



Cite this: *Chem. Soc. Rev.*, 2021, 50, 6684

A reaction-coordinate perspective of magnetic relaxation

Cassidy E. Jackson, Ian P. Moseley,  Roxanna Martinez, Siyoung Sung and Joseph M. Zadrozny *

Understanding and utilizing the dynamic quantum properties of metal ions is the frontier of many next generation technologies. One property in particular, magnetic relaxation, is a complicated physical phenomenon that is scarcely treated in undergraduate coursework. Consequently, principles of magnetic relaxation are nearly impenetrable to starting synthetic chemists, who ultimately design the molecules that fuel new discoveries. In this Tutorial Review, we describe a new paradigm for thinking of magnetic relaxation in metal complexes in terms of a simple reaction-coordinate diagram to facilitate access to the field. We cover the main mechanisms of both spin–lattice (T_1) and spin–spin (T_2) relaxation times within this conceptual framework and how molecular and environmental design affects these times. Ultimately, we show that many of the scientific methods used by inorganic chemists to study and manipulate reactivity are also useful for understanding and controlling magnetic relaxation. We also describe the cutting edge of magnetic relaxation within this paradigm.

Received 1st January 2021

DOI: 10.1039/d1cs00001b

rsc.li/chem-soc-rev

Key learning points

- (1) Magnetic relaxation is the process of a molecular magnetic moment flipping orientation while the spin system (either a single molecule or a bulk sample) returns to equilibrium.
- (2) Magnetic relaxation mechanisms are analogous to reaction pathways: they proceed from a high-energy “starting material” to a lower energy “product,” where the start and end points are different orientations.
- (3) The mechanisms of magnetic relaxation are controlled by varying chemical composition and structure.
- (4) The mechanisms of magnetic relaxation are also governed by extrinsic factors, such as applied magnetic field, temperature, and local environment.
- (5) The cutting edge of molecular magnetic relaxation research is designing species where relaxation processes are as slow as possible.

Introduction

Magnetic molecules are centerpiece components of numerous areas of research, ranging from quantum information processing^{1–3} and classical data storage^{1,4} to magnetic resonance imaging (Fig. 1)^{5–7} and spin-controlled reactions.⁸ Of these, open-shell transition metal complexes are particularly popular, with proposed applications as molecular qubits,^{2,3,9} spin-crossover sensors,^{10,11} single-molecule magnets,^{12,13} and molecular spintronic materials.¹⁴ Organic radicals are also studied for their biomedical imaging applications, polymer applications, and use in pharmaceuticals.^{15,16} Molecular compounds such as these two classes are notably advantageous for these applications because they offer a blank slate to design magnetic properties by harnessing synthetic chemistry.

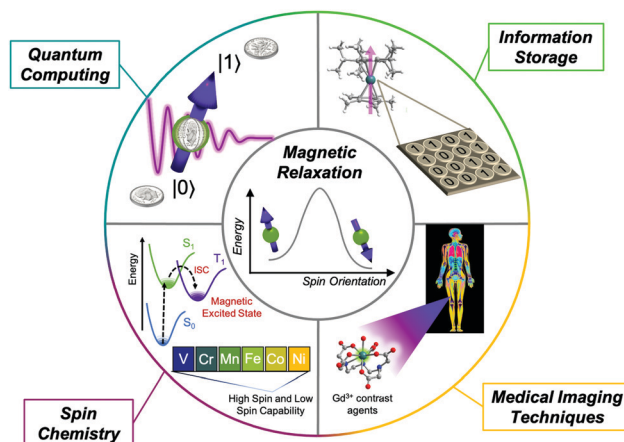


Fig. 1 Overview of areas where understanding and designing magnetic relaxation in molecules is important to modern and future technological developments.

Department of Chemistry, Colorado State University, Fort Collins, CO, USA.
 E-mail: joe.zadrozny@colostate.edu



A specific magnetic property that is important in the foregoing applications is magnetic relaxation (sometimes called spin relaxation), or the response of the magnetic moment after misalignment from an applied magnetic field (from an MRI scanner or nuclear magnetic resonance spectrometer, for example). A slow relaxation rate permits many important and exciting possibilities, *e.g.* the storage of information (quantum or classical) in the orientation of the magnetic molecule (here a spin-up orientation could be “0” *versus* spin-down orientation of “1” in analogy to a simple bit). Slow relaxation rates will also enable certain magnetic resonance spectroscopic experiments (*e.g.* to noninvasively detect local chemistry) while fast relaxation can make these measurements more challenging. Therefore, there is a clear necessity to understand how to use synthetic design to control and slow down magnetic relaxation – an end goal that many groups are still in pursuit of.

Magnetic relaxation is driven by interaction of the relaxing species with local electronic and magnetic fields. These fields

are influenced by intrinsic properties of the molecule. For example, metal-ion identity, spin state, oxidation state, ligand field, and geometry can all affect the relaxation rate of an open-shell metal ion. Extrinsic properties, such as counterion, concentration, temperature, and matrix (*e.g.* solvent or local chemical surroundings), are also important. Thus, there is considerable overlap between the molecular factors that control magnetic relaxation and those that dictate commonly approached properties by chemists, like reactivity.

Despite this conceptual overlap and the importance of dynamic magnetic properties in many cutting-edge fields, education in magnetic relaxation is typically absent from undergraduate chemical curricula.^{17–19} As a consequence, concepts in magnetic relaxation can be intimidating for chemistry researchers to utilize in their research. Yet it is precisely these chemists that are needed to make the molecules that drive the observations of new processes, developments of new theories,



Cassidy E. Jackson

Cassidy E. Jackson received her BS in Chemistry from James Madison University where she worked under Dr Donna Amenta and Dr John Gilje on the synthesis and characterization of porous materials. She is now a PhD candidate in the Department of Chemistry at Colorado State University working in the group of Dr Joseph Zadrozny. Her graduate research in molecular magnetism aims to understand how specific placements, arrangements, and patterns of nuclei surrounding an unpaired electron on a molecule impact magnetic relaxation. She is passionate about advocating for more accessible science literature. Her favorite element is silver due to its historical significance.



Ian P. Moseley

Ian Moseley received his undergraduate BS degree from the University of North Carolina at Chapel Hill in 2017. He then moved to Colorado to pursue his PhD in Inorganic Chemistry at Colorado State University. His research is focused on spin bath dynamics and understanding their impact on magnetic relaxation processes. If he had to choose, his favorite element would be boron – as a nonmetal sitting above the all-metallic group 13, it is found in everything from glassware to semiconductors.

Ian Moseley received his undergraduate BS degree from the University of North Carolina at Chapel Hill in 2017. He then moved to Colorado to pursue his PhD in Inorganic Chemistry at Colorado State University. His research is focused on spin bath dynamics and understanding their impact on magnetic relaxation processes. If he had to choose, his favorite element would be boron – as a nonmetal sitting above the all-metallic group 13, it is found in everything from glassware to semiconductors.



Roxanna Martinez

Roxanna Martinez received her BA in Chemistry from Skidmore college in 2019. She is currently pursuing her PhD in Chemistry at Colorado State University. Her research focuses on understanding the dynamics of nuclear spins in magnetic environments to develop fundamental design principles for molecular qubit systems. Her favorite element is bismuth due to its iridescence when exposed to oxygen and its use as an environmentally safe replacement for lead-based materials.



Siyoung Sung

Siyoung Sung obtained his Master's degree in Chemistry from Seoul National University in Seoul, South Korea. He then moved to US and completed his PhD in Chemistry at Texas A&M University, where he studied homogeneous electrocatalytic systems for CO₂ reduction. In 2020, he joined the Zadrozny group as a postdoctoral fellow to investigate synthetic control of spin properties in inorganic complexes. Siyoung's favorite element is 'Mn' because its complexes generate a fantastic 6-line pattern in EPR spectra.

Siyoung Sung obtained his Master's degree in Chemistry from Seoul National University in Seoul, South Korea. He then moved to US and completed his PhD in Chemistry at Texas A&M University, where he studied homogeneous electrocatalytic systems for CO₂ reduction. In 2020, he joined the Zadrozny group as a postdoctoral fellow to investigate synthetic control of spin properties in inorganic complexes. Siyoung's favorite element is 'Mn' because its complexes generate a fantastic 6-line pattern in EPR spectra.



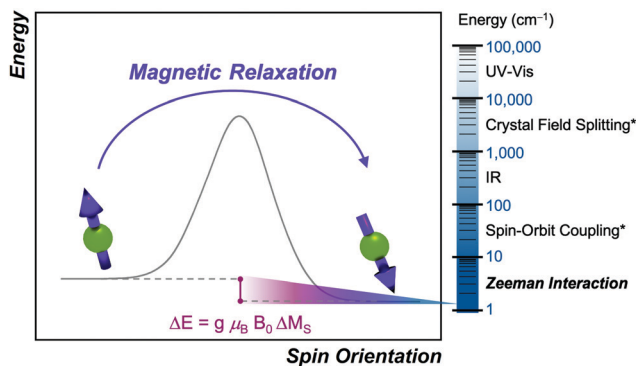


Fig. 2 Magnetic relaxation in analogy to a standard reaction-coordinate diagram. Here, the reaction coordinate corresponds to spin orientation, with the products aligned parallel to the applied field and the starting materials aligned against it. The energy difference (ΔE) between the starting materials and products is the Zeeman energy, as defined in the plot. The energy scaling on the right compares the Zeeman interaction to other important energies. Scaling in the reaction coordinate diagram is arbitrary. Asterisks denote energy scaling for 3d transition metals.

and surmount the diabolical challenges to designing slow magnetic relaxation processes. This Tutorial Review aims to assist researchers unfamiliar with magnetic relaxation and provide them with an overview of the field in an accessible way. To do so, we draw a direct analogy between the complex quantum mechanical processes of spin relaxation and one of the most intuitive phenomena to the synthetic chemist: the reaction-coordinate diagram (Fig. 2). We focus primarily on metal complexes, though many of the concepts discussed here can be applied to organic radicals as well.

