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Experimental investigations into the irregular synthesis of iron(III) terephthalate metal–organic frameworks MOF-235 and MIL-101†

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MOF-235(Fe) and MIL-101(Fe) are two well-studied metal–organic frameworks (MOFs) with dissimilar crystal structures and topologies. Previously reported syntheses of the former show that it has greatly varying surface areas, indicating a lack of phase purity of the products, *i.e.* the possible presence of both MOFs in the same sample. To find the reason for this, we have tested and modified the commonly used synthesis protocol of MOF-235(Fe), where a 3 : 5 molar ratio of iron(III) ions and a terephthalic acid linker is heated in a 1 : 1 DMF : ethanol solvent at 80 °C for 24 h. Using XRD and BET surface area (S_{BET}) measurements, we found that it is difficult to obtain a pure phase of MOF-235, as MIL-101 also appears to form during the solvothermal treatment. Comparison of the XRD peak height ratios of the synthesis products revealed a direct correlation between the MOF-235/MIL-101 content and surface area; more MOF-235 yields a lower surface area and *vice versa*. In general, using a larger (3 : 1) DMF : ethanol ratio than that reported in the literature and a stoichiometric (4 : 3) Fe(III) : TPA ratio yields a nearly pure MOF-235 product ($S_{\text{BET}} = 295 \text{ m}^2 \text{ g}^{-1}$, 67% yield). An optimized synthesis procedure was developed to obtain high-surface area MIL-101(Fe) ($S_{\text{BET}} > 2400 \text{ m}^2 \text{ g}^{-1}$) in a large yield and at a previously unreported temperature (80 °C vs. previously used 110–150 °C). *In situ* X-ray scattering was utilized to investigate the crystallization of MOF-235 and MIL-101. At 80 °C, only MOF-235 formed and at 85 and 90 °C, only MIL-101 formed.

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Introduction

There is currently a quest for finding materials with outstanding chemical and physical properties while still being economically and environmentally friendly. During the last two decades metal–organic frameworks (MOFs) have been intensively studied due to their unusually high surface areas and porosities. This enables their applications in many different important areas, *e.g.* in the areas of gas adsorption,^{3–5} water purification,⁷ molecular sensing,^{8–10} catalysis^{11,12} and drug delivery.^{13–15} MOFs consist of metal centres connected to organic ligands, forming extensive porous networks with potentially active sites in both the metal ions and ligands. The high degree of porosity enables large surface areas and to date, the MOF with the highest experimentally obtained surface area has been DUT-60 with a BET (after Brunauer–Emmett–Teller) surface area of $7839 \text{ m}^2 \text{ g}^{-1}$.¹⁶

One of the environmentally friendly MOF contenders is MOF-235(Fe), an iron(III) terephthalate MOF with the chemical formula $[\text{Fe}_3\text{O}(\text{1,4-BDC})_3(\text{DMF})_3][\text{FeCl}_4]$ (1,4-BDC: 1,4-benzenedioic acid/terephthalic acid; DMF: *N,N*-dimethylformamide). It was first synthesized by Sudik *et al.* (2005),² and since then several articles have been published where the exact same synthesis procedure has been utilized. In these publications, both determined XRD patterns and measured surface areas have differed at varying degrees. For example, the reported S_{BET} ranges from 9.6 to $540 \text{ m}^2 \text{ g}^{-1}$,^{17,18} yet no explanations for these deviations have been provided in the literature, to our knowledge.

Another well-known iron(III) terephthalate MOF is MIL-101(Fe) with the molecular formula $[\text{Fe}_3\text{O}(\text{OH})(\text{H}_2\text{O})_2(\text{1,4-BDC})_3]$. Both MOF-235 and MIL-101 are composed of non-toxic iron(III) oxo clusters connected *via* terephthalic acid (TPA) linker molecules. The main differences between these two MOFs are their topologies and pore structures; MIL-101 contains two different-sized nanocages ($d = 29$ and 34 \AA (ref. 6)) connected by windows ($d = 12$ and 16 \AA) forming a three-dimensional porous structure (Fig. 1(A)), while MOF-235 consists of one-dimensional hexagonal channels with a diameter of 6.7 \AA^2 (Fig. 1(B)). The S_{BET} of the reported MIL-101(Fe) lies in the

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range of 1000–3000 m² g⁻¹, making it a promising candidate for applications such as adsorption,^{20–23} catalysis^{19,24–28} and medicine.²⁹

MIL-101(Fe) was first synthesized³⁰ using microwave radiation (150 °C) and a 1 : 1 ratio of Fe(III) : TPA in a pure DMF solvent, resulting in a product yield of 20%. This means that to synthesize 100 mg of MIL-101 (MW = 728.169 g mol⁻¹), 90 mL of DMF, a toxic and potentially carcinogenic organic solvent, is required. Since DMF decomposes into dimethylamine and carbon monoxide during the solvothermal treatment,³¹ recycling of the solvent is not possible. In addition, a relatively high temperature of 110–150 °C is required for MIL-101 to form (see Table 4 for references). If it would be possible to use a lower amount of harmful solvent and decrease the synthesis temperature it would make the material more environmentally and economically sustainable for industrial and medical applications.

In this study we have tested and modified the most used synthesis protocol of MOF-235(Fe), where a 3 : 5 molar ratio of iron(III) ions and a terephthalic acid linker is heated in a 1 : 1 DMF : ethanol solvent at 80 °C for 24 h. We tested three different Fe(III) : TPA and DMF : ethanol ratios and investigated the influence of equilibration time on the end product. To determine the obtained products' phase purity, a combination of XRD and BET surface area measurements was used. X-ray photoelectron spectroscopy (XPS) was used as a complimentary technique. The crystal development during the solvothermal synthesis of MOF-235/MIL-101 was investigated using medium-angle X-ray scattering (MAXS). In addition, methylene blue adsorption efficiencies and zeta potentials were determined for the two most phase-pure products.

