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CORRECTION

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Correction: Stereoelectronic power of oxygen in control of chemical reactivity: the anomeric effect is not alone

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Correction for 'Stereoelectronic power of oxygen in control of chemical reactivity: the anomeric effect is not alone' by Igor V. Alabugin *et al.*, *Chem. Soc. Rev.*, 2021, DOI: 10.1039/d1cs00386k.

The authors regret that Fig. 4 and 116 contained errors in the labeling of the acceptor orbitals of the carboxylic acids. The corrected versions of these figures are presented here.

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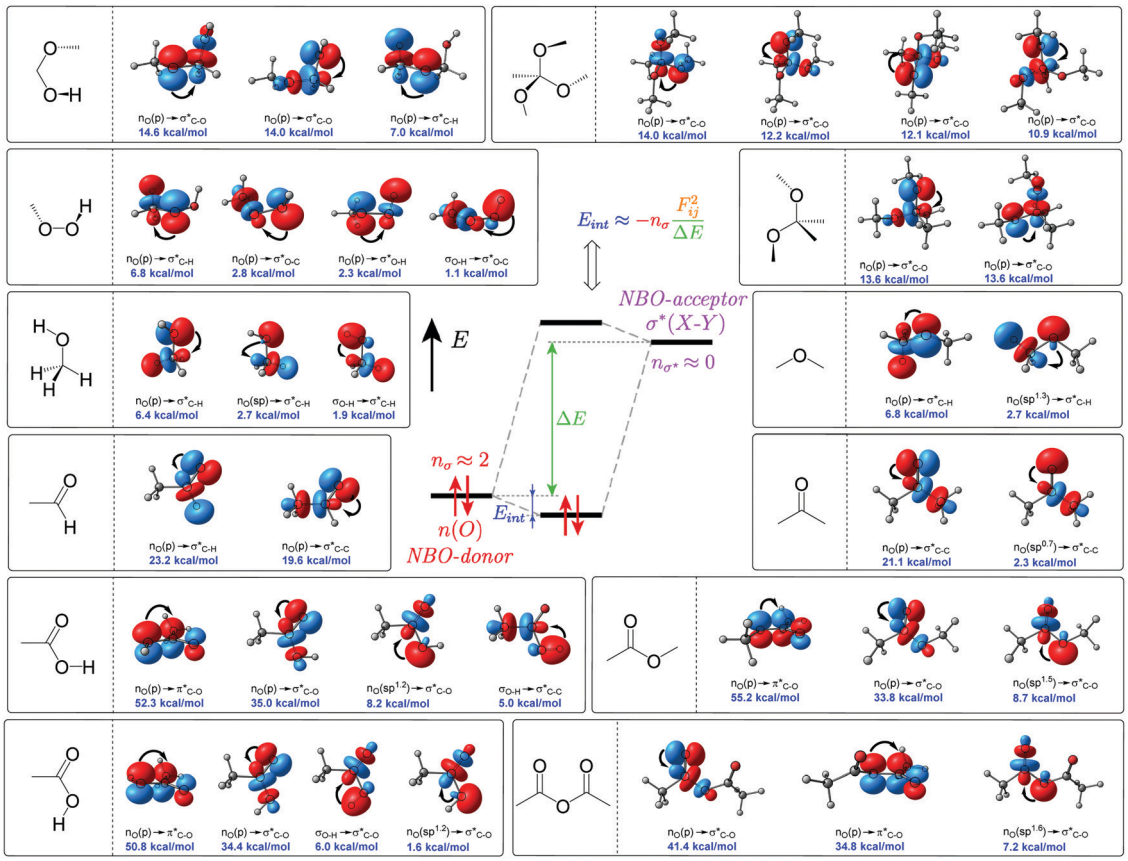

Fig. 4 A selection of donor–acceptor anomeric interactions in organic O-functionalities. E_{int} values are computed for the Natural Bond Orbital method (NBO) using the second-order perturbation approximation for the PBE0-D3BJ/aug-cc-pVTZ/CPCM(H_2O) level of theory. The orbital interaction diagram at the center of the scheme shows the equation used to calculate the interaction energy—E_{int}—and its contributors: F_{ij} is the Fock matrix element between natural bond orbitals i and j, ε_σ and ε_{σ^*} are the energies of the σ and σ* NBO's, and n_σ is the population of the donor orbital.[27]

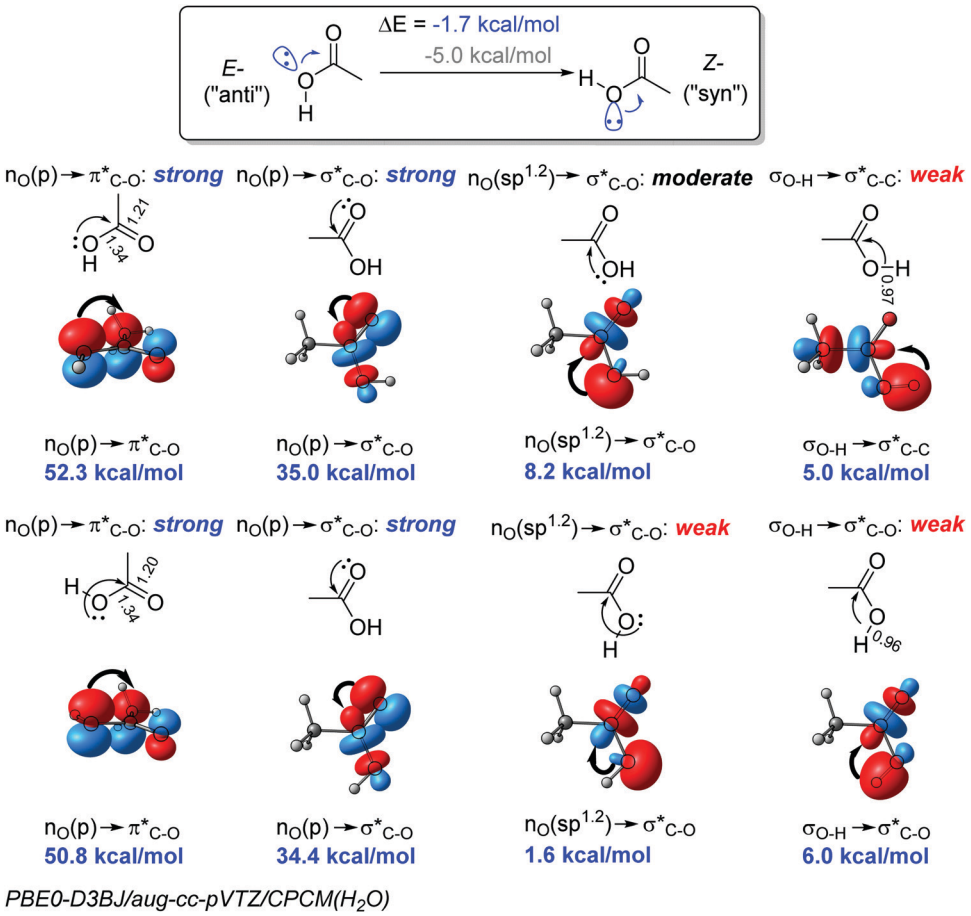

Fig. 116 Stereoelectronic "portrait" and conformational preferences of acetic acid.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