A reaction-coordinate picture of relaxation

Magnetic relaxation is the process of a magnetic moment (or a spin) coming to alignment with an applied magnetic field from some other orientation. This misaligned orientation is an



Joseph M. Zadrozny

Joseph (Joe) Zadrozny is an assistant professor of chemistry at Colorado State University. Joe got his BS in chemistry from Virginia Tech, then moved to UC Berkeley for his graduate studies. His subsequent postdoctoral studies proceeded at Northwestern University. He now leads a group focused on magnetism, inorganic chemistry, and the many exciting intersections between these fields. Joe's favorite element is iron because it is solid, reliable, and magnetic, like a good friend.

excited state while the ground state is when the spin is aligned with the field. For this reason, magnetic relaxation can be thought of analogous to an exergonic reaction that proceeds from relatively high energy reactants (the misaligned orientation) to lower-energy products (the aligned orientation). In the case of a single unpaired electron in an external magnetic field undergoing relaxation (Fig. 2), the starting materials could be “spin up”, corresponding to the spin quantum number $m_s = +\frac{1}{2}$. Likewise,

the products could be “spin down”, bearing the $m_s = -\frac{1}{2}$ spin quantum number. The relaxation process is then the “reaction” that enables the spin to convert orientation.

There are some important similarities between the relaxation process and a chemical reaction. Just like many reactions, the process of magnetic relaxation is thermodynamically favored.¹⁵ There are also often many possible mechanisms by which the system can proceed with relaxation, similar to how many different reaction pathways can exist in a catalytic system. Though there are many of these pathways available to a molecule, again just like with a reaction system, it is the fastest relaxation process that typically proceeds under a given set of conditions. Several relaxation processes have an activation energy, directly analogous to a chemical reaction. However, other relaxation mechanisms can proceed by tunneling through the activation barrier, just like proton-tunneling reactions that can be interrogated through kinetic isotope studies.²⁰ Owing to all the foregoing similarities, it should be no surprise that elucidating the operative magnetic relaxation processes is just as rich and intellectually rewarding as mechanistic investigations of reactions.

There are also some critical differences between magnetic relaxation and a typical chemical reaction. First, the energy difference separating the starting materials and products is extremely small, typically on the order of a few wavenumbers or less so the relaxation process is nearly thermoneutral. This small energy difference is because the energy separating the spin orientations in a typical applied magnetic field (from the “Zeeman” interaction, Fig. 2) is weak. This situation is in stark contrast to the multiple-kcal-energy magnitudes ($1 \text{ kcal mol}^{-1} = 350 \text{ cm}^{-1}$) separating starting materials and products in a general chemical reaction. Second, the transition states of molecules in the “relaxation reaction” can correspond to high-energy spin orientations, which are discrete and quantized energy levels with m_s (or M_s if $S > \frac{1}{2}$) values, or other quantum phenomena entirely. This point highlights a key contrast to a chemical reaction, where a transition state is a transient, high-energy, and distorted molecular geometry. Third, because these transition states are quantized, a given relaxation process does not proceed along a continuous potential-energy curve like molecular transformations, but instead by discrete jumps with emission/absorptions of energy.

We also note that, just like a chemical reaction, the rate of relaxation reflects the slowest step in the operative pathway. Indeed, there are often competing mechanisms, yet the slowest step in the fastest pathway is most important for the observed rate.



In this light, the “rate-determining step” for magnetic relaxation is whether the environment can provide the energy to drive the process or accommodate the energy emitted during relaxation.²¹ This process is in contrast to reactions where rate determining steps involve structural transformations or proton/electron transfers, and possibly isolable chemical intermediates.

The figure of merit for magnetic relaxation is the time constant of the process, the magnetic relaxation time, which is the inverse of the relaxation rate. The timescale of a magnetic relaxation time is highly variable, typically ranging from picosecond to minute timescales. There are two basic types of magnetic relaxation, spin–lattice relaxation, T_1 , and spin–spin relaxation, T_2 . Both of these parameters are vital for the applications in Fig. 1. In light of this importance, it is essential that we understand how to control the processes that govern these relaxation times and how they correlate to molecular structure. Below we describe the different relaxation mechanisms that govern these two relaxation times, again in analogy to the basic chemical reaction-coordinate paradigm described above.

Basic molecular spin properties relevant to relaxation

In evaluating the reactivity of a metal complex used in a reaction or as a catalyst, common considerations might be steric congestion or electron abundance/deficiency in the ligand scaffold. Similarly, for coarse-grain prediction of a magnetic relaxation property, there are several key magnetic parameters that are important to consider as a starting point. These parameters are the g factor, hyperfine coupling interactions, and zero-field splitting (Fig. 3). Comprehensive treatments of all the magnetic properties in metal complexes, which are diverse, can be found elsewhere.¹⁵

The first magnetic factor is the electron gyromagnetic ratio, or g factor. Electrons possess an intrinsic angular momentum with a magnetic moment that interacts with an applied magnetic field. The strength (or energy) of this interaction is proportional to the applied magnetic field (B_0), the Bohr magneton (μ_B , a fundamental constant), the m_S value of the given spin orientation, and the g factor: $E = g\mu_B B_0 m_S$. The part relevant to relaxation is the g factor, a proportionality constant that describes the sensitivity of the magnetic moment to an external magnetic field. Simply stated – a larger g factor

indicates a larger change in energy between different spin orientations in an applied magnetic field (Fig. 3).

The g factor often gives different values based on the orientation of a molecule in a magnetic field. The g value is often referred to as the g tensor for this point. For an organic radical, g will be close to 2.0023, which is the g factor for a free electron, and relatively independent of orientation. In contrast, metal complexes often have highly orientation dependent g values that can range from 0 to nearly 20 and are often anisotropic (*i.e.* g_x , g_y , and g_z are all different, x , y , and z here defined with respect to molecular axes). The largest and most anisotropic g factors tend to be found in rare-earth ions and low-coordinate transition metals.²² Generally speaking, anisotropic g values tend to produce faster relaxation rates, though there are many exceptions.

The second important factor is the magnetic interaction between an unpaired electronic spin and a nuclear spin on the same atom. This interaction is known as the hyperfine coupling interaction and has a strength denoted by the hyperfine coupling constant (A) (Fig. 3). An example of this interaction occurs in Co(II) complexes, from the coupling of the magnetic nucleus of ^{59}Co ($I = \frac{7}{2}$) and the unpaired electrons. Hyperfine interactions are very similar to J -coupling in proton nuclear magnetic resonance (NMR) spectroscopy. Hyperfine couplings can also be anisotropic, just like the g factor.

Magnetic nuclei are abundant in complexes beyond the spin-bearing ion and are often located in the ligand shell, counterions, and the surrounding matrix (*e.g.* proton-rich organic solvents, $I = \frac{1}{2}$ ^1H). Magnetic interactions with these latter three classes of nuclei is commonly referred to as “superhyperfine” coupling. Typically, stronger superhyperfine interactions with environmental magnetic species engender faster relaxation rates, though there is nuance to this statement that will be described later.

The final interaction that is noteworthy to highlight, known as the zero-field splitting, is a specific manifestation of spin-orbit coupling in metal ions with spin states greater than $\frac{1}{2}$ (as a result of bearing two or more unpaired electrons) (Fig. 3). For an electronic spin with $S > \frac{1}{2}$, there are $|2S + 1|$ accessible M_S levels. Each M_S level corresponds to a different alignment of the electronic spin relative to a molecular axis. A high $|M_S|$ value represents a spin precessing tightly around a molecular z axis, or closely aligned with that axis. A low $|M_S|$ value, in contrast, is a spin precessing far away from that same axis, commonly represented with perpendicular alignment to z . In the absence of zero-field splitting, or for a light atom (like an $S > \frac{1}{2}$ organic radical), the M_S levels are degenerate, or nearly so, at zero applied magnetic field. In the presence of zero-field splitting, which is common for heavy metal atoms (3d, 4d, 5d, 4f, and 5f elements) however, they are not, as shown in Fig. 3.

There are two parameters that describe the zero-field splitting. The first, the *axial* zero field splitting, or D , splits the energies of

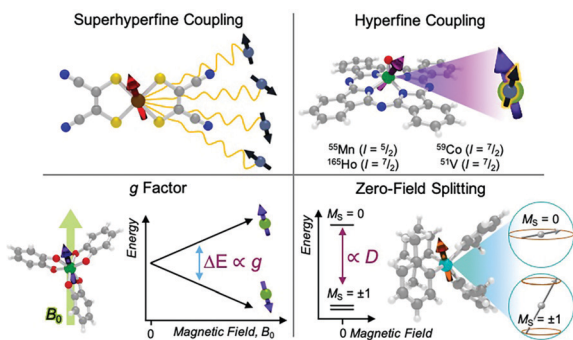


Fig. 3 Graphical overview of chemical features that impact relaxation processes.



$|M_S|$ levels away from one another (e.g. it will separate a pair of $M_S = \pm 1$ levels from $M_S = 0$ levels as in Fig. 3). For first-row transition metal complexes, values of D can range from less than 1 cm^{-1} to hundreds of cm^{-1} .²² The sign of D causes two limiting cases of M_S -level orderings. A positive D indicates a spin where the lowest energy level is $M_S = 0$ or $M_S = \pm \frac{1}{2}$, depending on whether S is integer (even number of electrons) or half-integer (odd number of electrons). A negative D indicates a spin where the M_S levels with largest $|M_S|$ are lowest in energy (e.g. $M_S = \pm 1$ for an $S = 1 \text{ Cr}^{4+}$ ion, as shown in Fig. 3). The second parameter, E , the *transverse* zero-field splitting or rhombic zero field splitting parameter, will cause an energy splitting of M_S levels in a pair (e.g., it will split $M_S = \pm 1$ levels from each other at zero field). The ratio $|E/D|$ is often referred to as the *rhombicity*.²³

The zero-field splitting is important for relaxation because individual M_S levels are the starting points, end points, and often transition states of the magnetic relaxation process. Thus, the zero-field splitting parameters directly modify the relaxation rates by dictating the relative energies of all involved steps in a relaxation pathway. The parameters D and E are also a direct result of the electronic structure of the metal complex, meaning that many intuitive molecular features, primarily symmetry and ligand field, can be modified to direct the sign and magnitude of D and E . This fact also means that general correlations between D , E , and relaxation times are challenging to make, and are best discussed on a case-by-case basis.

Spin–lattice relaxation (T_1)

Spin–lattice relaxation (also called longitudinal relaxation) refers to relaxation driven by interactions between the molecule and the environment, or “lattice,” that exchange energy. Spin–lattice relaxation specifically describes the process for a spin to relax from a starting spin-up (destabilized) orientation to the spin-down (stabilized) product orientation (Fig. 2). There is an enormous mechanistic diversity in how this process occurs depending on how a magnetic molecule exchanges energy with the lattice.

