Electron transfer pathways in photoexcited lanthanide(III) complexes of picoline ligands†

Daniel Kovacs, Daniel Kocsi, Jordann A. L. Wells, Salauat R. Kiraev and K. Eszter Borbas

A series of luminescent lanthanide(III) complexes consisting of 1,4,7-triazacyclononane frameworks and three secondary amide-linked carboxylate antennae were synthesized. The metal binding sites were augmented with two pyridylcarboxylate donors yielding octadentate ligands. The antennae carried methyl, methoxymethyl or trifluoromethyl substituents in their 4-positions, allowing for a range of excited state energies and antenna electronic properties. The 1H NMR spectra of the Eu(III) complexes were found to be analogous to each other. Similar results were obtained in the solid-state by single-crystal X-ray crystallography, which showed the structures to have nine-coordinate metal ions with heavily distorted tricapped trigonal prismatic geometries. Steady-state and time-resolved luminescence spectroscopy showed that the antennae could sensitize both Tb(III) and Eu(III), however, quantum yields were lower than in other octadentate complexes lacking pyridylcarboxylate. Complexes with more electron-poor pyridines were less emissive even when equipped with the same antenna. The oxidation and reduction potentials of the antennae and the pyridinecarboxylates, respectively, were determined by cyclic voltammetry. The obtained values were consistent with electron transfer from the excited antenna to the pyridine providing a previously unexplored quenching pathway that could efficiently compete with energy transfer to the lanthanide. These results show the crucial impact that photophysically innocent ligand binding sites can have on lanthanide luminescence.

Introduction

Pyridines are among the most versatile ligands for metal coordination. Transition metal pyridine and bipyridine complexes promote self-assembly1,2 and a variety of transformations,3–6 while their photophysical properties are used for both analyte detection7,8 and photocatalysis.9 The trivalent lanthanide (Ln) ions have coordination requirements that are distinct from those of the d-block metals. The 4f orbitals are shielded by the 5s and 5p orbitals, therefore, ligand binding is governed by coulombic forces. Since the confirmation of bidentate binding of α-picoline and its N-oxide to Ln(III) ions a large variety of pyridine-based ligands have been developed for Ln(III) coordination (Fig. 1).10–17 Ln(III) ions commonly have large coordination numbers, typically 8–9,16 with some notable exceptions.17,18 Complex stability is improved by the introduction of additional coordinating groups onto the α-picoline core, e.g. carboxylates,19 phosphonates,20,21 tetrazolates,22 and oxazolines23 (LⅠ–LⅥ, Fig. 1a). Integration of the pyridine into a polyazamacrocycle (LⅦ–LⅨ, Fig. 1a) can further increase stability,24–28 as demonstrated by the Eu(III) complex of locked-in cyclam LⅦ incorporating two bidentate picolinites which is stable for over 167 days in 2 M aqueous HCl; under the same conditions the analogous DOTA chelate has a half-life shorter than 7 h.29

Ln(III) emitters have now reached a level of maturity, indicated by the commercial availability, and industrial and medical application of several luminescent cryptates.30,31 However, this does not imply a full understanding of their photophysical behaviour. Ln(III) emissions originate from f-f transitions, and their direct excitation is inefficient. Indirect excitation is possible by placing a light-harvesting moiety (antenna) close to the Ln(III). Pyridine-based ligands allow for the introduction of sensitizing antennae in a variety of constellations (Fig. 1a). Simple pyridines, picolinites, and bipyridines are suitable sensitizers for multiple Lns,15,23,32–34 Cs3Eu(LⅠ)3 and Cs3Tb(LⅠ)3 have found use as standards for luminescent quantum yield (Φ) determinations35 and as invisible inks.36 Tbn(Ⅲ)3 has a luminescence quantum yield of 90%.37 The
heterocycle absorption spectrum can be red-shifted through judiciously chosen para-substituents, or by fusion of additional aryl rings onto the pyridine.21 The push–pull systems obtained by para-functionalization with electron-donating alkynes yield chromophores that are excellent sensitisers for Eu(III),38–42 and have large two-photon absorption cross-sections.43–47 Thus, the near infrared-emitting Yb⁷⁷Ⅲ and related structures can be excited in the red at λex = 760 nm.48,49 Sensitising antennae can also be grafted onto the pyridine using a linker,50–54 or both pyridine ligand and antenna can be assembled around a macrocyclic core.55 The coordination and sensitising properties of the ligand are often assumed to work independently of each other.

