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A structure activity relationship for ring closure reactions in unsaturated alkylperoxy radicals†

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Terpenoids are an important class of multi-unsaturated volatile organic compounds emitted to the atmosphere. During their oxidation in the troposphere, unsaturated peroxy radicals are formed, which may undergo ring closure reactions by an addition of the radical oxygen atom on either of the carbons in the C=C double bond. This study describes a quantum chemical and theoretical kinetic study of the rate of ring closure, finding that the reactions are comparatively fast with rates often exceeding 1 s^{-1} at room temperature, making these reactions competitive in low-NO_x environments and allowing for continued autoxidation by ring closure. A structure–activity relationship (SAR) is presented for 5- to 8-membered ring closure in unsaturated RO₂ radicals with aliphatic substituents, with some analysis of the impact of oxygenated substituents. H-migration in the cycloperoxide peroxy radicals formed after the ring closure was found to be comparatively slow for unsubstituted RO₂ radicals. In the related cycloperoxide alkoxy radicals, migration of H-atoms implanted on the ring was similarly found to be slower than for non-cyclic alkoxy radicals and is typically not competitive against decomposition reactions that lead to cycloperoxide ring breaking. Ring closure reactions may constitute an important reaction channel in the atmospheric oxidation of terpenoids and could promote continued autoxidation, though the impact is likely to be strongly dependent on the specific molecular backbone.

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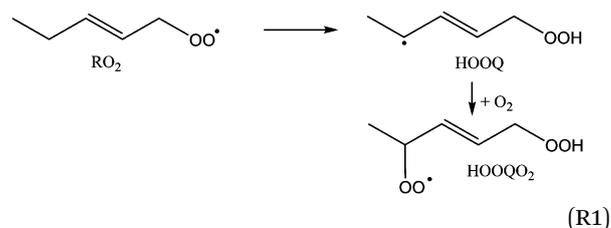
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1. Introduction

The chemistry of peroxy radicals (RO₂) has received increased interest over the last decade. In combustion systems, the H-migration in peroxy radicals has been shown to be a source of OH radicals in low-temperature combustion.^{1–3} Similar H-migration processes have been shown to be important in the atmosphere in regenerating OH radicals,⁴ the most important radical oxidizing volatile organic compounds (VOCs). H-Migration in RO₂, followed by O₂ addition on a product alkyl radical, is also the critical reaction sequence in the so-called autoxidation of VOCs, where successive H-migration/O₂-addition events can lead to highly oxidized molecules (HOMs).^{5–9} Such molecules typically have a low vapor pressure, and can readily transition to the aerosol phase, affecting air quality and climate change. The rate of H-shift

reactions depends strongly on the distance migrated, as well as on substituents in the molecule;^{10–14} particularly favorable are H-atoms from oxygenated groups (aldehydes, alcohols) and β-unsaturated moieties leading to allyl-resonance stabilized product radicals.¹¹



Due to the much stronger =C–H bond on carbons with a double bond, vinylic H-atoms do not migrate at a rate competitive to other atmospheric processes, and in autoxidation processes, the double-bonded carbons are thus more difficult to activate and oxidize, limiting the oxygen to carbon ratio O:C. In the atmosphere, however, the highest emissions of non-methane VOCs consist of isoprene (C₅H₈), monoterpenes (C₁₀H₁₆), and other terpenoids,¹⁵ most of which are (poly-)unsaturated compounds. Even for these compounds, highly oxidized compounds were measured with very high O:C ratios,^{7,9,16–19} indicating that all carbons can be oxidized without breaking the molecular backbone.

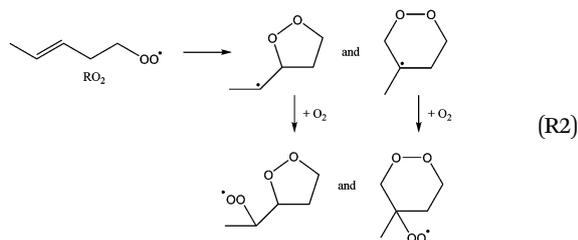
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† Electronic supplementary information (ESI) available: Details on the algorithm for the conformer search; the geometries, rovibrational characteristics and energies of all reactants, TS and products characterized in this study are available at the M06-2X level of theory. Where available, CCSD(T) energies and T1 diagnostics are also provided. See DOI: 10.1039/d1cp02758a



One of the pathways that enables the oxidation of double-bonded carbons in a molecule is a ring closure reaction:



Ring closure reactions in RO_2 and alkoxy radicals have been studied theoretically,^{20–27} and mechanisms incorporating these reactions for some compounds²⁸ have been shown to provide a better agreement with experimental data. Still, these studies focused on a subset of molecules, and no structure–activity relationship (SAR) is available to enable systematic inclusion of such ring closure reactions in atmospheric chemical mechanisms.

In this work, we describe a theory-based SAR for ring closure in unsaturated RO_2 radicals, examining 4- to 8-membered ring formation in linear and branched olefins. This SAR is directly applicable in mechanism development, can be readily implemented in software-assisted mechanism generation software such as GeckoA^{29,30} or SAPRC,³¹ and serves as a guide for further extensions examining a wider range of substituents. We also examine the ability of the product cycloperoxide radicals to undergo further autoxidation steps by H-migration, either at the peroxy radical or alkoxy radical stages of their subsequent chemistry.

2. Methodology

The geometries and rovibrational characteristics of all conformers of the reactants and transition states were first optimized at the M06-2X/cc-pVDZ level of theory,^{32,33} performing an exhaustive search of the conformational space of all internal rotors and ring structures (see ESI† for more details). The ring closure alkyl radical products were likewise characterized, though here the conformational space was not comprehensively covered in all cases as the alkyl radical products don't contribute to the reaction kinetics. All structures obtained thus were then re-optimized at the M06-2X/aug-cc-pVTZ (ring closure) or M06-2X-D3/aug-cc-pVTZ (H-migration) level of theory,^{32–34} with a harmonic frequency analysis performed at the same level of theory. ZPE corrections and partition function calculations are based on this data, using a wavenumber scaling factor of 0.971.^{35,36} Finally, the relative energy of the transition state and products was further improved by single point calculations at the CCSD(T)/aug-cc-pVTZ level of theory³⁷ on the lowest-energy conformers of reactant and TS. For molecules with multiple low-lying conformers, several of the lowest-lying conformers were examined at this level of theory to ensure the lowest energy structure at the coupled cluster level was used in calculating the barrier heights. T1 diagnostics (see ESI†) indicate that no strong multi-reference character is expected for any of the structures.^{38,39}

The rovibrational data at the M06-2X or M06-2X-D3/aug-cc-pVTZ level combined with ZPE-corrected CCSD(T) barrier heights were then used to predict the thermal rate coefficient in a rigid rotor harmonic oscillator multi-conformer transition state theory (MC-TST) approximation,⁴⁰ incorporating all conformers. Tunneling corrections are included based on an asymmetric Eckart barrier paradigm.^{41,42} As the reaction rates are low compared to the thermalization rate by collisions with the bath gas, no pressure dependence is expected under atmospheric conditions. As discussed below, the reverse reactions are expected to be too slow to compete, and no rate corrections were implemented for reverse reactions.

3. Ring closure reactivity trends

The barrier heights and rate coefficients for all the ring closure reactions are provided in Table 1. We have also collected calculations and literature data on ring closure in oxygenated RO_2 radicals in Table 2.

3.1 Ring size and site of attack on the double bond

The size of the ring structure is a key parameter in the rate of reaction, where especially smaller rings suffer from an energetic disadvantage due to ring strain. The ring strain is furthermore influenced by whether the ring closure occurs on the nearer or farther carbon of the double bond. Formation of 4-membered rings is energetically unfavorable, being endothermic by $\sim 20 \text{ kcal mol}^{-1}$ and requiring surmounting a high energy barrier of around 30 kcal mol^{-1} . All other ring closure reactions are roughly thermoneutral, within a few kcal mol^{-1} . 5-Membered ring closure in β -unsaturated RO_2 radicals, however, still has a high energy barrier exceeding 30 kcal mol^{-1} owing to the geometric disadvantage of having the double bond inside the tight TS ring structure. In contrast, γ - and δ -unsaturated RO_2 radicals, forming 5-/6- and 6-/7-membered rings respectively, have energy barriers to ring closure that are only half as high, and where the ring closure on the nearer unsaturated carbon is typically more favorable as it does not have the rigid olefinic moiety in the TS cycle. For even wider separations between the peroxy radical site and the double bond, *i.e.* ϵ -unsaturated RO_2 and beyond, the larger ring can accommodate the double bond more easily without undue ring strain, and ring closure on the outer carbon becomes more favorable.