The energy exchanged with the environment during spin–lattice relaxation is in discrete amounts. Here, energy is exchanged through phonons, which are collective, long-range vibrations in a solid, often spanning multiple molecules.²⁴ The availability of phonons in a system is dependent on the temperature of the system and the nature of the surrounding matrix (e.g. a crystalline vs. frozen solvent glass environment).²⁷

The rates of the different mechanisms of relaxation that we discuss below generally follow different temperature, magnetic field, or environmental dependencies. Furthermore, the rates of these processes may be impacted by the lattice just as much as the molecule itself. Thus, in any real system, unraveling the operative mechanisms requires measuring the dependence of the relaxation process on all of these parameters. This process, though applied to dynamic magnetic processes, is therefore comparable to how one might determine the full picture of potential reaction pathways in a catalytic system. Below we give

an overview of each of the known mechanisms of spin–lattice relaxation, their relation to the reaction analogy established earlier, and the dependence of the processes on the molecule and environment.

Measurement of T_1

Spin–lattice relaxation is measured primarily by two different instruments: a magnetometer such as a magnetic properties measurement system (MPMS), or a pulsed electron paramagnetic resonance (EPR) spectrometer. Magnetometers allow for the study of magnetic relaxation processes across a range of temperatures, time scales and magnetic field strengths with the technique of alternating current (ac) susceptibility. However, the limited frequency range for ac susceptibility measurements (typically 0.01–1500 Hz) means that quickly relaxing systems ($T_1 < ca. 0.1 \text{ ms}$) can not be easily analyzed with this instrument. An alternative analysis is by pulsed EPR spectroscopy, which can measure T_1 through inversion or saturation recovery experiments. This technique can measure much faster relaxation times ($T_1 \sim 1 \mu\text{s}$ lengths), but pulsed spectrometers are far less common than magnetometers. There are some slight differences in the time constant measured by the techniques, because ac susceptibility extracts T_1 from bulk magnetization while the EPR experiment is probing a specific transition.²⁵ However, at the microscopic level, the same process is occurring – a spin is flipping during the process of relaxation. We direct the interested reader to several key resources to learn about these techniques in deeper detail.^{4,26}

Mechanisms of spin–lattice relaxation

Below we discuss the collection of mechanisms that drive T_1 relaxation, starting from those that most closely resemble the reaction-coordinate analogy, then move on to processes that deviate from the analogy and demonstrate the complexity of relaxation phenomena in magnetic molecules.

Orbach process

The Orbach process proceeds by excitation from the starting spin orientation to a higher energy spin orientation with energy provided by a phonon. This higher-energy M_S level is the effective transition state for the Orbach process. Relaxation then proceeds from this transition state to the lower energy spin orientation by releasing energy (in the form of a phonon) to the lattice (Fig. 4).^{26,27} For this process to occur, there must be available phonons of the appropriate energy to excite to the transition state. The Orbach process is not limited to a single “transition state” and can involve multiple steps to a higher energy M_S level before relaxation *via* phonon emission.

High-spin metal ions frequently relax *via* the Orbach process, as other M_S levels within the $|2S + 1|$ manifold often serve as transition states. In contrast, $S = \frac{1}{2}$ molecules do not show the Orbach process, as there are only two m_s levels for an



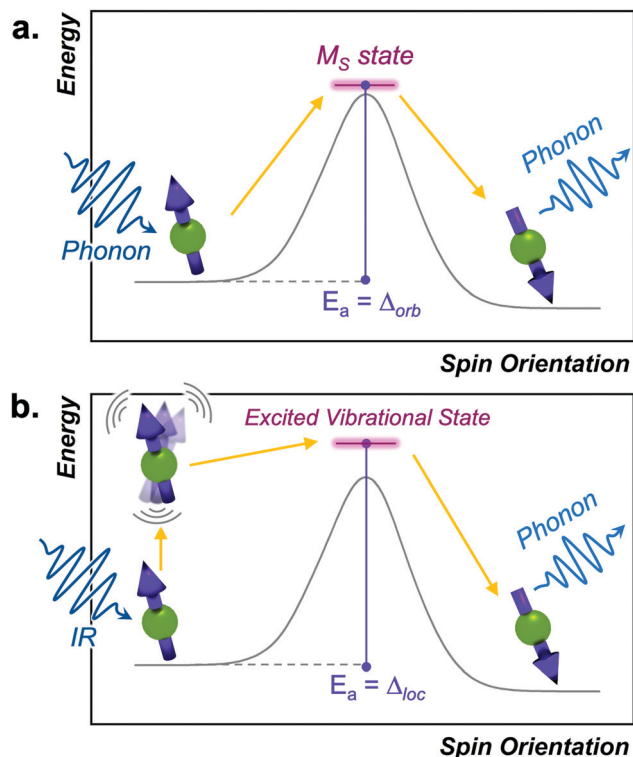


Fig. 4 (a) Depiction of the Orbach process in analogy to a reaction-coordinate diagram. The transition state is an actual high-energy M_S level for the relaxing spin. (b) Energy profile of the local mode process. Here, discrete vibrations in a molecular species promote the spin to overcome the barrier and enable relaxation.

$S = \frac{1}{2}$ system ($m_S = \pm\frac{1}{2}$) (and thus no transition state). The relaxation rate of the Orbach process is described by:

$$\frac{1}{T_1} = A_{\text{orb}} e^{(\Delta_{\text{orb}}/k_B T)} \quad (1)$$

This equation mirrors the form of the Arrhenius law, where $1/T_1$ is the rate constant. A_{orb} is the attempt frequency for the Arrhenius description and a larger A_{orb} means there is higher probability the mechanism will contribute to relaxation. A_{orb} is determined by several factors: it decreases in magnitude with increasing phonon availability (or energy availability) and generally increases as the activation energy increases.¹⁵ Hence, it is challenging to determine A_{orb} *a priori* on the basis of intuitive molecular considerations. Indeed, the basic theories for understanding relaxation (and the origins of A_{orb}) are developed for solid state defects, not molecular systems. As such, A_{orb} is frequently just extracted from experimental data, and is typically found with values of 10^{-2} to 10^{-10} s^{-1} .^{4,15} The term Δ_{orb} is the activation energy (E_a) to the M_S -level transition state. Finally, note that A_{orb} is denoted as τ_0 and Δ_{orb} as “ U_{eff} ” or “effective energy barrier” in the single-molecule magnet literature, which tends to focus on these parameters as the key figures of merit.⁴ An example molecule that exhibits the Orbach process is presented in the state-of-the-art section.

Local mode processes

The local mode relaxation mechanism utilizes discrete vibrations (“local modes”) on a molecule to facilitate relaxation.^{28,29} Here, incoming energy excites the spin system from the starting configuration to a transition state that is a local vibration. From this transition state, the spin can then relax to the product spin orientation while releasing energy (Fig. 4).³⁰ The process is in some ways like chemiluminescence, where the reaction product (the relaxed spin) is accompanied by an emission of energy, here in the form of a lattice vibration.

The temperature dependence of the relaxation rate of the local mode process is described by:

$$\frac{1}{T_1} = A_{\text{loc}} \frac{e^{\Delta_{\text{loc}}/k_B T}}{(e^{\Delta_{\text{loc}}/k_B T} - 1)^2} \quad (2)$$

where A_{loc} is related to the amplitude of the active mode, local strain at the molecule, and is, ultimately, hard to empirically predict for a given molecule, just like A_{orb} . Therefore, also like A_{orb} , A_{loc} is commonly extracted from experiment and considered as a “weight” for the contribution of this process to the overall relaxation rate. Δ_{loc} is the energy of the local mode. In some cases, the vibrations that enable the excitation can be assigned to molecular features using vibrational spectroscopy.^{31,32} One example of a species that exhibits a local mode process is MoO(TTP)(OEt) (TTP = tetraphenylporphyrin) tentatively driven by a *ca.* 240 cm^{-1} local mode when measured near 100 K by pulsed EPR. This molecule is notable because the relaxation rate of the local mode process is approximately an order of magnitude greater than the V congener, a reflection of the high sensitivity of high-SOC spins to vibrations.²⁸

Direct process

The direct process is the first one we discuss that does not have an activation energy. In this process, a spin directly proceeds from the starting to final orientation because of an energy match between the energy separation of the two levels and a lattice phonon.^{33,34} Importantly, the process effectively circumvents any possible transition states (Fig. 5). The direct process is analogous to the emission step of phosphorescence in photochemistry, where relaxation from an excited triplet to a ground singlet occurs without an activation energy, with a release of energy in the form of a phonon (a visible photon in the phosphorescence picture). However, unlike for phosphorescence emission, which involves a change in the number of unpaired electrons, the direct process is simply a change in spin orientation for the relaxing species.

The direct process has a distinctive magnetic field and temperature dependence that can enable diagnosis with relaxation studies. Indeed, the rate of relaxation for the direct process is described by the follow equation:

$$\frac{1}{T_1} = A_{\text{dir}} B^4 T \quad (3)$$

A_{dir} is related to many features of a relaxing molecule’s environment, specifically the number of phonons available that match



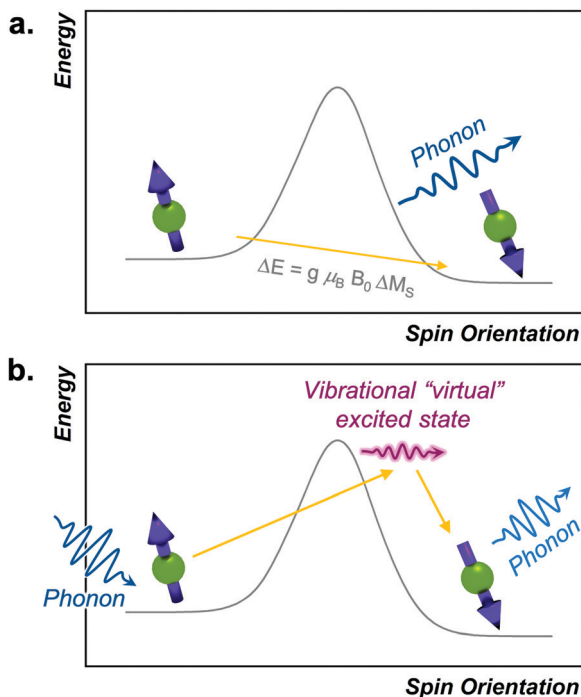


Fig. 5 (a) Depiction of the direct process in analogy to a reaction coordinate diagram. The spin flip energy is the Zeeman energy and relaxation emits a single phonon of that energy to the lattice. (b) Depiction of the Raman process in analogy to a reaction coordinate diagram. The “virtual” transition state is a superposition of vibrational states in the solid and is not an actual defined energy level. Relaxation involves two phonons in a simultaneous excitation and deexcitation of the spins, ultimately transferring energy to the environment of an amount equal to the ΔE between the starting and product spin orientations.

