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Photocatalytic CO$_2$ reduction is a highly vital process for converting CO$_2$ into valuable chemicals. However, the reaction always proceeds less efficaciously at low temperature. The combination of photo and thermal conditions is one of feasible approaches to achieve the reaction with high efficiency and has gained much attention recently. In the present work, trimetallic Co–Cu–Mn catalysts were developed in photothermal CO$_2$ hydrogenation to CH$_4$ and C$_2$ hydrocarbons. The structural and photoelectric characteristics, as well as adsorption behaviors of these catalysts affected the catalytic performances remarkably.

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Photothermal CO₂ hydrogenation to hydrocarbons over trimetallic Co–Cu–Mn catalysts†

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Photocatalytic CO₂ reduction is a highly vital process for converting CO₂ into valuable chemicals. However, the reaction always proceeds less efficaciously at low temperature. A combination of optical and thermal conditions is one of the feasible approaches to achieve the reaction with high efficiency and has gained much attention recently. In the present work, we prepared several Co–Cu–Mn trimetallic catalysts via a simple co-precipitation method, which were used in catalyzing photothermal CO₂ reduction to hydrocarbons. The metal composition and reduction temperature of the catalysts had important effects on their structural and photoelectrical characteristics and adsorption behaviors, further resulting in diverse catalytic performances. Among the prepared trimetallic catalysts, Co₇Cu₁Mn₁Oₓ (200), with a Co/Cu/Mn molar ratio of 7/1/1 and reduced at 200 °C in H₂ for 2 h, could produce CH₄ with an activity of 14.5 mmol g⁻¹ h⁻¹ in 10% CO₂/30% H₂/60% N₂, and CH₄ and C₂+ hydrocarbons with the activities of 15.9 and 7.5 mmol g⁻¹ h⁻¹ in 25% CO₂/75% H₂, respectively. The present strategy for constructing trimetallic oxide catalysts for the photothermal reaction not only provides a highly active catalyst for CO₂ utilization, but also offers a potential possibility for reducing the high temperature of conventional thermal reactions.

Introduction

CO₂ is the primary greenhouse gas, whose concentration in the atmosphere continues to rise.1 It is also an important C1 source and can be converted into many valuable chemicals such as CO,2 hydrocarbons,3–5 olefins,6 acids,7–9 alcohols,10–12 etc. Although many efforts have been invested in this area, it is still very difficult to convert CO₂ into chemicals due to its thermodynamic stability and chemical inertness.1 CO₂ hydrogenation to hydrocarbons represents an essential and green way for its utilization.13 However, the conventional thermal process is always conducted under harsh conditions, leading to the formation of carbon deposit, the increase of energy consumption, etc. The photocatalytic method is an ideal way for the reaction, which could be carried out under mild conditions and using solar energy. However, this route suffers from low activity especially at room temperature, impeding its further applications. Recently, the photothermal catalytic method has gained increasing attention, especially in CO₂ reduction.14,15

The photothermal routes always contain four types, including thermal-assisted photocatalysis, photo-assisted thermocatalysis, photothermal co-catalysis, and photo-driven thermocatalysis.16 Generally, these routes combine the advantages of both photocatalysis and thermal catalysis including high activity, proceeding under mild conditions, efficient energy utilization, etc. To date, many catalysts such as Ru/silicon nanowire,17 RuO₂/SrTiO₃,18 RuO₂/3D silicon photonic crystals,19 Ru@FL-LDHs (Ru/Mg-Al LDHs matrix),20 Ru/Al₂O₃,21 Ru/TiO₂,22 Pd@Nb₂O₅,23 CoFeAl-LDH nanosheet,13 Ni/ SiO₂-Al₂O₃,14 Cu₂O/Zn-MOF,24 Cu₃O/graphene,25 etc. have been used for photothermal CO₂ hydrogenation, and their catalytic performances are given in Table S1.† Significant advances have been made in recent years; however, most of the above catalysts suffer from drawbacks such as using expensive noble metals, working at high temperature, and exhibiting low activity. Besides, the obtained products mainly focused on C1 chemicals such as CO and CH₄. The synthesis of C₂+ products via a photothermal route is a more difficult but significant issue, which is rarely involved in the above developed catalytic systems.26 Considering the above aspects, to develop a cost-efficient, highly active and reusable catalyst for photothermal
CO₂ reduction into CH₄ and C₂+ hydrocarbons is particularly important. On the other hand, it has been well known that pure metals used in CO₂ hydrogenation always showed a low activity due to their low CO₂ binding capacity. Oxides could not only anchor metal atoms but also provide vacancies for the reaction, and in this way, they could participate in the reaction and enhance the catalytic performances. Trimetallic oxide catalysts show unique electronic and structural features different from their individual components. The complex composition of trimetallic catalysts could form multiple interfaces conducive to catalysis. In this aspect, trimetallic catalysts have been widely used in many reactions such as CO and CO₂ hydrogenation, ethanol steam reforming, ammonia oxidation, etc. with good performances. In particular, complicated metal–trimetallic oxides can provide many vacancies and active sites for catalysis, and have great potential for application in CO₂ reduction. Herein, a series of Co–Cu–Mn trimetallic oxides were prepared and used for photothermal CO₂ reduction, and the screened CoₓCuₓMnₓOₓ(200) catalyst, with a Co/Cu/Mn ratio of 7/1/1 and reduced at 200 °C in H₂, showed a high CH₄ formation activity (14.5 mmol g⁻¹ h⁻¹) even at low CO₂ (10%) and H₂ (30%) concentrations, which lies in the highest activity range in Table S1.† Interestingly, after increasing the CO₂ and H₂ concentrations to 25% and 75%, 15.9 and 7.5 mmol g⁻¹ h⁻¹ of CH₄ and C₂+ hydrocarbons were produced with selectivities of 65.4% and 30.8%, respectively. Moreover, the catalyst has a good temporal stability under continuous-flow conditions. The metal/oxides were confirmed to be very important for the catalytic performance, in which the Co⁰ species were the active sites, which together with the oxides acted as supports and semiconductors for providing a local heating environment via the photothermal effects. The Cu species could promote the reduction of Co and Mn oxides, and also promote the coupling of C–C bonds to produce C₂, products especially at a high CO₂ concentration. The Mn species could on one hand enhance the adsorption of CO₂ and H₂, and also provide the photothermal effect to produce a local heating effect to promote the reaction. The present strategy for constructing trimetallic catalysts for the photothermal reaction does not only provide a simple but efficient way to convert CO₂ into valuable hydrocarbons, but also offer a general method for reducing the reaction temperature of conventional high-temperature reactions.

