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High coordination number actinide-noble gas complexes; a computational study†

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The geometries, electronic structures and bonding of early actinide-noble gas complexes are studied computationally by density functional and wavefunction theory methods, and by *ab initio* molecular dynamics. AcHe_{18}^{3+} is confirmed as being an 18-coordinate system, with all of the He atoms accommodated in the primary coordination shell, and this record coordination number is reported for the first time for Th^{4+} and Th^{3+} . For Pa and U in their group valences of 5 and 6 respectively, the largest number of coordinated He atoms is 17. For AnHe_{17}^{q+} ($\text{An} = \text{Ac}$, $q = 3$; $\text{An} = \text{Th}$, $q = 4$; $\text{An} = \text{Pa}$, $q = 5$; $\text{An} = \text{U}$, $q = 6$), the average An–He binding energy increases significantly across the series, and correlates linearly with the extent of $\text{He} \rightarrow \text{An}^{q+}$ charge transfer. The interatomic exchange–correlation term V_{xc} obtained from the interacting quantum atoms approach correlates linearly with the An–He quantum theory of atoms-in-molecules delocalization index, both indicating that covalency increases from AcHe_{17}^{3+} to UHe_{17}^{6+} . The correlation energy in AnHe_{16}^{3+} obtained from MP2 calculations decreases in the order $\text{Pa} > \text{Th} > \text{U} > \text{Ac}$, the same trend found in V_{xc} . The most stable complexes of Ac^{3+} with the heavier noble gases Ar–Xe are 12 coordinate, best described as Ng_{12} cages encapsulating an Ac^{3+} ion. There is enhanced $\text{Ng} \rightarrow \text{Ac}^{3+}$ charge transfer as the Ng gets heavier, and Ac–Ng covalency increases.

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1. Introduction

The noble gas (Ng) elements, with their completely filled principal quantum shells, were considered unreactive for a long time. However, the 1962 report of the first stable Ng compound, $\text{Xe}^+[\text{PtF}_6]^-$ disproved that view,¹ and a new field of chemistry was opened. After that, many scientists devoted themselves to noble gas chemistry, and various new Ng compounds were reported from both experimental and theoretical studies.^{2–5} Among them, actinide–Ng complexes, $\text{CUO}(\text{Ar})_{4-n}(\text{Ng})_n$ ($\text{Ng} = \text{Kr–Xe}$, $n = 1–4$), were first reported by Andrews *et al.* in 2002.⁶ During the synthesis of CUO by laser ablation of U and CO in Ng matrices, it was found that the vibrational spectrum of CUO in an Ar matrix was different from that in a Ne matrix, with the spectral shift from Ne to Ar being much larger than the normal “matrix shift”. Combined with density functional theory (DFT) calculations, Andrews *et al.* assigned the spectrum in Ar to triplet CUO, but to singlet CUO in Ne. Such matrix-induced ground-state reversal suggested direct Ar–U bonding, and U–Ng bonds were also found in Kr and Xe. The interaction between U and Ng was attributed to the donation of Ng lone pair electrons into vacant

orbitals of U, suggesting that positively charged actinide compounds such as UO_2^{2+} would have stronger interaction with Ng elements, and the U–Ng complexes $[\text{UO}_2(\text{Ne})_6]^+$, $[\text{UO}_2(\text{Ng})_5]^+$ ($\text{Ng} = \text{Ar–Xe}$) and $\text{UO}_2(\text{Ng})_4$ ($\text{Ng} = \text{Ne, Ar}$) were indeed subsequently reported.^{7,8} The larger average U–Ng binding energy in $[\text{UO}_2(\text{Ng})_5]^+$ vs. $\text{UO}_2(\text{Ng})_4$ and $\text{CUO}(\text{Ng})_4$ (*e.g.* 23.6 kJ mol^{-1} , 14.6 kJ mol^{-1} and 16.2 kJ mol^{-1} respectively for $\text{Ng} = \text{Ar}$) supports the influence of charge on U–Ng interaction strength, also found in AuXe_n^{q+} .^{2,9}

Actinide–Ng (especially He) complexes are candidate molecules for achieving high coordination numbers. Coordination number (CN), originally defined as the total number of neighboring atoms directly bonded to the central atom in a molecule or ion, is a fundamental concept in coordination chemistry, and the experimental and theoretical search for compounds with high CN has a long history.^{10–13} Hermann *et al.* predicted the existence of PbHe_{15}^{2+} with a CN of 15 by DFT calculation,¹⁴ and that even higher CN systems may be achievable in charged actinide–He interactions. Motivated by this work, our group¹⁵ theoretically studied actinide ions coordinated by He atoms and reported 17-coordinated AcHe_{17}^{3+} , PaHe_{17}^{4+} and ThHe_{17}^{4+} complexes, a step forward as for several years the highest known CN was 16.^{11–13} UHe_n^{q+} ($q = 2–6$) and PaHe_n^{5+} were also studied, but no stable structures were found with $n > 15$. However, a year later, AcHe_{18}^{3+} was reported by Ozama *et al.*, using coupled cluster theory and path integral molecular dynamics (PIMD).¹⁶ According to their work, 18 He atoms can be accommodated in the first

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coordination shell, in a highly symmetrical D_{4d} structure, the increase in CN from our work potentially being due to omission of basis set superposition error (BSSE) in our CCSD(T) calculations. Although CN higher than 18 has recently been reported for $M(H_2)_{12}^{n+}$ ($M = Ac, Th, Pa, U, La, n = 3, 4$)¹⁷ and $An(BH)_{24}$ ($An = Th-Cm$),¹⁸ genuine CNs of 24 are debatable given the strong H-H interaction in $M(H_2)_{12}^{n+}$ and lack of direct evidence for true 24 M-B bonds in $An(BH)_{24}$. High coordination number is also observed in metal cluster-based compounds. $M(EH)_{12}$ ($M = Cr, Mo, W; E = Zn, Cd, Hg$)¹⁹ possess 12 M-ER bond paths which are characterized as 6 three-centre two-electron bonds by quantum theory of atoms-in-molecules (QTAIM) and MO analysis. Weak peripheral E-E bonding is also observed with lower bond order than the M-E bonds. By contrast, $[Pt@Pb_{12}]^{2-}$,^{20,21} which has similar geometry, is stabilized by strong interactions in the Pb icosahedron.

In this contribution, we study $AnHe_n^{q+}$ ($An = Ac-U$) using wavefunction theory, DFT and *ab initio* molecular dynamics (AIMD) calculations, and probe the nature of the An-He interaction. The latter is investigated *via* the QTAIM, interacting quantum atoms (IQA) and natural population analysis (NPA) methods. We also extend the helium work to the heavier Ng elements neon, argon, krypton, and xenon, systemically investigating the nature of the An^{q+} -Ng interaction, including the influence of Ng polarizability.

2. Computational details

All the geometrical structures, binding energies and thermodynamic stabilities of $AnNg_n^{q+}$ complexes were studied by dispersion-corrected density functional theory methods (DFT-D3),²² as implemented in TURBOMOLE 7.3.²³ Effective core potentials (ECPs) with 60 core electrons were used for the actinide elements along with the def-TZVPP valence basis sets,²⁴⁻²⁶ and the aug-cc-pV5Z basis set²⁷ was used for He. Stuttgart RLC basis sets and ECPs with 2, 10, 28, and 46 core electrons respectively were used for Ne, Ar, Kr, and Xe.^{28,29} To ensure valid comparisons with the He results, both the aug-cc-pV5Z and Stuttgart RLC ECP basis sets were used for benchmark $AcNe_n^{3+}$ calculations (Table S1, ESI[†]). The average Ac^{3+} -Ne distance and distance range obtained, as well as the binding energies, are close, suggesting that comparisons may indeed be justifiably made between the He and heavier Ng results, despite the difference in Ng basis sets employed. Harmonic vibrational frequency analysis³⁰ was conducted for all optimized structures to ensure that they are true minimum structures.

