPR effect. This intentional defect engineering is a promising way to the current challenges in solar energy conversion systems.

3.4 Two-photon upconversion

Cu_{2-x}S NPs demonstrate not only tunable LSPR but also a two absorption effect under infrared light excitation, offering a novel “all-in-one” platform for investigating exciton–plasmon



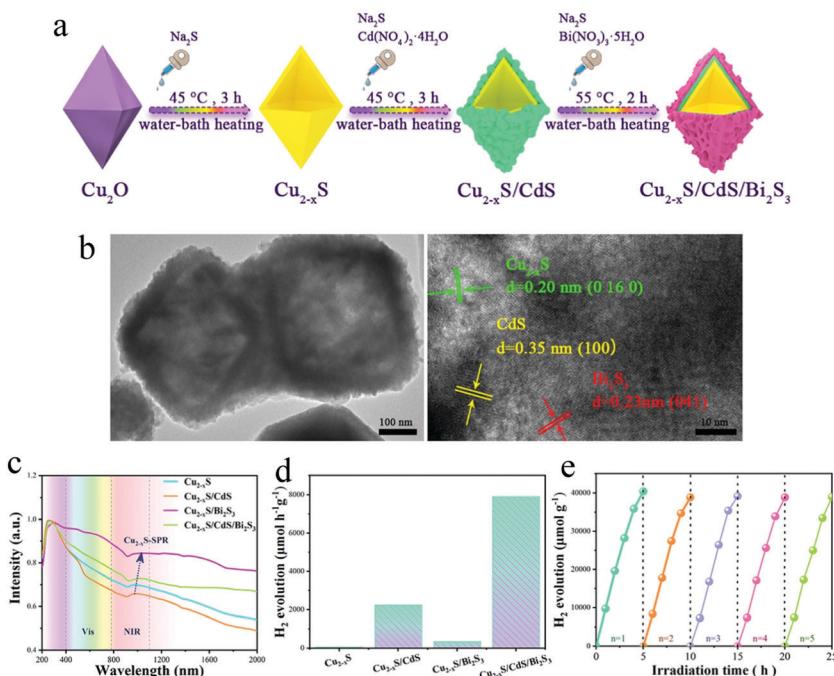


Fig. 12 (a) Fabrication process for the sandwich-layered Cu_{2-x}S @ CdS - Bi_2S_3 octahedral heterostructure. (b) TEM images of the prepared samples indicate the presence of large surface areas and intimate interfaces. (c) UV-vis-NIR absorption spectra of four test samples: Cu_{2-x}S , $\text{Cu}_{2-x}\text{S}/\text{CdS}$, $\text{Cu}_{2-x}\text{S}/\text{Bi}_2\text{S}_3$ and $\text{Cu}_{2-x}\text{S}/\text{CdS}/\text{Bi}_2\text{S}_3$. (d) Photocatalytic H_2 generation of four test samples and (e) H_2 generation stability tests of 100 mg $\text{Cu}_{2-x}\text{S}/\text{CdS}/\text{Bi}_2\text{S}_3$ under vis-NIR illumination ($\lambda > 420$ nm) in a methanol/ H_2O aqueous solution. (Reproduced with permission from ref. 113, copyright 2020 American Chemical Society.)

coupling without a physical or chemical interface. Marin *et al.* reported the synthesis of selenium doped CuS NDs and demonstrated the two-photon upconversion effect and the role of LSPR on fluorescence emission.¹¹⁵ Owing to the energy alignment of the CuS bandgap at twice the LSPR, the CuS NDs exhibit self-enhanced emission (Fig. 13a). The authors measured the emission intensity of CuS as a function of the power of the excitation laser. As expected, one-photon absorption gives a linear relationship, while multiphoton absorption gives rise to a nonlinear relationship between emission intensity and excitation power.¹¹⁶ For a perfect two-photon process, the emission intensity varies with the square of excitation power.¹¹⁷ A slope of 1.87 and 2.37 is observed upon increasing and decreasing of the laser power, which is closely correlated with the power-squared dependence (slope = 2) of a two-photon absorption process. Furthermore, the authors studied the effect of LSPR of CuS NDs on the emission intensity using selenium doping. A blue-shift of the LSPR peak toward the two-photon absorption edge of the Se doped CuS samples, a 3-fold enhancement of emission was observed, indicating that these CuS NPs hold great potential for applications such as bioimaging and photodynamic therapy which require multiphoton absorption. In addition to self-enhanced emission of Cu_{2-x}S NPs, Zhou *et al.* exploited upconversion luminescence enhancement of $\text{NaYF}_4:\text{Yb}^{3+},\text{Er}^{3+}$ NPs induced by Cu_{2-x}S NPs.¹¹⁸ The authors prepared a three-layered Cu_{2-x}S - MoO_3 - NaYF_4 hybrid film consisting of plasmonic Cu_{2-x}S NPs as the bottom layer, MoO_3 with a suitable thickness (8 nm) as a separator layer, and upconversion NPs as the top layer to

