PAPER

Capturing spatially resolved kinetic data and coking of Ga–Pt supported catalytically active liquid metal solutions during propane dehydrogenation *in situ*†

Moritz Wolf, [ID] [a] Narayanan Raman,[a] Nicola Taccardi,[a] Raimund Horn,[b] Marco Haumann [ID] [a] and Peter Wasserscheid*[ac]

Supported liquid phase catalysis has great potential to unify the advantages from both homogeneous and heterogeneous catalysis. Recently, we reported supported catalytically active liquid metal solutions (SCALMS) as a new class of liquid phase catalysts. SCALMS enable high temperature application due to the high thermal stability of liquid metals when compared to supported molten salts or ionic liquids. The highly dynamic liquid metal/gas interface of SCALMS allows for catalysis over single atoms of an active metal atom within a matrix of liquid gallium. In the present study, kinetic data is acquired along the catalyst bed in a compact profile reactor during propane dehydrogenation (PDH) over gallium–platinum SCALMS. The reactor design allows for the analysis of the temperature and gas phase composition along the catalyst bed with a high spatial resolution using a sampling capillary inside the reactor. The concentration profiles suggest enhanced deactivation of the catalyst at the end of the bed with a deactivation front moving from the end to the beginning of the catalyst bed over time on stream. Only minor amounts of side products, formed *via* cracking of propane, were identified, supporting previously reported high selectivity of SCALMS during alkane dehydrogenation. The acquired data is supported by *in situ* high-resolution thermogravimetry coupled with mass spectrometry to monitor the activity and coking behaviour of SCALMS during PDH. The results strongly suggest an enhanced formation of coke over Al_2O_3-supported SCALMS when compared to using SiO_2 as the support material.

[a]Friedrich-Alexander-Universität Erlangen-Nürnberg (FAU), Lehrstuhl für Chemische Reaktionstechnik (CRT), Egerlandstr. 3, 91058 Erlangen, Germany. E-mail: peter.wasserscheid@fau.de
[b]Technische Universität Hamburg (TUHH), Institut für Chemische Reaktionstechnik, V-2, Eißendorfer Str. 38, 21073 Hamburg, Germany
[c]Forschungszentrum Jülich, "Helmholtz-Institute Erlangen-Nürnberg for Renewable Energies" (IEK 11), Egerlandstr. 3, 91058 Erlangen, Germany

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0fd00010h
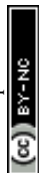
Introduction

Industrial application of Pt-based catalysts for technical propane dehydrogenation (PDH; eqn (1)) requires frequent regeneration of the catalysts due to strong deactivation. The deposition of carbonaceous species is generally accepted to be the culprit for the rapid decay of the catalyst activity.[1–4] Coke formation is mostly caused by overcracking of hydrocarbons.[5–8] Even though regeneration of the catalyst *via* oxidative treatments may be realised within short time frames,[3] the associated downtime of the reactor drastically reduces the process efficiency due to increased capital and operational expenditures. Hence, research on catalyst design for effective PDH focuses on high resistance against carbon deposition.[1,2,9–11]

$$C_3H_8 \leftrightarrow C_3H_6 + H_2 \tag{1}$$

We have recently proposed supported catalytically active liquid metal solutions (SCALMS) as a novel supported liquid phase catalysis concept.[12] Application of classical SLP catalysts with organic liquids, ionic liquids, or molten salts as the liquid phase on porous supports, is typically restricted to relatively low temperature applications ($\leq$300 °C) due to the limited thermal stability of the applied liquid phases.[13–16] Conversely, SCALMS employ liquid metals allowing for high temperature applications, because virtually no decomposition occurs for a solution of elementary metals.[10–12] While other concepts for catalysis over liquid metals require large volumes of liquid metal in a reactor,[17,18] SCALMS materials are composed of dispersed supported droplets of a liquid alloy consisting of a catalytically active metal and an excess of a low melting point metal, *e.g.* Ga.[10–12,19] The catalytic reaction in SCALMS occurs exclusively at the liquid metal/gas interface.[10,12,19–25] Contrary to conventional SLP catalysis, the reactants and products are insoluble in the liquid phase. In addition, the liquid metal/gas interface is highly dynamic on an atomic scale. The topmost layer of the Ga-rich alloy droplets is depleted of active metal atoms in the absence of substrates.[10,12,20,21] Nevertheless, the active metal atoms may diffuse to the surface of the droplets and interact with substrates *via* adsorption, which is a prerequisite for the catalytic activity of SCALMS. In the case of PDH over Ga–Rh SCALMS, *ab initio* molecular dynamics simulations suggest that the presence of a propane molecule at the liquid metal/gas interface may trigger the diffusion of Rh atoms to the surface of the liquid alloy droplet. Adsorption of the substrate then retains the Rh atom at the surface and is followed by C–H bond breakage resulting in two hydrogen atoms and propylene bound to a single Rh atom. Subsequently, diffusion of propylene to adjacent Ga atoms is suggested by the simulations. Lastly, propylene desorbs and the two hydrogen atoms bound to Rh combine and desorb as H_2, while the Rh immediately moves away from the surface of the droplet into the Ga matrix.[10] Hence, dehydrogenation of propane over SCALMS is suggested to require only a single active metal atom in a Ga matrix, which is in line with Biloen *et al.*[26] Consequently, side reactions, which require a second vicinal active site, may be suppressed during PDH over SCALMS.

Herein, we employ Ga–Pt SCALMS during PDH using Al_2O_3 and SiO_2 as carrier materials. We recently hypothesised that the coke formation during PDH over a related catalyst, namely $GaRh/Al_2O_3$ SCALMS, at increased reaction

temperatures of 550 °C is governed by the acidic sites of the support material.[11] Hence, the coking behaviour of both Ga–Pt SCALMS catalysts was compared by means of *in situ* high-resolution thermogravimetric analysis coupled with mass spectrometry (HRTGA-MS). Further, spatially resolved kinetic data was acquired along the catalyst bed in a compact profile reactor (CPR), which may provide the foundation for a comprehensive kinetic analysis of the activity and deactivation mechanisms of SCALMS during PDH.