To benchmark the density functional selection, CCSD(T)³¹ calculations were conducted in MOLPRO 2019³² using the 60 electron ECP along with associated ECP60MWB_SEG valence basis set for An (Ac, U)²⁴⁻²⁶ and aug-cc-pV5Z basis set for He.²⁷ HF³³ and MP2³⁴ calculations were performed in Gaussian 16³⁵ with the same basis sets as used for the CCSD(T) calculations. BSSE corrections were evaluated by the counterpoise correction method.³⁶

QTAIM and IQA³⁷ calculations were conducted using the AIMAll software.³⁸ The input files for the IQA calculations were

generated from DFT calculations based on the B3LYP-D3 density functional³⁹ because the BHLYP-D3 density functional (which is used for our other production DFT calculations) is not currently supported for IQA analysis in AIMAll. The details for the IQA calculations were as reported in our previous work.⁴⁰ Natural population analysis (NPA) was carried out using NBO 7.0⁴¹ to obtain the natural charges and electron configurations of the $AnNg_n^{q+}$ complexes.

To study the thermodynamic stability of the optimized $AnNg_n^{q+}$ structures, AIMD calculations were performed in TURBOMOLE, using the Nosé-Hoover thermostat.⁴² The time step was set to 1.21 fs and the total simulation time to 1.25 ps for Ac^{3+} -Ng compounds. For the Th^{4+} -He complexes, longer simulation times were used, 2.47 ps and 3.71 ps for temperatures of 3 K and 10 K respectively. The simulations for $AcHe_{18}^{3+}$ and $ThHe_{18}^{4+}$ were conducted twice, yielding similar results. Hence, simulations for the Ac^{3+} -Ng complexes were conducted only once.

3. Results and discussion

3.1. Potential energy curves for An^{3+} -He

As stated in the Introduction, Ozama *et al.* reported that BSSE has a significant influence on the Ac^{3+} -He potential energy curve at the ECP60MWB/CCSD(T) level.¹⁶ To verify this, we began by studying BSSE, calculating the potential energy curve for the Ac^{3+} -He interaction with CCSD(T), with and without BSSE correction. Both contracted and uncontracted valence basis sets were used for Ac. As shown in Fig. 1(a), the effect of BSSE depends on whether the Ac basis set is contracted or uncontracted. The contracted basis set without BSSE gives the deepest potential well of 35.13 kJ mol⁻¹, close to the 34.98 kJ mol⁻¹ and 35.0 kJ mol⁻¹ calculated by Ozama *et al.* and by us, respectively. However, the potential well without BSSE but using the uncontracted basis set is much shallower, 22.68 kJ mol⁻¹ at 2.61 Å. BSSE correction leads to a shallowing of the potential energy curve, to 21.93 kJ mol⁻¹ and 22.18 kJ mol⁻¹ for contracted and uncontracted basis set respectively, close to the value of 21.94 kJ mol⁻¹ predicted by Ozama *et al.* This suggests that BSSE correction has a marked influence on CCSD(T) calculations using contracted basis sets, but that the impact on uncontracted basis set calculations is slight. BSSE at the DFT level was also studied, using the PBE-D3 functional as a representative example (Fig. S1, ESI[†]). The results suggest that the effect is very small, with potential well depths of 27.40 kJ mol⁻¹ and 27.46 kJ mol⁻¹ with and without BSSE correction, respectively. On the basis of these results, we (i) use BSSE-corrected contracted basis set CCSD(T) data as a benchmark to determine the most appropriate density functional (contracted basis sets are used for DFT calculations in the TURBOMOLE software) and (ii) neglect BSSE corrections in our DFT calculations.

Ac^{3+} -He potential energy curves were calculated by various density functionals; the results from PBE-D3 and BHLYP-D3 are shown in Fig. 1(b) with the other data being plotted in Fig. S2 (ESI[†]). All functionals match reasonably well with BSSE-corrected CCSD(T), with hybrid functionals performing slightly better than GGA. Among the hybrids, the BHLYP curve is the best match for



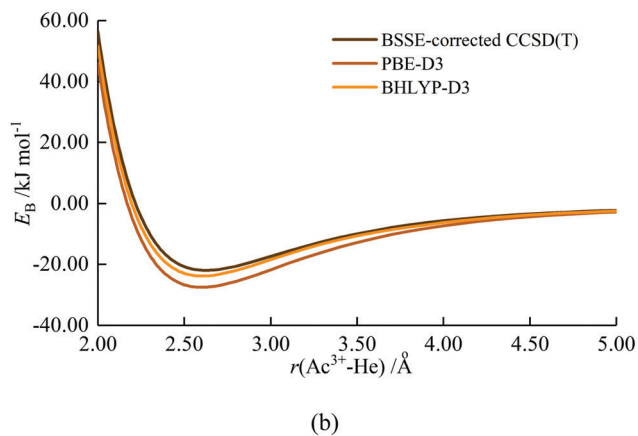
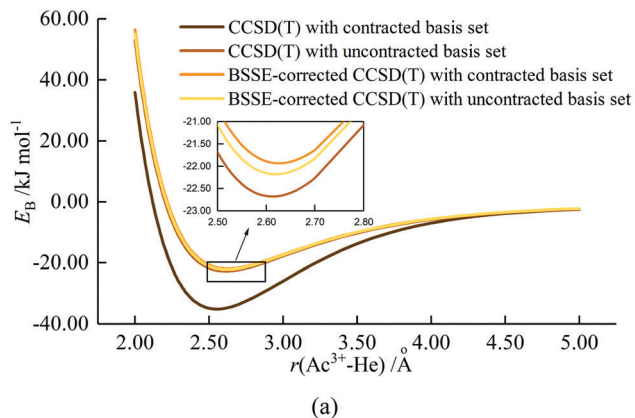


Fig. 1 Potential energy curves of Ac^{3+} -He by (a) CCSD(T) calculation and (b) DFT calculation using PBE-D3 and BHLYP-D3 density functionals.^{39,43}

the coupled cluster results, and hence BHLYP-D3 is used for geometry optimization and binding energy calculations of AnNg_n^{q+} . The only exception is UHe_n^{3+} , for which BHLYP-D3 gave a poor description of the UHe_{18}^{3+} average binding energy (Fig. S3, ESI[†]) and U^{3+} -He potential energy curve (Fig. S4, ESI[†]). Fig. S4 (ESI[†]) shows that the binding energies obtained from BHLYP-D3 are very large, leading to an extremely deep potential well which does not converge to 0 at large U^{3+} -He distance. Further examination indicates that the total energy of U^{3+} obtained by BHLYP-D3 is appreciably less negative than from the other methods employed (Table S2, ESI[†]), and also less negative than obtained from the same calculation performed in Gaussian 16, suggesting that TURBOMOLE does not converge to the correct ground state of U^{3+} with BHLYP. Given this observation and, as the curve from B3LYP-D3 matches better with that of CCSD(T), it is used for the study of U^{3+} -He complexes.