investigate the power-dependent red, green, and blue emission from the upconversion NPs. The extraordinary power-dependent behavior of the hybrid film was explained by the interaction mechanism between plasmonic Cu_{2-x}S NPs and $\text{NaYF}_4:\text{Yb}^{3+},\text{Er}^{3+}$ upconversion NPs, *i.e.*, plasmon scattering effect at lower excitation power intensity and electron diffusion by a two-photon interband transition in Cu_{2-x}S NP at higher power intensity.¹¹⁸ Subsequently, they synthesized $\text{Cu}_{2-x}\text{S}@\text{SiO}_2@\text{Y}_2\text{O}_3:\text{Yb}^{3+}/\text{Er}^{3+}$ core-shell nanocomposites and experimentally and theoretically investigated the LSPR induced upconversion enhancement. An enhancement of approximately 30-fold was obtained compared with the sample without Cu_{2-x}S , that is $\text{SiO}_2@\text{Y}_2\text{O}_3:\text{Yb}^{3+}/\text{Er}^{3+}$. This enhancement was induced by a combination effect of the amplification of the excitation field and an increase of resonance energy transfer rate from Yb^{3+} to Er^{3+} .¹¹⁹ Furthermore, the authors introduced photonic crystals (PCs) to manipulate the localized electric fields of Cu_{2-x}S and explored to enhance the upconversion luminescence of $\text{NaYF}_4:\text{Yb}^{3+},\text{Er}^{3+}@\text{NaYF}_4:\text{Yb}^{3+},\text{Nd}^{3+}$ NPs.¹²⁰ In the PMMA OPC/ $\text{Cu}_{2-x}\text{S}/\text{MoO}_3/\text{NaYF}_4$ hybrids, due to the extra contribution of the PC effect, upconversion enhancement factor reached to hundreds to thousands of fold, which is strongly dependent on the crystal structure of Cu_{2-x}S and excitation wavelength (Fig. 13b). Obviously, the $\text{Cu}_{1.8}\text{S}$ hybrid yielded much stronger enhancement factor than that of the CuS hybrid. The optimized enhancement factor under 980 nm laser excitation for the OPC/ $\text{Cu}_{1.8}\text{S}/\text{MoO}_3/\text{NaYF}_4:\text{Yb}^{3+},\text{Er}^{3+}@\text{NaYF}_4:\text{Yb}^{3+},\text{Nd}^{3+}$ hybrids was nearly 1500-fold for whole light emission, 2000-fold for green light emission and 1200-fold for red light emission. It is suggested