3.2 Substituents on the double bond

Alkyl substitution impacts ring closure by affecting the electron density in the attacked double bond, by stabilizing the product alkyl radical, and through steric hindrance in the cyclisation process. Generally, a more substituted double bond, and a more substituted product radical site yields higher rate coefficients, where a fully substituted double bond has a rate coefficient for ring closure 3 to 4 orders of magnitude faster at 298 K than an unsubstituted terminal olefine. Furthermore, the stereo-specificity of the substitution is important, as *cis*- and *trans*-substituted double bonds can have rate coefficients differing by



Table 1 Site- and stereo-specific ring closure reactions in unsaturated aliphatic RO₂ radicals. Indicated are the barrier height E_b (kcal mol⁻¹) and relative energy of the ring product E_{ring} (kcal mol⁻¹) at the CCSD(T)//M06-2X/aug-cc-pVTZ level of theory, as well as the rate coefficient at 298 K and as a Kooij expression $k(200-450\text{ K}) = A \times (T/K)^n \times \exp(-E_a/T)$ ($k(298\text{ K})$ and A in s⁻¹; E_a in K). If present, the first of the two stereo-indicators refers to the carbon bearing the peroxy radical group, the second to the >C=C< carbon atom forming the new C–O bond. IUPAC relative stereo-indicator (R^* , R^*) indicates the same stereo-indicator (i.e. (R,R) or (S,S)) on both chiral carbons, and (R^* , S^*) indicates differing stereo-indicators (i.e. (R,S) or (S,R)). "Total" rate coefficients for chiral compounds assume Boltzmann equilibrium between the enantiomers

Reactant	Product ring + stereo	E_b	E_{ring}	$k(298\text{ K})$	A	n	E_a
CH ₂ =CHCH ₂ OO*	4-Membered	30.8	19.0	4.6×10^{-11}	3.56×10^{-2}	4.38	13 532
	5-Membered	29.0	2.6	2.3×10^{-10}	5.15×10^1	3.08	13 032
CH ₂ =CHCH ₂ CH ₂ OO*	5-Membered	14.1	2.1	5.2×10^0	2.27×10^5	1.67	6019
	6-Membered	16.8	0.9	3.1×10^{-2}	1.31×10^7	0.96	7547
CH ₂ =CHCH ₂ CH(CH ₃)OO*	5-Membered (R^* , S^*)	15.1		1.3×10^0	3.78×10^4	1.92	6328
	5-Membered (R^* , R^*)	14.0		8.5×10^0	1.07×10^6	1.42	5919
	5-Membered (total)	14.0	2.0	9.7×10^0	7.71×10^4	1.87	5858
	6-Membered	16.8	0.8	8.4×10^{-2}	1.02×10^7	1.13	7462
CH ₂ =CHCH ₂ C(CH ₃) ₂ OO*	5-Membered	14.4	2.0	8.7×10^0	8.73×10^5	1.59	6135
	6-Membered	17.5	1.6	2.2×10^{-2}	1.48×10^8	0.71	7958
CH ₂ =CHCH(CH ₃)CH ₂ OO*	5-Membered	14.0		1.0×10^1	3.69×10^5	1.64	5908
	6-Membered	16.4		9.4×10^{-2}	6.43×10^6	1.10	7242
CH ₂ =CHC(CH ₃) ₂ CH ₂ OO*	5-Membered	14.0	1.1	1.0×10^1	1.19×10^5	1.83	5897
	6-Membered	16.1	-0.2	1.9×10^{-1}	8.61×10^6	1.15	7207
E-CH(CH ₃)=CHCH ₂ CH ₂ OO*	5-Membered	13.2	1.6	3.9×10^1	4.66×10^6	1.31	5713
	6-Membered	14.5	-1.3	1.4×10^0	2.90×10^7	0.85	6452
E-CH(CH ₃)=CHCH ₂ CH(CH ₃)OO*	5-Membered (R^* , S^*)	13.0		6.3×10^1	4.88×10^7	0.95	5656
	5-Membered (R^* , R^*)	14.1		8.5×10^0	3.78×10^6	1.31	6109
	5-Membered (total)	13.0	1.5	7.2×10^1	4.42×10^6	1.36	5600
	6-Membered (R^* , S^*)	15.8		1.7×10^{-1}	2.83×10^7	0.82	7031
Z-CH(CH ₃)=CHCH ₂ CH ₂ OO*	6-Membered (R^* , R^*)	14.5		1.6×10^0	6.48×10^7	0.71	6436
	6-Membered (total)	14.5	-1.2	1.7×10^0	5.90×10^6	1.12	6379
	5-Membered	13.2	0.6	3.7×10^1	1.81×10^7	1.06	5702
	6-Membered	16.0	-2.3	1.3×10^{-1}	5.48×10^7	0.66	7111
Z-CH(CH ₃)=CHCH ₂ CH(CH ₃)OO*	5-Membered (R^* , S^*)	13.9		1.2×10^1	8.43×10^5	1.61	6042
	5-Membered (R^* , R^*)	12.8		6.9×10^1	1.31×10^8	0.79	5653
	5-Membered (total)	12.8	0.5	8.1×10^1	1.96×10^6	1.49	5543
	6-Membered (R^* , S^*)	15.8		1.5×10^{-1}	5.57×10^7	0.69	7058
CH ₂ =C(CH ₃)CH ₂ CH ₂ OO*	6-Membered (R^* , R^*)	17.2		1.3×10^{-2}	1.74×10^7	0.86	7709
	6-Membered (total)	15.8	-2.2	1.6×10^{-1}	4.53×10^6	1.11	6991
	5-Membered	12.6	0.5	4.1×10^1	4.71×10^5	1.50	5335
	6-Membered	15.5	-2.4	3.4×10^{-1}	4.19×10^7	0.83	6953
E-CH(CH ₃)=C(CH ₃)CH ₂ CH ₂ OO*	5-Membered	11.0	-1.5	7.9×10^2	1.26×10^7	1.04	4647
	6-Membered	12.6	-2.8	7.0×10^1	2.60×10^8	0.64	5592
E-CH(CH ₃)=C(CH ₃)CH ₂ CH(CH ₃)OO*	5-Membered (R^* , S^*)	13.2		3.4×10^1	9.00×10^6	1.21	5770
	5-Membered (R^* , R^*)	11.0		1.3×10^3	2.22×10^8	0.69	4771
	5-Membered (total)	11.0	-1.4	1.3×10^3	1.64×10^7	1.10	4684
	6-Membered (R^* , S^*)	12.8		5.8×10^1	3.59×10^9	0.27	5811
Z-CH(CH ₃)=C(CH ₃)CH ₂ CH ₂ OO*	6-Membered (R^* , R^*)	14.6		2.1×10^0	3.23×10^9	0.22	6680
	6-Membered (total)	12.8	-2.7	6.0×10^1	6.66×10^8	0.55	5761
	5-Membered	11.4	-1.3	6.1×10^2	2.33×10^7	1.06	4947
	6-Membered	14.6	-2.4	1.9×10^0	1.42×10^8	0.69	6572
Z-CH(CH ₃)=C(CH ₃)CH ₂ CH(CH ₃)OO*	5-Membered (R^* , S^*)	12.7		6.9×10^1	4.60×10^5	1.69	5498
	5-Membered (R^* , R^*)	11.0		1.2×10^3	3.48×10^7	1.01	4773
	5-Membered (total)	11.0	-1.4	1.3×10^3	1.25×10^6	1.54	4668
	6-Membered (R^* , S^*)	16.7		5.1×10^{-2}	1.03×10^8	0.71	7594
C(CH ₃) ₂ =CHCH ₂ CH ₂ OO*	6-Membered (R^* , R^*)	14.3		3.3×10^0	7.67×10^7	0.80	6411
	6-Membered (total)	14.3	-2.7	3.4×10^0	1.90×10^7	1.02	6365
	5-Membered	11.4	-1.3	6.1×10^2	2.33×10^7	1.06	4947
	6-Membered	14.6	-2.4	1.9×10^0	1.42×10^8	0.69	6572
C(CH ₃) ₂ =CHCH ₂ CH(CH ₃)OO*	5-Membered (R^* , S^*)	13.0		6.0×10^1	1.37×10^7	1.25	5801
	5-Membered (R^* , R^*)	11.9		3.8×10^2	1.17×10^9	0.52	5340
C(CH ₃) ₂ =C(CH ₃)CH ₂ CH ₂ OO*	5-Membered (total)	11.9	0.4	4.4×10^2	1.77×10^7	1.22	5227
	6-Membered	13.5	-2.8	5.4×10^0	2.91×10^6	1.16	5897
	5-Membered	9.7	-2.7	7.5×10^3	2.85×10^8	0.64	4230
	6-Membered	11.0	-5.2	2.6×10^2	1.40×10^7	0.83	4657
CH ₂ =CHCH ₂ CH ₂ CH ₂ OO*	6-Membered	16.2	-0.7	1.6×10^{-1}	2.20×10^3	2.19	6559
	7-Membered	18.4	0.5	4.5×10^{-3}	2.35×10^6	1.23	8071
CH ₂ =C(CH ₃)CH ₂ CH ₂ CH ₂ OO*	6-Membered	14.5	-1.4	7.0×10^{-1}	7.64×10^2	2.27	5935
	7-Membered	16.6	0.2	6.0×10^{-2}	5.50×10^6	1.17	7448
E-CH(CH ₃)=CHCH ₂ CH ₂ CH ₂ OO*	6-Membered	15.4		4.3×10^{-1}	4.04×10^6	1.15	6736
	7-Membered	16.4		8.6×10^{-2}	1.01×10^8	0.69	7393
Z-CH(CH ₃)=CHCH ₂ CH ₂ CH ₂ OO*	6-Membered	15.0		1.9×10^0	3.93×10^8	0.46	6493
	7-Membered	17.3		1.8×10^{-2}	1.65×10^8	0.49	7669



