REVIEW

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Wide bandgap polymer donors for high efficiency non-fullerene acceptor based organic solar cells

Keqiang He, Pankaj Kumar, Yi Yuan and Yuning Li ⓘ *

In the past few years, the power conversion efficiency (PCE) of organic solar cells (OSCs) has improved rapidly with the milestone value exceeding 18%, primarily owing to the development of novel non-fullerene acceptors (NFAs) as well as matching polymer donors. The molecular structure of a polymer donor fundamentally determines its molecular packing (crystal structure and morphology) and optoelectronic properties, which influence the photovoltaic processes and the ultimate PCE of the OSC device. The structure–property–cell performance relationships of polymer donors with respect to the specific acceptor are very complex, involving numerous parameters, but are extremely important towards the development of high-performance polymer donors to achieve high PCE. This review provides a timely analysis of the top-performing wide bandgap (WBG) polymer donors that have been developed to match the three most representative narrow bandgap NFAs, ITIC, IT-4F, and Y6, in terms of their structural design, fine-tuning of their optoelectronic properties, and control of the morphology and crystallinity of their blends with NFAs. We hope that this article provides deeper insight into the structure–property–cell performance relationships of polymer donors and a collection of useful guidelines and strategies for the design and processing of novel polymer donors for matching with NFAs for achieving ultrahigh performance OSCs.

1 Introduction

Organic solar cells (OSCs) have attracted much attention as a promising technology to convert solar energy into electricity because of their advantages in fabricating flexible, lightweight, large-area, and low-cost solar cells.[1–4] Since the first report by Heeger et al.[5] in 1995, OSCs with a bulk heterojunction (BHJ) structure composed of a blend active layer, comprising a p-type conjugated polymer as a donor and an n-type organic semiconductor as an acceptor, have attracted tremendous attention due to their excellent solution processability and mechanical properties that allow high throughput, roll-to-roll manufacturing. Significant improvements in the OSC performance (with the highest PCE exceeding 18%) have been achieved, largely by the judicious design and delicate synthesis of matching polymer donor and acceptor materials.[6–10]

The first generation acceptor materials for BHJ OSCs comprise fullerene derivatives such as phenyl-C_{61} (or C_{71})-butyric acid methyl ester ($PC_{61}BM$ or $PC_{71}BM$), which were first developed by Wudl et al.[11] They have good solubility in organic solvents, high electron mobility (μ_e), and high electron

affinity.[12,13] Among various polymer donors developed to match these acceptors, a benzo[1,2-b:4,5-b′]dithiophene (BDT)-based polymer, PTB7, developed by Liang et al.[7] in 2010, showed a high PCE of 7.4% and 9.2% using a conventional and inverted device structure, respectively, when blended with $PC_{71}BM$.[7,14] The highest PCE of 11.7% for single-junction binary-blend fullerene-based OSCs was achieved using PffBT4T-C_9C_{13} as a donor by Zhao et al. in 2016.[15] Since then, the development of fullerene-based OSCs has become subdued because of their weak absorption in the visible spectral region, limited energy level tunability, and inadequate long-term stability of the devices caused by the susceptibility to dimerization and gradual aggregation.

To overcome the drawbacks of fullerene-based acceptors, non-fullerene acceptors (NFAs) have been developed.[16,17] ITIC, developed by Lin et al. in 2015, is one of the most efficient NFAs. It has a rigid indacenodithienothiophene (IDTT) central unit and a narrow bandgap of 1.59 eV.[8] OSCs using ITIC as an acceptor and PTB7-Th as a donor showed a moderate PCE of 6.8% due to their similar absorption range with poor absorption in the shorter wavelength region of the solar spectrum. The PCE of ITIC-based OSCs rapidly improved to ∼10% when WBG polymer donors with complementary absorption such as J51 (9.26%),[18] J61 (9.53%),[19] and PBDB-T (11.21%)[16] were used. Recently, a new WBG polymer donor, PBTA-PSF, was developed by Li et al.[20] to match with ITIC to realize complementary absorption. OSCs based on PBTA-PSF:ITIC

Department of Chemical Engineering and Waterloo Institute for Nanotechnology (WIN), University of Waterloo, 200 University Ave West, Waterloo, Ontario, N2L 3G1, Canada. E-mail: yuning.li@uwaterloo.ca; Fax: +1-519-888-4347; Tel: +1-519-888-4567 ext. 31105
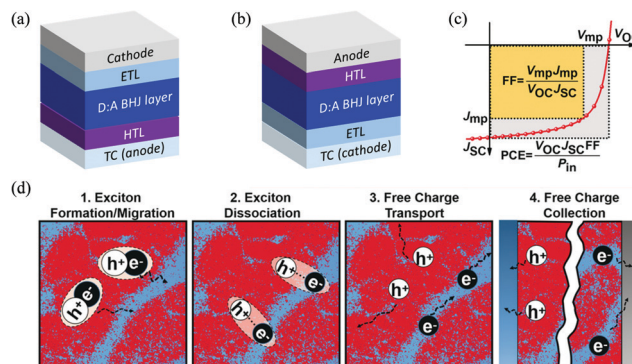
achieved a PCE of 13.91%, which is by far the highest among ITIC-based OSCs. At the same time, incorporation of electron-donating (*e.g.*, methyl in IT-M and IT-DM)[21] and electron-withdrawing (*e.g.*, fluorine in IT-4F)[22] groups, side chain engineering of the ITIC core structure (*e.g. m*-ITIC,[23] ITIC2,[24] ITIC-Th,[25] ITIC-Th1,[26] and IDIC[27]) and optimization of the central core (AOIC,[28] INIC,[29] FNIC,[30] and FOIC[31]) further improved the PCEs when appropriate matching polymer donors were used. In 2019, Yuan *et al.*[9] developed a new high performance NFA, Y6, which has a slightly electron-deficient dithienothiophen[3.2-*b*]-pyrrolobenzothiadiazole core that helps achieve a narrower bandgap of 1.33 eV compared to ITIC. Y6 achieved a very high PCE of 15.7% when blended with donor polymer PM6 owing to the largely improved photocurrent. The record PCE of 18.22% was obtained using Y6 as an acceptor and a novel WBG (1.98 eV) polymer D18 as a donor.[10]

Because BHJ OSCs utilize both donor and acceptor materials in the active layer blend, matching of the optoelectronic properties (frontier molecular orbital (FMO) energy levels, optical absorption, *etc.*) and morphological compatibility (miscibility, phase separation, crystallinity, *etc.*) between donor and acceptor materials are vital for achieving high photovoltaic performance. Compared to the previous review articles that focus on donors[4,32–35] and acceptors,[36–41] respectively, this review will provide a perspective from a different angle by placing emphasis on the matching between donors and acceptors to gain a better understanding of the relationships between the structures, properties, and device performances of representative donors with different prominent acceptor materials.

Firstly, we briefly discuss the working mechanism of OSCs and the general relationships between the molecular structure, nano-/microstructure, properties, and cell performance. Next, we select and classify some high-performance polymer donors that have been used to match the three most representative NFAs, ITIC, IT-4F, and Y6, to achieve PCEs above 10%. Then, we discuss the properties and photovoltaic performances of different donors and provide some guidelines for the design and processing of polymer donors to match with a certain acceptor to achieve high solar cell performance. Finally, a summary of the key findings in terms of the structure–property–cell performance relations of WBG high-performance polymer donors and an outlook for the future development of this type of materials and OSCs in general are provided.

2 Structure–property–cell performance relationships of polymer donors

A typical OSC device is made of an electron donor (commonly a polymer material) and an electron acceptor (commonly a small molecule material) that form a BHJ (photo)active layer with interpenetrating donor and acceptor phases at the nanometer scale.[5] This active layer is sandwiched between an anode and a cathode, which collect the holes and electrons generated in the active layer, respectively. Usually, a hole transport layer (HTL) such as PEDOT:PSS or MoO3 is inserted between the active layer

Fig. 1 (a) Conventional and (b) inverted OSC architectures; (c) solar cell figures-of-merit: short circuit current density (J_{SC}, mA cm^{-2}), open circuit voltage (V_{OC}, V), fill factor (FF) and PCE;[42] and (d) working mechanism of the BHJ layer in OSCs (red and blue areas represent donor and acceptor domains, respectively).[42] Reproduced from ref. 42 with permission from John Wiley and Sons, copyright 2019.

and the anode to facilitate the collection of holes and/or blocking of the electrons. A counterpart electron transport layer (ETL) such as ZnO is placed between the active layer and the cathode to extract electrons and/or block holes. At least one electrode, the cathode or anode, is made of a transparent conductor (TC) such as indium-doped tin oxide (ITO) to allow transmission of light to reach the active layer. The two most commonly adopted OSC architectures are shown in Fig. 1a and b; the conventional one has a TC anode (Fig. 1a), while the inverted one has a TC cathode (Fig. 1b).

The donor material in the active layer is a p-type semiconductor that has relatively high-lying highest occupied molecular orbital (E_{HOMO}) and lowest unoccupied molecular orbital (E_{LUMO}) energy levels, which can stabilize and transport the photogenerated holes. On the contrary, the acceptor material is an n-type semiconductor that has low E_{HOMO} and E_{LUMO} with respect to those of the donor material, which can stabilize and transport the photogenerated electrons. One or both of the donor and acceptor materials should strongly absorb sunlight.

The ultimate photovoltaic performance parameter of an OSC is the PCE or η, which is the percentage of electric energy generated by an OSC out of the total photo energy incident on the active layer conventionally under the standardized AM 1.5G solar spectrum with an intensity of 100 mW cm^{-2}. As shown in Fig. 1c, the PCE is contributed by three factors, the short circuit current density (J_{SC}, mA cm^{-2}), open circuit voltage (V_{OC}, V), and fill factor (FF), obtained from the current density–voltage (J–V) curve of an OSC based on the relationship PCE = ($J_{SC} \times V_{OC} \times$ FF)/P_{in}, where P_{in} is the input power density (100 mW cm^{-2}) of the light source under AM 1.5G conditions. The FF represents the ratio of the product of $J_{mp} \times V_{mp}$ (at the maximum power point on the J–V curve) over the product of J_{SC} and V_{OC}. The FF values are typically 0.5–0.7 for high-performing OSCs. Broader light absorption in the solar spectrum by the active layer increases J_{SC} of the OSC; however, absorption of long wavelength (lower energy) photons results in a lower V_{OC}. Therefore, it is preferred that an active layer absorbs the maximum number of photons in the low wavelength region of ∼400–925 nm

(or energies of ~3–1.34 eV) based on the Shockley–Queisser (S–Q) limit.[43] The portion of sunlight with wavelengths longer than 925 nm can be effectively utilized by developing a tandem OSC with two or more OSC devices combined.[44–47] Research into tandem OSCs has also attracted much attention, but faces numerous issues since deposition of an increased number of well-defined thin active layers and interlayers is very challenging and the charge transport in these multilayered devices is too complicated to control. Although tandem OSCs can achieve higher PCE intrinsically, single junction OSCs still perform better at the moment.

A typical OSC device undergoes photovoltaic processes as follows (Fig. 1d):[48–50] (1) excitons (hole–electron pairs) are produced in the BHJ active layer (in the donor, the acceptor, or both phases) upon absorption of photons from sunlight; (2) the formed excitons diffuse towards the donor–acceptor (D–A) interfaces; (3) the excitons are dissociated into free holes and electrons at the D–A interfaces; and (4) the free holes and electrons travel through the donor and acceptor phases and are collected at the anode and cathode, respectively.
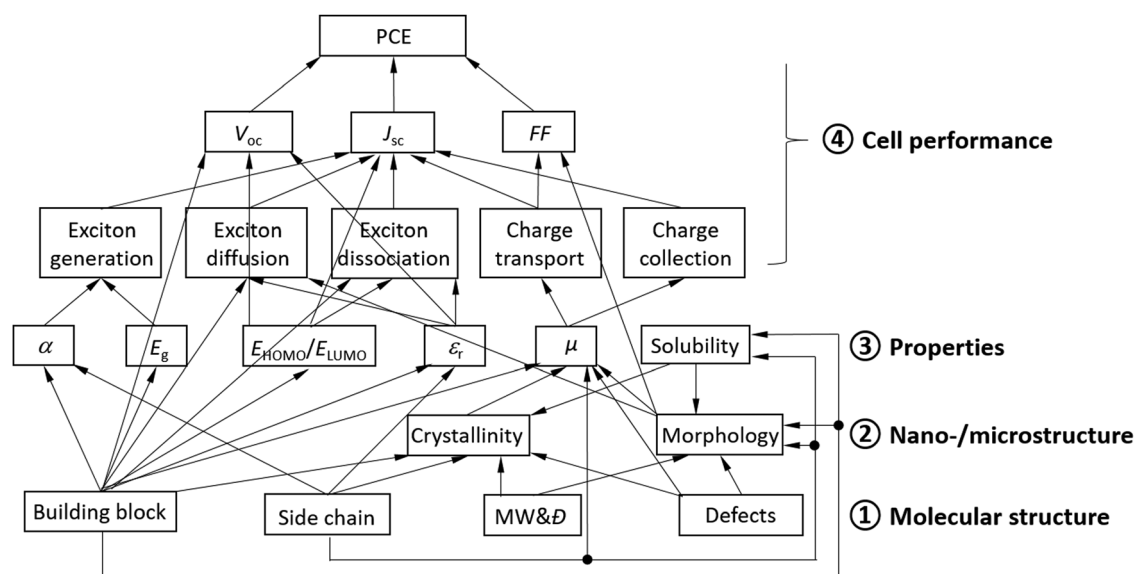
An OSC works through close collaboration of matching donor and acceptors. Fig. 2 illustrates the complex hierarchical relationships between the cell performance, properties, and the molecular and nano-/microstructures of polymer donor materials. Each of the cell performance parameters, J_{SC}, V_{OC}, or FF, is influenced by the photovoltaic processes and material properties, which are ultimately governed by the molecular structure of the polymer donor.

Specifically, J_{SC} is influenced by all the photovoltaic processes, where (1) the number of excitons generated (or photons absorbed) depends largely on the absorption coefficient (α) and the bandgap (E_g) of the donor; (2) exciton diffusion is influenced by the

electronic structure of the conjugated building block, the dielectric constant (ε_r), and the morphology (phase size) of the donor; (3) exciton dissociation is influenced by the electronic structure of the building block, the dielectric constant, and the HOMO and LUMO energy offsets, ΔE_{HOMO} and ΔE_{LUMO}, between the donor and the acceptor; and (4) transport and collection of charge carriers (holes) are determined by the hole mobility (μ_h) of the donor and its balance with the electron mobility (μ_e) of the acceptor. A higher dielectric constant can decrease the exciton binding energy, reducing exciton recombination events. If the donor phase is too large, excitons generated in a region with a distance to the donor and acceptor interface greater than the exciton diffusion length would recombine, leading to a reduction in J_{SC}.

V_{OC} is determined by the difference between E_{LUMO} of the acceptor and E_{HOMO} of the donor as well as the dielectric constant of the donor (and acceptor). A lower E_{HOMO} of the donor helps achieve a high V_{OC}, while a large dielectric constant can reduce the exciton binding energy, which would help reduce the required donor–acceptor energy offset ΔE_{LUMO} for exciton dissociation, achieving a higher V_{OC}. Additionally, the type of building block (electronic structure) often plays a critical role in determining the minimal energy offset required for exciton dissociation, which is directly related to V_{OC}.

The FF is critically influenced by μ_h of the donor and its balance with μ_e of the acceptor. High and balanced μ_h and μ_e of $>10^{-4}$ cm^2 V^{-1} s^{-1} are usually required to achieve a high FF. The morphology of the donor and acceptor blend film also exerts some influence on the FF.[51–53] The carrier mobility is influenced by several factors including the building block, side chain, crystallinity (degree of crystallinity and crystal orientation), film morphology, and amount of structural defects. A face-on lamellar packing

Fig. 2 Scheme of the structure−property−cell performance relationships of a polymer donor for OSCs, where α is the absorption coefficient, ε_r is the dielectric constant (or relative permittivity), μ is the mobility, and MW and Đ are the molecular weight and its distribution, respectively, of the polymer donor. Defects may include terminal groups, homo coupled units in a copolymer, random arrangements of comonomers in a copolymer, regio-irregular units, branching, lightly cross-linked units, oligomers, *etc*. Some or all of these relationships may apply to small molecule donors as well as polymer and small molecule acceptors.

motif for the polymer donor in an OSC is desirable because the charge transport is facilitated through the vertically aligned π–π stacks in this type of crystal motif. Structural defects refer to the terminal groups, homo coupled units in a copolymer, random arrangements of comonomers in a copolymer, regio-irregular units, branching, lightly cross-linked units, oligomers, *etc.* Most structural defects are difficult to determine but are often largely responsible for the poor cell performance since they have adverse effects on the crystallinity, morphology, and charge transport.

Additionally, although not shown in Fig. 2, the properties of the interfaces between the active layer and the HTL and/or ETL determine the efficiency of charge collection, which influences J_{SC}, V_{OC}, and/or the FF.

As can be clearly seen in Fig. 2, the molecular structure fundamentally dictates the nano-/microstructure, properties, and OSC performance of a polymer donor. In particular, the π-conjugated building block used to construct a donor determines all the properties of the donor and the performance parameters of its OSC.

Characterization of the structures and properties of donors, acceptors, and their blends is nontrivial and often encounters great challenges. The widely used methodologies for characterization of their key structural features and properties are briefly described as follows:

(1) E_g: UV-vis-NIR spectroscopy is a simple and the most widely used method to obtain the bandgap of the donor or acceptor using the absorption onset wavelength. However, many materials, particularly polymers, are non-monodisperse and contain disordered structures, resulting in a tail in their absorption spectra with ill-defined onset edges. Measurements using the intersection of the normalized absorption and emission spectra can overcome this issue.[54,55] Alternatively, E_g can be calculated from the onset of the EQE spectrum of the OSC device to minimize the influence of the film morphology.[55]

(2) E_{HOMO}/E_{LUMO}: due to its low-cost and easy operation, cyclic voltammetry (CV) is the most popular method used to estimate E_{HOMO} and E_{LUMO} of the donor or acceptor by using the onset oxidation and reduction potentials, respectively, with a reference having a known E_{HOMO} such as ferrocene (E_{HOMO} = −4.8 eV). However, most donors and acceptors only show the oxidation or the reduction process, respectively, in their CV diagrams. In such cases, the optical bandgap obtained by UV-vis-NIR and E_{HOMO} or E_{LUMO} obtained by CV are usually combined to calculate E_{LUMO} or E_{HOMO} of the donor or acceptor, respectively. The energy levels calculated in this way are often inaccurate because of the large exciton binding energy (up to ∼0.3–1 eV) associated with organic semiconductors. The energy levels determined by CV measurements also have large variations caused by different experimental conditions[56,57] and thus the values reported for the same materials in the literature often vary largely. Nonetheless, the E_{HOMO} and E_{LUMO} values obtained by CV under similar conditions are still very useful and quite reliable for comparing the energy levels of different materials. A combination of ultraviolet photoelectron spectroscopy (UPS) and (low-energy) inverse photoemission spectroscopy (IPES), which are much more expensive and complex to operate compared to the electrochemical setup used for the CV measurements, can also be used to more accurately measure the ionization potential (IP) and electron affinity (EA) values, which correspond to E_{HOMO} and E_{LUMO}, respectively.[56–58]

(3) Crystallinity: wide-angle X-ray diffraction (WAXD) is the standard characterization method to study the crystal structure. When this method is used for thin films (∼100 nm), a grazing incidence geometry is used and the method is termed as GIWAXS or GIXD. GIWAXS can be used to obtain information about the structure, orientation, and ordering of crystallites.[59] Particularly, two-dimensional (2-D) GIWAXS measured with a powerful synchrotron X-ray source can reveal the in-plane and out-of-plane crystal structures of the finely mixed BHJ blend film at the nanometer scale.[60–62]

(4) Morphology/phase separation: atomic force microscopy (AFM) is a surface analysis technique commonly used to probe morphological features (on the nanometer to micron-scale) such as film roughness, phase segregation, and domain size. Transmission electron microscopy (TEM) has also been used to extract information about phase segregation, grain size, and grain connectivity.[59] Resonant soft X-ray scattering (RSoXS) is another evolving method to investigate multi-component, multi-phase systems like the active layer of an OSC, where the contrast of the donor and acceptor components can be tuned by selection of the incident X-ray energy to quantitatively reveal the molecular orientation, domain purity, and domain spacing.[59,60,63,64]

(5) Exciton dissociation, collection, and recombination: the trend of photocurrent density J_{ph} ($J_{ph} = J_L - J_D$, where J_L is the current density under illumination and J_D is the current density in the dark) *versus* effective voltage ($V_{eff} = V - V_0$) gives insights into the charge generation and exciton dissociation characteristics of the device. Here V_0 is called the compensation voltage or the voltage at which J_{ph} = 0 and V is the applied voltage.[65,66] J_{ph} reaches a saturation value (J_{sat}) with increasing V_{eff}, which means that all the photogenerated excitons are dissociated into free carriers and collected by the electrodes with the assistance of large reverse bias. Thus, the exciton dissociation probability, defined as $P_{diss} = J_{SC}/J_{sat}$, reflects the efficiency of exciton dissociation, charge transport, and charge collection.[65] The photoinduced charge transfer efficiency (from the donor to the acceptor or *vice versa*) can also be probed *via* photoluminescence (PL) quenching and decay measurements of the blend films relative to the neat films.[67,68]

The light intensity dependence of V_{OC} can directly indicate the role of trap-assisted recombination (or SRH recombination) *versus* bimolecular recombination under open circuit conditions. For this analysis the slope S_1 (or ideality factor) is defined as $S_1 = \dfrac{q\partial V_{OC}}{kT\partial \ln(P_{light})}$, where q is the elementary charge, k is the Boltzmann constant, T is the absolute temperature, and P_{light} is the light intensity.[69–71] A value of S_1 close to 1 indicates more ideal recombination, whereas values >1 indicate more trap-assisted/SRH recombination in the device.[65,71] Moreover, J_{SC} of an OSC follows a power law dependence with respect to the light intensity (P_{light}) (*i.e.* $J_{SC} \propto P_{light}^{S_2}$), where S_2 is the exponential factor. A value

of S_2 close to unity indicates negligible bimolecular recombination during sweep-out under short circuit conditions.[72–74]

(6) Mobility: the methods for measuring the carrier mobility along the direction vertical to the film (relevant to the charge transport in OSC devices) include space-charge-limited current (SCLC),[75–77] time of flight (TOF),[78,79] carrier extraction by linearly increasing voltage (CELIV),[80,81] photogenerated charges in CELIV (photoCELIV),[82–84] and impedance spectroscopy (IS).[85,86] The SCLC method is most widely used.

(7) Dielectric constant: the exciton binding energy (E_b) estimated by the coulomb interaction depends inversely on the dielectric constant, ε_r: $E_b = q^2/4\pi\varepsilon_0\varepsilon_r r$, where q is the elementary charge, ε_0 is the permittivity of a vacuum, and r is the electron and hole separation distance.[87–92] Therefore, to facilitate exciton diffusion and dissociation, semiconductors with high dielectric constant values are preferred. The capacitance–voltage (C–V) measurement at different frequencies (impedance spectroscopy) of a film sandwiched between two electrodes can be performed to obtain the dielectric constant value.[93–98]

3 High performance WBG polymer donors for NFAs

On the basis of the optical bandgap, polymer donors can be divided into narrow-bandgap or low-bandgap (LBG, bandgap <1.6 eV), medium-bandgap (MBG, 1.6 eV < bandgap < 1.8 eV) and wide-bandgap (WBG, bandgap >1.8 eV).[35] To form complementary absorption with narrow-bandgap NFAs, polymer donors with medium- or wide-bandgaps are needed. Regioregular, head-to-tail poly(3-hexylthiophene-2,5-diyl) (P3HT), which can be synthesized at low cost and shows high μ_h, has been extensively studied as a donor for OSCs based on fullerene acceptors.[50,99,100] However, P3HT has a rather high E_{HOMO} ($ca.$ −5.1 eV), which leads to a low V_{OC}, limiting its solar cell performance.[101–105] On the other hand, donor–acceptor (D–A) polymers containing alternating D and A building blocks on the backbone (Fig. 3) have been extensively developed as donor materials since their optical bandgaps and FMO ($i.e.$ HOMO and LUMO) energy levels can be conveniently tuned through intramolecular charge transfer (ICT) by choosing different D and A units.[32,35,49,106–108] D–A polymers also have

enhanced μ_h due to intermolecular D–A interactions. D–A polymers have shown far better solar cell performances than P3HT in both fullerene and NFA-based OSCs.

The D building blocks mostly used to construct high-performing D–A polymer donors for NFA-based OSCs are thiophene-containing moieties such as thiophene (T), thieno[3,2-b]thiophene (TT), and benzo[1,2-b:4,5-b']dithiophene (BDT) (Fig. 3), owing to their appropriate electron-donating effect to tune E_{HOMO}, lower steric effect to maintain backbone coplanarity, rich chemistry, and good chemical stability. Two or more D blocks are often used to constitute a D unit in a D–A polymer. The frequently used A building blocks include 3-fluorothieno[3,4-b]thiophene-2-carboxylate (FC-TT),[109] thieno[3.4-c]pyrrole-4,6-dione (TPD),[110] benzo[1,2-c:4,5-c']dithiophene-4,8-dione (BDD),[111] difluorobenzotriazole (FBTA),[112] quinoxaline (Qx),[113] and benzothiadiazole (BT),[114] which have an electron-withdrawing ester, imide, fluorine, diketone, and N-containing heterocycle substituting or fused to the thiophene or benzene ring. One or more linear or branched alkyl side chains (R_1 and R_2) are anchored to D, A, or both units to render the polymer soluble and control the packing of polymer chains in the solid state.