Energy transfer (EnT) can occur from the ligand triplet or singlet via several alternative mechanisms, and both the antenna and the lanthanide excited states are subject to quenching by a variety of processes. We have recently reported that photoinduced electron transfer (PeT) from the antenna to the Ln(III) is a prominent quencher for several lanthanides.56 PeT quenching was in some cases comparable in effect to that of the well-known X–H quenching.57 Vibronic quenching of the lanthanide excited states due to coupling to X–H (X = O, N, C) oscillators has long been known, and provides the basis for the determination of the hydration state of lanthanide complexes.58–60 In an effort to eliminate a coordinated water quencher from octadentate DO3A-type Eu complexes we prepared a nonadentate ligand (Fig. 1b, left).55 Much to our surprise no improvement was seen in the Eu(III) luminescence quantum yield (ΦEu) compared to the octadentate parent. Analysis of the sensitization efficiency (ηsens) and the residual antenna fluorescence quantum yield (ΦL) suggested the presence of an additional quenching process depleting the antenna. We hypothesised that the pyridine engaged in a competitive PeT from the excited antenna. This suggestion was somewhat supported by re-analysis of the photophysical properties of lanthanide complexes in picolinate-linked oxycoumarin antennae.51,52 Specifically, we noted a marked decrease in the coumarin emission when the ligand contained a picolinate binding site (Fig. 1b, right). This difference was observable irrespective of the nature of the lanthanide, including in non-photoactive, redox-inactive Gd(Ⅲ), photophysically active, redox-inactive Tb(Ⅲ), and luminescent and reducible Eu(Ⅲ). Interestingly, in the nonadentate ligands binding Tb(Ⅲ) or Gd(Ⅲ) there was only a minor difference (10 and 5%, respectively) in ΦL compared to the octadentate ones (Fig. 1b, left).

Fig. 1  (a) A selection of pyridine-containing ligands for Ln(Ⅲ) coordination and sensitisation. (b) Putative electron transfer steps in photoexcited Ln(Ⅲ) complexes carrying carbostyril (left)⁵⁵ or coumarin (right)⁵² light-harvesting antennae. (c) Complexes prepared in this work for the investigation of antenna-pyridine ligand interactions.
Here, we investigate the fate of photoexcited carbostyril antennae in ligands containing pyridines. The structures and the compound numbering are shown in Fig. 1c. All ligands are octadentate, the Ln(n) coordination sphere is completed by a water molecule. As metal-bound water molecules quench the Ln excited state, keeping this number constant facilitates comparison of the lanthanide-based emissions. Sensitisation is performed by carbostyril antennae carrying either electron-donating Me, or slightly or strongly electron-withdrawing methoxymethyl (MOM) or CF3 groups in the 4-positions. The abovementioned class of ligands contains two unsubstituted carboxypyridine groups and the electronically different carbostyrils (LH,Me, LH,MOM, LH,CF3). In the second set of ligands, four donor atoms are provided by two carboxypyridines with electron-donating (OMe, LOMe,MOM) and electron-neutral (H, LH,OMOM) and electron-withdrawing (Cl, LCl,MOM and CF3, LCf3,MOM) para-substituents; the antenna in these ligands is only the MOM-substituted carbostyril. These ligands test the electron-accepting ability of the pyridines. In both sets of ligands the antenna and the coordination sphere are linked through a secondary amide.

Results and discussion

Synthesis

Complexes were synthesised as shown in Scheme 1. Additional details, experimental procedures, and characterization for all new compounds are given in the ESI†. Triazacyclononane (TACN) 1 was dialkylated with picolinate benzylic bromide 2. Alkylation of the secondary nitrogen in the macrocycle was possible using the appropriately substituted carbostyril chloroacetate 3. Basic hydrolysis of the methyl esters yielded the ligands L, which could be complexed by exposure to LnCl3 (Ln = Eu, Tb, Gd) in a warm EtOH : H2O mixture. The complexes were isolated in quantitative yield as white solids.

To investigate the hypothesised communication between the excited antenna and the pyridines, a series of complexes was prepared where the picolinate arms were para-substituted with electron-donating OMe, or electron-withdrawing Cl or CF3 groups (Scheme 1, right). TACN 1 was monoalkylated with the MOM substituted carbostyril chloroacetate 3b yielding intermediate 5. Dialkylation of the second secondary nitrogens was possible using p-substituted picolinate benzylic bromides, giving 6X,MOM. Methyl ester hydrolysis and complexation were carried out as described above, and the complexes were isolated in 73% to quantitative yield as white solids.