Table 1 (continued)

Reactant	Product ring + stereo	E_b	E_{ring}	$k(298\text{ K})$	A	n	E_a
$C(CH_3)_2=CHCH_2CH_2CH_2OO^\bullet$	6-Membered	13.7		1.8×10^1	4.88×10^9	0.14	6027
	7-Membered	15.7		2.0×10^{-1}	3.78×10^8	0.34	6955
$E\text{-}CH(CH_3)=C(CH_3)CH_2CH_2CH_2OO^\bullet$	6-Membered	13.4		8.0×10^0	3.12×10^6	1.13	5757
	7-Membered	14.1		2.2×10^0	6.66×10^6	1.08	6284
$Z\text{-}CH(CH_3)=C(CH_3)CH_2CH_2CH_2OO^\bullet$	6-Membered	13.8		7.9×10^0	5.74×10^8	0.40	6071
	7-Membered	15.3		4.4×10^{-1}	9.96×10^6	0.96	6677
$C(CH_3)_2=C(CH_3)CH_2CH_2CH_2OO^\bullet$	6-Membered	12.2		8.6×10^1	6.52×10^8	0.43	5455
	7-Membered	12.8		1.2×10^1	2.08×10^5	1.54	5518
$CH_2=CHCH_2CH_2CH_2CH_2OO^\bullet$	7-Membered	17.7		3.4×10^{-3}	4.78×10^4	1.65	771
	8-Membered	16.4		9.9×10^{-3}	1.22×10^2	2.37	6825
$E\text{-}CH(CH_3)=CHCH_2CH_2CH_2CH(CH_3)OO^\bullet$	7-Membered (R^*,S^*)	16.7		2.0×10^{-2}	7.02×10^6	0.95	7479
	7-Membered (R^*,R^*)	16.7		1.2×10^{-2}	1.22×10^4	1.88	7316
	7-Membered (total)	16.7	0.1	3.2×10^{-2}	5.57×10^5	1.41	7370
	8-Membered (R^*,S^*)	14.7		6.8×10^{-2}	1.58×10^3	1.80	6053
	8-Membered (R^*,R^*)	14.7		7.2×10^{-2}	2.31×10^2	2.14	6038
	8-Membered (total)	14.7	1.0	1.4×10^{-1}	1.03×10^3	1.99	6040
$Z\text{-}CH(CH_3)=CHCH_2CH_2CH_2CH(CH_3)OO^\bullet$	7-Membered (R^*,S^*)	16.3		2.9×10^{-2}	4.29×10^5	1.26	7061
	7-Membered (R^*,R^*)	16.1		2.1×10^{-2}	5.51×10^2	2.14	6664
	7-Membered (total)	16.1	-0.8	4.9×10^{-2}	3.38×10^4	1.69	6871
	8-Membered (R^*,S^*)	15.9		1.5×10^{-2}	1.78×10^4	1.47	6677
	8-Membered (R^*,R^*)	16.0		1.0×10^{-2}	7.07×10^2	1.91	6563
	8-Membered (total)	15.9	0.1	2.5×10^{-2}	9.45×10^3	1.65	6631

Table 2 Site- and stereo-specific ring closure reactions in oxygenated unsaturated RO_2 radicals. Indicated are the barrier height E_b (kcal mol⁻¹) and the rate coefficient at 298 K and as a Kooij expression $k(T) = A \times (T/K)^n \times \exp(-E_a/T)$ for the temperature range 200–450 K (260–320 K for Chen *et al.*), with $k(298\text{ K})$ and A in s⁻¹ and E_a in K. Literature data on ring closure in cyclic RO_2 are not listed

Reactant	Product ring	E_b	$k(298\text{ K})$	A	n	E_a	Ref.
$CH_2=CHCH_2C(=O)OO^\bullet$	5-Membered	13.1	1.3×10^2	8.05×10^8	0.65	5757	This work
	6-Membered	17.4	1.1×10^{-1}	3.05×10^{11}	-0.16	8271	This work
$E\text{-}CH(OH)=CHCH_2CH_2OO^\bullet$	6-Membered	12.9	2.8×10^1	7.15×10^9	0.06	5879	Novelli <i>et al.</i> ²⁴
$Z\text{-}CH(OH)=CHCH_2CH_2OO^\bullet$	6-Membered	15.1	2.1×10^0	3.48×10^{16}	-2.28	7260	Novelli <i>et al.</i> ²⁴
$E\text{-}CH(OH)=C(CH_3)CH_2CH_2OO^\bullet$	6-Membered	10.6	9.2×10^2	6.24×10^8	0.40	4689	Novelli <i>et al.</i> ²⁴
$Z\text{-}CH(OH)=C(CH_3)CH_2CH_2OO^\bullet$	6-Membered	13.0	7.5×10^1	2.08×10^{14}	-1.42	6124	Novelli <i>et al.</i> ²⁴
$CH_2=C(CH_3)CH(OH)CH_2OO^\bullet$	6-Membered	16.2	9×10^{-1}	1.52×10^{16}	-1.71	8235 ^a	Vereecken and Peeters ²⁰
$CH_2=CHC(OH)(CH_3)CH_2OO^\bullet$	6-Membered	17.0	1×10^{-1}	1.29×10^{16}	-1.75	8710 ^a	Vereecken and Peeters ²⁰
$Z\text{-}C(CH_3)(CH_2ONO_2)=CHCH_2OO^\bullet$	5-Membered	27.5	2.1×10^{-9}	2.07×10^9	0.42	13062	Vereecken <i>et al.</i> ²⁵
$E\text{-}C(CH_3)(CH_2ONO_2)=CHCH_2OO^\bullet$	5-Membered	28.6	7.2×10^{-11}	3.43×10^4	1.90	13292	Vereecken <i>et al.</i> ²⁵
$CH_2=CHC(CH_3)(CH_2ONO_2)OO^\bullet$	5-Membered	26.6	2.2×10^{-8}	3.87×10^4	2.13	12028	Vereecken <i>et al.</i> ²⁵
$Z\text{-}CH(CH_2ONO_2)=C(CH_3)CH_2OO^\bullet$	5-Membered	27.4	3.1×10^{-9}	9.51×10^8	0.50	12845	Vereecken <i>et al.</i> ²⁵
$E\text{-}CH(CH_2ONO_2)=C(CH_3)CH_2OO^\bullet$	5-Membered	26.7	2.0×10^{-9}	5.01×10^7	0.74	12514	Vereecken <i>et al.</i> ²⁵
$CH_2=C(CH_3)C(CH_2ONO_2)OO^\bullet$	5-Membered	27.7	5.1×10^{-9}	1.03×10^7	1.31	12722	Vereecken <i>et al.</i> ²⁵
$CH_2=CHC(CH_3)(ONO_2)CH_2OO^\bullet$	5-Membered	14.4	7.7×10^0	1.04×10^7	1.19	6218	Vereecken <i>et al.</i> ²⁵
	6-Membered	14.7	2.1×10^0	2.69×10^7	0.91	6420	Vereecken <i>et al.</i> ²⁵
$CH_2=C(CH_3)CH(ONO_2)CH_2OO^\bullet$	5-Membered	13.4	4.1×10^1	9.60×10^9	0.12	5951	Vereecken <i>et al.</i> ²⁵
	6-Membered	15.7	1.5×10^0	2.47×10^{13}	-1.00	7384	Vereecken <i>et al.</i> ²⁵
$(R,R)\text{-}CH_2=C(CH_3)CH(C_2H_5O)CH_2CH(C_2H_5O)OO^\bullet$	6-Membered	14.1	3.9×10^0	2.73×10^{-5}		4087	Chen <i>et al.</i> ²⁷
	7-Membered	16.4	6.6×10^{-1}	3.62×10^{-4}		5384	Chen <i>et al.</i> ²⁷
$(R,S)\text{-}CH_2=C(CH_3)CH(C_2H_5O)CH_2CH(C_2H_5O)OO^\bullet$	6-Membered	15.1	3.0×10^0	1.75×10^{10}		6568	Chen <i>et al.</i> ²⁷
	7-Membered	16.5	3.2×10^{-1}	6.22×10^{10}		7609	Chen <i>et al.</i> ²⁷
$CH_2=C(CH_3)CH(C_2H_5O)CH_2CH_2C(=O)CH_2OO^\bullet$	9-Membered	16.2	3.5×10^{-2}	1.98×10^9		7297	Chen <i>et al.</i> ²⁷
$(S,S)\text{-}CH_2=C(CH_3)CH(C_4H_7O)CH(CHO)OO^\bullet$	5-Membered	12.8	2.3×10^2	1.04×10^{11}		5640	Chen <i>et al.</i> ²⁷
	6-Membered	14.1	2.0×10^1	5.50×10^{10}		6178	Chen <i>et al.</i> ²⁷
$(S,R)\text{-}CH_2=C(CH_3)CH(C_4H_7O)CH(CHO)OO^\bullet$	5-Membered	13.3	4.5×10^1	4.14×10^{10}		5943	Chen <i>et al.</i> ²⁷
	6-Membered	16.0	1.4×10^0	1.88×10^{11}		7440	Chen <i>et al.</i> ²⁷
$(S,S)\text{-}CH_2=C(CH_3)CH(CH_3)CH(CHO)OO^\bullet$	5-Membered	11.7	3.3×10^2	3.03×10^6	1.29	4906	This work
	6-Membered	14.4	1.2×10^1	1.25×10^{11}	-0.11	6692	This work
$(S,R)\text{-}CH_2=C(CH_3)CH(CH_3)CH(CHO)OO^\bullet$	5-Membered	10.5	1.7×10^3	7.73×10^6	1.16	4482	This work
	6-Membered	12.8	4.9×10^1	2.68×10^8	0.69	5793	This work

