




Cite this: *Phys. Chem. Chem. Phys.*,  
2021, **23**, 7271

# A test of *ab initio*-generated, radial intermolecular potential energy functions for five axially-symmetric, hydrogen-bonded complexes B...HF, where B = N<sub>2</sub>, CO, PH<sub>3</sub>, HCN and NH<sub>3</sub>

Anthony C. Legon 

The radial potential energy functions for five axially symmetric, hydrogen-bonded complexes B...HF (B = N<sub>2</sub>, CO, PH<sub>3</sub>, HCN and NH<sub>3</sub>) have been calculated *ab initio* at the explicitly correlated level of theory CCSD(T)(F12c)/cc-pVTZ-F12 as a function of the hydrogen-bond distance  $r(\text{Z}\cdots\text{H})$ , where Z is the hydrogen-bond acceptor atom of B. The remaining geometric parameters were optimised at each point in the calculation. The functions so generated were then used to estimate the spectroscopic constants  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$ ,  $\alpha_\sigma$ , and the  $\nu_\sigma = 1 \leftarrow 0$  transition wavenumber associated with the intermolecular stretching mode  $\nu_\sigma$  of each B...HF by using two equivalent approaches. Both involved the assumption that the vibrational modes of the B and HF molecules were sufficiently stiff relative to the intermolecular stretching mode that B...HF could be treated in a pseudo-diatomic approximation. One approach used derivatives of the potential evaluated at the distance  $r = r_e$  while the other used the potential constants obtained by non-linear regression fits of three analytical functions (Morse, Rydberg and Hulbert–Hirschfelder) to the *ab initio* calculated points. The two approaches would give exactly the same results if the functions were a perfect fit. The H–H function gave the best fit. The determined spectroscopic constants were found to be in reasonable agreement with the limited number available by experiment. The efficacy of the approach was tested for the diatomic molecule H<sup>35</sup>Cl by taking advantage of both an accurate RKR-type potential and an accurate set of spectroscopic constants. It was also established that the relationship  $D_e = k_\sigma/(2a^2)$  between two measures ( $D_e$  and  $k_\sigma$ , both calculated *ab initio*) of the strength of the hydrogen bond in the B...HF complexes (and required if the H–H function were an accurate representation of the B...HF potential functions) holds to an excellent level of approximation, and supports the conclusion that this function is appropriate to represent the hydrogen bond in the complexes investigated.

Received 5th January 2021,  
Accepted 22nd February 2021

DOI: 10.1039/d1cp00045d

rsc.li/pccp

## 1. Introduction

The aim of this article is to examine the radial potential energy function, that is the variation of the energy as a function of intermolecular distance, for some simple hydrogen-bonded complexes in the family B...HF, where B is a simple Lewis base. These functions will be calculated *ab initio* at a good level of theory, that is the CCSD(T)-F12c/cc-pVTZ-F12 level, which gives equilibrium dissociation energies  $D_e$  not too far from the complete basis set limit. The PE functions thereby obtained will then be fitted, *via* non-linear regression, to three simple analytical functions (initially proposed for diatomic molecules but used here with the internuclear distances of the diatomic molecules replaced by the intermolecular distances of the weakly bound complexes), namely the Morse function,<sup>1</sup> the Rydberg function<sup>2</sup> and the

Hulbert–Hirschfelder (H–H) function.<sup>3</sup> The aim is to find which gives the most accurate fit and therefore which is best as a model function for the Z...H hydrogen bond. The forms of these functions are given in eqn (1)–(3) for the Morse, Rydberg, and Hulbert–Hirschfelder versions, respectively.

$$V(x) = D_e[1 - e^{-ax}]^2 \quad (1)$$

$$V(x) = D_e - D_e[1 + bx]e^{-bx} \quad (2)$$

$$V(x) = D_e[(1 - e^{-ax})^2 + ca^3x^3e^{-2ax}(1 + abx)] \quad (3)$$

in which  $a$ ,  $b$  and  $c$  are constants,  $x = (r - r_e)$ ,  $r_e$  is the equilibrium internuclear distance Z...HF (Z is the acceptor atom in B) and  $D_e$  is the equilibrium dissociation energy for the process B...HF = B + HF.

The hydrogen-bonded systems investigated were chosen for their simplicity and, importantly, because each is either a linear or a symmetric-top molecule with HF lying along the unique axis at equilibrium. This ensures that for each B...HF there is a

School of Chemistry, University of Bristol, Cantock's Close, Bristol, BS8 1TS, UK.  
E-mail: a.c.legon@bristol.ac.uk



clearly defined path along the extension of the symmetry axis that leads to dissociation into B and HF. The complexes selected were  $N_2 \cdots HF$ ,<sup>4–6</sup>  $OC \cdots HF$ ,<sup>7–9</sup>  $H_3P \cdots HF$ ,<sup>10</sup>  $HCN \cdots HF$ <sup>11–13</sup> and  $H_3N \cdots HF$ .<sup>14</sup> Each has been investigated spectroscopically in isolation in the gas phase, all by rotational spectroscopy and some by high-resolution vibration–rotation spectroscopy, as revealed in the cited papers. The most fully examined is  $HCN \cdots HF$  and it is this complex for which the calculations presented here are best tested by comparing a number of spectroscopically determined constants with those calculated from the analytical potential energy functions fitted to the numerical version obtained *ab initio*. It was noted<sup>15,16</sup> previously that  $D_e$  for such complexes is, to reasonable approximation, directly proportional to the intermolecular stretching quadratic force constant  $k_\sigma$ . This is in fact a necessary consequence of all three functions if the constants  $a$  of Morse and Hulburt–Hirschfelder and  $b$  of Rydberg functions, do not vary much from molecule to molecule.

The fitting procedure and the level of theory used for the  $B \cdots HF$  complexes will be tested initially against the accurate experimental potential energy function of the Rydberg–Klein Rees type for  $H^{35}Cl$  as determined by Coxon and Hajigeorgiou.<sup>17</sup>

## 2. Theoretical methods

The energy of each of the complexes  $N_2 \cdots HF$ ,  $OC \cdots HF$ ,  $H_3P \cdots HF$ ,  $HCN \cdots HF$  and  $H_3N \cdots HF$  was calculated as a function of the distance  $r$  between the nucleus of acceptor atom Z of B and the proton of HF in increments of 0.05 Å, with either  $C_{\infty v}$  or  $C_{3v}$  symmetry, as appropriate, maintained throughout. The calculations were carried out by using the MOLPRO package<sup>18,19</sup> and employed the explicitly correlated CCSD(T)-F12c level of theory<sup>20,21</sup> with cc-pVTZ-F12 basis functions.<sup>22</sup> Tests using the counterpoise keyword showed that basis set superposition error (BSSE) is very small for such calculations. For example, the change in  $D_e$  for  $H_3N \cdots HF$  (the most strongly bound member of the series by far) from this source was only 0.6 kJ mol<sup>-1</sup>. Accordingly, BSSE corrections to the potential function calculations were not made. The calculation of the PE function for HCl was truncated at  $r = 3.2$  Å to avoid spin contamination problems that occur as the Coulson–Fisher point<sup>23</sup> in the internuclear distance is approached and thereby to allow the same level of theory to be used for all calculations reported here.

The intermolecular quadratic force constant  $k_\sigma$  was obtained for each complex  $B \cdots HF$  by fitting the potential energy curve  $V(r - r_e)$  versus  $r - r_e$  within  $\pm 0.1$  Å of the equilibrium intermolecular distance  $r_e$  by a sixth-order polynomial and then evaluating the second derivative with respect to  $(r - r_e)$  at  $r = r_e$ .