the ΔE separation of the starting/end orientations and speed of sound of the matrix.^{35,36} Like A_{loc} and A_{orb} , A_{dir} is generally treated as the weight of the contribution of this process to relaxation at a given field and temperature. B is the applied magnetic field and T is the temperature. The linear dependence of relaxation rate on temperature and B^4 field dependence are characteristic of the direct process. Furthermore, the involved phonons must be a direct energy match to the splitting of the starting materials and products of the spin system, which is small (usually $1\text{--}2\text{ cm}^{-1}$ or less) because the Zeeman energy is very small. As a result, this process is typically dominant only at low temperatures ($<10\text{ K}$), when the only available phonons from the lattice are small in energy. Examples of the direct process are abundant in copper(II) square planar complexes which tend to exhibit relaxation through the direct process below 20 K .³⁷

Raman process

The Raman relaxation process is more common at higher temperatures than the direct process. The Raman relaxation process proceeds when a spin system simultaneously absorbs and emits phonons of differing energies. This action is fundamentally different than the direct process, which emits only one phonon, or the Orbach process, which involves successive

absorption then emission. The difference in energies between the two phonons of the Raman process must match the difference in energy between the starting and final orientations. The fact that these phonons have defined energies suggests that there is absorption to and from a well-defined transition state. However, for the Raman process, that “state” is a superposition of lattice vibrations and does not actually exist, and is commonly referred to as a “virtual state”^{15,33} (Fig. 5b). This relaxation mechanism therefore does not have a clear connection to the classical reaction-coordinate picture developed earlier in this tutorial.¹² Nevertheless, the observation of the Raman process is incredibly common in relaxation studies. One example of a molecule exhibiting this process is the cobalt complex $[\text{Co}(\text{acac})_2(\text{H}_2\text{O})_2]$, which demonstrates the Raman process above 3 K when measured by ac susceptibility under ca. 1000 G magnetic fields.³⁸

The rate of relaxation *via* the Raman process is described by the following equation:

$$\frac{1}{T_1} = A_{\text{Ram}} \left(\frac{T}{\theta_{\text{D}}} \right)^n J_8 \left(\frac{\theta_{\text{D}}}{T} \right) \quad (4)$$

where A_{Ram} is an experimentally determined factor like A_{loc} , A_{dir} , and A_{orb} (and controlled by many similar parameters).³⁵ Like those coefficients, it also describes the contribution from the Raman process to the overall relaxation rate. B is the applied magnetic field, T is the temperature, θ_{D} is the Debye temperature (which corresponds to the energy above which phonons do not exist in a solid, converted to an energy by $1\text{ K} \approx 0.7\text{ cm}^{-1}$), n is an exponent dependent on the relaxing system, and J_8 is the transport integral.[†] The transport integral describes the energy distribution of phonons in the matrix. This integral and the characteristic temperature of the solid, or the Debye temperature (θ_{D}), depend on the physical composition of the solid.³⁹ Qualitatively, these two parameters signify that an environment with a larger distribution of accessible lattice vibrations will better facilitate the simultaneous absorption/emission process that drives the Raman process.

The Raman process produces a characteristic temperature dependence because of the exponent n . In theory, the rate of relaxation should scale with T^9 ($n = 9$) for species with a half-integer spin state and T^7 ($n = 7$) for integer spin state. In practice, however, multiple processes are often active and thus the apparent exponent extracted from fitting variable-temperature T_1 data is a non-integer value or ranges down to $n = 2$ to 3 .⁴⁰ The Raman process commonly occurs at temperatures where the lattice does not have available phonons of sufficient energy to excite the spin system to an actual transition state for the relaxation process. For this reason, this process “undercuts” the barrier, just like the direct process.

[†] The transport integral is defined by the following equation:

$$J_8 \left(\frac{\theta_{\text{D}}}{T} \right) = \int_0^{\theta_{\text{D}}/T} x^8 \frac{e^x}{(e^x - 1)^2} dx$$

where T is the temperature and θ_{D} is the Debye temperature.



Quantum tunneling processes

All previous mechanisms require energy input/release *via* interacting with the phonon system of the environment. However, there are two processes that circumvent this requirement by tunneling straight through the activation barrier for relaxation. It is for this reason that these processes are referred to as quantum tunneling of the magnetization (QTM) processes (Fig. 6).²³

There are two QTM-based processes. One occurs directly between the starting and final spin orientations, without any input energy, and is referred to as “ground-state” QTM (Fig. 6). This type of tunneling mechanism is operative if there is a magnetic interaction between the wavefunctions of the starting and final spin orientations, which opens a tunnel-splitting energy gap in the barrier (Δ_T , Fig. 6). Just like an electron tunnels because its spatial wavefunction can exist on both sides of an impenetrable barrier, the spin wavefunction of a system can exist on both sides of the activation barrier for reorientation. A stronger interaction between the starting and final orientations will generate a larger Δ_T and enable more efficient tunneling. Ground-state QTM is most facile near (or at) zero applied magnetic field, because interactions that drive QTM are strongest when the energy difference between the start and endpoint is smallest. Furthermore, the relaxation rate is generally temperature-independent because the tunneling

mechanism does not require energy from phonons to ascend over a barrier. Tunneling is typically dominant only at the lowest temperatures, when the available phonons lack sufficient energy to enable any of the other processes. For example, the molecule $[\text{Ph}_4\text{P}]_2[\text{Co}(\text{SPh})_4]$ displays tunneling below 2.5 K when measured by ac susceptibility under zero applied field.⁴¹

The second tunneling mechanism, thermally assisted QTM, requires exchange of energy with the lattice (Fig. 6). This mechanism involves an initial promotion (*via* phonon) of the starting spin orientation to a higher-energy M_s level. For the thermally assisted process, this level is below the highest-energy M_s level within a given manifold of an $S > \frac{1}{2}$ species (Fig. 6). Once promoted, the system then tunnels through the barrier to another M_s level, from which it can finally relax to the product. The thermally activated QTM mechanism behaves analogously to the Orbach process because of the promotion and phonon emission. However, this mechanism proceeds with an activation energy that is lower than the theoretical maximum defined by the zero-field splitting. An example of this phenomenon is illustrated with the $S = 2$ species $[(\text{TPA}^{\text{tBu}})\text{Fe}]^-$, which exhibits a Δ_{Orb} of 65 cm^{-1} , which is lower than the theoretical maximum of 192 cm^{-1} for this molecule.⁴²

Thermally activated processes

Thermally activated processes are relaxation processes that are driven by thermal motion in the environment, *e.g.* methyl or amino group rotations, wherein magnetic nuclei are located on the moving structure (Fig. 6). This process is common at higher temperatures, when the surrounding matrix of a magnetic molecule softens to allow physical motion to occur, when they would otherwise be frozen at low temperature.^{28,33,43} A given local motion will affect relaxation if the correlation time of that motion (τ_c , or the inverse of the rate of the local motion) approaches the energy difference of the starting/ending orientation (ω , in frequency).⁴³ The rates of thermally activated processes follow this equation:

$$\frac{1}{T_1} = A_{\text{therm}} \left(\frac{2\tau_c}{1 + \omega^2\tau_c^2} \right) \quad (5)$$

Here, A_{therm} correlates to the amplitude of the local magnetic field fluctuations from the thermally activated motion (but is generally treated as a weight for this process' contribution to T_1 , like A_{orb} , A_{loc} , A_{Ram} , and A_{dir}), τ_c^\ddagger is the correlation time of the thermally activated process, and ω is the frequency of the energy gap between the starting/ending spin configurations.⁴³ τ_c is temperature-dependent but depends on the activated motion and molecule, as τ_c needs to approach $1/\omega$. As such, there is no diagnostic temperature dependence for this process in the absence of other information about the molecular structure. One example of a molecule that displays

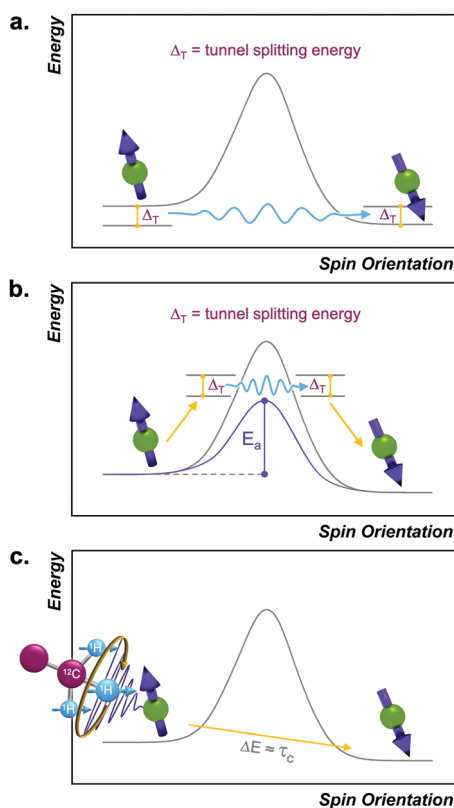


Fig. 6 Reaction-coordinate depictions for (a) ground-state and (b) thermally assisted quantum tunneling of the magnetization processes and (c) thermally activated processes.

\ddagger The correlation time for a thermally activated motion is described by the following equation: $\tau_c = \tau_0 e^{E_a/k_B T}$ where the E_a is the activation energy for the thermally activated process, τ_0 the preexponential factor, k_B is the Boltzmann constant, and T is the temperature.



this type of thermally activated process is in RbC₆₀ fulleride, where a metal-to-insulator phase transition triggers fast relaxation (measured by pulsed EPR) at *ca.* 25 K and 3400 G.⁴³

Intrinsic and extrinsic factors that impact T_1

There are many chemical and physical factors that impact T_1 relaxation in 3d transition metal ions. First, we describe the important (and synthetically tunable) factors for T_1 that are intrinsic to molecules then move on to extrinsic parameters that impact T_1 .