^a Temperature-dependent rate coefficients re-evaluated from original B3LYP/6-31G(d,p) quantum chemical data and reported barrier heights.

over an order of magnitude. The strong dependence of the ring closure on the substitution pattern suggests that this reaction class will also be strongly influenced by hetero-substitution on the double bond. Novelli *et al.*²⁴ calculated

rates for ring closure reactions in enol- RO_2 radicals, where the predicted rates (Table 2) indicate that an OH group enhances the reaction rate by an order of magnitude compared to a methyl group.



3.3 Substituents adjacent to the peroxy radical moiety

The presence of alkyl groups adjacent to the peroxy radical moiety has comparatively little influence. For ring closure to 5- or 6-membered rings, we find the spread of the rate coefficients for primary, secondary, and tertiary alkylperoxy radicals to be within a factor of four at 298 K, *i.e.* the rate coefficients are within a factor of two of the averaged rate coefficient for $-\text{CH}_2\text{OO}^\bullet$, $-\text{CHCH}_3\text{OO}$, and $-\text{C}(\text{CH}_3)_2\text{OO}^\bullet$ moieties. This is in strong contrast to *e.g.* H-migration in RO_2 radicals, where the impact of alkyl substitution around the peroxy radical moiety on the reaction rate can exceed a factor 20.¹¹

The small impact of alkyl substitution around the radical moiety allows for a significant reduction of the complexity of the SAR, and concomitantly the needed number of calculations to derive the SAR. For larger molecules, we have thus only explicitly characterized the $-\text{CH}_2\text{OO}^\bullet$ moieties, and derive the SAR to predict the expected average across primary, secondary, and tertiary peroxy radicals.

The oxo group in acylperoxy radicals, $-\text{C}(=\text{O})\text{OO}^\bullet$, was reported to accelerate RO_2 radical H-migration reactions by an order of magnitude.^{11,43,44} Exploratory calculations on acylperoxy radicals (Table 2) indicate that the acyl group also accelerates ring closure by up to a factor 25.

3.4 Spectator substituents

The scarce data in Table 1 suggests that substituents on carbons not implanted on the double bond or the carbon bearing the peroxy radical group have only a moderate impact on the predicted rate coefficients. As a first approximation, it appears one can consider substituents away from the active sites as spectators, and ignore their impact on the reaction. It is currently unclear whether longer-range effects such as H-bonding or strong electron-donating/withdrawing groups can invalidate this simplification. Vereecken and Peeters²⁰ reported ring closure reactions in isoprene-derived unsaturated RO_2 radicals with spectator OH-substitutions. Despite the much lower level of theory in that study, the predicted rate coefficients are within a factor 3 of the current predictions for RO_2 when ignoring the OH spectator, bringing further support for the low impact of spectator substituents. The results by Vereecken *et al.*²⁵ on nitrate-substituted RO_2 radicals also partially support the low impact of spectator groups. However, one 6-membered ring closure with a nitrate group shows a larger difference of a factor of 50 relative to the non-oxygenated RO_2 studied in this work. Data on β -carbonyl-substituted RO_2 radicals by us and by Chen *et al.*²⁷ (see Table 2) likewise suggests an impact of over an order of magnitude. Finally, some unsaturated RO_2 radicals with a second unsaturation in β -position to the peroxy radical, $\text{CH}_2=\text{CH}-\text{C}(=\text{CH}_2)\text{CH}_2\text{OO}^\bullet$, are reported to be less reactive²⁶ towards ring closure than suggested by the SAR. Our current treatment of spectator groups may thus be too simplified and further work on the impact of substituents is needed to refine the SAR.

3.5 Structure–activity relationship

The reactivity trends described above are used to derive a structure–activity relationship for ring closure in unsaturated

RO_2 radicals with aliphatic substituents. The reaction rates are considered to be mainly determined by ring size, site-specificity on the double bond, and substitution on the double bond. Compared to the uncertainty of the predictions, the impact of other substituents is considered to be not yet sufficiently clear to warrant a more complex SAR with additional parameters at this time. Finally, we assume that the impact of an alkyl group is not strongly dependent on the length of the substituent alkyl chain, as was already found in SARs for other reaction classes.^{4,11,25,45,46}

For ring closure on the inner carbon of the double bond, the stereo-specificity on the outer double bonded carbon has little influence and is not considered explicitly in the SAR. For these reactions, the reactivity trend prediction could potentially be even further simplified by considering only 4 template categories, *i.e.* a double bond with 0, 1, 2 or 3 additional alkyl substituents. However, this reduction does not work well for ring closure on the outer carbon, where the stereo-substitution and positioning of the substituents have a larger impact. To keep the SAR setup similar for both types of ring closure, we choose at this time to use categories with site-specific substitution on the double bond for both closures. This will also aid in future extensions to oxygenated compounds, where the impact of site-specificity is likely to be more pronounced.

Table 3 summarizes the SAR as a lookup table, listing room temperature $k(298\text{ K})$ and rate coefficients $k(T)$ for $T = 200\text{--}450\text{ K}$. Fig. 1 compares the explicit rate predictions in Table 1 against the SAR predictions; to our knowledge there are no direct experimental data available against which we can validate the SAR. The SAR captures the reactivity trends in the source data well (goodness of fit). Considering the uncertainty on the theoretical source data (of the order of a factor 3 to 5), and the uncertainty caused by omitting some of the minor influences on the rate coefficient, we estimate an overall uncertainty on the SAR predictions of a factor of 10 for aliphatic unsaturated RO_2 . For oxygenated RO_2 , Table 2 suggests that the deviation can be significantly larger, and care should be taken when trying to apply the SAR outside its field of applicability.