1H NMR spectroscopy

The ligands with the variously p-substituted picolinates had similar 1H NMR spectra. The chemical shifts of H-3 and H-5 of the protected ligands 6X,MOM were indicative of the electronic properties of the pyridine.51 H-3 were observed at 7.38, 7.78–7.89, 7.88, and 8.05 ppm, and H-5 at 7.12, 7.57–7.66, 7.70, and 7.95 ppm for the increasingly electron-withdrawing OMe, H, Cl, and CF3 substituents, respectively (Fig. S1†).

Crystallography

Single crystals suitable for X-ray diffraction analysis were obtained by vapour diffusion of glyme into concentrated aqueous solutions of GdLH,MOM,Cl, TbLCF3,MOM,Cl, EuLCF3,MOM,Cl, and EuLCf3,MOM,Cl. The fluoride structures were obtained from solutions containing one equiv. KF, which was added to facilitate crystallization. All four compounds have a nine-coordinate Ln centre with (heavily) distorted tricapped trigonal...
prismatic geometry (Fig. 2 and S16–S18†). The planes of the trigonal prism are formed by the triazacyclononane N-donors (N$_{3PL}$), and the pyridine N- and antenna amide O-donors (NNOPL). The remaining two carboxylate ligands and a fluoride or chloride ion cap the trigonal prism (Fig. 2). The two planes are not co-planar with N$_{3PL}$–Ln–NNOPL angles ranging 114–120°, owing to the significantly distorted geometry. The lanthanide centres are just below NNOPL (∼0.3 Å), and the Ln–N$_{3PL}$ distances range 2.017(4) to 2.065(4) Å. The complexes are racemic in the solid state, with both Δ and Λ isomers present in the unit cell.

Overall, the bond distances are comparable across this series of complexes, with no significant changes between fluoride- or chloride-bound complexes. The Ln–N$_{TACN}$ range 2.630(6)–2.671(7) Å, comparing well with related lanthanide complexes reported previously. The Ln–N$_{py}$ distances for GdL$_{X}$, MOM and TbL$_{X}$, MOM (2.557(2) Å and 2.532(6) Å, respectively) are comparable to those of related complexes (mean 2.533 Å) while those for GdL$_{X}$, MOM–F and EuL$_{X}$, MOM–Cl are slightly elongated (2.581(5) Å and 2.594(8) Å, respectively). The carboxylate Ln–O distances range 2.403(2)–2.436(3) Å, which is slightly longer than those reported for related TACN complexes (mean 2.338 Å), and could be due to the steric demand of the carbostyril antenna and the smaller TACN ligand platform. The antenna Ln–O1 distances range 2.450(5)–2.474(2) Å, comparing well to related carbostyril-substituted cyclohexane complexes (mean 2.444 Å). The Ln–F distances are shorter than in the square antiprismatic Eu cyclen complex previously reported (range 2.117(3)–2.185(2) Å versus 2.225(2) Å), and are shorter than the Ln–Cl (TACN) and Ln–O (cyclen) distances of related complexes.†

**Electrochemistry**

The ability of the pyridines to accept electrons from the photoexcited antenna was evaluated by determining the reduction potentials of the $p$-substituted pyridines and the oxidation potential of the MOM-substituted carbostyril antenna. The reduction potentials of the unsubstituted picolinate fragments were measured by cyclic voltammetry in the model compounds GdL$_X$ (X = CF$_3$, Cl, H, OMe, Fig. 3) to account for the effect of Ln($III$) binding but avoid potential interference from the antenna. The CF$_3$-substituted pyridines were expected to be the easiest to reduce providing a well characterised irreversible reduction wave at −1.13 V (vs NHE), while the Cl-substituted ones being less reducible showed diminished response at −1.21 V (vs NHE). Along the series the H-substituted analogues revealed faint reduction wave at −1.29 V (vs NHE), while the reduction of the OMe-substituted one was outside of the solvent window (Fig. 4a).