3.2. An^{q+} -He complexes

3.2.1 Ac^{3+} -He complexes. Using the BHLYP-D3 functional, we were able to locate a AcHe_{18}^{3+} true minimum structure. This is an increase in CN of one from our previous work, which employed the HFS functional,¹⁵ and located only AcHe_{17}^{3+} . The data in the previous section indicate that this difference arises from the omission of BSSE in our previous benchmarking CCSD(T) calculations, which generated an Ac^{3+} -He well which

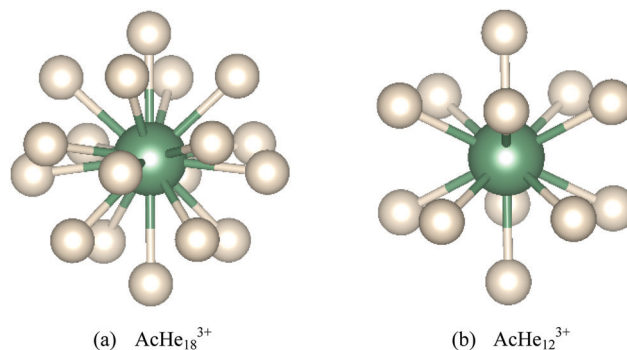


Fig. 2 Optimized geometries of AcHe_{18}^{3+} and AcHe_{12}^{3+} .

was too deep, and hence led to our using the HFS approach. The optimized structure of AcHe_{18}^{3+} is shown in Fig. 2(a); it is close to a distorted tetrakis cuboctahedron and has D_{4d} symmetry, as reported by Ozama *et al.*¹⁶ Efforts to coordinate 19 and 20 He atoms in the primary shell all failed.

The stability of AcHe_{18}^{3+} was studied by AIMD at 10 K, starting from the initial geometry shown in Fig. 2(a). The results are shown in Fig. 3; the lowest energy structure over the course of the simulation is the starting structure. Furthermore, the root mean square deviations of the He atoms from the initial structure are less than 0.15 Å, also indicating the stability of the AcHe_{18}^{3+} structure, with only slight relaxation of the He atoms in the first coordination shell.

Geometrical parameters of Ac^{3+} -He complexes from $n = 1$ to 18 are listed in Table 1. The average Ac^{3+} -He distance increases by only 0.109 Å from $n = 1$ to 17, with the bond distance range maximising at 0.039 Å for the 16-coordinate system. However, the bond distance range jumps to 0.108 Å at $n = 18$, with the average bond distance being 0.027 Å longer than that at $n = 17$, suggesting a crowded distribution of He atoms in AcHe_{18}^{3+} . However, the range of 0.108 Å is still small enough to assign all 18 He atoms to the primary coordination shell. That we find bond critical points (BCPs) between Ac^{3+} and all 18 He atoms also supports this conclusion. Note also the very small bond length range for $n = 12$ (Fig. 2(b)), 0.001 Å, rising to 0.026 Å at $n = 13$, due to disruption of the stable AcHe_{12}^{3+} icosahedral structure. The average Ac-He distance of AcHe_{17}^{3+} is 2.719 Å,

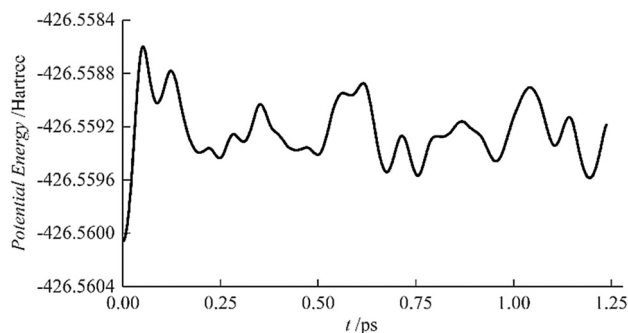
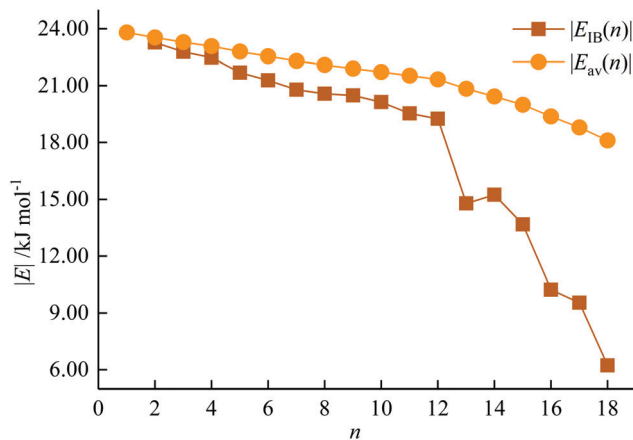


Fig. 3 Evolution of AcHe_{18}^{3+} potential energy with time at 10 K.



Table 1 Average Ac^{3+} -He distance (r_{av}) and distance range (Δr) in AcHe_n^{3+} ($n = 1-18$)

n	$r_{\text{av}}/\text{\AA}$	$\Delta r/\text{\AA}$	n	$r_{\text{av}}/\text{\AA}$	$\Delta r/\text{\AA}$
1	2.610	—	10	2.638	0.016
2	2.606	0.001	11	2.642	0.013
3	2.611	0.005	12	2.644	0.001
4	2.613	0.006	13	2.660	0.026
5	2.620	0.011	14	2.671	0.018
6	2.628	0.012	15	2.683	0.037
7	2.631	0.013	16	2.701	0.039
8	2.635	0.005	17	2.719	0.037
9	2.638	0.011	18	2.746	0.108

**Fig. 4** Incremental $|E_{\text{IB}}(n)|$ and average $|E_{\text{av}}(n)|$ binding energies of AcHe_n^{3+} ($n = 1-18$).

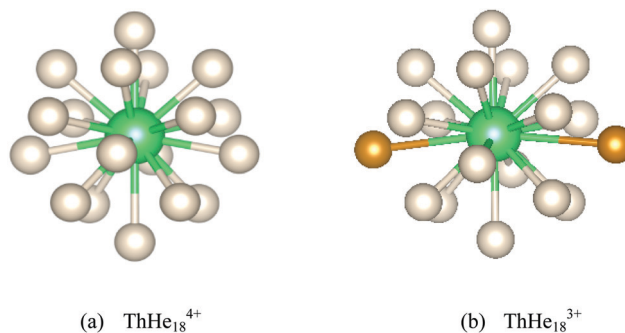
larger than the 2.671 Å that we reported before,¹⁵ consistent with the smaller revised Ac^{3+} -He well depth.

The incremental $|E_{\text{IB}}(n)|$ and average $|E_{\text{av}}(n)|$ binding energies of AcHe_n^{3+} ($n = 1-18$) are plotted in Fig. 4. $|E_{\text{IB}}(n)|$ is defined as the difference between the total energy of AcHe_n^{3+} and the sum of the total energy of AcHe_{n-1}^{3+} and He, and represents the energy gain for each He atom attachment. $E_{\text{av}}(n)$ is calculated by equation

$$E_{\text{av}}(n) = (E(n) - E(\text{Ac}^{3+}) - n \times E(\text{He}))/n$$

in which $E(n)$, $E(\text{Ac}^{3+})$ and $E(\text{He})$ are the total energy of AcHe_n^{3+} , Ac^{3+} , and He. As shown in Fig. 4, $|E_{\text{av}}(n)|$ decreases from $n=1$ to 18, which is as expected considering more He atoms are attached to the Ac^{3+} centre. $|E_{\text{IB}}(n)|$ is similar to our PBE results in ref. 15, while smaller than the HFS result. It decreases steadily from $n = 1$ to 12, and then suddenly drops at $n = 13$ before slightly increasing at 14. The energy change from $n = 12$ to 13 is consistent with the trend in Ac^{3+} -He distance range, and the increase at $n = 14$ is due to the establishment of a new stable structure as reported by Hermann and us.^{14,15} It is worth noting that the $|E_{\text{IB}}(n)|$ at $n = 18$ is only 6.24 kJ mol^{-1} which, together with the large Ac-He distance range of AcHe_{18}^{3+} , explains the failure to coordinate more than 18 He atoms.