Results and discussion

Preparation and characterization of catalysts

Trimetallic Co–Cu–Mn catalysts were prepared by a simple coprecipitation method, and the details are given in the ESI.† The prepared catalysts were denoted as CoₓCuₓMnₓOₓ(T), in which a/b/c and T represent the Co/Cu/Mn molar ratio and the reduction temperature, respectively. CoₓCuₓMnₓOₓ means unreduced catalysts. The morphologies of the catalysts were examined by TEM tests. Initially, the Co₇Cu₁Mn₁Oₓ catalyst mainly showed the morphology of nanoparticles with the size range of 6–15 nm (Fig. 1a and b). The lattice spacings of 0.186, 0.246, 0.219, 0.152, 0.197, and 0.135 nm in the HRTEM image...
could be successively indexed to the (311) and (311) planes of CoO, the (411) and (611) planes of MnO, and the (−112) and (−221) planes of CuO (Fig. 1c).

After reduction, the CoCuMnOx(200) catalyst maintained the nanoparticle morphology in the size range of 6–12 nm (Fig. 1d and e). CoO, MnO, and MnO2 were also detected, as shown in Fig. 1f. The HAADF-STEM and EDX elemental mappings of the CoCuMnOx(200) catalyst showed that Co, Cu, Mn, and O elements are homogeneously dispersed over the whole grain (Fig. 1g to k).

For comparison, the TEM images of catalysts with different Co/Cu/Mn ratios were also studied, and the results are shown in Fig. S1.† With the increasing Co content, the morphology changed from nanosheets to nanoparticles. For example, the CoCuMnOx(200) and CoCuMnOx(200) catalysts were nanosheets, while the CoCuMnOx(200) and CoCuMnOx(200) catalysts were mainly nanoparticles, which were beneficial for exposing more active sites.

The molar ratios of the bimetallic and trimetallic catalysts were tested by ICP-OES, and the results are given in Table S2.† The molar ratios of metals were consistent with the feed amounts.

XRD examinations showed that all the tested unreduced catalysts such as CoCuMnOx, CoMnOx, and CoCuO showed the XRD patterns of CoO (JCPDS file no. 73-1701). No obvious CuO or MnO peaks were found, implying that they were easily incorporated into the CoO crystals and were well dispersed (Fig. 2a). The CoCuMnOx(200) catalyst showed the same peaks as that of the CoCuMnOx catalyst, but with lower intensities, indicating the crystallinity decrease of the CoO component after reduction.

The XRD pattern of the CoCuMnOx(200) catalyst was almost the same as that of the unreduced catalyst, and it is mainly attributed to the CoO phase (Fig. 2b). However, two small peaks at 43.0° and 61.8° were carefully detected, which could be assigned to the CoO(200) and (220) planes (JCPDS file no. 70-2855). No Co3+, Cu3+, Mn3+, or Mn2O3 species were detected due to their low amounts. The CoMnOx(200) catalyst showed three peaks at 36.8°, 42.7°, and 61.8°, which could be assigned to the CoO (111), (200), and (220) planes (JCPDS file no. 78-0431), respectively. Two phases of CoO fcc and hcp were identified in the CoCuMnOx(200) catalyst, and the formation of two phases is commonly observed in the synthesis of cobalt nanostructures.† The peaks at 41.5°, 44.4°, 47.4°, and 75.9° could be indexed to CoO hcp (100), (002), (101), and (110) facets (JCPDS file no. 01-071-4239), respectively. The peaks at 44.4° and 51.5° could be contributed to the CoO fcc (111) and (200) planes (JCPDS file no. 01-071-4651).† The results indicated that CoO could be reduced to Co0 by the promotion of Cu species, and this will be further verified by the following H2-TPR tests. No MnO or CuO species were detected due to their good dispersion. It is worth noting that although Cu could promote the reduction of CoO species to Co0, no Co0 was detected in the bulk phase of the CoCuMnOx(200) catalyst.