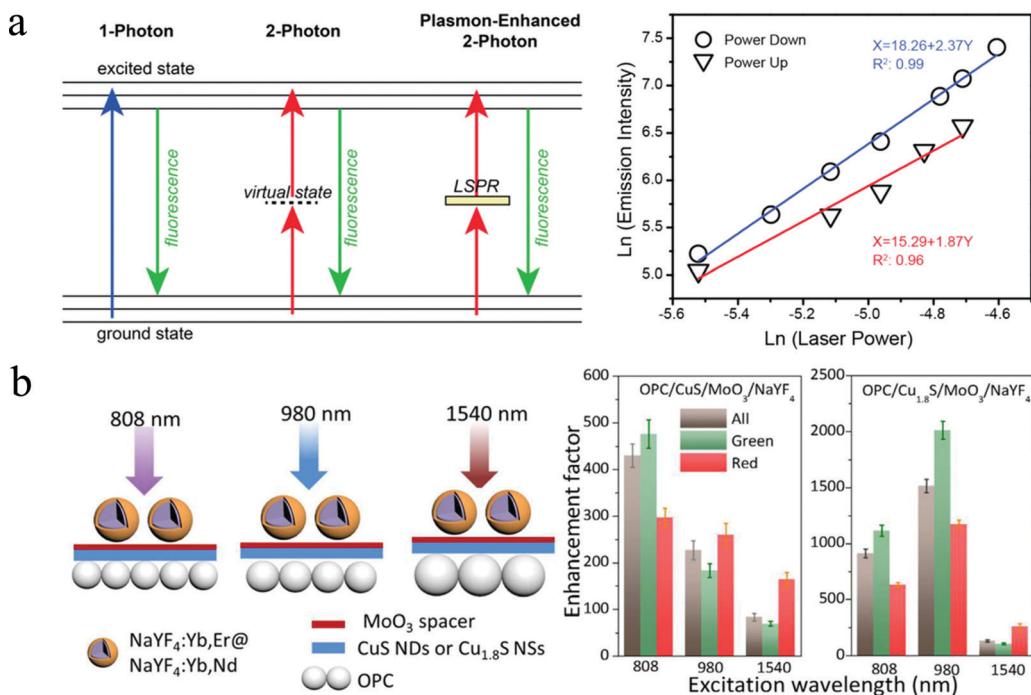


Fig. 13 LSPR coupling with excitonic transitions. (a) Scheme of LSPR enhanced upconversion and log-log plot of the power-square dependence of emission intensity of CuS NDs. Nonlinear relationship between emission and incident laser power as well as slope being close to 2 on a log-log scale indicates the possibility of two photon upconversion process. (Reproduced with permission from ref. 115, copyright 2016 American Chemical Society.) (b) Scheme of Cu_{2-x}S crystal structure dependent excitation and emission of two-photon upconversion OPC/Cu_{2-x}S/MoO₃/NaYF₄ hybrids (808 nm light at 1.25 W mm⁻², 980 and 1540 nm light at 0.15 W mm⁻²). (Reproduced with permission from ref. 120, copyright 2017 American Chemical Society.)

that such a high enhancement for the above hybrid can be ascribed to the amplified local electric field by PC effect and LSPR effect, as well as energy transfer from Cu_{2-x}S to NaYF₄:Yb³⁺,Er³⁺@NaYF₄:Yb³⁺,Nd³⁺ resulted from the two-photon absorption of Cu_{2-x}S NPs.¹²⁰

3.5 SERS probe

Recently, Raman imaging using surface-enhanced Raman scattering (SERS) probes has attracted numerous attention and a large number of efforts have been made on designing highly sensitive SERS probes, including most common noble metallic gold or silver NPs, and several semiconducting materials.^{121–124} However, most of the reported SERS probes are not degradable and therefore they cannot be completely removed from the imaged tissue. To circumvent the current limitation, Qiu *et al.* explored a photodegradable SERS probe using hollow CuS NPs with an outer encapsulation of poly (ethylene glycol) (PEG). By incorporation with Raman-active reporters such as 3,3'-diethylthiatricarbocyanine iodide (DTTC), the biocompatible DTTC-hollow CuS NPs can serve as a novel *in vivo* SERS probe (Fig. 14a).¹²⁵ Under 785 nm laser irradiation, a significant Raman signal intensity enhancement was observed for the DTTC-hollow CuS NPs, while it is faint for the DTTC adsorbed CuS single crystal (Fig. 14b). Furthermore, the DTTC-hollow CuS probe is stable during the incubation in the serum, demonstrating their potential applications in oncology surgery. Firstly, the DTTC-hollow CuS NPs were preoperatively injected into tumor tissues and after resection of the visible ones, the authors used 785 nm laser (10 mW) to

