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Pickup and reactions of molecules on clusters relevant for atmospheric and interstellar processes

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In this perspective, we review experiments with molecules picked up on large clusters in molecular beams with the focus on the processes in atmospheric and interstellar chemistry. First, we concentrate on the pickup itself, and we discuss the pickup cross sections. We measure the uptake of different atmospheric molecules on mixed nitric acid–water clusters and determine the accommodation coefficients relevant for aerosol formation in the Earth's atmosphere. Then the coagulation of the adsorbed molecules on the clusters is investigated. In the second part of this perspective, we review examples of different processes triggered by UV-photons or electrons in the clusters with embedded molecules. We start with the photodissociation of hydrogen halides and Freon CF₂Cl₂ on ice nanoparticles in connection with the polar stratospheric ozone depletion. Next, we mention reactions following the excitation and ionization of the molecules adsorbed on clusters. The first ionization-triggered reaction observed between two different molecules picked up on the cluster was the proton transfer between methanol and formic acid deposited on large argon clusters. Finally, negative ion reactions after slow electron attachment are illustrated by two examples: mixed nitric acid–water clusters, and hydrogen peroxide deposited on large Ar_N and (H₂O)_N clusters. The selected examples are discussed from the perspective of the atmospheric and interstellar chemistry, and several future directions are proposed.

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1 Introduction

Clusters of atoms and molecules have been investigated in molecular beams since the conception of the molecular beam technique and became a very broad field covered in many reviews and textbooks.^{1–8} One of the original motivations to study clusters stems from the aspiration to understand bulk materials based on the properties of the individual atoms and molecules, since the clusters can bridge the gas and condensed phases.^{5,6} From another perspective, clusters can provide insight into the solvent effects in chemistry and physics at a molecular level.^{9,10} In addition, clusters have their own intrinsic properties, different from the gas and condensed phases, that give rise to the cluster-specific physics and chemistry,^{4,6,11} which find many applications, *e.g.*, in nanocatalysis.^{12,13} The modern cluster research spans different disciplines from physics and chemistry, to biology, materials science and astronomy.¹⁴

The clusters in molecular beams can be doped by other molecules and used as nanomatrices or nanoreactors for molecular spectroscopy and chemical reactions. In this respect, large helium clusters known as He-nanodroplets have been largely exploited and represent a rather special case. They can efficiently capture molecules and provide an extremely cold (0.37 K)¹⁵ and superfluid environment for molecular spectroscopy and reactions, as outlined in many studies and recent reviews.^{16–21}

The pickup of molecules and their reactions especially on large argon clusters were implemented to investigate the solvent effect on reaction dynamics, and the pioneering technique was termed Cluster Isolated Chemical Reactions (CICR) and reviewed by Mestdagh *et al.* more than 20 years ago.²² For the product detection, most of the early CICR experiments considered chemiluminescence, namely the prototypical reaction of the barium atom with N₂O yielding the chemiluminescent product BaO.²³ Further reactions of Ba with CH₄, Cl₂, O₂, CO₂ and SF₆ were studied.^{24–29} The CICR technique has been further exploited for investigations of reactions,³⁰ spectroscopy,³¹ photo-induced Ca + HBr and Ca + CH₃F reactions^{32,33} or more recent investigation of solvation dynamics.^{34,35} The rare gas clusters were also implemented in the investigation of the solvent effect on the

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photodissociation dynamics in Buck's group, where the molecules were deposited in the clusters and dissociated by UV lasers.^{36–38} Our present work, in a sense, follows up on the CICR methodology, and extends it towards applications in atmospheric and interstellar chemistry.

From the point of view of astrochemistry, ice and dust grains offer surfaces for chemical reactions in the environment of interstellar medium, where collisions between molecules are infrequent and inefficient for reactions to proceed.^{39–41} The processes on interstellar ice and dust nanoparticles can be mimicked by reactions of molecules on clusters in molecular beam experiments. From the perspective of atmospheric chemistry, the cluster research can provide detailed insight into aerosol formation and reactions.^{42,43} Despite huge experimental and theoretical efforts, aerosol particles still belong to the least known players in atmospheric chemistry.^{42,44–47} Nevertheless, their importance is unquestionable; among many other effects, they provide surfaces for heterogeneous chemistry and photochemistry.^{44,48,49} Even small pure water clusters were found to mediate chemistry in the atmosphere.⁵⁰ Clusters in molecular beams can represent proxies for atmospheric and interstellar aerosols in laboratory experiments. Early studies by Castleman's group investigated atmospherically relevant phenomena, such as nitric acid solvation and dissolution in water clusters,^{51,52} and further cluster ion reactions.^{53,54} Many recent examples exist in the literature exploiting mass spectrometry,^{43,55–57} spectroscopy,^{58–61} photodissociation^{62,63} and other molecular beam experiments with atmospherically and astrochemically relevant clusters.^{43,64}