By analysing the chemical structures of high-performing D–A polymer donors, one would notice that the majority of them contain the fused ring structure BDT as the D building block. Therefore, in this review, the polymer donors are classified as BDT-based polymers and non-BDT based polymers. The BDT-based polymers are further classified according to the A units: (1) ester or imide substituted building blocks (FC-TT and TPD), (2) diketone derivatives (BDD), (3) N-heterocycles containing sp^2 nitrogen (FBTA and Qx), and (4) other A units (such as BT) (Fig. 3).[4,49,106,114] The structures of the polymer donors discussed in this review are shown in Fig. 4 and 5 based on this classification. Their key optoelectronic properties are listed in Table 1.

As mentioned in the previous section, good matching between donor and acceptor materials is vital for achieving high OSC performance. Therefore, developing a polymer donor material having its photophysical properties matching with those of a certain NFA material on purpose is the most adopted and effective strategy. From a large number of high performance NFAs,[36–38,40] we deliberately select the three most representative ones, ITIC, IT-4F, and Y6, which have been widely

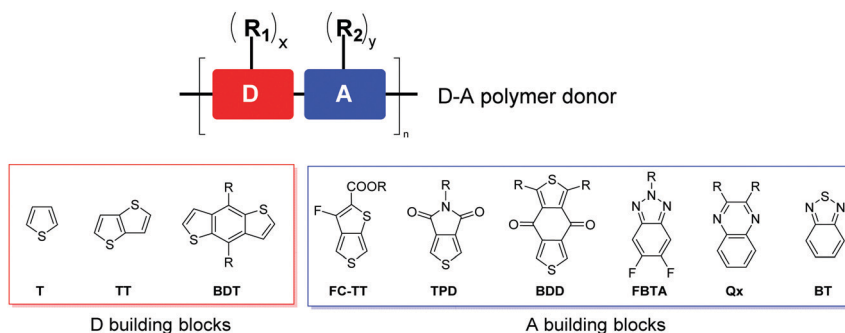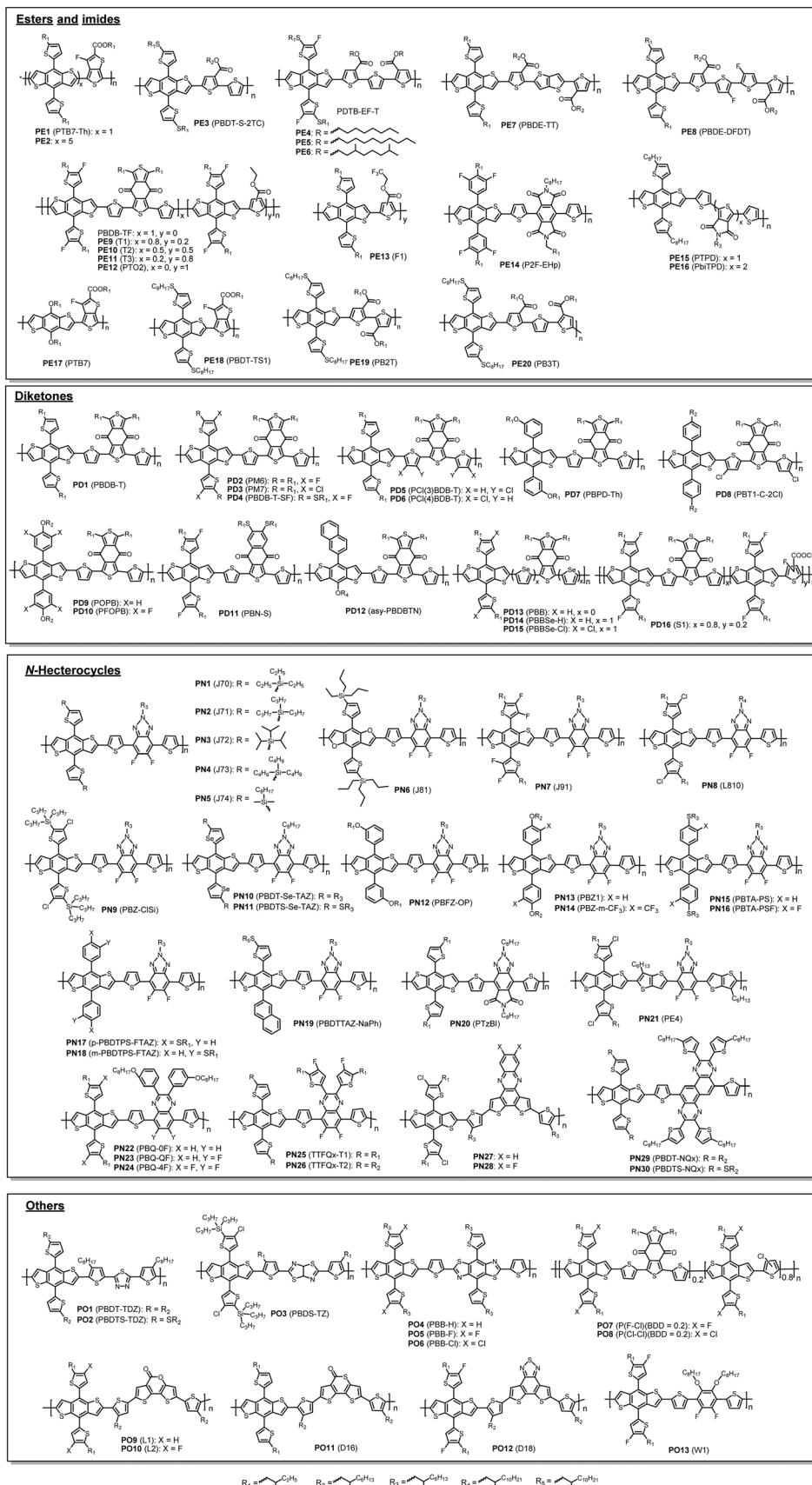**Fig. 3** Chemical structures of representative D and A building blocks used for high performance D–A polymers, where R₁ or R₂ is a side chain chosen from alkyl, alkoxy, alkylthio, and aryl groups, $etc.$, and x or y is an appropriate integer.

Fig. 4 Chemical structures of BDT-based polymer donors discussed in this review. Their properties are listed in Table 1. Names in parentheses are used in the original publications.
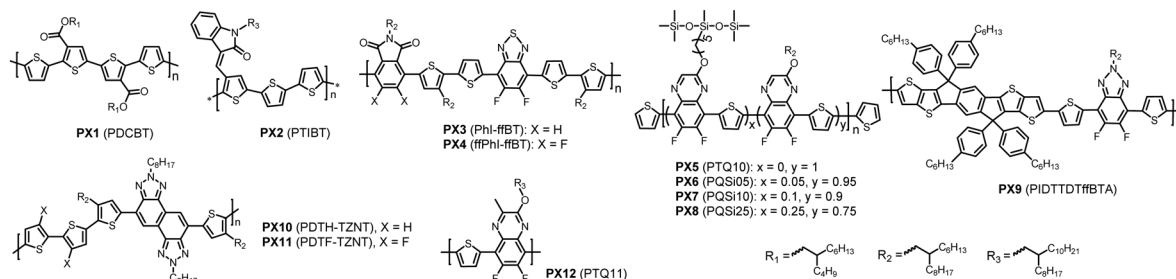
Fig. 5 Chemical structures of non-BDT-based polymer donors discussed in this review. Their properties are listed in Table 2. Names in parentheses are used in the original publications.

used for developing ultrahigh performance (PCE >10%) polymer donors. Their chemical structures are shown in Fig. 6, while their key optoelectronic properties are summarized in Table 2. It should be noted that the reported E_{LUMO} and E_{HOMO} values of the three NFAs, particularly ITIC, vary largely in the literature due to the different measurement conditions as mentioned earlier.

3.1 High performance polymer donors matching with ITIC

As a milestone for NFAs, ITIC was developed in 2015 by Lin et al.,[8] which has an electron-donating indacenodithienothiophene (IDTT) central unit and two electron-withdrawing 2-methylene-(3-(1,1-dicyanomethylene)-indanone) (IC) end groups. Hexyl phenyl side chains are introduced on IDTT to render the material soluble and to suppress the excessive aggregation of molecules in the solid state. ITIC has a good SCLC μ_e of 3.0×10^{-4} cm^2 V^{-1} s^{-1}, which approaches that of fullerene-based acceptors ($\sim 10^{-4}$–10^{-2} cm^2 V^{-1} s^{-1}).[179,180] Its rather high E_{LUMO} (-3.83 eV) helps achieve a high V_{OC}. However, ITIC has a relatively large bandgap (E_g = 1.59 eV), absorbing sunlight below 800 nm, which restricts J_{SC} when a WBG polymer donor is used. Nonetheless, ITIC has been widely used as a model acceptor for the study and development of various polymer donors to provide significant insights into their structure–property–cell performance relationships.

3.1.1 BDT-based polymer donors using ester or imide substituted building blocks as A units. PE1 (PTB7-Th) developed by Liao et al.,[109] which has 2-ethylhexyl-thienyl substituted BDT and FC-TT as D and A units, respectively, is one of the prominent D–A polymer donor materials for OSCs. PE1 has a rather narrow bandgap of 1.58 eV, which was designed as a donor material mainly for matching fullerene-based acceptors that absorb sunlight poorly. When ITIC is used as an acceptor and PE1 as a donor, their absorption profiles are mostly overlapped, which is undesirable for efficient light harvesting. Consequently, the devices based on the PE1:ITIC blend showed a moderate J_{SC} of 14.21 mA cm^{-2} and PCE of 6.8%.[8] In order to improve the light absorption, Kim et al.[116] developed a series of polymer donors, modulating the energy levels and absorption spectra by controlling the BDT/FC-TT (D/A) ratio in the polymer backbone, and found that as the D/A ratio increases, E_{HOMO} downshifts and E_g widens. PE2 with a D/A ratio of 5 : 1 has the optimal E_{HOMO} of -5.37 eV and E_g of 1.74 eV to form

complementary absorption with ITIC. As a result, a higher J_{SC} (17.24 mA cm^{-2}) along with a higher V_{OC} (0.89 V) and FF (67%) and thus a higher PCE of 10.27% were achieved compared to polymers with other D/A ratios. An et al.[117] developed a polymer PE3 (PBDT-S-2TC) having a weaker electron-withdrawing ester substituted thiophene as A units. A non-substituted thiophene spacer was inserted between the ester substituted thiophene and the BDT unit to maintain the planarity of the polymer backbone. In addition, alkylthio chains were introduced onto BDT to lower E_{HOMO} (-5.47 eV). A wide bandgap of 1.94 eV was obtained, which matches that of ITIC very well for complementary light absorption. The devices based on PE3:ITIC showed an improved J_{SC} and a very high V_{OC} of 0.96 V, resulting in a PCE of 10.12%.

3.1.2 BDT-based polymer donors using diketone derivatives as A units. Diketone substituted thiophene building block BDD is a relatively weak electron-accepting (A) unit. The D–A polymers based on BDT and BDD demonstrated excellent photovoltaic performance when paired with narrow bandgap NFAs. PD1 (PBDB-T or PBDTBDD) was the first polymer using BDD developed by Qian et al.,[111] which achieved a PCE of 6.67% when paired with PC$_{61}$BM. Later, Zhao et al.[16] utilized PD1 as a donor to match ITIC since PD1 has a relatively wide bandgap of 1.80 eV. Compared with PE1:ITIC, PD1:ITIC exhibited much improved EQE values in the short wavelength range (300–500 nm) due to the strong absorption by PD1 (Fig. 7b). Therefore, a higher J_{SC} (16.81 mA cm^{-2} for PD1:ITIC vs. 14.21 mA cm^{-2} for PE1:ITIC) was obtained. The improved and more balanced μ_h and μ_e of PD1:ITIC (Table 3) led to the increased FF compared to PE1:ITIC (74.2% vs. 59.1%). In addition, although PD1 showed a similar E_{HOMO} (~ -5.2 eV) to PE1, a higher V_{OC} (0.899 vs. 0.810 V) was obtained owing to its smaller energy loss (E_{loss}). Combining all these improvements, the PD1:ITIC-based device displayed a much higher PCE of 11.21%.[16]

Since then, PD1 was widely used as a star polymer donor and the PCE of devices based on PD1 was further improved to above 13% with a largely improved J_{SC} of up to $\sim$20 mA cm^{-2} through rational design of new NFAs with narrower bandgaps.[181–185]

Fan et al.[133] synthesized a novel polymer donor PD7 (PBPD-Th) with a low-lying E_{HOMO} of -5.42 eV by replacing the 5-alkylthienyl in PD1 with m-alkoxyphenyl on the BDT units. OSCs based on PD7:ITIC have a smaller E_{loss} and therefore achieved a high V_{OC} of up to 1.01 V. Moreover, the polymer also has a wide bandgap of 1.90 eV, resulting in complementary absorption with ITIC. High EQE values (up to $\sim$80%), especially in the short wavelength range

Table 1 Key optoelectronic properties of polymer donors discussed in this review

Polymer	E_{HOMO} (eV)	E_{LUMO} (eV)	E_{opt}^{g} (eV)	λ_{max}^{a} (nm)	λ_{onset}^{a} (nm)	$\alpha_{f\text{-max}}$ ($\times 10^5$ cm^{-1})	$\alpha_{s\text{-max}}$ ($\times 10^5$ M^{-1} cm^{-1})	μ_{h} ($\times 10^{-4}$ cm^2 V^{-1} s^{-1})	Ref.
PE1	−5.22	−3.64[b]	1.58	625, 690	785	1.0[b]		11.1[115]	8
PE2	−5.37	−3.63	1.74	549	711				116
PE3	−5.47	−3.58	1.94		538	1[b]			117
PE4	−5.51	−3.59	1.93	540	642	0.79		1[b]	118
PE5	−5.50	−3.59	1.93	544	642	0.81		1.2[b]	118
PE6	−5.54	−3.6[b]	1.94	536	639	0.72		0.73	118
PE7	−5.49		1.97	538	629	0.726[b]	0.607[b]	1.11[b]	119
PE8	−5.5		1.94	544	639	0.758	0.65[b]	3.4	119
PE9	−5.48	−3.63	1.83		678				120
PE10	−5.51	−3.63	1.87		663				120
PE11	−5.55	−3.62	1.93		642				120
PE12	−5.59	−3.67	1.99	533	623	0.769			121
PE13	−5.50	−3.53	1.97	577	630	1.25			122
PE14	−5.46	−3.13	1.82		700				123
PE15	−5.05	−3.19	1.86	562, 613	667				110
PE16	−5.20	−3.45	1.78	586, 630	670				110
PE17	−5.15[7]	−3.31[7]	1.60	614, 674	775			5.8	124
PE18	−5.33	−3.52	1.51	655, 705	820			109[125]	126
PE19	−5.63	−3.63	2.24	416	554		0.362	0.028	127
PE20	−5.39	−3.37	1.96	540	633		0.557	0.94	127
PD1	−5.23	−3.18	1.80	581, 622	689	0.86[128]		1.11[128]	111
PD2	−5.50	−3.56	1.80	570, 614	689			2.97	129
PD3	−5.52	−3.57	1.79	577, 606[128]	692	0.99[128]		6.67[128]	130
PD4	−5.40	−3.60	1.80	626	688	1.08			22
PD5	−5.54	−3.27	2.11	479	588				131
PD6	−5.48	−3.47	1.78	622	697				131
PD7	−5.42	−3.36	1.90	556	652	0.697	0.331	7.56[132]	133
PD8	−5.59	−3.53	1.85	627	671			13.4	134
PD9	−5.38	3.57	1.81	577	685	0.531		3.96	135
PD10	−5.5	−3.64	1.86	612	667	0.845		15.1	135
PD11	−5.48	−3.52	1.75	632, 625	710				136
PD12	−5.41	−3.58	1.83	621	678	1			137
PD13	−5.54	−3.59	1.95	551	636		2.63		138
PD14	−5.16	−3.45	1.71	657	726		5.54		138
PD15	−5.36	−3.63	1.73	645	718		7.37		138
PD16	−5.52	−3.72	1.80		689				139
PN1	−5.37	−2.91	1.99	538, 580	623	0.89			140
PN2	−5.40	−3.24	1.96	528, 573	633	0.96			141
PN3	−5.42	−3.29	1.98	533, 578	626	0.7			140
PN4	−5.46	−2.92	1.98	535, 576	626	0.69			140
PN5	−5.56	−3.06	1.99	526	623	0.68			140
PN6	−5.42	−2.93	1.93	553, 600	642				142
PN7	−5.50	−3.02	2.00	536	620	0.98			143
PN8	−5.57	−3.58	1.99		622	0.838			144
PN9	−5.56	−3.50	1.94	536	640	1.08		10.2	145
PN10	−5.23	−3.43	1.92	372, 550, 597	646		0.564		146
PN11	−5.29	−3.49	1.9	374, 550, 602	653		0.854		146
PN12	−5.33	−3.01	1.99	530, 575	623			7.28	147
PN13	−5.27	−3.06	1.96	539	632	0.523		7.23	148
PN14	−5.49	−3.22	1.99	533	623	0.651		7.86	148
PN15	−5.34	−3.4	1.94		639		0.73	5.59	20
PN16	−5.52	−3.54	1.98	537, 577	626		0.7	6.01	20
PN17	−5.32	−3.36	1.96	544, 589	633		0.62		149
PN18	−5.40	−3.40	2.00	542, 579	621		0.60		149
PN19	−5.44	−3.49	1.95		636				150
PN20	−5.34	−3.46	1.81	600	685	0.7			151
PN21	−5.42	−3.49	1.93	590	642		1.1		152
PN22	−5.18	−3.48	1.70	636	755	0.545	0.454		113
PN23	−5.34	−3.62	1.72	600	737	0.731	0.515		113
PN24	−5.49	−3.7	1.79	583	716	0.859	0.617		113
PN25	−5.31	−3.62	1.69	361, 435, 652	736	4.73			153
PN26	−5.38	−3.67	1.71	361, 443, 635	724	4.64			153
PN27	−5.33	−2.94	1.86		667			2.00	154
PN28	−5.46	−3.16	1.79		693			7.74	154
PN29	−5.24	−3.42	1.80	374, 450, 632	689		29		155
PN30	−5.31	−3.46	1.81	368, 452, 636	685		34.5		155
PO1	−5.35	−2.78	2.07		599		25		156
PO2	−5.39	−2.79	2.09		593		31		156
PO3	−5.41	−3.46	1.95	550	636	1.13			157

Table 1 (continued)

Polymer	E_{HOMO} (eV)	E_{LUMO} (eV)	E_{opt}^g (eV)	$\lambda_{max}{}^a$ (nm)	$\lambda_{onset}{}^a$ (nm)	$\alpha_{f\text{-max}}$ ($\times 10^5$ cm^{-1})	$\alpha_{s\text{-max}}$ ($\times 10^5$ M^{-1} cm^{-1})	μ_h ($\times 10^{-4}$ cm^2 V^{-1} s^{-1})	Ref.
PO4	−5.47	−3.47	2.00	580	620		1.53	0.614	158
PO5	−5.47	−3.57	2.00	579	620		1.39	0.143	158
PO6	−5.64	−3.64	2.00	580	620		1.27	0.27	158
PO7	−5.62	−3.74	1.88	361, 565	660		0.896		159
PO8	−5.67	−3.77	1.90	364, 565	653		0.944		159
PO9	−5.45	−2.79	1.96		634			7.20[160]	161
PO10	−5.52	−2.91	1.96		634				161
PO11	−5.48	−2.83	1.95	551, 588	636			11.9	160
PO12	−5.51	−2.77	1.98	584, 559	626			15.9	10
PO13	−5.36	−2.81	2.16	502, 535	574				162
PX1	−5.31	−3.00	1.90	551	653	1.05			163 and 164
PX2	−5.41	−3.61	1.80	576	689				165
PX3	−5.55	−3.80	1.75	583	708				166
PX4	−5.63	−3.86	1.77	580	700				166
PX5	−5.54	−2.98	1.92		646				167
PX6	−5.67	−3.74	1.93		642				168
PX7	−5.62	−3.71	1.92		646				168
PX8	−5.61	−3.69	1.92		646				168
PX9	−5.25	−3.63	2.02	583	627	0.818	1.2		169
PX10	−5.16	−3.10	1.94	541	639		0.804		170
PX11	5.24	−3.23	1.84	559, 617	674		1.28		170
PX12	−5.52	−2.76	1.95		637				171

^{a} Obtained from film absorption, $\alpha_{f\text{-max}}$ represents the maximum extinction coefficient in a thin film, and $\alpha_{s\text{-max}}$ represents the maximum extinction coefficient in solution. ^{b} The numbers of significant digits of the data are adopted from the original references and may be different.
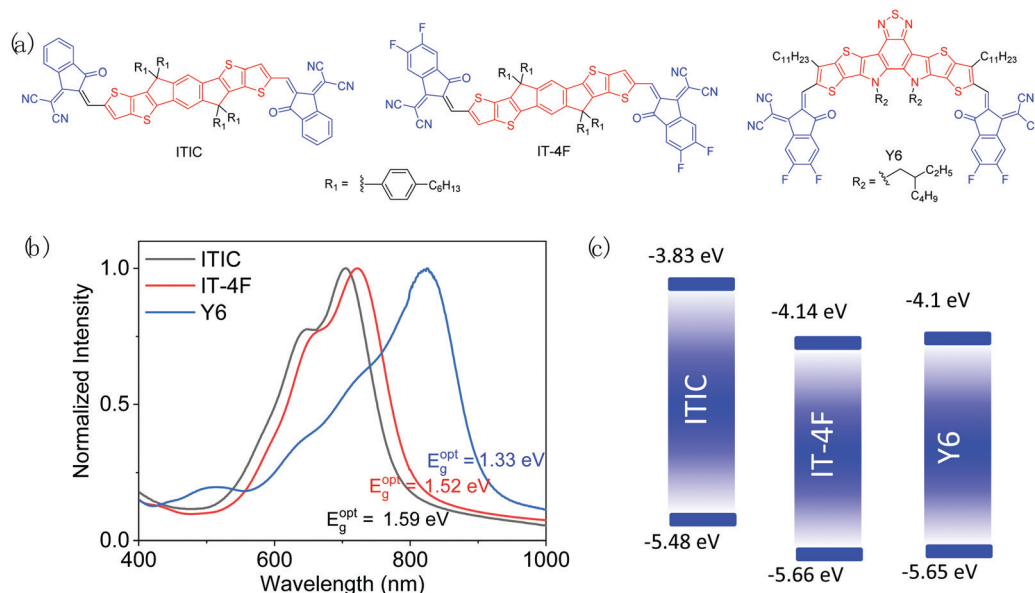

Fig. 6 Chemical structures (a), thin-film UV-vis-NIR spectra (b), and energy diagrams (c) of ITIC, ITIC-4F, and Y6.

(400–550 nm, ∼70% on average), were obtained, which led to a high J_{SC} (18.1 mA cm^{-2}) and thus a high PCE of 10.8%.

