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CORRECTION

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Correction: Stereoelectronic power of oxygen in control of chemical reactivity: the anomeric effect is not alone

Igor V. Alabugin,*[a] Leah Kuhn,[a] Michael G. Medvedev,[bc] Nikolai V. Krivoshchapov,[bd] Vera A. Vil',[b] Ivan A. Yaremenko,[b] Patricia Mehaffy,[a] Meysam Yarie,[e] Alexander O. Terent'ev[b] and Mohammad Ali Zolfigol[e]

Correction for 'Stereoelectronic power of oxygen in control of chemical reactivity: the anomeric effect is not alone' by Igor V. Alabugin *et al.*, *Chem. Soc. Rev.*, 2021, DOI: 10.1039/d1cs00386k.

The authors regret that Fig. 4 and 116 contained errors in the labeling of the acceptor orbitals of the carboxylic acids. The corrected versions of these figures are presented here.

[a] *Department of Chemistry and Biochemistry, Florida State University, Tallahassee, FL 32306, USA. E-mail: alabugin@chem.fsu.edu*
[b] *N. D. Zelinsky Institute of Organic Chemistry, Russian Academy of Sciences, 47 Leninsky prosp., 119991 Moscow, Russian Federation*
[c] *A. N. Nesmeyanov Institute of Organoelement Compounds, Russian Academy of Sciences, 28 Vavilova St., 119991 Moscow, Russian Federation*
[d] *Lomonosov Moscow State University, Leninskie Gory 1 (3), Moscow, 119991, Russian Federation*
[e] *Department of Organic Chemistry, Faculty of Chemistry, Bu-Ali Sina University, Hamedan 65167, Iran*