Experimental section

Preparation of supported catalytically active liquid metal solutions (SCALMS)

At first, $(Et_3N)GaH_3$ was synthesised according to the published procedure,[27] while triethylammonium chloride (Sigma-Aldrich) was used instead of trimethylammonium chloride. The compound was not isolated and used as ethereal solution. The Ga content of this solution was determined as follows: 1 mL of ethereal solution was quenched in 10 mL of HCl ($\sim$2 M) and the resulting mixture shortly boiled to evaporate the diethyl ether. The homogeneous solution was diluted to 500 mL and the Ga content determined by inductively coupled plasma-atomic emission spectroscopy (ICP-AES). Given the low stability of gallane complexes, the $(Et_3N)GaH_3$ solution was analysed and used immediately after its preparation. The second part of the synthesis procedure includes the decoration of the support material with Ga according to a previously published procedure.[10,12] Either 20 g of SiO_2 (Sigma-Aldrich; particle size: 200–500 μm; BET surface area: 500 m^2 g^{-1}) or 100 g of Al_2O_3 (Sigma-Aldrich; grade: Brockmann I, activated, standard; particle size: 50–150 μm; BET surface area: 155 m^2 g^{-1}; pH: 7.0 ± 0.5) were employed as carrier materials and dried by heating under vacuum (1 mbar) at 350 °C overnight in a Schlenk flask. After cooling-down, the support was suspended in dry diethyl ether (50 mL for SiO_2 and 200 mL for alumina) under an inert argon atmosphere. The amount of an ethereal solution of $(Et_3N)GaH_3$ added to this suspension was that required in order to obtain the targeted loading of Ga (*ca.* 5–6 wt% of Ga) with respect to the support. After complete removal of diethyl ether under vacuum at *ca.* −30 °C, the flask was heated up to 300 °C (*ca.* 10 °C min^{-1}) until no gaseous products were observed indicating the termination of the gallane decomposition. The resulting grey solid was further heated to 300 °C under vacuum (1 mbar) overnight and stored under Ar after cooling-down to room temperature. In the last step of the synthesis of Ga–Pt SCALMS, 10 g of Ga decorated material were suspended in 50 mL of iso-propanol in a 100 mL round flask under stirring. The required volume of a stock solution of H_2PtCl_6 (nominal Pt concentration of 3.6 mg mL^{-1}) in distilled water was added to this solution to obtain the targeted atomic Ga/Pt ratio. The flask was connected to a rotary evaporator and the solvent evaporated at 50 °C and 25 mbar. The resulting solid was calcined at 500 °C for 120 min.

Metal content analysis

The Ga and Pt loading and the corresponding Ga/Pt ratios of the SCALMS were determined by ICP-AES using a Ciros CCD (Spectro Analytical Instruments GmbH). The solid samples were digested in 3 : 1 : 1 volumetric ratio of concentrated HCl : HNO_3 : HF using microwave heating up to 220 °C for 40 min. The

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instrument was calibrated for Pt (214.423 nm) and Ga (417.206 nm) with standard solutions of the particular elements prior to the analyses.

High-resolution thermogravimetric analysis coupled with mass spectrometry (HRTGA-MS)

PDH over SCALMS was also characterised by means of *in situ* HRTGA-MS using a XEMIS sorption analyser (Hiden Isochema).[28] The XEMIS sorption analyser has a superior resolution of ± 0.1 μg when compared to classical thermogravimetric analysers and can be operated under a vacuum or pressures up to 200 bar and temperatures up to 500 °C. A total of 200 mg of as prepared $GaPt/Al_2O_3$ or $GaPt/SiO_2$ SCALMS was loaded and dried under a flow of He (200 mL_N min^{-1}; 500 °C for 6 h with a heating ramp of 5 °C min^{-1}). Desorption of physisorbed and chemisorbed water during *in situ* measurements would deter the measurement of the net weight change making this first step inevitable. Subsequently, PDH was conducted at 450 °C or 500 °C for 6 h using a feed gas composition of 10% C_3H_8/He (overall flow rate: 200 mL_N min^{-1}). The last experimental step during analysis by HRTGA-MS was temperature programmed oxidation (TPO) in 21% O_2/He (100 mL_N min^{-1}), which was introduced immediately after catalytic application and cooling-down to 100 °C under He. TPO was conducted after an initial isotherm at 100 °C (0.5 h) during a ramp to 500 °C at 1 °C min^{-1} with a subsequent isotherm at 500 °C (12 h). A mass spectrometer (Hiden Analytical) continuously analysed the off-gas in the mass-to-charge (m/z) range of 1–50.

Propane dehydrogenation (PDH) in a compact profile reactor (CPR)

SCALMS were applied in PDH using a CPR (Reacnostics, Germany). The catalyst was placed in-between two quartz wool plugs inside a quartz tube reactor (length: 180 mm; outer diameter: 6 mm; inner diameter: 4 mm). A stainless steel capillary with four orifices (diameter: 100 μm) at a defined position is placed through the catalyst bed inside the quartz tube (Fig. 1). The reactor can be heated up to 550 °C and is mounted on a sledge allowing for movement of the reactor in the axial direction along the capillary. This relative movement enables spatially resolved sampling of the gas phase *via* the orifices, probing of the temperature inside the catalyst bed using a thermocouple inside the capillary (Fig. 1), as well as spatially resolved Raman spectroscopy *via* an external optical window. The latter was realised using a 532 nm solid state laser Raman set-up (Avantes: AvaRaman-532HERO-EVO, AvaSpec-HERO, AvaRaman-PRB-532) with a laser power of 15 mW, an exposure time of 30 s and averaging of 5 repetitions. A total of 333.3 mg of catalyst was loaded into the fixed-bed reactor. The resulting catalyst bed of 55 mm lies within the isothermal zone of the heating chamber ($\geq$60 mm). The catalyst was heated to the desired temperature (450, 500, or 550 °C) at 10 °C min^{-1} under a flow of He (Air Liquide, purity 4.6). After 35 min at the final temperature, the feed stream was changed to 12% C_3H_8 (Air Liquide, purity 3.5) in He for PDH. The total gas hourly space velocity (WHSV) of 4500 mL_N mL_{cat}^{-1} h^{-1} was constant throughout the experiments. The reaction gas mixture was actively sampled through the orifices of the capillary by the pump of an I-GRAPHX PR micro gas chromatograph (I-GraphX GmbH, Germany).
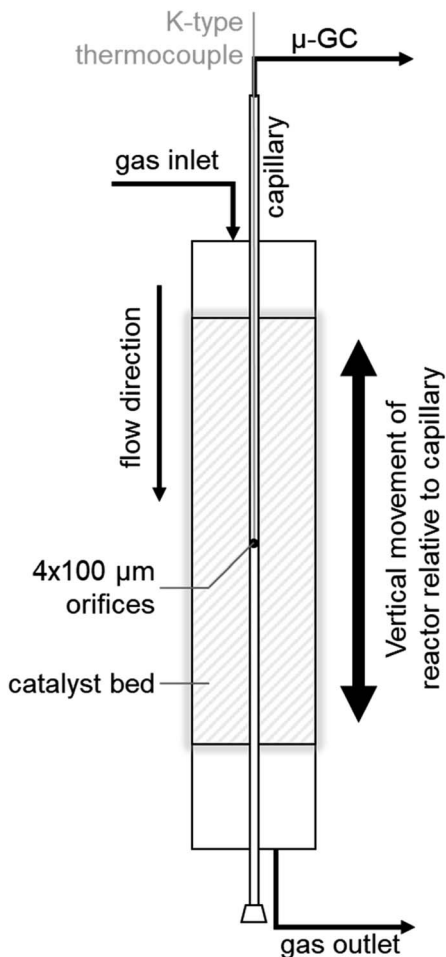
Fig. 1 Graphical depiction of the compact profile reactor (CPR) allowing for spatially resolved concentration profiles along the catalyst bed inside a quartz tube reactor.