By breaking the symmetry of the BDT unit, *i.e.*, introducing different side chains on two sides of BDT, Li *et al.*[137] developed an asymmetric polymer donor **PD12** (asy-PBDBTN). Alkoxyl side chains were introduced to modulate the absorption spectra and solubility of the polymers (achieving enhanced light-harvesting ability and compatibility with acceptor materials, simultaneously), while a β-position linked naphthalene unit with high

ionization potential and low electron density was introduced to lower E_{HOMO}. **PD12** showed a low-lying E_{HOMO} of −5.41 eV, a wide bandgap of 1.83 eV, and high light-harvesting ability with an $\alpha_{f\text{-max}}$ of up to 1×10^5 cm^{-1}. OSCs based on the **PD12:ITIC** blend displayed a high PCE of 10.5%.[137]

3.1.3 BDT-based polymer donors using N-heterocycles as A units. By copolymerizing BDT with FBTA, Li's group developed a J-series (J40 to J91) of polymer donors.[19,23,142,143,186–188] Through optimization of the backbone conjugation dimension, variation

Table 2 Key optoelectronic properties of NFAs discussed in this review

NFA	E_{HOMO} (eV)	E_{LUMO} (eV)	E_{opt}^{g} (eV)	$\lambda_{max}{}^{a}$ (nm)	$\lambda_{onset}{}^{a}$ (nm)	α_{f-max} ($\times 10^5$ cm^{-1})	α_{s-max} ($\times 10^5$ M^{-1} cm^{-1})	$\mu_e{}^{b}$ ($\times 10^{-4}$ cm^2 V^{-1} s^{-1})	Ref.
ITIC	−5.48	−3.83	1.59	702	780	1.10	1.72	3 (150 °C)	8
	−5.60	−3.85							172
	−5.64	−4.04							173
	−5.70	−4.00							174
IT-4F	−5.66	−4.14	1.52	726	803	1.16	2.10	5.05 (150 °C)	22
	−5.71	−4.15							175
	−5.69	−4.17							176
Y6	−5.65	−4.10	1.33	821	931	1.07	2.39	2.35 (100 °C), 14.8 (150 °C), 33.0 (200 °C)	9
	−5.70	−4.08							177
	−5.60	−4.10							130
	−5.65	−4.11							175

^{a} Obtained from film absorption. ^{b} SCLC mobility obtained at the annealing temperature shown in brackets.
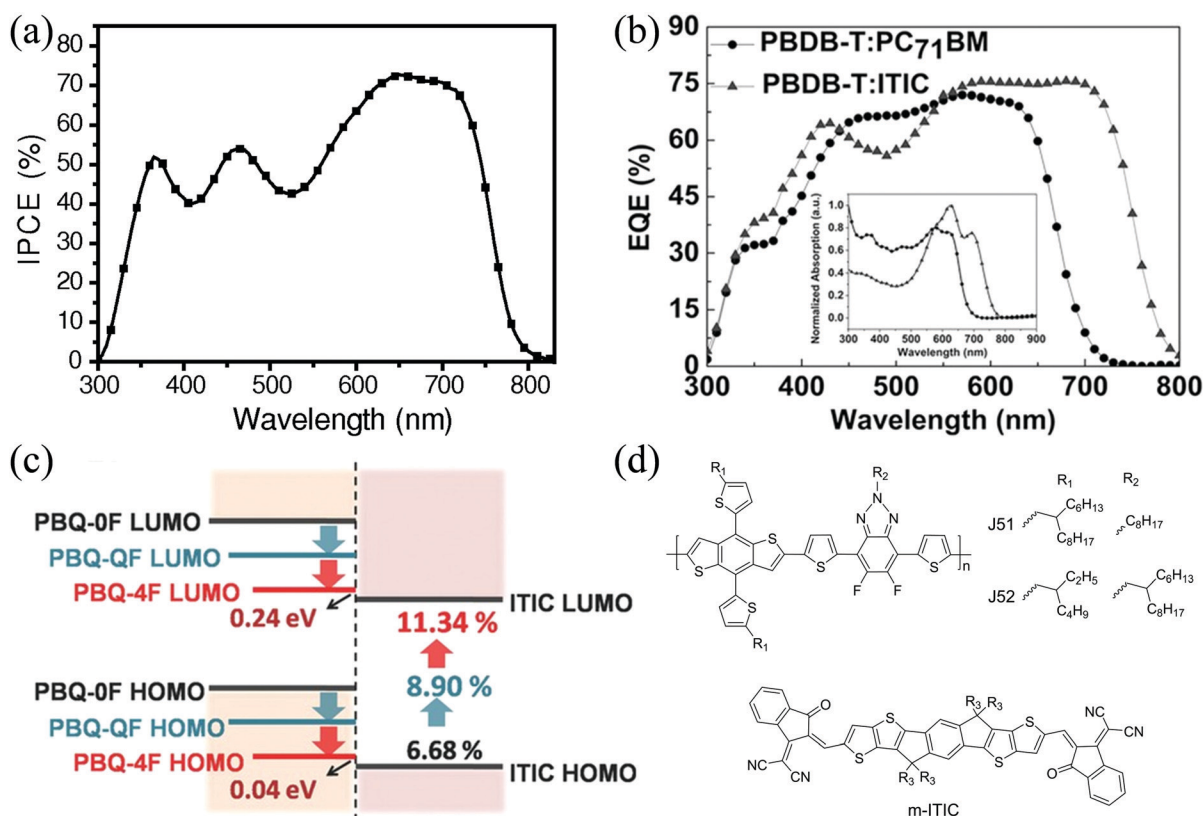

Fig. 7 (a) IPCE spectrum of the **PE1**:ITIC-based OSC, reproduced from ref. 8 with permission from John Wiley and Sons, copyright 2015; (b) EQE curve of the **PD1**:ITIC-based OSC, reproduced with permission,[16] copyright 2016, John Wiley and Sons; (c) energy level schematic of **PN22**, **PN23**, **PN24**, and ITIC, reproduced from ref. 113 with permission from John Wiley and Sons, copyright 2016; and (d) chemical structures of J51, J52 and m-ITIC.

of side chains, and fluorination of BDT units, the PCE was increased from 6.48% to 9.53% for polymers J40 to J61.[18,19,186] Furthermore, trialkylsilyl substituents were used to replace the alkyl chains on the BDT unit in J52 (Fig. 7d) to produce **PN2** (J71), where the bond interaction of the low-lying σ* orbital of the Si atom with the π* orbital of the aromatic units would result in stabilization of E_{LUMO} and lowering of E_{HOMO} (−5.40 eV).[141] **PN2** also showed a higher extinction coefficient of 0.96 × 10^5 cm^{-1} than that (0.73 × 10^5 cm^{-1}) of its analogue J52 with alkyl chains. OSCs based on the

PN2:ITIC blend showed a J_{SC} of 17.32 mA cm^{-2}, a V_{OC} of 0.94 V, and a FF of 69.77%, resulting in a high PCE of 11.41%. The annealed blend film showed high and balanced μ_h and μ_e of 3.78 × 10^{-4} cm^2 V^{-1} s^{-1} and 3.07 × 10^{-4} cm^2 V^{-1} s^{-1}, respectively, which contributed to the high J_{SC} and high FF. It should be noted that ΔE_{HOMO} between **PN2** and ITIC is only 0.11 eV, which suggests a high hole transfer efficiency from the acceptor to the donor.

The same research group systemically investigated the effects of the size and configuration of side chains (**PN1–PN5**

Table 3   Photophysical properties and OSC performance of polymer donor:ITIC blends

Blend	μ_h, μ_e ($\times 10^{-4}$ cm^2 V^{-1} s^{-1}), $\mu_{e(h)}/\mu_{h(e)}$[a]	EQE[b] (%)	V_{OC} (V)	J_{SC} (mA cm^{-2})	FF (%)	PCE (%)	S_1[c]	S_2[d]	P_{diss} (%)	Ref.
PE1:ITIC	0.43, 1.1, 2.63	72.6[e]	0.81[e]	14.21[e]	59.1[e]	6.8[e]				8
PE2:ITIC	0.298, 0.346, 1.16		0.89	17.24	67[e]	10.27[e]	1.15	0.927		116
PE3:ITIC	2.11, 2.83, 1.34		0.96	16.4[e]	64.3	10.12				117
PD1:ITIC	2.1, 3.13, 1.49		0.899[e]	16.81	74.2	11.21				16
PD7:ITIC	2.59, 2.25, 1.15	79[e]	1.01	18.1	59	10.8	1.01		86.8[e]	133
PD12:ITIC	2.91, 2.67, 1.09	76	0.942	16.81	66.3	10.5				137
PN2:ITIC	3.78, 3.07, 1.23	76.5	0.94	17.32	69.77[e]	11.41				141
PN1:m-ITIC	1.93, 8.4, 4.34	81.5	0.92	18.09	69.82	11.62				140
PN2:m-ITIC	1.09, 4.66, 4.28		0.944	18.09	70.59	12.05				140
PN3:m-ITIC	0.46, 3.77, 8.2		0.962	16.35	65.03	10.23				140
PN4:m-ITIC	0.38, 3.12, 8.2		0.974	16.45	66.87	10.71				140
PN5:m-ITIC	0.24, 4.0, 16.66		0.99	15.89	61.18	9.63				140
PN6:ITIC	3.66, 4.06, 1.11		0.95	15.27	73.08	10.6		0.94[e]	96[e]	142
PN6:m-ITIC	4.08, 4.67, 1.15		0.96	16.48	69.83	11.05		1.01	97	142
PN7:m-ITIC	1.016, 3.002, 2.95		0.984	18.03	65.54	11.63		1.028[e]		143
PN10:ITIC	13.3, 8.7, 1.53	78	0.81	18.63	66.7	10.07				146
PN11:ITIC	24.7, 20.3, 1.22	80	0.84	19.51	75.1	12.31				146
PN12:ITIC	3.25, 2.62, 1.24	82	0.91	18.7	61.8	10.5	1.26	0.94		147
PN13:ITIC	4.36, 2.98. 1.46	72	0.74	15.7	49.8	5.8	1.8	0.93		148
PN14:ITIC	5.08, 3.96, 1.28	84	0.94	18.4	60.2	10.4	1.32	0.99		148
PN15:ITIC	4.94, 3.81, 1.3		0.94	18.23	69.19	11.85		0.93	93	20
PN16:ITIC	5.46, 5.25 1.04		1.01	18.51	74.4	13.91		0.97	95.4	20
PN17:ITIC	1.15, 3.93, 3.42		0.89	18.12	67.37	10.86		0.92	93	149
PN18:ITIC	4.16, 3.16, 1.32		0.95	18.76	73.85	13.16		0.96	95	149
PN19:ITIC	2.132, 3.397, 1.59	83	0.936	18.21	67.8	11.56				150
PN20:ITIC	0.86, 0.23, 3.74		0.87	18.29	64.34	10.24		0.971		178
PN22:ITIC	7.8, 8.2, 1.05		0.69	16.16	59.91	6.68	1.33	0.952	88	113
PN23:ITIC	8.6, 8.3, 1.04		0.83	17.16	62.49	8.9	1.23	0.981	95	113
PN24:ITIC	10, 10, 1		0.95	17.87	66.8	11.34				113
PN25:ITIC	1.13, 1.6, 1.42	75	0.9	16.88	69.24	10.52		0.96	97.2	153
PN26:ITIC	0.139, 0.68, 4.89	66.4	0.94	13.75	56.11	7.22		0.9	96.3	153
PN29:ITIC	1.8, 0.56, 3.2	71	0.87	16.21	64.6	9.11	1.37	0.92	89	155
PN30:ITIC	2.5, 1.3, 1.9	78	0.92	17.86	69.8	11.47	1.25	0.96	94	155
PO1:ITIC	8.5, 6.8, 1.25		1.01	17.15	67.7	11.72				156
PO2:ITIC	12, 7.2, 1.67		1.1	17.78	65.4	12.8				156
P3HT:ITIC	1.4, 1.68, 1.2		0.52	4.22	56.93	1.25	1.14	0.99	92	163
PX1:ITIC	2.4, 16.3. 6.79	75	0.94	16.5	65.67	10.16				163
PX2:ITIC	1.79, 1.20, 1.49	72.5	0.66	15.19	57	5.72				165

[a] μ_e/μ_h or μ_h/μ_e, whichever is >1. [b] Maximum value of EQE. [c] Obtained from $V_{OC} \sim S_1 \log(P_{light})$. [d] Obtained from $J_{ph} \sim S_2 \log(P_{light})$. [e] The numbers of significant digits of the data are adopted from the original references and may be different.

(J70–J74)) on the OSC performance of their blends with an ITIC isomer, m-ITIC (Fig. 7d), which has *meta*-alkyl-phenyl substitution and shows a higher degree of self-organization and crystallinity.[140] Among these five polymers, they found that the film absorption coefficients and hole mobilities decreased, while the E_{HOMO} values downshifted, with increasing length and complexity (linear to branched) of the alkyl substituents, which resulted in decreased J_{SC} and FF, but enhanced V_{OC}. Among this series of polymers, **PN2** showed the highest PCE of 12.05%.

When the thiophene rings in the BDT units were replaced with furan rings to form benzo[1,2-b:4,5-b']difuran (BDF), the resultant polymer **PN6** (J81), which is analogous to **PN1**, showed a lower E_{HOMO} of −5.42 eV, achieving a high PCE of 11.05% when paired with m-ITIC.[142] On the other hand, when the thiophene side chains on the BDT units in **PN1** were disubstituted with electron-withdrawing fluorine atoms, the formed **PN7** (J91) exhibited a further down-shifted E_{HOMO} (−5.50 eV).[143] OSCs based on **PN7**:m-ITIC showed an even higher PCE of 11.63%, resulting from the increased V_{OC} and J_{SC}.

Compared with thiophene, selenophene shows lower aromaticity and enhanced ground-state quinoid resonance character, which would result in improved planarity, increased effective conjugation length, and a lowered bandgap of the polymer. In addition, improved interchain interactions and charge carrier transport could also be observed for selenophene-containing polymers because of the stronger heteroaromatic interactions induced by selenium atoms.[189,190]

When the thiophene side chains on BDT units of J51 (Fig. 7d)[18] were replaced with selenophene side chains, the obtained **PN10** (PBDT-Se-TAZ) exhibited similar E_{HOMO} (−5.23 vs. −5.26 eV) and E_g (1.92 vs. 1.91 eV) to those of J51.[146] However, the **PN10**:ITIC blend showed much higher μ_h and μ_e of 13.3 × 10^{-4} cm^2 V^{-1} s^{-1}/8.7 × 10^{-4} cm^2 V^{-1} s^{-1} compared to J51 (4.32 × 10^{-4} cm^2 V^{-1} s^{-1}/3.74 × 10^{-4} cm^2 V^{-1} s^{-1}), indicating much improved charge transport in **PN10** due to the presence of selenophene side chains. Consequently, the devices based on **PN10**:ITIC showed improved performance (J_{SC} = 18.63 mA cm^{-2}; V_{OC} = 0.81 V; FF = 66.7%; and PCE = 10.07%) compared to the devices based on J51:ITIC (J_{SC} = 16.47 mA cm^{-2}; V_{OC} = 0.82 V; FF = 0.69; and PCE = 9.26%).

Once the alkyl selenophene side chains in **PN10** were changed to alkylthio selenophene side chains, the resulting **PN11** (PBDTS-Se-TAZ) showed a slightly lower E_{HOMO} (-5.29 *vs.* -5.23 eV), a slightly narrower E_g (1.90 *vs.* 1.92 eV), and a higher α_{s-max} (the maximum extinction coefficient in solution) (8.54×10^4 *vs.* 5.64×10^4 M^{-1} cm^{-1}) compared to **PN10**.[146] The **PN11**:ITIC blend showed greatly enhanced and more balanced μ_h and μ_e (2.47×10^{-3} cm^2 V^{-1} s^{-1}/2.03×10^{-3} cm^2 V^{-1} s^{-1}), and thus less recombination and higher exciton dissociation and charge collection efficiencies (Table 3). Therefore, **PN11**:ITIC displayed higher V_{OC} (0.84 V), J_{SC} (19.51 mA cm^{-2}), FF (75.1%), and PCE (12.31%).[146]

Replacing thiophene side chains on BDT units with benzene side chains is a strategy to lower E_{HOMO} as mentioned previously (in the case of **PD7**). Li *et al.*[147] synthesized **PN12** (PBFZ-OP) based on *meta*-alkoxy-phenyl-substituted BDT (BDT-*m*-OP) and FBTA units. E_{HOMO} was lowered to -5.33 eV compared with J52 with thiophene side chains ($E_{HOMO} = -5.21$ eV).[19] **PN12** also showed a wider bandgap of 1.99 eV than that of J52 (1.96 eV), which could form a more complementary absorption profile when paired with ITIC. Much improved cell performance ($V_{OC} = 0.91$ V; $J_{SC} = 18.7$ mA cm^{-2}; FF = 0.618; and PCE = 10.50%) was obtained for the **PN12**:ITIC blend compared with the J52:ITIC blend ($V_{OC} = 0.73$ V; $J_{SC} = 13.1$ mA cm^{-2}; FF = 0.578; and PCE = 5.50%).

Later, the same group changed the alkoxy substituent of the phenyl side chains on BDT from the *meta* to the *para* position (**PN13** (PBZ1)) and introduced a CF_3 group on the *meta* position (**PN14** (PBZ-*m*-CF_3)) to further lower E_{HOMO}.[148] Compared with **PN12**, both polymers showed a similar bandgap (1.96 eV for **PN13** and 1.99 eV for **PN14**). **PN13** showed a slightly higher E_{HOMO} of -5.27 eV, while **PN14** showed a much lower E_{HOMO} of -5.49 eV. Between these two polymers, **PN14** exhibited a higher α_{f-max} (6.51×10^4 *vs.* 5.23×10^4 cm^{-1}), smaller π–π distance (3.60 *vs.* 3.69 Å), and slightly higher μ_h (7.86×10^{-4} *vs.* 7.23×10^{-4} cm^2 V^{-1} s^{-1}). The **PN14**:ITIC blend showed a much higher PCE of 10.4% than **PN13**:ITIC (5.8%).

Li *et al.*[20] synthesized another similar pair of polymers, using *p*-alkylthiophenyl without or with fluorine at the *meta* position as the side chains to BDT units, **PN15** (PBTA-PS) and **PN16** (PBTA-PSF), respectively. A wider bandgap of 1.98 eV and a much lower E_{HOMO} of -5.52 eV were obtained for **PN16** compared to **PN15** ($E_g = 1.94$ eV; and $E_{HOMO} = -5.34$ eV). When paired with ITIC, both showed a high V_{OC} of ~ 1 V and high J_{SC} of ~ 18 mA cm^{-2}. The device based on **PN16**:ITIC showed a much improved FF of 74.4% (*vs.* 69.19% for **PN15**:ITIC) due to its higher and more balanced μ_h and μ_e, less recombination and higher exciton dissociation and charge collection efficiencies (Table 3). Therefore, while **PN15**:ITIC showed a PCE of 11.85%, the **PN16**:ITIC blend exhibited a much higher PCE of 13.91%, which is the highest value achieved among ITIC-based OSCs by far. The results indicate that alkylthiophenyl substitution on BDT is very beneficial for the improvement of the OSC performance of the resulting polymer. On other hand, fluorination on the phenyl ring could further enhance the cell performance.

Chen *et al.*[149] conducted a further study to compare the effects of the position of alkylthio chains on phenyl (*para* or

meta) on the OSC performance of this type of polymers. Compared with the polymer with alkylthio chains at the *para* position (**PN17** (*p*-PBDTPS-FTAZ)), the one with the chains at the *meta* position, **PN18** (*m*-PBDTPS-FTAZ), showed a lower E_{HOMO} and blue-shifted absorption. When blended with ITIC, higher and more balanced μ_h and μ_e, less recombination, and higher exciton dissociation and charge collection efficiencies were observed. As a result, a higher PCE of 13.16% was achieved for the **PN18**:ITIC blend compared to **PN17**:ITIC (10.86%).

Liu *et al.*[150] developed a series of polymer donors having the BDT unit asymmetrically substituted with an alkylthiolthienyl group on one side and a phenyl and naphthyl (or *p*-biphenyl) on the other. It was found that the long alkylthiolthienyl substituent can lower E_{HOMO}, broaden the absorption spectrum, and improve the solubility. The non-substituted rigid aryl substituent (benzene, naphthalene and *p*-biphenyl) can be regarded as a lever arm to disturb the ITIC phase and weaken the polymer chain entanglements, which helps realize a favorable morphology even without post-treatment. When the polymer with the naphthyl substituent, **PN19** (PBDTTAZ-NaPh), was blended with ITIC, the obtained film exhibited a homogeneous morphology with finely separated domains. A high PCE of 11.56% was obtained with this blend film.

Fan *et al.*[151] used alkylimide to replace the two fluorine atoms on benzotriazole in FBTA to form an acceptor unit, TZBI, which was used to make a copolymer **PN20** (PTZBI) with BDT. **PN20** possesses a more delocalized electron distribution across the conjugated backbone, which lowers E_{HOMO} and enhances the carrier mobility. A high PCE of 10.24% was obtained with the **PN20**:ITIC blend.[178]

Qx is another moderate electron-accepting N-heterocylic building block that is widely used in constructing WBG conjugated D–A polymer donors for OSCs.[34,191] Zheng *et al.*[113] conducted a systematic study on the correlations between the energy level alignment, performance, and device physics by finely tuning the FMO energy levels through introducing different numbers of F atoms. Three polymers with 0, 2 and 4 F atoms in one repeat unit, **PN22** (PBQ-0F), **PN23** (PBQ-QF) and **PN24** (PBQ-4F), respectively, were synthesized. With an increasing number of F atoms, E_{LUMO} and E_{HOMO} downshifted gradually. **PN24** showed the lowest E_{HOMO} of -5.49 eV, and so achieved the highest V_{OC} of 0.95 V when blended with ITIC. It should be noted that although the energy offsets between **PN24** and ITIC are very small ($\Delta E_{HOMO} = 0.04$ eV and $\Delta E_{LUMO} = 0.24$ eV), the devices exhibited efficient charge generation, which is an outstanding advantage over fullerene-based OSCs, which require a large energy offset of > 0.3 eV for efficient exciton dissociation.[87,121,192–195] A high J_{SC} (17.87 mA cm^{-2}) and FF (66.80%) were also obtained owing to the stronger interchain interaction, resulting in higher carrier mobility. Therefore, this blend afforded a high PCE of 11.34%. A mixture of non-halogenated solvents THF and IPA was used for processing the active layer, demonstrating the potential for environment-friendly fabrication of OSCs.

The thiophene donor building block possesses higher electron density and smaller steric hindrance, providing better

absorption and charge transport to the resulting polymer than the benzene building block. Xu et al.[153] replaced the phenyl side chains of Qx units with alkyl substituted fluorothiophene and introduced alkyl chains with different lengths on BDT units to form polymers **PN25** (TTFQx-T1) and **PN26** (TTFQx-T2). When paired with ITIC, the blend based on **PN25** with short alkyl chains on BDT showed higher and more balanced μ_h and μ_e, less recombination, and higher exciton dissociation and charge collection efficiencies (Table 3). Therefore, it exhibited a much improved J_{SC} (16.88 vs. 13.75 mA cm^{-2}) and FF (69.24% vs. 56.11%) and thus a higher PCE (10.52% vs. 7.22%) compared with the **PN26**:ITIC blend. This result indicated that the alkyl chain length has notable influences on the photovoltaic properties.

Building blocks with an enlarged planar aromatic structure are favorable for ordered molecular packing and thus improved photovoltaic performance.[196,197] In addition, alkylthiothiophene side chains can further reduce E_{HOMO}. Based on these principles, Yu et al.[155] used quinoxalino[6,5-f]quinoxaline (NQx) as A units and introduced alkylthiophene and alkylthiothiophene as side chains of BDT units to form **PN29** (PBDT-NQx) and **PN30** (PBDTS-NQx), respectively. The latter showed a lower E_{HOMO} (-5.31 vs. -5.24 eV) and higher $\alpha_{s\text{-max}}$ (3.45×10^6 vs. 2.9×10^6 M^{-1} cm^{-1}). The devices based on **PN30**:ITIC showed a higher PCE of 11.47% than those based on **PN29**:ITIC (9.11%).

3.1.4 Other BDT-based polymers and non-BDT-based polymers. 1,3,4-Thiadiazole (TDZ) is highly polarizable with large electron affinity, which can be used as an excellent A building block for constructing high-performance D–A copolymers for OSCs.[198] Xu et al.[156] copolymerized TDZ with BDT units to form **PO1** (PBDT-TDZ) and **PO2** (PBDTS-TDZ), which employed alkylthiophene and alkylthiothiophene as side chains of BDT units, respectively. **PO2** showed a slightly lower E_{HOMO} (-5.39 vs. -5.35 eV) and higher $\alpha_{s\text{-max}}$ (3.1×10^6 vs. 2.5×10^6 M^{-1} cm^{-1}) than **PO1**. The **PO2**:ITIC blend showed higher crystallinity and higher μ_h and μ_e (Table 3). As a result, a higher PCE of 12.8% for **PO2**:ITIC was obtained compared to **PO1**:ITIC (11.72%).

The primary driving force for OSC research is the potential of manufacturing OSC products at a much lower cost than Si solar cells. The synthetic complexity (SC), which considers the number of reaction steps, purification methods, reaction yields, use of hazardous materials, etc., can be used to estimate the synthetic cost of the materials.[199] Although BDT-based polymers showed high cell performance, the tedious synthesis of the BDT tin monomer alone would result in a high SC ($>30\%$). P3HT can be synthesized with a very low SC (7.75%), but its OSC performance is rather low mainly due to its high E_{HOMO}. To overcome the disadvantageous high-lying E_{HOMO} of P3HT, Zhang et al.[163] introduced electron-withdrawing carboxylate substituents on half of the thiophene rings in the backbone of polythiophene to form **PX1** (PDCBT). E_{HOMO} of **PX1** down-shifted to -5.26 eV, much lower than that of P3HT (~ -5.0 eV). Therefore, a much higher V_{OC} for the **PX1**:ITIC blend was obtained (0.94 V for **PX1**:ITIC vs. 0.52 V for P3HT:ITIC). Furthermore, in contrast to P3HT, **PX1** predominantly adopted a face-on crystal orientation in the blend film, which is beneficial for the charge transport in the vertical direction. Combined with lower recombination and higher exciton dissociation and charge collection efficiencies, the devices based on **PX1**:ITIC showed a largely improved PCE of up to 10.16% compared to the PCE of 1.25% for P3HT:ITIC, demonstrating the potential of polythiophenes for achieving high photovoltaic performance. Another strategy to improve the photovoltaic performance of polythiophenes is to design novel D–A polymer donors with conjugated D and A units located at the backbone and side chains, respectively, which was demonstrated by He et al.[165] with a polymer **PX2** (PTIBT) that has an electron-rich polythiophene backbone and electron-accepting indolin-2-one side chains. The polymer with the so-called donor-backbone–acceptor-side-chain structure showed a very high dielectric constant of 7.70 due to the largely separated donor and acceptor units, which is beneficial for the exciton diffusion to the donor:acceptor interface and subsequent dissociation into free electrons and holes. Although the PCE was moderate (5.72%) for the **PX2**:ITIC blend, this novel polymer design showed potential for achieving high performance through increasing the dielectric constant of the material. It should be mentioned that this polymer could be synthesized with a very low SC of 23.5%.

3.2 High performance polymer donors matching with IT-4F

In the previous section it was pointed out that ITIC has a rather large bandgap (1.59 eV), which can only absorb sunlight below 800 nm. In 2017, Zhao et al.[22] introduced F atoms on the end IC groups of ITIC to form IT-4F (Fig. 6). Compared with ITIC, IT-4F has a narrower bandgap (1.52 eV) with red-shifted absorption and higher $\alpha_{s\text{-max}}$ and $\alpha_{f\text{-max}}$ values along with lower E_{LUMO} (-4.14 eV) and E_{HOMO} (-5.66 eV). OSCs based on IT-4F have demonstrated improved photovoltaic performance compared to ITIC-based devices.

3.2.1 BDT-based polymer donors using ester or imide substituted building blocks as A units. We have mentioned earlier that in addition to F or Cl atoms, ester groups can also downshift E_{HOMO} and enhance the interchain interactions. Li et al.[118] developed a novel polymer donor by using BDT and ester substituted thiophene (EST) as the D and A units, respectively, to form PDTB-EF-T polymers and investigated the influence of different alkyl chains of the ester groups on the cell performance. Three polymers with different alkyl chains, **PE4**, **PE5** and **PE6**, showed similar FMO energy levels and bandgaps, while the polymer with linear decyl chains (**PE5**) exhibited the strongest absorption and highest μ_h due to the most planar backbone and strongest intermolecular interactions. When using IT-4F as an acceptor, similar V_{OC} values were obtained for the three blends, but the **PE5**:IT-4F blend showed the highest chain ordering, the highest and most balanced μ_h and μ_e, the most suppressed recombination, and the highest exciton dissociation and charge collection efficiencies (Table 4), resulting in the highest J_{SC}, FF, and PCE (14.2%).