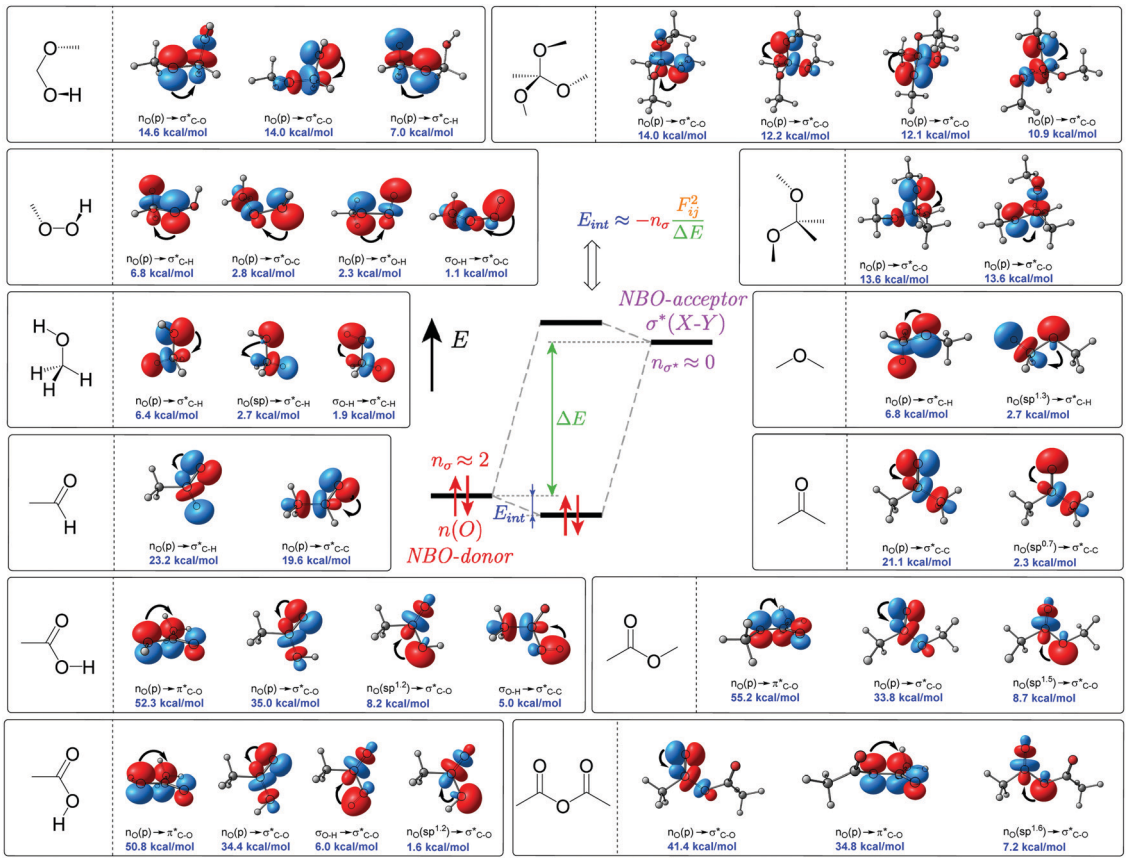

Fig. 4 A selection of donor–acceptor anomeric interactions in organic O-functionalities. E_{int} values are computed for the Natural Bond Orbital method (NBO) using the second-order perturbation approximation for the PBE0-D3BJ/aug-cc-pVTZ/CPCM(H_2O) level of theory. The orbital interaction diagram at the center of the scheme shows the equation used to calculate the interaction energy—E_{int}—and its contributors: F_{ij} is the Fock matrix element between natural bond orbitals i and j, ε_σ and $\varepsilon_{\sigma*}$ are the energies of the σ and σ* NBO's, and n_σ is the population of the donor orbital.[27]

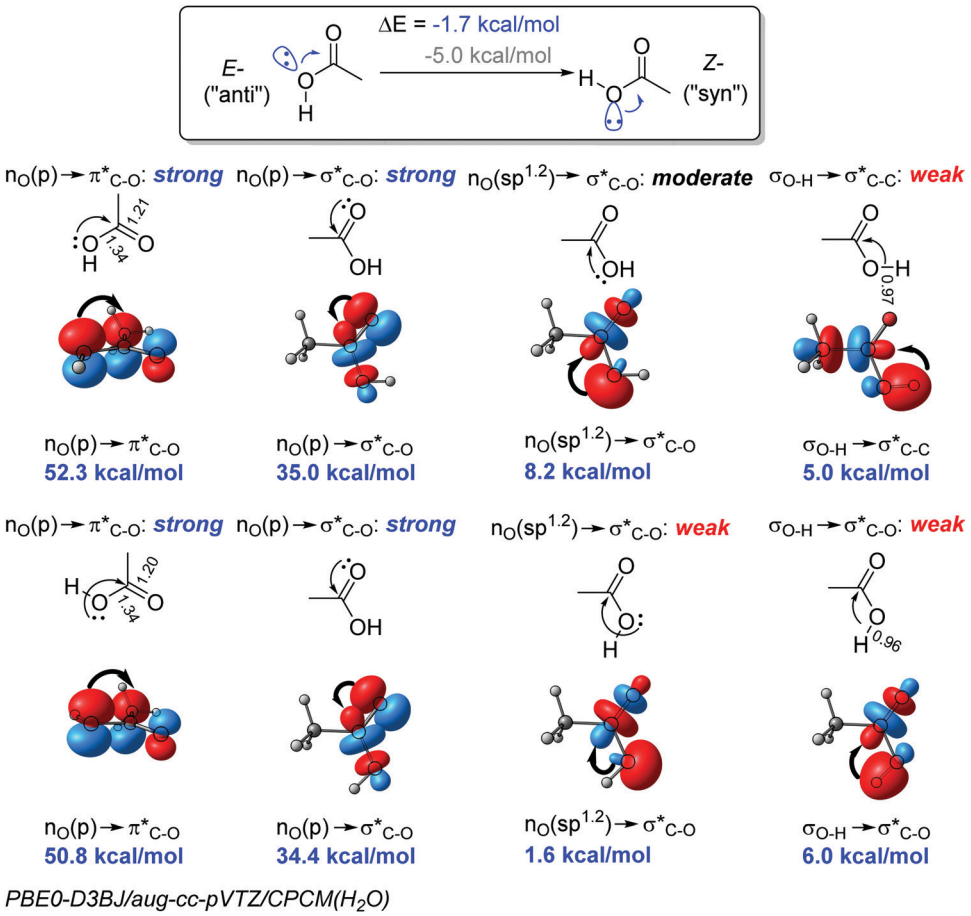

Fig. 116 Stereoelectronic "portrait" and conformational preferences of acetic acid.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