Experimental

Chemicals and instruments

All chemicals were commercially obtained and used without further purification. Solvothermal syntheses were conducted

in a Memmert UF30 Plus convection oven. Powder X-ray diffraction (P-XRD) analyses were performed by using Bruker D8 Advance Diffractometer with a Cu K α radiation source ($\lambda = 1.54 \text{ \AA}$) and a LYNXEYE PSD detector. Additional XRD analyses were performed with a Cr K α radiation source ($\lambda = 2.29 \text{ \AA}$) using a Bruker D8 Advance Diffractometer. X-ray photoelectron spectroscopy (XPS) was performed using a PHI5000 VersaProbe III instrument with a monochromatic Al X-ray source ($E = 1486.6 \text{ eV}$) with a beam size of 100 μm . BET surface areas were measured with a Micromeritics TriStar3000 instrument and the calculations were performed using data points in the 0.05–0.20 p/p_0 range. Before the adsorption study and the surface area measurements, the samples were activated in a vacuum oven (70 °C, >12 h, <50 mbar). In addition, the samples were dried in two steps in their sample holders (90 °C for >1 h, and then 110 °C for >2 h, N_{2(g)} atmosphere) immediately before starting the surface area measurements. All X-ray scattering analyses were performed with a Mat:Nordic instrument from SAXSLAB with a Rigaku 003 + Cu K α source and a Pilatus (300 K/100 K) detector in a vacuum environment. SEM images were obtained with a FEI Quanta200 ESEM in a high vacuum mode and at a 10.0 kV operating voltage.

Synthetic procedure

MOF-235 was first synthesized using a conventional solvothermal synthesis protocol.² Terephthalic acid (1.23 mmol, C₆H₄(COOH)₂) was dissolved in 30 mL of DMF in a 100 mL PTFE-lined stainless-steel autoclave. Iron(III) chloride hexahydrate (0.738 mmol, FeCl₃·6H₂O) was added during stirring together with 30 mL of ethanol (95%). The autoclave was tightly sealed and placed in a convection oven (80 °C, 24 h). The MOF dispersion was filtered and washed with 1 : 1 DMF : ethanol and pure ethanol, subsequently.

The original synthesis process reported in the literature did not yield a pure MOF-235 product; additional peaks were observed in the XRD spectrum, belonging to MIL-101 (Fig. 1). Several possible adjustments were tested: Only using comple-

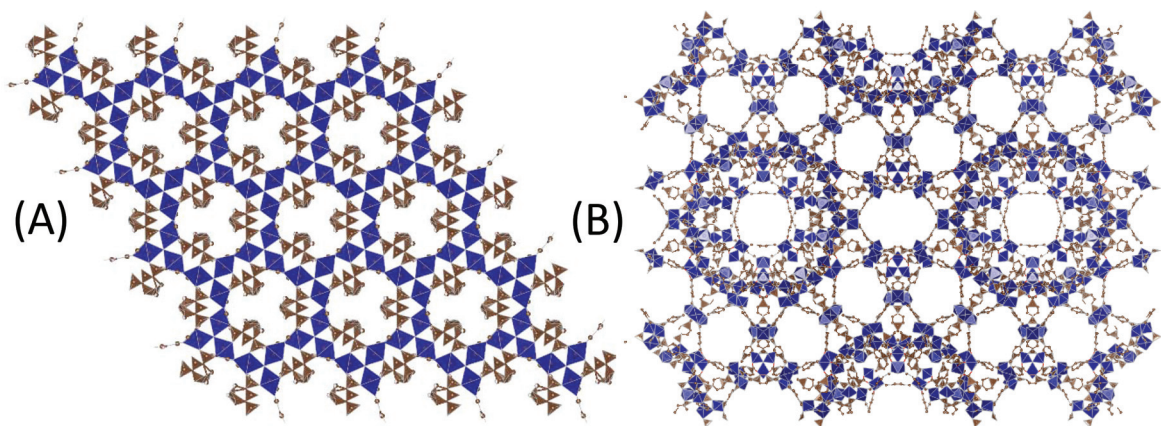


Fig. 1 Illustrations of (A) MOF-235 and (B) MIL-101 displaying their different porosities. Images were obtained using VESTA¹ with crystallographic information files from Sudik (2005)² and Lebedev (2005).³ Some atoms are omitted for clarity.



tely dry solvents (drying agent: molecular sieve silica, 3 Å beads), using newly purchased chemicals, replicating the synthesis several times, conducting the thermal treatment in other types of vessels and drying the product at lower and higher temperatures. None of the above-mentioned changes appeared to make a difference. Eventually, it was found that either lowering the synthesis temperature to 70 °C (Fig. S1†) or increasing the amount of DMF resulted in the formation of MOF-235.

To further investigate the factors affecting whether MOF-235 or MIL-101 is formed, different synthesis conditions were tested, *i.e.* different Fe(III):TPA ratios, different DMF:ethanol solvent ratios and different equilibration times (t_{eq}). The latter resembles the time between the solvothermal synthesis and the filtration of the product dispersion, which was found to affect both the product yield and phase. To simplify the whole procedure, all new syntheses were conducted in 100 mL glass bottles and a PTFE tape was used to keep the system sealed during the heat treatment. Before and after the synthesis, the pH of the solutions was measured using a combined pH electrode (Metrohm Unitrode easyClean). The details of the syntheses are shown in Table 1. The ns1.X syntheses used an equimolar ratio of Fe(III):TPA (1:1, 0.615:0.615 mmol), the ns 2.X syntheses used a stoichiometric ratio (4:3, 0.820:0.615 mmol) and the ns3.X syntheses used the ratio found in the literature to synthesize MOF-235^{2,17,32–35} (3:5, 0.369:0.615 mmol). In addition, performing the synthesis in a pure DMF solvent at the same temperature was tested, but this yielded no solid product within 24 h at 80 °C.