Despite the high performance of EST-based polymer donors, the steric effects of the ester groups lead to twisted backbone structures. To improve the planarity of this kind of polymers, Li et al.[119] inserted a TT unit and a difluorinated 2,2′-bithiophene (DFDT) unit as a π-spacer between two EST units to form **PE7**

Table 4 OSC device performances based on blends of polymer donors and IT-4F

Blend	μ_h, μ_e ($\times 10^{-4}$ cm^2 V^{-1} s^{-1}), $\mu_{e(h)}/\mu_{h(e)}$[a]	EQE[b] (%)	V_{OC} (V)	J_{SC} (mA cm^{-2})	FF (%)	PCE (%)	S_1[c]	S_2[d]	P_{diss} (%)	Ref.
PE4:IT-4F	1.51, 4.12, 2.7		0.896[e]	20.05[e]	64	11.5		1[e]	93[e]	118
PE5:IT-4F	3.46, 3.95, 1.1	82[e]	0.9[e]	20.73	76	14.2		0.91[e]	95	118
PE6:IT-4F	0.632, 2.35, 3.7		0.904	20.31	61	11.2			93	118
PE7:IT-4F	0.553, 0.918, 1.7		0.883	19.34	65	11.1	1.11	0.95	86	119
PE8:IT-4F	1.94, 2.21, 1.1	82	0.865	21.83	75	14.16			96	119
PE9:IT-4F (CB/DIO)			0.899	21.5[e]	78	15.1				120
PE9:IT-4F (THF/NMP)			0.887	21.1	76	14.2	1.31	0.91		120
PE12:IT-4F	16, 6.4, 2.5	87	0.91[e]	21.5	75	14.7				121
PE13:IT-4F	2.64, 2.21, 1.19	83	0.933	20.6	70	13.5				122
PD4:IT-4F	3.25, 4.32, 1.33	83	0.88	20.5	71.9	13	1.2	0.999[e]	93	22
PD2:IT-4F	9.76, 7.15, 1.37		0.84	22.2	72.5	13.5			97.8[e]	200
PD3:IT-4F			0.86	21.8	77	14.4				201
PD5:IT-4F			0.88	0.88	23.74	0.18				131
PD6:IT-4F	0.0574, 0.0516, 1.1		0.84	20.6	71.09	12.33		0.93		131
PD8:IT-4F	10.7, 5.4, 1.98		0.85	19.74	76	12.7			97	134
PD9:IT-4F	3.15, 1.86, 1.69		0.72	16.8	51.2	6.2	1.23	0.999	86.4	135
PD10:IT-4F	4.43, 3.46, 1.28		0.89	20.4	64.5	11.7	1.19	0.977	93.7	135
PD11:IT-4F	2.87, 1.02, 2.81	82.2[e]	0.891	21.03	69.9	13.1		0.931	98	136
PN8:IT-4F	8.98, 4.53, 1.98	80	0.790	20.76	73.5	12.1	1.24	0.998	96	144
PN9:IT-4F	2.11, 2.57, 1.22	81	0.93	19.2	71.5	12.8	1.11	0.935	93.7	145
PO3:IT-4F	5.01, 2.71, 1.85		0.804	20.05	73.1	12.01		1.01	95.7	157
PO4:IT-4F	0.84, 4.1, 5		0.82	21.5	67.8	12.0	1.37	0.96	98	158
PO5:IT-4F	3.4, 2.9, 1.17		0.92	22.4	71.1	14.7	1.21	0.98	98	158
PO6:IT-4F	7.0, 5.7, 1.22		0.94	21.8	72.5	14.8	1.24	0.99	98	158
PO7:IT-4F	4.91, 5.60, 1.14	85	0.813	24.06	65.0	12.70				159
PO8:IT-4F	6.76, 6.11, 1.11	85	0.891	23.40	67.0	13.97				159
PX3:IT-4F	14.2, 5.49, 2.59		0.91	19.41	76	13.31				166
PX4:IT-4F	5.61, 2.21, 2.54		0.94	19.01	71	12.74		0.95		166
PX5:IT-4F			0.931	20.52	66.96	12.79		0.938	95.9	168
PX6:IT-4F	2.60, 2.24, 1.16		0.911	21.42	69.48	13.56		0.929	96.4	168
PX7:IT-4F	1.31. 1.81, 1.38		0.911	21.12	69.28	13.32		0.924	96.3	168
PX8:IT-4F	0.7, 1.64, 2.4		0.891	21.07	69.75	13.1		0.94	95.5	168

[a] μ_e/μ_h or μ_h/μ_e, whichever is >1. [b] Maximum value of EQE. [c] Obtained from $V_{OC} \sim S_1 \log(P_{light})$. [d] Obtained from $J_{ph} \sim S_2 \log(P_{light})$. [e] The numbers of significant digits of the data are adopted from the original references and may be different.

(PBDE-TT) and **PE8** (PBDE-DFDT). The latter had much better coplanarity due to the F⋯S non-bonding interaction, which enhanced the optical absorption, strengthened the interchain π–π interaction, and shortened the π–π distance, resulting in much improved μ_h (3.4 × 10^{-4} cm^2 V^{-1} s^{-1} for **PE8** vs. 1.11 × 10^{-4} cm^2 V^{-1} s^{-1} for **PE7**). When blended with IT-4F, more compact and ordered packing, higher and more balanced μ_h and μ_e, suppressed recombination, and higher exciton dissociation and charge collection efficiencies were obtained for the **PE8**:IT-4F blend. As a result, a higher PCE of 14.16% was obtained for **PE8**:IT-4F compared to **PE7**:IT-4F (11.1%).

PE12 (PTO2) having BDT and EST building blocks as D and A units, respectively, was developed by Yao et al.[121] **PE12** showed a low-lying E_{HOMO} of −5.56 eV, so a high V_{OC} of 0.91 V was achieved when blended with IT-4F. ΔE_{HOMO} between **PE12** and IT-4F is very small (0.07 eV), but highly efficient exciton dissociation and charge separation were still observed. Through DFT calculations, they found that **PE12** and IT-4F have a very large difference in molecular electrostatic potential (ESP) and the induced intermolecular electric field (IEF) may assist the separation of excitons. A high J_{SC} and FF and thus a high PCE of up to 14.7% were achieved for **PE12**:IT-4F.

Although using ternary blends can enhance the optical absorption and thus the PCE, a third component would lead to a much more complicated blend morphology and mixing of incompatible materials would result in severe molecular disorder, hampering the device performance. As an alternative, incorporation of an additional building block to form a ternary polymer (terpolymer) is a feasible method to optimize the optical absorption and energy levels. Cui et al.[120] introduced different amounts of the EST building block into **PD2** to tune the electron-withdrawing property of the A units. The introduction of EST led to larger steric effects, but the desired downshifted E_{HOMO} and broadened absorption. When **PE9**, which contained 20% EST in the A units, was used as a donor to blend with IT-4F, PCEs of 15.1% for the blend film processed from a chlorobenzene (CB)/1,8-diiodooctane (DIO) system and 14.2% for the blend film processed from a non-halogenated tetrahydrofuran (THF)/N-methyl pyrrolidone (NMP) system were obtained.

Compared with F atoms, CF$_3$ groups also have strong electron-withdrawing ability, and can be more easily introduced. Meng et al.[122] introduced CF$_3$ groups at the terminal of the ester groups of **PE12**, but eliminating F atoms on the thiophene side chains, to form **PE13** (F1). Compared with the polymer without the CF$_3$ groups (F0), **PE13** showed a lower E_{HOMO}, increased absorption, and stronger intermolecular interactions. DFT results also suggested that the introduction
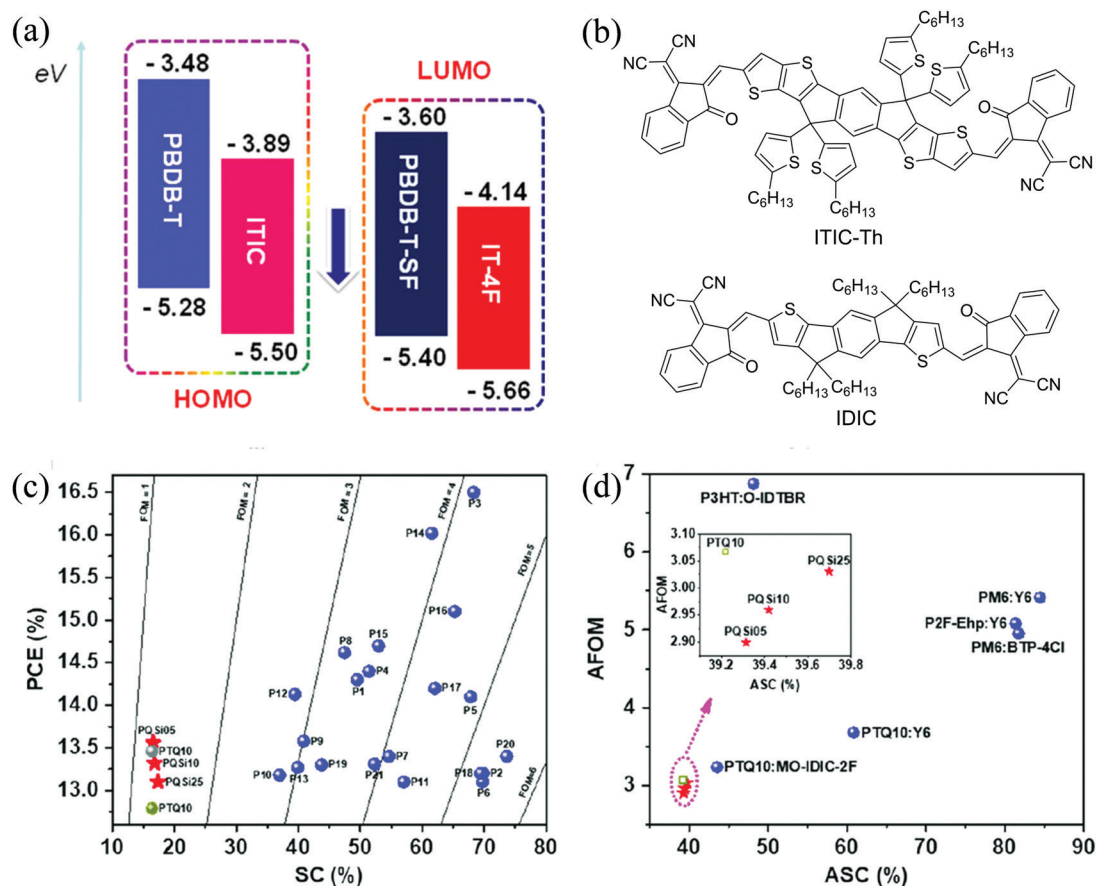
of CF$_3$ groups did not lead to larger steric hindrance. OSC devices based on the **PE13:IT-4F** blend showed a much higher PCE of 13.5% than those based on the polymer without CF$_3$ groups (4.9%).

3.2.2 BDT-based polymer donors using diketone derivatives as A units. To match with IT-4F, a polymer donor **PD4** (PBDB-T-SF) was developed by simultaneously introducing F and alkylthio chains on the thienyl substituents of BDT in **PD1** (Fig. 8a).[22] **PD4** has a low-lying E_{HOMO} of −5.40 eV and a wide bandgap of 1.80 eV. The **PD4:IT-4F** blend showed a broader optical absorption band, a higher absorption coefficient, and more ordered intermolecular arrangements in comparison with the **PD1:ITIC** blend. As a result, a high PCE of 13% was obtained for **PD4:IT-4F**. In addition, the devices maintained a high PCE of 11.99% after storage for 1700 h in an N$_2$-filled glovebox, demonstrating the excellent device stability. Fan *et al.*[200] used **PD2** as a donor to blend with IT-4F. The strong crystallinity of **PD2** is beneficial for forming a favorable blend morphology to suppress recombination, which led to a high PCE of 13.5%.

Although fluorination is an effective way to tune the FMO energy levels and molecular packing, the tedious synthesis and low yield would limit further industrial application. Zhang

et al.[201] replaced the F atoms in **PD2** with Cl atoms to form **PD3** (PM7). The lower material cost, much shorter synthetic route, and simpler purification of the intermediates made the synthetic cost of **PD3** potentially much lower than that for **PD2**. In addition, introduction of Cl atoms resulted in higher dipole moments and a more delocalized HOMO of the polymer. Except for a slightly lowered E_{HOMO}, the absorption profile, aggregation effect in solution, and film morphology of **PD3** were similar to those of **PD2**. However, a very high PCE of 14.4% was obtained for the **PD3:IT-4F** blend. Furthermore, the same group introduced Cl atoms at different positions on the thiophene spacers of **PD1** to form **PD5** and **PD6**.[131] These differences had dramatic effects on the molecular packing and thus the photovoltaic performance. **PD6** showed a large redshift of 140 nm in its absorption spectrum with a narrower bandgap of 1.78 eV than **PD5** (E_g = 2.11 eV). On the other hand, **PD5** showed much weaker aggregation behaviour, poorer crystallinity, and almost no SCLC mobility due to the larger steric effects. Consequently, an extremely low PCE of 0.18% was obtained for **PD5:IT-4F**, while a much higher PCE of 12.33% was obtained for **PD6:IT-4F**.

Ye *et al.*[134] developed a similar polymer donor, **PD8** (PBT1-C-2Cl), using *p*-alkylphenyl groups as the side chains of BDT units

Fig. 8 (a) Molecular energy levels of PBDB-T, ITIC, PBDB-T-SF and IT-4F, reproduced from ref. 22 with permission from the American Chemical Society, copyright 2017; (b) chemical structures of ITIC-Th and IDIC; (c) the maximum PCE *versus* SC of some polymers; and (d) the AFOM *versus* ASC for the active layers showing high PCEs in single-junction OSCs. Reproduced from ref. 168 with permission from John Wiley and Sons, copyright 2017.

and introducing Cl atoms on the 4-position of the thiophene spacer. Compared with **PD6**, **PD8** exhibited a lower-lying E_{HOMO} of -5.59 eV and a slightly wider bandgap of 1.85 eV. The devices based on **PD8**:IT-4F showed much enhanced μ_h and μ_e (Table 4), and therefore a higher FF of 76% than that of **PD6**:IT-4F (70.5%). Thus, a slightly higher PCE of 12.7% was obtained for **PD8**:IT-4F.

Li et al.[135] used fluorinated alkoxyphenyl groups as side chains of BDT to form **PD10** (PFOPB). Compared with **PD9** (POPB) without fluorination, **PD10** possessed a deeper E_{HOMO}, higher extinction coefficient, and higher μ_h. Higher crystallinity with a smaller phase separation size was obtained for the **PD10**:IT-4F blend. As a result, a higher PCE of 11.7% was obtained. They also replaced the BDD unit with a naphtho[2,3-c]thiophene-4,9-dione (NTDO) unit as the latter possesses a planar structure and offers several loci in the phenyl ring for functionalization.[202] Alkylthio side chains were also introduced to lower E_{HOMO}. The resultant **PD11** (PBN-S) showed a low-lying E_{HOMO} of -5.48 eV, high crystallinity, and ordered molecular packing. Benefitting from the efficient charge separation, high and balanced μ_h and μ_e, ordered molecular packing, appropriate aggregations, and complementary absorption in the active layer, the devices based on **PD11**:IT-4F demonstrated a high PCE of 13.10%. The devices showed good storage stability, retaining 88% of their initial PCE after 100 days storage in a nitrogen-filled glove-box.[136] A PCE of up to 10.21% with an active area of 100 mm^2 was obtained, demonstrating the area scalability of these OSCs.

3.2.3 BDT-based polymer donors using N-heterocycles as A units. Polymers based on BDT and FBTA units showed high cell performance when blended with ITIC. Liao et al.[144] optimized the side chains on FBTA units and introduced Cl atoms onto BDT units to synthesize **PN8** in order to optimize the planarity, aggregation property, light-harvesting ability, energy levels, molecular face-on orientation, and packing distance. The devices based on **PN8** (L810):IT-4F showed a high PCE of 12.1%. Based on the success of chlorination and alkylsilyl substitutions (**PN1**–**PN5**), Su et al.[145] employed these two strategies to obtain **PN9** (PBZ-ClSi). This polymer showed a lower E_{HOMO}, an increased absorption coefficient, and improved charge mobility compared to its parent polymer PBZ.[203] The **PN9**:IT-4F blend showed high and balanced μ_h and μ_e, less recombination, and higher charge separation and collection efficiencies, which led to a high PCE of 12.8% in OSCs.

3.2.4 Other BDT-based polymers and non-BDT-based polymers. Although alkylsilyl substituted BDT-based polymers showed a high PCE over 10%, the devices based on this kind of polymers always suffered from low FF. The thiazolo[5,4-d]thiazole (TZ) building block with excellent planarity and electron-withdrawing properties has been used to construct polymer donors with a high FF of ~75% when blended with NFAs.[204] Xu et al.[157] synthesized **PO3** (PBDS-TZ) using alkylsilyl BDT and TZ as the D and A unit, respectively, to lower E_{HOMO} and improve the backbone planarity and charge transport. The **PO3**:IT-4F blend film showed high and balanced μ_h and μ_e, resulting in suppressed recombination and high exciton

dissociation and charge collection efficiencies. A high FF of 73.1% was achieved, which contributed largely to a high PCE of 12.01% for the **PO3**:IT-4F-based OSCs.

Benzo[1,2-d:4,5-d']bis(thiazole) (BBT) is a weak electron-withdrawing building block for structural modification to improve μ_h and the stability of polymer donors. Wen et al.[158] employed BDT and BBT both with alkylthiophene side chains as the D and A units, respectively, and introduced different halogen atoms (F and Cl) on the thiophene side chains of the BDT units to form polymers **PO4** (PBB-H), **PO5** (PBB-F), and **PO6** (PBB-Cl). When blended with IT-4F, similarly high J_{SC} values were obtained for these three polymers due to their high extinction coefficient and favourable face-on orientation. Fluorination and chlorination lowered the energy levels, improved the charge transport, and suppressed recombination, which led to a higher V_{OC} and FF. Among them, the devices based on **PO6**:IT-4F showed the highest PCE of 14.8%.

Jeon et al.[205] synthesized a simple, potentially low cost polymer donor P(Cl) by using BDT and 3-chlorothiophene as D and A units, respectively. P(Cl) showed a high PCE of 12.14% when blended with ITIC-Th (Fig. 8b). Although fluorination or chlorination on BDT units can improve the device performance, the molecular weight and solubility would decrease with the introduction of F or Cl atoms. To achieve a balance between molecular weight and solubility, they introduced a small amount of BDD as A units to form two ternary polymers, **PO7** (P(F–Cl)(BDD = 0.2)) and **PO8** (P(Cl–Cl)(BDD = 0.2)).[159] Both polymers had higher molecular weights and showed better solubility than those without BDD units (P(F–Cl) and P(Cl–Cl)). High PCE values of 12.7% and 13.97% were obtained for **PO7**:IT-4F and **PO8**:IT-4F-based OSCs, respectively, when processed from eco-friendly solvents. For the BDT-based polymers, the π-orbitals of BDT show a high degree of localization, which is not beneficial for intramolecular charge carrier delocalization, resulting in low mobilities.[49] Replacing the BDT units with imide-oligothiophene can greatly improve the carrier mobility.[206] Yu et al.[166] developed two polymer donors with D–A₁–D–A₂ structures using thiophenes as D units and difluorobenzothiadiazole (ffBT) and phthalimide (PhI) or difluoro phthalimide (ffPhI) as the two A units to form **PX3** (PhI-ffBT) and **PX4** (ffPhI-ffBT), respectively. These polymers are highly crystalline and show a high field effect μ_h of 0.6–0.9 cm^2 V^{-1} s^{-1}. **PX4** has predominantly edge-on polymer backbone orientation, which is not beneficial for the charge transport in the vertical direction in an OSC. Therefore, the **PX4**:IT-4F blend showed a lower SCLC μ_h than that of **PX3**:IT-4F. The **PX3**:IT-4F blend exhibited tighter π–π stacking and less recombination and higher exciton dissociation and charge collection efficiencies. A higher PCE of 13.31% was obtained for **PX3**:IT-4F compared to **PX4**:IT-4F (12.74%).

Sun et al.[167] synthesized **PX5** (PTQ10) using thiophene and difluorinated quinoxaline as D and A units, respectively, in only two reaction steps with a high overall yield of 87.4%. **PX5** possessed a low E_{HOMO} of -5.54 eV and a wide bandgap of 1.92 eV and showed a high PCE of 12.70% when an ITIC derivative, IDIC (Fig. 8b), was used as an acceptor. Yuan et al.[168] adopted a combinatory side chain strategy to keep the low synthetic cost while achieving high efficiencies.

Siloxane terminated alkoxy and alkoxy groups with different ratios were incorporated to modify **PX5** to obtain **PX6** (PQSi05), **PX7** (PQSi10), and **PX8** (PQSi25). Compared with **PX5**, these polymers showed similar bandgaps, but increased E_{HOMO} with an increasing amount of siloxane-containing side chains, which resulted in a decrease in V_{OC} of OSCs based on their blends with IT-4F. Among them, the **PX6**:IT-4F blend showed the highest and most balanced μ_h and μ_e, least recombination and highest exciton dissociation and charge collection efficiencies (Table 4). Consequently, the devices based on the **PX6**:IT-4F blend showed the highest PCE of 13.56%. In addition, all these polymers have a low SC and high figure-of-merit (FOM) (a parameter that is related to the manufacturability of OSCs using this material) as well as a low average SC (ASC) and average FOM (AFOM) when considering the cost of acceptors (Fig. 8c and d).[168]

3.3 High performance polymer donors matching with Y6

In 2019, Yuan et al.[9] developed a novel NFA, Y6, which incorporates a BT-based fused-unit as the central core and 2FIC as the end groups (Fig. 6). Compared with ITIC and IT-4F, Y6 has a narrower bandgap of 1.33 eV with a red-shifted optical absorption onset at 931 nm, and a higher α_{s-max} of 2.39×10^5 M^{-1} cm^{-1}. The bandgap of Y6 is almost ideal for achieving the maximum PCE based on the Shockley–Queisser limit for single junction solar cells (at an E_g of 1.34 eV).[43] The E_{LUMO} of −4.1 eV was between those of ITIC and IT-4F. Y6 is the acceptor that has achieved the record PCE of 18.22% so far.[10]

3.3.1 BDT-based polymer donors using imide derivatives as A units. Since benzotriazole functionalized with an imide moiety (TzBI) can lower E_{HOMO}, Ding et al.[123] developed a novel polymer donor, **PE14** (P2F-EHp), using this building block, which showed a low-lying E_{HOMO} of −5.46 eV and a wide bandgap of 1.82 eV. They also synthesized a new NFA based on Y6 with a shorter central core, BTPT-4F, which has a wider bandgap of 1.45 eV than that of Y6 (1.33 eV). Therefore, the blend of **PE14**:Y6 showed much wider absorption (with an onset at about 900 nm) than that of **PE14**:BTPT-4F. In addition, the former blend exhibited more ordered packing with a face-on orientation and smaller domain size for efficient exciton

dissociation. A much higher PCE of 16.02% was obtained for **PE14**:Y6 than that for **PE14**:BTPT-4F (1.09%). Another widely used electron-accepting building block containing the imide group is thieno[3,4-c]pyrrole-4,6-dione (TPD). Huang et al.[110] synthesized two polymers, **PE15** (PTPD) and **PE16** (PbiTPD), using BDT and one or two TPD units in one repeat as the D and A units, respectively. The biTPD unit has a larger planar structure and stronger electron-withdrawing ability. **PE16** showed a lower E_{HOMO} (−5.20 vs. −5.05 eV) and red-shifted absorption. The blend of **PE16**:Y6 showed favourable face-on backbone orientation and higher crystallinity. A significantly enhanced PCE of 14.2% was obtained for **PE16**:Y6 compared to **PE15**:Y6 (5.9%).

3.3.2 BDT-based polymer donors using diketone derivatives as A units. **PD2** is the first donor used to match Y6.[9] The blend showed a wide EQE curve with the onset at about 900 nm due to the small bandgap of Y6, resulting in a high J_{SC} of 25.3 mA cm^{-2} (Fig. 9a). Together with a high V_{OC} of 0.83 V and a high FF of 74.8%, a high PCE of 15.7% was obtained. The PCE is relatively independent of the blend film thickness with a high PCE of 13.6% even at a thickness of 300 nm. Although **PD2**:Y6 achieved a high PCE over 15%, V_{OC} was relatively low due to the small difference between E_{HOMO} of **PD2** and E_{LUMO} of Y6. Since **PD3** with chlorine substitution had a lower E_{HOMO} than **PD2** (−5.52 vs. −5.50 eV), the devices based on **PD3**:Y6 showed a higher V_{OC} of 0.897 V, leading to a much higher PCE of 17.037%.[130]

As **PD8** had shown high performance when blended with IT-4F, Xie et al.[207] employed Y6 as an acceptor to blend with **PD8**. Compared with the IT-4F blend, the Y6 blend showed much wider EQE curves due to the red-shifted absorption of Y6. A much improved J_{SC} of 26.1 mA cm^{-2} was obtained for **PD8**:Y6 than that for **PD8**:IT-4F (19.74 mA cm^{-2}), contributing to the higher PCE of 16.1% of the former.