3.2.2 Th^{q+} -He ($q = 3, 4$) complexes. The search for the highest CN of charged actinide-Ng complexes was extended to other

**Fig. 5** Optimized (a) ThHe_{18}^{4+} and (b) ThHe_{18}^{3+} geometries. The furthest two He atoms in the latter are highlighted in orange in (b).**Table 2** Geometrical parameters and average binding energies of AnHe_n^{q+} in their highest CN structure

	$r_{\text{av}}/\text{\AA}$	$\Delta r/\text{\AA}$	$ E_{\text{av}} /\text{kJ mol}^{-1}$
ThHe_{18}^{4+}	2.541	0.151	42.001
ThHe_{18}^{3+}	2.729	0.538	19.452
PaHe_{17}^{5+}	2.360	0.053	90.604
PaHe_{16}^{3+}	2.623	0.213	22.257
UHe_{17}^{6+}	2.296	0.082	179.917
UHe_{16}^{3+}	2.623	0.083	22.908

early actinide elements from Th to U. Both their group valence and trivalent oxidation states were considered. The 18-coordinated structure was also located for ThHe_{18}^{q+} ($q = 3, 4$) with the geometries shown in Fig. 5 and Table 2. The geometry and symmetry of ThHe_{18}^{4+} are similar to AcHe_{18}^{3+} with a smaller Th-He distance (2.541 Å), while the average Th^{4+} -He binding energy is 42.001 kJ mol^{-1} , more than twice that of Ac^{3+} -He, likely due to the increased charge on Th^{4+} vs. Ac^{3+} .

To study the thermodynamic stability of ThHe_{18}^{4+} , AIMD calculations were conducted at 3 K and 10 K. As shown in Fig. 6, ThHe_{18}^{4+} remains as the minimum potential energy structure at 3 K. However, as the temperature is increased to 10 K, the minimum potential energy structure changes to ThHe_{17}^{4+} , with the 18th He atom located in the second shell, *i.e.* the system is better represented as ThHe_{17+1}^{4+} . The Th-He distance range in the ThHe_{17}^{4+} part is quite small (0.040 Å), and the distance difference between the two shells is 1.213 Å. Compared with AcHe_{18}^{3+} , the primary coordination shell of ThHe_{18}^{4+} is more crowded, which leads to a reduction in the CN at higher temperature.

Although a stable ThHe_{18}^{3+} structure was obtained, the range of Th^{3+} -He distances is quite large (2.645–3.183 Å). As shown in Fig. 5(b), there are two He atoms, highlighted in orange, with an average Th^{3+} -He distance of 3.178 Å, farther out than the other 16 He atoms (average Th^{3+} -He distance = 2.673 Å), making the CN 18 description debatable. However, QTAIM analysis shows BCPs between Th^{3+} and all 18 He atoms (Fig. 7). The presence of a BCP can be considered as the indicator of a chemical bond. The average electron density at the BCPs between Th^{3+} and the two more distant He atoms is only 0.005 a.u., while the value is 0.013–0.015 a.u. between Th^{3+}



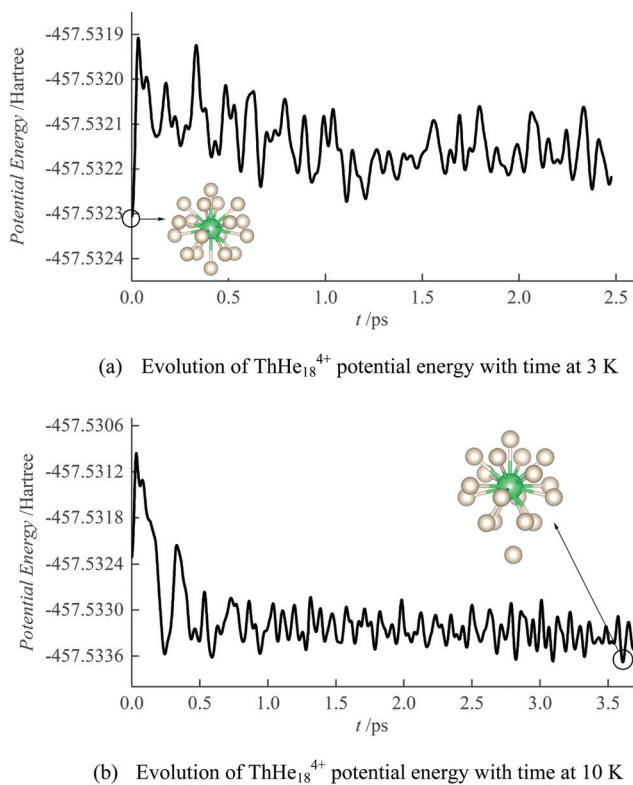


Fig. 6 Evolution of ThHe_{18}^{4+} potential energy by AIMD at (a) 3 K and (b) 10 K. The initial structures are as shown in Fig. 5(a).

and the other 16 He atoms. Therefore, we conclude that ThHe_{18}^{3+} is genuinely 18-coordinated, although two of the bonds are extremely weak. Note that the average Th^{3+} -He binding energy ($19.452 \text{ kJ mol}^{-1}$) is slightly larger than the $18.098 \text{ kJ mol}^{-1}$ of Ac^{3+} -He, but significantly less than that of Th^{4+} -He.

To further understand the actinide-Ng interaction within these highly coordinated complexes, QTAIM and NPA analysis for ThHe_{18}^{q+} ($q = 3, 4$) were conducted, and the results are

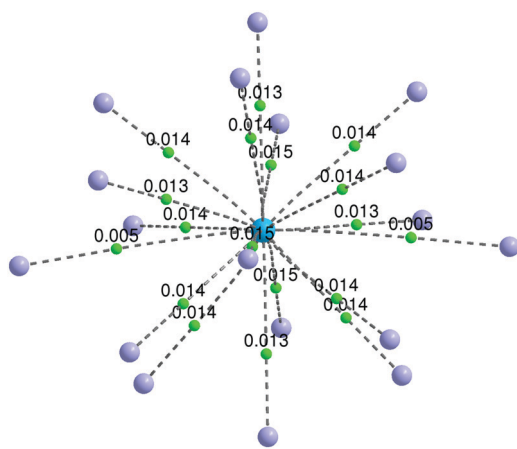


Fig. 7 Bond paths (dashed lines) and bond critical points (green dots) of ThHe_{18}^{3+} . The electron densities (a.u.) at the BCPs are displayed. He = purple spheres, Th = blue sphere.

presented in Table 3. The data for AcHe_{18}^{3+} are also given. Different from orbital-based analysis, QTAIM is based on the topology of the electron density.⁴⁴ BCP properties such as the electron density (ρ_{BCP}), Laplacian of the electron density ($\nabla^2\rho_{\text{BCP}}$), the ratio of kinetic and potential energy densities ($-G_{\text{BCP}}/V_{\text{BCP}}$), and the energy density (H_{BCP}) are all useful descriptors of chemical bonds. A general rule is that $\rho_{\text{BCP}} > 0.2 \text{ a.u.}$ and $H_{\text{BCP}} < 0$ are features of covalent interaction, while $\rho_{\text{BCP}} < 0.1 \text{ a.u.}$ represents closed-shell interactions (ionic bonding, hydrogen bonding, van der Waals bonding).⁴⁵ However, for chemical bonds involving actinides, ρ_{BCP} is rarely larger than 0.2 and usually less than 0.1 a.u. $\nabla^2\rho_{\text{BCP}}$ describes whether the electron density is concentrated ($\nabla^2\rho_{\text{BCP}} < 0$) or depleted ($\nabla^2\rho_{\text{BCP}} > 0$) at the BCP. For highly polarized bonds, positive $\nabla^2\rho_{\text{BCP}}$ is possible. For BCPs with positive $\nabla^2\rho_{\text{BCP}}$, $-G_{\text{BCP}}/V_{\text{BCP}}$ between 0.5 and 1 is a feature of covalent character, while $-G_{\text{BCP}}/V_{\text{BCP}} > 1$ is considered to indicate non-covalent interaction.⁴⁰ Moreover, the delocalization indices (δ), which represent the number of electrons shared between two atomic basis, is a widely used measure of bond order, and also a good indicator of covalence.