The cyclic voltammograms of GdL$_X$, MOM showed poorly defined irreversible reduction waves at −1.13 V, −1.21 V, −1.36 V, and −1.43 V (vs NHE) for X = CF$_3$, Cl, H, and OMe, respectively (Fig. 4b). Irreversible reductions at comparable potentials (−1.54 V vs Ag/AgCl/KCl (sat.), −1.34 V vs NHE) have been observed during CV analysis of pyridine under acidic conditions. Therefore, these waves were assigned to picolinate-based reduction events. The influence of the $p$-substituent on the pyridine reduction potential mirrors what has been observed in Pd($II$)–C^N^C pincer complexes.61 The first

![Fig. 3 Model compounds for electrochemistry.](image)

**Fig. 3** X-ray crystal structure of GdL$_X$, MOM–F (top) and the Gd coordination environment (bottom). H atoms and solvent molecules are omitted for clarity. The first coordination sphere of the metal centre is depicted with ellipsoids at 50% probability, the rest of the atoms are displayed as capped sticks.
reduction event in the latter series occurred at ∼0.2 V more positive values than in the Gd complexes, which is consistent with their lacking the fully negatively-charged carboxylate. The reduction potential of the model compound \( \text{GdLH} \) was more positive by ∼70 mV than that of the analogous complex with the antenna, \( \text{GdLH,MOM} \). Variations in the amide substituents could influence the reduction potentials of the encapsulated metals, and likely for the rest of the ligand. However, the effect of such substitution is expected to be much smaller than seen here and of the opposite direction.67

The excited state oxidation potential of the MOM-substituted carbostyril was estimated from the singlet excited states of the ligand (3.53 eV, *vide infra*) and the oxidation potential of the ground state antenna (1.76 V vs. NHE, *AcCS*MOM, Fig. 3 and S19–S21†). Using eqn (1), the free energy of the electron transfer (\( \Delta G_{ET} \)) can be calculated from the oxidation potential of the carbostyril (\( E_{ox} \)), the reduction potential of the pyridine (\( E_{red} \)), the excited state of the antenna (\( E_S \)), and the attraction between the radical ion pair (\( e_0^2/\varepsilon \)). The last term is ∼0.15 eV for an exciplex,68 and this is the value we will use.

\[
\Delta G_{ET} = (E_{ox} - E_{red}) - E_S - \frac{e_0^2}{\varepsilon} \quad (1)
\]

Negative \( \Delta G_{ET} \) was calculated for PeT from \( S_1 \) for \( \text{GdL}^\text{H} \) (−0.63 eV) and the \( \text{GdL\text{X,MOM}} \) (−0.79, −0.71, −0.56, −0.49 eV for \( \text{X} = \text{CF}_3, \text{Cl}, \text{H}, \text{and OMe, respectively} \), PeT from the first triplet excited state (\( T_1 \)) was not thermodynamically favoured. Linear correlations were found between the substituent constants described by Hammett and the measured reduction potentials of the pyridine units. The picolinate electron accepting ability increases with increasing electron withdrawing ability of the \( \text{para} \)-substituent (Fig. 5). This observation supports our hypothesis that as the electron accepting ability of the pyridines increase, greater PeT quenching of the antenna is observed. Therefore, lower sensitization efficiency, and thus weaker Ln luminescence was expected.

**Photophysical studies**

The expected effect of the pyridine electronic properties on Ln(III) luminescence were investigated using UV-Vis absorption, and steady-state and time-resolved emission spectroscopies. The photophysical properties of \( \text{LnL} \) were determined in PIPES-buffered aqueous solutions (pH 6.5) at complex concentrations in the 10–15 μM range. Changes in the absorption spectra were observed both upon changing the picolinate \( p \)-substituent, and the 4-substituent of the antenna. Pyridine substitution caused absorption changes only in the 250–310 nm range. Above 310 nm only the carbostyril antennae absorb, thus, emission spectra collected with \( \lambda_{ex} > 330 \text{ nm} \) are due to carbostyril-based sensitisation. Absorption (\( \lambda_{abs} \)) and emission (\( \lambda_{em} \)) maxima were red-shifted with increasing electron-withdrawing ability (\( \text{CF}_3 > \text{MOM} > \text{Me} \)) of the carbos-
Table 1  Antenna photophysical properties of LnL