4. Subsequent chemistry of the cycloperoxide alkyl product radicals

In general, two reaction routes are accessible for the cycloperoxide alkyl radicals formed after ring closure. The first is recombination with O_2 , forming a cycloperoxide peroxy radical (called cycOO- RO_2 hereafter). The rate coefficient for alkyl + O_2 recombination reactions has been measured to be 10^{-12} to $10^{-11}\text{ cm}^3\text{ molecule}^{-1}\text{ s}^{-1}$,⁴⁷ which under atmospheric conditions yields a pseudo-first order rate coefficient of $\sim 10^7\text{ s}^{-1}$. An alternative reaction (R3) is re-arrangement to an epoxide and an alkoxy radical:

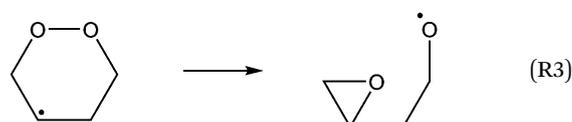


Table 3 Structure–activity relationship for ring closure reactions in unsaturated RO₂ radicals forming 5- to 8-membered rings, for various substitutions on the double bond (CH₃ is considered equivalent to an alkyl chain of any length). Indicated are the barrier height E_b (kcal mol⁻¹) and the rate coefficient at 298 K and as a Kooji expression $k(200\text{--}450\text{ K}) = A \times (T/K)^n \times \exp(-E_a/T)$ ($k(298\text{ K})$ and A in s⁻¹; E_a in K). Ring closure in β -unsaturated RO₂ is slow and not competitive

Olefinic substituents	$k(298\text{ K})$	A	n	E_a	Olefinic substituents	$k(298\text{ K})$	A	n	E_a
γ -Unsaturated RO ₂ radicals (CH ₂ =CHCH ₂ CH ₂ OO)					6-Membered ring closure				
5-Membered ring closure					6-Membered ring closure				
CH ₂ =CH-	6.6×10^0	2.84×10^5	1.67	6019	CH ₂ =CH-	3.9×10^{-2}	1.64×10^7	0.96	7547
CH(CH ₃)=CH-	4.8×10^1	1.15×10^7	1.19	5708	<i>E</i> -CH(CH ₃)=CH-	1.8×10^0	3.64×10^7	0.85	6452
C(CH ₃) ₂ =CH-	3.3×10^2	1.47×10^8	0.91	5414	<i>Z</i> -CH(CH ₃)=CH-	1.3×10^{-1}	6.88×10^7	0.66	7111
CH ₂ =C(CH ₃)-	5.2×10^1	5.91×10^5	2.27	5935	C(CH ₃) ₂ =CH-	4.8×10^0	2.19×10^8	0.50	6100
CH(CH ₃)=C(CH ₃)-	8.7×10^2	2.15×10^7	1.05	4797	CH ₂ =C(CH ₃)-	4.3×10^{-1}	5.26×10^7	1.17	7448
C(CH ₃) ₂ =C(CH ₃)-	9.4×10^3	3.57×10^8	0.64	4230	<i>E</i> -CH(CH ₃)=C(CH ₃)-	8.8×10^1	3.26×10^8	0.64	5592
					<i>Z</i> -CH(CH ₃)=C(CH ₃)-	2.4×10^0	1.78×10^8	0.69	6572
					C(CH ₃) ₂ =C(CH ₃)-	3.2×10^2	1.76×10^7	0.83	4657
δ -Unsaturated RO ₂ radicals (CH ₂ =CHCH ₂ CH ₂ CH ₂ OO)					7-Membered ring closure				
6-Membered ring closure					7-Membered ring closure				
CH ₂ =CH-	2.0×10^{-1}	2.76×10^3	2.19	6559	CH ₂ =CH-	5.6×10^{-3}	2.95×10^6	1.23	8071
CH(CH ₃)=CH-	1.1×10^0	5.00×10^7	0.81	6614	<i>E</i> -CH(CH ₃)=CH-	1.1×10^{-1}	1.27×10^8	0.69	7393
C(CH ₃) ₂ =CH-	2.3×10^1	6.12×10^9	0.14	6027	<i>Z</i> -CH(CH ₃)=CH-	2.2×10^{-2}	2.07×10^8	0.49	7669
CH ₂ =C(CH ₃)-	8.7×10^{-1}	9.59×10^2	2.27	5935	C(CH ₃) ₂ =CH-	2.4×10^{-1}	4.74×10^8	0.34	6955
CH(CH ₃)=C(CH ₃)-	1.0×10^1	5.30×10^7	0.77	5914	CH ₂ =C(CH ₃)-	7.6×10^{-2}	6.90×10^6	1.17	7448
C(CH ₃) ₂ =C(CH ₃)-	1.1×10^2	8.18×10^8	0.43	5455	<i>E</i> -CH(CH ₃)=C(CH ₃)-	2.8×10^0	8.35×10^6	1.08	6284
					<i>Z</i> -CH(CH ₃)=C(CH ₃)-	5.6×10^{-1}	1.25×10^7	0.96	6677
					C(CH ₃) ₂ =C(CH ₃)-	1.5×10^1	2.61×10^5	1.54	5518
ϵ -Unsaturated RO ₂ radicals (CH ₂ =CHCH ₂ CH ₂ CH ₂ CH ₂ OO)					8-Membered ring closure				
7-Membered ring closure					8-Membered ring closure				
CH ₂ =CH-	4.2×10^{-3}	6.00×10^4	1.65	7714	CH ₂ =CH-	1.2×10^{-2}	1.53×10^2	2.37	6825
CH(CH ₃)=CH-	2.7×10^{-2}	2.64×10^7	0.82	7550	<i>E</i> -CH(CH ₃)=CH-	4.1×10^{-1}	1.86×10^2	2.33	5788
C(CH ₃) ₂ =CH-	3.5×10^{-1}	2.30×10^9	0.21	7098	<i>Z</i> -CH(CH ₃)=CH-	4.5×10^{-2}	2.53×10^3	1.85	6400
CH ₂ =C(CH ₃)-	2.6×10^{-2}	4.42×10^4	1.63	7036	C(CH ₃) ₂ =CH-	1.0×10^0	1.73×10^3	1.90	5432
CH(CH ₃)=C(CH ₃)-	3.7×10^{-1}	1.21×10^7	0.89	6663	CH ₂ =C(CH ₃)-	1.5×10^{-1}	4.09×10^2	2.27	6215
C(CH ₃) ₂ =C(CH ₃)-	4.0×10^0	5.19×10^7	0.70	6078	<i>E</i> -CH(CH ₃)=C(CH ₃)-	1.8×10^1	9.19×10^2	2.18	4861
					<i>Z</i> -CH(CH ₃)=C(CH ₃)-	9.8×10^{-1}	4.40×10^1	2.55	5471
					C(CH ₃) ₂ =C(CH ₃)-	7.2×10^1	4.23×10^0	2.78	3870

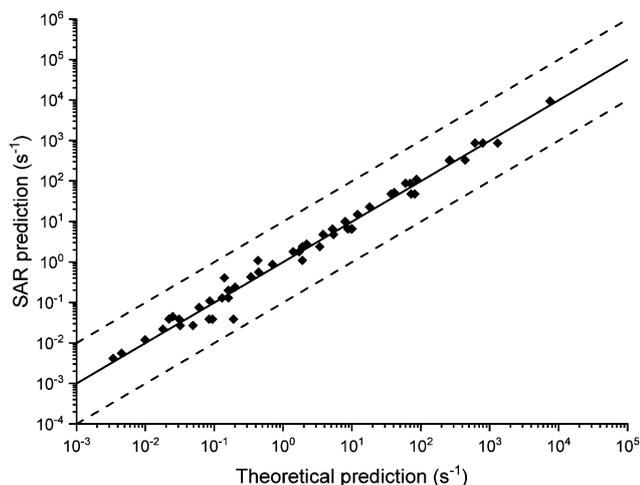


Fig. 1 Goodness of fit for the structure–activity relationship relative to the calculated (total) rate coefficients in Table 1. The solid line depicts 1:1 correspondence, the dashed lines delineate a factor of 10 uncertainty.

These rearrangements have been theoretically characterized by Møller *et al.*⁴⁸ to have high reaction barriers of 13 to 17 kcal mol⁻¹ for the substitution patterns in this work, and Vereecken and Peeters²⁰ argue that the geometric constraints caused by the ring structure will further hamper this reaction,

with a calculated energy barrier exceeding 25 kcal mol⁻¹ for a 6-membered ring. The theoretical work by Møller *et al.*⁴⁸ predicts rate coefficients $< 10^2$ s⁻¹ for non-cyclic non-oxygenated β -hydroperoxy alkyl radicals, making epoxide-alkoxy radical formation negligible compared to O₂ addition under atmospheric conditions. For multi-substituted β -hydroperoxy alkyl radicals, however, rates as high as 10^7 to 10^{10} s⁻¹ are predicted, which would dominate recombination with O₂. Møller *et al.* summarize the structural requirements for fast epoxidation reactions, concluding that an OH group is needed on the alkyl radical site, and another oxygenated group. For ring closure reactions, this implies that epoxidation would only be important for RO₂ with a double bond carrying an OH substituent (*i.e.* an enol) and simultaneously another oxygenated group on the other carbon of the double bond. Such substitution patterns make up only a very small fraction of the atmospherically relevant unsaturated RO₂ radicals, and are not considered further at this time.