Strength and symmetry of the ligand field

The effects of the ligand field on relaxation most commonly manifest through zero-field splitting in high-spin ions. This tendency is for two main reasons. First, the magnitudes of D and E , which are governed by the ligand field, can shift the energies of the M_S levels involved in relaxation. Thus, D and E will dictate the activation energies for both the Orbach and thermally assisted QTM mechanisms. Second, a nonzero value of E will induce ground-state QTM as an efficient relaxation pathway. Generally, the magnitudes of D and E are greater when the ligand field for a high-spin metal ion is weaker. However, the dependence of D (both in sign and magnitude) on the ligand field is extremely intricate, and so many exceptions to this rule exist. Furthermore, the magnitude of E is heavily dependent on symmetry. Coordination geometries that have one principal axis of rotation and adhere to nearly idealized uniaxial symmetry tend to have lower-magnitude E parameters and thus, suppressed ground-state tunneling. We refer the reader to excellent reviews of these parameters, which demonstrate the power of molecular design for tuning relaxation processes.^{12,13,44,45}

Spin state

The effects of integer (“non-Kramers”) *versus* half-integer spin (“Kramers”) systems of unpaired electrons are evident most clearly in the QTM relaxation pathways. Here, a species with an integer spin is far more likely to exhibit relaxation by ground-state QTM than a half-integer-spin complex, because the term E is more effective at opening a tunneling gap for integer spin than half-integer-spin species.⁴ A second place the spin state impacts relaxation is the difference in temperature dependence of the Raman process, as described earlier.

Isotopic identity

Many metal ions possess isotopes that have non-zero nuclear spin, *e.g.* ⁵⁹Co ($I = \frac{7}{2}$, 100% natural abundance) or ¹⁶⁵Ho ($I = \frac{7}{2}$, also 100% natural abundance). The impact of a non-zero nuclear spin is primarily observed at zero field, low temperatures, and primarily affects quantum tunneling mechanisms.⁴⁶ The origin of this effect is the hyperfine coupling to the

metal-ion nuclear spin, which can open a tunneling gap and increase the rate of QTM.

Spin-orbit coupling

Spin-orbit coupling (SOC) is an intrinsic property of a molecule. This fundamental electronic structure feature is broadly important in dictating relaxation, but has two different effects, generally depending on whether relaxation occurs at high temperature or low temperature. The high temperature impact is because the SOC interaction ties the energies of the M_S levels (*e.g.* potential activation energies for the Orbach process) directly to the spin-bearing orbital energies of the metal ion. Consequently, small changes in the structure of a molecule (*e.g.* from a vibration that modulates metal–ligand bond distances) impact the spin-bearing orbitals, when then modulate the M_S -level energies *via* the SOC to facilitate relaxation. The stronger the SOC for a given system, the more efficient this effect, and the shorter T_1 tends to become. For this reason, light-element species, *e.g.* organic radicals, tend to have T_1 values (often μ s) that are longer at higher temperature than metal-ion systems. At lower temperature, the effect of a large SOC is different, particularly if the Orbach mechanism is active. In this case, a large SOC can push transition states to higher energies, enhancing the activation energy and ultimately slowing relaxation. The Dy-containing molecule in the state-of-the-art section later in the manuscript is a prime example of this point.

The importance of the SOC also means that the relaxation time can be dependent on precisely what orbital an electron resides in. Indeed, this effect is seen for the $[\text{V}(\text{C}_6\text{H}_4\text{O}_2)_3]^{2-}$ *versus* $[\text{VO}(\text{C}_6\text{H}_4\text{O}_2)_2]^{2-}$ molecules, which possess a single d electron in the d_{z^2} *versus* $d_{x^2-y^2}$ orbital, respectively. T_1 is different by approximately an order of magnitude between these two molecules owing to this difference.^{47–49}

Local magnetic species

The local magnetic species surrounding a molecule (the “spin bath”) also exert important effects on relaxation. This factor is readily altered by the chemist, who can choose, for example, to co-crystallize a molecule of interest with other molecules or dilute by dissolution in different organic solvents.⁴ These studies show that relaxation times generally increase when local magnetic content decreases, specifically the concentration of open-shell molecules. This general observation is because the presence of nearby magnetic units can hasten several of the above relaxation processes. For example, magnetic coupling to nearby electron spin systems will produce discrete spin levels that are relatively low-lying above the energies of the starting/ending spin orientations. Thus, the activation energy is lowered, and the relaxation rate accelerates *via* an Orbach process. Another common observed impact is through quantum tunneling, wherein proximate magnetic species enable the ground state QTM *via* a tunneling gap created by dipolar interactions.

Temperature

Temperature is one of the main extrinsic mechanisms of controlling relaxation. A given spin system will relax *via* the



most efficient mechanism available. Thus, over a given temperature regime, one process is typically dominant, and cooling/heating the system can transition the system to relaxation *via* a different mechanism. The exact ordering of these domains with temperature is highly dependent on the studied system. Nevertheless, a general ordering scheme is still possible. Ground-state QTM and the direct process are often observed at the lowest temperatures, where high-energy phonons are unavailable. A system will then usually transition to a regime where the Raman process is active upon warming. With further increasing temperature, where high-energy phonons become available in the solid, the Orbach process (or the analogous thermally assisted QTM) will typically take over. At even higher temperatures, local-mode and thermally activated mechanisms become active. A complete picture of magnetic relaxation for a molecule requires scanning an abundance of temperatures, just like understanding the kinetics of a chemical reaction.

Magnetic field

The applied magnetic field is the second most-commonly varied extrinsic factor for studying relaxation. Most relaxation processes exhibit a field dependence, because the magnetic field will vary the energies of the starting and final spin orientations. The only relaxation process with a rate that has a field dependence explicitly written into the equation is the direct process, which displays a B^4 field dependence. However, in the case of processes that involve an M_S -level transition state, the applied field can also vary the energy of the transition state, thereby modifying the activation energies and rates, though this impact is relatively small. A changing applied field can also affect the efficiency of the thermally activated process, by modulating the difference between the ΔE for the starting/ending spin configurations and the correlation times of environmental motions. Applied fields are also enormously impactful in the quantum tunneling, as they split the M_S levels nominally involved in tunneling, effectively killing the process and inducing slower magnetic relaxation. The effect of a magnetic field is much stronger on the ground-state tunneling process than the thermally assisted one, as larger M_S levels have energies that are more sensitive to changes in the magnetic field. The foregoing discussion justifies the widespread nature of using variable fields for studying relaxation: it can enhance T_1 by orders of magnitude and is thus a powerful handle for optimization.

In summary, there are many dials available to the experimentalist to turn when controlling the operative T_1 processes. However, the foregoing points are not the only factors that affect T_1 . Indeed, there are many other effects stemming from specifics of measurement. We direct the enthusiast to further reading for deeper information.¹⁵

The need for benchmarking

Frequently, a molecule will exhibit different relaxation processes depending on the aforementioned extrinsic factors. Because it is often too time-consuming to measure relaxation rates under every conceivable condition, the picture of magnetic relaxation in a molecule is often incomplete, and sometimes magnetic relaxation

mechanisms are misassigned. Thus, numerical comparisons of the impacts of the foregoing extrinsic/molecular features can be quite challenging. We contend that a larger discussion about benchmarking relaxation parameters is desperately needed in this field to overcome this difficulty and pave the way to true understanding.

Spin–spin relaxation (T_2)

Spin–spin relaxation, or T_2 , describes relaxation of a magnetic moment oriented in the plane perpendicular to the applied magnetic field (Fig. 7). The spin–spin relaxation time is one of the key figures of merit for quantum bits (“qubits”) in quantum computing, quantum sensing, and other quantum-information based fields.^{2,11} The given orientation-based description of T_2 is useful for instruction but note: the orientation of a spin is quantized and thus can only orient up or down in a magnetic field. The sideways depiction of the spin is in reality a unique quantum state known as a spin superposition, existing as both spin-up and spin-down orientations simultaneously. This quantum state is the cornerstone of the computational advantages for spin-based qubits.² A spinning-coin based analogy of the superposition is depicted in Fig. 7. Here, a spinning coin is neither heads nor tails but both simultaneously. Similarly, the spin up or spin down orientations are heads or tails, while the superposition is both.

In general, the superposition is an extremely delicate state and will decay into its constituent states very quickly after it is generated. The lifetime of this decay is the spin–spin relaxation time, T_2 . These values can range from very short (on the order of a few nanoseconds) to extremely long (on the order of milliseconds). For many applications of qubits, T_2 needs to be 100 μs , and pursuit of ever-larger T_2 values is a growing area of work.^{1,2,50} This time constant is known by several other

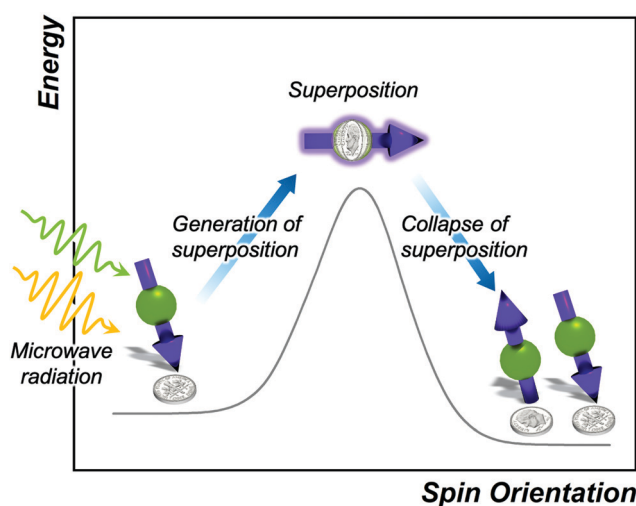


Fig. 7 Reaction-coordinate depiction of spin–spin relaxation. Spin up and spin down are shown to illustrate that the superposition can collapse to either spin-up or spin-down states. Coins depict the commonly used analogy that a spinning coin is a superposition of both heads and tails, or spin-up and spin-down state, respectively.



names in the literature: coherence time, transverse relaxation time, and phase memory relaxation time. We refer the interested reader to further references for the key differences between them.^{15,33} In general, however, these terms are used interchangeably to describe the same relaxation process. In this review, we will use the term “spin–spin relaxation time” and T_2 to describe all forms of this type of relaxation.

The spin–spin relaxation process

The process of spin–spin relaxation is fundamentally different than spin–lattice relaxation because of the superposition state, which does not require thermal activation to relax. Instead, any interaction of the superposition with the environment that perturbs the energy of the spin will initiate relaxation. We refer the interested reader to deeper descriptions of the spin–spin relaxation process and superposition collapse,^{33,51} as the ultimate goal here is to understand how local chemistry controls T_2 . Importantly, because there are many ways that the environment can interact with the spin superposition, there are likewise many factors that affect T_2 .