The XRD patterns of the catalysts with different Co/Cu/Mn ratios also showed the main peaks of CoO and CoO (Fig. S2†); however, the CoCuMnOx(200) and CoCuMnOx(200) catalysts showed a wide peak at around 20°, indicating that they possess amorphous structures.

The results of the N2 adsorption/desorption tests are shown in Fig. S3 and Table S3.† The reduction temperature affects the catalyst structure remarkably, and the BET surface areas of the tested trimetallic catalysts declined gradually with the increase of the reduction temperature, which is probably because the high reduction temperature leads to catalyst sintering. The bimetallic Co–Cu catalyst shows the lowest BET surface area, while the Co–Mn catalyst presents the highest. Indeed, the surface areas of the CoO and CoCuMnOx(100, 200, and 300) catalysts did not differ greatly from one another, thus their diverse performances did not stem from the difference in their surface areas.

To distinguish the valences of the Co, Cu, and Mn species on the surface of the catalysts, the XPS of Co 2p, Cu 2p, Mn 2p and 3s, and AES of Cu LMs, Mn Kα were analyzed (Fig. 3). In the XPS of the CoCuMnOx catalyst, the Co element was composed of Co3+ and Co2+ species, which was confirmed by the peaks at 779.5 eV and 780.9 eV, which can be attributed to the Co3+ and Co2+ 2p3/2 peaks (Fig. 3a). The Cu 2p XPS showed two peaks at 934.5 eV and 954.2 eV, which could be assigned to Cu2+ 2p3/2 and 2p1/2, respectively (Fig. 3b and c). The XPS peaks of Co species in the CoCuMnOx(200) catalyst were the same as those in the unreduced catalyst; however, the Cu species was found to be Cu+. The Mn XPS indicated that the catalyst contains Mn3+ oxidation state (Fig. 3d and e). The XPS results indicated that the surface of the CoCuMnOx(200) catalyst mainly contained Co2+, Co3+, Cu+, a small amount of Cu2+, and Mn3+, respectively, but no Co0 or Cu0 species were detected for air oxidation.

For comparison, the XPS spectra of the catalysts reduced at different temperatures were also tested. The Co element in these catalysts was mainly composed of Co2+ and Co3+ species.
The Co\textsuperscript{2+} component increased with the increase of the reduction temperature. No obvious Co\textsuperscript{0} species was found until the reduction temperature reached 500 °C. As for the Cu XPS, the area of the satellite peak around 942 eV decreased with increasing reduction temperature, indicating that Cu\textsuperscript{0/+} 2p\textsubscript{3/2} increased. Besides, all the reduced catalysts possess the dominant Cu\textsuperscript{+} component. The Mn 3s XPS showed that the Co\textsubscript{7}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x}(500) catalyst is composed of Mn\textsuperscript{2+} species, i.e., MnO, while the others showed Mn\textsuperscript{3+} species.

The XPS spectra of the Co–Cu–Mn trimetallic catalysts with different molar ratios were also investigated, and the results are given in Fig. S4.\textsuperscript{†} All the elements had the same valences but contained different contents, such as Co\textsuperscript{2+}, Co\textsuperscript{3+}, Mn\textsuperscript{3+}, Cu\textsuperscript{+}, and a small amount of Cu\textsuperscript{2+}, which are similar to the Co\textsubscript{7}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x}(200) catalyst, indicating that they are probably active in the reaction. This was confirmed by the following catalytic activity studies (Table 1).

The Raman spectrum of the Co\textsubscript{7}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x}(200) catalyst showed a peak at 680 cm\textsuperscript{-1}, which could be assigned to the A\textsubscript{1g} mode of CoO\textsubscript{4} (Fig. S5).\textsuperscript{39} This profile is similar to those of the Co\textsubscript{2}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{4} and Co\textsubscript{2}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{4} (100, 300, and 400) catalysts, indicating that CoO\textsubscript{4} exists in these catalysts. Interestingly, the Co\textsubscript{2}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{4}(500) catalyst showed a wide peak at 540 cm\textsuperscript{-1}, assigned to the one-phonon LO mode of CoO\textsubscript{4},\textsuperscript{40} which was formed by the partial oxidation of Co\textsuperscript{0} by local laser heating.\textsuperscript{45} No obvious Cu or Mn species were detected, indicating that they have permeated into the CoO\textsubscript{4} crystal lattice.

The above characterization studies showed that the Co\textsubscript{2}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{4}(200) catalyst mainly possesses CoO\textsubscript{4} in the bulk phase, while the surface of the catalyst was composed of Co\textsuperscript{3+}, Co\textsuperscript{2+}, Mn\textsuperscript{3+}, Cu\textsuperscript{+}, and a small amount of Cu\textsuperscript{2+} species, which are in the forms of Co\textsubscript{2}O\textsubscript{4}, Mn\textsubscript{3}O\textsubscript{4}, Cu\textsubscript{2}O, and CuO. The catalysts reduced at 100, 200, and 300 °C, and the catalysts with different Co/Cu/Mn molar ratios contained similar components but of different concentrations, indicating that they possess similar structures and properties.