<table>
<thead>
<tr>
<th>Complex</th>
<th>$\lambda_{abs}$ [nm]</th>
<th>$\lambda_{em}$ [nm]</th>
<th>$E_{00}(S_1)$ [cm$^{-1}$]</th>
<th>$E_{00}(T_1)$ [cm$^{-1}$]</th>
</tr>
</thead>
<tbody>
<tr>
<td>GdLH,Me</td>
<td>329</td>
<td>366</td>
<td>28 900</td>
<td>22 900</td>
</tr>
<tr>
<td>GdLOMe,MOM</td>
<td>331</td>
<td>375</td>
<td>28 500</td>
<td>22 500</td>
</tr>
<tr>
<td>GdLMe,MOM</td>
<td>331</td>
<td>375</td>
<td>28 500</td>
<td>22 500</td>
</tr>
<tr>
<td>GdLCl,MOM</td>
<td>331</td>
<td>375</td>
<td>28 500</td>
<td>22 500</td>
</tr>
<tr>
<td>GdCl,MOM</td>
<td>331</td>
<td>375</td>
<td>28 500</td>
<td>22 500</td>
</tr>
<tr>
<td>GdLH,CF3</td>
<td>342</td>
<td>391</td>
<td>27 400</td>
<td>21 700</td>
</tr>
</tbody>
</table>

*a* In aqueous PIPES buffer (10 mM), pH 6.5. $b$ $\lambda_{em}$ = 329 nm (GdLH,Me), 331 nm (GdLOMe,MOM, GdLCl,MOM). $c$ Calculated from the 0-0 phonon transitions observed in the Gd-complex at 77 K.

The $T_1$ energies of the antennae were determined from the 77 K steady-state emission spectra of the Gd-complexes (Table 1 and Fig. S42†). The CF$_3$-substituted antenna had the lowest energy, while the highest value was found for the Me-substituted one. The antenna fluorescence excitation spectra matched the absorption spectrum attributed to the carbostyril unit (Fig. S30–S41†). The excitation spectra of the phosphorescence bands, however, matched the absorption spectrum of the entire complex. The pyridine pendant arms have observable triplet energy levels in this region (Fig. S44†). The antenna triplets are reasonably well-placed for energy transfer to Eu(m), as they are more than 2000 cm$^{-1}$ above its $^5D_0$ excited state. The $T_1$ of the CF$_3$-substituted antennae at 21 700 cm$^{-1}$ (Table 1) were within 2000 cm$^{-1}$ above the Tb(m) $^5D_1$ excited state at 20 490 cm$^{-1}$.

Antenna excitation at $\lambda_{ex}$ = 330 nm resulted in Tb(m) and Eu(m) emission. For Tb the $^5D_4 \rightarrow ^7F_j$ ($J = 6 \rightarrow 0$) transitions were seen at 488, 543, 582, 620, 652, 668 and 678 nm, respectively. The $J = 5$ transition was found to be more than twice as high as the second most intense peak. In the Eu spectra the $^5D_0 \rightarrow ^7F_j$ ($J = 0 \rightarrow 5$) transitions were located at 579, 593, 614, 649, 693 and 751 nm, along with the hypersensitive and higher energy $^5D_1 \rightarrow ^7F_j$ ($J = 0 \rightarrow 1$) transitions (537, 554 nm, respectively) (Fig. 6). The shape of the spectra and intensity of the peaks were different from those of the cyclen-based complexes carrying the same antenna/linker due to the different ligand environment (Fig. S49 and S50†). The most intense peak corresponds to the $\Delta J = 2$ transition in contrast to the cyclen-based Eu complexes, where the $\Delta J = 4$ transition is the strongest.

The overall luminescence quantum yields of the complexes were determined in 10 mM aqueous PIPES buffer (pH = 6.5) using quinine sulfate ($\Phi = 0.59$) in H$_2$SO$_4$ (0.05 M) as the reference (Table 2). Tb(m) quantum yields were 25.4–30.6% for the complexes carrying the 4-Me or the 4-MOM-substituted carbostyril and OMe, Cl and H-substituted pyridines. A CF$_3$-substituent in either the antenna 4-position or in the pyridine para-position resulted in diminished $\Phi_{Tb}$. The Tb(m) emission intensity of TblH,CF3 increased from $\Phi_{Tb} = 3.3\%$ upon deoxygenation; ligand emission was not affected. The oxygen sensitivity of $\Phi_{Tb}$ is consistent with back energy transfer from the Tb(m) excited state to the antenna $T_1$ and the quenching of the antenna $T_1$ by dissolved $^1O_2$ (Fig. S54 and S55†). $\Phi_{Tb}$ values were in the 0.8–8.0% range, which is comparable to complexes sensitised by carbostyril antennae in D03A frameworks.
The intensity of Ln(III) luminescence was dependent on the pyridine donors in the ligand. For complexes carrying the same antenna Φr decreased dramatically with decreasing pyridine electron density (Fig. S51–S53†). The ligand emission was only a third (2.1%) in GdL[CF3,MOM] compared to the Φr of GdL[OMe,MOM] (6.9%). The trend was reproduced in the Tb(III) complexes. Within the series, a smaller Φr was accompanied by smaller Φan further confirming that the pyridines promote quenching of the antenna. The data for Eu(III) follow the same trend, although Φr of EuL[HMOM] is somewhat larger than that of EuL[OMe,MOM]. Given the small Φr of the Eu(III) complexes, these variations carry a large uncertainty.