The immediate question, then, is how to apply the reaction-coordinate paradigm to understanding T_2 . However, the typical reaction-coordinate diagram is insufficient to describing the process, because the superposition is, in some sense, the transition state itself. As such, the activation energy is the input from the experimentalist, often with pulses of microwaves and the technique of pulsed EPR, to generate the superposition in the first place (Fig. 7). This fact is why T_2 relaxation does not follow the typical temperature-dependent rules that T_1 does. As one final point, T_2 is often much faster than T_1 . Whereas T_1 can approach second- and minute-long magnitudes for molecules, in most cases T_2 is at most hundreds of microseconds (rarely so), and more commonly tens of microseconds or less.¹¹

Measurement of T_2

The most common way of measuring the T_2 of a magnetic molecule is through the use of a pulsed EPR spectrometer. This instrument can apply a brief pulse of microwaves (typically 10-to-100 ns in length) to “tilt” the spin perpendicular in the applied magnetic field and generate the superposition. Then, after a delay, a second pulse of microwaves is applied, which triggers an emissive magnetic response, a *Hahn echo*, from the superpositions that have not decayed since the initial pulse. The time dependence of the echo intensity then provides a direct measurement of the lifetime of the superposition, which is T_2 . T_2 values typically have to be at least tens of nanoseconds in length to measure with commercial instrumentation. We direct the interested reader to several key volumes with more thorough descriptions of the techniques and instrumentation for analyzing T_2 .³³

Intrinsic and extrinsic factors that impact T_2

There are many different ways that the local chemical environment will affect T_2 (Fig. 8). For metal complexes, it is important

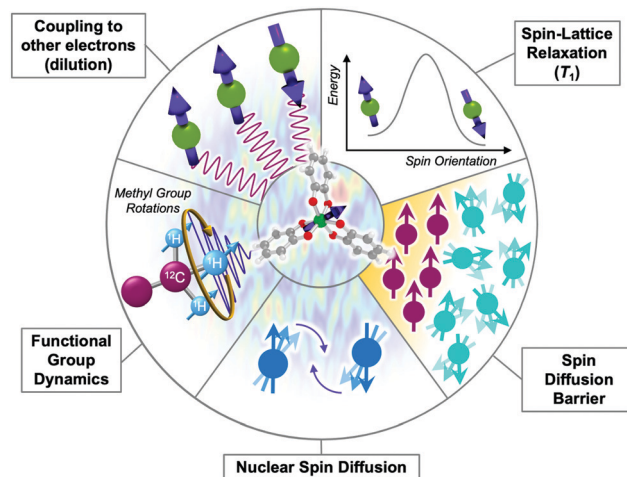


Fig. 8 Graphical depiction of intrinsic and extrinsic processes that affect T_2 . Processes that hasten T_2 relaxation via the creation of noise in the local magnetic field possess a “noisy” background.

to note that the “local environment” does not exclude the molecule itself. Indeed, specific molecular features, as detailed below, can and do commonly control T_2 by imposing specific local environmental effects. First, as for the T_1 section, we describe specific molecular features that control T_2 that can be manipulated by the synthetic chemist. Then we cover a select number of important extrinsic tunable features in the local environment that will affect T_2 . The following description is not exhaustive of all of the features that can affect T_2 , some of which can be induced by details of the measurement process itself. We refer the reader to other sources for descriptions of these effects.¹⁵

Electronic structure

The main ways that the electronic structure of a metal ion affects T_2 is by enhancing or suppressing environmental interactions. High-spin states (and the oxidation states and ligand fields that enable such spin configurations) tend to have shorter T_2 values because the larger magnetic moments of these spins display stronger dipolar interactions with environmental magnetism. This tendency is why molecular systems with the longest T_2 values tend to be $S = \frac{1}{2}$ ions, e.g. V(IV) and Cu(II).^{48,52} In the final section of this review, we detail two additional cutting-edge methods of manipulating electronic structure to suppress environmental interactions.

The second way that electronic structure will affect T_2 is through manipulating spin–lattice relaxation. Spin–lattice relaxation is the upper limit of T_2 for a molecule, because relaxation in the plane perpendicular to the magnetic field will be rapid if the spin is also rapidly realigning with the applied magnetic field direction.²⁶ When T_1 is short enough, T_2 is effectively controlled by all of the molecular factors described in the T_1 section. At the lowest temperatures, however, T_1 is usually orders of magnitude longer than T_2 . In these situations, T_1 does not limit T_2 , but with increasing temperatures, T_1 will



frequently shorten, eventually becoming the dominant contributor to T_2 relaxation.^{2,15}

Functional groups, counterions, and their dynamics

Functional groups are important for governing T_2 , like for many other chemical properties of molecules. However, those functional groups have to exert a chaotic, noisy fluctuation in the local magnetic field to affect T_2 . For example, the conventional “steric bulk” from a *tert*-butyl group will have an influence on a metal-ion T_2 not by impeding substitution-based reactivity, but because of a superhyperfine interaction with the nine magnetic protons (^1H , $I = \frac{1}{2}$, $\mu = 2.79\mu_{\text{N}}$ where μ is the nuclear magnetic moment in units of μ_{N} , the nuclear magneton) contained in the functional group.

There are two key points to make about the impacts of ligand and counterion-based magnetic nuclei. First, the motion of magnetic nuclei on functional groups in ligands and functional groups, such as rotation (*e.g.* in methyl groups), will generate potent changes in the local magnetic field that will shorten T_2 . Molecular motion shows its greatest effect on T_2 when the rate of a given motion is on the timescale of the measurement of T_2 and the dynamic group is closer to the spin.⁵³ The impact of motion on T_2 does not have a specific temperature dependence like T_1 , but instead will vary substantially at the temperatures where the timescale of a given motion approaches that of measurement. Second, the type, relative position, and number of magnetic nuclei on the ligand/counterion are all tunable handles to adjust T_2 . For example, substitution of proton nuclei with lower-magnetic-moment nuclei (*e.g.* ^2H , $I = 1$, $\mu = 0.86\mu_{\text{N}}$) will result in longer T_2 values because those substitutions suppress the amplitude of fluctuations in the local magnetic field. Finally, modifying the relative interactions between magnetic nuclei and the magnetic ion by varying the substitutional pattern on ligands, or changing the distance separating the metal ion and the magnetic nuclei, will also affect T_2 . There is considerable power in manipulating T_2 *via* molecular design through these species, though there is much to learn in this area.

Librational motion and orientation dependence

In many molecules, the relative orientation of the molecule to an applied magnetic field will affect the T_2 magnitude. This sensitivity stems from small molecular motions called librations (wagging/stretching of an entire molecule, not individual functional-group vibrations) that can occur at low temperatures. These librations will slightly change the orientation of the molecule in the applied magnetic field. If the g factor is anisotropic, as it is for many molecules, then the librations perturb the interactions with the applied magnetic field. This change in interaction drives T_2 relaxation. For many species, T_2 is greatest where a particular x , y , or z axis of the molecule is orientated parallel to the applied field, as slight changes at these orientations produce relatively small changes in the interaction with the applied field than “off-axis” orientations. There are two important practical outcomes of this effect.

First, there is no single T_2 for a molecule, instead, a given molecule will often exhibit multiple different T_2 values depending on its orientation relative to in an applied external magnetic field. Second, because of the orientation dependence, orientation is a critical design concern for any proposed molecule-based quantum computing architecture, specifically surface-mounted ones.⁵⁴

Local magnetic species

Fluctuations in local magnetism induced by other molecules in the environment also hasten T_2 relaxation. The strongest impacts come from proximate open-shell molecules, which exert their own local magnetic fields to hasten T_2 . The impact of local magnetic species is also more prominent when there are more of them, and T_2 can generally be observed to decrease with increasing concentration of open-shell species.^{15,33} Conversely, dilution increases T_2 and thus most reported T_2 times for molecular complexes are measured in millimolar (or less) concentrations.^{2,52,53,55,56}

A separate source local magnetic species is the collection of magnetic nuclei on surrounding molecules and matrix. The impacts of environmental nuclei can exert similar impacts on T_2 as nuclei in the ligand-shell. Hence, dynamic motion such as methyl rotations in the local matrix will shorten T_2 just as if it were part of the metal complex itself. A second mechanism through which nuclei can generate local magnetic noise and shorten T_2 is nuclear spin diffusion. Recall that the chemical shift of a magnetic nucleus is the energy required to flip that magnetic nucleus in a magnetic field. Spin diffusion occurs when two oppositely oriented nuclear spins (with identical chemical shifts) undergo a simultaneous flip, a “flip flop”, which produces changes to the local magnetic field (Fig. 8, bottom).

Nuclear spin diffusion is energy conserving, meaning that it cannot be frozen out by cooling to low temperatures, in contrast to other factors that affect T_2 . Therefore, nuclear spin diffusion tends to become the main contributor to T_2 at the lowest temperatures. The impact of this process can be minimized somewhat if the environment is full of lower-moment magnetic nuclei, *e.g.* a frozen deuterated solvent matrix instead of a protiated one.^{2,52,56}

Spin diffusion barrier

The requirement of matching chemical shifts between spin-diffusion-active nuclei generates a unique feature known as the spin diffusion barrier, which surrounds a relaxing open-shell molecule. Environmental magnetic nuclei interact with the magnetic molecule and that superhyperfine coupling strength grows as the magnetic nuclei approach the electronic spin. At a certain radius, even adjacent nuclei experience substantially different strengths of these superhyperfine interactions. As a consequence, the two nuclei now have significantly different chemical shifts from each other and the other magnetic nuclei of the bath. Spin diffusion is thus deactivated for nuclei in close proximity to the electron, even if the two nuclei are right next to each other. Consequently (and counterintuitively) nuclei that are the closest to an electronic spin will not generate



substantial fluctuations in the local magnetic field to shorten T_2 . The radius at which the T_2 -shortening effect of magnetic nuclei is deactivated is the spin diffusion barrier.⁵⁸ Existing studies place the edge of the barrier at about 3–10 Å.^{57,59} However, the actual radius is dependent on the system, and does not need to be spherical.^{2,60} Finally, the barrier also precludes spin diffusion as a mechanism by which the hyperfine coupling to a metal-ion nuclear spin will impact T_2 .⁶¹

State of the art in molecular magnetic relaxation

The length of the relaxation time for a system directly translates into potential utility. For example, realizing a long T_1 is a path toward storing and processing information in a system,¹³ as well as controlling environmental ^1H dynamics for magnetic resonance imaging.⁶² Producing long T_2 values enables quantum information processing⁵² and enables the performance of advanced magnetic resonance detection protocols, as could be leveraged in, *e.g.* electron paramagnetic resonance imaging.⁶³ Below we highlight the forefront of this area of work, describing how the combined use of synthetic chemistry and electronic structure design strategies are producing new pathways to slowing magnetic relaxation.