As mentioned previously, the π-spacer between D and A units plays a key role in improving the coplanarity of the polymer backbone. As a π-spacer, selenophene with lower aromaticity than thiophene can endow a more planar backbone and longer conjugation. Chao et al.[138] synthesized three polymers, **PD13** (PBB), **PD14** (PBBSe-H), and **PD15** (PBBSe-Cl), to
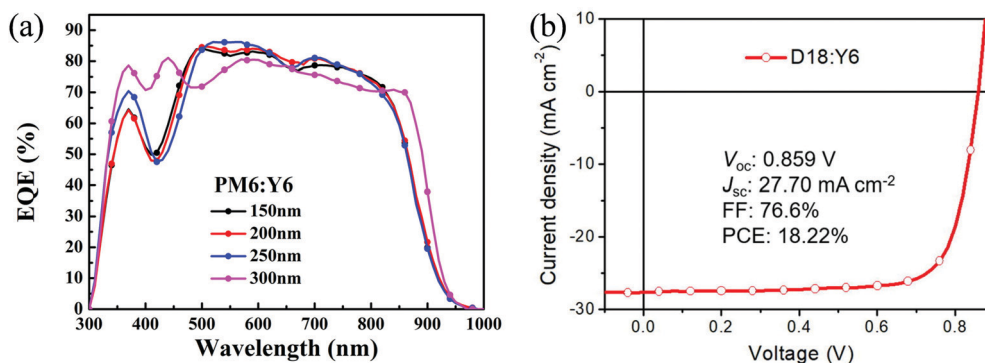

Fig. 9 (a) EQE curves of **PD2**:Y6-based OSCs, reproduced from ref. 9 with permission from Elsevier, copyright 2019; and (b) J–V curve for a **PO12**:Y6-based OSC, reproduced from ref. 10 with permission from Elsevier, copyright 2020.

investigate the influence of selenophene π-spacers as well as chlorination on the performance of OSCs. It was found that **PD13** without π-spacers and chlorination showed the most twisted structure with blue-shifted absorption. On the other hand, **PD15** with π-spacers and chlorination demonstrated the strongest sunlight harvesting capability with a medium band-gap of 1.73 eV, a deep E_{HOMO}, the largest dipole moment, the most planar polymer backbone, the strongest aggregation, and the highest charge mobility. Therefore, a high PCE of 14.44% was obtained when **PD15** was blended with Y6.[138]

As both F atoms and ester groups can downshift the FMO energy levels and enhance the planarity and hence the charge transport, Sun et al.[139] simultaneously introduced these two types of substituents onto the β positions of thiophene to form FE-T as A units and made several terpolymers with **BDT** and BDD units via random copolymerization. The resultant polymer **PD16** (S1) had a low-lying E_{HOMO} of −5.52 eV and enhanced aggregation/crystallinity. Moreover, this ternary strategy enabled a favorable balance between crystallinity and miscibility. When this polymer was blended with Y6, a high PCE of 16.42% was obtained.

3.3.3 BDT-based polymer donors using N-heterocycles as A units.

The molecular conformation is crucial for the molecular packing and crystallinity, and thus the photovoltaic performance of the polymer donor. Since many D–A polymer donors employ π-spacers between the D and A units, the type of π-spacer would influence the molecular conformation. Tang et al.[152] utilized BDT with chlorothiophene side chains and FBTA as D and A units, respectively, and 3-hexylthieno[3,2-b]thiophene as π-spacers to develop a polymer **PN21** (PE4). Compared with the polymer using thiophene as the π-spacers (J52-Cl), the molecular conformation of **PN21** using this new spacer changed from a zigzag conformation to a linear conformation, which led to strong π–π stacking, ordered lamellar

packing, small phase domains, and high mobility. OSC devices based on the **PN21**:Y6 blend showed a high PCE of up to 14.02%.

Phenazine (Pz) can be considered as a Qx derivative with extended conjugation. Ding et al.[154] employed dithieno[3,2-a:2′,3′-c]phenazine (DTPz) and 9,10-difluorodithieno[3,2-a:2′,3′-c]-phenazine (FDTPz) as A units, BDT as D units, and thiophene as π-spacers to form two novel polymer donors, **PN27** and **PN28**, respectively. Both polymers formed complementary absorption with Y6, while the fluorinated polymer **PN28** showed a lower E_{HOMO} of −5.46 eV than that of **PN27** (−5.33 eV). The devices based on the **PN28**:Y6 blend showed a higher V_{OC} (0.867 vs. 0.777 V) and a similar J_{SC} compared with **PN27**:Y6. Moreover, due to the higher and more balanced μ_h and μ_e (Table 5), a higher FF (65.8% vs. 61.6%) was obtained for **PN28**:Y6. Thus, the **PN28**:Y6 blend showed a higher PCE of 15.14% than that of **PN27**:Y6 (12.3%).

3.3.4 Other BDT-based polymers and non-BDT-based polymers.

Ding et al.[161] used a fused-ring aromatic lactone (FRAL) with strong electron-withdrawing capability and an extended molecular plane as a building block to develop D–A polymer donors. The resulting **PO9** (L1) and **PO10** (L2) showed complementary absorption spectra with Y6. The higher and more balanced μ_h and μ_e (Table 5) of **PO9**:Y6 contributed to a higher J_{SC} (23.93 vs. 20.52 mA cm^{-2}) and FF (74.7% vs. 70.2%) and thus a higher PCE (14.36% vs. 12.57%) compared to **PO10**:Y6. They further replaced the lactone unit with the thiolactone unit to form **PO11** (D16), which showed good planarity and thus stronger intermolecular π–π stacking and charge carrier transport, leading to a higher PCE of 16.22%.[160]

Later, Ding et al.[10] used dithieno[3′,2′:3,4;2″,3″:5,6]benzo[1,2-c][1,2,5]thiadiazole (DTBT) as the A units and BDT as the D units to synthesize a polymer donor, **PO12** (D18), which showed a higher μ_e of 1.59×10^{-4} cm^2 V^{-1} s^{-1} than that

Table 5 OSC device performances based on blends of polymer donors and **Y6**

Blend	μ_h, μ_e ($\times 10^{-4}$ cm^2 V^{-1} s^{-1}), $\mu_{e(h)}/\mu_{h(e)}$[a]	EQE[b] (%)	V_{OC} (V)	J_{SC} (mA cm^{-2})	FF (%)	PCE (%)	S_1[c]	S_2[d]	P_{diss} (%)	Ref.
PE14:Y6	1.18, 0.83, 1.42	82	0.81	26.68	74.11	16.02	1.16	0.99		123
PE15:Y6	0.87, 1.32, 1.52	<70	0.66	19.5	46	5.9[e]	1.27	0.98		110
PE16:Y6	5.61, 2.42, 2.32	83	0.83	25.6	66.7	14.2	1.59	0.9[e]		110
PD2:Y6	5.90, 2, 2.95		0.83	25.3	74.8	15.7		0.975		9
PD3:Y6	7.46, 4.66, 1.60		0.897	25.644	74	17.037			96.7	130
PD8:Y6	7.1, 5.1, 1.39		0.84	26.1	72.5	16.1				207
PD13:Y6	0.84, 4.1, 5		0.84	4.25	42.36	1.51		0.9		138
PD14:Y6	3.4, 2.9, 1.17		0.65	22.23	56.4	8.17		0.95		138
PD15:Y6	7.0, 5.7, 1.22		0.82	24.07	73.16	14.44		0.96		138
PD16:Y6	7.01, 4.43, 1.67		0.877	25.402	73.7	16.421		0.978	95.8	139
PN21:Y6	8.04, 1.14, 7.05		0.84	22.21	75.43	14.02	1.27	1		152
PN27:Y6	1.33, 0.681, 1.95		0.777	25.7	61.6	12.3		0.933	87.9	154
PN28:Y6	3.42, 2.42, 1.41		0.867	26.53	65.8	15.14		0.958	94.9	154
PO9:Y6	5.62, 4.77, 1.18	76	0.8	23.93	74.7	14.63		0.954	96.7	161
PO10:Y6	1.60, 2.05, 1.28	68	0.87	20.52	70.2	12.57		0.951	94.7	161
PO11:Y6	2.82, 2.81, 1.01	83	0.85	25.41	74.9	16.22		0.981	97.8	160
PO12:Y6	1.49, 1.40, 1.06	87	0.859	27.70	76.6	18.22				10
PO13:Y6	3.23, 1.51, 2.14	85	0.88	25.87	70.7	16.16		0.959	94.6	162
PX5:Y6	32.6, 13.5, 2.41		0.826	26.65	75.1	16.53		0.986	99.4	208
PX9:Y6	2.82, 1.25, 2.26		0.74	22.7	65.73	11.05				169

[a] μ_e/μ_h or μ_h/μ_e, whichever is >1. [b] Maximum value of EQE. [c] Obtained from $V_{OC} \sim S_1 \log(P_{light})$. [d] Obtained from $J_{ph} \sim S_2 \log(P_{light})$. [e] The numbers of significant digits of the data are adopted from the original references and may be different.

of **PO11** (1.19×10^{-4} cm^2 V^{-1} s^{-1}) since DTBT has a larger molecular plane. While the high and balanced μ_h and μ_e of the **PO12**:Y6 blend were responsible for the high FF (76.6%), the complementary absorption contributed to the very high J_{SC} of 27.7 mA cm^{-2} (Table 5). A PCE of 18.22% was achieved (Fig. 9b), which is the highest value reported for single junction OSCs. They also used fluorine- and alkoxyl-substituted benzene (FAB) as A units to synthesize a new polymer donor **PO13** (W1) with a low E_{HOMO} of −5.36 eV and a wide bandgap of 2.16 eV.[162] The **PO13**:Y6 blend achieved a PCE of 16.16%.

PX5 was synthesized with very low cost and showed high performance when blended with IDIC or IT-4F.[167,168] When **PX5** was blended with Y6, much wider absorption and EQE curves were obtained owing to the red-shifted absorption of Y6.[208] In addition, higher μ_h and μ_e were also obtained. A largely improved J_{SC} of 26.65 mA cm^{-2} was achieved, leading to a high PCE of 16.53%.

When J52 was blended with Y6, a low V_{OC} (0.65 V) resulting from the high E_{HOMO} (−5.12 eV) of J52 and a low FF (53.82%) due to the imbalanced μ_h and μ_e were obtained, leading to a moderate PCE of 7.15%.[169] When IDTT was used to replace BDT in J52, the resulting **PX9** (PIDTT-DTffBTA) had a lower E_{HOMO} of −5.25 eV.[169] When blended with Y6, more balanced μ_h and μ_e were obtained for the blend film. Consequently, a higher V_{OC} of 0.74 V, an improved FF of 65.73%, and an enhanced PCE of 11.05% were obtained.

3.4 General guidelines for the design of high-performance WBG polymer donors

Based on the above discussions, some useful guidelines for the design of high-performance WBG polymer donors for matching narrow bandgap NFAs are obtained, which are briefly summarized as follows:[42,209–212]

(1) A combination of suitable D and A building blocks to form D–A polymers has been almost exclusively used as the strategy to develop WBG polymer donors for matching narrow bandgap NFAs since desired FMOs and bandgaps can be readily obtained by fine selection of donor and acceptor units with different abilities of electron withdrawing or donating properties.[35] Thiophene-containing building blocks have been used in the D units due to their appropriate electron-donating ability to obtain the desired E_{HOMO} and their lower steric effects on the neighbouring units to maintain backbone planarity to achieve high hole mobility. Among them, BDT has been most frequently used due to its large and planar fused-ring extended π-conjugated structure that can facilitate hole transport and more importantly induce face-on crystal orientation for efficient charge transport along the vertical direction in the active layer. Moderately electron-withdrawing building blocks that can maintain a relatively high E_{LUMO} have been utilized as the A units to construct WBG polymer donors. The electron-withdrawing groups among these building blocks are mainly esters, imides, ketones, N-heterocyles (containing sp^2 nitrogen), and their combinations. In addition, an electron donating, sterically less demanding π-spacer has often been inserted between the D and A units in a D–A type polymer donor to maintain the backbone planarity and the distance between the

D and A units to increase the electronic polarization, which possibly helps stabilize the excitons.[213–216]

(2) Several strategies can be used to lower E_{HOMO} of the polymer donors: (i) introducing electron-withdrawing groups such as fluorine, chlorine, ester, trifluoromethyl, and alkoxy groups;[121,147,148,217,218] (ii) replacing thienyl side chains with alkoxy or alkylthio substituted phenyl side chains;[133,219] and (iii) replacing the alkyl side chains with alkylthio groups.[220]

(3) A coplanar and rigid backbone is needed for high charge carrier mobility, which can be achieved by: (i) extending the conjugation area of the backbone by using larger fused ring building blocks;[196] (ii) replacing thiophene with selenophene, which has lower aromaticity and enhanced ground-state quinoid resonance character, resulting in improved planarity and increased effective conjugation length;[138] (iii) fluorination or chlorination, which enhances the intermolecular interactions;[217,218] and (iv) strengthening intermolecular interactions through side chain engineering[221] and introduction of π-spacers.[138]

3.5 Morphology control in WBG polymer donor:NFA systems

The morphology of the BHJ blend active layer directly impacts the exciton dissociation, charge separation, and transport processes. A bicontinuous network of the donor and acceptor with domain sizes between 10 and 20 nm could provide a large interfacial area between the two domains to facilitate exciton dissociation, while forming a continuous and relatively pure percolating pathway to transport free charges to the corresponding electrodes.[222–228] The extent of this phase separation is critical since domains that are too large (greater than the typical exciton diffusion length on the order of tens of nanometres) will favour geminate recombination, whereas too-mixed domains can increase the rate of non-geminate recombination.[4,229–231] A favourable morphology can be achieved by tuning the donor:acceptor miscibility and crystallinity.[229]

Herein, we discuss how changing the donor properties, choosing a certain acceptor in the BHJ blend, and modifying the processing conditions would impact the morphology of the donor:acceptor blend and hence the performance of the OSC device.

3.5.1 Donor and acceptor interaction and miscibility. In most high-performance WBG donor:NFA systems, the as-cast active layer is in the metastable state and the film formation is strongly dependent on the thermodynamics and kinetics of thin film formation from the blend solution.[232–234] The Flory–Huggins interaction parameter (χ) has been used to predict the morphology evolution in the active layer in a number of studies. Three distinct miscibility regions have been defined by Ye *et al.*[228] as shown in Fig. 10 in the χ–ϕ phase diagram, where ϕ is the polymer volume fraction; an ideal morphology is neither too pure nor too mixed. Parameter χ can be calculated by differential scanning calorimetry (DSC) measurements (melting point depression method) or surface energy measurements of the donor and acceptor (details of the equations used for the calculation can be found elsewhere).[222,228,232,233,235,236]

A computational approach to studying the charge separation efficiency between polymer donors and NFAs was proposed by
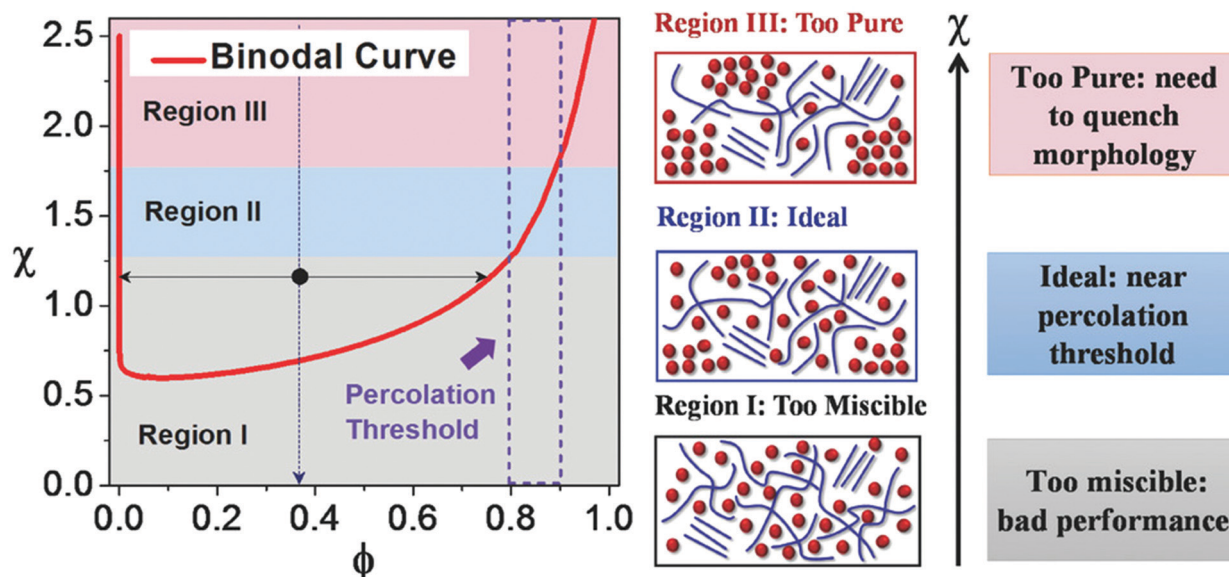
Fig. 10 Schematic illustration of three regimes in the Flory–Huggins interaction parameter–polymer volume fraction (χ–ϕ) phase diagram. Regime I, II, and III represent high χ, medium χ (where the purity of the mixed domain is close to the small molecule acceptor percolation threshold), and low χ, respectively. The dashed rectangle marks the reported percolation threshold values where the polymer volume fractions are 0.8–0.9. The composition corresponding to the lower limit for electron transport is termed as the percolation threshold. Reproduced from ref. 228 with permission from John Wiley and Sons, copyright 2018.

Yao et al.[121] They found that the larger electrostatic potential difference between a WBG polymer donor (**PE12**) and NFA (IT-4F) was responsible for a high PCE of 14.7%, even though the driving force for charge separation (energy level offsets) in the blend was quite small.[121]

3.5.2 Active layer morphology optimization methods. During film casting, the intermolecular forces among polymer chains play a critical role in determining the properties of the resulting condensed state. These properties include the aggregation, morphology, domain purity, molecular orientation, etc. of the polymer in the blend film. For example, rigid polymers (usually polymers with rigid BDT and BDD units) with strong π–π stacking lead to ordered BHJ films with increased domain purity. On the other hand, polymers with relatively weak π–π stacking and less pronounced molecular ordering in solution (such as **PX1**, **PE17** (PTB7), **PE1** and **PE18** (PBDT-TS1)) tend to form much smaller and random polymer domains upon casting.[163,236–238]

The morphology of the donor:acceptor blend can be controlled by pre- and post-deposition treatments of the as-cast film. These morphology optimization methods aim to achieve high average domain purity and highly ordered packing of both the donor and acceptor to minimize charge recombination in NFA-based OSCs.[223]

Processing solvent additives are well known for regulating the BHJ active layer morphology. Wang et al.[239] studied the effect of different concentrations of a common solvent additive DIO (1,8-diiodooctane) in the **PD1**:ITIC system, using a vertical stratification study of the donor and acceptor. They found an optimal DIO content of 0.5%. It is noteworthy that while an increase in the DIO content (to 1%) leads to abruptly enhanced face-on packing of **PD1** (beneficial in OSCs), excessive vertical component distribution and

over crystallization of ITIC occur, which results in reduced vertical carrier transport paths for both **PD1** and ITIC and a reduced PCE. The device with no additive demonstrated an average PCE of 8.74%, J_{SC} of 15.88 mA cm^{-2}, V_{OC} of 0.892 V, and FF of 61.7%. With 0.5% DIO an increased J_{SC} and FF of 17.23 mA cm^{-2} and 69.3%, respectively, resulted in a high PCE of 10.75%. With 1% DIO, both J_{SC} and the FF dropped to 15.23 mA cm^{-2} and 66.8%, respectively, thereby leading to a lower PCE of 9.27%.

Solvent vapor annealing (SVA) is an efficient post-processing technique to control the morphology of the active layer. It serves as a method for molecular level rearrangement by penetrating solvent vapor into the active layer.[240,241] Jiao et al.[240] used four common annealing solvents, methylene chloride (CH$_2$Cl$_2$), chloroform (CHCl$_3$), tetrahydrofuran (THF), and carbon disulfide (CS$_2$), to study the effect of active layer SVA treatment on the performance of **PD3**:IT-4Cl-based OPVs, achieving significantly improved performance compared with that of the as-cast OPVs. SVA (especially with CS$_2$) enhanced the charge generation, transport, and collection efficiency. The **PD3**:IT-4Cl device with CS$_2$ solvent treatment showed the highest PCE of 13.76% with a J_{SC} of 20.53 mA cm^{-2}, a V_{OC} of 0.87 V and a FF of 77.05%, while the control device showed a lower J_{SC} (19.28 mA cm^{-2}) and FF (70.36%), and a similar V_{OC} of 0.88 V with a PCE of 11.94%.

Thermal annealing is another widely used simple method for morphology optimization, mostly due to the temperature-dependent aggregation behaviour of donor–acceptor phases.[223,233,242–244] For example, Bin et al.[141] reported a more than 25% improvement in the PCE of the **PN2**:ITIC device upon annealing at 150 °C for 10 min due to enhanced crystalline domains with prominent face-on orientation leading to higher

SCLC mobility (improved charge transport) as compared to unannealed devices. The as-prepared **PN2**:ITIC device showed a PCE of 9.03% with a J_{SC} of 14.81 mA cm^{-2}, a V_{OC} of 0.96 V and a FF of 63.63%, while, upon annealing, the device showed a significantly higher J_{SC} (17.32 mA cm^{-2}) and FF (69.77%) and a similar V_{OC} (0.94 V), resulting in a PCE of 11.41%. Especially for the high performance WBG polymer **PD1**, with temperature dependent aggregation behavior, thermal annealing is necessary to achieve the optimal photovoltaic performance.[141] The critical parameters that determine the effectiveness of thermal annealing for a polymer:NFA system are the glass transition temperature (T_g) and miscibility of the components (χ).[233]

4 Desirable properties of WBG polymer donors for high efficiency NFA-based OSCs

4.1 Molecular packing and charge transport

The choice of polymer donor will affect not only μ_h of the blend but also the molecular ordering of the NFAs significantly and thus μ_e and the domain purity (miscibility).[245] A coplanar polymer with reduced steric hindrance in the backbone is favorable for ordered molecular packing and intermolecular charge transport and is correlated with higher μ_h of neat and blend films. For example, Li *et al.*[119] employed covalent and noncovalent (F···S conformational lock) backbone rigidification strategies, respectively, to design two BDT and ester-modified oligothiophene-based copolymers (**PE7** and **PE8**). Hydrogen-bonding interactions induced in the DFDT-based unit in polymer **PE8** resulted in stronger interchain π–π interactions, a smaller π–π distance with a longer coherence length (L_c) and thus a significantly higher μ_h compared to its TT counterpart. A V_{OC} of 0.86 V, a high J_{SC} of 21.83 mA cm^{-2} and a remarkable FF of 75%, and thus a high PCE of 14.16% were obtained for the **PE8**:IT-4F device, whereas the **PE7**:IT-4F device showed a significantly lower J_{SC} (19.34 mA cm^{-2}) and FF (65%), and thus a low PCE of 11.10%. Liu *et al.*[127] designed two new WBG polymers (**PE19** (PB2T) and **PE20** (PB3T)) based on **BDT** and oligothiophene units. **PE20** exhibits a more planar geometry, whereas **PE19** has a highly twisted backbone conformation; therefore, a PCE of 11.9% was achieved for the **PE20**:IT-M (Fig. 11) device compared to the very poor performance (0.01%) of the **PE19**:IT-M device. Tang *et al.*[170] used a backbone fluorination strategy to improve the crystallinity

(lamellar and π–π packing coherence length) and thus SCLC mobilities, making use of the introduced F···F, C–F···H, and F···S interactions in **PX11** (PDTF-TZNT) compared to non-fluorinated counterpart **PX11**. Consequently, the **PX10**:IT-M device showed a relatively low PCE of 4.42% with a J_{SC} of 10.15 mA cm^{-2}, a V_{OC} of 0.75 V, and a FF of 58.1%, whereas the **PX11**:IT-M device obtained a higher PCE of 10.05% with an improved J_{SC} of 17.33 mA cm^{-2}, a V_{OC} of 0.80 V, and a significantly high FF of 72.5%.