Here, we focus on the special experiments with clusters, where different molecules are deposited on the clusters in pickup processes, and subsequently, reactions between the molecules on/in the clusters can be triggered by photons or electrons, as illustrated schematically in Fig. 1. We review some of our work and the related experiments of other groups in the broader context of atmospheric and interstellar chemistry. In the spirit of the perspective article, we also report some unpublished results, and suggest future directions in this field.

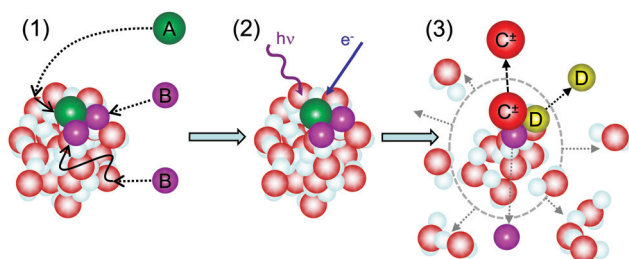


Fig. 1 Schematic illustration of the cluster processes reviewed here from the perspective of atmospheric and interstellar physics and chemistry: (1) pickup of molecules A and B on clusters and their coagulation (Section 3), and (2) reactions triggered by electrons and/or photons $A + B + e^- (hv) \xrightarrow{\text{cluster}} C^\pm + D$ (Section 4); (3) the reaction products can be ejected or the clusters decay upon excitation and the positively or negatively charged products C^\pm are detected.

The article is organized as follows: Section 2 briefly describes our cluster beam apparatus. The pickup technique is introduced in Section 3, and exploited for cluster characterization (Section 3.1), and for mimicking the processes of atmospheric aerosol nucleation (Section 3.2) using pure ice nanoparticles (Section 3.2.1) and nitric acid–water clusters (Section 3.2.2). In Section 3.3, pickup on clusters relevant for interstellar chemistry is briefly discussed. The molecules adsorbed on large clusters can coagulate, which is discussed in Section 3.4. Reactions on the clusters are discussed in Section 4, starting with photodissociation (Section 4.1). Further, we show examples of reactions on the clusters between the adsorbed molecules after their ionization or excitation by photons or electrons (Section 4.2). Finally, we review some reactions initiated by slow electron attachment (Section 4.3). We close by summarizing the general observations and possible future directions.

2 CLUB experiment

The cluster pickup experiments reviewed here were mostly performed using the cluster beam apparatus (CLUB) in Prague (Fig. 2).⁴³ The CLUB is based on the apparatus used originally for photodissociation experiments in the group of U. Buck,³⁶ which was extended in our laboratory as a versatile tool for various cluster experiments. The different experiments can be done with the same cluster beam probing the same species and thus the results can provide information about different properties of the same clusters.

The clusters are produced in the nozzle chamber (NC) by supersonic continuous expansion through a conical nozzle (typically 50–130 μm in diameter, full opening angle of 30 deg, length 2 mm). The nozzle can be attached to different cluster sources. A simple rare gas cluster source consists essentially of a tube kept at a controlled temperature *via* external heating and cooling. More sophisticated sources are used for evaporation of liquid and solid samples such as polycyclic hydrocarbons and simple biomolecules. A special case is the recently developed source of microsolvated molecules described in ref. 65. Fig. 2 shows the CLUB with the

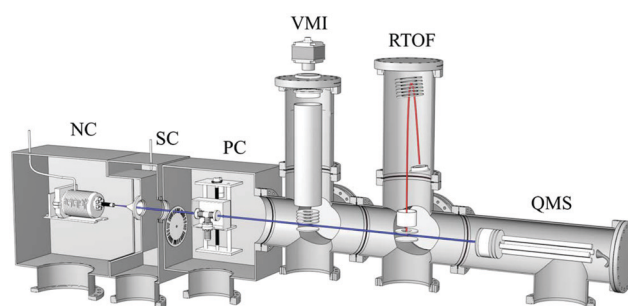


Fig. 2 Schematic overview of the cluster beam apparatus (CLUB): nozzle chamber (NC), scattering chamber (SC), pickup chamber (PC) with a pickup cell and a pseudorandom chopper, velocity map imaging (VMI), reflectron time-of-flight mass spectrometer (RTOF), and quadrupole mass spectrometer (QMS).



sketch of a water cluster source in the NC. The continuous expansion allows achieving stable and well reproducible cluster generation, and the conical nozzle shape ensures efficient clustering and sufficient beam densities along the entire cluster beam path of about 2 m.