5. H-Migration in the product cycloperoxide-alkylperoxy radicals

To study the subsequent fate of the cycOO-RO₂ radicals, we examined the migration of H-atoms implanted on the cycloperoxide



ring towards the radical oxygen. CycOO-RO₂ radicals formed from unsaturated-RO₂ by ring closure on the outer olefinic carbon have the peroxy radical group implanted directly on the ring, while ring closure on the inner carbon has the peroxy group on the first carbon of a ring substituent. One should also distinguish between migration of an aliphatic H-atom, *versus* abstraction of an H-atom adjacent to the peroxide functionality.¹¹ As already shown earlier, product radicals with an α -OOR alkyl radical site are unstable,^{49,50} and the peroxide ring can be expected to break after such H-migrations. At this time, we do not quantify the impact of substituents on H-migration, though it is briefly discussed below.

5.1 Reactivity trends

The barrier heights and rate coefficients for migrations of H-atoms on the cycloperoxide ring to the peroxy radical site are provided in Table 4, while some molecular variations are listed in Table 5. 1,3- and 1,4-H-migrations have too high a barrier¹¹ and are not considered. The accessible H-atoms are all on the same side as the peroxy radical moiety (*syn* side), except for a few structures where 1,5-shift of *anti*-H-atoms is geometrically accessible. H-Migration of an α -OOR H-atom, *i.e.* next to the peroxide functionality in the ring, leads to ring breaking forming a carbonyl-alkoxy radical with an energy 35 to 40 kcal mol⁻¹ below the cycOO-RO₂. In most cases this

ring breaking is spontaneous, but for cycOO-RO₂ with the -OO[•] peroxy group implanted on the ring, our current level of theory suggests that the α -OOR alkyl radical product may have a very shallow potential energy well. We have not examined whether this is an artifact of the methodology, as breaking of the ring is clearly much faster than any competing reactions, including O₂ addition on the radical site, rendering the issue moot. Except for the formation of carbonyl-alkoxy radicals by migration of α -OOR H-atoms and their subsequent fragmentation, all alkyl radical products formed in the reactions studied here are expected to rapidly recombine with O₂, forming a new cyclic RO₂ with the radical moiety implanted directly on the ring.

The rate of reaction can be seen to be strongly dependent on the size of the ring, the position of abstraction in that ring, and the span of the H-migration. Compared to similar H-migrations in non-cyclic hydrocarbons,^{10,11} the predicted rate varies by as much as 2 orders of magnitude due to the differences in ring strain, barrier height and tunneling. Fig. 2 shows the ratio $k_{\text{cyclic}}(T)/k_{\text{noncyclic}}(T)$ of the predicted rates for the cycOO-RO₂ (Table 4) over the predictions by the SAR by Vereecken and Nozière,¹¹ where we used the SAR-predicted rate of α -OOH-substituted H-atoms as a proxy for the α -OOR H-atoms in the cycOO-RO₂, and where the SAR predictions were corrected by a factor of 2 to account for the inaccessible *anti*-H-atoms. As can be seen, the cycOO-RO₂ have significantly different rate

Table 4 Site- and stereo-specific H-migration reactions in cycloperoxide-alkylperoxy radicals (cycOO-RO₂). Indicated are the barrier height E_b (kcal mol⁻¹) and relative energy of the product E_{product} (kcal mol⁻¹) at the CCSD(T)//M06-2X-D3 level of theory (products at M06-2X/aug-cc-pVTZ level of theory), as well as the rate coefficient at 298 K and as a Kooij expression $k(200-450 \text{ K}) = A \times (T/K)^n \times \exp(-E_a/T)$ ($k(298 \text{ K})$ and A in s⁻¹; E_a in K)

Reactant	H-migration span	E_b	E_{product}	$k(298 \text{ K})$	A	n	E_a
	1,5-H _a (<i>syn</i> -side)	23.8	15.4	4.3×10^{-5}	5.48×10^{-16}	8.39	6771
	1,6-H _b (α -peroxide)	21.2	8.4	1.2×10^{-2}	1.17×10^{-36}	14.93	1998
	1,5-H _a (<i>syn</i> -side)	24.5	14.1	2.5×10^{-5}	2.92×10^{-29}	12.84	5370
	1,6-H _b	25.0	13.6	5.0×10^{-6}	1.55×10^{-33}	13.91	4735
	1,7-H _c (α -peroxide)	26.2	-34.6	1.9×10^{-5}	1.73×10^{-113}	39.78	-6613
	1,5-H _a (<i>syn</i> -side)	22.5	14.4	1.3×10^{-3}	1.15×10^{-23}	10.97	4864
	1,5-H _a (<i>anti</i> -side)	28.9		3.4×10^{-7}	3.07×10^{-55}	21.18	2988
	1,6-H _b	23.7	11.9	9.9×10^{-5}	4.45×10^{-36}	14.83	3670
	1,7-H _c	25.3	13.7	4.1×10^{-6}	1.42×10^{-30}	12.79	4924
	1,7-H _d (α -peroxide)	26.4	-37.0	1.7×10^{-5}	1.17×10^{-112}	39.42	-6621
	1,5-H _a (<i>syn</i> -side)	23.8	12.8	1.4×10^{-4}	2.20×10^{-35}	14.83	4041
	1,5-H _a (<i>anti</i> -side)	26.0		7.2×10^{-6}	1.00×10^{-44}	17.86	3650
	1,6-H _b	21.2	12.3	1.9×10^{-3}	5.31×10^{-19}	9.19	4923
	1,7-H _c	24.7	10.1	1.9×10^{-5}	5.15×10^{-45}	17.31	2242
	1,8-H _d	27.1	11.8	5.3×10^{-7}	2.04×10^{-47}	18.05	2910
	1,7-H _e (α -peroxide)	28.1	-37.4	1.2×10^{-6}	6.70×10^{-121}	41.96	-7163
	1,5-H _a (α -peroxide)	28.1	14.6	1.6×10^{-7}	1.02×10^{-45}	18.19	4670
	1,5-H _a	24.2	13.8	4.5×10^{-5}	1.54×10^{-27}	12.23	5340
	1,6-H _b (α -peroxide)	28.3	15.2 ^a	2.0×10^{-7}	1.03×10^{-39}	16.12	5219
	1,5-H _a	25.1	13.9	2.1×10^{-5}	1.65×10^{-31}	13.64	5241
	1,6-H _b	26.2	15.6	4.3×10^{-6}	5.50×10^{-23}	10.54	6308
	1,7-H _c (α -peroxide)	26.3	11.2 ^a	3.1×10^{-6}	1.17×10^{-48}	18.73	2688

^a The α -OOR alkyl radical has a very small barrier to decomposition at the current level of theory.



Table 5 Site- and stereo-specific H-migration reactions in structural variations of cyc-RO₂, cycOO-RO₂ and cycOO-RO radicals; these compounds may not have any atmospheric formation pathway. Indicated are the barrier height E_b (kcal mol⁻¹) and relative energy of the product E_{product} (kcal mol⁻¹) at the CCSD(T)//M06-2X-D3 level of theory (products at M06-2X-D3/aug-cc-pVTZ level of theory), as well as the rate coefficient at 298 K and as a Kooij expression $k(200-450 \text{ K}) = A \times (T/K)^n \times \exp(-E_a/T)$ ($k(298 \text{ K})$ and A in s⁻¹; E_a in K)