The product yields were calculated after the vacuum treatment with respect to TPA and the molar weights used were adjusted from the respective products' MOF-235:MIL-101 ratios obtained from the XRD analyses, as they differ (1093.16 g mol⁻¹ vs. 728.17 g mol⁻¹).

Table 1 The various synthesis conditions investigated in this study. $T = 80$ °C and $t = 24$ h for all the syntheses

Synthesis	Fe(III): TPA ratio	DMF : EtOH ratio	t_{eq} (days)
ns1.1	1.0	1 : 1	1
ns2.1	1.3	1 : 1	1
ns3.1	0.6	1 : 1	1
ns1.2	1.0	1 : 1	0
ns2.2	1.3	1 : 1	0
ns3.2	0.6	1 : 1	0
ns1.3	1.0	2 : 1	1
ns2.3	1.3	2 : 1	1
ns3.3	0.6	2 : 1	1
ns1.4	1.0	2 : 1	0
ns2.4	1.3	2 : 1	0
ns3.4	0.6	2 : 1	0
ns1.5	1.0	3 : 1	1
ns2.5	1.3	3 : 1	1
ns3.5	0.6	3 : 1	1
ns1.6	1.0	3 : 1	0
ns2.6	1.3	3 : 1	0
ns3.6	0.6	3 : 1	0
ns1.7	1.0	1 : 1	2
ns2.7	1.3	1 : 1	2
ns3.7	0.6	1 : 1	2

Characterization techniques

In situ MAXS. Due to the inconclusive results from the analyses of the synthesis products, we decided to further investigate the crystallization process of MOF-235 and MIL-101 at three different temperatures: 80, 85 and 90 °C. Time-resolved *in situ* MAXS was performed on the ns2.2 synthesis protocol (a 1:1 DMF:ethanol ratio and stoichiometric 4:3 Fe(III):TPA ratio). Every 15 minutes, the sample was scanned in the q range of 0–10 Å⁻¹. A small volume of the reagent solution was added to a narrow glass capillary which was sealed with glue. During the analysis, the temperature was kept constant, and heating was carried out only with a small volume where the X-ray beam passed through. The diameter of the beam was 1 μm. A reference sample, containing only the solvent, was analysed and the resulting spectrum was used for background correction.

A reoccurring problem was phase separation in the glass capillary holding the reaction mixture (Fig. S2(A)†); we had to change the experimental setup for the subsequent analyses to obtain scattering data. Instead of focusing the X-ray beam on the middle of the capillary, it was now focused on the bottom (Fig. S2(B) and (C)†).

Ex situ XRD. For *ex situ* XRD analyses, solvothermal syntheses under the same conditions as those of nsX.2 and ns X.6, *i.e.* at 1:1 and 3:1 DMF:EtOH ratios were performed. The starting volume of each synthesis was 150 mL and at times 6, 12 and 24 h, aliquots of 50 mL were taken from each batch. The aliquots were immediately filtered, washed and vacuum dried in the same way as regular solvothermal syntheses. The products were characterized with XRD and BET surface area measurements.

Dye adsorption. To compare the adsorption capacities of the synthesized MOFs, aqueous phase adsorption studies of methylene blue (MB) were performed at a solid mass concentration of 0.1 mg mL⁻¹ and a MB concentration of 625 μM (corresponding to 200 ppm). Two of the synthesis products were chosen by their apparent phase purity determined by XRD. The respective MOFs, ns2.7 and ns2.5, were dispersed in water ($V_{tot} = 10$ mL) and equilibrated for 1 h before MB was added and the adsorption took place in a shaking incubator (25 °C, 250 rpm) for 6 h. The solutions were centrifuged (5000 rpm, 10 min) and a part of the supernatant was diluted 100 times and analyzed by UV-Vis spectroscopy ($\lambda = 664$ nm). External standard solutions of MB were used to obtain a calibration curve which was used to quantify the amount of MB that was adsorbed (Fig. S3†).

Results and discussion

XRD

During the initial standard synthesis of MOF-235(Fe), XRD data showed that a pure MOF-235 product was difficult to obtain. Instead of obtaining a product having a single large peak at 9.4°, which is the main (1 0 1) peak of MOF-235, several additional peaks were found in the 5–10° region



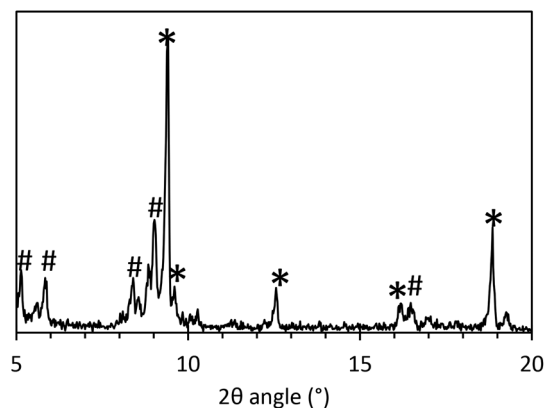


Fig. 2 XRD of the product of the original MOF-235 synthesis process performed in a 100 mL PTFE lined steel autoclave at 80 °C for 24 h. MOF-235 peaks are marked with * and MIL-101 peaks are marked with #.