Achieving high FF requires efficient charge extraction capability facilitated by high mobilities of both holes and electrons. As seen in Fig. 12a, for high performance systems, blend carrier mobilities in excess of 10^{-4} cm^2 V^{-1} s^{-1} can achieve a FF of >70%. To achieve such moderate to high mobilities in the device, the polymer donor is preferred to have a planar backbone with moderate aggregation capability, predominantly face-on orientation, and covalent and noncovalent backbone rigidification induced by strategies such as fluorination, conformation locking, and/or side chain modification. In addition, the mobility ratio, crystal sizes, and coherence length of the π–π stacking order of the polymer in the blend are also crucial (OSC performance dependencies on the mobility ratio and (010) coherence length are plotted in Fig. 12b and c, respectively). For perspective, most high-performance systems utilize polymers with neat film SCLC mobilities in excess of 10^{-4} cm^2 V^{-1} s^{-1} for achieving a high FF of >60% (Fig. 12d). Importantly, an overly crystalline polymer donor can also lead to decreased miscibility with the NFA in the blend, leading to large pure aggregated domains, which is detrimental to charge transport.[236]

4.2 Matching energy levels for minimizing E_{loss}

For a long time, it has been considered that a sufficient energy level offset of ≥ 0.3 eV between organic donor and acceptor materials, ΔE_{HOMO} or ΔE_{LUMO}, is required for efficient charge separation mostly based on observations for fullerene-based OSCs.[87,121,192–195] This is due to the nature of organic semiconductors, which typically have much larger exciton binding energies (~ 0.3–1 eV) than inorganic semiconductors. However, due to the advent of efficient NFAs, several studies have demonstrated highly efficient OSCs with efficient and fast charge separation at a ΔE_{HOMO} (and/or ΔE_{LUMO}) much smaller than 0.3 eV. For example, Sun *et al.*[171] used finely tuned WBG polymer donor:NFA acceptor system **PX12** (PTQ11):TPT10 (Fig. 11) to achieve highly efficient exciton dissociation and charge transfer, resulting in a high PCE of 16.32% with a high
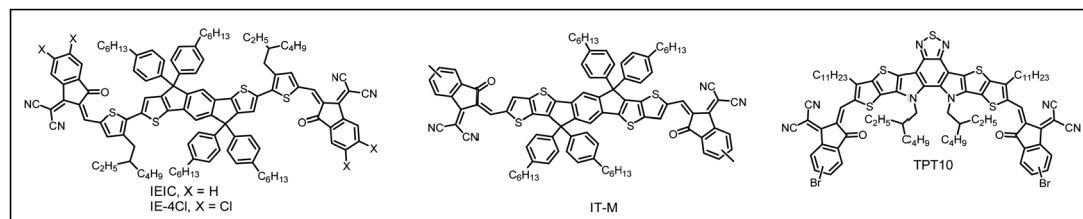

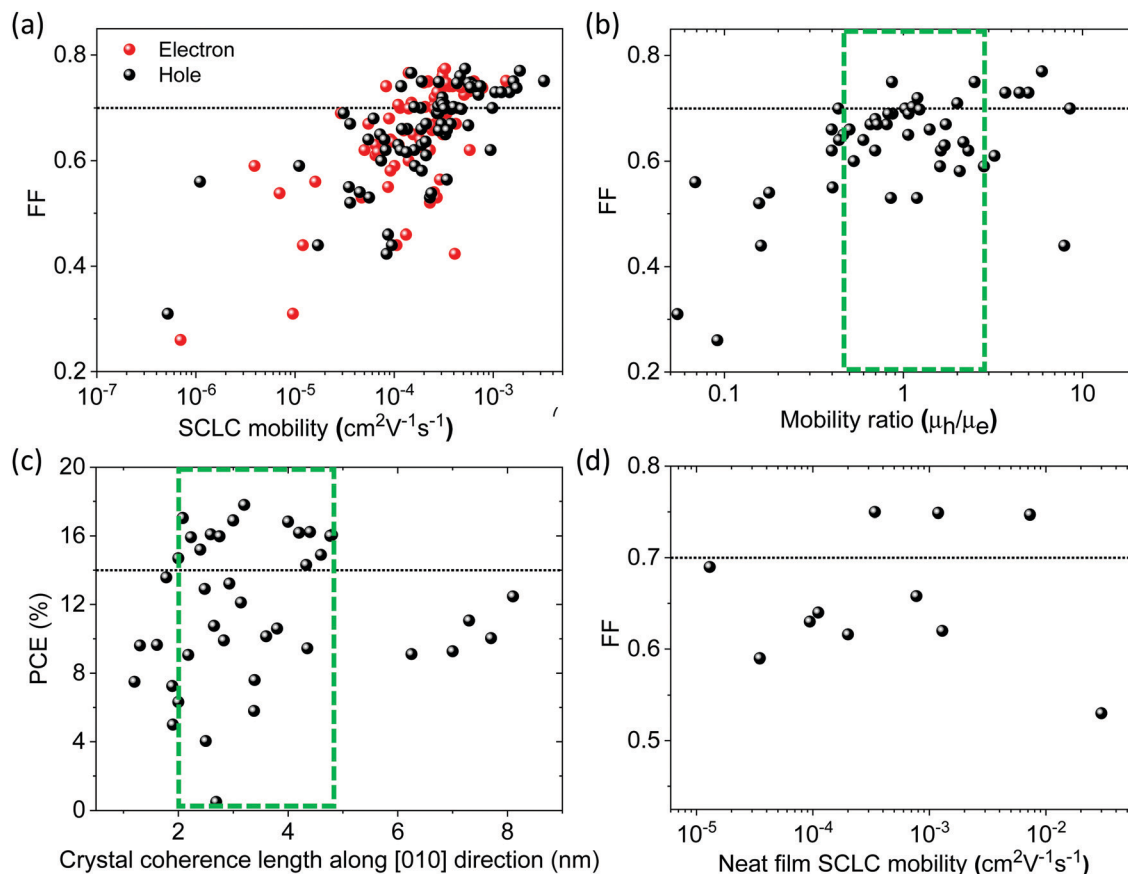**Fig. 11** Chemical structures of additional NFAs.

Fig. 12 Plots of (a) SCLC electron and hole mobilities measured in blend films *vs.* FF, (b) SCLC mobility ratio (μ_h/μ_e) *vs.* the corresponding FF, (c) coherence length along the [010] direction of crystal packing in blend films *vs.* the respective PCE, and (d) neat film SCLC mobility *vs.* the corresponding FF in OSCs with WBG polymer donors and NFAs discussed in this review.

V_{OC} of 0.88 V, a large J_{SC} of 24.79 mA cm^{-2}, and a high FF of 74.8%, despite the zero ΔE_{HOMO} value. Similar instances of a high J_{SC} (high EQE) despite an incredibly low or zero ΔE_{HOMO} or ΔE_{LUMO} have been reported elsewhere.[65,246–250]

The voltage loss (V_{loss}), calculated from the difference between the optical bandgap of the solar cell active layer and V_{OC} of the device ($V_{loss} = E_g/q - V_{OC}$), is the key concern for obtaining the maximum performance from a given pair of a WBG polymer donor and NFA.[248,250] This voltage loss (also referred to as energy loss, $E_{loss} = eV_{loss}$) can be minimized by increasing the difference between E_{LUMO} of the acceptor and E_{HOMO} of the donor as V_{OC} of OSCs is proportional to this difference. Therefore, if fast and efficient charge separation can be achieved in blends with almost no energy offsets, the LUMO and HOMO of the donor polymer could align with the acceptor to minimize ΔE_{HOMO} and ΔE_{LUMO}, effectively decreasing V_{loss}.[171] This was one of the two design criteria proposed by Qian *et al.*,[248] where they used spectroscopic and quantum-chemistry approaches to identify key rules for minimizing voltage losses: (1) a low energy offset between the donor and acceptor molecular states and (2) high luminescence of the low-bandgap component.

Nevertheless, several studies have reported the need to fine tune the molecular energy alignment to provide a "certain

minimum" driving force for balancing the exciton dissociation probability and V_{OC} loss. Yang *et al.*[65] used structurally similar donor and acceptor combinations (with similar blend morphologies and charge transport properties) to study the performance variations corresponding to their different energy-level offsets. They found that the energy offset has no impact on the charge transport process, but it does affect the exciton dissociation process in NFA-based OSCs. A larger ΔE_{HOMO} between the donor and acceptor can provide sufficient driving force, which contributes to a higher exciton dissociation probability (P_{diss}) and PCE. Although minimizing ΔE_{HOMO} can effectively reduce E_{loss} (or improve V_{OC}), the decreasing driving force will diminish the exciton dissociation process. Moreover, as ΔE_{HOMO} falls below a threshold, P_{diss} drops sharply, reducing the overall performance of the device.[65] In yet another study investigating the relationship between E_{loss} and the device performance of NFA-based OSCs, Ling *et al.*[249] found that donor and acceptor blends with an E_{loss} below 0.6 eV demonstrated inefficient charge transfer between the polymer donor and NFA, resulting in a low EQE and PCE (an example of the trade-off between device performance and energy loss). The **PD1**:IE-4Cl (Fig. 11)-based device afforded a PCE of 11.1% (with a high J_{SC} of 21.49 mA cm^{-2} and a moderate FF of 60%) and E_{loss} of 0.64 eV, whereas the **PD1**:IEIC based device obtained a low

PCE of 3.8% (with a low J_{SC} of 7.65 mA cm^{-2} and low FF of 45%) despite the lower E_{loss} of 0.52 eV.

The reasons as to why some polymer donor:acceptor blend systems require a very low ΔE_{HOMO} or ΔE_{LUMO} for efficient charge separation and show a very low E_{loss} are still not well understood, but seem to correlate with both the donor and acceptor materials. More examples are needed to probe the relationships of the structures (chemical and morphological) and optoelectronic properties of the polymer donors with E_{loss} of the OSC devices, which are extremely important for maximizing the potential of OSCs.

4.3 Characteristics of top performing polymer donors and OSCs

The OSC performance parameters, V_{OC}, J_{SC}, FF, and PCE, of the WBG polymer donors paired with ITIC, IT-4F, and Y6 discussed in this review are plotted in Fig. 13. It can be seen that V_{OC} decreases and J_{SC} increases from ITIC to IT-4F to Y6, while the FF is similar among the three NFAs even though the IT-4F based devices show a slightly higher FF. From ITIC to IT-4F to Y6, the FMO energy levels downshift and the absorption spectra red-shift, which leads to a narrowed optical bandgap, resulting in reduced V_{OC}, but increased J_{SC}. The overall PCE of

the OSCs follows the order of ITIC < IT-4F < Y6, which is due to the relatively large contribution of J_{SC} that supersedes the reduction in V_{OC}. The PCE values of the top five high-performing polymer donors and the photovoltaic parameters of the best polymer donor for each NFA are shown in Fig. 14.

Some distinct characteristics of these high-performance polymer donors (in Fig. 14) can be observed. Firstly, the E_{HOMO} values of the donors matching with ITIC are -5.3 to -5.5 eV, while those of the donors matching with IT-4F and Y6 are -5.5 to -5.6 eV, which are decided by E_{HOMO} of the NFAs since a sufficient ΔE_{HOMO} for efficient dissociation of excitons generated in the NFA phase is required. Secondly, these WBG polymer donors have E_g values of 1.9–2.1 eV that cover the short wavelength region (where the NFAs absorb poorly) to form complementary absorption to achieve high J_{SC}. Lastly, these polymers have SCLC hole mobilities higher than 10^{-4} cm^2 V^{-1} s^{-1}, which could achieve high and balanced μ_h and μ_e in their blends with NFAs, leading to high FF. All these blends have a μ_h/μ_e ratio very close to 1 except for **PX5:Y6**, which has a μ_h/μ_e ratio (2.41) greater than 2, where both μ_h and μ_e are higher than 10^{-3} cm^2 V^{-1} s^{-1}, which would probably not impede charge transport even though μ_h and μ_e are not well balanced in this case.
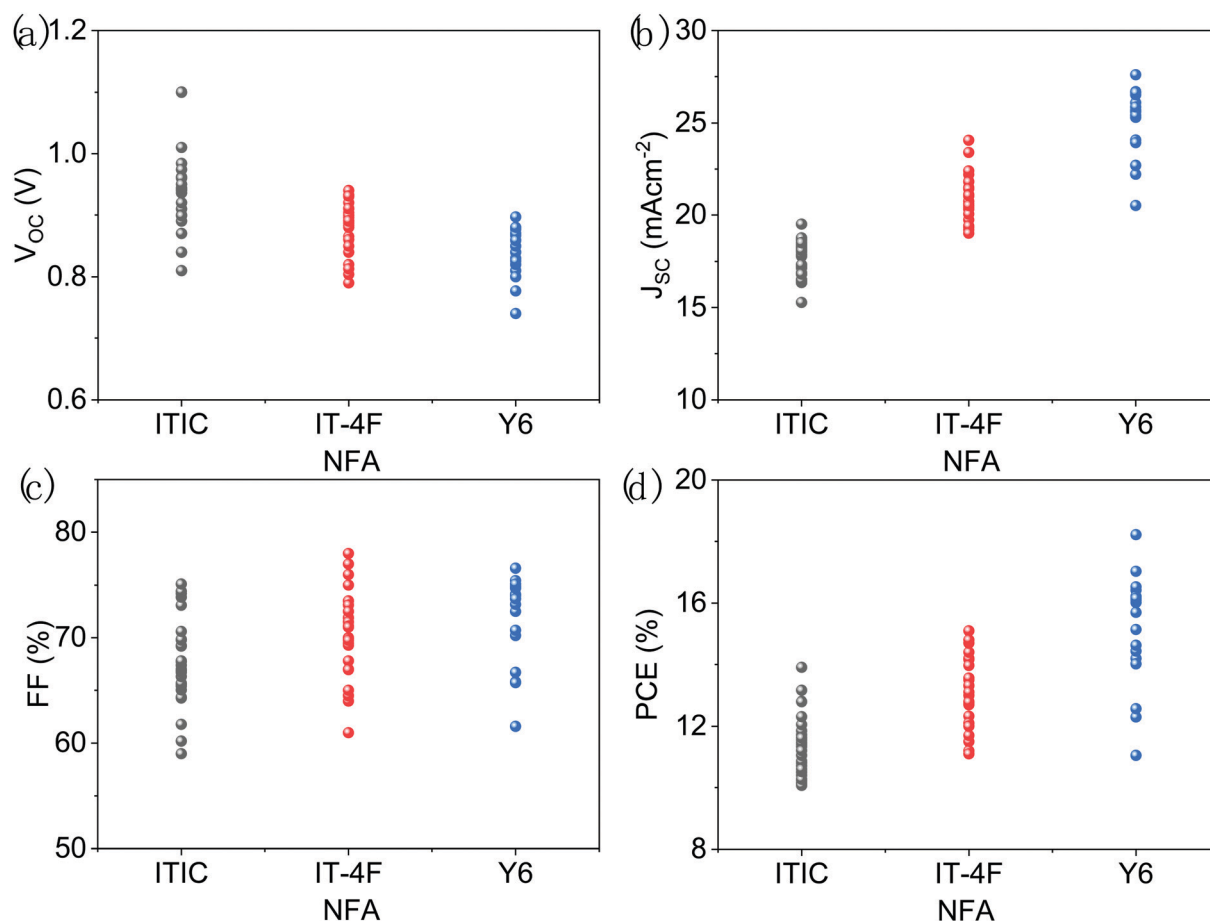

Fig. 13 V_{OC} (a), J_{SC} (b), FF (c), and PCE (d) of high-performance OSCs using WBG polymer donors and different NFAs, ITIC, IT-4F and Y6, discussed in this review.
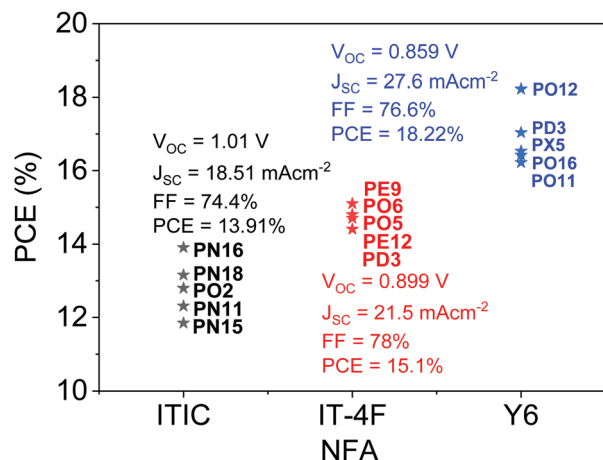
Fig. 14 Device performances of the top performing polymer donors matching with ITIC, IT-4F and Y6, where the photovoltaic parameters shown are from the best performing **PN16**:ITIC, **PE9**:IT-4F and **PO12**:Y6 systems.

5 Conclusions and outlook

In the past few years, tremendous progress has been made in the development of WBG polymer donors that match the emerging high-performance narrow bandgap NFAs, achieving remarkable PCEs of up to 18.22%, which has beaten the predicted PCE limit of ~10% for OSCs based on fullerene acceptors.[251] In this review, we have examined the most representative WBG polymer donors that have been used to match three typical NFAs: ITIC, IT-4F, and Y6. The structure–property–cell performance relationships are probed by analysing the structural design strategies to control and finely tune the FMO energy levels, adjust the light absorption profile, and enhance the hole mobility to match those of different NFAs. The morphology control of the polymer blends to create desirable phase separation and crystallinity for achieving high device performance is also discussed.

Despite these recent advances in the development of polymer donors and OSCs in general, there is room for further improvement of the cell performance, while some challenges and issues need to be overcome in order to make OSCs a commercial reality.

(1)   Further boosting the PCE of OSCs

Theoretical limit of the PCE for OSCs. According to Shockley and Queisser,[43] the maximum PCE of a single-junction solar cell is 33.7% (the S–Q limit) achievable with a semiconductor having a bandgap of 1.34 eV. Hou et al.[17] predicted that a PCE of >20% for OSCs is possible if V_{loss} can be lowered to <0.45 V, a realistic value considering that a V_{loss} of 0.3 V has been reported recently for perovskite solar cells,[252] while the current state-of-the art OSCs show a V_{loss} of ca. 0.6 V.[253,254] Therefore, theoretically there is still much room for further enhancement of the PCE for OSCs.

Reducing the energy loss. Organic materials with large dielectric constants in the same range as silicon (~11) could reduce the coulombic interaction in excitons, or the exciton binding energy, to enable highly efficient exciton dissociation.[87,90,95,255] Consequently, thicker layers could be used to maximize the absorption of light (and to ease large-scale fabrication). Moreover, the largely reduced exciton binding energy would only require very small ΔE_{HOMO} and ΔE_{LUMO} for exciton dissociation, which could greatly reduce the energy loss. Furthermore, the reduced exciton binding energy would enlarge the exciton diffusion distance, which may enable bilayer heterojunction OSCs without delicate and tedious morphological control of the blend layer to realize a nanoscale BHJ structure. This would be very desirable for single component-organic solar cell devices (SC-OSCs), where the polymer in the active layer having both donor and acceptor components could greatly simplify the device fabrication process and achieve high morphological stability.[98] Very recently, Jiang et al.[256] achieved a record PCE of 8.40% in SC-OSCs using a double-cable conjugated polymer.

(2)   Long-term stability

Despite the considerable efforts made over the past few decades, OSCs have not become a commercial reality yet. The stability of OSCs is one of the main obstacles in the path towards commercialization. Solar cells work under various operational stresses such as solar illumination, increased temperature, and exposure to air.[102,257–262] The intrinsic degradation mechanisms of the active layer in OSCs include degradation in the dark caused by the morphological changes (rearrangement of donor and acceptor domains) that occur over time and photo-chemical reactions of the organic semiconductor materials in the presence of illumination.[263–267] In terms of morphological stability, all-polymer solar cells have attracted much attention since a polymer donor and a polymer acceptor can be entangled to form a much more metastable blend, leading to superior thermal and mechanical stabilities compared to small molecule-based OSCs.[268–274] Notable progress has been made in improving the lifetime of OSCs with extrapolated lifetimes (with 80% PCE retention) of over 20 years being demonstrated recently.[263,275]

(3)   Fabrication of large area devices

Another big obstacle towards the commercialization of OSCs is the poor scalability of OSC devices. Most, if not all, of the OSCs showing high PCEs greater than 10% are reported with an active area of less than 1 cm². Increasing the device active area usually results in a significant reduction in the PCE. Although promising large area modules have been reported,[276,277] the lab-to-manufacturing translation seems still very challenging and requires urgent attention.[278]

(4)   Low cost

Commercialization of OSCs will require a competitive low manufacturing cost-to-PCE ratio along with energy payback and embodied energy metrics comparable to existing silicon solar cells.[279,280] It has been predicted that for OSCs having unique merits such as flexibility, lightness, and high throughput manufacturability, a module efficiency of 10% with an

operational lifetime of 10 years and a production cost of < \$0.5 per W_p would be enough to reach the threshold of commecialization.[281] For the donor:acceptor combination to be fully scalable, it should possess a high value of the industrial figure of merit (i-FOM).[282–284] This parameter takes account of the synthetic complexity (SC) of the donor and acceptor materials as well as the PCE and lifetime of the OSC device. Most of the high PCE polymer donors require tedious synthesis and use of toxic substances, resulting in very high SC and thus insufficient i-FOM. The importance of developing polymer donors with low SC has been increasingly realized. On the other hand, small molecule donor materials have found increased interest in the last couple of years for their industrial scalability owing to the advantages of well-defined chemical structures, synthetic ease, high purity of materials, and outstanding repeatability from batch to batch synthesis.[285,286]

(5) Environmentally friendly manufacturing

Most of the high performance OSCs are fabricated using halogenated solvents such as o-dichlorobenzene, chlorobenzene, or chloroform, which are environmental hazards. These solvents might be suitable for proof-of-concept lab-based research, but for large scale production alternative eco-friendly solvents (green solvents) need to be used.[287,288] As discussed, some polymer donors that can be processed using green solvents and yet show high PCE have been reported.[120,289,290]

Conflicts of interest

There are no conflicts to declare.

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Notes and references

1 G. Li, R. Zhu and Y. Yang, Nat. Photonics, 2012, 6, 153–161.
2 O. Inganäs, Adv. Mater., 2018, 30, 1800388.
3 N. Matsuhisa, X. Chen, Z. Bao and T. Someya, Chem. Soc. Rev., 2019, 48, 2946–2966.
4 C. Cui and Y. Li, Energy Environ. Sci., 2019, 12, 3225–3246.
5 G. Yu, J. Gao, J. C. Hummelen, F. Wudl and A. J. Heeger, Science, 1995, 270, 1789–1791.
6 Y. Tong, Z. Xiao, X. Du, C. Zuo, Y. Li, M. Lv, Y. Yuan, C. Yi, F. Hao, Y. Hua, T. Lei, Q. Lin, K. Sun, D. Zhao, C. Duan, X. Shao, W. Li, H.-L. Yip, Z. Xiao, B. Zhang, Q. Bian, Y. Cheng, S. Liu, M. Cheng, Z. Jin, S. Yang and L. Ding, Sci. China: Chem., 2020, 63, 758–765.
7 Y. Liang, Z. Xu, J. Xia, S.-T. Tsai, Y. Wu, G. Li, C. Ray and L. Yu, Adv. Mater., 2010, 22, E135–E138.
8 Y. Lin, J. Wang, Z.-G. Zhang, H. Bai, Y. Li, D. Zhu and X. Zhan, Adv. Mater., 2015, 27, 1170–1174.
9 J. Yuan, Y. Zhang, L. Zhou, G. Zhang, H.-L. Yip, T.-K. Lau, X. Lu, C. Zhu, H. Peng, P. A. Johnson, M. Leclerc, Y. Cao, J. Ulanski, Y. Li and Y. Zou, Joule, 2019, 3, 1140–1151.
10 Q. Liu, Y. Jiang, K. Jin, J. Qin, J. Xu, W. Li, J. Xiong, J. Liu, Z. Xiao, K. Sun, S. Yang, X. Zhang and L. Ding, Sci. Bull., 2020, 65, 272–275.
11 J. C. Hummelen, B. W. Knight, F. LePeq, F. Wudl, J. Yao and C. L. Wilkins, J. Org. Chem., 1995, 60, 532–538.
12 J. Peet, J. Y. Kim, N. E. Coates, W. L. Ma, D. Moses, A. J. Heeger and G. C. Bazan, Nat. Mater., 2007, 6, 497–500.
13 Y. He, H.-Y. Chen, J. Hou and Y. Li, J. Am. Chem. Soc., 2010, 132, 1377–1382.
14 Z. He, C. Zhong, S. Su, M. Xu, H. Wu and Y. Cao, Nat. Photonics, 2012, 6, 591–595.
15 J. Zhao, Y. Li, G. Yang, K. Jiang, H. Lin, H. Ade, W. Ma and H. Yan, Nat. Energy, 2016, 1, 15027.
16 W. Zhao, D. Qian, S. Zhang, S. Li, O. Inganäs, F. Gao and J. Hou, Adv. Mater., 2016, 28, 4734–4739.
17 J. Hou, O. Inganäs, R. H. Friend and F. Gao, Nat. Mater., 2018, 17, 119–128.
18 L. Gao, Z.-G. Zhang, H. Bin, L. Xue, Y. Yang, C. Wang, F. Liu, T. P. Russell and Y. Li, Adv. Mater., 2016, 28, 8288–8295.
19 H. Bin, Z.-G. Zhang, L. Gao, S. Chen, L. Zhong, L. Xue, C. Yang and Y. Li, J. Am. Chem. Soc., 2016, 138, 4657–4664.
20 X. Li, G. Huang, N. Zheng, Y. Li, X. Kang, S. Qiao, H. Jiang, W. Chen and R. Yang, Sol. RRL, 2019, 3, 1900005.
21 S. Li, L. Ye, W. Zhao, S. Zhang, S. Mukherjee, H. Ade and J. Hou, Adv. Mater., 2016, 28, 9423–9429.
22 W. Zhao, S. Li, H. Yao, S. Zhang, Y. Zhang, B. Yang and J. Hou, J. Am. Chem. Soc., 2017, 139, 7148–7151.
23 Y. Yang, Z.-G. Zhang, H. Bin, S. Chen, L. Gao, L. Xue, C. Yang and Y. Li, J. Am. Chem. Soc., 2016, 138, 15011–15018.
24 J. Wang, W. Wang, X. Wang, Y. Wu, Q. Zhang, C. Yan, W. Ma, W. You and X. Zhan, Adv. Mater., 2017, 29, 1702125.
25 Y. Lin, F. Zhao, Q. He, L. Huo, Y. Wu, T. C. Parker, W. Ma, Y. Sun, C. Wang, D. Zhu, A. J. Heeger, S. R. Marder and X. Zhan, J. Am. Chem. Soc., 2016, 138, 4955–4961.
26 F. Zhao, S. Dai, Y. Wu, Q. Zhang, J. Wang, L. Jiang, Q. Ling, Z. Wei, W. Ma, W. You, C. Wang and X. Zhan, Adv. Mater., 2017, 29, 1700144.
27 Y. Lin, Q. He, F. Zhao, L. Huo, J. Mai, X. Lu, C.-J. Su, T. Li, J. Wang, J. Zhu, Y. Sun, C. Wang and X. Zhan, J. Am. Chem. Soc., 2016, 138, 2973–2976.
28 B. Jia, J. Wang, Y. Wu, M. Zhang, Y. Jiang, Z. Tang, T. P. Russell and X. Zhan, J. Am. Chem. Soc., 2019, 141, 19023–19031.
29 S. Dai, F. Zhao, Q. Zhang, T.-K. Lau, T. Li, K. Liu, Q. Ling, C. Wang, X. Lu, W. You and X. Zhan, J. Am. Chem. Soc., 2017, 139, 1336–1343.
30 J. Wang, J. Zhang, Y. Xiao, T. Xiao, R. Zhu, C. Yan, Y. Fu, G. Lu, X. Lu, S. R. Marder and X. Zhan, J. Am. Chem. Soc., 2018, 140, 9140–9147.