excite the Raman signals of DTTC-hollow CuS NPs for residual tumor visualization. Then, they applied 980 nm laser (800 mW, 3 min) to start the hyperthermia for residual tumor ablation. Finally, by continuous irradiation for 5 min, the large hollow CuS NPs were degraded into small individual crystals, accelerating their self-clearance from imaged tissues (Fig. 14c). However, the low enhancement factor of Cu_{2-x}S SERS probe in some cases is a major obstacle limiting their practical applications. Therefore, it is desirable to explore the facile synthetic strategies for Cu_{2-x}S NPs and reveal the SERS enhancement mechanism. Chen *et al.* developed a facile method to synthesize copper sulfide based nanostructures with different Cu_{7.2}S₄ contents by the annealing of CuS nanosheets at different temperatures.¹²⁶ Compared with pure CuS or pure CuSO₄ (annealing temperature at 350 °C), the sample with a higher Cu_{7.2}S₄ content exhibited lower detection limit of R6G molecules, indicating the SERS performance enhancement by an introduction of nonstoichiometric Cu_{7.2}S₄. Furthermore, the DFT theoretical calculations indicate that the excellent SERS behavior by Cu_{7.2}S₄ is dependent on the charge transfer resonance.¹²⁶ In addition to the CuS SERS biological probe, Bao *et al.* presented an isotropic shell growth of homogeneous CuS on plasmonic Au NPs, exhibiting an efficient SERS-based probe for detection of trace heavy-metal ions. For example, when the CuS shell was immersed in a Hg²⁺-contained solution, the Cu⁺ in the shell can capture and react with Hg²⁺ through cation exchange, *in situ* forming a new HgS shell on the Au surface (Fig. 14d).¹²⁷ Because of the distinctive vibration energy between CuS and HgS, the authors can thus realize the Hg²⁺



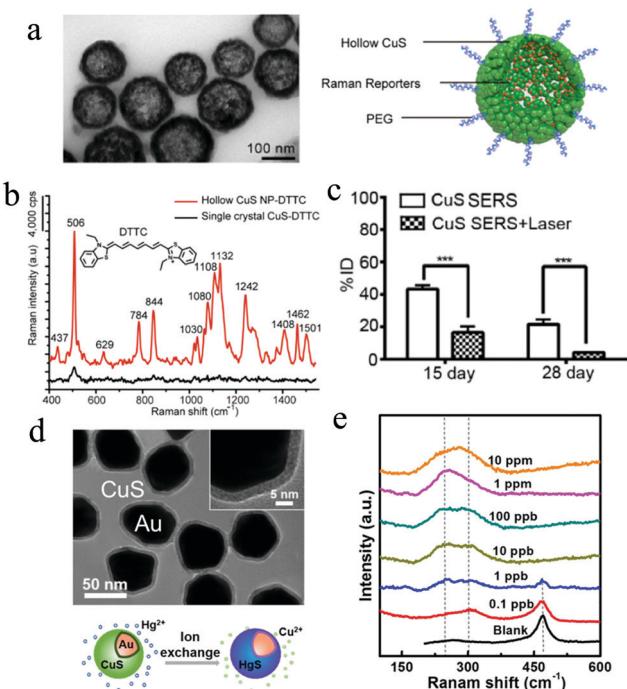


Fig. 14 (a) SERS probe of DTTC & PEG-hollow CuS NPs and (b) SERS signal enhancement of DTTC molecules by hollow CuS NPs. (c) *In vivo* photodegradation of above hollow CuS NPs under laser irradiation. The laser-induced photothermal effect accelerated the clearance of hollow CuS probes from tumors over 15 day and 28 day. (Reproduced with permission from ref. 125, copyright 2019 American Chemical Society.) (d) Au@CuS SERS probe for trace heavy-metal ions detection by cation exchange. (e) Evolution of SERS signals of the Au@CuS NPs to the Hg^{2+} -contained solutions with different concentrations. (Reproduced with permission from ref. 127, copyright 2019 American Chemical Society.)

detection by subsequent SERS measurement. The CuS shell shows a peak at 470 cm^{-1} , which originates from the disulfide bond vibration.^{128,129} After immersion in Hg^{2+} solution, the initial 470 cm^{-1} peak vanished and two new peaks at 250 and 305 cm^{-1} evolved, which matches well with the first-order A_1 and E vibration in HgS , respectively.¹³⁰ With the SERS signal enhancement by Au and CuS NPs, the limit of Hg^{2+} detection of Au@CuS reaches to $<0.1\text{ ppb}$, which is better compared with the reported values (Fig. 14e). More important, this detection technique has strong anti-interference to the environments, even with a concentrated common metal ions or high acidity, demonstrating good practical applications.