Reactant	H-Migration span	E_b	E_{product}	$k(298 \text{ K})$	A	n	E_a
	1,5-H _a	27.7		6.4×10^{-6}	6.79×10^{-112}	39.98	-4849
	1,6-H _b	27.5		1.8×10^{-5}	1.19×10^{-104}	37.25	-4820
	1,5-H _a	29.2	15.1	3.6×10^{-8}	1.43×10^{-46}	18.50	5048
	1,6-H _b (α-peroxide)	29.1	-38.9	1.4×10^{-6}	1.84×10^{-126}	44.21	-7218
	1,5-H _a	25.7	16.1	4.5×10^{-6}	4.57×10^{-22}	10.54	6915
	1,6-H _b	30.4	17.7	5.9×10^{-9}	1.75×10^{-39}	16.31	6747
	1,6-H _b (α-peroxide)	33.4	-40.0	5.5×10^{-8}	6.55×10^{-131}	45.12	-7752
	1,5-H _a (α-peroxide)	12.4		6.8×10^3	5.19×10^{-30}	13.39	17
	1,5-H _a	9.0		7.6×10^5	8.17×10^{-14}	7.97	518
	1,6-H _b (α-peroxide)	13.5		1.8×10^2	1.19×10	3.85	5044
	1,5-H _a	10.9		4.4×10^4	5.68×10^{-15}	8.50	1474
	1,6-H _b	11.0		2.7×10^4	7.66×10^{-24}	11.13	-3
	1,7-H _c (α-peroxide)	12.8		3.9×10^2	4.17×10^{-3}	4.43	4114

coefficients than non-cyclic RO₂, where especially below 350 K the rate coefficient ratio is strongly temperature dependent due to the differences in energy, entropy, and tunneling between cyclic and non-cyclic RO₂. Even within a specific reaction class, the ratio $k_{\text{cyclic}}(T)/k_{\text{noncyclic}}(T)$ varies significantly, with e.g. 1,5- or 1,6-H-migration rates in cycloperoxide-CH₂OO• radicals being

significantly faster, about similar, or significantly slower than their non-cyclic counterpart (Fig. 2). Likewise, the temperature-dependence of the ratio is very different between α-OOR-1,6-H-migrations compared to α-OOR-1,7-H-migrations. For cycloperoxide-OO• radicals, there can be many orders of magnitude difference between the rates at the lower temperatures, even within a given

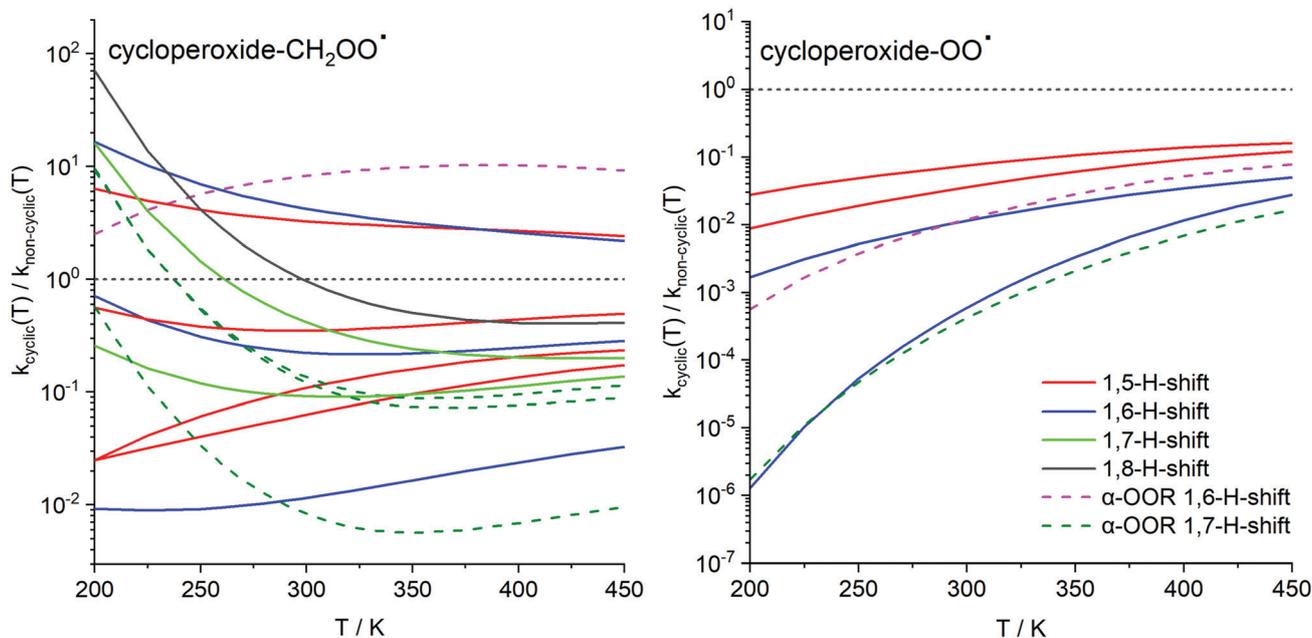


Fig. 2 Ratio of the rate coefficient for H-migration in cycloperoxide-RO₂ radicals against SAR-predictions¹¹ for analogous H-migrations in oxygenated non-cyclic RO₂. Dashed line: 1:1 correspondence. Left panel: -CH₂OO• radical site implanted on the ring. Right panel: Peroxy radical -OO• implanted directly on the ring.



reaction class. Hence, while on average the rate coefficients for cycOO-RO₂ are only a factor ~4 below those of non-cyclic RO₂, the large scatter of a factor 25 indicates that reactivity trends are not transferable between cyclic and non-cyclic RO₂.

To investigate the impact of the peroxide group in the ring, we examined some molecular isomers of the cycOO-RO₂ radicals, shown in Table 5; note that these compounds can not be formed by the mechanism of RO₂ ring closure described above, and these structures may not even have atmospheric formation channels. The position of the peroxide group in the ring changes which H-migration spans lead to an α-OOR radical product, which directly affects the rate of reaction. Comparing the data from Table 4 against Table 5, we also find an order of magnitude difference in rate for the 1,5-alkyl-H-migration out of a seven-membered ring that differs only in the position of the (remote) peroxide group, indicating that the position of the peroxide group affects the barrier height even if not directly involved in the reaction. Comparing against a cyclohexylperoxy radical, we find that the corresponding 6-membered cycloperoxide-peroxy radical isomers both have higher barrier heights, likely caused by the increased ring strain induced by the different bond lengths and bond angles around the endo-cyclic oxygen atoms. We conclude that the presence of a peroxide functionality, and its precise position in the ring, can have a large impact on the reaction rate; at the present time we do not have sufficient data to systematically characterize such influence.

5.2 Towards a structure–activity relationship

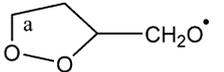
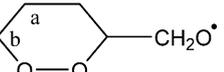
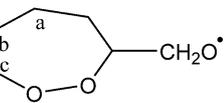
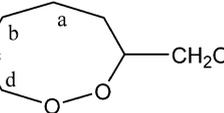
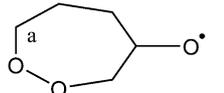
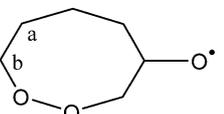
The ratios shown in Fig. 2 do not show a correlation with ring size or migration span that can be used to derive an easily quantifiable

trend. As such, it does not appear feasible to build an SAR for migration of H-atoms implanted on the ring of cycOO-RO₂ based on the reactivity trends for non-cyclic RO₂, at least not below ~400 K. At this moment we also don't have enough data available to fully examine the impact of the position of the two O-atoms in the ring, or to quantify the impact of substituents. We surmise that an additional α-alkyl, α-OH or α-OOH substituent would have a similar impact on cyclic RO₂ as in aliphatic RO₂.¹¹ Such a substituent must be in anti-position to avoid blocking H-migration, *i.e.* stereo-specificity must be rigorously accounted for. In contrast, reactivity trends¹¹ for *e.g.* double bonds or carbonyl groups which affect ring strain and require the appropriate relative orientation to allow allylic/vinoxy stabilization of the product radical, will likely never be transferable between cyclic and non-cyclic RO₂. As migration of H-atoms implanted on the ring appears on average slower than in non-cyclic RO₂, one should consider that H-abstraction from the (oxygenated) substituents, if present, instead of from the ring itself could be the more favorable pathway. If so, a useful SAR for cyclic RO₂ must be based on an extensive set of data on substituted rings, which is outside the scope of this paper. Finally, we note that the slow H-migration pathways could make bimolecular loss processes dominant; this includes reaction with R'O₂, HO₂, NO, NO₂, NO₃, OH, *etc.* For many of these processes, predictive structure–activity relationships are available.^{51–53}

6. Fate of cycloperoxide-alkoxy radicals

Given the rather slow reaction rates of autoxidation predicted above for the product cycOO-RO₂ radicals, they have a higher

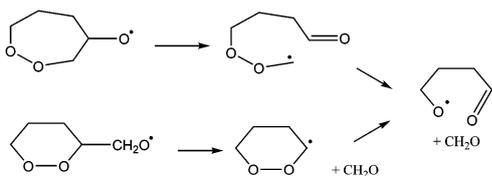
Table 6 Site- and stereo-specific H-migration reactions in cycloperoxide-alkoxy radicals (cycOO-RO). Indicated are the barrier height E_b (kcal mol⁻¹) at the CCSD(T)//M06-2X-D3 level of theory (products at M06-2X-D3/aug-cc-pVTZ level of theory), as well as the rate coefficient at 298 K and as a Kooij expression $k(200–450\text{ K}) = A \times (T/K)^n \times \exp(-E_a/T)$ ($k(298\text{ K})$ and A in s⁻¹; E_a in K)