**Catalytic performances**

Photothermal CO\textsubscript{2} hydrogenation was carried out in a tubular reactor (Fig. S6),\textsuperscript{†} and the procedure details are given in the ESI.\textsuperscript{†} The products of photothermal CO\textsubscript{2} reduction include CO, CH\textsubscript{4}, and C\textsubscript{2+} hydrocarbons (mainly C\textsubscript{2} to C\textsubscript{6}) in the present work. Initially, the catalytic performances over different catalysts are given in Table 1. No products were detected in the absence of a catalyst, and the sole metal catalysts or the bimetallic catalysts showed a very low activity in the reaction. The Co\textsubscript{7}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x}(200) catalyst showed the best catalytic performances with a CH\textsubscript{4} formation activity of 14.5 mmol g\textsubscript{cat}\textsuperscript{-1} h\textsuperscript{-1} and a selectivity of 85.3% (entries 1–7 vs. 8). These activities were achieved under the irradiation of a 300 W Xe lamp (234 mW cm\textsuperscript{-2}, 300–1100 nm) and heat generated by an additional electric heater (Fig. S6).\textsuperscript{†} Unlike the CuO\textsubscript{x}(200) and MnO\textsubscript{x}(200) catalysts, the CoO\textsubscript{x}(200) catalyst could show a very low CH\textsubscript{4} formation activity (only 3.1 mmol g\textsubscript{cat}\textsuperscript{-1} h\textsuperscript{-1}) (entries 2–4), demonstrating that the Co species is probably the active species, while the CuO\textsubscript{x} and MnO\textsubscript{x} species are the promoters for the Co species. Indeed, the Cu and Mn species are always used as promoters or active sites for the Co-catalyzed CO\textsubscript{2} reduction.\textsuperscript{41} The activities of the bimetallic Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x} or Co\textsubscript{2}Mn\textsubscript{1}O\textsubscript{x} catalysts confirmed that Cu or Mn are not the active sites (entries 5–7). Interestingly, the Co-Cu\textsubscript{2}O\textsubscript{4}(200) catalyst showed higher activities for the formation of CH\textsubscript{4} and C\textsubscript{2+}.

Fig. 3 XPS spectra of the unreduced Co\textsubscript{7}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x} catalyst and the Co\textsubscript{7}Cu\textsubscript{1}Mn\textsubscript{1}O\textsubscript{x}(100, 200, 300, and 500) catalysts. (a) Co 2p XPS, (b) Cu 2p XPS, (c) Cu L\textsubscript{3}M\textsubscript{4,5}M\textsubscript{4,5} AES spectra, (d) Mn 2p XPS, and (e) Mn 3s XPS, respectively.

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![Figure 3](image-url)
hydrocarbons than those of the pure CoO2(200) and CuO2(200) catalysts, indicating that the Cu species could not only enhance the activity of the Co-based catalyst in CO2 reduction to CH4, but also promote the formation of C2+ products. For comparison, other trimetallic catalysts with different ratios were also assessed; however, all of them possessed lower activities than that of the Co7Cu1Mn1O2(200) catalyst (entries 8 vs. 9–13).

The activities achieved in CO2/H2/N2 and CO2/H2/Ar were very similar (entries 8 vs. 14), indicating that the catalytic performances were not affected by the type of diluent gas. In addition, the C2+ activities of the Co0.2Cu0.8Mn1.0O2(200), Co0.2Cu0.8Mn1.0O2(200), Co0.2Cu1.0Mn0.0O2(200), and Co0.2Cu1.0Mn0.0O2(200) catalysts were only 0.9, 1.4, 0.8, and 0.5 mmol gcat⁻¹ h⁻¹, respectively, and inconsistent with the change rule of the catalyst composition. Because the formation of C2+ products is closely related to the Cu species, the uneven distribution of Cu species on the catalyst surface is probably the reason for the discrepancy between the C2+ activity and the Co/Cu/Mn molar ratio of the catalyst.

The Co species are the active sites; however, the promoters of Cu and Mn species are also very important for the reaction. To confirm this, the catalysts with the same component of Co element were investigated, and the results indicated that the Cu and Mn concentrations affected the catalytic performances remarkably (entries 15–18). The Co0.2Cu0.25Mn1.25O2(200) and Co0.2Cu1.25Mn0.75O2(200) catalysts showed high CH4 formation activities, while the Co0.2Cu0.25Mn1.75O2(200) and Co0.2Cu1.75Mn0.25O2(200) catalysts offered relatively low values, indicating that the low concentrations of Cu and Mn are not beneficial for catalysis.

The catalytic performance studies showed that the coexistence of Co, Cu, and Mn in the Co/Cu/Mn trimetal is very important for the reaction, and their molar ratios and the synergistic effects of these multicomponents are closely related to the catalytic performances, and a similar phenomenon could be found in the previous report.42

### Effect of the reduction temperature of the CoCuxMn3Ox catalysts

The above characterization results indicated that the reduction temperature affected the metal valence compositions of the catalysts, and it was proved to be able to affect their catalytic performances.43,44 To verify this, the catalysts reduced at different temperatures were studied, and the results are shown in Fig. 4. Indeed, the CoCuxMn3Ox and CoCuxM3O2x(100, 200, and 300) catalysts showed good catalytic performances during the reaction, indicating that they have similar catalysts.
compositions and properties. However, the Co7Cu1Mn1Ox(400 and 500) catalysts showed decreased catalytic activities, which has a great relationship with the high concentration of Co0 in the bulk phase but less Co3O4 and Mn2O3.