The QTAIM atomic charges in Table 3 indicate that the overall $q+$ charge partially redistributes between An and He due to electron transfer from the He atoms to the An centres. That said, the An QTAIM charges are close to the formal oxidation state in all cases, and the He charges are small. These charges indicate closed-shell bonding in Th^{q+} -He and Ac^{3+} -He; this has been previously characterized as a charge-induced dipole interaction.¹⁵ The BCP data and small δ values further confirm this. Although the An/He charges calculated by NPA are not as disparate as those from the QTAIM, they are also suggestive of ionic bonding. Charge transfer, as well as ρ_{BCP} and δ , for ThHe_{18}^{4+} are larger than for ThHe_{18}^{3+} and AcHe_{18}^{3+} (which are similar to one another) consistent with the average An $^{q+}$ -He binding energy being mainly determined by the charge on the metal centre, as noted above.

3.2.3 An $^{q+}$ -He (An = Pa, q = 3, 5; An = U, q = 3, 6) complexes. The highest CN obtained for Pa^{5+} -He and U^{6+} -He complexes is 17, and 16 for Pa^{3+} -He and U^{3+} -He complexes. It is not unexpected to find lower CN of Pa^{q+} and U^{q+} than Th^{q+} and Ac^{3+} , given the smaller ionic radii of the former and the increased An $^{q+}$ -He distance range from Ac^{3+} to Th^{q+} . The CNs of 17 and 16 reported here are higher than the 15 suggested in our previous work for Pa^{5+} -He and U^{q+} -He ($q = 2-6$),¹⁵ as well as in other reported actinide complexes.^{46,47}

18-Coordinated structures were also located for PaHe_{18}^{3+} and UHe_{18}^{3+} , but were found to have imaginary frequencies. To obtain true minima, the PaHe_{18}^{3+} and UHe_{18}^{3+} structures were distorted along the largest imaginary mode and re-optimized, resulting in lower energy An He_{16}^{3+} (An = Pa, U) structures with the remaining two He atoms located in the second shell. The geometries of these complexes are given in Fig. 8 and Table 2. PaHe_{17}^{5+} and UHe_{17}^{6+} have similar symmetry and geometry to one another, with the average An $^{q+}$ -He distances being 2.360 \AA and 2.296 \AA , respectively. The average Pa^{5+} -He and U^{6+} -He binding energies are much larger than that of Th^{4+} -He, which



Table 3 QAIM metrics (a.u.) and NPA charges for optimized AcHe_{18}^{3+} and ThHe_{18}^{q+} ($q = 3, 4$)

	Atomic charge				BCP properties				
	QAIM		NPA		ρ_{BCP}	$\nabla^2\rho_{\text{BCP}}$	$-(G_{\text{BCP}}/V_{\text{BCP}})$	H_{BCP}	δ
	q_{An}	q_{He}	q_{An}	q_{He}					
AcHe_{18}^{3+}	2.843	0.009	2.211	0.048	0.012	0.055	1.278	0.002	0.048
ThHe_{18}^{3+}	2.827	0.010	1.766	0.069	0.013	0.061	1.247	0.002	0.053
ThHe_{18}^{4+}	3.613	0.021	2.226	0.099	0.020	0.083	1.154	0.002	0.077

we attribute in part to the reduced steric effect induced by the reduction in CN, but primarily to the increased metal charge of Pa^{5+} and U^{6+} , as found in the comparison of ThHe_{18}^{4+} and AcHe_{18}^{3+} . For PaHe_{16}^{3+} and UHe_{16}^{3+} , the average An³⁺–He distances are the same with the distance range of the latter being smaller. The average binding energy of UHe_{16}^{3+} is slightly larger than that of PaHe_{16}^{3+} , but this increase is small compared with that from PaHe_{17}^{5+} to UHe_{17}^{6+} , as found for ThHe_{18}^{q+} complexes as noted above. This further indicates the significant influence of metal charge on An^{q+}–He interaction.

3.2.4 The nature of the An^{q+}–He interaction. In order to study the dependence of the nature of the An^{q+}–He interaction on An and oxidation state, QAIM, IQA, and NPA analyses were conducted on AnHe_{17}^{q+} and AnHe_{16}^{3+} (An = Ac–U), selected as they form families of closed-shell and trivalent AnHe_n^{q+} complexes, respectively.

AnHe₁₇^{q+}. As shown in Table 4, the average An–He distance in AnHe_{17}^{q+} decreases significantly from Ac to U, while the average binding energy increases substantially. The atomic

charges on the actinide ions deviate from their formal oxidation states, suggesting that part of the overall positive charge is acquired by the He atoms. This charge transfer increases across the series from Ac to U, with the partial charge on Ac(III) close to its formal oxidation state but increasingly less so from Th(IV) to U(VI). The average binding energy correlates linearly with this charge transfer ($R^2 = 0.999$). We noted above that the Th^{q+}–He ($q = 3, 4$) and Ac^{3+} –He interaction is primarily charge-induced dipole, with little charge transfer and small delocalization indices. This description is also suitable for the Pa–He interaction in PaHe_{17}^{5+} , given the small ρ_{BCP} value, positive H_{BCP} , and $-G_{\text{BCP}}/V_{\text{BCP}}$ greater than 1. However, considering the charge loss of 1.73 for U(VI) and small negative H_{BCP} , this description is arguably less appropriate for U–He, which shows more Lewis acid–base interaction characteristics. $-G_{\text{BCP}}/V_{\text{BCP}}$ between 0.5 and 1.0 further evidences this. Furthermore, ρ_{BCP} and δ rise from Ac to U, suggesting increasing covalency.

The exchange–correlation energy (V_{xc}) between An and He, obtained from IQA analysis, is also listed in Table 4. IQA is a real-space energy partitioning method based on topology theory, with good performance in quantifying chemical bonding, which has only recently begun to be employed in the 5f series.^{40,48} According to the IQA approach, the interatomic energy (V_{int}) can be decomposed into electrostatic (V_{Elec}) and exchange–correlation terms (V_{xc}).³⁷ The former is composed of the nuclear–nuclear repulsive energy, electron–electron coulombic repulsion energy, and the electron–nuclear attraction energy. V_{xc} is a good descriptor of the covalent contribution to the interatomic energy, and it can be seen from Table 4 that V_{xc} becomes more negative as the series is crossed. Considering that the more negative V_{xc} , the higher the degree of covalency, IQA gives the same trend as the other QAIM metrics, *i.e.* that covalency increases from Ac to U in this closed-shell family. As shown in Fig. 9(a), V_{xc} shows excellent correlation with δ ($R^2 = 0.991$), which is a more widely used indicator of covalency. This further suggests that the IQA method can be a useful tool in the study of covalency in the 5f series.