4. Conclusion and outlook

In this review, we have briefly summarized the LSPR of Cu_{2-x}S NPs and their plasmonic applications. In particular, the tunable hole density, various crystal structures and stoichiometries, controllable morphology, hot carrier generation, elevated local temperature as well as locally enhanced electric field make the Cu_{2-x}S NPs suitable for a wide range of applications, ranging from plasmonic heating, photocatalysis, two-photon upconversion, SERS probe and many others. In spite of substantial

progress made in understanding the LSPR of Cu_{2-x}S NPs and their applications, there still remain numbers of open concerns relating to the NP synthesis, chemistries, physics, optical modulation and industrial applications. We list some of representative questions as follows:

(i) In Cu_{2-x}S with a special crystal structure, it is of great importance to investigate the structural variations upon chemical-(electro) redox post-treatment, which may largely affect the parameters of its dielectric function. That is vitally important for a more reasonable estimation of the hole density of Cu_{2-x}S NPs from their optical spectra.

(ii) Direct *in situ* mapping the number of copper vacancy and its distribution in Cu_{2-x}S NP is also important to accurately determine its plasmon oscillation mode. In general, the oscillation mode is determined by the simulation methods such as discrete dipole approximation (DDA) and finite difference time domain (FDTD), which may lead to disagreement with the experimental optical spectra of Cu_{2-x}S NPs with different crystal structures and morphologies. The direct copper vacancy mapping is challenging because of the limited spatial resolution of TEM and the low hole density. Further difficulty is the high mobility of Cu^+ in the NP, which is prone to be damaged by the electron beam during analysis.

(iii) The local electric field enhancement in the near surface region of Cu_{2-x}S NP is still low. Optimizing the size, morphology, hole density and its spatial distribution of the NP may increase its enhancement factor. Another feasible approach is to employ the plasmon coupling effect to generate hot-spots, where the ultrahigh electric field intensity can greatly improve excitation efficiency of the LSPR related applications.

(iv) The surface chemistry of interaction between surface ligands and surface free holes of Cu_{2-x}S is not well investigated. For photocatalysis, ionic ligands with positive/negative charges are usually needed to ascertain the Cu_{2-x}S NPs high dispersity in aqueous solutions. How do interfacial electrostatic attraction/repulsion affect the LSPR behavior of Cu_{2-x}S NPs and thus their photocatalytic performances?

(v) Cu_{2-x}S NP used as plasmonic heater: practical considerations such as cost, heating efficiency, NP stability, biosafety and NP clearance should be carefully investigated for any possible practical biological applications. Moreover, a comprehensive study considering both performance and practicality metrics should be conducted over a wide selection of plasmonic materials including not only the Cu_{2-x}S , but also metal oxides, metal nitrides, doped silicon and many others, making optical matches for specialized applications.

(vi) The heterostructure interface between Cu_{2-x}S domain and metal domain or semiconductor domain, which plays an important role in determining the hole transfer efficiency for photocatalysis, is not well studied. Which is the crystal planes of Cu_{2-x}S favoring the hole accumulation? How can selectively load the metal or semiconductor domain on these hole-preferred planes to build intimate interfaces? Furthermore, construction of two-dimensional interface between Cu_{2-x}S and other semiconductor for efficient hole transfer may be another promising strategy. In addition to the interface, the directional dependence



of plasmon oscillation mode in Cu_{2-x}S component also plays a major role.

(vii) Large scale synthesis of Cu_{2-x}S NPs with desired LSPR to meet application demands in future may need development in manufacturing techniques with quality control. How can industrial scale (post)synthetic strategies be innovated while without losing the rigid control over particle size, morphology, doping level as well as crystal structure for a manifold of industrial applications?

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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