The cluster beam passes through a skimmer into a differentially pumped scattering chamber (SC), which is used as the first pickup chamber in the present pickup experiments. In the next vacuum chamber (PC), further two pickup cells can be inserted in the cluster beam. In Fig. 2, only one pickup cell is drawn, which has recently been developed for the pickup of hydrogen peroxide.⁶⁶ A pseudorandom chopper is located in this vacuum chamber as well, which allows for accurate velocity measurements of the clusters, exploited in the cross section measurements discussed below.^{67–69}

The following chamber is devoted to the photodissociation experiments using velocity map imaging (VMI). Our VMI setup and its first implementation in the CLUB are described elsewhere.^{62,70} Various UV lasers are available in the laboratory for the photolysis and resonance enhanced multiphoton ionization (REMPI) of the photodissociation fragments, and IR lasers can be also used for IR-UV experiments.⁷¹

The reflectron time-of-flight mass spectrometer (RTOF) in the next chamber was first described in ref. 72 and 73. The clusters can be ionized by different methods: electron ionization introduced in the first studies,^{72,73} multiphoton processes using UV lasers,^{74,75} or by the special method of Na-doping and subsequent electron photodetachment.^{72,76,77} We use also electron attachment and negative cluster ion mass spectrometry.⁷⁸ These different ionization methods can produce quite different mass spectra from the same clusters, and information about the original neutral clusters can be recovered by their combination. An example represents the investigation of the mixed nitric oxide–water clusters by electron ionization, Na-doping, and electron attachment.⁷⁹

Some of our mass spectrometry experiments have recently been reviewed.⁴³ The power of the combination of VMI with high-resolution RTOF mass spectrometry has been illustrated for a few examples investigating the photodissociation dynamics of molecules in clusters, where the complex information about the nature of the clusters could not be revealed by VMI or mass spectrometry experiments alone.^{80,81}

A quadrupole mass spectrometer (QMS) with continuous electron ionization is mounted in the last chamber. It can be used to monitor the molecular beam, or to measure the beam velocity by the time-of-flight in connection with the chopper in the PC chamber.^{67,68}

3 Uptake of molecules on clusters

The pickup technique has been used since the 1980s⁸² for doping clusters with different molecules as the cluster beam passes through a diluted gas and the gas molecules are embedded in the clusters upon sticking collisions. It has been exploited especially in experiments with He-nanodroplets, where

numerous examples exist, *e.g.*, for spectroscopy,^{15,18,19,83,84} reactions,^{20,21,85–90} and generation of clusters⁹¹ in the cold superfluid environment of He-nanodroplets. Interesting applications have recently been developed, such as the decoration and visualization of superfluid vortices with atoms in He-nanodroplets.^{92,93}

Here we focus on large water clusters $(\text{H}_2\text{O})_N$ (ice nanoparticles) and other molecular species such as, *e.g.*, nitric acid–water clusters. We often perform analogous pickup experiments also with the Ar_N clusters for comparison of the molecular cluster with the inert environment of the rare-gas cluster. The pickup of molecules on Ar_N was reviewed in the context of the CICR method.²² Huisken and Stemmler used Ar_N for infrared spectroscopy of adsorbed methanol clusters.⁹⁴ The pickup of hydrogen halide molecules on large rare gas clusters was implemented in Buck's group for the photodissociation studies.^{36–38} Among molecular clusters, especially water clusters were used for uptake of different molecules.^{95–98} A rather special example is the pickup of Na atoms,^{99,100} which was exploited for mass spectrometry and size resolved investigations of large water, ammonia, and other clusters.^{101–104}

An interesting implementation of the pickup technique was exploited to develop the Exchange Metal Cluster source (EXMEC):¹⁰⁵ argon clusters were first generated by supersonic expansion, which then entered the pickup cell with metal vapor or other atoms or molecules; the picked up species eventually replaced the Ar atoms in the cluster completely, and clusters composed purely of metal atoms or other species then left the source. This method was used to produce small neutral clusters of several different elements and compounds, especially alkali halide clusters.^{106–109} Water clusters were used in these experiments as the primary clusters to investigate the solvation effects on alkali-halides.^{110,111}