31 T. Li, S. Dai, Z. Ke, L. Yang, J. Wang, C. Yan, W. Ma and X. Zhan, *Adv. Mater.*, 2018, **30**, 1705969.

32 Y. Ma, Z. Kang and Q. Zheng, *J. Mater. Chem. A*, 2017, **5**, 1860–1872.

33 H. Hu, P. C. Y. Chow, G. Zhang, T. Ma, J. Liu, G. Yang and H. Yan, *Acc. Chem. Res.*, 2017, **50**, 2519–2528.

34 D. Gedefaw, M. Prosa, M. Bolognesi, M. Seri and M. R. Andersson, *Adv. Energy Mater.*, 2017, **7**, 1700575.

35 L. Sun, X. Xu, S. Song, Y. Zhang, C. Miao, X. Liu, G. Xing and S. Zhang, *Macromol. Rapid Commun.*, 2019, **40** 1900074.

36 A. Wadsworth, M. Moser, A. Marks, M. S. Little, N. Gasparini, C. J. Brabec, D. Baran and I. McCulloch, *Chem. Soc. Rev.*, 2019, **48**, 1596–1625.

37 S. Dey, *Small*, 2019, **15**, 1900134.

38 G. Zhang, J. Zhao, P. C. Y. Chow, K. Jiang, J. Zhang, Z. Zhu, J. Zhang, F. Huang and H. Yan, *Chem. Rev.*, 2018, **118**, 3447–3507.

39 S. Suman and S. P. Singh, *J. Mater. Chem. A*, 2019, **7**, 22701–22729.

40 H. Wang, J. Cao, J. Yu, Z. Zhang, R. Geng, L. Yang and W. Tang, *J. Mater. Chem. A*, 2019, **7**, 4313–4333.

41 C. Yan, S. Barlow, Z. Wang, H. Yan, A. K.-Y. Jen, S. R. Marder and X. Zhan, *Nat. Rev. Mater.*, 2018, **3**, 18003.

42 G. Wang, F. S. Melkonyan, A. Facchetti and T. J. Marks, *Angew. Chem., Int. Ed.*, 2019, **58**, 4129–4142.

43 W. Shockley and H. J. Queisser, *J. Appl. Phys.*, 1961, **32**, 510–519.

44 L. Meng, Y. Zhang, X. Wan, C. Li, X. Zhang, Y. Wang, X. Ke, Z. Xiao, L. Ding, R. Xia, H.-L. Yip, Y. Cao and Y. Chen, *Science*, 2018, **361**, 1094–1098.

45 T. Ameri, N. Li and C. J. Brabec, *Energy Environ. Sci.*, 2013, **6**, 2390.

46 G. Li, W.-H. Chang and Y. Yang, *Nat. Rev. Mater.*, 2017, **2**, 17043.

47 P. Cheng, G. Li, X. Zhan and Y. Yang, *Nat. Photonics*, 2018, **12**, 131–142.

48 K. Ramki, N. Venkatesh, G. Sathiyan, R. Thangamuthu and P. Sakthivel, *Org. Electron.*, 2019, **73**, 182–204.

49 H. Yao, L. Ye, H. Zhang, S. Li, S. Zhang and J. Hou, *Chem. Rev.*, 2016, **116**, 7397–7457.

50 U. Mehmood, A. Al-Ahmed and I. A. Hussein, *Renewable Sustainable Energy Rev.*, 2016, **57**, 550–561.

51 M.-H. Jao, H.-C. Liao and W.-F. Su, *J. Mater. Chem. A*, 2016, **4**, 5784–5801.

52 S. Liu, J. Yuan, W. Deng, M. Luo, Y. Xie, Q. Liang, Y. Zou, Z. He, H. Wu and Y. Cao, *Nat. Photonics*, 2020, **14**, 300–305.

53 R. S. Gurney, W. Li, Y. Yan, D. Liu, A. J. Pearson and T. Wang, *J. Energy Chem.*, 2019, **37**, 148–156.

54 K. Vandewal, J. Benduhn and V. C. Nikolis, *Sustainable Energy Fuels*, 2018, **2**, 538–544.

55 Y. Wang, D. Qian, Y. Cui, H. Zhang, J. Hou, K. Vandewal, T. Kirchartz and F. Gao, *Adv. Energy Mater.*, 2018, **8**, 1801352.

56 J. Sworakowski, J. Lipiński and K. Janus, *Org. Electron.*, 2016, **33**, 300–310.

57 B. Dandrade, S. Datta, S. Forrest, P. Djurovich, E. Polikarpov and M. Thompson, *Org. Electron.*, 2005, **6**, 11–20.

58 R. E. M. Willems, C. H. L. Weijtens, X. de Vries, R. Coehoorn and R. A. J. Janssen, *Adv. Energy Mater.*, 2019, **9**, 1803677.

59 J. Rivnay, S. C. B. Mannsfeld, C. E. Miller, A. Salleo and M. F. Toney, *Chem. Rev.*, 2012, **112**, 5488–5519.

60 F. Liu, Y. Gu, X. Shen, S. Ferdous, H.-W. Wang and T. P. Russell, *Prog. Polym. Sci.*, 2013, **38**, 1990–2052.

61 J. Liu, Y. Sun, X. Gao, R. Xing, L. Zheng, S. Wu, Y. Geng and Y. Han, *Langmuir*, 2011, **27**, 4212–4219.

62 J. Liu, M. Arif, J. Zou, S. I. Khondaker and L. Zhai, *Macromolecules*, 2009, **42**, 9390–9393.

63 J. H. Litofsky, Y. Lee, M. P. Aplan, B. Kuei, A. Hexemer, C. Wang, Q. Wang and E. D. Gomez, *Macromolecules*, 2019, **52**, 2803–2813.

64 Y. Gu, C. Wang and T. P. Russell, *Adv. Energy Mater.*, 2012, **2**, 683–690.

65 C. Yang, J. Zhang, N. Liang, H. Yao, Z. Wei, C. He, X. Yuan and J. Hou, *J. Mater. Chem. A*, 2019, **7**, 18889–18897.

66 V. D. Mihailetchi, L. J. A. Koster, J. C. Hummelen and P. W. M. Blom, *Phys. Rev. Lett.*, 2004, **93**, 216601.

67 M. E. Ziffer, S. B. Jo, H. Zhong, L. Ye, H. Liu, F. Lin, J. Zhang, X. Li, H. W. Ade, A. K.-Y. Jen and D. S. Ginger, *J. Am. Chem. Soc.*, 2018, **140**, 9996–10008.

68 D. Baran, N. Gasparini, A. Wadsworth, C. H. Tan, N. Wehbe, X. Song, Z. Hamid, W. Zhang, M. Neophytou, T. Kirchartz, C. J. Brabec, J. R. Durrant and I. McCulloch, *Nat. Commun.*, 2018, **9**, 2059.

69 S. R. Cowan, A. Roy and A. J. Heeger, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2010, **82**, 245207.

70 L. J. A. Koster, V. D. Mihailetchi, R. Ramaker and P. W. M. Blom, *Appl. Phys. Lett.*, 2005, **86**, 123509.

71 K. Jiang, G. Zhang, G. Yang, J. Zhang, Z. Li, T. Ma, H. Hu, W. Ma, H. Ade and H. Yan, *Adv. Energy Mater.*, 2018, **8**, 1701370.

72 R. Sun, Q. Wu, J. Guo, T. Wang, Y. Wu, B. Qiu, Z. Luo, W. Yang, Z. Hu, J. Guo, M. Shi, C. Yang, F. Huang, Y. Li and J. Min, *Joule*, 2020, **4**, 407–419.

73 L. Arunagiri, Z. Peng, X. Zou, H. Yu, G. Zhang, Z. Wang, J. Y. Lin Lai, J. Zhang, Y. Zheng, C. Cui, F. Huang, Y. Zou, K. S. Wong, P. C. Y. Chow, H. Ade and H. Yan, *Joule*, 2020, **4**, 1790–1805.

74 L. J. A. Koster, V. D. Mihailetchi, H. Xie and P. W. M. Blom, *Appl. Phys. Lett.*, 2005, **87**, 203502.

75 P. Mark and W. Helfrich, *J. Appl. Phys.*, 1962, **33**, 205–215.

76 D. Natali and M. Sampietro, *J. Appl. Phys.*, 2002, **92**, 5310–5318.

77 J. A. Röhr, X. Shi, S. A. Haque, T. Kirchartz and J. Nelson, *Phys. Rev. Appl.*, 2018, **9**, 044017.

78 R. Mauer, M. Kastler and F. Laquai, *Adv. Funct. Mater.*, 2010, **20**, 2085–2092.

79 W. E. Spear, *J. Non-Cryst. Solids*, 1969, **1**, 197–214.

80 G. Juška, K. Arlauskas, M. Viliūnas and J. Kočka, *Phys. Rev. Lett.*, 2000, **84**, 4946–4949.

81 G. Juška, K. Arlauskas, M. Viliūnas, K. Genevičius, R. Österbacka and H. Stubb, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2000, **62**, R16235–R16238.

82 A. Kokil, K. Yang and J. Kumar, *J. Polym. Sci., Part B: Polym. Phys.*, 2012, **50**, 1130–1144.

83 A. J. Mozer, C.-Q. Ma, W. W. H. Wong, D. J. Jones, P. Bäuerle and G. G. Wallace, *Org. Electron.*, 2010, **11**, 573–582.

84 K. Genevičius, R. Österbacka, G. Juška, K. Arlauskas and H. Stubb, *Synth. Met.*, 2003, **137**, 1407–1408.

85 T. Okachi, T. Nagase, T. Kobayashi and H. Naito, *Thin Solid Films*, 2008, **517**, 1331–1334.

86 H. C. F. Martens, H. B. Brom and P. W. M. Blom, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **60**, R8489–R8492.

87 T. M. Clarke and J. R. Durrant, *Chem. Rev.*, 2010, **110**, 6736–6767.

88 M. Engel, F. Kunze, D. C. Lupascu, N. Benson and R. Schmechel, *Phys. Status Solidi RRL*, 2012, **6**, 68–70.

89 B. P. Rand, D. P. Burk and S. R. Forrest, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2007, **75**, 115327.

90 L. Zhu, Y. Yi and Z. Wei, *J. Phys. Chem. C*, 2018, **122**, 22309–22316.

91 S. Chen, S.-W. Tsang, T.-H. Lai, J. R. Reynolds and F. So, *Adv. Mater.*, 2014, **26**, 6125–6131.

92 S. Y. Leblebici, T. L. Chen, P. Olalde-Velasco, W. Yang and B. Ma, *ACS Appl. Mater. Interfaces*, 2013, **5**, 10105–10110.

93 B. Bernardo, D. Cheyns, B. Verreet, R. D. Schaller, B. P. Rand and N. C. Giebink, *Nat. Commun.*, 2014, **5**, 1–7.

94 J. A. Carr and S. Chaudhary, *Appl. Phys. Lett.*, 2012, **100**, 213902.

95 S. Leblebici, J. Lee, A. Weber-Bargioni and B. Ma, *J. Phys. Chem. C*, 2017, **121**, 3279–3285.

96 J. Brebels, J. V. Manca, L. Lutsen, D. Vanderzande and W. Maes, *J. Mater. Chem. A*, 2017, **5**, 24037–24050.

97 H.-W. Li, Z. Guan, Y. Cheng, T. Lui, Q. Yang, C.-S. Lee, S. Chen and S.-W. Tsang, *Adv. Electron. Mater.*, 2016, **2**, 1600200.

98 N. Cho, C. W. Schlenker, K. M. Knesting, P. Koelsch, H.-L. Yip, D. S. Ginger and A. K.-Y. Jen, *Adv. Energy Mater.*, 2014, **4**, 1301857.

99 P. Schilinsky, C. Waldauf and C. J. Brabec, *Appl. Phys. Lett.*, 2002, **81**, 3885–3887.

100 G. Li, V. Shrotriya, J. Huang, Y. Yao, T. Moriarty, K. Emery and Y. Yang, *Nat. Mater.*, 2005, **4**, 864–868.

101 S. Holliday, R. S. Ashraf, A. Wadsworth, D. Baran, S. A. Yousaf, C. B. Nielsen, C.-H. Tan, S. D. Dimitrov, Z. Shang, N. Gasparini, M. Alamoudi, F. Laquai, C. J. Brabec, A. Salleo, J. R. Durrant and I. McCulloch, *Nat. Commun.*, 2016, **7**, 11585.

102 D. Baran, R. S. Ashraf, D. A. Hanifi, M. Abdelsamie, N. Gasparini, J. A. Röhr, S. Holliday, A. Wadsworth, S. Lockett, M. Neophytou, C. J. M. Emmott, J. Nelson, C. J. Brabec, A. Amassian, A. Salleo, T. Kirchartz, J. R. Durrant and I. McCulloch, *Nat. Mater.*, 2017, **16**, 363–369.

103 X. Jia, Z. Chen, C. Duan, Z. Wang, Q. Yin, F. Huang and Y. Cao, *J. Mater. Chem. C*, 2019, **7**, 314–323.

104 X. Xu, G. Zhang, L. Yu, R. Li and Q. Peng, *Adv. Mater.*, 2019, **31**, 1906045.

105 C. Yang, S. Zhang, J. Ren, M. Gao, P. Bi, L. Ye and J. Hou, *Energy Environ. Sci.*, 2020, **13**, 2864–2869.

106 B. Zheng, L. Huo and Y. Li, *NPG Asia Mater.*, 2020, **12**, 3.

107 W. Li, K. H. Hendriks, M. M. Wienk and R. A. J. Janssen, *Acc. Chem. Res.*, 2016, **49**, 78–85.

108 Y.-W. Su, Y.-C. Lin and K.-H. Wei, *J. Mater. Chem. A*, 2017, **5**, 24051–24075.

109 S.-H. Liao, H.-J. Jhuo, Y.-S. Cheng and S.-A. Chen, *Adv. Mater.*, 2013, **25**, 4766–4771.

110 J. Zhao, Q. Li, S. Liu, Z. Cao, X. Jiao, Y.-P. Cai and F. Huang, *ACS Energy Lett.*, 2020, **5**, 367–375.

111 D. Qian, L. Ye, M. Zhang, Y. Liang, L. Li, Y. Huang, X. Guo, S. Zhang, Z. Tan and J. Hou, *Macromolecules*, 2012, **45**, 9611–9617.

112 S. C. Price, A. C. Stuart, L. Yang, H. Zhou and W. You, *J. Am. Chem. Soc.*, 2011, **133**, 4625–4631.

113 Z. Zheng, O. M. Awartani, B. Gautam, D. Liu, Y. Qin, W. Li, A. Bataller, K. Gundogdu, H. Ade and J. Hou, *Adv. Mater.*, 2017, **29**, 1604241.

114 Y. Wang and T. Michinobu, *J. Mater. Chem. C*, 2016, **4**, 6200–6214.

115 C. Chang, W. Li, X. Guo, B. Guo, C. Ye, W. Su, Q. Fan and M. Zhang, *Org. Electron.*, 2018, **58**, 82–87.

116 D. H. Kim, T. T. Trang Bui, S. Rasool, C. E. Song, H. K. Lee, S. K. Lee, J.-C. Lee, W.-W. So and W. S. Shin, *ACS Appl. Mater. Interfaces*, 2019, **11**, 2189–2196.

117 Y. An, X. Liao, L. Chen, J. Yin, Q. Ai, Q. Xie, B. Huang, F. Liu, A. K.-Y. Jen and Y. Chen, *Adv. Funct. Mater.*, 2018, **28**, 1706517.

118 S. Li, L. Ye, W. Zhao, H. Yan, B. Yang, D. Liu, W. Li, H. Ade and J. Hou, *J. Am. Chem. Soc.*, 2018, **140**, 7159–7167.

119 S. Li, W. Zhao, J. Zhang, X. Liu, Z. Zheng, C. He, B. Xu, Z. Wei and J. Hou, *Chem. Mater.*, 2020, **32**, 1993–2003.

120 Y. Cui, H. Yao, L. Hong, T. Zhang, Y. Xu, K. Xian, B. Gao, J. Qin, J. Zhang, Z. Wei and J. Hou, *Adv. Mater.*, 2019, **31**, 1808356.

121 H. Yao, Y. Cui, D. Qian, C. S. Ponseca, A. Honarfar, Y. Xu, J. Xin, Z. Chen, L. Hong, B. Gao, R. Yu, Y. Zu, W. Ma, P. Chabera, T. Pullerits, A. Yartsev, F. Gao and J. Hou, *J. Am. Chem. Soc.*, 2019, **141**, 7743–7750.

122 C. Yao, Y. Zhu, K. Gu, J. Zhao, J. Ning, D. F. Perepichka, Y.-L. Loo and H. Meng, *J. Mater. Chem. A*, 2020, **8**, 12149–12155.

123 B. Fan, D. Zhang, M. Li, W. Zhong, Z. Zeng, L. Ying, F. Huang and Y. Cao, *Sci. China: Chem.*, 2019, **62**, 746–752.

124 U. Huynh, T. Basel, T. Xu, L. Lu, T. Zheng, L. Yu and V. Vardeny, in *Physical Chemistry of Interfaces and Nanomaterials Xiii*, ed. N. Banerji, S. C. Hayes and C. Silva, Spie-Int Soc Optical Engineering, Bellingham, 2014, vol. 9165, p. UNSP91650Z.

125 S. Zhang, M. A. Uddin, W. Zhao, L. Ye, H. Y. Woo, D. Liu, B. Yang, H. Yao, Y. Cui and J. Hou, *Polym. Chem.*, 2015, **6**, 2752–2760.

126 L. Ye, S. Zhang, W. Zhao, H. Yao and J. Hou, *Chem. Mater.*, 2014, **26**, 3603–3605.

127 D. Liu, B. Yang, B. Jang, B. Xu, S. Zhang, C. He, H. Y. Woo and J. Hou, *Energy Environ. Sci.*, 2017, **10**, 546–551.

128 Q. Fan, Q. Zhu, Z. Xu, W. Su, J. Chen, J. Wu, X. Guo, W. Ma, M. Zhang and Y. Li, *Nano Energy*, 2018, **48**, 413–420.

129 M. Zhang, X. Guo, W. Ma, H. Ade and J. Hou, *Adv. Mater.*, 2015, **27**, 4655–4660.

130 R. Ma, T. Liu, Z. Luo, Q. Guo, Y. Xiao, Y. Chen, X. Li, S. Luo, X. Lu, M. Zhang, Y. Li and H. Yan, *Sci. China: Chem.*, 2020, **63**, 325–330.

131 Y. Wu, C. An, L. Shi, L. Yang, Y. Qin, N. Liang, C. He, Z. Wang and J. Hou, *Angew. Chem., Int. Ed.*, 2018, **57**, 12911–12915.

132 Z. Xu, Q. Fan, X. Meng, X. Guo, W. Su, W. Ma, M. Zhang and Y. Li, *Chem. Mater.*, 2017, **29**, 4811–4818.

133 Q. Fan, Z. Xu, X. Guo, X. Meng, W. Li, W. Su, X. Ou, W. Ma, M. Zhang and Y. Li, *Nano Energy*, 2017, **40**, 20–26.

134 L. Ye, Y. Xie, K. Weng, H. S. Ryu, C. Li, Y. Cai, H. Fu, D. Wei, H. Y. Woo, S. Tan and Y. Sun, *Nano Energy*, 2019, **58**, 220–226.

135 W. Li, G. Li, H. Guo, X. Guo, B. Guo, Q. Zhu, Q. Fan, W. Ma, M. Zhang and Y. Li, *J. Mater. Chem. A*, 2019, **7**, 1307–1314.

136 Y. Wu, H. Yang, Y. Zou, Y. Dong, J. Yuan, C. Cui and Y. Li, *Energy Environ. Sci.*, 2019, **12**, 675–683.

137 Y. Li, D. Liu, J. Wang, Z.-G. Zhang, Y. Li, Y. Liu, T. Zhu, X. Bao, M. Sun and R. Yang, *Chem. Mater.*, 2017, **29**, 8249–8257.

138 P. Chao, M. Guo, Y. Zhu, H. Chen, M. Pu, H.-H. Huang, H. Meng, C. Yang and F. He, *Macromolecules*, 2020, **53**, 2893–2901.

139 H. Sun, T. Liu, J. Yu, T.-K. Lau, G. Zhang, Y. Zhang, M. Su, Y. Tang, R. Ma, B. Liu, J. Liang, K. Feng, X. Lu, X. Guo, F. Gao and H. Yan, *Energy Environ. Sci.*, 2019, **12**, 3328–3337.

140 H. Bin, Y. Yang, Z. Peng, L. Ye, J. Yao, L. Zhong, C. Sun, L. Gao, H. Huang, X. Li, B. Qiu, L. Xue, Z.-G. Zhang, H. Ade and Y. Li, *Adv. Energy Mater.*, 2018, **8**, 1702324.

141 H. Bin, L. Gao, Z.-G. Zhang, Y. Yang, Y. Zhang, C. Zhang, S. Chen, L. Xue, C. Yang, M. Xiao and Y. Li, *Nat. Commun.*, 2016, **7**, 13651.

142 H. Bin, L. Zhong, Y. Yang, L. Gao, H. Huang, C. Sun, X. Li, L. Xue, Z.-G. Zhang, Z. Zhang and Y. Li, *Adv. Energy Mater.*, 2017, **7**, 1700746.

143 L. Xue, Y. Yang, J. Xu, C. Zhang, H. Bin, Z.-G. Zhang, B. Qiu, X. Li, C. Sun, L. Gao, J. Yao, X. Chen, Y. Yang, M. Xiao and Y. Li, *Adv. Mater.*, 2017, **29**, 1703344.

144 Z. Liao, Y. Xie, L. Chen, Y. Tan, S. Huang, Y. An, H. S. Ryu, X. Meng, X. Liao, B. Huang, Q. Xie, H. Y. Woo, Y. Sun and Y. Chen, *Adv. Funct. Mater.*, 2019, **29**, 1808828.

145 W. Su, G. Li, Q. Fan, Q. Zhu, X. Guo, J. Chen, J. Wu, W. Ma, M. Zhang and Y. Li, *J. Mater. Chem. A*, 2019, **7**, 2351–2359.

146 Z. Li, X. Xu, G. Zhang, T. Yu, Y. Li and Q. Peng, *Sol. RRL*, 2018, **2**, 1800186.

147 W. Li, G. Li, X. Guo, B. Guo, Z. Bi, H. Guo, W. Ma, X. Ou, M. Zhang and Y. Li, *J. Mater. Chem. A*, 2017, **5**, 19680–19686.

148 W. Li, G. Li, X. Guo, Y. Wang, H. Guo, Q. Xu, M. Zhang and Y. Li, *J. Mater. Chem. A*, 2018, **6**, 6551–6558.

149 W. Chen, G. Huang, X. Li, Y. Li, H. Wang, H. Jiang, Z. Zhao, D. Yu, E. Wang and R. Yang, *ACS Appl. Mater. Interfaces*, 2019, **11**, 33173–33178.

150 D. Liu, K. Zhang, Y. Zhong, C. Gu, Y. Li and R. Yang, *J. Mater. Chem. A*, 2018, **6**, 18125–18132.

151 L. Lan, Z. Chen, Q. Hu, L. Ying, R. Zhu, F. Liu, T. P. Russell, F. Huang and Y. Cao, *Adv. Sci.*, 2016, **3**, 1600032.

152 A. Tang, Q. Zhang, M. Du, G. Li, Y. Geng, J. Zhang, Z. Wei, X. Sun and E. Zhou, *Macromolecules*, 2019, **52**, 6227–6233.

153 S. Xu, X. Wang, L. Feng, Z. He, H. Peng, V. Cimrová, J. Yuan, Z.-G. Zhang, Y. Li and Y. Zou, *J. Mater. Chem. A*, 2018, **6**, 3074–3083.

154 W. Li, Q. Liu, K. Jin, M. Cheng, F. Hao, W.-Q. Wu, S. Liu, Z. Xiao, S. Yang, S. Shi and L. Ding, *Mater. Chem. Front.*, 2020, **4**, 1454–1458.

155 T. Yu, X. Xu, G. Zhang, J. Wan, Y. Li and Q. Peng, *Adv. Funct. Mater.*, 2017, **27**, 1701491.

156 X. Xu, T. Yu, Z. Bi, W. Ma, Y. Li and Q. Peng, *Adv. Mater.*, 2018, **30**, 1703973.

157 G. Xu, L. Chen, H. Lei, Z. Liao, N. Yi, J. Liu and Y. Chen, *J. Mater. Chem. A*, 2019, **7**, 4145–4152.

158 S. Wen, Y. Li, N. Zheng, I. O. Raji, C. Yang and X. Bao, *J. Mater. Chem. A*, 2020, **8**, 13671–13678.

159 S. J. Jeon, Y. W. Han and D. K. Moon, *Small*, 2019, **15**, 1902598.

160 J. Xiong, K. Jin, Y. Jiang, J. Qin, T. Wang, J. Liu, Q. Liu, H. Peng, X. Li, A. Sun, X. Meng, L. Zhang, L. Liu, W. Li, Z. Fang, X. Jia, Z. Xiao, Y. Feng, X. Zhang, K. Sun, S. Yang, S. Shi and L. Ding, *Sci. Bull.*, 2019, **64**, 1573–1576.

161 J. Liu, L. Liu, C. Zuo, Z. Xiao, Y. Zou, Z. Jin and L. Ding, *Sci. Bull.*, 2019, **64**, 1655–1657.

162 T. Wang, J. Qin, Z. Xiao, X. Meng, C. Zuo, B. Yang, H. Tan, J. Yang, S. Yang, K. Sun, S. Xie and L. Ding, *Sci. Bull.*, 2020, **65**, 179–181.

163 Y. Qin, M. A. Uddin, Y. Chen, B. Jang, K. Zhao, Z. Zheng, R. Yu, T. J. Shin, H. Y. Woo and J. Hou, *Adv. Mater.*, 2016, **28**, 9416–9422.