The luminescent lifetimes of EuL[X,R] and TbL[X,R] were determined using time-resolved luminescence spectroscopy (Table 3). In H2O EuL[X,R] had lifetimes (τH2O) ~0.50 ms, in D2O these (τD2O) were lengthened to ~1.30 ms. The number of Ln(III)-coordinated water molecules (q) were calculated using the following equations: q = (5 ms)/(1/τH2O − 1/τD2O − 0.06 ms⁻¹) for Tb, and q = (1.2 ms)/(1/τH2O − 1/τD2O − 0.25 ms⁻¹ − m×0.075 ms⁻¹) for Eu, where m is the number of nearby N–H oscillators.58,60 The complexes had one Ln(III)-bound water molecule, which is in line with an expected nine-coordinate Ln(III)-coordinated water molecules (Table 3). In H2O D2O [ms] were lengthened to 0.06 ms, 0.25 ms, respectively, as are τH2O, τD2O.

The intensity of Ln(III) luminescence was dependent on the antenna triplet is taking place in both. Given the high error of ±10%.60

The luminescent lifetimes of EuL[X,R] and TbL[X,R] were next investigated. ΦL of Eu and Tb complexes in the current complexes were determined from the 77 K luminescence spectra of GdL[X] (Fig. S44†). Due to the lack of vibrational structure in the spectra T1 were located by deconvolution of the emission bands into Gaussian functions. This analysis yielded T1 at 25 600, 25 600, 25 400, and 25 100 cm⁻¹, for GdL[X] (for X = OMe, H, Cl, and CF3, respectively; Fig. S45–S48†). These values are higher than that of the antenna T1, therefore any energy back transfer would likely involve the latter (located at 22 500 cm⁻¹). Thus, Tb(III) to picolinate EnT is unlikely.

The reasons behind the different Φr and Φan in LnL[X,R] were next investigated. ΦLn is a product of the efficiency of the sensitization (ηsens) and the intrinsic quantum yield (ΦLn(T) of the Ln (eqn (2)).73 The former is influenced by processes that are disruptive to the antenna excited state, e.g. X–H quenching of the antenna or PeT to or from the antenna. The intrinsic quantum yield is mostly determined by the coordination environment of the Ln(III). For Eu(III), ηsens and ΦLn can be calculated from the corrected Eu(III) luminescence spectrum using eqn (2) and (3).74 Here, ηrad is the Eu(III) radiative lifetime, A0,0 and AMD are the integrated full spectrum (521–800 nm) and the area for the 5D0 → 7F1 transition (618–630 nm), respectively. A0,0 (14.65 s⁻¹) is the spontaneous emission probability for the 5D0 → 7F1 transition of Eu(III) in vacuo. τobs is the same as τH2O. The refractive index of the solution is taken as n = 1.333 in H2O.75 The results are summarised in Table 4.

\[ Φ_{Ln} = η_{sens} \cdot Φ_{Ln} = η_{sens} \cdot \frac{τ_{obs}}{τ_{rad}} \]  
\[ \frac{1}{τ_{rad}} = A_{MD,0} \cdot η_{sens} \cdot \frac{I_{0,0}}{I_{MD}} \] (2)

Table 3 Luminous lifetimes and calculated inner-sphere water molecules of Eu and Tb complexes