## 3. Results

### 3.1 A test of the Morse, Rydberg and Hulburt–Hirschfelder functions with the RKR potential function of $H^{35}Cl$

A very accurate potential energy function for  $H^{35}Cl$  has been evaluated by Coxon and Hajigeorgiou<sup>17</sup> by using the RKR

method to give the energies at the classical turning points of the vibration for the vibrational states  $\nu = 0$  to  $\nu = 20$ , the latter lying within only 18 cm<sup>-1</sup> of the dissociation limit at  $D_e = 445.39$  kJ mol<sup>-1</sup> (37232.0 cm<sup>-1</sup>). This potential energy  $V(r - r_e)$  is plotted against  $(r - r_e)$ , where  $r_e = 1.27454677(6)$  Å,<sup>17</sup> in Fig. 1. The open circles mark the energies at the turning points. Also shown in Fig. 1 are the non-linear regression fits to the RKR points by a Morse function (dashed line) and a Rydberg function (continuous green line). It is clear that the Rydberg function gives a better fit than the Morse function, but that both functions begin to diverge from the RKR-type points beyond  $r = r_e$  and this becomes more significant beyond  $(r - r_e) > 1$  Å. The constants  $D_e$ ,  $a$  and  $b$  that result from these fits are recorded in Table 1 together with their standard errors. The largest residuals are: 9.1 kJ mol<sup>-1</sup> for the Rydberg function fit and 12.0 kJ mol<sup>-1</sup> for the Morse function fit. Also included in Table 1 are the values of  $D_e$  and  $r_e$  for  $H^{35}Cl$  determined by Coxon and Hajigeorgiou.<sup>17</sup> It is clear from Table 1 that both analytical functions overestimate  $D_e$  by about 2–5%. Fig. 2 shows the corresponding diagram for the fit of the  $H^{35}Cl$  RKR-type data by the Hulburt–Hirschfelder function. The continuous green curve is the fitted H–H function.

The determined parameters  $D_e$ ,  $a$ ,  $b$  and  $c$  from the non-linear regression fit are included in Table 1. All fitted points except the final one at  $(r - r_e) = 3.5$  Å have residuals lying in the range  $\pm 5$  kJ mol<sup>-1</sup> for the H–H function fit, confirming its superior quality. This is reinforced by the standard errors of fitted parameters and  $R^2$ , which is the usual goodness-of-fit measure for regression analyses, (see Table 1).

### 3.2 A fit of the Morse, Rydberg and Hulburt–Hirschfelder functions to the potential energy curve of HCl calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory

Having tested the three types of analytical potential energy function with the aid of the high-quality RKR-type  $H^{35}Cl$  experimental function, the version obtained *ab initio* by calculation at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory

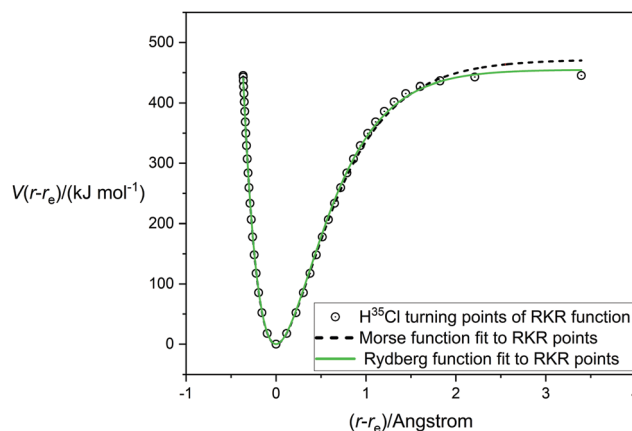


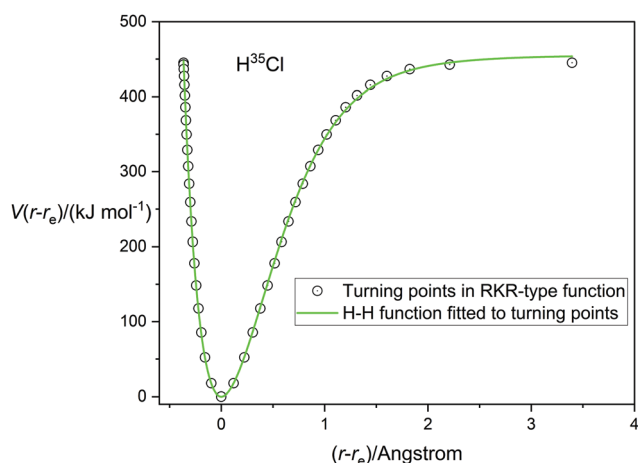
Fig. 1 The RKR-type potential energy function for  $H^{35}Cl$  due to Coxon and Hajigeorgiou<sup>17</sup> (open circles) fitted by a Morse function (dashed line) and a Rydberg function (green line).



**Table 1** Parameters of the Morse, Rydberg and Hulbert–Hirschfelder function fits to the Coxon–Hajigeorgiou RKR-type potential energy function<sup>a</sup> for H<sup>35</sup>Cl

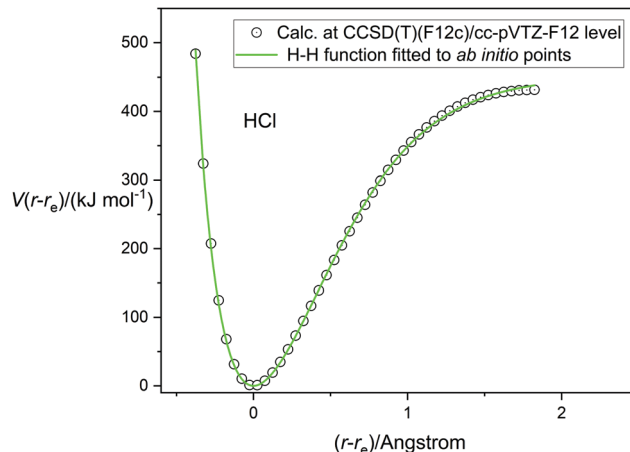
| Parameter                  | Morse function | Rydberg function | Hulbert–Hirschfelder function |
|----------------------------|----------------|------------------|-------------------------------|
| $D_e/(\text{kJ mol}^{-1})$ | 471.9(25)      | 454.8(1.7)       | 455.1(1.7)                    |
| $a/\text{\AA}^{-1}$        | 1.854(5)       | —                | 1.838(8)                      |
| $b/\text{\AA}^{-1}$        | —              | 2.704(5)         | 3.51(68)                      |
| $c$                        | —              | —                | 0.042 (7)                     |
| $R^2$                      | 0.9982         | 0.9991           | 0.9997                        |

<sup>a</sup> The best available value of  $D_e$  for H<sup>35</sup>Cl is 445.39 kJ mol<sup>-1</sup>, with  $r_e = 1.27454677(6)$  \AA from ref. 17.



**Fig. 2** The RKR-type potential energy function for H<sup>35</sup>Cl due to Coxon and Hajigeorgiou<sup>17</sup> (open circles) fitted by a Hulbert–Hirschfelder function (continuous green line).