The catalysts after being used for 9 h and 21 h were characterized by XPS and XRD in Fig. 5, and the results showed that Co0 and Cu0 species could be found during the reaction. Co0 was widely recognized as the active site for CO2 hydrogenation. The XRD patterns of the catalysts used for 9 h showed similar peaks to those of Co3O4 and CoO, revealing that Co3O4 still exists in the bulk phase. Compared with the Co7Cu1Mn1Ox(200) catalyst in Fig. 2, the peak at 36.6° decreased remarkably. The catalyst used for 21 h showed similar results but with more Co0 and Cu0 species. The XPS and XRD characterization studies showed that the catalyst compositions were stable during the reaction.

The highest CH4 activity was achieved over the Co7Cu1Mn1Ox(200) catalyst. Besides, the Co7Cu1Mn1Ox(100 and 300) catalysts also showed acceptable CH4 formation activities. Generally, in thermal catalysis, Co0 species are recognized as the active components. However, the Co7Cu1Mn1Ox(500) catalyst showed the lowest activity in the reaction, which mainly contained the Co0 species but no Co3O4 or Mn2O3. It should be noted that in the present photothermal catalysis, the oxides including Co3O4, Mn2O3, etc. could promote the reaction via the interaction effect between the Co species and other oxides.

The H2-TPR tests were carried out to study the interaction effect of different metal oxides, and the results are shown in Fig. 6. Pure Co3O4 showed two peaks at 273 °C and 336 °C, which could be assigned to the reduction peaks of Co3O4 to CoO and CoO to Co0, respectively.46 Mn2O3 showed two peaks located at 292 °C and 388 °C, which could be assigned to the reduction of Mn2O3 to Mn3O4 and Mn3O4 to MnO, respectively.47,48 The Co7Cu1Ox catalyst had two peaks at lower temperatures of 156 °C and 210 °C indicating that Cu could obviously promote the reduction of Co3O4 and CoO, which is consistent with previous reports.45 However, the Co7Mn1Ox catalyst showed two main peaks at 280 °C and 428 °C, which could be assigned to the reduction of (Co,Mn)3O4 to (Co,Mn)O and (Co,Mn)O to Co0 and MnO,5,49,50 respectively, revealing that Mn species impedes the reduction of CoOx. The Co7Cu1Mn1Ox catalyst had three peaks at 128, 183, and 284 °C, which could be attributed to the reduction peaks of CuO to Cu0, Co3O4 to CoO, and CoO to Co0 and MnO, respectively.50 The H2-TPR results indicated that the Co, Cu, and Mn species show strong interaction effects, and the previous report showed that the presence of Cu could alter the electronic interactions with Co and Ni in the Co–Cu–Ni trimetallic catalysts.30

The surface adsorption capacities could significantly affect the catalytic activity. Thus, we conducted the CO2-TPD and H2-TPD tests for the diverse trimetallic catalysts and the results indicated that with the increase of the reduction temperature, the adsorption amounts of CO2 and H2 on the catalysts decreased. The Co7Cu1Mn1Ox(100 and 200) catalysts contained higher CO2 and H2 adsorption amounts than the others,
which is beneficial for them to catalyze the CO₂ photothermal reduction.

To investigate the effects of the Mn and Cu species on the Co-based catalysts, the CO₂-TPD tests for Co–Mn and Co–Cu bimetallic catalysts were performed (Fig. S7†). The Mn species was found to be able to improve CO₂ and H₂ adsorption for the Co-based catalysts, while the Cu species reduced their adsorption since the Co–Cu bimetallic catalyst could be deeply reduced at 200 °C. Even so, the Co₇Cu₁Oₓ(200) catalyst exhibited a strong CO₂ adsorption peak at around 350 °C, showing that the Cu species could promote strong adsorption of CO₂ on the catalyst.

We investigated the catalytic performances of the Co–Cu and Co–Mn bimetallic catalysts reduced at different temperatures, and the results are given in Table S4.† With the increase of the reduction temperature, the activities of the Co–Cu and Co–Mn catalysts increased initially but were then reduced at high reduction temperatures of 400 and/or 500 °C. The highest activities were lower than that of the Co₇Cu₁Mn₁Oₓ(200) catalyst (Table 1, entry 8), indicating that the Cu and Mn species were both needed for the photothermal CO₂ reduction.

**Effect of the reaction conditions**

The above investigation showed the composition and pretreatment of the catalyst affected the activity remarkably. In addition, the reaction conditions also had a great influence on the catalytic performance. The effects of the reaction conditions on the catalytic performance were also investigated, and the results are shown in Fig. 7. Almost no product was detected in the absence of CO₂ or H₂, indicating that the products are generated from CO₂. This was further confirmed by a control experiment of adding ¹³CO₂ in the reaction, in which ¹³CH₄ was detected with a mass spectrometer (Fig. S8†).