Results and discussion

Elemental analyses of the as prepared SCALMS samples by means of ICP-AES resulted in a Ga loading of 5.80 and 4.14 wt% on SiO_2 and Al_2O_3, respectively. Together with a Pt loading of 0.29 and 0.14 wt%, molar ratios of Ga/Pt of 55 and 86 are obtained for SiO_2 and Al_2O_3, respectively. Both Ga-rich alloy compositions are expected to be fully liquid at temperatures exceeding 200 °C.[29]

The performance of GaPt/SiO_2 and GaPt/Al_2O_3 SCALMS during PDH at 450 and 500 °C was qualitatively evaluated by means of *in situ* HRTGA-MS using a XEMIS sorption analyser. We recently hypothesised that the coking behaviour of the catalysts may be dominated by the applied carrier materials.[11] Coking can be monitored using *in situ* HRTGA-MS as well, but at least three processes may affect the sample weight during PDH over SCALMS: gas–metal interaction such as adsorption, formation of carbonaceous deposit, and reduction of oxidic gallium

species (GaO$_x$) by the dehydrogenation product H$_2$.[11] The latter species are present in the as-prepared SCALMS due to the passivation of metallic Ga during the synthesis of SCALMS or the subsequent exposure to air. Abstraction of oxygen from these Ga$_x$O species is enhanced by the presence of Pt atoms[10,30,31] and consequently decreases the sample weight. The activity of the catalysts during PDH was qualitatively monitored *via* mass spectrometry (MS). While the fragmentation patterns of propane and propylene show high similarities, the mass-to-charge ratios (m/z) of 1 and 2 can be assigned to propylene (Fig. S1†). During PDH, H$_2$ is formed at an equimolar ratio together with propylene (eqn (1)) and has a parent ion peak at $m/z = 2$, which will dominate this mass-to-charge ratio during PDH. Equal signal strengths are expected for the fragments of H$_2$ and propylene for $m/z = 1$ (Fig. S1†). Hence, $m/z = 2$ resembles the formation of H$_2$ from propane, while $m/z = 1$ describes the formation of propylene and H$_2$. Normalisation of the ion signal strengths to the parent ion peak of propane ($m/z = 29$)
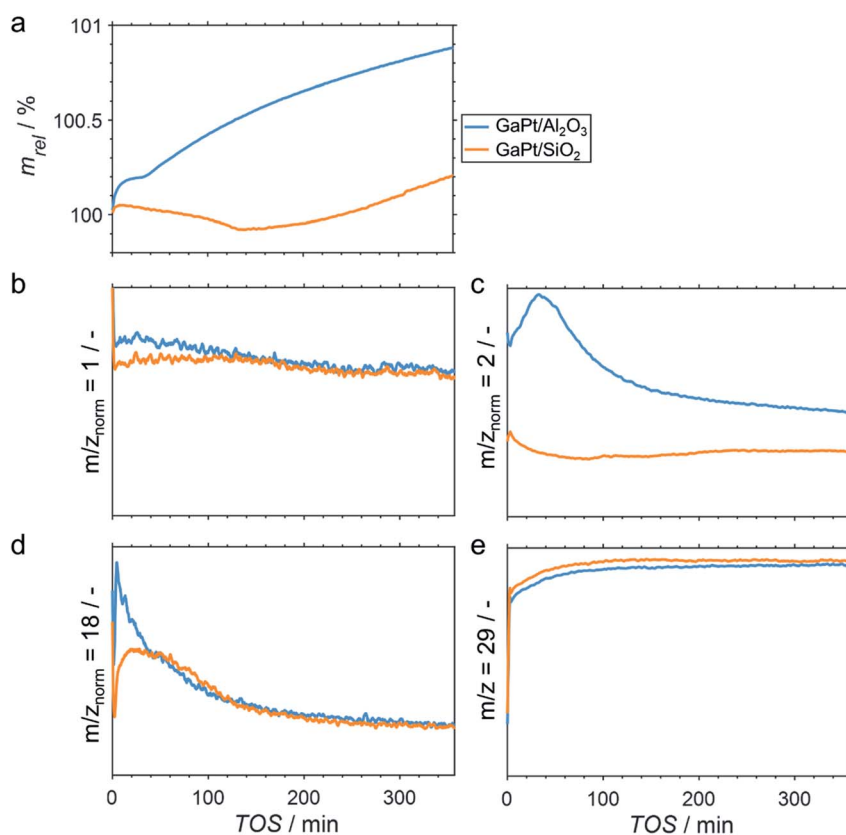

Fig. 2 (a) Sample weight relative to the weight of the dried SCALMS during PDH over GaPt/SiO$_2$ and GaPt/Al$_2$O$_3$ SCALMS with atomic Ga/Pt ratios of 55 and 86, respectively, at 500 °C; (b–d) mass-to-charge ratios of 1 (propylene and H$_2$), 2 (almost exclusively H$_2$), and 18 (exclusively H$_2$O) relative to the mass-to-charge ratio of 29 (exclusively propane); and (e) mass-to-charge ratio of 29 (exclusively propane) as monitored *via in situ* high-resolution thermogravimetry coupled with mass spectrometry. Conditions of the experiment: 180 mL$_N$ min^{-1} He; 20 mL$_N$ min^{-1} C$_3$H$_8$; WHSV 60 000 mL$_N$ g^{-1} h^{-1}.

isolates the signal from flow effects due to the large dead volume of the HRTGA-MS.