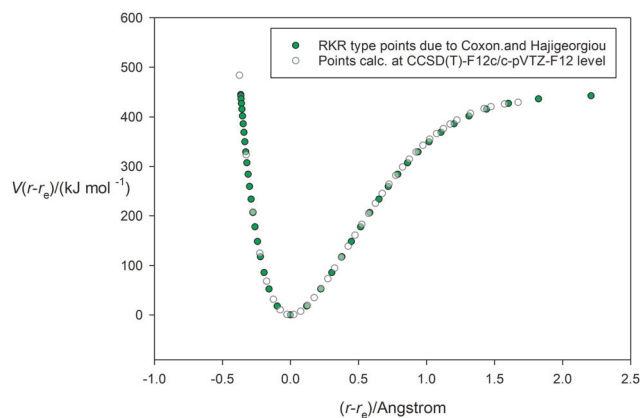
can be similarly scrutinized. Fig. 3 display this function, with points calculated every 0.05 \AA in the range  $(r - r_e) = -0.4$  to 1.8 \AA and indicated by the open circles. The green continuous line again represents the non-linear regression fit of the Hulbert–Hirschfelder function to these points. The parameters of the fit are included in Table 2, together with those derived from the corresponding fits of the Morse function and the Rydberg function to the same calculated points. Recall that the CCSD(T)(F12c) calculation was not extended beyond  $(r - r_e) = 1.8$  \AA to avoid the effects of spin contamination that begin at the Coulson–Fischer point. Fig. 3 indicates that the Hulbert–Hirschfelder function provides a good fit to the *ab initio* calculated points. As with the fits to the RKR-type points for H<sup>35</sup>Cl, the obtained parameters for the Morse and Rydberg functions were less precise, as indicated by the  $R^2$  values. It is clear from Table 2 that although the  $D_e$  values have similar errors for all three analytical functions, those values are higher than the best experimental value from ref. 17. Fig. 4 makes the direct comparison of the RKR-type points for H<sup>35</sup>Cl with those calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level and illustrates visually that the *ab initio* potential function is in excellent agreement with the best available function.



**Fig. 3** The potential energy function of HCl calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory (open circles) fitted by the Hulbert–Hirschfelder analytical function (continuous green line).

**Table 2** Parameters of the Morse, Rydberg and Hulbert–Hirschfelder function fits to the potential function of HCl calculated *ab initio* at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory

| Parameter                  | Morse function | Rydberg function | Hulbert–Hirschfelder function |
|----------------------------|----------------|------------------|-------------------------------|
| $D_e/(\text{kJ mol}^{-1})$ | 484.4(16)      | 461.4(17)        | 455.3(24)                     |
| $a/\text{\AA}^{-1}$        | 1.848(4)       | —                | 1.818(8)                      |
| $b/\text{\AA}^{-1}$        | —              | 2.710(7)         | 4.37(78)                      |
| $c$                        | —              | —                | 0.045(7)                      |
| $R^2$                      | 0.9993         | 0.9991           | 0.9997                        |



**Fig. 4** Comparison of the potential energy curve for HCl calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level (unfilled circles) and that of the RKR-type due to Coxon and Hajigeorgiou<sup>17</sup> (green-filled circles).

### 3.3 Calculation of spectroscopic properties of H<sup>35</sup>Cl from the RKR-type and CCSD(T)(F12c)/cc-pVTZ-F12 potential energy functions using derivatives of $V(r - r_e)$ with respect to $r$ , evaluated at $r = r_e$

It can be shown<sup>24</sup> that, for diatomic molecules, the equilibrium spectroscopic wavenumbers  $\omega_e$ , and  $\omega_e x_e$  and the vibration–rotation



constant  $\alpha_e$  are related to various derivatives of the potential energy function  $V(r - r_e)$  by eqn (4)–(6):

$$\omega_e = (2\pi c)^{-1} [V''(r_e)/\mu]^{\frac{1}{2}} \quad (4)$$

$$\omega_e x_e = \left( \frac{\hbar}{\pi c \mu} \right) \left\{ \frac{5}{96} \left[ \frac{V'''(r_e)}{V''(r_e)} \right]^2 - \frac{1}{32} \left[ \frac{V''''(r_e)}{V''(r_e)} \right] \right\} \quad (5)$$

$$\alpha_e = - \left[ \frac{V'''(r_e)}{3V''(r_e)} r_e + 1 \right] 6 \frac{B_e^2}{\omega_e} \quad (6)$$

In these equations,  $V''$  is the second derivative of the potential function with respect to  $r$  evaluated at  $r_e$ , and so on,  $\mu$  is the reduced mass of the diatomic molecule and  $c$  is the speed of light. The various derivatives were determined by fitting the points, calculated at 0.05 Å intervals, between 0.2 and  $-0.2$  Å with a 7th order polynomial and differentiating the result with respect to  $r$ . The resulting values of  $\omega_e$ ,  $\omega_e x_e$  and  $\alpha_e$  are given in Table 3 for the diatomic molecule  $H^{35}Cl$ , both for the potential energy function of the RKR type due to Coxon and Hajigeorgiou<sup>17</sup> and for that calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level. The purpose of this approach is to provide a calibration in the case where the equations are applied to a diatomic molecule. Experimental values of  $\omega_e$ ,  $\omega_e x_e$  and  $\alpha_e$  for  $H^{35}Cl$  are accurately known<sup>26</sup> and are reproduced in Table 3. It is clear from Table 3 that, as expected, the Coxon–Hajigeorgiou RKR-type function leads to almost perfect reproduction of the experimental values of  $\omega_e$ ,  $\omega_e x_e$  and  $\alpha_e$  for  $H^{35}Cl$ . The *ab initio* calculated function also yields reasonably good agreement with the experimental values of  $\omega_e$  and  $\omega_e x_e$  but significantly underestimates  $\alpha_e$ .

### 3.4 Calculation of spectroscopic properties of the five complexes $B \cdots HF$ available from their CCSD(T)(F12c)/cc-pVTZ-F12 potential energy functions by using derivatives of $V(r - r_e)$ with respect to $r$ , evaluated at $r = r_e$

The potential energy curves  $V(r - r_e)$  versus  $(r - r_e)$  for the five complexes  $B \cdots HF$  ( $B = N_2, CO, PH_3, HCN$  and  $NH_3$ ) calculated *ab initio* at the CCSD(T)(F12c)/cc-pVTZ-F12 level are collected

together in Fig. 5. The points, which were calculated at 0.05 Å intervals in  $r$  with geometry optimization at each, are not shown explicitly but are connected by a spline function to give continuous curves for clarity. The dissociation energies cover the range from  $\sim 50$  kJ mol<sup>-1</sup> ( $H_3N \cdots HF$ ) to  $\sim 10$  kJ mol<sup>-1</sup> (for  $N_2 \cdots HF$ ). The spectroscopic properties  $\omega_e$ ,  $\omega_e x_e$  and  $\alpha_e$  for the five complexes can be estimated by use of eqn (4)–(6), respectively. The required derivatives  $V''$  etc. were obtained by the method used for  $H^{35}Cl$ , but a fit to a 6th order polynomial was sufficient. In applying eqn (4)–(6) to the  $B \cdots HF$  complexes, there is an implicit assumption that the component molecules  $B$  and  $HF$  are sufficiently rigid, relative to the intermolecular hydrogen bond of the complexes examined, for these equations can be applied if  $\mu = (m_B m_{HF}) / (m_B + m_{HF})$  is used in place of the diatomic molecule reduced mass. In that case, the determined properties are written as  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$  and  $\alpha_\sigma$  where the subscripts now indicate that they refer to the intermolecular stretching mode  $\nu_\sigma$  of the complexes. This assumption is reasonable given the simplicity of the complexes and the large vibrational wavenumbers of  $B$  and  $HF$  compared with those calculated for the hydrogen-bond stretching mode  $\nu_\sigma$ . The results obtained under this assumption are included in Table 3. Their comparison with such experimental quantities as are available for the  $B \cdots HF$  will be postponed until Section 3.4.