The best performances for CH₄ formation were achieved in CO₂/H₂ with a molar ratio of 1/3 (Fig. 6a). Surprisingly, changing the CO₂ and H₂ concentrations slightly affected the activity of CH₄ formation (around 15 mmol gcat⁻¹ h⁻¹), but had an influence on C₂+ hydrocarbon formation. In the absence of a dilute gas, the highest C₂+ activity reached 7.5 mmol gcat⁻¹ h⁻¹ with a 30.8% selectivity. Generally, C₂+ hydrocarbon formation involves CO₂ reduction and C–C coupling steps, and is much more difficult to achieve under photothermal conditions. Thus, the present Co₇Cu₁Mn₁Oₓ(200) catalyst could catalyze the C₂+ hydrocarbon formation, but only the CO₂ and H₂ concentrations need to be enhanced. The C₂+ formation is strongly linked to the amount of surface adsorbed carbon species such as CO, –CH₂–, etc. A high amount of carbon species can be coupled to generate C₂+ hydrocarbons over a suitable catalyst, thus an increase of the CO₂ concentration could enhance the C₂+ selectivity. However, generally, to obtain C₂+ products, high pressure and/or CO₂/H₂ molar ratio are always needed.42 The present Co₇Cu₁Mn₁Oₓ(200) catalyst has a good CO₂ adsorption capacity, which was confirmed by the CO₂-TPD tests. The high C₂+ selectivity indicates that the CO₂/Cu₇Mn₁O_x(200) catalyst possesses good CO₂ reduction and C–C coupling abilities even under ambient pressure and a low CO₂/H₂ molar ratio (1/3). These multi-functional applications of synthesis of CH₄ and C₂+ hydrocarbons stemmed from the multiple interfaces created by the complicated compositions of Co, CoO, Mn₂O₃, CuO and Cu₂O species, etc.

The reaction temperature has a significant influence on the catalytic performance, which was monitored using a thermocouple positioned above the catalyst layer (Fig. S6†). As shown in Fig. 7c, high temperature promoted the reaction remarkably. With the increase of the reaction temperature, the activity of CH₄ formation increased. The C₂+ selectivity increased initially and then slightly decreased at a temperature higher than 200 °C. This is probably because at high temperature, the catalyst had an induction period. During this period, the catalyst was sintered and reduced, which affected the product distribution especially with the CH₄ selectivity increased and C₂+ products decreased.

The evolution of activity on reaction time was also studied and the results are shown in Fig. 7d. In the first 30 min of reaction, CH₄ was formed with a very low activity, indicating that the catalyst had an induction period. During this period, the surface Co species, especially the surface CoO was reduced to the active Co²⁺ species. With prolonged time, the CH₄ formation activity increased remarkably. During our experiments, we found that the pressure of the reaction gas decreased remarkably after 3 h, but further prolonging the time did not show any remarkable decrease, indicating that the reaction has reached equilibrium. Thus, we chose 3 h as the optimum reaction time. In addition, control experiments were conducted to investigate the effect of optical and thermal conditions. The reaction could hardly occur under light irradiation but without extra heating, and the sole thermal conditions offered a very low CH₄ formation activity (only one-sixth that of
the photothermal reaction). However, their combination is beneficial for the CO_2 reduction, indicating that the excellent catalytic performances stemmed from their synergistic promotion effect.

As mentioned above, the present catalytic system could achieve the photothermal CO_2 reduction to C_2 hydrocarbons, but only needs to enhance the concentrations of CO_2 and H_2 to 25% and 75%. Fig. S9† shows the catalytic performances of the reaction. Under the conditions of 200 °C and 3 h, the highest C_2 selectivity and activity reached 30.8% and 7.5 mmol g_{cat}^{-1} h^{-1} obtained after 3 h of reaction.

The results indicated that the multifunctional applications of the trimetallic Co–Cu–Mn catalysts derive from not only the multicomponents of the catalysts, but also the interaction effects between the components.

**Photoelectric properties of the Co–Cu–Mn catalysts**

The study of the catalytic performances confirmed that Co_3O_4 and Mn_2O_3 oxides are very important for the reaction, and this differs from the fully reduced Co-based catalysts in H_2 such as Co/CoO_{29} and Co/MnO_x catalysts in CO_2 thermal hydrogenation. This diversification is probably because photothermal catalysis needs more light-sensitive components such as Co_3O_4 and Mn_2O_3 due to their semiconductor nature, while the thermal catalyst favours low-valent metals such as Co_0.

To further investigate the light-sensitive properties of the Co–Cu–Mn trimetallic catalysts, the UV-vis DRS spectra were investigated, and the results are shown in Fig. S10a.† The unreduced catalysts including Co-Mn_2O_3, Co-Cu_2O_x, and Co-Cu-Mn_2O_3 showed higher absorption than the corresponding reduced catalysts throughout the UV-vis region. This was because the former mainly contain Co_3O_4, which has strong absorption of d–d transitions. Upon reduction, the Co_3O_4 amount decreased, leading to the decrease of absorption accordingly. As for the Co-Cu-Mn_2O_3(200, 300, 400, and 500) catalysts, with the increase of the reduction temperature, the absorption strength decreased gradually for the deep reduction of Co_3O_4 to CoO and Co^0 species. Both the Co-Cu_2O_x(200) and Co-Cu-Mn_2O_3(500) catalysts showed a lower absorbance than the other samples because they did not contain the Co_3O_4 species, which was confirmed by the XRD tests.