To further analyze the electronic structures of these complexes, the natural charges and natural electron configurations were explored using NPA, and are listed in Table 5, together with the Wiberg bond indices (WBIs). The trend in the NPA charges of He in AnHe_{17}^{q+} is the same as that obtained from QAIM but with more charge being acquired by He, similar to the behaviour noted above, indicating less closed-shell interaction compared with QAIM. Like V_{xc} , WBI is very strongly correlated with δ ($R^2 = 0.985$). Given the charge transfer, it is unsurprising that the natural electron configurations show

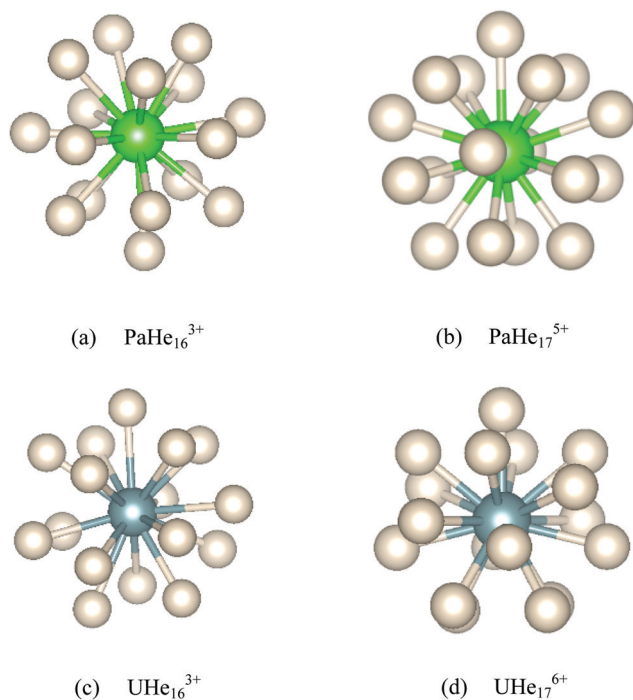
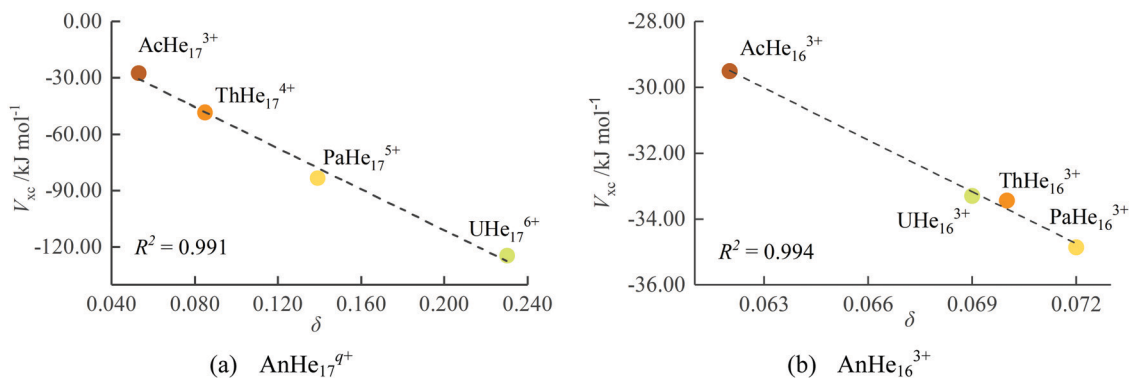


Fig. 8 Optimized AnHe_n^{q+} (An = Pa, $q = 3, 5$; An = U, $q = 3, 6$) configurations.



Table 4 Average An^{q+}–He distance, An^{q+}–He binding energies, QTAIM (a.u.) and IQA metrics for the optimized AnHe_n^{q+} complexes

	$r_{av}/\text{\AA}$	$E_{av}/\text{kJ mol}^{-1}$	q_{An}	q_{He}	ρ_{BCP}	$\nabla^2\rho_{BCP}$	$-\frac{G_{BCP}}{V_{BCP}}$	H_{BCP}	δ	$V_{xc}/\text{kJ mol}^{-1}$
AcHe ₁₇ ³⁺	2.719	-18.796	2.802	0.009	0.013	0.058	1.263	0.002	0.053	-27.543
ThHe ₁₇ ⁴⁺	2.511	-44.121	3.604	0.023	0.022	0.090	1.138	0.002	0.085	-48.296
PaHe ₁₇ ⁵⁺	2.360	-90.604	4.144	0.050	0.033	0.118	1.023	0.001	0.139	-83.375
UHe ₁₇ ⁶⁺	2.296	-179.917	4.259	0.102	0.040	0.123	0.958	-0.001	0.230	-124.580
AcHe ₁₆ ³⁺	2.701	-19.374	2.802	0.012	0.014	0.059	1.247	0.002	0.062	-29.507
ThHe ₁₆ ³⁺	2.650	-21.270	2.779	0.014	0.015	0.067	1.213	0.002	0.070	-33.444
PaHe ₁₆ ³⁺	2.623	-22.257	2.761	0.015	0.016	0.071	1.198	0.003	0.072	-34.859
UHe ₁₆ ³⁺	2.623	-22.908	2.754	0.015	0.015	0.069	1.198	0.002	0.069	-33.302

Fig. 9 The correlation between V_{xc} and δ for AnHe_n^{q+}.

significant deviation from the formal oxidation states, in which the valence orbitals are expected to be empty. The occupancies of all valence orbitals increase from AcHe₁₇³⁺ to UHe₁₇⁶⁺, although the details differ. For AcHe₁₇³⁺, the electrons mainly occupy the 6d orbitals, with the 5f occupancy being 0.08. Although the 6d populations are also larger than 5f for ThHe₁₇⁴⁺ and PaHe₁₇⁵⁺, the 5f increases significantly and, for UHe₁₇⁶⁺, the occupation of 5f exceeds that of 6d. The increasing occupation of the 5f orbitals in part reflects their stabilization across the actinide series, as reported in previous theoretical work.^{49,50}

Table 5 Wiberg bond index (WBI) of An^{q+}–He, natural charge, and natural electronic configuration of An^{q+} in AnHe₁₇^{q+} and AnHe₁₆³⁺. For the electronic configurations, the values listed are those above the formal populations (shown in parentheses). As the 7p populations are negligible, they are not shown. Note that there are 0.22, 0.29, and 0.15 metal Rydberg electrons in ThHe₁₆³⁺, PaHe₁₆³⁺ and UHe₁₆³⁺ respectively

	WBI	q_{An}	$q_{He}(av)$	Natural electronic configuration		
				7s	5f	6d
AcHe ₁₇ ³⁺	0.054	2.251	0.044	(0) 0.18	(0) 0.08	(0) 0.48
ThHe ₁₇ ⁴⁺	0.119	2.338	0.098	(0) 0.29	(0) 0.44	(0) 0.92
PaHe ₁₇ ⁵⁺	0.187	2.292	0.159	(0) 0.36	(0) 1.02	(0) 1.31
UHe ₁₇ ⁶⁺	0.286	1.789	0.248	(0) 0.37	(0) 2.23	(0) 1.58
AcHe ₁₆ ³⁺	0.053	2.304	0.044	(0) 0.17	(0) 0.17	(0) 0.45
ThHe ₁₆ ³⁺	0.138	1.827	0.073	(0) 0.21	(1) 0.10	(0) 0.65
PaHe ₁₆ ³⁺	0.151	1.714	0.080	(0) 0.22	(2) 0.11	(0) 0.67
UHe ₁₆ ³⁺	0.136	1.846	0.072	(0) 0.21	(3) 0.12	(0) 0.67

AnHe₁₆³⁺. As discussed in Section 3.1, the B3LYP-D3 functional failed to convincingly describe U³⁺-based complexes, and hence we employed B3LYP-D3 for calculations on UHe₁₆³⁺. To systematically study trivalent An–He, QTAIM and NBO analyses for AnHe₁₆³⁺ (An = Ac–U) were all conducted from single-point calculations using B3LYP-D3, and the results are collected in Table 4. The average An–He distance shortens across the early actinide series, following the An trivalent radius contraction; that $r(\text{Pa–He})$ and $r(\text{U–He})$ are the same as one another is likely due to the very similar ionic radii of Pa³⁺ (1.04 Å) and U³⁺ (1.03 Å). The average binding energy increases slightly from AcHe₁₆³⁺ to UHe₁₆³⁺, and the values for the Th, Pa and U systems are significantly smaller than in the corresponding AnHe₁₇^{q+}. Moreover, the energy difference between AnHe₁₆³⁺ and the corresponding closed-shell AnHe₁₇^{q+} rises from Ac to U, further illustrating the strong influence of metal charge on the An^{q+}–He interaction.