### 3.5 The fit of the Morse, Rydberg and Hulbert–Hirschfelder functions to the potential energy curves of the $B \cdots HF$ calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory

The three analytical potential energy functions tested for  $HCl$  were next fitted to the *ab initio* calculated points for each of the five axially symmetric hydrogen-bonded complexes  $B \cdots HF$ . The fits of each of the three functions (Morse, Rydberg and Hulbert–Hirschfelder) to the *ab initio* calculated points for  $HCN \cdots HF$  are displayed in Fig. 6(a)–(c), respectively. The quality of fit to the Rydberg function is worse than that of the Morse function. As found for  $HCl$ , the Hulbert–Hirschfelder function provides the best fit to the CCSD(T)(F12c)/cc-pVTZ-F12 calculated points. This is clear from Table 4, which records the

**Table 3** Spectroscopic properties associated with vibration in  $H^{35}Cl$  and with the intermolecular stretching mode of the axially symmetric hydrogen-bonded complexes  $B \cdots HF$ , where  $B = N_2, CO, PH_3, HCN$  and  $NH_3$ . These properties were calculated from the derivatives of the  $V(r - r_e)$  function, the latter calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory, in each case. For  $H^{35}Cl$ , those obtained from the RKR-type function and experiment are also included

| Complex                         | Spectroscopic property          |                       |   |   |                              |   |
|---------------------------------|---------------------------------|-----------------------|---|---|------------------------------|---|
|                                 | $B_e$ /MHz (calc.) <sup>a</sup> | $r_e$ <sup>a</sup> /Å | $\omega_e$ <sup>b</sup> /cm <sup>-1</sup> | $\omega_e x_e$ <sup>b</sup> /cm <sup>-1</sup> | $\alpha_e$ <sup>b</sup> /MHz | $\nu = 1 \leftarrow 0$<br>wavenumber/cm <sup>-1</sup> |
| $H^{35}Cl$ (experiment)         | 317582.622                      | 1.274552 <sup>c</sup> | 2990.946 <sup>c</sup>                     | 52.8186 <sup>c</sup>                          | 9209.1 <sup>c</sup>          | 2885.31   |
| $H^{35}Cl$ (RKR)                | 317582.622                      | 1.274552              | 2991.1                                    | 53.2  | 9190.0                       | 2884.0  |
| $H^{35}Cl$ ( <i>ab initio</i> ) | 316880.128                      | 1.275964              | 2999.6                                    | 51.6  | 9127.5                       | 2896.4  |
| $N_2 \cdots HF$                 | 3302.814                        | 2.087212              | 115.4                                     | 4.66  | 51.4                         | 106.1   |
| $OC \cdots HF$                  | 3125.996                        | 2.090646              | 132.7                                     | 3.98  | 35.4                         | 124.7   |
| $H_3P \cdots HF$                | 3550.382                        | 2.362105              | 134.2                                     | 2.97  | 45.9                         | 128.3   |
| $HCN \cdots HF$                 | 3619.792                        | 1.853759              | 181.5                                     | 3.99  | 28.9                         | 173.5   |
| $H_3N \cdots HF$                | 7437.592                        | 1.699216              | 254.0                                     | 4.19  | 59.3                         | 245.6   |

<sup>a</sup> Calculated from the geometry optimised at the CCSD(T)(F12c)/cc-pVTZ-F12 level, except for  $H^{35}Cl$  (RKR), which uses the experimental values of  $r_e$  and  $B_e$  from ref. 25. <sup>b</sup> The subscript e (rather than  $\sigma$ ) is appropriate for  $H^{35}Cl$  quantities. <sup>c</sup> See ref. 25 for a collection of experimentally observed constants for  $H^{35}Cl$ .



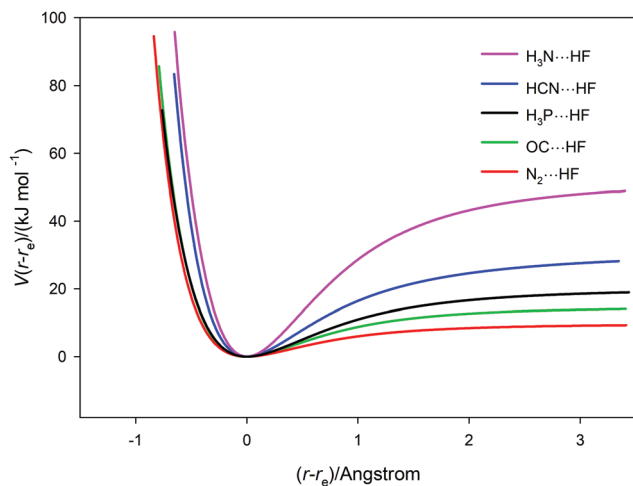


Fig. 5 Potential energy curves  $V(r - r_e)$  versus  $(r - r_e)$  for the five hydrogen-bonded complexes  $B \cdots HF$  ( $B = N_2, CO, PH_3, HCN$  and  $NH_3$ ) calculated at 0.05 Å intervals, with geometry optimization at each point. The points (not shown explicitly) are joined by a spline function.

parameters obtained from a non-linear regression fit for each  $B \cdots HF$  complex to each of the 3 analytical functions.

In applying the Morse, Rydberg and Hulburt–Hirschfelder functions, all of which were proposed originally for diatomic molecules, to the five complexes  $B \cdots HF$ , the assumption is again that the B and HF molecules are sufficiently rigid relative to the intermolecular hydrogen bond to allow the complexes to be treated in reasonable approximation as diatomic molecules. The fact that in each case the intermolecular non-covalent bond lies along the symmetry axis is a useful simplification. Fig. 7 compares the potential energy curve for  $HCN \cdots HF$ , calculated as a function of the  $(r - r_e)$  (applying geometry optimization at each point), with that calculated when HCN and HF are frozen at the geometries they possess in the optimized equilibrium version of the complex. Both curves were calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level. It is clear that, apart for small differences on the repulsive part of the curves at small distance  $r(N \cdots F)$ , the two sets of points are almost coincident for  $(r - r_e) > -0.7$  Å and *de facto* lie very close to the H–H fit to the relaxed function shown in Fig. 6(c).

To decide which, if any, of the three analytical potential energy functions is suitable as a model for the hydrogen bond in these axially symmetric complexes  $B \cdots HF$ , one criterion must be the goodness of their fits to the *ab initio*-calculated functions. However, another important criterion is how well the fitted analytical functions predict spectroscopic properties associated with the intermolecular stretching mode  $\sigma$ , namely the equilibrium stretching wavenumbers  $\omega_\sigma$  and  $\omega_\sigma x_\sigma$ , the vibration–rotation constant  $\alpha_\sigma$  and the dissociation energy  $D_e$ , all of which can in principle be measured spectroscopically.  $HCN \cdots HF$  is the most investigated of the five complexes  $B \cdots HF$  selected for consideration here, but some of the spectroscopic constants are also available for  $OC \cdots HF$  and  $N_2 \cdots HF$ . Eqn (4)–(6), as set out in Section 3.2, allow  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$  and  $\alpha_\sigma$ , respectively, to be obtained from the fitted parameters of each of the three analytical potential functions for each of the five

complexes  $B \cdots HF$ . As indicated earlier, these equations involve derivatives of the potential function with respect to  $r$  when evaluated at  $r = r_e$ . It has been established that the Hulburt–Hirschfelder function gives a better fit to the potential energy curves of the five  $B \cdots HF$  complexes obtained through *ab initio* calculations at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory. Therefore, only the relationships between  $a$ ,  $b$  and  $c$  of this function and  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$  and  $\alpha_\sigma$ , will be considered here.