During PDH at 500 °C, the sample weight of the SCALMS increases upon first exposure of the pre-dried sample to 10% propane in He (Fig. 2a). The weight increase is due to the first interaction of propane with the catalytically active metal atoms of the alloy at the liquid–gas interface of SCALMS.[10–12] Adsorbed propane may not only be dehydrogenated to propylene and H_2, but also result in initial coke formation contributing to this first increase of the sample weight. The weight increase is more pronounced for the GaPt/Al_2O_3 SCALMS than for the GaPt/SiO_2 SCALMS (0.19 vs. 0.04%). As the former catalyst contains lower quantities of Pt when compared to the SiO_2-supported SCALMS (0.14 vs. 0.29 wt%), this may be a first indication of enhanced coking over GaPt/Al_2O_3. Within the first 30 min time on stream (TOS), the sample weight stabilises (GaPt/Al_2O_3) or even decreases (GaPt/SiO_2) due to the in situ reduction of GaO_x species by the dehydrogenation product H_2.[11] This reduction is also evidenced when analysing the off-gas by means of MS. The simultaneous formation of H_2O ($m/z = 18$) and a consumption of in situ formed H_2 ($m/z = 2$) is identified for both SCALMS (Fig. 2c and d). A continuous formation of H_2O with an exponential decay over TOS is evidenced for the GaPt/Al_2O_3 sample. The formation of H_2 reaches a maximum after 30 min TOS, which may indicate initial consumption of H_2 during an almost spontaneous reduction of GaO_x species. Contrarily, the formation of H_2O peaks after 20 min TOS and the formation of H_2 decreases for the SiO_2-supported catalyst suggesting an initially hindered reduction of GaO_x species.

The net weight change of the GaPt/Al_2O_3 SCALMS is dominated by the formation of coke after 30 min TOS (Fig. 2a) even though the reduction of the GaO_x species is not completed (Fig. 2d). The formation of coke seemingly decelerates with extended TOS, which is in contrast to previously observed constant coke formation over a GaRh/Al_2O_3 SCALMS.[11] When using SiO_2 as a support material, the weight only increases after 140 min TOS indicating a higher resistance of this SCALMS against coking. A linear weight increase after 100 min TOS may be exclusively due to the continuous deposition of coke. Similar results were obtained when analysing PDH over the same SCALMS materials in situ at 450 °C (Fig. S2†). However, the duration of the reduction of GaO_x species was extended due to the reduced reaction temperature. In general, 450 °C is within a rather moderate temperature range for PDH. Coke formation dominates the weight increase after 180 min TOS for the GaPt/Al_2O_3 sample, while the weight decreases throughout PDH for the GaPt/SiO_2 catalyst suggesting low or zero coke formation.

The performance of SCALMS was analysed by MS. The conversion of propane to the desired dehydrogenation product propylene ($m/z = 1$ normalised to the parent ion peak of propane $m/z = 29$) decreases continuously upon first exposure to propane (Fig. 2b). Initial deactivation of catalysts during PDH is well established in the literature[1–3] and has recently been reported for GaRh/Al_2O_3 SCALMS.[10,11] Blockage of active sites by carbonaceous deposits may be at play, which may (in part) lead to the initial weight increase of the SCALMS (Fig. 2a). However, the herein observed deactivation is less pronounced, which can be assigned to the low conversion during HRTGA-MS. Both SCALMS materials display a relatively stable performance even though a continuous build-up of coke is indicated by the weight increase for the GaPt/Al_2O_3 catalyst. Comparison of $m/$
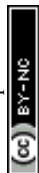
$z = 1$ and 2 allows for analysis of the ratio of formed propylene and H_2, which theoretically forms in an equimolar ratio during PDH (eqn (1)). The profile of H_2 formation ($m/z = 2$) differs from the one of propylene and H_2 ($m/z = 1$; Fig. 2c), which is in part due to the aforementioned reduction of GaO_x species consuming H_2.[11] Further, H_2 is a side product of coking. Significantly enhanced formation of H_2 over $GaPt/Al_2O_3$ strongly suggests enhanced coking when compared to the $GaPt/SiO_2$ sample (Fig. 2c), which is in line with the observed weight increase (Fig. 2a).

After cooling-down the reactor to 100 °C under a continuous flow of He, the potential formation of coke during PDH was analysed by TPO in 21% O_2/He. The weight increase upon first exposure to the oxidative atmosphere is mostly due to adsorption of O_2 to form GaO_x species (Fig. 3a).[11] In addition, the increase may in part originate from the formation of oxygen-containing functionalities ($C(O_x)$) on the surface of coke.[11,32–36] The temperature increase to 500 °C at 1 °C min^{-1} results in an enhanced weight increase of the spent catalysts (Fig. 3a), which is followed by decomposition of $C(O_x)$ functionalities. This volatilisation of coke results in a net decrease of weight and can be monitored via the formation of CO_2 by means of MS (Fig. 3b). As expected, the $GaPt/SiO_2$ sample displayed a less pronounced weight loss of 0.25%, while the sample weight of the $GaPt/Al_2O_3$ catalyst after PDH at 500 °C decreased by 0.89%. The same trend can be observed for the formation of CO_2. Hence, the formation of coke over Al_2O_3-supported SCALMS is clearly enhanced when compared to SiO_2 as the carrier material. Once again, similar results were obtained during TPO of SCALMS after PDH at 450 °C (Fig. S3†). The
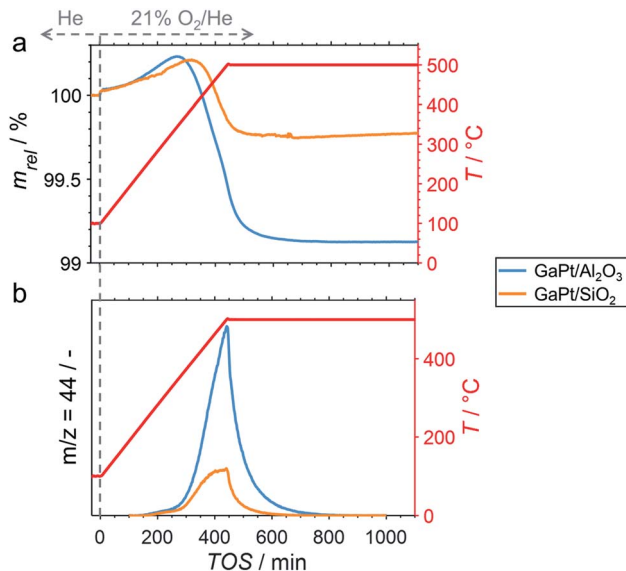

Fig. 3 (a) Sample weight relative to the weight prior to exposure to 21% O_2/He at 100 °C, and (b) formation of CO_2 during TPO (1 °C min^{-1}) of spent $GaPt/SiO_2$ and $GaPt/Al_2O_3$ SCALMS with atomic Ga/Rh ratios of 55 and 86, respectively, after PDH at 500 °C for 24 h as monitored via in situ high-resolution thermogravimetry coupled with mass spectrometry. Conditions of the experiment: 100 mL$_N$ min^{-1} He (TOS < 0); 79 mL$_N$ min^{-1} He and 21 mL$_N$ min^{-1} O_2 (TOS > 0); WHSV 30 000 mL$_N$ g^{-1} h^{-1}.

major reason for this pronounced affinity towards coking may be the increased concentration of acidic sites in the Al_2O_3 material. It is well established in the literature, that such acidic sites promote cracking of hydrocarbons and the formation coke.[4–8]