3.1 Pickup cross section

In a pickup experiment, the cluster flies through a very diluted gas and collides with the gas molecules, which can stick to the cluster. This process can be characterized by the cluster pickup cross section. The pickup cross section depends on the cluster size, *i.e.*, on its geometrical cross section, and on the size of the molecule (see Fig. 3). The geometrical pickup cross section can be written as $\sigma_{\text{pu}}^{\text{g}} = \pi \cdot (R_N + r_m)^2$, where R_N and r_m are the radius of the cluster and the molecule, respectively. However, the actual pickup cross section can be significantly larger than the geometrical one, since the molecules can be attracted to the cluster due to the long-range interactions between the cluster and the molecules even from distances outside of the area corresponding to $\sigma_{\text{pu}}^{\text{g}}$. On the other hand, the pickup cross section includes also the sticking probability of the molecule to the cluster, which can be smaller than one. In addition, the collision and sticking probabilities depend on the collision energy. Thus the pickup is a complex process and some of these points were investigated in the references reviewed in this section and will be discussed below.

One of the early implementations of the pickup technique was to determine the size of the large neutral clusters.¹¹² The measurement of the cluster size in a molecular beam is difficult.



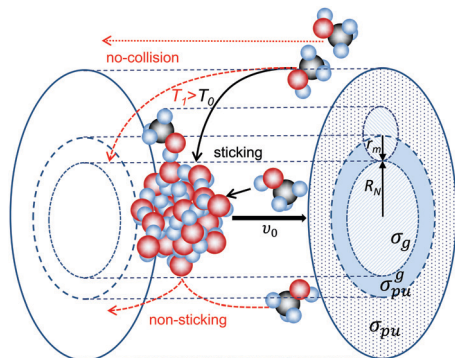


Fig. 3 Schematic picture of the pickup processes illustrating the cluster geometrical cross section $\sigma_g = \pi R_N^2$, geometrical pickup cross section $\sigma_{pu}^g = \pi(R_N + r_m)^2$, and pickup cross section σ_{pu} . The pickup cross section depends also on the probability of the sticking collision, and the collision probability is a function of the collision energy T .

Mass spectrometry involves ionization, which can cause severe cluster fragmentation.^{43,76} Even the threshold photoionization, which can be relatively soft, can cause fragmentation in some cases.^{113,114} For large clusters, other methods were used, *e.g.* high energy electron diffraction,^{115–119} helium atom diffraction,¹²⁰ and cluster beam scattering by a buffer gas.¹²¹ Some special methods were developed, such as elastic scattering^{122–124} and sodium doping with subsequent photodetachment of electrons,^{101,102} the latter of which was later proposed as the sizer for neutral weakly bound ultrafine aerosol particles.^{103,104} Clusters can be also deflected according to their sizes in strong inhomogeneous electric fields due to their size-dependent dipole moments and polarizability,^{125–127} which is, however, limited to small complexes. None of the above methods is universal. Elastic scattering is fairly elaborate and applicable only for small clusters, and Na-doping worked nicely with water and ammonia,¹⁰¹ and some other clusters,^{103,104,113} yet, it turned out not to be applicable for clusters containing molecules reacting with the solvated electron.^{72,77,79,128}

Alternatively, the cluster size can be determined by the methods based on the pickup. Historically, there have been essentially two approaches: in the first one, sticking collisions of molecules with clusters lead to momentum transfer, and thus by measuring the variations of the average cluster beam velocity with the number of collisions (the pickup gas pressure) the mean cluster size could be determined.¹¹² The second approach relies on measuring the Poisson distributions of the adsorbed molecules.¹²⁹ We have compared and combined these two methods for the pickup of different molecules on Ar_N clusters, and established a method to determine the pickup cross sections.⁶⁷

The original methods^{112,129} required certain knowledge about the pickup cross section σ_{pu} in order to evaluate the cluster mean size \bar{N} . The simplest assumption was to put the pickup cross section of a large cluster equal to its geometrical pickup cross section σ_{pu}^g , which could be simply determined from the hard sphere geometrical cross section σ_g assuming a spherical cluster shape and close packing of the molecules in