The results of evaluating eqn (4) and (5) for the Hulburt–Hirschfelder function are as follows:

$$k_\sigma = 2a^2 D_e, \quad (7)$$

$$\omega_\sigma = \frac{1}{2\pi c} \left( \frac{k_\sigma}{\mu} \right)^{\frac{1}{2}} = \frac{1}{2\pi c} \left( \frac{2a^2 D_e}{\mu} \right)^{\frac{1}{2}} \quad (8)$$

$$\omega_\sigma x_\sigma = \frac{\hbar a^2}{4\pi c \mu} \left[ \frac{15(c-1)^2}{8} - \frac{7}{8} + \frac{3c}{1} - \frac{3bc}{2} \right] \quad (9)$$

Eqn (9) transforms to the corresponding expression obtained from the Morse function (see ref. 23) when  $b = c = 0$ , as required.

It can be shown<sup>3,26</sup> by use of eqn (6) that the vibration–rotation constant  $\alpha_e$  for a diatomic molecule is related to the constants  $a$  and  $c$  of the Hulburt–Hirschfelder potential energy function by

$$\alpha_e = 6 \frac{B_e^2}{\omega_e} [ar_e(1-c) - 1], \quad (10)$$

where  $B_e$  is the equilibrium rotational constant. If the components B and HF of the complex can be assumed rigid and eqn (10) is applied in that approximation, a route to  $\alpha_\sigma$  is available because  $B_e$  is known from the *ab initio* calculations described here. Table 5 gives the values of  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$ ,  $\alpha_\sigma$ , and the  $\nu_\sigma = 1 \leftarrow 0$  transition wavenumber for the hydrogen-bond stretching mode  $\nu_\sigma$ . They were calculated from eqn (8)–(10) using the parameters  $a$ ,  $b$ ,  $c$  of the Hulburt–Hirschfelder function (see Table 4) obtained by fitting the *ab initio*-generated potential energy curves for each of the five  $B \cdots HF$ . Also in Table 5 are the corresponding quantities for  $H^{35}Cl$  calculated (1) from fit of the H–H function to the Coxon–Hajigeorgiou RKR-type function,<sup>17</sup> and (2) from the fit of the H–H function to the CCSD(T)(F12c)/cc-pVTZ-F12 points calculated here. The experimental values of the  $H^{35}Cl$  properties are also included in Table 5 for convenience.<sup>25</sup> The agreement of these quantities with those obtained by fitting the RKR-type potential with the H–H function (see Table 3) is reasonable, although  $D_e$  is overestimated by about 2% while  $\omega_e$  is underestimated by 0.5%. As might be expected the terms  $\omega_e x_e$  and  $\alpha_e$  for  $H^{35}Cl$ , which account for differences between equilibrium and zero-point vibrational and rotational constants, respectively, are more significantly in error, by 5% and 9%. The set of quantities for  $H^{35}Cl$  calculated from the potential energy function generated *ab initio* and fitted by the H–H function, also given in Table 4, are somewhat more in error, as might be expected.

The spectroscopic constants of the  $B \cdots HF$  have been estimated by two equivalent approaches, both of which depend



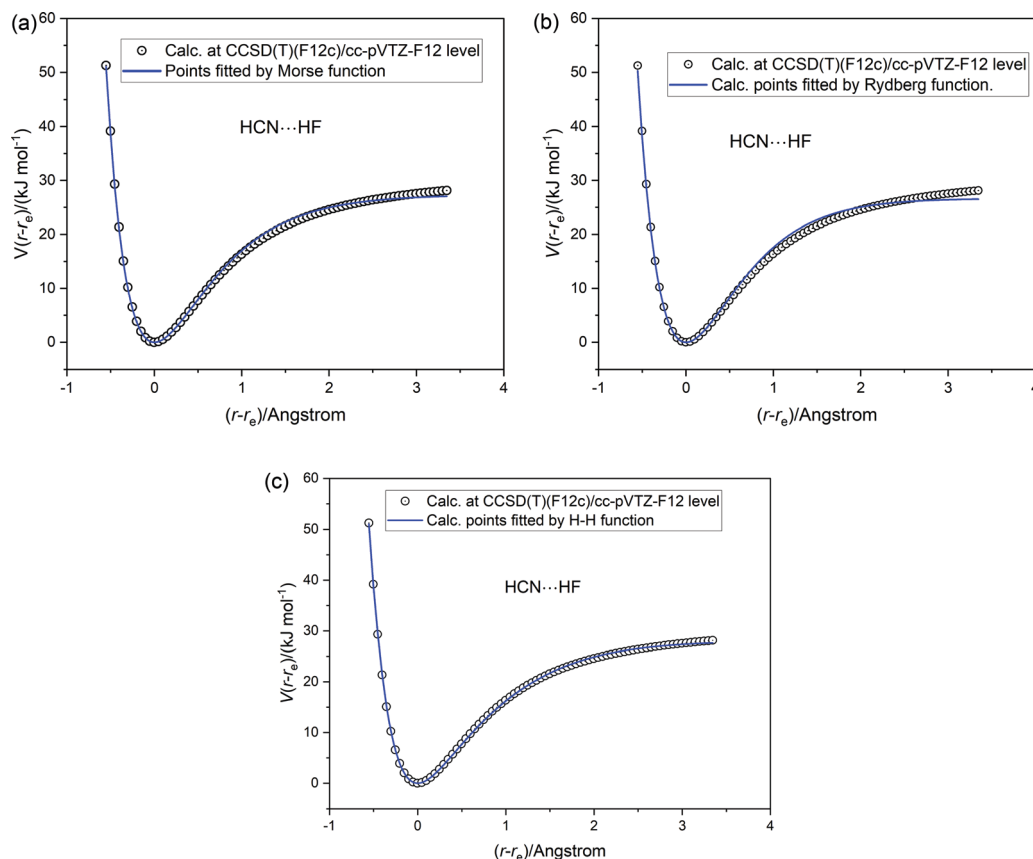


Fig. 6 Potential energy  $V(r - r_e)$  versus  $(r - r_e)$  for HCN...HF calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level (open circles). The continuous blue line is a non-linear regression fit to the *ab initio* calculated points by (a) a Morse function (b) a Rydberg function and (c) a Hulbert–Hirschfelder function. The parameters of the fits are included in Table 4.

Table 4 Constants of the Morse, Rydberg and Hulbert–Hirschfelder analytical potential energy functions obtained by fitting the  $V(r - r_e)$  versus  $(r - r_e)$  points generated at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory for axially-symmetric complexes B...HF

| Complex               | Morse function             |                     | Rydberg function           |                     | Hulbert–Hirschfelder function |                     |           |           |
|-----------------------|----------------------------|---------------------|----------------------------|---------------------|-------------------------------|---------------------|-----------|-----------|
|                       | $D_e/(\text{kJ mol}^{-1})$ | $a/\text{\AA}^{-1}$ | $D_e/(\text{kJ mol}^{-1})$ | $b/\text{\AA}^{-1}$ | $D_e/(\text{kJ mol}^{-1})$    | $a/\text{\AA}^{-1}$ | $b$       | $c$       |
| N <sub>2</sub> ...HF  | 9.07(2)                    | 1.740(2)            | 8.84(4)                    | 2.557(3)            | 9.20(1)                       | 1.720(5)            | 0.90(7)   | −0.067(4) |
| OC...HF               | 13.96(5)                   | 1.583(2)            | 13.56(5)                   | 2.334(3)            | 14.02(2)                      | 1.603(3)            | 1.29(10)  | −0.035(3) |
| H <sub>3</sub> P...HF | 18.87(3)                   | 1.441(2)            | 18.35(6)                   | 2.101(3)            | 19.02(2)                      | 1.447(2)            | 1.69(12)  | −0.026(3) |
| HCN...HF              | 27.35(9)                   | 1.556(4)            | 26.64(14)                  | 2.263(11)           | 28.04(5)                      | 1.545(5)            | 1.20(9)   | −0.084(4) |
| HCN...HF <sup>a</sup> | —                          | —                   | —                          | —                   | 28.21(5)                      | 1.581(7)            | 0.76(9)   | −0.120(5) |
| H <sub>3</sub> N...HF | 49.92(17)                  | 1.354(3)            | 48.41(12)                  | 1.977(4)            | 49.21(7)                      | 1.440(3)            | −1.14(16) | 0.042(3)  |