Due to the lower affinity towards the formation of coke during PDH, the performance of the $GaPt/SiO_2$ catalyst with an atomic Ga/Pt ratio of 55 was evaluated using the CPR (Fig. 1) at reaction temperatures of 450, 500, and 550 °C. The initial conversion of propane at the end of the catalyst bed (length: 55 mm) increases with temperature from ~13 to ~22% (Fig. 4). The initial conversion levels at given conditions in the CPR are in the range of the thermodynamic equilibrium conversion for pure propane,[37,38] but the heavy dilution of the feed stream with inert gas shifts the thermodynamic equilibrium in the present study to higher conversion levels. A rapid decay of the initial high activity is observed for all reaction temperatures. The initial activity of Ga–Pt SCALMS reduces between 41–45% within a start-up time of 200 min. The final activity after >900 min TOS is reduced by 59, 55, and 49% of the initial activity for a reaction temperature of 450, 500, and 550 °C, respectively. Hence, the SCALMS may be operated at a relatively stable operation point after the initial start-up period. Such deactivation behaviour of catalysts during PDH has often been described in the literature.[1–3] The herein conducted *in situ* HRTGA-MS measurements suggest a moderate deposition of coke during PDH over $GaPt/SiO_2$ SCALMS at 450–500 °C (Fig. 2). However, even small amounts of monoatomic carbon may result in strong deactivation. Al_2O_3-supported SCALMS was shown to deactivate with a similar profile[10,11] and has been demonstrated to be prone to coking due to the increased acidity of the support material (see above). On the other hand, structural reorganisation with consequent morphology change of the liquid metal droplets may be another, potentially major, culprit for the observed deactivation of the SCALMS samples.[10] However, SEM imaging of such materials has not yet produced conclusive results,[10,12] *i.e.*, the exact mechanisms at play are still under investigation. Finally, a contribution of Pt–GaO_x species, a known catalyst for PDH,[39,40] may also explain the higher initial activity. The reduction of said species by *in situ* formed H_2 to the
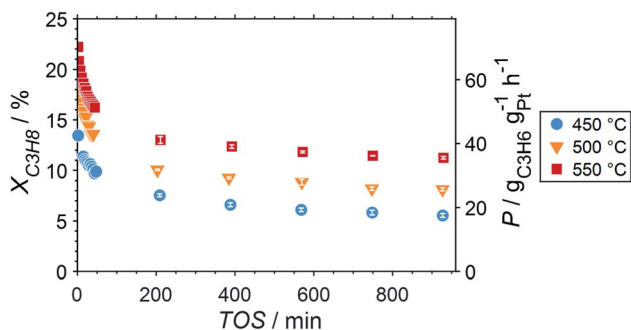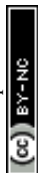

Fig. 4 Conversion of propane and productivity at the end of the catalyst bed during PDH over $GaPt/SiO_2$ SCALMS with an atomic ratio of Ga/Pt of 55 at 450–550 °C. The error bars represent the standard deviation of 6 consecutive measurements in-between profile acquisitions. Conditions of the experiment: 22 mL_N min^{-1} He; 3 mL_N min^{-1} C_3H_8; WHSV 4500 mL_N g^{-1} h^{-1}.

metallic state upon exposure to propane has been observed by means of *in situ* HRTGA-MS measurements (Fig. 2).[11] Apparently, this process transforms the initial SCALMS material into less active species and consequentially exacerbates the activity of the catalysts. Nevertheless, the rather stable operation after the initial start-up period results in productivity (P) values that are comparable with the literature for Pt catalysed PDH (Fig. 4).[41–45]

The CPR (Fig. 1) allows for analysis of spatially resolved concentration profiles during PDH over the applied novel GaPt/SiO$_2$ SCALMS. Only small amounts of catalyst are required to obtain valuable and meaningful kinetic information on activity and deactivation alike. Herein, only 333.3 mg of catalyst were loaded for each reaction temperature (450, 500, and 550 °C), which remains the only parameter to be varied to generate valuable kinetic data. As expected, a decreasing gas phase concentration of propane (Fig. 5a, c and e) and an increasing concentration of propylene (Fig. 5b, d and f) are monitored along the catalyst bed at all reaction temperatures and throughout the experiments. The profiles of the gas fraction of propane ($y_{C_3H_8}$) and the molar flow of propane ($F_{C_3H_8}$) can be mathematically described by a simple power function (eqn (2) and (3)) for all profiles at all three reaction temperatures (Fig. 6). This behaviour exemplarily
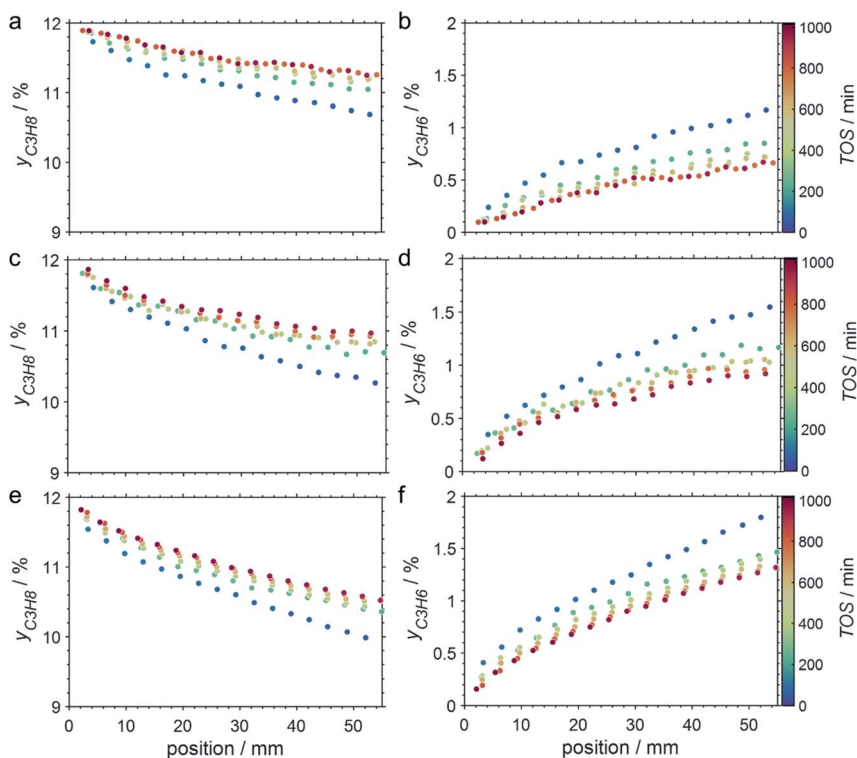