the cluster. This assumption reduced the unknown parameters to only the cluster size \bar{N} , and thus \bar{N} could be determined from the experiments. However σ_{pu} could be significantly different from σ_{pu}^g , as discussed below. Already in their original paper, Cuvelier *et al.*¹¹² used collision dynamics simulations to account for a more realistic attractive interaction between the molecule and the Ar_N cluster yielding a correction to σ_{pu}^g . Another theoretical study of the role of long-range forces in the cluster–molecule collisions for Ar_N resulted also in the capture cross sections larger than the geometrical ones.¹³⁰ In our experiments,⁶⁷ we worked with the Ar_N clusters, for which the mean size \bar{N} was known from the expansion conditions based on semiempirical Hagena's scaling laws,^{131–133} confirmed experimentally by helium atom diffraction.¹²⁰ The independent determination of \bar{N} for Ar_N clusters allowed the evaluation of the pickup cross sections σ_{pu} from our velocity measurements, and we have confirmed our measured cross sections with molecular dynamics (MD) and Monte-Carlo simulations.⁶⁷

The water $(H_2O)_N$ clusters represent another example, for which the mean cluster size \bar{N} was determined by independent experiments.¹⁰¹ Thus, we could measure the pickup cross sections for uptake of different molecules on large water clusters.⁶⁸ The pickup cross sections were measured for Ar_N and $(H_2O)_N$ in the size range of $\bar{N} \approx 50$ –600 and compared to MD simulations as well as to an analytical model.⁶⁹ The pickup cross sections for both Ar_N and $(H_2O)_N$ were larger than the corresponding geometrical pickup cross sections; however, there was also a difference in the cross section size dependence. While the pickup cross sections for Ar_N were consistent with their assumed spherical shapes, the cross sections of $(H_2O)_N$ departed from the spherical cluster model for $\bar{N} \geq 300$ towards larger values. The larger cross sections could be justified by assuming irregular shapes. These clusters occur in supersonic expansions under the conditions where large clusters are generated by smaller cluster coagulation rather than by the addition of the individual molecules. Thus the pickup experiments delivered not only the cluster pickup cross sections σ_{pu} but also indirect information about the cluster shape.⁶⁹

The velocity measurements for cluster size determination were implemented also by Kresin *et al.*¹³⁴ They used the momentum transfer in a single sticking collision, which was proved by the mass spectrometry of isotopically labeled species, to determine the neutral $(H_2O)_N$ and $(D_2O)_N$ cluster sizes of smaller clusters ($N \leq 50$). These results confirmed the extensive water cluster fragmentation after electron ionization observed previously,⁷⁶ emphasizing the conclusion that the simple mass spectrometry might be misleading for cluster size determination.

3.2 Mimicking atmospheric aerosol nucleation

3.2.1 Ice nanoparticles. The pickup cross sections reviewed above are important parameters in the modeling of aerosol formation and growth in the atmosphere. We have measured the pickup cross sections of ice nanoparticles $(H_2O)_N$ of mean



size $\bar{N} \approx 260$ for various atmospherically relevant molecules (H_2O , HCl , HBr , NO , NO_2 , CH_3OH , $\text{CH}_3\text{CH}_2\text{OH}$, CH_4).⁶⁸ The pickup cross sections for the uptake of H_2O molecules on $(\text{H}_2\text{O})_N$ are useful for modeling of the uptake of water molecules on aqueous particles, although the homogeneous nucleation of water is unlikely to occur in the Earth's atmosphere. The spherical cluster radius corresponds to $R_N = r_0 \cdot N^{1/3}$, where $r_0 \approx 2 \text{ \AA}$ could be obtained from the theoretical calculations.¹³⁵ The corresponding geometrical pickup cross section was $\sigma_{\text{pu}}^{\text{g}} = 685 \text{ \AA}^2$; however, the experimentally measured cross section was significantly larger, $\sigma_{\text{pu}} = 1020 \pm 100 \text{ \AA}^2$. This was due to the attractive interaction between the water molecule and the cluster as shown by our MD simulations using realistic interaction potentials for water, which yielded a cross section of 950 \AA^2 .⁶⁸

The pickup cross section depends also on the molecular mass and collision velocity. Our measured pickup cross section for the uptake of H_2O molecules on ice nanoparticles was larger than the simple geometrical cross section by at least a factor of 1.5. However, it can be further enhanced under atmospheric conditions by lower collision speed ($\sim 500 \text{ ms}^{-1}$) compared to our cluster beam experiment ($v_0 = 1450 \text{ ms}^{-1}$). This could enhance the pickup cross section by another factor of ≈ 1.4 .⁶⁸ The fact that the real pickup cross sections can be larger than the geometrical ones by a factor of more than 2 can have a pronounced effect in atmospheric aerosol modeling, where the geometrical cross sections are often considered. Finally, it should be noted that the evaluation of the pickup cross section required the assumption of sticking collisions and the total momentum transfer. These assumptions were justified for H_2O molecules by our MD simulations; however, similar simulations would be needed for the other molecules picked up by the ice nanoparticles.