<sup>a</sup> The values in this row refer to the Hulbert–Hirschfelder function fitted to the *ab initio* potential energy points generated with the HCN and HF geometries frozen at their values in the equilibrium complex (see text for discussion).

ultimately on eqn (4)–(6). Those in Table 3 result from the application of eqn (4)–(6) directly, while those in Table 5 were obtained by first fitting the H–H function to the *ab initio* calculated points and then applying eqn (8)–(10). If the H–H function were a perfect fit to the *ab initio* points, the two sets of values for the B...HF in Tables 3 and 5 would be identical. In fact the two sets are in reasonable agreement, an observation which confirms that the H–H function provides a good fit to the *ab initio*-generated potential functions.

Some of the properties of the B...HF given in Tables 3 and 5 can be compared with those available from experiment.

Those members of the B...HF series for which gas-phase experimental investigations have been published are N<sub>2</sub>...HF,<sup>6</sup> OC...HF<sup>8,9</sup> and HCN...HF.<sup>11,12</sup> Various of the calculated constants in Tables 3 and 5 have been accurately measured by rotational and high-resolution vibrational spectroscopy and their values are included in Table 5. The agreement between the calculated values and those obtained by experiment is satisfactory, except for  $\alpha_\sigma$  for HCN...HF, which, as was also noted for HCl in the preceding paragraph, is significantly underestimated. It thus appears that *ab initio* calculations at the CCSD(T)(F12c)/cc-pVTZ-F12 explicitly correlated level lead to a reasonably good



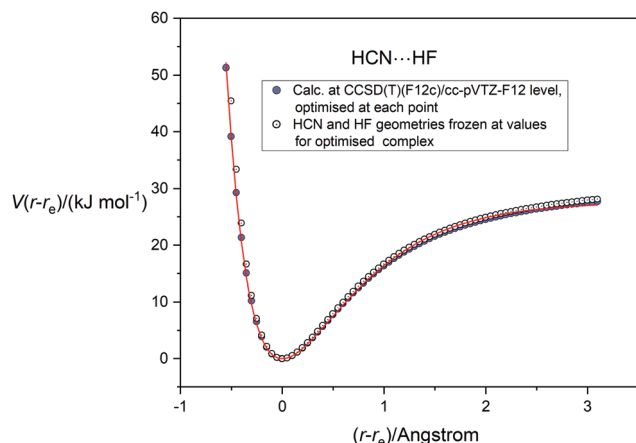


Fig. 7 Comparison of potential energy curve for HCN...HF calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level with geometry relaxation at each point (blue-filled circles) with that calculated at the same level of theory but with the HCN and HF geometries frozen at those appropriate to the equilibrium complex (open circles). The continuous red line is the Hulbert–Hirschfelder function fitted to the relaxed PE points.

representation of the radial potential energy function for the hydrogen-bond stretching mode in those axially symmetric complexes B...HF considered here. Certainly, the H–H function provides a good fit to the *ab initio* curves. If the same percentage errors in the various calculated properties found for HCl were to apply to the B...HF complexes, the agreement would appear better just because the quantities are generally smaller for the complexes.

An alternative approach to fitting the *ab initio*-calculated potential functions of the B...HF complexes by means of the Hulbert–Hirschfelder analytical function is worth investigating. This approach involves holding  $D_e$  fixed at its value calculated from the energies  $E_B$ ,  $E_{HF}$  and  $E_{B...HF}$  of the B, HF and B...HF geometries optimized at the CCSD(T)(F12c)/cc-pVTZ-F12 level.  $D_e$  is related to these energies by the equation

$$D_e = E_{B...HF} - E_B - E_{HF}, \quad (11)$$

Table 6 Constants of the Hulbert–Hirschfelder analytical potential energy function obtained by fitting the  $V(r - r_e)$  versus  $(r - r_e)$  points generated by calculations at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory. In these fits,  $D_e$  was held fixed at the value given by the *ab initio* calculations (see footnote) and  $a$ ,  $b$  and  $c$  released

| Complex               | Hulbert–Hirschfelder function    |                     |           |            |
|-----------------------|----------------------------------|---------------------|-----------|------------|
|                       | $D_e^a$ /(kJ mol <sup>-1</sup> ) | $a/\text{\AA}^{-1}$ | $b$       | $c$        |
| N <sub>2</sub> ...HF  | 9.46                             | 1.683(11)           | 0.81(11)  | -0.115(7)  |
| OC...HF               | 14.49                            | 1.558(7)            | 0.873(74) | -0.092(7)  |
| H <sub>3</sub> P...HF | 19.84                            | 1.404(8)            | 1.01(11)  | -0.094(7)  |
| HCN...HF              | 30.70                            | 1.433(22)           | 0.81(15)  | -0.232(16) |
| H <sub>3</sub> N...HF | 51.61                            | 1.392(10)           | 3.76(81)  | -0.030(8)  |

<sup>a</sup> Calculated from the CCSD(T)(F12c)/cc-pVTZ-F12 energies of B...HF, B and HF, after counterpoise correction, using the equation  $D_e = E_{B...HF} - E_B - E_{HF}$ .

Counterpoise correction was applied to correct for BSSE. With  $D_e$  so fixed, the remaining three parameters of the H–H function were fitted to the *ab initio* potential curve for each B...HF to give the values of the constants  $a$ ,  $b$  and  $c$  recorded in Table 6. These values have larger standard errors than those obtained by fitting all four parameters  $D_e$ ,  $a$ ,  $b$ , and  $c$  (see Table 4), as expected. However, the numerical spectroscopic constants calculated from the them (Table 7) are in good agreement with those shown in Table 5, which were calculated from the H–H constants given in Table 4.

### 3.6 The relationship between $D_e$ and $k_\sigma$ for the B...HF complexes

Eqn (7) shows the Hulbert–Hirschfelder function requires that  $D_e = k_\sigma/2a^2$ . Accurate  $k_\sigma$  values are available, given (1) the definition  $k_\sigma = V''(r_e)$  (*i.e.* the second derivative of the potential evaluated at the minimum  $r = r_e$ ) and (2) the fact that the region close to the minimum (*i.e.*  $\pm 0.1$  Å about  $r_e$  in steps of 0.025 Å) of