Fig. 5 Concentration profiles of propane (a, c, and e) and propylene (b, d and f) as a function of the catalyst bed length at various points in time during PDH over GaPt/SiO$_2$ SCALMS with an atomic ratio of Ga/Pt of 55 at (a and b) 450 °C, (c and d) 500 °C, and (e and f) 550 °C. Conditions of the experiment: 22 mL$_N$ min^{-1} He; 3 mL$_N$ min^{-1} C$_3$H$_8$; WHSV 4500 mL$_N$ g^{-1} h^{-1}.
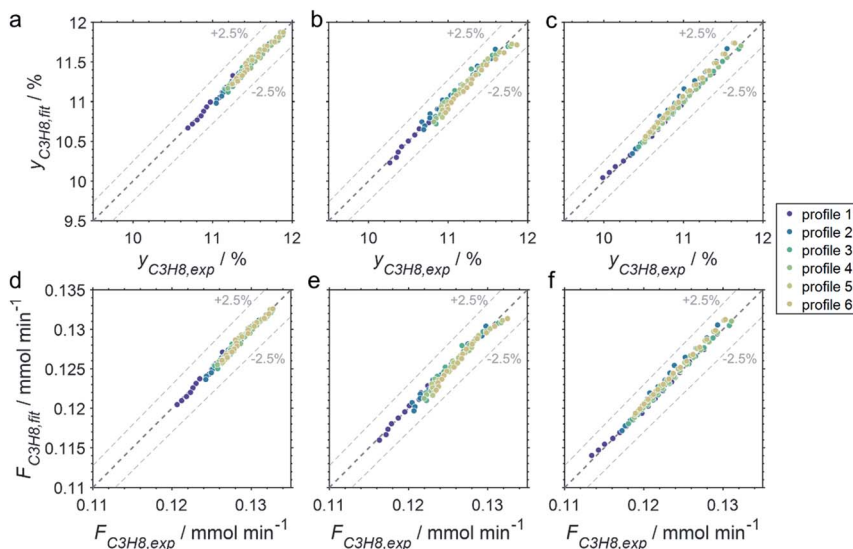
Fig. 6 Parity plots of the experimental and fitted (a–c) gas fractions and (d–f) molar flows of propane of the six profiles acquired at approx. 82, 259, 437, 614, 791, and 969 min time on stream during PDH over GaPt/SiO$_2$ SCALMS with an atomic ratio of Ga/Pt of 55 at (a and d) 450, (b and e) 500, and (c and f) 550 °C.

demonstrates the high consistency and significance of the data acquired in the CPR.

$$y_{C_3H_8}(x) = 12\% - a \times x^{b}\%^{1-b} \tag{2}$$

$$F_{C_3H_8}(x) = (0.134 - c \times x^d \, \text{mm}^{-d}) \, \text{mmol min}^{-1} \tag{3}$$

with x being the position in the catalyst bed in mm, $4.59 \times 10^{-2} < a < 2.33 \times 10^{-1}$, $0.54 < b < 0.72$, $4.68 \times 10^{-4} < c < 2.32 \times 10^{-3}$, and $0.54 < d < 0.72$.

SCALMS have been reported to display a superior alkene selectivity during alkane dehydrogenation when compared to conventional solid-phase
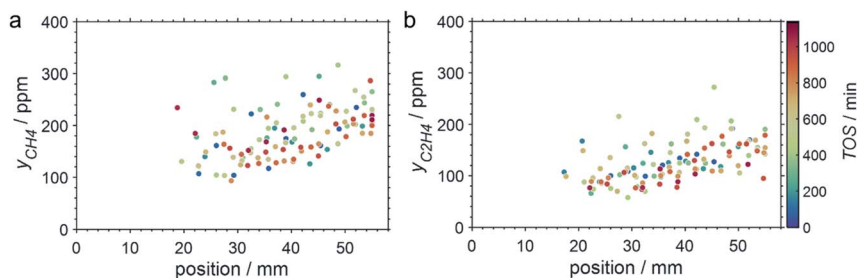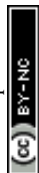

Fig. 7 Concentration profiles of (a) methane and (b) ethylene as a function of the catalyst bed length at various points in time during PDH over GaPt/SiO$_2$ SCALMS with an atomic ratio of Ga/Pt of 55 at 550 °C. Conditions of the experiment: 22 mL$_N$ min^{-1} He; 3 mL$_N$ min^{-1} C$_3$H$_8$; WHSV 4500 mL$_N$ g^{-1} h^{-1}.

heterogeneous catalysts.[10,12] Herein, only negligible amounts of side products were identified in the μ-GC. In fact, reasonable signal-to-noise ratios were obtained for methane and ethylene during PDH at 550 °C only and exclusively in the second half of the catalyst bed. Both compounds are expected side products of PDH and may form *via* the cracking of propane (eqn (4)).[5] However, concentrations as low as 100–250 ppm were identified (Fig. 7). Hence, the (gas phase) selectivity towards propylene from propane can be expected to be close to 100% for 450–500 °C. At the highest reaction temperature of 550 °C, the propylene selectivity still exceeds 97% after 900 min TOS.

$$C_3H_8 \leftrightarrow CH_4 + C_2H_4 \tag{4}$$

The captured, spatially resolved concentration profiles of the reactant propane and product propylene, indicate a continuous increase of the conversion of propane along the catalyst bed length with increased residence time (Fig. 5). However, deactivation over TOS is suggested by the steady increase and decrease of the particular profiles of gas fractions. In particular the first profile deviates from the second. This change in catalytic performance can be easily described with the corresponding conversion of propane (eqn (5)) during PDH (Fig. 8a, c and
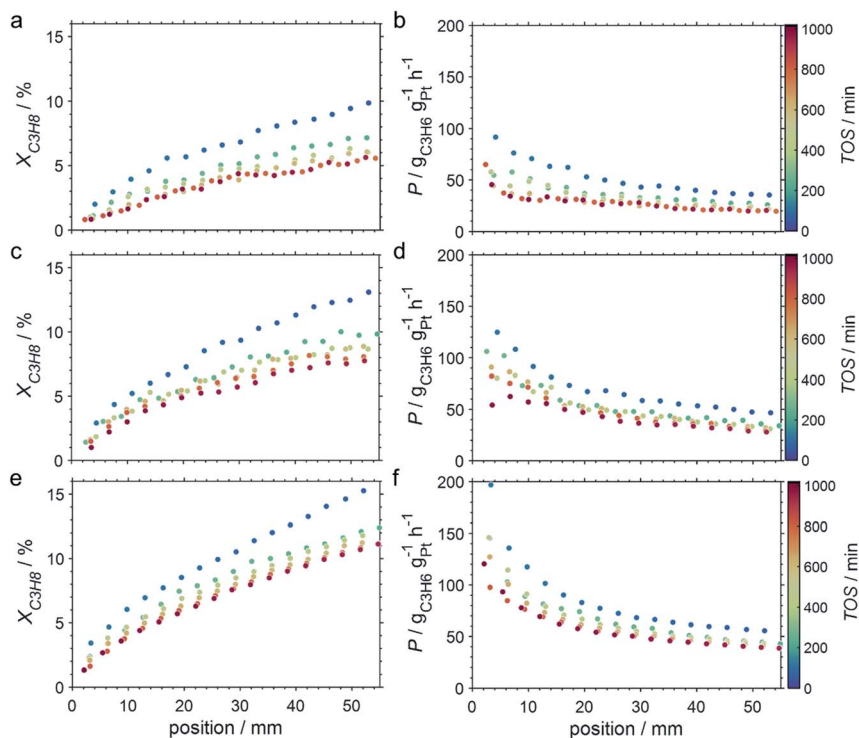