Complexes with slow spin–lattice relaxation

One prominent class of molecular species with long spin–lattice relaxation times, single molecule magnets, features a single metal ion (transition-metal or lanthanide) with carefully crafted ligand shells to facilitate a long T_1 . Many of these design strategies rely on controlling the activation energies of spin–lattice relaxation through the Orbach or thermally assisted QTM process and suppressing the non-thermally activated processes. With a suitably high activation energy and the absence of these other processes, one may expect sufficiently long relaxation to enable information storage in a molecular magnetic moment. In pursuit of this goal, in 2017, Layfield, Chilton, Mills, and coworkers reported new Ln-based metallocene complexes that show a slow relaxation process up to and above liquid nitrogen temperatures (Fig. 9).^{20,64} Analyses of the temperature dependence of T_1 in these species revealed that one molecule, $[(\text{Cp}^{\text{IPr5}})\text{Dy}(\text{Cp}^*)]^+$, exhibited an activation energy for relaxation of $1541(11) \text{ cm}^{-1}$ for a thermally assisted QTM process, the highest reported value for any molecule. There are two key aspects of the molecule that enable this remarkable observation. First, the molecule possesses a Dy^{3+} ion, which leverages a large spin state and spin–orbit coupling to produce a large potential activation energy. The spin orbit coupling in particular leads to a $^6\text{H}_{15/2}$ term ground state and magnetic orientations that are described with M_J values up to $\pm\frac{15}{2}$ (stemming from $J = L + S$, where J is the total angular momentum, L the orbital angular momentum, and S the spin).⁴⁴ Second, the species also possesses a highly pseudo-axial molecular symmetry. These two properties suppress

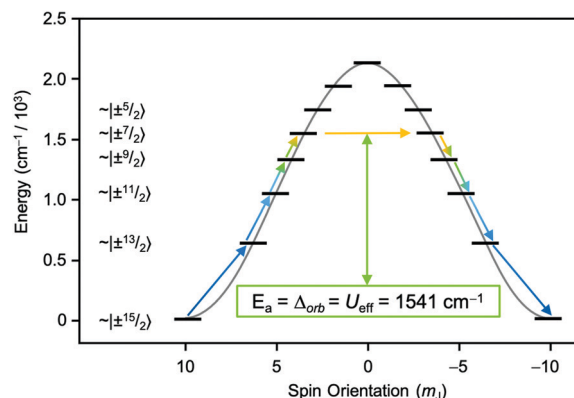
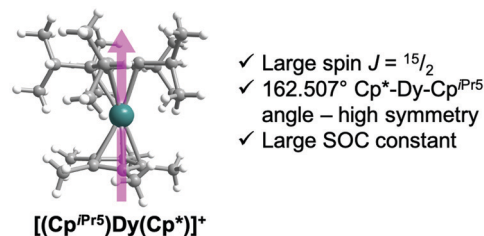


Fig. 9 Recent ground-breaking system for targeting a long T_1 . The highly axial geometry of the metal complex suppresses tunneling, which combines with the high spin state and large spin–orbit coupling to enable an exceptionally high activation energy to relaxation. Cp^* = pentamethylcyclopentadiene; Cp^{IPr5} = pentaisopropylcyclopentadiene. The levels in the bottom plot are labeled in “ket” notation, which here simply give the M_J values of the $^6\text{H}_{15/2}$ Dy^{3+} ion, which reach up to $M_J = \pm\frac{15}{2}$.

tunneling in the ground state and lower lying levels. The result is a thermally assisted tunneling mechanism that is nearly the maximum possible activation energy for the molecule.

Complexes with slow spin–spin relaxation

Long T_2 values are common for open-shell defects in solid-state materials like the nitrogen vacancy center in diamond⁶⁵ or the double-vacancy site in silicon carbide.⁶⁶ All of these species exhibit extremely large T_2 values (~ 1 ms) because isotopic enrichment removes all environmental nuclear spins. Metal complexes are emerging as powerful alternative platforms to understand how to lengthen superposition lifetimes using synthetic chemistry, a strategy that is substantially more challenging in the solid state.

Of these, first-row transition metal complexes hold recent records for the longest spin–spin relaxation times. In 2015, Freedman and coworkers demonstrated the impact of using synthetic design to achieve near-complete removal of nuclear spins from the coordination shell of a metal ion. The resulting molecule, $(\text{d}^{20}\text{-Ph}_4\text{P})_2[\text{V}(\text{C}_8\text{S}_8)_3]$, uses nearly nuclear-spin free ligands (^{12}C : $I = 0$, 98.9%; $^{32/34/36}\text{S}$: $I = 0$, 99.25%), deuterated counterions, and a remarkable solubility in CS_2 , a relatively nuclear spin-free solvent. Together, these factors create a magnetically quiet environment, and hence a groundbreaking, millisecond T_2 is observed (Fig. 10). This finding (the first for a metal complex) shows that controlling the properties of magnetic nuclei in the environment is critical to lengthen T_2 , but



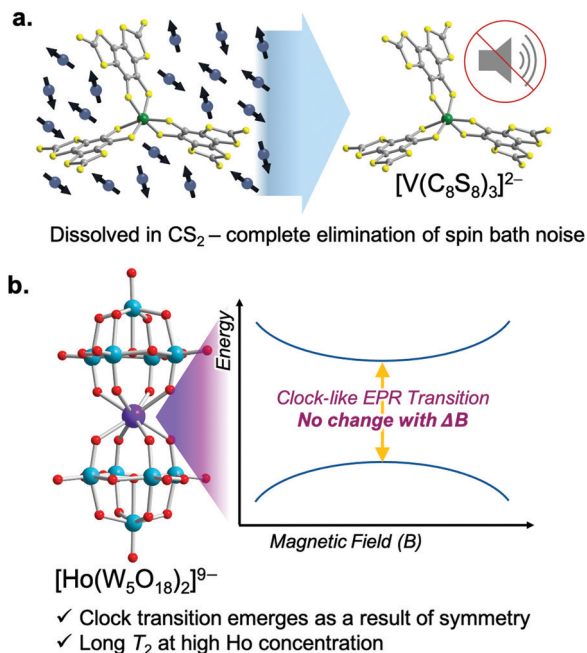


Fig. 10 State-of-the-art molecules for lengthening T_2 . (a) $[\text{V}(\text{C}_8\text{S}_8)_3]^{2-}$ is a predominantly nuclear-spin free complex that displays millisecond T_2 values when dissolved in the nuclear-spin-free solvent CS_2 . (b) $[\text{Ho}(\text{W}_5\text{O}_{18})_2]^{9-}$ exploits an almost perfect D_4 molecular symmetry to generate a clock-like EPR transition. These transitions have zero change in frequency with respect to changes in local magnetic field, and thus yield long T_2 values.

there is much to learn about the exact role of ligand-based nuclei, and cutting-edge efforts reflect that.^{48,53}

Another exciting strategy to lengthen T_2 is to design electronic structures that suppress environmental sensitivity. One method to do so is to create an EPR transition at an avoided-crossing point (Fig. 10). An avoided crossing point is where two M_S values would cross in energy at a given magnetic field. For an avoided crossing, however, there is an interaction between the wavefunctions of the M_S levels that strengthens as the levels become closer in energy. As a result, the M_S levels never actually cross and instead bow away from each other with increasing magnetic field. The result is an energy gap that is accessible with EPR spectroscopy. However, the transition is unconventional as the energies of the M_S levels (M_J in the case of the Ho^{3+}) are field-independent at the crossing point. The field independence of the energies translates into a relative immunity toward local magnetism, and the superposition created at the avoided crossing point is consequently relatively long-lived. In fact, this mechanism for designing environmental immunity engenders the extraordinary frequency stability of the time-keeping transition of an atomic clock. It is for this reason that the EPR transition is “clock-like”. In the case of the molecule shown in Fig. 10, the symmetry, orbital angular momentum, and crystal field of the Ho^{3+} are what create the clock transition. However, other magnetic interactions in molecules can be used to create a clock transition, *e.g.* the hyperfine interaction.⁶⁷ It remains to be seen if this strategy will be able to achieve the near-millisecond T_2 values in molecules reported

by the nuclear-spin-free strategy. Nevertheless, these studies represent a fascinating and rare intersection of coordination chemistry and atomic-clock physics.

Outlook

Magnetic molecules are promising components of next-generation applications spanning from quantum information processing to magnetic resonance imaging. But to realize that potential, we need a comprehensive understanding of how molecular structure and the environment control magnetic relaxation. In this review, we contextualized the relaxation process in a new analogy to a reaction-coordinate diagram, which enabled us to intuit relaxation phenomena in an accessible manner.

The question then is, where to next? There are of course many exciting directions, and we touched on a select few of these in the last section. We highlight one final area of particular opportunity: studying relaxation in highly magnetic and dynamic environments. Indeed, many of the proposed applications for magnetic metal complexes require long relaxation times in proton-rich biological environments, room-temperature solutions, or the stray-magnetic-field-rich interiors of electronic devices. Yet, as touched on herein, current studies of magnetic molecules focus almost exclusively on electron- and nuclear-spin-free conditions to suppress magnetic noise and very low temperatures to freeze out structural dynamics. These studies revealed that extraordinarily long relaxation times are achievable. One pressing goal, then is the discovery of how to translate those proof-of-concept observations into noisy conditions. We are excited to see, then, more fundamental studies into controlling the interactions between spin baths and molecules *via* structural design.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

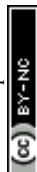
We thank C. D. Charles and D. A. Corbin for helpful feedback. We thank the NSF (CHE-1836537), the NIH (R21EB027293), and Colorado State University (CSU) for financial support. CSU acknowledges, with respect, that the land the university is on today is the traditional and ancestral homelands of the Arapaho, Cheyenne, and Ute Nations and peoples. I. P. M. is supported by the National Science Foundation Graduate Research Fellowship Program under Grant No. (006784-0002).

References

- 1 G. Aromí, D. Aguilà, P. Gamez, F. Luis and O. Roubeau, *Chem. Soc. Rev.*, 2012, **41**, 537–546.
- 2 M. J. Graham, J. M. Zadrozny, M. S. Fataftah and D. E. Freedman, *Chem. Mater.*, 2017, **29**, 1885–1897.