Table 5 Spectroscopic constants associated with the intermolecular stretching mode of the axially symmetric hydrogen-bonded complexes B...HF, where B = N<sub>2</sub>, CO, PH<sub>3</sub>, HCN and NH<sub>3</sub>. The constants were calculated from the parameters of the Hulbert–Hirschfelder potential energy function fitted to the  $V(r - r_e)$  versus  $(r - r_e)$  curve, the latter generated at the CCSD(T)(F12c)/cc-pVTZ-F12 level of theory, in each case

| Complex                                 | Spectroscopic property  |                                |                                     |   |                        |   |
|---|-------------------------|--------------------------------|-------------------------------------|---|------------------------|---|
|   | $B_e$ /MHz              | $D_e$ /(kJ mol <sup>-1</sup> ) | $\omega_\sigma^a$ /cm <sup>-1</sup> | $\omega_{\sigma\sigma}^a$ /cm <sup>-1</sup> | $\alpha_\sigma^a$ /MHz | $\nu_\sigma = 1 \leftarrow 0$ wavenumber <sup>a</sup> /cm <sup>-1</sup> |
| H <sup>35</sup> Cl (fit to RKR)         | 317582.622 <sup>b</sup> | 455.3                          | 2974.5                              | 43.7  | 8410                   | 2887.0  |
| H <sup>35</sup> Cl ( <i>ab initio</i> ) | 316880.128              | 455.1                          | 2941.4                              | 38.3  | 8295                   | 2864.9  |
| N <sub>2</sub> ...HF                    | 3302.814                | 9.20(1)                        | 114.7                               | 4.91  | 53.9                   | 104.9   |
| OC...HF                                 | 3125.996                | 14.02(2)                       | 131.9                               | 4.07  | 36.6                   | 123.8   |
| H <sub>3</sub> P...HF                   | 3550.382                | 19.02(2)                       | 133.6                               | 3.04  | 47.3                   | 127.5   |
| HCN...HF                                | 3619.792                | 28.04(5)                       | 180.5                               | 4.29  | 30.5                   | 171.9   |
| H <sub>3</sub> N...HF                   | 7437.592                | 49.21(7)                       | 250.2                               | 3.97  | 59.5                   | 242.3   |
| Experimental                            |                         |                                |                                     |   |                        |   |
| H <sup>35</sup> Cl                      | 317582.622 <sup>a</sup> | 445.39 <sup>c</sup>            | 2990.946 <sup>b</sup>               | 52.8186 <sup>b</sup>                        | 9209.1 <sup>b</sup>    | 2885.31   |
| N <sub>2</sub> ...HF                    | —                       | 8.97 <sup>d</sup>              | —                                   | —   | —                      | 79.5 <sup>e</sup>   |
| OC...HF                                 | —                       | 15.67 <sup>f</sup>             | 108.7–114.5 <sup>g</sup>            | 3.79  | —                      | 107.99(2) <sup>f</sup>  |
| HCN...HF                                | —                       | 26.1(16) <sup>g</sup>          | —                                   | —   | 61.79(3) <sup>h</sup>  | 197(15) <sup>g</sup> , 168.3(2) <sup>i</sup>                            |

<sup>a</sup> The subscript e (rather than  $\sigma$ ) is appropriate for H<sup>35</sup>Cl quantities. <sup>b</sup> Ref. 25. <sup>c</sup> Ref. 17. <sup>d</sup> Ref. 6.  $D_e$  is semi-experimental estimate. <sup>e</sup> Ref. 6. <sup>f</sup> Predicted using a high-quality morphed potential function. Ref. 8. <sup>g</sup> Ref. 9. <sup>h</sup> From absolute intensity measurements in the rotational spectrum of HCN...HF. See ref. 11. <sup>i</sup> Ref. 12.



**Table 7** Spectroscopic constants associated with the intermolecular stretching mode of the axially symmetric hydrogen-bonded complexes B...HF. The constants were calculated from the parameters  $a$ ,  $b$  and  $c$  of Hulburt–Hirschfelder potential energy function fitted to the  $V(r - r_e)$  versus  $(r - r_e)$  curve with  $D_e$  fixed at its value calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level (see Table 6)

| Complex               | Spectroscopic property |                                   |                                    |   |                       |   |
|-----------------------|------------------------|-----------------------------------|------------------------------------|---|-----------------------|---|
|                       | $B_e^a$ /MHz           | $D_e^b$ / (kJ mol <sup>-1</sup> ) | $\omega_\sigma$ / cm <sup>-1</sup> | $\omega_\sigma x_\sigma$ / cm <sup>-1</sup> | $\alpha_\sigma$ / MHz | $\nu_\sigma = 1 \leftarrow 0$ wavenumber/cm <sup>-1</sup> |
| N <sub>2</sub> ...HF  | 3302.814               | 9.46                              | 113.8                              | 5.12  | 56.0                  | 103.6   |
| OC...HF               | 3125.996               | 14.49                             | 130.4                              | 4.23  | 38.3                  | 121.9   |
| H <sub>3</sub> P...HF | 3550.382               | 19.84                             | 132.3                              | 3.24  | 77.3                  | 125.8   |
| HCN...HF              | 3619.792               | 30.70                             | 175.8                              | 4.70  | 33.9                  | 166.4   |
| H <sub>3</sub> N...HF | 7437.592               | 51.61                             | 247.5                              | 4.28  | 64.2                  | 238.9   |

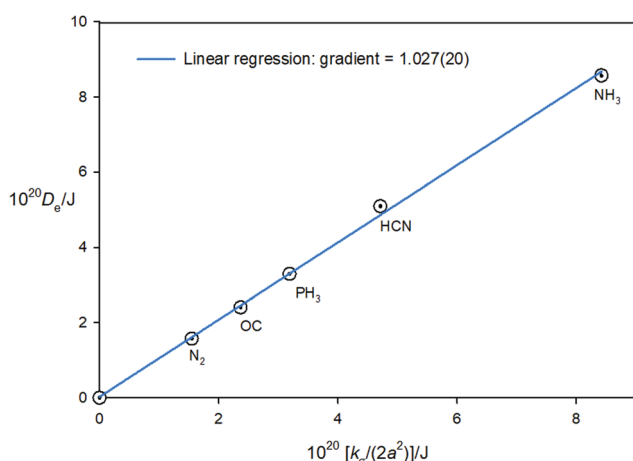
<sup>a</sup> Value of equilibrium rotational constant of most abundant isotopologue calculated at CCSD(T)(F12c)/cc-pVTZ-F12 level. <sup>b</sup> Calculated from the CCSD(T)(F12c)/cc-pVTZ-F12 energies of B...HF, B and HF, after counterpoise correction, using the equation  $D_e = E_{B...ClF} - E_B - E_{ClF}$ .

**Table 8** Two measures ( $D_e$  and  $k_\sigma$ ) of binding strength for axially-symmetric, hydrogen-bonded complexes B...HF calculated at the CCSD(T)(F12c)/cc-pVTZ-F12 level

| Complex               | $D_e^a$ /kJ mol <sup>-1</sup> | $k_\sigma^b$ /(N m <sup>-1</sup> ) |
|-----------------------|-------------------------------|------------------------------------|
| N <sub>2</sub> ...HF  | 9.46                          | 9.15                               |
| OC...HF               | 14.49                         | 12.11                              |
| H <sub>3</sub> P...HF | 19.84                         | 13.36                              |
| HCN...HF              | 30.70                         | 22.31                              |
| H <sub>3</sub> N...HF | 51.61                         | 34.96                              |

<sup>a</sup> Calculated from the CCSD(T)(F12c)/cc-pVTZ-F12 energies of B...HF, B and HF, after counterpoise correction, using the equation  $D_e = E_{B...ClF} - E_B - E_{ClF}$ . <sup>b</sup> Calculated from  $V''(r_e) = k_\sigma$  by fitting a polynomial around the minimum of the *ab initio*-calculated curves and differentiating.

the CCSD(T)(F12c)/cc-pVTZ-F12 potential function of each B...HF complex can be fitted precisely to a 6th order polynomial (see Section 3.2) and  $k_\sigma$  can then be obtained by differentiation. Moreover,  $D_e$  values at this level of theory have



**Fig. 8** A plot of the equilibrium dissociation energy  $D_e$  versus  $k_\sigma/(2a^2)$  for five axially-symmetric, hydrogen-bonded complexes B...HF, where B = N<sub>2</sub>, CO, PH<sub>3</sub>, HCN and NH<sub>3</sub>.

been evaluated as the difference between the energies of B...HF and B + HF, [see eqn (11)] after counterpoise correction (which is very small). The two sets of quantities are recorded in Table 8 for the five B...HF complexes investigated. According to eqn (11), if the Hulburt–Hirschfelder function were an exact fit for each of the five B...HF complexes, a plot of  $D_e$  versus  $k_\sigma/(2a^2)$  would be a straight line of gradient 1.000 and intercept 0.000. Fig. 8 shows that such a plot is an excellent straight line, for which the linear regression fit leads to the gradient =  $6.17(13) \times 10^{23} \text{ mol}^{-1}$  or 1.02(2) per molecule, with the intercept 0.005(15) kJ mol<sup>-1</sup> and  $R^2 = 0.9986$ . This result establishes with some accuracy that  $D_e = k_\sigma/2a^2$  for the family of the five axially symmetric, hydrogen-bonded complexes B...HF, (B = N<sub>2</sub>, CO, PH<sub>3</sub>, HCN, NH<sub>3</sub>).