Fig. 8 Conversion profiles of propane, and productivity as a function of the catalyst bed length at various points in time during PDH over GaPt/SiO$_2$ SCALMS with an atomic ratio of Ga/Pt of 55 at (a and b) 450 °C, (c and d) 500 °C, and (e and f) 550 °C. Conditions of the experiment: 22 mL$_N$ min^{-1} He; 3 mL$_N$ min^{-1} C$_3$H$_8$; WHSV 4500 mL$_N$ g^{-1} h^{-1}.

e). Strong deactivation of SCALMS over TOS is observed throughout the catalyst bed in the first three profile acquisitions. Only the first profiles of the conversion of propane have a linear dependency of conversion and bed length from 15–40 mm of the catalyst bed. It is noted, that the profiles were acquired from the end of the catalyst bed to the beginning, which enhances potential effects of rapid deactivation during acquisition of the profile. At 450 and 500 °C, the obtained profiles are comparable from the third acquisition onwards (Fig. 8a and c). During PDH at 550 °C, already the second profile resembles the others, while the initial slope decreases with TOS resulting in deviation between the profiles at 15–25 mm bed length (Fig. 8e). This observation points towards a deactivation front moving from the end to the beginning of the catalyst bed during the experiments, *i.e.*, enhanced deactivation of the catalyst along the catalyst bed over TOS. This spatial dependency of the deactivation of SCALMS can be demonstrated with the integral productivity along the catalyst bed length (eqn (6)), which describes the effective utilisation of Pt atoms for the conversion of propane to propylene (Fig. 8b, d and f). Firstly, a continuous decrease along the catalyst bed at all reaction temperatures, once again, indicates lower efficiency of the catalyst at the end of the catalyst bed. On the one hand, this behaviour is anchored in the kinetics of PDH if the reaction rate is dependent on the partial pressure of the reactant or products. However, the point of divergence in consecutive profiles of the productivity shifts from the end to the beginning of the catalyst bed over TOS indicating the dependency of deactivation on TOS and the position in the catalyst bed. For example, the point of divergence between the second and third profile at 550 °C is identified at approx. 40 mm bed length, while it shifts to approx. 22 mm for the fourth and fifth profile (Fig. 8f). Hence, the catalytic performance of the remaining catalyst bed is comparable and the increase in conversion is only due to the different activity levels of the SCALMS in the beginning of the catalyst bed.

$$X_{C_3H_8}(x) = \frac{F_{C_3H_8}(x)}{F_{C_3H_8,in}} \times 100\% \tag{5}$$

$$P(x) = \frac{F_{C_3H_6}(x) \times 60 \times M_{C_3H_6}}{m_{Pt} \times x/x_{bed}} \tag{6}$$

with x being the position in the catalyst bed in mm, F_i being the molar flow of compound i in mol min^{-1}, $F_{C_3H_8,in}$ being the molar flow of propane in the feed stream in mol min^{-1}, $M_{C_3H_6}$ being the molar weight of propylene in g mol^{-1}, m_{Pt} being the mass of Pt in the loaded SCALMS (0.3333 g × 0.29/100), and x_{bed} being the total length of the catalyst bed (55 mm).

As demonstrated using HRTGA-MS (Fig. 2, 3, S2 and S3†), coking of the SCALMS catalyst is feasible,[10,11] even for SiO$_2$ supported GaPt droplets. Carbonaceous deposits are generally known to have the potential to exacerbate the activity of catalysts due to chemical or physical blockage of active sites.[46,47] Enhanced coke formation during PDH at 550 °C is supported by the formation of considerable amounts of the cracking products methane and ethylene (Fig. 7), but coking may also be at play at lower reaction temperatures. Coking is generally expected to be pronounced under the (relatively) high conversion environment at the end of the catalyst bed,[48] which is also indicated by the concentration profiles acquired using the CPR suggesting a coking front moving from the end to the beginning of the catalyst bed over TOS. Such a spatial dependency of coke

formation has been reported for oxidative dehydrogenation of ethane to ethylene over a MoO_3/Al_2O_3 catalyst.[49] Geske *et al.* employed a similar reactor set-up and analysed the formation of coke *via in situ* Raman spectroscopy. Catalytic partial oxidation of methane using Pt coated, cylindrical α-Al_2O_3 foam monoliths has also been reported to result in an enhanced deposition of carbonaceous species in the axial direction of the reactor.[50] Once again, the formation of coke was spatially analysed by means of Raman spectroscopy after catalytic testing. In the present study, the strong fluorescent character of SiO_2 prohibited a spatially resolved
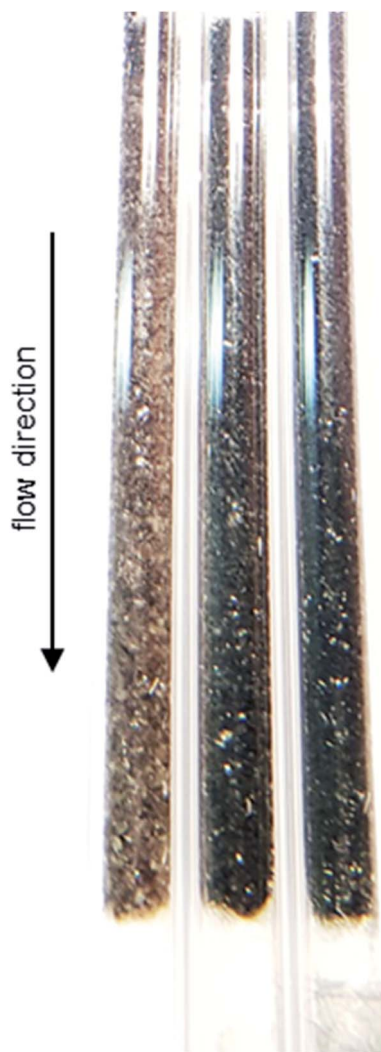

Fig. 9   Picture of the three quartz capillary reactors after PDH over GaPt/SiO$_2$ SCALMS with an atomic ratio of Ga/Pt of 55 at 450 (left), 500 (middle), and 550 °C (right) visualising the temperature dependent deposition of carbonaceous species on the catalyst. It is noted, that small amounts of coke also formed on the second part of the quartz wool plug at the end of the catalyst bed after operation at 550 °C. The first part of this quartz wool plug is seemingly not affected, most likely as it is still within the isothermal zone of the fixed-bed reactor.
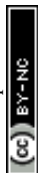
analysis of coke deposits in the $GaPt/SiO_2$ samples. However, coke deposition is clearly identified by the colour change of the SCALMS catalysts in the quartz capillary reactors after catalytic application (Fig. 9). After PDH at 550 °C, the catalyst has the darkest shade of brown/grey and even some carbonaceous residuals are identified in the quartz wool plug at the end of the catalyst bed demonstrating the high coking affinity of the reaction gas mixture during PDH at 550 °C.