<table>
<thead>
<tr>
<th>Complex</th>
<th>τH2O [ms] b</th>
<th>τD2O [ms] b</th>
<th>q</th>
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<tbody>
<tr>
<td>TbL[H,Me]</td>
<td>1.12</td>
<td>1.67</td>
<td>1.16</td>
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<td>TbL[OMe,MOM]</td>
<td>0.76</td>
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<td>1.29</td>
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<td>TbL[HMOM]</td>
<td>0.82</td>
<td>1.08</td>
<td>1.19</td>
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<tr>
<td>TbL[ClMOM]</td>
<td>0.73</td>
<td>0.99</td>
<td>1.48</td>
</tr>
<tr>
<td>TbL[CF3,MOM]</td>
<td>0.60</td>
<td>0.94</td>
<td>—</td>
</tr>
<tr>
<td>TbL[CF3,MOM]</td>
<td>0.09</td>
<td>0.16</td>
<td>—</td>
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<tr>
<td>EuL[HM]</td>
<td>0.50</td>
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<td>1.06</td>
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<td>EuL[CF3]</td>
<td>0.51</td>
<td>1.28</td>
<td>1.04</td>
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</table>

a In PIPES buffered (10 mM, pH 6.3), non-deaerated aqueous solutions at nominally 10 μM complex concentrations. bLifetime values were averaged from three independent measurements and are subject to an error of ±10%.60

Table 4 Eu(III)-Centred photophysical properties

<table>
<thead>
<tr>
<th>Complex</th>
<th>τrad [ms]</th>
<th>τobs [ms]</th>
<th>ηsens [%]</th>
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<tbody>
<tr>
<td>EuL[HM]</td>
<td>2.87</td>
<td>0.50</td>
<td>17.4</td>
</tr>
<tr>
<td>EuL[OMe,MOM]</td>
<td>2.67</td>
<td>0.51</td>
<td>19.0</td>
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<td>EuL[HMOM]</td>
<td>2.86</td>
<td>0.50</td>
<td>17.5</td>
</tr>
<tr>
<td>EuL[ClMOM]</td>
<td>2.69</td>
<td>0.49</td>
<td>18.2</td>
</tr>
<tr>
<td>EuL[CF3,MOM]</td>
<td>2.85</td>
<td>0.48</td>
<td>16.9</td>
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<tr>
<td>EuL[CF3]</td>
<td>2.87</td>
<td>0.51</td>
<td>17.8</td>
</tr>
</tbody>
</table>

a Measurements were performed in 10 mM aqueous PIPES buffer solutions at pH 6.5 at 10 μM concentrations.
affecting the excited antenna. More electron-rich antennae are more reducing, and are thus more likely to be quenched by PeT to Eu(III) or other oxidants,\(^{56,57,76}\) which could explain some of the \(\eta_{\text{sens}}\) differences seen between EuL\(^{H,Me}\) (5%), EuL\(^{H,MOM}\) (14%), and EuL\(^{H,CF3}\) (46%). For complexes carrying identical antennae the variation in \(\eta_{\text{sens}}\) must be caused by the rest of the ligand.

The ligand may influence \(\eta_{\text{sens}}\) indirectly, by modulating the Eu(III) reduction potential, or via a separate quenching pathway that does not involve the Ln(III). More electron-rich ligands would be expected to stabilize Eu(III), and thus result in less PeT. Thus, for the same antenna \(\eta_{\text{sens}}\) should decrease in the following order: EuL\(^{OMe,MOM}\) > EuL\(^{H,MOM}\) > EuL\(^{CL,MOM}\) > EuL\(^{CF3,MOM}\), which is what is seen, with \(\eta_{\text{sens}}\) = 20, 14, 8, and 5%, respectively. However, electron transfer to Eu(III) does not explain the \(\Phi_L\) and \(\Phi_{\text{Tb}}\) variations in Tbl\(^{X,MOM}\), which decrease in the same order from 5.3% to 1.8% and 30.6% to 13.1%, respectively. As Tb(III) is one of the most difficult to reduce Ln(III),\(^{77}\) and is stable under these conditions, these data are consistent with the picolinates directly quenching the excited antenna, which was shown above to be thermodynamically favourable. Additional support is provided by the \(\Phi_L\) values of Gdl\(^{X,MOM}\), which follow the same trend, going from 7.0% to 2.2% with increasingly electron poor picolimates (Table 2).