164 Y.-S. Wu, Y.-C. Lin, S.-Y. Hung, C.-K. Chen, Y.-C. Chiang, C.-C. Chueh and W.-C. Chen, *Macromolecules*, 2020, **53**, 4968–4981.

165 K. He, X. Li, H. Liu, Z. Zhang, P. Kumar, J. H. L. Ngai, J. Wang and Y. Li, *Asian J. Org. Chem.*, 2020, **9**, 1301–1308.

166 J. Yu, P. Chen, C. W. Koh, H. Wang, K. Yang, X. Zhou, B. Liu, Q. Liao, J. Chen, H. Sun, H. Y. Woo, S. Zhang and X. Guo, *Adv. Sci.*, 2018, 1801743.

167 C. Sun, F. Pan, H. Bin, J. Zhang, L. Xue, B. Qiu, Z. Wei, Z.-G. Zhang and Y. Li, *Nat. Commun.*, 2018, **9**, 743.

168 D. Yuan, F. Pan, L. Zhang, H. Jiang, M. Chen, W. Tang, G. Qin, Y. Cao and J. Chen, *Sol. RRL*, 2020, 2000062.

169 F. Li, A. Tang, B. Zhang and E. Zhou, *ACS Macro Lett.*, 2019, **8**, 1599–1604.

170 D. Tang, J. Wan, X. Xu, Y. W. Lee, H. Y. Woo, K. Feng and Q. Peng, *Nano Energy*, 2018, **53**, 258–269.

171 C. Sun, S. Qin, R. Wang, S. Chen, F. Pan, B. Qiu, Z. Shang, L. Meng, C. Zhang, M. Xiao, C. Yang and Y. Li, *J. Am. Chem. Soc.*, 2020, **142**, 1465–1474.

172 W. Su, Q. Fan, X. Guo, X. Meng, Z. Bi, W. Ma, M. Zhang and Y. Li, *Nano Energy*, 2017, **38**, 510–517.

173 H. Hwang, D. H. Sin, C. Park and K. Cho, *Sci. Rep.*, 2019, **9**, 12081.

174 Q. Li, Y. Sun, X. Xue, S. Yue, K. Liu, M. Azam, C. Yang, Z. Wang, F. Tan and Y. Chen, *ACS Appl. Mater. Interfaces*, 2019, **11**, 3299–3307.

175 M.-A. Pan, T.-K. Lau, Y. Tang, Y.-C. Wu, T. Liu, K. Li, M.-C. Chen, X. Lu, W. Ma and C. Zhan, *J. Mater. Chem. A*, 2019, **7**, 20713–20722.

176 J. Wang, J. Xu, N. Yao, D. Zhang, Z. Zheng, S. Xie, X. Zhang, F. Zhang, H. Zhou, C. Zhang and Y. Zhang, *J. Phys. Chem. Lett.*, 2019, **10**, 4110–4116.

177 Q. Yue, H. Wu, Z. Zhou, M. Zhang, F. Liu and X. Zhu, *Adv. Mater.*, 2019, **31**, 1904283.

178 B. Fan, K. Zhang, X.-F. Jiang, L. Ying, F. Huang and Y. Cao, *Adv. Mater.*, 2017, **29**, 1606396.

179 Y. He and Y. Li, *Phys. Chem. Chem. Phys.*, 2011, **13**, 1970–1983.

180 C. Zhang, S. Langner, A. V. Mumyatov, D. V. Anokhin, J. Min, J. D. Perea, K. L. Gerasimov, A. Osvet, D. A. Ivanov, P. Troshin, N. Li and C. J. Brabec, *J. Mater. Chem. A*, 2017, **5**, 17570–17579.

181 W. Liu, J. Zhang, Z. Zhou, D. Zhang, Y. Zhang, S. Xu and X. Zhu, *Adv. Mater.*, 2018, **30**, 1800403.

182 J. Sun, X. Ma, Z. Zhang, J. Yu, J. Zhou, X. Yin, L. Yang, R. Geng, R. Zhu, F. Zhang and W. Tang, *Adv. Mater.*, 2018, **30**, 1707150.

183 C. Huang, X. Liao, K. Gao, L. Zuo, F. Lin, X. Shi, C.-Z. Li, H. Liu, X. Li, F. Liu, Y. Chen, H. Chen and A. K.-Y. Jen, *Chem. Mater.*, 2018, **30**, 5429–5434.

184 W. Gao, T. Liu, R. Ming, Z. Luo, K. Wu, L. Zhang, J. Xin, D. Xie, G. Zhang, W. Ma, H. Yan and C. Yang, *Adv. Funct. Mater.*, 2018, **28**, 1803128.

185 J. Cao, S. Qu, J. Yu, Z. Zhang, R. Geng, L. Yang, H. Wang, F. Du and W. Tang, *Mater. Chem. Front.*, 2020, **4**, 924–932.

186 H. Bin, L. Zhong, Z.-G. Zhang, L. Gao, Y. Yang, L. Xue, J. Zhang, Z. Zhang and Y. Li, *Sci. China: Chem.*, 2016, **59**, 1317–1322.

187 J. Min, Z.-G. Zhang, S. Zhang and Y. Li, *Chem. Mater.*, 2012, **24**, 3247–3254.

188 Y. Li, L. Zhong, J.-D. Lin, F.-P. Wu, H.-J. Bin, Z. Zhang, L. Xu, Z.-Q. Jiang, Z.-G. Zhang, F. Liu, T. P. Russell, Y. Li, L.-S. Liao and S. R. Forrest, *Sol. RRL*, 2017, **1**, 1700107.

189 A. Patra and M. Bendikov, *J. Mater. Chem.*, 2010, **20**, 422–433.

190 S. S. Zade, N. Zamoshchik and M. Bendikov, *Chem. – Eur. J.*, 2009, **15**, 8613–8624.

191 L. Liu, G. Zhang, B. He, S. Liu, C. Duan and F. Huang, *Mater. Chem. Front.*, 2017, **1**, 499–506.

192 Y. Xu, H. Yao, L. Ma, J. Wang and J. Hou, *Rep. Prog. Phys.*, 2020, **83**, 082601.

193 A. C. Jakowetz, M. L. Böhm, A. Sadhanala, S. Huettner, A. Rao and R. H. Friend, *Nat. Mater.*, 2017, **16**, 551–557.

194 S. Gélinas, A. Rao, A. Kumar, S. L. Smith, A. W. Chin, J. Clark, T. S. van der Poll, G. C. Bazan and R. H. Friend, *Science*, 2014, **343**, 512–516.

195 B. A. Gregg, *J. Phys. Chem. Lett.*, 2011, **2**, 3013–3015.

196 Y. Dong, X. Hu, C. Duan, P. Liu, S. Liu, L. Lan, D. Chen, L. Ying, S. Su, X. Gong, F. Huang and Y. Cao, *Adv. Mater.*, 2013, **25**, 3683–3688.

197 M. Wang, X. Hu, P. Liu, W. Li, X. Gong, F. Huang and Y. Cao, *J. Am. Chem. Soc.*, 2011, **133**, 9638–9641.

198 Y. Li, J. Geng, Y. Liu, S. Yu and G. Zhao, *ChemMedChem*, 2013, **8**, 27–41.

199 R. Po, G. Bianchi, C. Carbonera and A. Pellegrino, *Macromolecules*, 2015, **48**, 453–461.

200 Q. Fan, W. Su, Y. Wang, B. Guo, Y. Jiang, X. Guo, F. Liu, T. P. Russell, M. Zhang and Y. Li, *Sci. China: Chem.*, 2018, **61**, 531–537.

201 S. Zhang, Y. Qin, J. Zhu and J. Hou, *Adv. Mater.*, 2018, **30**, 1800868.

202 C. Cui, X. Fan, M. Zhang, J. Zhang, J. Min and Y. Li, *Chem. Commun.*, 2011, **47**, 11345.

203 Q. Fan, W. Su, X. Meng, X. Guo, G. Li, W. Ma, M. Zhang and Y. Li, *Sol. RRL*, 2017, **1**, 1700020.

204 B. Guo, W. Li, X. Guo, X. Meng, W. Ma, M. Zhang and Y. Li, *Adv. Mater.*, 2017, **29**, 1702291.

205 S. J. Jeon, Y. W. Han and D. K. Moon, *Sol. RRL*, 2019, **3**, 1900094.

206 X. Guo, N. Zhou, S. J. Lou, J. Smith, D. B. Tice, J. W. Hennek, R. P. Ortiz, J. T. L. Navarrete, S. Li, J. Strzalka, L. X. Chen, R. P. H. Chang, A. Facchetti and T. J. Marks, *Nat. Photonics*, 2013, **7**, 825–833.

207 Y. Xie, Y. Cai, L. Zhu, R. Xia, L. Ye, X. Feng, H. Yip, F. Liu, G. Lu, S. Tan and Y. Sun, *Adv. Funct. Mater.*, 2020, 2002181.

208 Y. Wu, Y. Zheng, H. Yang, C. Sun, Y. Dong, C. Cui, H. Yan and Y. Li, *Sci. China: Chem.*, 2020, **63**, 265–271.

209 Y. Lin, F. Zhao, Y. Wu, K. Chen, Y. Xia, G. Li, S. K. K. Prasad, J. Zhu, L. Huo, H. Bin, Z.-G. Zhang, X. Guo, M. Zhang, Y. Sun, F. Gao, Z. Wei, W. Ma, C. Wang, J. Hodgkiss, Z. Bo, O. Inganäs, Y. Li and X. Zhan, *Adv. Mater.*, 2017, **29**, 1604155.

210 Y. Geng, A. Tang, K. Tajima, Q. Zeng and E. Zhou, *J. Mater. Chem. A*, 2019, **7**, 64–96.

211 J. Zhang, L. Zhu and Z. Wei, *Small Methods*, 2017, **1**, 1700258.

212 H. Lee, C. Park, D. H. Sin, J. H. Park and K. Cho, *Adv. Mater.*, 2018, **30**, 1800453.

213 S. Namuangruk, S. Jungsuttiwong, N. Kungwan, V. Promarak, T. Sudyoadsuk, B. Jansang and M. Ehara, *Theor. Chem. Acc.*, 2016, **135**, 14.

214 B. F. Minaev, G. V. Baryshnikov and V. A. Minaeva, *Dyes Pigm.*, 2012, **92**, 531–536.

215 Z. Sun, Y. Jiang, L. Zeng and L. Huang, *ChemSusChem*, 2019, **12**, 1325–1333.

216 S. Jungsuttiwong, T. Yakhanthip, J. Surakhot, J. Khunchalee, T. Sudyoadsuk, V. Promarak, N. Kungwan and S. Namuangruk, *J. Comput. Chem.*, 2012, **33**, 1517–1523.

217 H. Yao, J. Wang, Y. Xu, S. Zhang and J. Hou, *Acc. Chem. Res.*, 2020, **53**, 822–832.

218 Q. Zhang, M. A. Kelly, N. Bauer and W. You, *Acc. Chem. Res.*, 2017, **50**, 2401–2409.

219 J. Yuan, L. Qiu, Z.-G. Zhang, Y. Li, Y. Chen and Y. Zou, *Nano Energy*, 2016, **30**, 312–320.

220 Y. Wu, Y. Zou, H. Yang, Y. Li, H. Li, C. Cui and Y. Li, *ACS Appl. Mater. Interfaces*, 2017, **9**, 37078–37086.

221 Y. Li, N. Zheng, L. Yu, S. Wen, C. Gao, M. Sun and R. Yang, *Adv. Mater.*, 2019, **31**, 1807832.

222 X. Jiao, L. Ye and H. Ade, *Adv. Energy Mater.*, 2017, **7**, 1700084.

223 L. Ye, W. Zhao, S. Li, S. Mukherjee, J. H. Carpenter, O. Awartani, X. Jiao, J. Hou and H. Ade, *Adv. Energy Mater.*, 2017, **7**, 1602000.

224 J. Song, M. Zhang, M. Yuan, Y. Qian, Y. Sun and F. Liu, *Small Methods*, 2018, **2**, 1700229.

225 T. Liu, L. Huo, S. Chandrabose, K. Chen, G. Han, F. Qi, X. Meng, D. Xie, W. Ma, Y. Yi, J. M. Hodgkiss, F. Liu, J. Wang, C. Yang and Y. Sun, *Adv. Mater.*, 2018, **30**, 1707353.

226 J. Chen, Z. Bi, X. Xu, Q. Zhang, S. Yang, S. Guo, H. Yan, W. You and W. Ma, *Adv. Sci.*, 2019, **6**, 1801560.

227 L. Yang, Z. Hu, Z. Zhang, J. Cao, H. Wang, J. Yu, F. Zhang and W. Tang, *J. Mater. Chem. A*, 2020, **8**, 5458–5466.

228 L. Ye, B. A. Collins, X. Jiao, J. Zhao, H. Yan and H. Ade, *Adv. Energy Mater.*, 2018, **8**, 1703058.

229 H. Fu, Z. Wang and Y. Sun, *Angew. Chem., Int. Ed.*, 2019, **58**, 4442–4453.

230 S. Xiao, Q. Zhang and W. You, *Adv. Mater.*, 2017, **29**, 1601391.

231 S. Holliday, Y. Li and C. K. Luscombe, *Prog. Polym. Sci.*, 2017, **70**, 34–51.

232 L. Ye, S. Li, X. Liu, S. Zhang, M. Ghasemi, Y. Xiong, J. Hou and H. Ade, *Joule*, 2019, **3**, 443–458.

233 M. Ghasemi, H. Hu, Z. Peng, J. J. Rech, I. Angunawela, J. H. Carpenter, S. J. Stuard, A. Wadsworth, I. McCulloch, W. You and H. Ade, *Joule*, 2019, **3**, 1328–1348.

234 L. Ye, H. Hu, M. Ghasemi, T. Wang, B. A. Collins, J.-H. Kim, K. Jiang, J. H. Carpenter, H. Li, Z. Li, T. McAfee, J. Zhao, X. Chen, J. L. Y. Lai, T. Ma, J.-L. Bredas, H. Yan and H. Ade, *Nat. Mater.*, 2018, **17**, 253–260.

235 N. Yi, Q. Ai, W. Zhou, L. Huang, L. Zhang, Z. Xing, X. Li, J. Zeng and Y. Chen, *Chem. Mater.*, 2019, **31**, 10211–10224.

236 X. Xue, K. Weng, F. Qi, Y. Zhang, Z. Wang, J. Ali, D. Wei, Y. Sun, F. Liu, M. Wan, J. Liu and L. Huo, *Adv. Energy Mater.*, 2019, **9**, 1802686.

237 H. Cha, G. Fish, J. Luke, A. Alraddadi, H. H. Lee, W. Zhang, Y. Dong, S. Limbu, A. Wadsworth, I. P. Maria, L. Francàs, H. L. Sou, T. Du, J.-S. Kim, M. A. McLachlan, I. McCulloch and J. R. Durrant, *Adv. Energy Mater.*, 2019, **9**, 1901254.

238 Z. Zheng, H. Yao, L. Ye, Y. Xu, S. Zhang and J. Hou, *Mater. Today*, 2020, **35**, 115–130.

239 L.-M. Wang, Q. Li, S. Liu, Z. Cao, Y.-P. Cai, X. Jiao, H. Lai, W. Xie, X. Zhan and T. Zhu, *ACS Appl. Mater. Interfaces*, 2020, **12**, 24165–24173.

240 C. Jiao, C. Pang and Q. An, *Int. J. Energy Res.*, 2019, **43**, 8716–8724.

241 M. Zhang, F. Zhang, Q. An, Q. Sun, W. Wang, X. Ma, J. Zhang and W. Tang, *J. Mater. Chem. A*, 2017, **5**, 3589–3598.

242 M. B. Upama, N. K. Elumalai, M. A. Mahmud, M. Wright, D. Wang, C. Xu and A. Uddin, *Sol. Energy Mater. Sol. Cells*, 2018, **176**, 109–118.

243 W. Li, J. Cai, Y. Yan, F. Cai, S. Li, R. S. Gurney, D. Liu, J. D. McGettrick, T. M. Watson, Z. Li, A. J. Pearson, D. G. Lidzey, J. Hou and T. Wang, *Sol. RRL*, 2018, **2**, 1800114.

244 W. Ma, C. Yang, X. Gong, K. Lee and A. J. Heeger, *Adv. Funct. Mater.*, 2005, **15**, 1617–1622.

245 H. Hu, K. Jiang, P. C. Y. Chow, L. Ye, G. Zhang, Z. Li, J. H. Carpenter, H. Ade and H. Yan, *Adv. Energy Mater.*, 2018, **8**, 1701674.

246 S. Li, L. Zhan, Y. Jin, G. Zhou, T. Lau, R. Qin, M. Shi, C. Li, H. Zhu, X. Lu, F. Zhang and H. Chen, *Adv. Mater.*, 2020, 2001160.

247 J. Yuan, T. Huang, P. Cheng, Y. Zou, H. Zhang, J. L. Yang, S.-Y. Chang, Z. Zhang, W. Huang, R. Wang, D. Meng, F. Gao and Y. Yang, *Nat. Commun.*, 2019, **10**, 570.

248 D. Qian, Z. Zheng, H. Yao, W. Tress, T. R. Hopper, S. Chen, S. Li, J. Liu, S. Chen, J. Zhang, X.-K. Liu, B. Gao, L. Ouyang, Y. Jin, G. Pozina, I. A. Buyanova, W. M. Chen, O. Inganäs, V. Coropceanu, J.-L. Bredas, H. Yan, J. Hou, F. Zhang, A. A. Bakulin and F. Gao, *Nat. Mater.*, 2018, **17**, 703–709.

249 L. Hong, H. Yao, R. Yu, Y. Xu, B. Gao, Z. Ge and J. Hou, *ACS Appl. Mater. Interfaces*, 2019, **11**, 29124–29131.

250 S. M. Menke, N. A. Ran, G. C. Bazan and R. H. Friend, *Joule*, 2018, **2**, 25–35.

251 G. Dennler, M. C. Scharber and C. J. Brabec, *Adv. Mater.*, 2009, **21**, 1323–1338.

252 M. Jeong, I. W. Choi, E. M. Go, Y. Cho, M. Kim, B. Lee, S. Jeong, Y. Jo, H. W. Choi, J. Lee, J.-H. Bae, S. K. Kwak, D. S. Kim and C. Yang, *Science*, 2020, **369**, 1615.

253 Y. Ji, L. Xu, X. Hao and K. Gao, *Sol. RRL*, 2020, **4**, 2000130.

254 Z.-H. Chen, P.-Q. Bi, X.-Y. Yang, M.-S. Niu, K.-N. Zhang, L. Feng and X.-T. Hao, *J. Phys. Chem. C*, 2019, **123**, 12676–12683.

255 J. H. Bombile, M. J. Janik and S. T. Milner, *Phys. Chem. Chem. Phys.*, 2019, **21**, 11999–12011.

256 X. Jiang, J. Yang, S. Karuthedath, J. Li, W. Lai, C. Li, C. Xiao, L. Ye, Z. Ma, Z. Tang, F. Laquai and W. Li, *Angew. Chem., Int. Ed.*, 2020, **59**, 2–12.

257 K. Wang, Y. Li and Y. Li, *Macromol. Rapid Commun.*, 2020, **41**, 1900437.

258 W. R. Mateker and M. D. McGehee, *Adv. Mater.*, 2017, **29**, 1603940.

259 P. Cheng and X. Zhan, *Chem. Soc. Rev.*, 2016, **45**, 2544–2582.

260 M. Jørgensen, K. Norrman, S. A. Gevorgyan, T. Tromholt, B. Andreasen and F. C. Krebs, *Adv. Mater.*, 2012, **24**, 580–612.

261 M. Giannouli, V. M. Drakonakis, A. Savva, P. Eleftheriou, G. Florides and S. A. Choulis, *ChemPhysChem*, 2015, **16**, 1134–1154.

262 G. A. dos, R. Benatto, B. Roth, M. Corazza, R. R. Søndergaard, S. A. Gevorgyan, M. Jørgensen and F. C. Krebs, *Nanoscale*, 2016, **8**, 318–326.

263 X. Xu, J. Xiao, G. Zhang, L. Wei, X. Jiao, H.-L. Yip and Y. Cao, *Sci. Bull.*, 2020, **65**, 208–216.

264 J. Xiao, M. Ren, G. Zhang, J. Wang, D. Zhang, L. Liu, N. Li, C. J. Brabec, H.-L. Yip and Y. Cao, *Sol. RRL*, 2019, **3**, 1900077.

265 J. Guo, Y. Wu, R. Sun, W. Wang, J. Guo, Q. Wu, X. Tang, C. Sun, Z. Luo, K. Chang, Z. Zhang, J. Yuan, T. Li, W. Tang, E. Zhou, Z. Xiao, L. Ding, Y. Zou, X. Zhan, C. Yang, Z. Li, C. J. Brabec, Y. Li and J. Min, *J. Mater. Chem. A*, 2019, **7**, 25088–25101.

266 X. Du, T. Heumueller, W. Gruber, A. Classen, T. Unruh, N. Li and C. J. Brabec, *Joule*, 2019, **3**, 215–226.

267 N. Y. Doumon, M. V. Dryzhov, F. V. Houard, V. M. Le Corre, A. Rahimi Chatri, P. Christodoulis and L. J. A. Koster, *ACS Appl. Mater. Interfaces*, 2019, **11**, 8310–8318.

268 Z. Li, W. Zhong, L. Ying, F. Liu, N. Li, F. Huang and Y. Cao, *Nano Energy*, 2019, **64**, 103931.

269 Y. Zhang, Y. Xu, M. J. Ford, F. Li, J. Sun, X. Ling, Y. Wang, J. Gu, J. Yuan and W. Ma, *Adv. Energy Mater.*, 2018, **8**, 1800029.

270 K. Zhang, R. Xia, B. Fan, X. Liu, Z. Wang, S. Dong, H.-L. Yip, L. Ying, F. Huang and Y. Cao, *Adv. Mater.*, 2018, **30**, 1803166.

271 Y.-Y. Yu, C.-H. Chen, C.-C. Chueh, C.-Y. Chiang, J.-H. Hsieh, C.-P. Chen and W.-C. Chen, *ACS Appl. Mater. Interfaces*, 2017, **9**, 27853–27862.

272 T. Kim, R. Younts, W. Lee, S. Lee, K. Gundogdu and B. J. Kim, *J. Mater. Chem. A*, 2017, **5**, 22170–22179.

273 T. Kim, J. Choi, H. J. Kim, W. Lee and B. J. Kim, *Macromolecules*, 2017, **50**, 6861–6871.

274 T. Kim, J.-H. Kim, T. E. Kang, C. Lee, H. Kang, M. Shin, C. Wang, B. Ma, U. Jeong, T.-S. Kim and B. J. Kim, *Nat. Commun.*, 2015, **6**, 8547.

275 Q. Burlingame, X. Huang, X. Liu, C. Jeong, C. Coburn and S. R. Forrest, *Nature*, 2019, **573**, 394–397.

276 X. Meng, L. Zhang, Y. Xie, X. Hu, Z. Xing, Z. Huang, C. Liu, L. Tan, W. Zhou, Y. Sun, W. Ma and Y. Chen, *Adv. Mater.*, 2019, **31**, 1903649.

277 S. H. Park, S. Park, S. Lee, J. Kim, H. Ahn, B. J. Kim, B. Chae and H. J. Son, *Nano Energy*, 2020, **77**, 105147.

278 P. Meredith and A. Armin, *Nat. Commun.*, 2018, **9**, 1–4.

279 J. Guo and J. Min, *Adv. Energy Mater.*, 2019, **9**, 1802521.

280 A. Gambhir, P. Sandwell and J. Nelson, *Sol. Energy Mater. Sol. Cells*, 2016, **156**, 49–58.

281 N. Li and C. J. Brabec, *Nat. Energy*, 2017, **2**, 772–773.

282 N. Li, I. McCulloch and C. J. Brabec, *Energy Environ. Sci.*, 2018, **11**, 1355–1361.

283 J. Min, Y. N. Luponosov, C. Cui, B. Kan, H. Chen, X. Wan, Y. Chen, S. A. Ponomarenko, Y. Li and C. J. Brabec, *Adv. Energy Mater.*, 2017, **7**, 1700465.

284 E. Bundgaard, F. Livi, O. Hagemann, J. E. Carlé, M. Helgesen, I. M. Heckler, N. K. Zawacka, D. Angmo, T. T. Larsen-Olsen, G. A. dos R. Benatto, B. Roth, M. V. Madsen, M. R. Andersson, M. Jørgensen, R. R. Søndergaard and F. C. Krebs, *Adv. Energy Mater.*, 2015, **5**, 1402186.

285 D. Hu, Q. Yang, H. Chen, F. Wobben, V. M. L. Corre, R. Singh, T. Liu, R. Ma, H. Tang, L. J. A. Koster, T. Duan, H. Yan, Z. Kan, Z. Xiao and S. Lu, *Energy Environ. Sci.*, 2020, **13**, 2134–2141.

286 R. Ilmi, A. Haque and M. S. Khan, *Org. Electron.*, 2018, **58**, 53–62.

287 S. Dong, K. Zhang, B. Xie, J. Xiao, H.-L. Yip, H. Yan, F. Huang and Y. Cao, *Adv. Energy Mater.*, 2019, **9**, 1802832.

288 K. An, W. Zhong and L. Ying, *Org. Electron.*, 2020, **82**, 105701.

289 S. Strohm, F. Machui, S. Langner, P. Kubis, N. Gasparini, M. Salvador, I. McCulloch, H.-J. Egelhaaf and C. J. Brabec, *Energy Environ. Sci.*, 2018, **11**, 2225–2234.

290 Y. Qin, L. Ye, S. Zhang, J. Zhu, B. Yang, H. Ade and J. Hou, *J. Mater. Chem. A*, 2018, **6**, 4324–4330.