## 4. Conclusions

In this article, several conclusions have been established about the suitability of explicitly correlated *ab initio* calculations at the CCSD(T)(F12c)/cc-pVTZ-F12 level for generating the potential energy changes as a function of the intermolecular distance  $r$  in the series of axially symmetric, hydrogen-bonded complexes B...HF, when B = N<sub>2</sub>, CO, PH<sub>3</sub>, HCN and NH<sub>3</sub>. The aim was to predict spectroscopic properties of the complexes under the assumption that the component molecules B and HF were sufficiently rigid relative to the intermolecular radial mode  $\nu_\sigma$  to allow them to be treated as diatomic molecules in reasonable approximation. Two approaches were followed: (1) eqn (4)–(6) relating the spectroscopic properties  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$  and  $\alpha_\sigma$  to various derivatives of the potential energy with respect to  $r$ , evaluated at  $r = r_e$  were applied to the *ab initio* generated functions, and (2) the potential function calculated *ab initio* for each B...HF was fitted by Morse, Rydberg and Hulburt–Hirschfelder functions in turn. Analysis of the fits revealed that the four-constant, H–H function provided statistically the best fit in each case. The same sets of spectroscopic properties as considered in (1) were then evaluated from the parameters of the fitted H–H functions. The constants  $\omega_\sigma$ ,  $\omega_\sigma x_\sigma$  and  $\alpha_\sigma$  determined by approaches (1) and (2) were in good agreement, and would be identical if the H–H function were a perfect fit. When compared with the small number of available experimental versions of these constants, the predicted values were in satisfactory agreement with experiment, indicating that potential energy functions calculated in the diatomic approximation for the chosen complexes might provide a guide to the (often demanding) experiments required to measure such quantities.

As a prelude to the discussion of the B...HF, the same procedures were applied to the diatomic molecule H<sup>35</sup>Cl, for which an accurate RKR-type potential is available. The conclusion for H<sup>35</sup>Cl was that the CCSD(T)(F12c)/cc-pVTZ-F12 calculations provided an encouragingly good representation of the RKR function and yielded spectroscopic properties in satisfactory (but, of course, not perfect) agreement with the very accurately known experimental values.



It was also established that the relationship  $D_e = k_{\sigma}/(2a^2)$  between the two measures ( $D_e$  and  $k_{\sigma}$ , both calculated *ab initio*) of the strength of the hydrogen bond in the  $B \cdots HF$  complexes (and required if the H–H function were an accurate representation of the  $B \cdots HF$  potential functions) holds to an excellent level of approximation, and supports the conclusion that the H–H function is appropriate to represent the hydrogen bond in the complexes investigated.

## Conflicts of interest

There are no conflicts of interest to report.

## Acknowledgements

I am pleased to acknowledge the award of a University Senior Research Fellowship by the University of Bristol, during the tenure of which the work described here was carried out.

## References

- 1 P. M. Morse, *Phys. Rev.*, 1929, **34**, 57–65.
- 2 R. Rydberg, *Z. Phys.*, 1932, **73**, 376–385.
- 3 H. M. Hulburt and J. O. Hirschfelder, *J. Chem. Phys.*, 1941, **9**, 61–69, and *J. Chem. Phys.*, 1961, **35**, 1901.
- 4 P. D. Soper, A. C. Legon, W. G. Read and W. H. Flygare, *J. Chem. Phys.*, 1982, **76**, 292–300.
- 5 C. M. Lovejoy and D. J. Nesbitt, *J. Chem. Phys.*, 1987, **86**, 3151–3165.
- 6 P. Jankowski, S. N. Tsang, W. Klemperer and K. Szalewicz, *J. Chem. Phys.*, 2001, **114**, 8948–8963.
- 7 A. C. Legon, P. D. Soper and W. H. Flygare, *J. Chem. Phys.*, 1981, **74**, 4944–4950.
- 8 L. A. Rivera-Rivera, Z. Wang, B. A. McElmurry, R. R. Lucchese, J. W. Bevan and G. Kanschä, *Chem. Phys.*, 2011, **390**, 42–50.
- 9 G. T. Fraser and A. S. Pine, *J. Chem. Phys.*, 1988, **88**, 4147–4152.
- 10 A. C. Legon and L. C. Willoughby, *Chem. Phys.*, 1983, **74**, 127–136.
- 11 A. C. Legon, D. J. Millen and S. C. Rogers, *Proc. R. Soc. London, Ser. A*, 1980, **370**, 213–237.
- 12 D. Bender, M. Eliades, D. A. Danzeiser, M. W. Jackson and J. W. Bevan, *J. Chem. Phys.*, 1987, **86**, 1225–1234.
- 13 B. A. Wofford, S. G. Lied and J. W. Bevan, *J. Chem. Phys.*, 1987, **87**, 4478–4486.
- 14 S. W. Hunt, K. J. Higgins, M. B. Craddock, C. S. Brauer and K. R. Leopold, *J. Am. Chem. Soc.*, 2003, **125**, 13850–13860.
- 15 A. C. Legon, *Phys. Chem. Chem. Phys.*, 2014, **16**, 12415–12421 and 25199.
- 16 I. Alkorta and A. C. Legon, *New J. Chem.*, 2018, **42**, 10548–10554.
- 17 J. A. Coxon and P. G. Hajigeorgiou, *J. Quant. Spectrosc. Radiat. Transfer*, 2015, **151**, 133–154.
- 18 H.-J. Werner, P. J. Knowles, G. Knizia, F. R. Manby and M. Schütz, ‘Molpro: a general-purpose quantum chemistry program package’, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2012, **2**, 242–253.
- 19 MOLPRO 2019.2, a package of ab initio programs, H.-J. Werner, P. J. Knowles, G. Knizia, F. R. Manby and M. Schütz, M, 2012.
- 20 C. Hättig, D. P. Tew and A. Köhn, *J. Chem. Phys.*, 2010, **132**, 231102.
- 21 C. Hättig, W. Klopper, A. Köhn and D. P. Tew, *Chem. Rev.*, 2012, **112**, 4–74.
- 22 J. G. Hill and K. A. Peterson, *J. Chem. Phys.*, 2014, **141**, 094106.
- 23 C. A. Coulson and I. Fischer, *Philos. Mag.*, 1949, **40**, 386–393.
- 24 See, for example, G. W. King, *Spectroscopy and Molecular Structure*, Holt, Rinehart and Winston Inc., New York, 1964, pp. 164–171.
- 25 Constants of Diatomic Molecules, compiled by K. P. Huber and G. H. Herzberg, <https://webbook.nist.gov>.
- 26 A. Heidaria, O. A. Bégb and M. Ghorbania, *Russ. J. Phys. Chem. A*, 2013, **87**, 216–224.