The extent to which the pyridine electronic properties were affecting the excited state were then analysed by plotting \(\Phi_L\), \(\Phi_{\text{Tb}}\), and \(\eta_{\text{sens}}\) for LnL\(^{X,MOM}\) (Ln = Eu, Tb, Gd) vs. the picolinate reduction potential (Fig. 7 and S56–S60).\(^{\dagger}\) Good linear relationships with \(R^2\) close to 1 were found for \(\Phi_L\) (Gdl\(^{X,MOM}\), \(\Phi_{\text{Tb}}\), and \(\eta_{\text{sens}}\) (EuL\(^{X,MOM}\)), and a somewhat poorer one for \(\Phi_{\text{Tb}}\). No correlation was found between \(\Phi_L\) (EuL\(^{X,MOM}\)) and the pyridine reduction potential, which is likely due at least in part to the large uncertainty associated with measuring such low values. However, there are at least 2 competing electron transfer pathways in Eu complexes, from the antenna to pyridine, and from the antenna to Eu(III). The electron poorer pyridines presumably destabilize Eu(III) thus also increasing PeT to Eu(III). The combination of these may cause a deviation from the trend seen in Gdl\(^{X,MOM}\) and TbL\(^{X,MOM}\).

A conceivable photochemical reaction would be the oxidation of the photoexcited Tb(III) to Tb(IV) by the electron-poor pyridine. Tb(IV),\(^{78-80}\) along with Ce(IV)\(^{81-84}\) and Pr(IV)\(^{85}\) is more stable than the tetravalent ions of the other Ln. However, the Tb(III)/Tb(IV) oxidation potential is still quite positive, 1.3 V vs. NHE in concentrated aqueous carbonate solution,\(^{86}\) and most examples of Tb(IV) are reported in the solid state rather than as coordination complexes.\(^{87}\) Applying the values of 1.3 V for \(E_{\text{ox}}\) (Tb\(^{III}/Tb^{IV}\) oxidation potential), –1.25 V for \(E_{\text{red}}\) (picolinate reduction potential), 2.54 eV for \(E_{\Phi}\) (Tb(III) excited state) to eqn (1), a \(\Delta G_{\text{ET}}\) of 0.01 eV is obtained without any coulombic stabilization. For more electron-poor picolimates \(\Delta G_{\text{ET}} < 0\) eV may be possible, which could account for the deviation from linearity for \(\Phi_L\) (TbL\(^{X,MOM}\)) vs. \(E_{\text{pc(py/py^-)}}\) and the lack of such deviation for the \(\Phi_L\) (TbL\(^{X,MOM}\)) vs. \(E_{\text{pc(py/py^-)}}\) relationship. However, given the variation in measured \(E_{\text{ox}}\) (Tb\(^{III}/Tb^{IV}\)) and the expected large influence of the ligand, Tb(III) is unlikely to be easily oxidized in +1 charged Tbl\(^{X,MOM}\).

**Conclusions**

A set of structurally related TACN-based Ln(m) complexes equipped with carbostyril sensitising antennae were prepared with the aim of studying the effect picolinate binding sites have on the luminescent properties. The similarities of the solution and solid-state geometries of the species carrying H, OMe, Cl, or CF\(_3\) substituents in the pyridine \(p\)-positions was supported by paramagnetic \(^1\)H NMR spectroscopy and single crystal X-ray crystallography. Further support was provided by the similar radiative and observed lifetimes (average values: 2.80 ms and 0.50 ms, respectively), and intrinsic quantum yields (mean value: 17.8%) for all the Eu(III)-based emitters irrespective of pyridine or carbostyril substituents.

Despite the similarities in structure and certain photophysical properties, the complexes showed remarkable differences in their photon outputs. It was revealed that lowering the electron density on the pyridine binding units promote their electron accepting ability from the excited antenna decreasing the overall quantum yields. The phenomenon is thermodynamically favoured from the antenna \(S_1\), and in Eu(III) complexes it competes with processes such as energy transfer and ligand-to-metal PeT. This intra-ligand electron transfer also impacts the emission of non-redox active Ln complexes, as shown by the large degree of quenching in residual ligand fluorescence of Gd(m) and Tb(m) complexes.

Pyridine-bearing TACN-based ligands enjoy increasing popularity due to their ease of synthesis, versatility, and stability. These systems can trace their origins to tris-picolinate-based emitters, which are among the best understood and most widely employed Ln chelates. This study demonstrates...
that the pyridine structural units may participate in unintended background reactions. Notably, many of these quenching processes can be difficult to notice without in-depth structure-property relationship analyses, and their relevance is likely dependent on the arrangement of the pyridines vis-a-vis the sensitizing antenna, as suggested by the excellent luminescence properties of several picolinate-carrying TACN-based Eu [III] chelates.

Conflicts of interest

There are no conflicts to declare.

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Notes and references