(Fig. 2). A literature search revealed that these peaks belong to MIL-101(Fe),^{21,25,30} which has diffraction peaks at 2θ values of 5.1, 5.8, 8.4 and 9.0° in the 5–10° region, as well as a minor peak at 16.5°. Reference diffractograms obtained from crystallographic information files can be found in the ESI, Fig. S4 and 5.†

The new synthesis schemes yielded products containing both phases of MOF-235 and MIL-101 at varying degrees. For all 5–10° XRD scans, see Fig. S6 and 7.† To examine the influence of the solvent composition and reagent ratio, a bar chart of the relative XRD peak heights of the 9.0° (main peak of MIL-101 in this range) and 9.4° peaks is shown in Fig. 2. A dominantly black bar shows that mainly MIL-101 has formed, while a dominantly green bar means that mainly MOF-235 has formed. Note that the actual value of the ratio can seem to be misleading, as the MOF-235 peak is inherently stronger in intensity than the MIL-101 peak. The synthesis that gives the most pure MOF-235 appears to be the ns2.5 synthesis, where a 3 : 1 DMF : ethanol ratio and a stoichiometric 4 : 3 Fe(III) : TPA ratio were used with an equilibration time of 1 day. This also resulted in a product yield of 65% with respect to TPA (146 mg of the solid product after drying). With no equilibration time (ns2.6), a smaller amount of MOF-235 was formed with a lower yield (52%, 116 mg). It is clear that using a larger DMF : ethanol ratio yields a more pure MOF-235 product under the given synthesis conditions. In the 2 : 1 DMF : ethanol syntheses (nsX.3 and nsX.4), a longer t_{eq} seems to favor the MIL-101 formation. On the other hand, the 3 : 1 DMF : ethanol syntheses (nsX.5 and nsX.6) show the opposite trend, except for when a 3 : 5 Fe(III) : TPA ratio is used (ns3.5 and ns3.6).

A few studies have reported the formation of both MOF-235 and MIL-101 in the same system. Goesten *et al.*^{36–38} discovered that during the solvothermal synthesis of NH₂-MIL-101(Al), NH₂-MOF-235(Al) was initially formed. From density functional theory (DFT) modelling,³⁹ it was found that DMF has a stabilizing effect during the MOF-235 formation. NMR ana-

lyses by the same group revealed that DMF also promotes the transformation of NH₂-MOF-235(Al) to NH₂-MIL-101(Al). More specifically, by using chloride, DMF deprotonates one of the water molecules combined to Al³⁺ in MOF-235, facilitating its transformation to MIL-101 while creating a H-Cl-DMF complex. A large energy is required for this reaction (+272 kJ mol⁻¹), but the favorable ligand exchange in the Al³⁺ complex, *i.e.* Cl⁻ is exchanged for OH⁻, decreases the over-all reaction energy to +2 kJ mol⁻¹.³⁹ SAXS measurements³⁷ during the NH₂-MIL-101(Al) solvothermal synthesis in pure DMF at 130 °C showed that the MOF-235 phase starts to form after about 8 minutes and the MIL-101 phase is not seen until after 25 minutes. DFT and molecular dynamics have shown that DMF and other more hydrophobic solvents effectively solvate and stabilize the Al₃O(BDC)₆(DMF)₃⁺ clusters formed during the MOF-235(Al) crystallization process.⁴⁰

Looking at our results, it seems like the process of formation of MOF-235(Fe) and MIL-101(Fe) depends on several factors, not only on time. The most crucial factor appears to be the DMF : ethanol ratio of the solvent. In addition, the most phase-pure products (ns2.7 and ns2.5) were obtained using Fe (III) : TPA at a 4 : 3 ratio, indicating that an excess of iron ions benefits the formation of both phases. Another benefit was an over-all increased yield (Table 2), which could be an effect of pH and solubility and/or due to a higher concentration of the metal clusters which are required for MOF formation. It appears as if given enough thermal energy and/or time to equilibrate, the final product will be MIL-101. The role of ethanol in the heat treatment is unclear; no product was observed in the absence of ethanol at 80 °C and the conventional synthesis method of MIL-101 in pure DMF requires at least 110 °C. The type of solvothermal vessel also appears to dictate the

Table 2 Results from the BET surface area measurements of the ns1.X syntheses. The MOF-235 : MIL-101 ratios were calculated from the peak height ratios of the 9.4° and 9.0° XRD peaks. The yields were calculated with respect to TPA

Product	MOF-235 : MIL-101 ratio	S _{BET} (m ² g ⁻¹)	Yield (%)
ns1.1	0.54	2430	74
ns2.1	2.5	1080	80
ns3.1	2.0	285	39
ns1.2	0.98	1530	77
ns2.2	0.45	2320	84
ns3.2	0.26	1690	50
ns1.3	2.3	1030	52
ns2.3	1.8	1300	60
ns3.3	1.3	1180	43
ns1.4	9.2	313	44
ns2.4	3.4	736	44
ns3.4	2.8	937	32
ns1.5	7.1	502	58
ns2.5	26	295	67
ns3.5	1.1	1480	47
ns1.6	1.5	1160	49
ns2.6	6.8	420	55
ns3.6	3.8	530	30
ns1.7	2.5	37.3	73
ns2.7	0.35	2570	95
ns3.7	0.42	2170	48



MOF-235 : MIL-101 ratio (Fig. S8†); in the stainless-steel autoclave more pure MOF-235 was obtained (XRD peak ratio = 2.64 vs. 0.46) using the original synthesis procedure, *i.e.* ns3.2. This could be because the glass bottle has an overall higher heat-permeability than the autoclave, promoting the formation of MIL-101 over MOF-235.

The *ex situ* XRD experiments gave results which are not matching the regular solvothermal results (Fig. S9 and S10†). MOF-235 was only obtained in small amounts in two of the products after 24 h; ns1.6 (1 : 1 Fe(III) : TPA ratio) and ns2.6 (4 : 3 Fe(III) : TPA ratio). Both these products were acquired using a 3 : 1 DMF : EtOH solvent ratio. All other products only exhibited MIL-101 peaks. It is worth noting that the BET surface areas of the two products acquired after 24 h using a 1 : 1 DMF : EtOH ratio are some of the largest ones reported so far for MIL-101(Fe) (almost 3300 m² g⁻¹). One reason for the lack of MOF-235 in the products is that when taking aliquots from the reagent bottles, involving cooling down of the solution and gas and pressure release upon opening the bottles, the MOF formation process is hindered. Another reason could be the larger reagent volume (150 mL).