The band gap energies (E_g) for the tested catalysts were calculated using the Kubelka–Munk (K–M) model. The Co-Cu-Mn_2O_3(200) catalyst had a lower E_g value (1.86 eV) than those of the catalysts reduced at higher temperatures (Fig. S10b†). The Co-Cu-Mn_2O_3(500) catalyst, however, had the highest E_g (3.26 eV) value and showed a very low activity in the reaction. The low E_g value of the Co-Cu-Mn_2O_3(200) catalyst led to the feature that the electrons in the semiconductor could be easily excited under light irradiation. The lower band gap is conducive to visible light excitation, which was confirmed by the control experiments using different types of light including full irradiation (300–1100 nm), UV (300–420 nm), and visible light (420–800 nm). The catalytic performances obtained under visible light illumination are almost similar to those under UV-vis and full irradiation, as shown in Table S5.†

The photocurrent–time curves of the Co_Cu_Mn_2O_3(100, 200, 300, and 400) catalysts are shown in Fig. S10c.† All of these catalysts showed obvious photocurrent response behavior, indicating that they could be excited to generate electron–hole pairs.

Co_3O_4, Mn_2O_3, Cu_2O, and CoO are p-type semiconductors; however, after reduction, oxygen vacancies were created and regulated the p-type to a n-type semiconductor, which is confirmed by the Mott–Schottky curves shown in Fig. S10d.† All the catalysts showed dipolar (p-type and n-type) semiconductor properties, resulting from the multiple interfaces of the complex trimetallic catalyst.

The EIS Nyquist plots of the tested samples including Co_Cu_Mn_2O_3(100, 200, 300, and 500) showed similar arc radii, revealing that they have similar impedance values and electron transfer rates (Fig. S10e).†

The photothermal effect was investigated via testing the stable temperature of the samples under vacuum (Fig. S11†). Under similar conditions, the catalyst temperatures were monitored using an infrared camera. Among the tested catalysts, the MnO_2 catalyst showed the highest temperature (104.6 °C) under the full irradiation, while the CuO catalyst offered the lowest (88.3 °C). The Co-Cu_Mn_2O_3(200) catalyst showed a higher temperature (102.2 °C) than that of the Co-Cu_Mn_2O_3(500) (95.8 °C), indicating that the Co-Cu_Mn_2O_3(200) catalyst could provide a higher local heating environment. The results revealed that the presence of Mn could provide a higher local temperature for the catalyst.

In summary, the results obtained due to the photoelectric characteristics indicated that the Co-Cu_Mn_2O_3(200) catalyst is a p-type and n-type semiconductor, showing a low band gap. All of these properties are closely related to the excellent catalytic activity in photothermal CO_2 reduction. However, it should be noted that the unique catalytic performances of the Co-Cu_Mn_2O_3(200) catalyst were derived from the synergistic effect of the adsorption effect, light-response properties, and photothermal effect.

**Light effect and the proposed mechanism for photothermal CO_2 reduction**

In the present catalytic system, the synergistic effect between light and heat had a great influence on the catalytic activity. Under light irradiation, the electrons of the Co-Cu_Mn_2O_3(200) catalyst could be excited to generate the electron–hole pairs. Generally, in photocatalysis, the excited electrons and holes can participate in reactions. However, under photothermal conditions, the E_g value of the semiconductor is normally found to be lower, and the excited electrons recombine more easily with the photogenerated holes than under the traditional photocatalysis without heating.

To investigate the effects of the generated electron–hole pairs, we carried out two control experiments by adding sacrificial electron donors including p-xylene (~2.18 V vs. Ag/AgCl_{sat}) and anisole (1.92 V vs. Ag/AgCl_{sat}) in the absence of H_2. As shown in Fig. S12,† both the reactions gave a very low CH_4
yield despite the fact that p-xylene and anisole could quench the photogenerated holes more efficiently than H₂. These results indicated that photocatalysis plays a minor role in the present photothermal CO₂ reduction. In this aspect, similar to the Ru-, 18 Rh-, 58 Au-, 58 and Ag-based 59 catalysts, the charge separation and recombination under light irradiation resulted in a local thermal effect at the surface of the catalyst, offering more energy to promote CO₂ reduction.

The mechanism for the photothermal CO₂ reduction to hydrocarbons is still unclear, and two reaction routes are generally accepted. 60 The first involves a CO intermediate, in which CO₂ is initially converted to CO, and CO could be further hydrogenated to hydrocarbons. 61 The other route is the direct CO₂ conversion. In this route, CO₂ is hydrogenated to carbonate, formate, and methoxy intermediates, and then hydrogenated to hydrocarbons. In this work, the CO formation activity was very low. To study the reaction pathway, we conducted the control experiments of replacing CO₂ by CO under the conditions given in Table 1, entry 8. Interestingly, the activities of CH₄ and C₂⁺ were 4.9 and 7.3 mmol g⁻¹ h⁻¹, and their selectivities were 40.2% and 59.8%, respectively (Table 1, entry 8, Fig. S13, S14 vs. Table S6, entry 2, Fig. S15†). To exclude that the high CO concentration may occupy the active sites on the catalyst surface, a mixed gas of 2%CO/8%CO₂ (total 10%) was also reacted (Table S6, entry 3 and Fig. S16†). However, the activities of CH₄ and C₂⁺ were 7.2 and 2.8 mmol g⁻¹ h⁻¹ with the selectivities of 72.0% and 28.0%, respectively. The results implied that the presence of CO could not improve but reduce the CH₄ activity, meaning that in the photothermal CO₂ reduction, CH₄ is probably not formed via the CO intermediate route. However, CO was found to be beneficial for the C₂⁺ synthesis, revealing that it is probably an intermediate for the formation of the C₂⁺ hydrocarbon product.