Turning to the QTAIM metrics, the small values of ρ_{BCP} and δ , and the positive $\nabla^2\rho_{BCP}$ and H_{BCP} , all indicate charge-induced dipole interaction for all AnHe₁₆³⁺ complexes. The charge difference between metal and He decreases slightly from Ac to U, although this change is small compared with the closed-shell AnHe₁₇^{q+} complexes. The BCP electron densities are very similar for all four complexes, with that for Ac–He being slightly smaller, similar to the charge transfer trend. $-G_{BCP}/V_{BCP}$ also suggests that the Ac system is the most ionic, as does δ , from which the covalency trend is Pa > Th > U > Ac. V_{xc} indicates the same trend and once again shows excellent



correlation with δ ($R^2 = 0.994$, Fig. 9(b)). Notably, the difference in δ and V_{xc} between the trivalent AnHe_{16}^{3+} and group valent AnHe_{17}^{q+} complexes increase very significantly from Ac to U, suggesting significantly greater covalency in the higher oxidation state systems. To place the QTAIM metrics for the An–He complexes in context, the values of ρ_{BCP} , δ , and V_{xc} as a function of the shortening and elongation of the An–He bond length are studied for An^{3+} –He diatomic systems (Fig. S5, ESI†).

The natural charge and WBI data (Table 5) show the same trend as the QTAIM and IQA metrics, and once again the correlation of WBI with δ is very strong ($R^2 = 0.978$). As discussed above, the enhanced charge transfer predicted by NPA suggests less of a closed-shell interaction nature. The natural electronic configuration data show that the largest enhancement in all valence orbitals is in the 6d, indicating that the An 6d orbital is the principal acceptor of He electron density in AnHe_{16}^{3+} .

In order to study the effect of correlation on the An^{3+} –He bonding, HF and MP2 calculations were performed at the DFT optimized geometries, and the results are plotted in Fig. S6 (ESI†). Although the average binding energies are smaller at the HF level than the DFT, the trend as function of actinide is similar, with a slight increase from Ac to U. At the MP2 level, the binding energies are extremely close to the DFT values for Th to U, but that for AcHe_{16}^{3+} is rather larger than from DFT. As with the CCSD(T) data presented in Section 3.1, correction for BSSE significantly reduces the depth of the potential energy curve of Ac^{3+} –He at the MP2 level, bringing it close to the DFT value and resulting in a trend similar to that from DFT and HF. BSSE has less of an effect on the MP2 data for the later members of our target series. The correlation energy, obtained as the difference between the MP2 and HF results, is largest for Pa (8.62 kJ mol^{-1}) and smallest for Ac (5.52 kJ mol^{-1}) with an overall trend of $\text{Pa} > \text{Th} > \text{U} > \text{Ac}$. This is the same trend as V_{xc} , further demonstrating the reliability of the IQA method. The strengthening of the An^{3+} –He bond with the inclusion of MP2-level correlation is reminiscent of the increase in the closed-shell metallophilic interaction in $[\text{Cl-M-PH}_3]_2$ ($\text{M} = \text{Cu, Ag, Au, Rg}$) from HF to MP2.⁵¹

3.2.5 Ac^{3+} –Ng (Ng = Ne–Xe) complexes. To investigate the influence of noble gas species on the structure and bonding of AcNg_n^{q+} complexes, we extended our work to the heavier noble gas elements Ne–Xe. DFT search for the highest CN of AcNg_n^{3+} (Ng = Ne–Xe) complexes resulted in an n of 15, 10, 10, and 8 for Ng = Ne, Ar, Kr, and Xe, respectively, with the structures shown

in Fig. 10. The highest CN reduces down the Ng group, which is not unexpected considering the increasing atomic radii of the Ng elements. The highest CNs of Ac^{3+} –Ar and Ac^{3+} –Kr are both 10, with similar structure, and the average metal–ligand distance of the latter (3.383 \AA) is larger than the former (3.173 \AA). AIMD calculations were performed to confirm the thermodynamic stability of these structures (Fig. S7, ESI†), and all of them remain as the potential energy minimum during the whole relaxation time at 10 K.

During the AIMD relaxations, an interesting effect was discovered. It is logical that, for a compound consisting of n Ng atoms and an An ion, if n is larger than the highest CN, the system will have the highest-coordinated geometry with the remaining Ng atoms accommodated in the second shell. If n is less than the highest CN, an n -coordinated structure is expected. But this is not true for AcAr_{12}^{3+} (Ng = Ar–Xe). The optimized structures of AcAr_n^{3+} ($n = 10$ –12, 14) are shown in Fig. 11. It can be seen that for $n = 10, 11$ and 14, 10-coordinated structures with the remaining Ar atoms in the second shell are obtained after AIMD relaxation at $T = 10 \text{ K}$. For $n = 12$, however, the minimum potential energy structure is the icosahedral 12-coordinated one, which is stable even at $T = 100 \text{ K}$ (as shown in Fig. S8, ESI†). The 10-coordinated structure with 2 more Ar atoms in the second shell was also obtained, but its energy is $11.64 \text{ kJ mol}^{-1}$ higher than that of AcAr_{12}^{3+} . This 12-coordinated structure is also located for Kr and Xe.

To further study these AcNg_{12}^{3+} structures, QTAIM analysis was conducted, with the results being presented in Table 6. QTAIM analysis of the actinide-free Ng_{12} clusters were also conducted (Table 6). BCPs are found between Ac^{3+} and Ng atoms (Ng = He–Xe) in all AcNg_{12}^{3+} , confirming direct Ac^{3+} –Ng bonding. For the Ng_{12} clusters, BCPs are present between neighbouring Ng atoms in all systems (Ng = He–Xe). In AcHe_{12}^{3+} and AcNe_{12}^{3+} , the Ng–Ng BCPs disappear, suggesting that the stability of these two complexes is mainly due to the interaction between Ac and Ng. However, for AcNg_{12}^{3+} (Ng = Ar–Xe), BCPs are retained between neighbouring Ng atoms, indicating that Ng–Ng interactions also contribute to the stability of these structures. As shown in Table 6, the electron densities at these Ng–Ng BCPs (Ar–Xe) are uniform down the group, even though the average Ng–Ng distance increases from 3.44 \AA for Ar to 4.09 \AA for Xe. The delocalization indices also increase from Ar to Xe, indicating enhanced interaction between Ng and Ng. Moreover, the Ng–Ng BCP data in AcNg_{12}^{3+} (Ng = Ar–Xe) are slightly larger than that in the bare Ng_{12} . It seems that the 12 Ng atoms in AcNg_{12}^{3+}

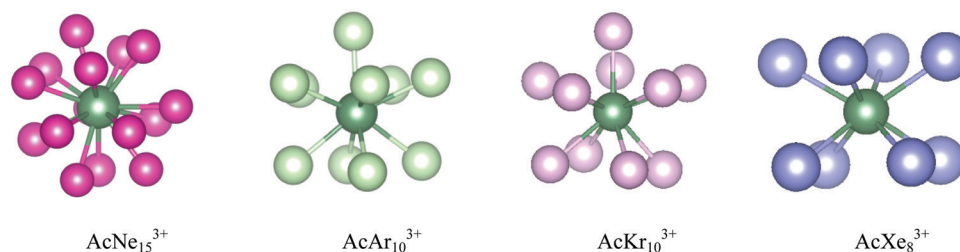


Fig. 10 Optimized highest CN structures of Ac^{3+} –Ng.