The influence of the equilibration time on the formation of MOFs has not previously been investigated, to our knowledge. In Fig. S11,† the relationship between t_{eq} and the surface area is displayed for the syntheses, showing that a longer t_{eq} does not always favor a larger surface area. In some cases, it appears as if the crystallization/transformation of the MOF continues at room temperature after the synthesis has been completed. In other cases, a longer t_{eq} appears to yield a less crystalline product. The bond between the iron(III) and oxygen containing ligands in these kinds of complexes becomes unstable at lower pH as the solubility of iron increases with decreasing pH.

From the XRD analyses at different temperatures (Fig. S36 and 37†), it was discovered that the crystal structure of MOF-235(Fe) is stable up to 350 °C, while MIL-101(Fe) decomposes at 300 °C. Thermogravimetric analyses by others have

indicated that MIL-101 exhibits a lower degree of thermal stability than MOF-235. While both start to decompose at 300 °C, the mass loss is 66.7% between 300 and 350 °C for MIL-101⁴¹ and 27.68% between 350 and 400 °C for MOF-235.¹⁸

The ambiguity of characterization of MOF-235 using only XRD and FTIR stems from several factors. One of the largest peaks of MIL-101 at 9.0° could be mistaken as noise and therefore becomes background subtracted and there is also the possibility of peak overlapping with the main peak of MOF-235 at 9.4°. Since the main diffraction peaks of MIL-101 lie below 5°, these peaks are easily missed. Fluorescence from iron when using a Cu K α radiation source can be significant (Fig. S12†), which further complicates the XRD analysis.⁴² Results from the XRD analyses using Cr K α radiation ($\lambda = 2.29$ Å), which eliminates the fluorescence signal, are shown in Fig. 4. Compared to the diffractograms with a Cu K α source (Fig. S13†), the signal-to-noise ratio is significantly improved (2.3 ± 0.41 for Cu vs. 6.5 ± 2.6 for Cr). The MOF-235 peaks at 2θ angles of 14.1° and 28.3° are especially prominent. The MIL-101 peaks at 7.8° and 13.6° are weaker than those with Cu K α radiation, yet visible. For example, from the S_{BET} of ns1.4 and ns1.1, it can be concluded that the latter contains a large amount of MIL-101. Despite this, their diffractograms look similar, apart from the weak MIL-101 peak at 13.6°, present in ns1.1, but not in ns1.4.

In situ MAXS

In the time-resolved MAXS analyses at 80 °C (Fig. 3(A)), only MOF-235 has formed. It takes approximately 4 hours for the major peak (9.4°) to appear and the intensity increases almost linearly during the entire analysis. No MIL-101 peaks are visible. At 85 °C (Fig. 3(B)), MIL-101 peaks (3.3° and 9.0°) appear after 2 h. The peaks increase in intensity during the entire analysis. At 90 °C (Fig. S14†) the MIL-101 peaks reach an intensity maximum at 11 h, before starting to decrease. No MOF-235 peaks are visible in the diffractograms for the syn-

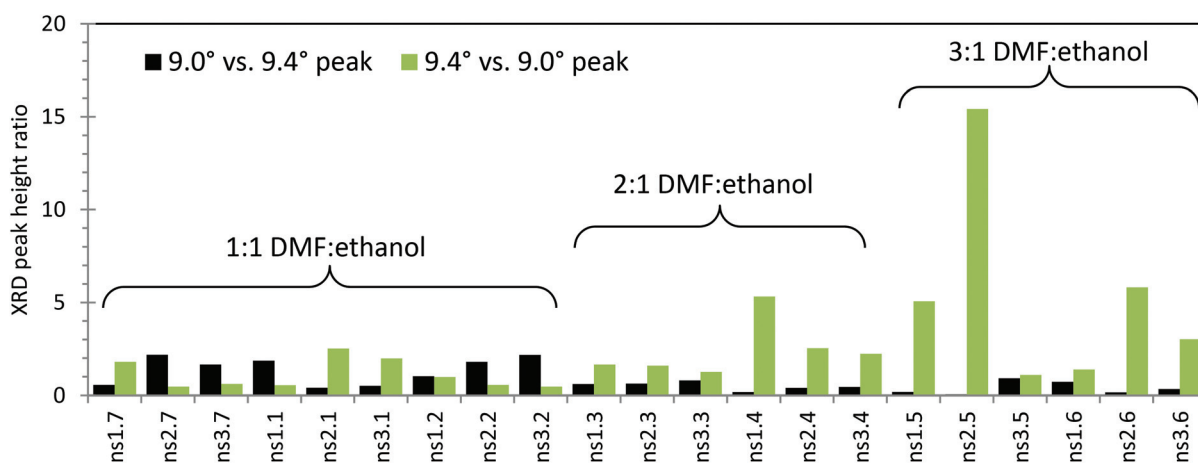


Fig. 3 XRD peak ratios of the synthesis products indicating their crystal phase purity. The black bars represent the ratio of the 9.0° peak, belonging to the MIL-101 phase, vs. the 9.4° peak, belonging to the MOF-235 phase. The green bars represent the ratio of the 9.4° peak vs. the 9.0° peak. The Fe(III) : TPA ratios of the different syntheses are as follows: ns1.X = 1 : 1; ns2.X = 4 : 3; ns3.X = 3 : 5. All the syntheses were performed at 80 °C for 24 h.