- 3 A. Gaita-Ariño, F. Luis, S. Hill and E. Coronado, *Nat. Chem.*, 2019, **11**, 301–309.
- 4 D. Gatteschi, R. Sessoli and J. Villain, *Molecular Nanomagnets*, Oxford University Press, Oxford, New York, 2006.
- 5 J. Wahsner, E. M. Gale, A. Rodríguez-Rodríguez and P. Caravan, *Chem. Rev.*, 2019, **119**, 957–1057.
- 6 G. Tircs and Z. Baranyai, *The Chemistry of Contrast Agents in Medical Magnetic Resonance Imaging Stability and Toxicity of Contrast Agents*, John Wiley & Sons, 2013.
- 7 M. C. Heffern, L. M. Matosziuk and T. J. Meade, *Chem. Rev.*, 2014, **114**, 4496–4539.
- 8 A. L. Buchachenko and V. L. Berdinsky, *Chem. Rev.*, 2002, **102**, 603–612.
- 9 D. Aravena and E. Ruiz, *Dalton Trans.*, 2020, **49**, 9916–9928.
- 10 S. Brooker, *Chem. Soc. Rev.*, 2015, **44**, 2880–2892.
- 11 E. Coronado, *Nat. Rev. Mater.*, 2020, **5**, 87–104.
- 12 S. T. Liddle and J. van Slageren, *Chem. Soc. Rev.*, 2015, **44**, 6655–6669.
- 13 G. A. Craig and M. Murrie, *Chem. Soc. Rev.*, 2015, **44**, 2135–2147.
- 14 L. Bogani and W. Wernsdorfer, *Nat. Mater.*, 2008, **7**, 179–186.
- 15 L. J. Berliner, S. S. Eaton and G. R. Eaton, *Distance Measurements in Biological Systems by EPR*, Springer, US, 2000.
- 16 *In Vivo EPR (ESR): Theory and Application*, ed. L. J. Berliner, Springer, US, 2003.
- 17 B. A. Reisner, S. R. Smith, J. L. Stewart, J. R. Raker, J. L. Crane, S. G. Sobel and L. L. Pesterfield, *Inorg. Chem.*, 2015, **54**, 8859–8868.
- 18 J. R. Raker, B. A. Reisner, S. R. Smith, J. L. Stewart, J. L. Crane, L. Pesterfield and S. G. Sobel, *J. Chem. Educ.*, 2015, **92**, 980–985.
- 19 L. J. Fox and G. H. Roehrig, *J. Chem. Educ.*, 2015, **92**, 1456–1465.
- 20 F.-S. Guo, B. M. Day, Y.-C. Chen, M.-L. Tong, A. Mansikkamäki and R. A. Layfield, *Science*, 2018, **362**, 1400–1403.
- 21 L. Tesi, A. Lunghi, M. Atzori, E. Lucaccini, L. Sorace, F. Totti and R. Sessoli, *Dalton Trans.*, 2016, **45**, 16635–16643.
- 22 J. M. Frost, K. L. M. Harriman and M. Murugesu, *Chem. Sci.*, 2016, **7**, 2470–2491.
- 23 D. Gatteschi and R. Sessoli, *Angew. Chem., Int. Ed.*, 2003, **42**, 268–297.
- 24 M. Vallone, *Phys. Status Solidi B*, 2020, **257**, 1900443.
- 25 M. S. Fataftah, M. D. Krzyaniak, B. Vlasisavljevich, M. R. Wasielewski, J. M. Zadrozny and D. E. Freedman, *Chem. Sci.*, 2019, **10**, 6707–6714.
- 26 A. Schweiger and G. Jeschke, *Principles of Pulse Electron Paramagnetic Resonance*, Oxford University Press, 2001.
- 27 A. Lunghi, F. Totti, R. Sessoli and S. Sanvito, *Nat. Commun.*, 2017, **8**, 14620.
- 28 Y. Zhou, B. E. Bowler, G. R. Eaton and S. S. Eaton, *J. Magn. Reson.*, 1999, **139**, 165–174.
- 29 D. W. Feldman, J. G. Castle and J. Murphy, *Phys. Rev.*, 1965, **138**, A1208–A1216.
- 30 P. G. Klemens, *Phys. Rev.*, 1962, **125**, 1795–1798.
- 31 C. B. Harris, R. M. Shelby and P. A. Cornelius, *Phys. Rev. Lett.*, 1977, **38**, 1415–1419.
- 32 N. V. Vugman and M. R. Amaral, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1990, **42**, 9837–9842.
- 33 D. Goldfarb and S. Stoll, *EPR Spectroscopy: Fundamentals and Methods*, John Wiley & Sons, 2018.
- 34 J. Murphy, *Phys. Rev.*, 1966, **145**, 241–247.
- 35 A. Abragam and B. Bleaney, *Electron Paramagnetic Resonance of Transition Ions*, Oxford University Press, Oxford, New York, 2012.
- 36 K. J. Standley, *Electron Spin Relaxation Phenomena in Solids*, Springer, US, 1969.
- 37 A. J. Fielding, S. Fox, G. L. Millhauser, M. Chattopadhyay, P. M. H. Kroneck, G. Fritz, G. R. Eaton and S. S. Eaton, *J. Magn. Reson.*, 2006, **179**, 92–104.
- 38 S. Gómez-Coca, A. Urtizberea, E. Cremades, P. J. Alonso, A. Camón, E. Ruiz and F. Luis, *Nat. Commun.*, 2014, **5**, 4300.
- 39 W. M. Rogers and R. L. Powell, *Tables of Transport Integrals*, U.S. Government Printing Office, 1958.
- 40 L. Gu and R. Wu, *Phys. Rev. B*, 2021, **103**, 014401.
- 41 J. M. Zadrozny and J. R. Long, *J. Am. Chem. Soc.*, 2011, **133**, 20732–20734.
- 42 W. H. Harman, T. D. Harris, D. E. Freedman, H. Fong, A. Chang, J. D. Rinehart, A. Ozarowski, M. T. Sougrati, F. Grandjean, G. J. Long, J. R. Long and C. J. Chang, *J. Am. Chem. Soc.*, 2010, **132**, 18115–18126.
- 43 V. A. Atsarkin, V. V. Demidov and G. A. Vasneva, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1997, **56**, 9448–9453.
- 44 J.-L. Liu, Y.-C. Chen and M.-L. Tong, *Chem. Soc. Rev.*, 2018, **47**, 2431–2453.
- 45 B. N. Figgis and M. A. Hitchman, *Ligand field theory and its applications*, Wiley-VCH, 2000.
- 46 E. Moreno-Pineda, M. Damjanović, O. Fuhr, W. Wernsdorfer and M. Ruben, *Angew. Chem., Int. Ed.*, 2017, **56**, 9915–9919.
- 47 M. Atzori, E. Morra, L. Tesi, A. Albino, M. Chiesa, L. Sorace and R. Sessoli, *J. Am. Chem. Soc.*, 2016, **138**, 11234–11244.
- 48 C.-J. Yu, M. J. Graham, J. M. Zadrozny, J. Niklas, M. D. Krzyaniak, M. R. Wasielewski, O. G. Poluektov and D. E. Freedman, *J. Am. Chem. Soc.*, 2016, **138**, 14678–14685.
- 49 S. R. Cooper, Y. B. Koh and K. N. Raymond, *J. Am. Chem. Soc.*, 1982, **104**, 5092–5102.
- 50 M. Atzori and R. Sessoli, *J. Am. Chem. Soc.*, 2019, **141**, 11339–11352.
- 51 M. A. Schlosshauer, *Decoherence: and the Quantum-To-Classical Transition*, Springer-Verlag, Berlin Heidelberg, 2007.
- 52 K. Bader, D. Dengler, S. Lenz, B. Endeward, S.-D. Jiang, P. Neugebauer and J. van Slageren, *Nat. Commun.*, 2014, **5**, 5304.
- 53 C. E. Jackson, C.-Y. Lin, S. H. Johnson, J. van Tol and J. Zadrozny, *Chem. Sci.*, 2019, **10**, 8447–8454.
- 54 M. Atzori, L. Tesi, E. Morra, M. Chiesa, L. Sorace and R. Sessoli, *J. Am. Chem. Soc.*, 2016, **138**, 2154–2157.
- 55 J. M. Zadrozny, J. Niklas, O. G. Poluektov and D. E. Freedman, *ACS Cent. Sci.*, 2015, **1**, 488–492.
- 56 K. Bader, M. Winkler and J. van Slageren, *Chem. Commun.*, 2016, **52**, 3623–3626.
- 57 E. R. Canarie, S. M. Jahn and S. Stoll, *J. Phys. Chem. Lett.*, 2020, **11**, 3396–3400.



- 58 W. E. Blumberg, *Phys. Rev.*, 1960, **119**, 79–84.
- 59 J. Chen, C. Hu, J. F. Stanton, S. Hill, H.-P. Cheng and X.-G. Zhang, *J. Phys. Chem. Lett.*, 2020, **11**, 2074–2078.
- 60 J. P. Wolfe, *Phys. Rev. Lett.*, 1973, **31**, 907–910.
- 61 R. Hussain, G. Allodi, A. Chiesa, E. Garlatti, D. Mitcov, A. Konstantatos, K. S. Pedersen, R. De Renzi, S. Piligkos and S. Carretta, *J. Am. Chem. Soc.*, 2018, **140**, 9814–9818.
- 62 J. P. Klare, *Biomed. Spectrosc. Imaging*, 2012, **1**, 101–124.
- 63 R. Rahimi, H. J. Halpern and T. Takui, in *Oxygen Transport to Tissue XXXIX*, ed. H. J. Halpern, J. C. LaManna, D. K. Harrison and B. Epel, Springer International Publishing, Cham, 2017, pp. 335–339.
- 64 C. A. P. Goodwin, F. Ortu, D. Reta, N. F. Chilton and D. P. Mills, *Nature*, 2017, **548**, 439–442.
- 65 F. Casola, T. van der Sar and A. Yacoby, *Nat. Rev. Mater.*, 2018, **3**, 1–13.
- 66 A. Lohrmann, B. C. Johnson, J. C. McCallum and S. Castelletto, *Rep. Prog. Phys.*, 2017, **80**, 034502.
- 67 J. M. Zadrozny, A. T. Gallagher, T. D. Harris and D. E. Freedman, *J. Am. Chem. Soc.*, 2017, **139**, 7089–7094.