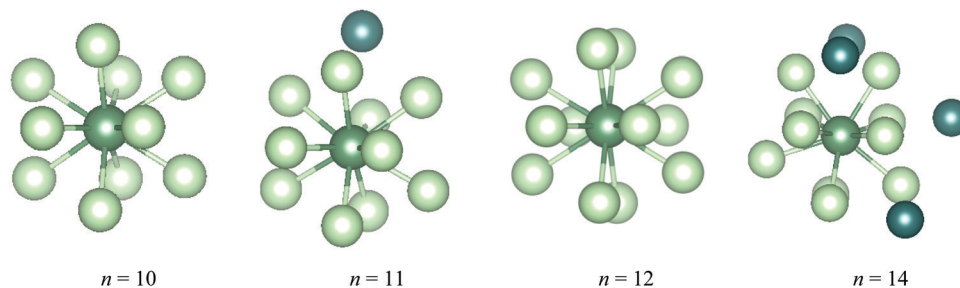


Fig. 11 The lowest potential energy structures of AcAr_n^{3+} determined by AIMD at $T = 10$ K. The Ar atoms in the second shell are highlighted in dark green.

Table 6 Structural, energetic, and QTAIM parameters (a.u.) for Ac–Ng and Ng–Ng BCPs in AcNg_{12}^{3+} complexes. The QTAIM parameters for Ng–Ng BCPs in bare Ng_{12} are listed in parentheses

	$r_{\text{av}}/\text{\AA}$	$ E_{\text{av}} /\text{kJ mol}^{-1}$	q_{Ac}	$q_{\text{Ng}}(\text{av})$	Ac–Ng			Ng–Ng		
					ρ_{BCP}	$\nabla^2\rho_{\text{BCP}}$	δ	ρ_{BCP}	$\nabla^2\rho_{\text{BCP}}$	δ
AcHe_{12}^{3+}	2.644	21.338	2.848	0.013	0.016	0.069	0.072	— (0.002)	— (0.009)	— (0.007)
AcNe_{12}^{3+}	2.779	33.445	2.800	0.017	0.019	0.092	0.092	— (0.003)	— (0.021)	— (0.015)
AcAr_{12}^{3+}	3.272	66.859	2.550	0.038	0.015	0.047	0.118	0.007 (0.006)	0.026 (0.025)	0.052 (0.048)
AcKr_{12}^{3+}	3.521	78.037	2.478	0.043	0.012	0.033	0.120(7)	0.007 (0.006)	0.023 (0.022)	0.067 (0.061)
AcXe_{12}^{3+}	3.887	88.278	2.403	0.050	0.009	0.019	0.121(3)	0.006 (0.006)	0.017 (0.016)	0.084 (0.075)

(Ng = Ar–Xe) form a cage, trapping Ac^{3+} within and enhancing the binding strength of Ng_{12} in return. This phenomenon is also found in ThNg_{12}^{4+} (geometrical and QTAIM parameters are listed in Table S3, ESI[†]). In summary, we suggest that the highest CN for AcNg_n^{3+} (Ng = Ar–Xe) is 12 considering the BCPs found between Ac^{3+} and all Ng atoms, although not all AcNg_n^{3+} structures are stable from $n = 1$ to 12 (e.g. AcAr_{11}^{3+} , AcKr_{11}^{3+} , AcXe_n^{3+} , $n = 9$ –11). The bonding situation in these complexes are similar to the $\text{M}(\text{EH})_{12}$ ¹⁹ which possess strong radial M–EH bonding and weak peripheral E–E bonding.

The partial positive charge found on the Ng increases down the group (Table 6), with a corresponding decrease in q_{Ac} , reflecting the increasing polarizability of the heavier Ng elements, and hence greater transfer of electrons from $\text{Ng} \rightarrow \text{Ac}$. The small values of ρ_{BCP} and δ , as well as the positive $\nabla^2\rho_{\text{BCP}}$, indicate the closed-shell nature of the Ac^{3+} –Ng interaction. It is interesting to note that the Ac^{3+} –Ng ρ_{BCP} data increase from He to Ne before reducing in the heavier members of group 18. This is inconsistent with the metal–Ng interaction trend reported in previous theoretical work, which finds the largest BCP data at metal–Xe,⁵² but the difference in behaviour likely arises from the direct Ng–Ng interaction in AcNg_{12}^{3+} (Ng = Ar–Xe). The delocalization indices rise from He to Xe, although the change from Ar to Xe is quite small compared to the increase in binding energies, also due to the Ng–Ng interaction.

The effect of outer coordination on the properties of AcNg_{12}^{3+} are also studied. AIMD simulation of an Ac^{3+} – Ar_{42} cluster clearly shows separate first, second and outer Ar atom shells (Fig. S9, ESI[†]) with the former two containing 12 and 20 Ar atoms, respectively. The structure of Ac^{3+} –Ar with two Ar shells is shown in Fig. S10 (ESI[†]). It can be seen that the second

shell Ar has the same symmetry as that of the first shell. This structure is also located for Ne and Kr. The structural and QTAIM parameters for these compounds are shown in Table S4 (ESI[†]). Comparison with Table 6 shows that the average Ac–Ng bond distance decreases slightly as the second shell is included. Moreover, the cluster size has no effect on the trend in Ac–Ng and first shell Ng–Ng BCP properties. It is worth noting there are very weak bonds between the Ng atoms in the second shell even though the distances are large (e.g. ca. 4 Å for Ar).

4. Conclusions

In this contribution, we have studied the geometric and electronic structures and bonding of a range of complexes involving early actinide ions and the noble gases. We began by identifying the most appropriate form of DFT to study these systems, benchmarking the An^{3+} –He interaction energy curves against BSSE-corrected coupled cluster data, and concluding that dispersion-corrected B3LYP is a good choice for $\text{An} = \text{Ac}$ –Pa and for U(vi), with B3LYP–D3 being better for U(III). Using this approach, we verified that 18 He atoms can be accommodated in the primary coordination shell of Ac^{3+} , and checked the stability of this system by AIMD. The very low incremental binding energy for the 18th He supports our conclusion that a coordination number of 18 is the maximum. We also report new 18-coordinate Th^{4+} –He and Th^{3+} –He complexes. The former is stable at 3 K by AIMD, but increasing the temperature to 10 K leads to one He migrating into the second coordination shell. In ThHe_{18}^{3+} , two of the He are ca. 0.5 Å further from the metal than



the other 16, but QTAIM analysis supports the CN = 18 description. Further QTAIM and NPA of the 18 coordinated systems indicates closed shell bonding, supporting previous conclusions of charge-induced dipole interactions.

For Pa⁵⁺-He and U⁶⁺-He complexes, the highest CN is 17, whereas for Pa³⁺ and U³⁺ it is 16. As expected, the average An-He binding energies in the 17 coordinate group valent species is significantly larger than in all four An(III) systems, and there is greater transfer of charge from He → An^{q+} in the more highly charged complexes. The average An-He binding energy correlates linearly with this charge transfer. The interatomic exchange-correlation energy V_{xc} determined by the IQA approach correlates linearly with the more widely used QTAIM delocalization index covalency metric, both indicating that An-He covalency increases from AcHe₁₇³⁺ to UHe₁₇⁶⁺. NPA of these 17 CN systems shows that the He → An^{q+} charge transfer is primarily into metal 6d orbitals for Ac, Th and Pa, but mainly into 5f for U.

QTAIM analysis of the bonding in AnHe₁₆³⁺ suggests charge-induced interactions in all cases. HF and MP2 calculations on these systems indicate that the trend in correlation energy is Pa > Th > U > Ac; the same trend is seen in V_{xc} .

Finally, we studied complexes of Ac³⁺ with the heavier noble gases. DFT searches show that the maximum coordination number decreases down the group, being 15, 10, 10 and 8 for Ne, Ar, Kr and Xe respectively. These complexes are confirmed as being stable at 10 K by AIMD. However, further AIMD analysis suggest that the highest coordination number for the Ar, Kr and Xe systems is actually 12 as the 12 Ac³⁺-Ng bond paths are confirmed by QTAIM. Comparison of the bonding in AcNg₁₂³⁺ (Ng = He-Xe) using QTAIM shows enhanced Ng → Ac³⁺ charge transfer as the Ng gets heavier, and the Ac-Ng delocalization index increases.

Conflicts of interest

There are no conflicts to declare.

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