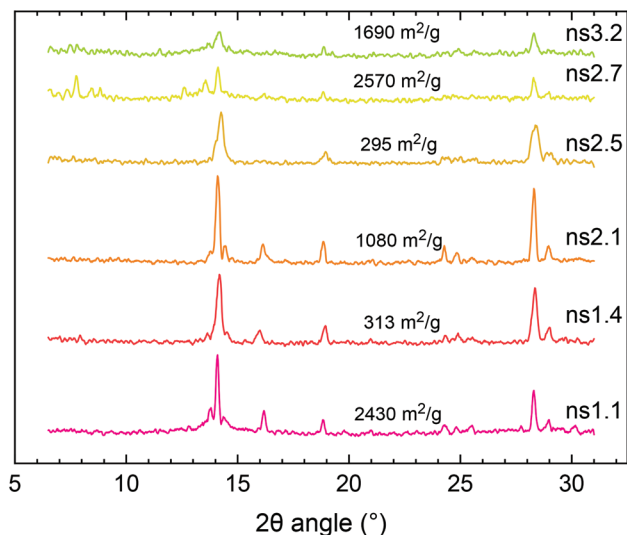


Fig. 4 XRD patterns using Cr K α radiation of the selected synthesis products in this study. Their respective BET surface areas are included in the plot.

thesis at 85 and 90 °C. These results, in combination with the *ex situ* XRD analyses, indicate that no transformation between MOF-235(Fe) and MIL-101(Fe) occurs during the solvothermal process. Instead, both phases appear to be formed independently of each other.

BET surface area

The surface areas of the synthesis products in this study are displayed in Table 2 and the N₂ isotherms are shown in Fig. S27–30.† The difference in the pore structure between MOF-235 and MIL-101 makes their respective surface areas differ substantially. MOF-235 has quite narrow hexagonal one-dimensional channels, which limit the surface area. On the

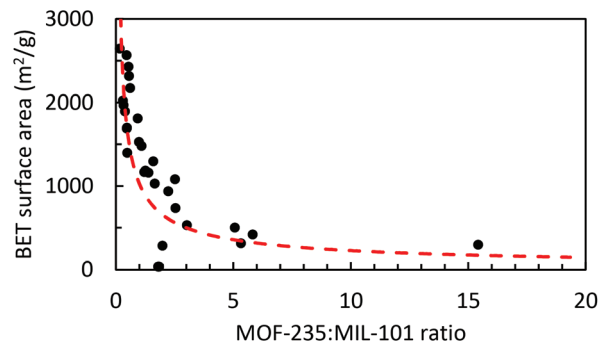


Fig. 6 Graph showing the determined BET surface areas of the synthesis products plotted against their MOF-235:MIL-101 XRD peak height ratios. A power trendline (the red dashed line) shows the relationship between the MOF-235/MIL-101 content and surface area.

other hand, the large nano-cages through which MIL-101 is made up of enable a substantially larger surface area. In addition, the MOF-235's channels are potentially more easily clogged, also affecting the measured effective surface area. The ns2.2 sample was synthesized exactly according to the literature yet exhibits a previously unreported large surface area in addition to the MIL-101 XRD peaks. Table 2 also shows the calculated yields with respect to TPA which are overall larger than earlier reports.

By utilizing the BJH desorption isotherm method, the total volume of the pore size range was obtained for ns2.2–24 h (MIL-101, Fig. S31†) and ns2.6 (MOF-235, Fig. S32†). Not surprisingly, the pore volume of MIL-101 (2.11 cm³ g⁻¹) is significantly larger than that of MOF-235 (0.169 cm³ g⁻¹).

The BET surface areas of the synthesized products are plotted against their respective MOF-235 : MIL-101 peak height ratios, as obtained by using the XRD peak height ratios, Fig. 6. The correlation between the MOF composition and surface

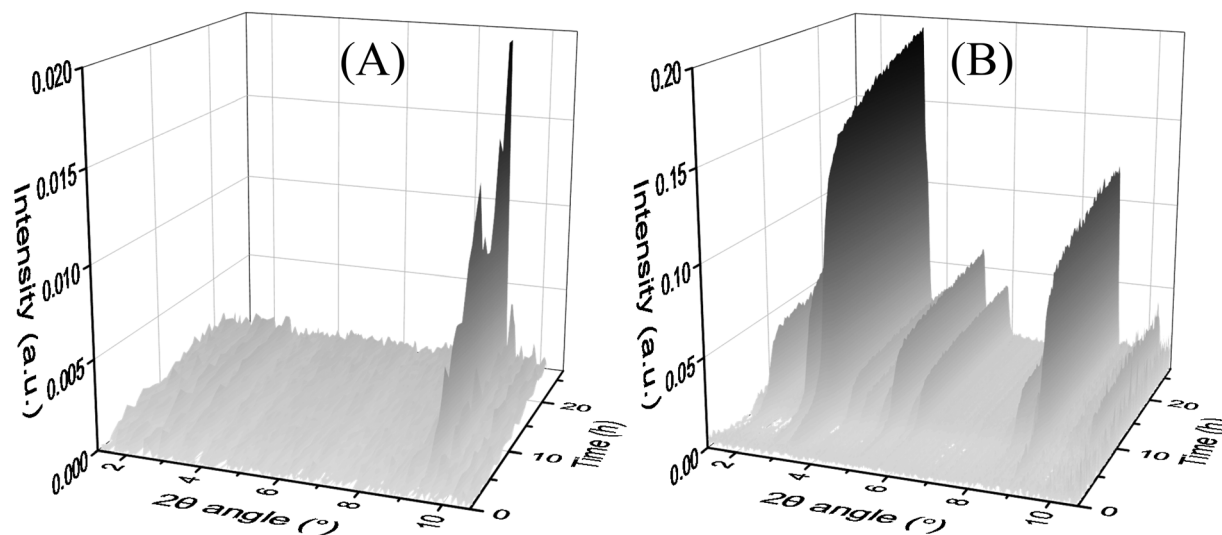


Fig. 5 Time-resolved *in situ* MAXS measurements of the solvothermal synthesis at (A) 80 °C and (B) 85 °C. A 1 : 1 DMF : ethanol ratio and a 4 : 3 Fe (iii) : TPA ratio were used in both syntheses.



area is clear; a higher MOF-235 : MIL-101 ratio yields a lower surface area. The product with the largest MOF-235 content (ns2.5) also has the lowest surface area ($295 \text{ m}^2 \text{ g}^{-1}$). Both ns2.7 and ns3.2 have the lowest MOF-235 content and show similar diffraction patterns (Fig. S15†), yet their surface areas are different (2570 and $1690 \text{ m}^2 \text{ g}^{-1}$). This could be due to a lower over-all amount of the crystalline product in the sample and/or the presence of impurities clogging the pores and decreasing the surface area. There is one clear outlier in Fig. 5, as the ns1.7 synthesis product's BET surface area was only $37.3 \text{ m}^2 \text{ g}^{-1}$, which is due to its low degree of crystallinity. It appears that solely using XRD to determine the MOF purity and quality is insufficient, but it does give an indication of the obtained product's surface area. Reference surface areas of MIL-101(Fe) are given in Table 4.