Reactant	H-Migration span	E_b	$k(298\text{ K})$	A	n	E_a
	1,5-H _a (α-peroxide)	12.7	5.7×10^3	2.78×10^{-28}	12.74	153
	1,5-H _a 1,6-H _b (α-peroxide)	12.9 15.9	5.4×10^3 9.1×10^0	2.63×10^{-33} 7.12×10^{-28}	14.42 12.72	-440 2307
	1,5-H _a 1,6-H _b 1,7-H _c (α-peroxide)	10.9 14.3 16.8	8.8×10^4 4.3×10^2 1.1×10^1	1.32×10^{-17} 8.79×10^{-33} 1.39×10^{-51}	9.24 14.31 20.27	703 488 -1194
	1,5-H _a 1,6-H _b 1,7-H _c (α-peroxide) 1,6-H _d	8.8 15.6 14.2 15.2	1.1×10^6 5.7×10^1 1.5×10^2 1.4×10^1	7.77×10^{-9} 6.84×10^{-44} 6.67×10^{-30} 1.28×10^{-23}	6.28 17.90 13.21 11.09	937 -443 918 2345
	1,5-H _a (α-peroxide)	10.2	1.1×10^5	9.41×10^{-10}	6.81	1909
	1,5-H _a 1,6-H _b (α-peroxide)	9.1 8.8	3.3×10^6 1.2×10^6	2.87×10^{-14} 2.31×10^{-8}	8.26 6.32	246 1310



likelihood of undergoing bimolecular reactions than non-cyclic RO₂. In their reaction with R'O₂, HO₂, OH, NO, or NO₃ radicals, this would (partially) lead to the formation of cycloperoxide-alkoxy radicals, called cycOO-RO radicals henceforth. Alkoxy radicals with a sufficiently long chain can undergo fast H-migration or decomposition reactions, with rate coefficients often exceeding the rate of reaction with O₂ by several orders of magnitude.^{45,54,55} For the compounds studied here, the main competition for the fate of the cycOO-RO radicals would be between H-migration and decomposition, with HO₂ formation by reaction with O₂ being too slow.⁵⁴ To assess the impact of H-migration and hence further autoxidation in cycOO-RO radicals, we present a set of calculations in Table 6 for alkoxy radicals derived from RO₂ ring closure reactions, as well as some structural variations in Table 5. The rate coefficients for migration of H-atoms on the ring, below 10⁴ s⁻¹ for most H-migrations studied, are somewhat low compared to those in non-cyclic alkoxy radicals; for comparison, the 1,5-H-shift in 1-butoxy has a rate coefficient of ~3 × 10⁵ s⁻¹.^{40,55-57} The lower rates are mostly due to the increased ring strain in the bicyclic transition state, exacerbated by the position of the radical oxygen implanted either directly on the ring or with only 1 intervening carbon. Consequently, the highest rates were found for the largest rings and hence lowest ring strain. We have not investigated the impact of substituents, but the H-migration SAR by Vereecken and Peeters⁴⁵ may offer some transferable information on the impact of substituents, e.g. for α-OH substituents. Similar to H-migration in cyc-RO₂ radicals, we surmise that H-migration from any (oxygenated) substituents, if present, may be more favorable than from the carbons in the cycle, but did not explore this further. The H-migration is followed by O₂ addition under atmospheric conditions, forming a hydroxyperoxyradical.

For the alternative fate by decomposition, rates for cycOO-RO are expected to be well predicted by the SAR by Vereecken and Peeters,⁵⁸ and its recent update in Novelli *et al.*,⁵⁹ this SAR includes the impact of the ring. For the cycOO-RO in Table 6 that decompose by CH₂O elimination, barrier heights ≤ 7.5 kcal mol⁻¹ and rates $k(298\text{ K}) \geq 8 \times 10^7\text{ s}^{-1}$ are predicted, while cycOO-RO with the radical O-atom implanted on the ring preferentially break the •OC-COO bond, with SAR-predicted barriers of ≤ 8.5 kcal mol⁻¹ and $k(298\text{ K}) \geq 1 \times 10^7\text{ s}^{-1}$. In both cases, an α-OOR alkyl radical is formed, which are known to decompose without (significant) barrier to a carbonyl and an alkoxy radical;^{49,50} both classes of cycOO-RO in Table 6 will thus lead to two fragments: HCHO and a non-cyclic carbonyl-alkoxy radical, as shown below.



Given the high rate of decomposition compared to H-migration in cycOO-RO, the main fate of most cycOO-RO in Table 6 is fragmentation and decyclization of the ring originally formed in

the unsaturated RO₂ ring closure. Only for the largest rings studied, where ring strain in the H-migration TS is lowest, H-migration may contribute for a non-negligible fraction.

Comparing the data for cycOO-RO radical isomers different only by the position of the peroxide group in the ring (Tables 5 and 6), we find that the site-specificity of the peroxide group can affect the rate coefficient by over an order of magnitude, with sometimes large differences in barrier height for otherwise comparable processes. As for the H-migration in cycOO-RO₂ radicals, we thus conclude that the rate coefficients are isomer-specific and do not readily generalize to generic cycloperoxides.

7. Conclusions

The ring closure reaction in aliphatic mono-unsaturated RO₂ radicals was examined by theoretical calculations, and a SAR was formulated to predict rate coefficients for use in mechanism development. From these calculations and SAR, we found that ring closure reactions in unsaturated RO₂ radicals are fast, exceeding 1 s⁻¹ for many substitution patterns and ring sizes. These reactions are typically faster than aliphatic RO₂ H-migrations and comparable in rate to RO₂ allylic H-migrations, and should readily compete with bimolecular reactions of RO₂ even at moderately elevated concentrations of NO, HO₂, and RO₂. The cyclic peroxide alkyl radical products are expected to recombine with O₂ in the atmosphere, forming cycloperoxide RO₂ radicals.

The fastest reactions are found for formation of 5 and 6-membered rings. For addition on the inner carbon, larger rings have an entropic disadvantage and the rate coefficient gradually decreases as the cyclic chain becomes longer. For addition on the outer carbon, which is more affected by the geometric impact of the double bond in the TS ring structure, this reduction of the rate for longer chains is not as pronounced, and for rings up to 8 atoms the energetic advantage of the longer chain compensates for the increasing entropic disadvantage. The rate coefficient is most affected by substitution on the double bond, with higher substitution leading to faster ring closure reactions. The stereo-specificity of the substitution on the double bond is most important for ring closure on the outer carbon, where it affects the rate coefficients at 298 K by an order of magnitude. Preliminary calculations and literature data on oxygenated substituents suggest that the latter can greatly enhance the rate of ring closure.

H-Migration in the product cycloperoxide RO₂ radicals was examined for H-atoms implanted directly on the ring structure, including migration of α-OOR H-atoms which leads to ring decomposition. The reaction rates were found to be highly dependent on the ring size and migration span, and are poorly correlated to the H-migration rate in non-cyclic RO₂ especially below 350 K. On average, the rate of migration is slower than the equivalent migration in non-cyclic RO₂. It is proposed that H-migration in substituted cyclic RO₂ could be driven mainly by migration of substituent H-atoms.

H-Migration in the cycloperoxide alkoxy radicals formed from the product cycloperoxide RO₂ was also examined, and



compared to SAR-predictions for decomposition reactions. The migration of H-atoms on the ring is typically slower than for non-cyclic alkoxy radicals, owing to the ring strain in the H-migration TS. For the specific molecular skeleton resulting from the ring closure in the parent unsaturated RO₂ radical, we find that the dominant fate of the cycloperoxide alkoxy radicals is formation of a carbonyl fragment and a non-cyclic carbonyl-alkoxy radical. It is proposed that H-migration in cyclic RO could become important when migrating substituent H-atoms instead of ring H-atoms.

As multi-unsaturated compounds such as isoprene and the monoterpenes make up the bulk of the non-methane organic matter emitted to the atmosphere, unsaturated RO₂ radicals are common in the atmosphere, and ring closure reactions could have a significant impact on the oxidation of some volatile organic compounds. For some compounds it was already shown that RO₂ ring closure reactions have a large impact on the product distribution.^{20–23,28} However, the overall impact of these reactions on the atmospheric oxidation of such compounds can not be quantified in this work, as it is likely to differ greatly between the individual compounds and thus requires detailed analysis of each specific oxidation mechanism. Specifically, the impact of substituents is not covered in the current work, and could have profound impact on the chemical mechanisms of the individual compounds.

Conflicts of interest

There are no conflicts to declare.

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