In addition to the SiO_2-supported catalysts, the CPR was also employed for studying PDH over $GaPt/Al_2O_3$ SCALMS for the medium operation temperature of 500 °C. However, the increased coking affinity of this catalyst (Fig. 2 and 3) prevented sampling of reaction gas mixture *via* the orifices in the capillary over prolonged TOS, which was most likely caused by clogging of the 100 μm openings by carbon deposits from the catalyst bed. As aforementioned, the employment of the Al_2O_3 carrier material is hypothesised to cause this enhanced formation of carbon when compared to the SiO_2 carrier, which enabled the spatially resolved analysis of the coking behaviour of the $GaPt/Al_2O_3$ SCALMS by means of Raman spectroscopy. In contrast to the SiO_2-supported catalyst, a sufficient amount of carbon was deposited on this catalyst to quench the moderate fluorescence of the Al_2O_3 support material. The acquired spatially resolved Raman spectra strongly support the herein hypothesised enhanced coking at the end of the catalyst bed during PDH (Fig. 10a). Firstly, a peak evolves along the catalyst bed at the typical Raman shift for the D band of carbon ($\sim$1360 cm^{-1}),[51] which is accompanied by a generally increased signal-to-noise ratio of the spectra. Secondly, the full width at half maximum (FWHM) of the G band ($\sim$1605 cm^{-1})[51] decreases along the catalyst bed (Fig. 10b) indicating a spatially dependent increase in the quantity or a change in the nature of the carbon deposits, *e.g.* an increasing degree of graphitisation. Even though the low signal strength avoids a more detailed analysis of the Raman spectra, the results demonstrate the spatial dependency of
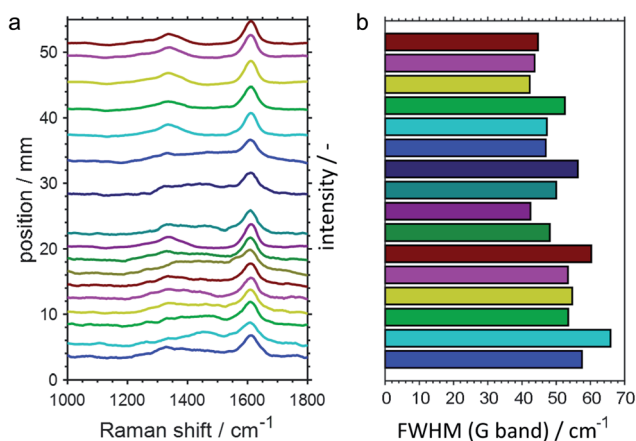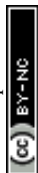

Fig. 10   (a) Spatially resolved Raman spectra of the catalyst bed after propane dehydrogenation over $GaPt/Al_2O_3$ SCALMS with an atomic ratio of Ga/Pt of 86 with (b) the corresponding full width at half maximum (FWHM) of the G band of carbon at a Raman shift of 1605 cm^{-1}. Conditions of the experiment: 500 °C; 22 mL$_N$ min^{-1} He; 3 mL$_N$ min^{-1} C$_3$H$_8$; WHSV 4500 mL$_N$ g^{-1} h^{-1}.

the coking behaviour in a GaPt/Al$_2$O$_3$ SCALMS catalyst bed during PDH, which is most likely also the case for the SiO$_2$-supported catalyst as indicated by the catalytic data (Fig. 7 and 8).

The acquired data demonstrates the great potential of the CPR for acquisition of kinetic data describing spatially resolved activity, selectivity, and deactivation alike. However, the spatial dependency of deactivation of the catalyst on the TOS, results in superimposition of intrinsic kinetic data with deactivation. Hence, the acquired profiles are strongly affected by the enhanced deactivation over the catalyst bed length resulting in a perpetual change of the apparent kinetics. Nevertheless, the data may represent the foundation for a comprehensive study of intrinsic kinetics together with deactivation kinetics, which exceeds the scope of the present study.

Summary and conclusion

The *in situ* characterisation of novel SCALMS during PDH by means of HRTGA-MS provides valuable insights on the effect of the carrier material on coke formation. Only small amounts of coke (~0.25 wt%) were formed during PDH over Ga–Pt SCALMS at 500 °C employing SiO$_2$ as the carrier material. In contrast, more than three times the amount of coke were formed during PDH over Al$_2$O$_3$-supported SCALMS due to the enhanced acidity of the carrier material. The coke is seemingly formed continuously during PDH over both SCALMS. In addition, spatially resolved kinetic data was acquired during PDH over GaPt/SiO$_2$ SCALMS using an innovative CPR. The performance of the catalyst was relatively stable after a start-up period of ~200 min with rapid deactivation of the catalyst. Structural rearrangement of the liquid Ga-rich GaPt droplets and/or rapid initial carbon deposition may cause this decay in activity. After this initial period, the concentration profiles are highly consistent, which may enable the future development of kinetic models including deactivation terms in more comprehensive studies. The results indicate an enhanced deactivation of the SCALMS at the end of the catalyst bed. In fact, a deactivation front moving from the end to the beginning of the catalyst bed over time on stream is highly likely, which may be related to accelerated coking of the catalyst under high conversion environments. Such a spatially dependent deposition of carbon deposits was successfully demonstrated by coupling the CPR with Raman spectroscopy for the GaPt/Al$_2$O$_3$ SCALMS due to the increased coking affinity of this catalyst. Only minor amounts of side products from propane cracking were identified supporting the previously reported superior selectivity of SCALMS during alkane dehydrogenation.

Conflicts of interest

There are no conflicts to declare.

Note after first publication

The data in Fig. 10, and the related discussion, was presented at the Faraday Discussions meeting and added to the manuscript post-meeting.
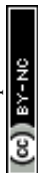
Acknowledgements

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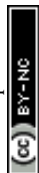
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