Several studies found in the literature have reported the synthesis MOF-235(Fe) under similar synthesis conditions. However, the measured surface areas differ to a great extent (Table 3). The reason for this could be that, in some of the high-surface area reference studies, MIL-101 has formed simultaneously with the formation of MOF-235, contributing to an increased porosity and surface area. It is difficult to account for the very low surface areas determined in some of the published studies; it could be due to clogging by the species not removed from the narrow pores of MOF-235. Most references activated the MOFs in a vacuum, which should eliminate any trapped solvent. Another possibility is the presence of precipitated terephthalic acid or iron oxide, which at low concentrations would not be detectable by XRD/FTIR/XPS measurements. Insufficient washing with DMF and ethanol of the sample before drying could cause this to occur. In addition, the drying procedure can cause agglomeration and damage to the pore structure.⁴⁴

Another factor affecting the surface area and crystallinity is the material's stability over time. The SEM images taken a year

Table 3 Literature values of the BET surface areas of synthesized MOF-235 and the corresponding Fe(III) : TPA and DMF : ethanol ratios used in the synthesis. A 1 : 1 DMF : ethanol ratio was used in all the references. V: vacuum oven

Reference	Fe(III) : TPA	Activation	SA _{BET} ($\text{m}^2 \text{ g}^{-1}$)
Haque, 2011 ³²	0.60	V: 100 °C ^c	24.5 ^d
Sudik, 2005 ²	0.60	N/A	N/A
Duo, 2019 ¹⁷	0.60	V	9.64
Li, 2016 ^{33a}	0.60	V: 150 °C, 5 h	148
Li, 2016 ^{33b}	0.60	V: 150 °C, 5 h	135
Anbia, 2012 ³⁴	0.60	V ^c	147 ^e
Le, 2016 ¹⁸	1.00	V: 150 °C, 3 h	540
De Smedt, 2015 ³⁵	0.60	V: 150 °C	8.09 ^d
Wang, 2017 ⁴³	0.60	V	918 ^f

^a Microwave-assisted synthesis. ^b Solvothermal synthesis. ^c Not explicitly stated, but their products were vacuum oven dried for other purposes. ^d Only the N₂ adsorption isotherms were published, the given surface areas were calculated by us. ^e The calculation from the N₂ isotherm to the BET surface area in the article was incorrect. The re-calculated value obtained using the data from the published isotherm is $147 \text{ m}^2 \text{ g}^{-1}$. ^f FTIR was the only characterization method, *i.e.* no XRD data were presented in the reference.

Table 4 References, synthesis conditions and determined BET surface areas of MIL-101(Fe). All the syntheses were performed in a DMF solvent at 110 °C, unless stated otherwise

Reference	Fe(III) : TPA	<i>t</i> (h)	SA _{BET} ($\text{m}^2 \text{ g}^{-1}$)
Taylor-Pashow, 2009 ^{30a}	1.0	0.17	2500 ^c
Maksimchuk, 2012 ²⁴	2.0	20	3200
Wu, 2013 ^{48b}	1.3	24	N/A
Barbosa, 2017 ²⁵	2.0	20	N/A
Xie, 2017 ²⁰	2.0	20	2350
Simsek, 2017 ²³	2.0	24	N/A
He, 2019 ²⁷	1.0	20	601.3
Gecgel, 2019 ²⁸	2.0	20	2865
Li, 2019 ²¹	2.0	20	1172.5

^a Microwave-assisted synthesis, $T = 150$ °C. ^b Using amino-terephthalic acid as a linker molecule, $T = 150$ °C. ^c Only the Langmuir surface area was given in the article, the BET surface area was calculated from the presented nitrogen adsorption isotherm.

after the synthesis revealed small amounts of sheet-like impurities in the ns2.5 sample (Fig. 9(A) and Fig. S35†), likely to be recrystallized TPA.⁴⁵ BET analysis confirmed that the SA has decreased from 295 to $9.2 \text{ m}^2 \text{ g}^{-1}$ (Fig. S33†).

XPS analyses

The chemical composition of the MOF samples *via* XPS measurements, on average, consists of 58 at% of carbon (C), 28 at% of oxygen(O), 11 at% of iron (Fe), <2 at% of chloride (Cl) and <1.0 at% of nitrogen (N) (Fig. S16†). The XPS spectra of the core level of iron, oxygen and carbon in the pure MOF-235 sample (ns2.5) are displayed in Fig. 8. In the core Fe 2p line, the binding energies of the spin-doublet Fe 2p_{3/2} and Fe 2p_{1/2} are located at 712.0 eV and 725.5 eV with the energy separation of 13.5 eV. These positions are higher than that in ferric oxide (Fe₂O₃, hematite), *i.e.* 711.0 eV.⁴⁶ This shift is due to the polarization effect on the iron by the surrounding ligands which causes the de-shielding of the inner orbitals, *i.e.* the electron density decreases and the corresponding core electron-nucleus interaction increases. The associated shake-up satellites are strong (*i.e.* each area is 45% of the corresponding main peak) and +5.9 eV separated from each main line. The characteristic features of the spin-doublet refer to a ferric state (Fe³⁺) and those of the satellites refer to strong metal-to-ligand charge transfer and strong covalent bonding characteristics. Therefore, a paramagnetic ferric (Fe³⁺) ion complex is determined in both the MIL-101 and MOF-235 structures.⁴⁷