The characterization results confirmed that the trimetallic Co₇Cu₁Mn₁O₅(200) catalyst contains CoO, Co₃O₄, Cu₂O, Mn₂O₃, etc. During the reaction, the CoO species could be reduced to Co²⁺, which was recognized as the active site. The CoO and Cu₂O species could be reduced by H₂ to Co⁰ and Cu⁰ during the reaction, which was confirmed by the XPS results (Fig. 5). In the present photothermal CO₂ hydrogenation, metal oxides such as Co₃O₄ and Mn₂O₃ were very important for the reaction because they are semiconductors, which could generate electrons and holes under light illumination. A quick recombination of electrons and holes provides a local heating environment for the catalysts, and further promotes the reaction. 5 In addition, Mn₂O₃ could offer Lewis acidic sites to enhance CO₂ and H₂ adsorption capacity of the catalyst.

Cu species could promote the partial reduction of cobalt oxides to CoO and Co⁰ and maintain the presence of CoO and Co₃O₄ prior to and during the reaction, and reduce the temperature of catalyst reduction. Besides, the CO₂-TPD tests and the control experiments confirmed that the Cu species or the Cu–Co interfaces could provide strong CO₂ adsorption sites to convert CO₂ to CO, and further form C₂⁺ hydrocarbons.

It should be mentioned that, although CH₄ was probably not formed via the CO route, CO could be produced with an activity of 1.1 mmol g⁻¹ h⁻¹ (entry 8, Table 1). Importantly, the higher concentration of CO could make it easily coupled into the C₂⁺ products over the same Co₇Cu₁Mn₁O₅(200) catalyst. Thus, C₂⁺ hydrocarbons could be formed by increasing the CO₂ concentration.

In addition, the present protocol could be extended to photothermal CO hydrogenation. We conducted the reaction over the Co₇Cu₁Mn₁O₅(200) catalyst under the similar conditions given in Table 1, entry 8, and the results are shown in Fig. S17.† The photothermal activity is about 7 times higher than that of the thermal activity, and the selectivity of the C₂⁺ products is higher than that of CH₄, which is similar to CO₂ hydrogenation. The results indicated that the present trimetallic Co–Cu–Mn catalyst has multifunctional applications as shown in photothermal CO hydrogenation.

The reusability and stability studies of the Co₇Cu₁Mn₁O₅(200) catalyst

The reusability of the Co₇Cu₁Mn₁O₅(200) catalyst was investigated, and the results are shown in Fig. 8a. For comparison, the CO₂ reduction reactions were alternatively conducted under photothermal and thermal conditions. The results revealed that the present catalyst possesses good reusability in the reaction, and after using for 18 times, it still offered an activity of higher than 10 mmolCH₄ g⁻¹ h⁻¹. Besides, the results also confirmed that the activities under photothermal conditions were much higher than those under thermal conditions. The used catalysts were characterized by XRD and XPS, and the results are shown in Fig. 5, respectively.

The activity loss is probably caused by catalyst agglomeration, which was confirmed by the TEM images of the catalyst after being used 18 times (Fig. S18).† The size of the used catalyst increased in the range 40–90 nm.

Fig. 8b shows the temporal stability of the Co₇Cu₁Mn₁O₅(200) catalyst in a continuous flow operation (details are given in the ESI†). The catalyst offered approximately 20 mmol g⁻¹ h⁻¹ of CH₄ after a short induction period of 3 h, and maintained higher than 18 mmol g⁻¹ h⁻¹ activity.
even after 25 hours of reaction. The result showed that the present Co-Cu$_x$Mn$_{1-x}$O$_2$(200) catalyst has a long-term temporal stability, confirming that it has a potential industrial application.

Conclusions

In summary, a trimetallic Co–Cu–Mn catalytic system for photothermal CO$_2$ reduction was constructed by a simple co-precipitation method. The metal compositions and reduction temperatures of the catalysts affected their activities remarkably, and the screened Co$_7$Cu$_1$Mn$_1$O$_2$(200) catalyst could offer 14.5 mmol g$_{cat}^{-1}$ h$^{-1}$ of CH$_4$ at low CO$_2$ (10%) and H$_2$ (30%) concentrations. Importantly, the catalyst system could produce C$_2$, hydrocarbons with an activity of 7.5 mmol g$_{cat}^{-1}$ h$^{-1}$ and a selectivity of 30.8% by only improving the CO$_2$ and H$_2$ concentrations to 25% and 75%, respectively. High activities stemmed from the multicomponents of the catalyst, especially the metallic Co$_0$, Cu$_0$, and Co$_2$O$_3$, Cu$_2$O, and Mn$_3$O$_4$ semiconductors. The Cu species could tune the reduction characteristics, and further tune the balance of Co$_2^+$ and Co$_3^+$. The Mn species could enhance the CO$_2$ and H$_2$ adsorption, and also offer a high local temperature for the nanocatalyst. The present strategy of partial reduction of the Co–Cu–Mn trimetallic catalyst not only highlights the route for using solar energy to produce valuable chemicals under mild conditions especially low CO$_2$ concentration, but also constructs the highly efficient trimetallic catalysts with multifunctional applications in photothermal catalysis.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

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