Carbon (Fig. 8(C)) in the form of C–C bonds (284.9 eV) and the carboxylic group (288.9 eV) are stemming from terephthalic acid (TPA). The N 1s signal (Fig. 7(B)) centered at 400.2 eV shows the presence of amine, most likely contributed from the DMF solvent. In the core O 1s and Cl 2p lines (Fig. 8(B) and Fig. 7(C) and (D)), there are Fe–O bonds (O 1s = 530.4 eV) and O–H bonding (O 1s = 533.1 eV) as well as chloride (Cl 2p_{3/2} = 198.1 eV). The large ratio of the C–O vs. C=O peak area ($66.7/10.5 \approx 6.4$) indicates that the C...O– bond in the carboxylic



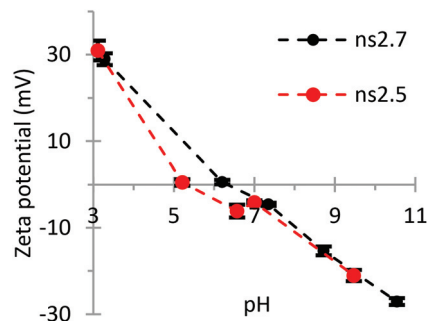


Fig. 10 Zeta potentials at different pH values of ns2.7 (MIL-101) and ns2.5 (MOF-235). Each sample was analyzed six times and the error bars represent the calculated standard deviations.

5.0 to 220 mg g⁻¹ at pH 10.5 ($c_0 = 30$ ppm). Considering the molecular size of MB ($17 \times 7.6 \times 3.3 \text{ \AA}^3$ (ref. 49)) and the previously reported channel diameter of MOF-235 (6.7 \AA^2), it is remarkable that MOF-235 exhibits such a high adsorption capacity. MIL-101(Cr) showed relatively low adsorption capacities for MB; 6.78 mg g^{-1} .⁵⁰ Doping MIL-101(Cr) with iron oxide (Fe_3O_4) resulted in a higher adsorption efficiency: 80 mg g^{-1} ($c_0 = 60$ ppm).⁵¹ This indicates that the presence of iron(III) enhances the adsorption of MB in water. Chromium (III) is considered a hard Lewis acid in HSAB (hard and soft acids and bases) theory, while iron(III) is considered as an intermediate one. This could explain the difference in their adsorption capacities. However, amino-functionalized MIL-101 (Al) showed an even greater adsorption capacity of 762 mg g^{-1} ($c_0 = 40$ ppm, $30 \text{ }^\circ\text{C}$, 0.05 mg mL^{-1} solid concentration).⁵² Since Al^{3+} is a hard acid, the enhanced adsorption could be due to some other factors.

The zeta potential measurements in Fig. 10 show that there is little difference between the surface charge behaviors of MOF-235 and MIL-101. The ZP measurements of the particles carrying little or no charge, *i.e.* when pH approaches the isoelectric point (pH_{IEP}), are often inaccurate as a low particle charge inhibits the particle movement during the analysis. This could explain the small difference in the ZP of the two materials around their IEP.

Conclusions

Despite obvious differences in the reported surface areas of MOF-235 in the literature, no-one has previously investigated the possibility of the simultaneous formation of MOF-235(Fe) and MIL-101(Fe) during the synthesis of MOF-235. We have shown, using extensive experimental methods such as XRD and *in situ* MAXS, that it appears to be difficult to obtain a pure phase of both MOFs. The formation of the respective phase depends on several factors. A larger DMF solvent fraction and lower temperature appear to favor MOF-235. Using a non-insulated vessel like a glass bottle facilitates more MIL-101, the more thermodynamically favored phase.

To our knowledge, no one before has used non-synchrotron X-ray scattering for the *in situ* investigations of the crystallization of MOFs. In this study we have shown that the acquisition of time-resolved data is possible with an in-house Mat:Nordic device. *In situ* MAXS at different temperatures (80 , 85 , and $90 \text{ }^\circ\text{C}$) showed no phase transformation during the solvothermal synthesis. At $80 \text{ }^\circ\text{C}$ only MOF-235 was formed, while the syntheses at 85 and $90 \text{ }^\circ\text{C}$ yielded only MIL-101.

With the combination of the afore-mentioned techniques, we concluded that the primary decisive factors through which the iron(III) TPA MOF product is formed are the synthesis temperature and DMF:ethanol solvent ratio. The Fe(III):TPA reagent ratio also plays a role in the synthesis; a higher ratio gives a larger yield and better crystallinity. More specifically, a near-pure MIL-101 phase with a BET surface area of $2600 \text{ m}^2 \text{ g}^{-1}$ and a yield of 95% was obtained by using a 4:3 ratio of Fe(III):TPA, a 1:1 ratio of DMF:ethanol, and an equilibration time of 2 days (“ns2.7”). To obtain MOF-235, the DMF solvent content has to be increased, resulting in a product with a lower degree of porosity and subsequently a lower surface area: $295 \text{ m}^2 \text{ g}^{-1}$, in a 67% yield (“ns2.5”). XPS analysis revealed that MIL-101 contained smaller amounts of nitrogen and chlorine than MOF-235. Both products showed enhanced adsorption efficiencies of methylene blue: 425 and 276 mg g^{-1} for MIL-101 and MOF-235, respectively.

To our knowledge, the synthesis of MIL-101 has never previously been performed in an ethanol:DMF solvent mixture at a temperature as low as $80 \text{ }^\circ\text{C}$. The possibility of creating a high-surface area MOF with a smaller amount of DMF and at a lower temperature provides further applicability to these kinds of materials.

Conflicts of interest

There are no conflicts to declare.

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