In the context of the aerosol nucleation, it is interesting to compare our pickup cross sections with the recent work of Signorelli's group.¹³⁶ Their Laval-nozzle experiments allowed determination of the monomer association rates $k_{1,N}$ for H_2O molecules with water clusters of different sizes N . In their work,¹³⁶ $k_{1,N}$ for the clusters above the dimer up to $N = 30$ were compared to the free molecular collision rates obtained assuming the geometrical cross section of spherical clusters. The measured association rates were 4–5 times larger than the collision rates. This was attributed to the long range forces, e.g., dipole–dipole interactions and dispersion forces. Although we cannot provide a direct quantitative comparison between the association rates and our pickup cross sections, there is a clear correlation between them, and the larger measured association rates qualitatively agree with our pickup cross sections being larger than the geometrical ones.

From another point of view, the uptake of water molecules on mass selected protonated water clusters was measured by Zamith *et al.*^{137–140} The mass-spectrometric experiments with the protonated clusters provided the advantage of the known single cluster size, and the sticking collisions could be determined from the mass spectra. On the other hand, they were limited to smaller clusters and somewhat higher collision

energies. For their upper size limit of $N = 250$, the measured pickup cross section was $\approx 800 \text{ \AA}^2$. Comparison with our neutral $(\text{H}_2\text{O})_N$ pickup cross section suggests that the charge on the clusters does not seem to increase the pickup cross section, which calls for theoretical simulations.

Despite an extensive effort in the atmospheric aerosol research, there are still many open questions concerning the new particle formation and aerosol growth. Nucleation theories rely on macroscopic properties and gas kinetic arguments.^{141–143} Classical nucleation theory (CNT), which represents the current benchmark,¹⁴³ can provide reasonable water vapor nucleation rates in some cases;^{143–146} nevertheless, empirical corrections and modifications along with refined theories are implemented in some cases to account for the differences between the experimental results and CNT predictions.^{136,147,148} From the perspective of the molecular-level approach, the nucleation starts with the individual molecules colliding and sticking to small clusters. Typically, hard sphere collision rates and geometrical cross sections are considered in the modeling.^{46,149,150} The experiments mentioned above demonstrate that the actual pickup cross sections can be significantly larger, and more realistic values should be utilized.

3.2.2 Mixed nitric acid–water clusters. Aerosol nucleation in the atmosphere often involves the hydrated acid clusters, especially the sulfuric acid,^{151–154} and also nitric acid in the stratosphere.¹⁵⁵ A recent extension of our method for pickup cross section measurements to acid clusters could provide relevant data for the atmospheric aerosol nucleation models. We used the hydrated nitric acid clusters $(\text{HNO}_3)_M(\text{H}_2\text{O})_N$ as the model for a general hydrated acid cluster.^{72,156,157} The hydrated sulfuric acid clusters were investigated as well;⁵⁷ nevertheless, they led to frequent nozzle clogging, and the pickup experiments have not been accomplished yet.

In the experiments with the large $(\text{H}_2\text{O})_N$ clusters above, many collisions in the pickup cell led to multiple pickup and momentum transfer processes, and the dependence of the cluster velocity on the pickup pressure was measured. Here, the experimental method was modified. Relatively small $(\text{HNO}_3)_M(\text{H}_2\text{O})_N$ clusters with an average size of $\bar{M} \approx 2$ and $\bar{N} \approx 6$ passed through the pickup gas at a fixed pressure, which was carefully controlled so that the clusters of mean size underwent only a few (about 3) collisions in the pickup cell. Under these conditions, we could distinguish two mass peak series in the mass spectra measured after the pickup of a molecule X: one corresponding to the pure protonated cluster fragments $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$, and the other one corresponding to the fragments with the adsorbed molecules $\text{X} \cdot (\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$. Both series had almost identical character and differed only by their intensities. This is illustrated for our test case of methanol, $\text{X} = \text{CH}_3\text{OH}$, in Fig. 4. The top panel of Fig. 4a shows the spectrum of the pure $(\text{HNO}_3)_M(\text{H}_2\text{O})_N$ clusters analyzed elsewhere.^{51,72} The middle panel of Fig. 4b shows the spectrum after the pickup of methanol. Clearly, new weaker series are observed in Fig. 4b as indicated, which are analogous to $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ fragment ions and shifted by the mass of



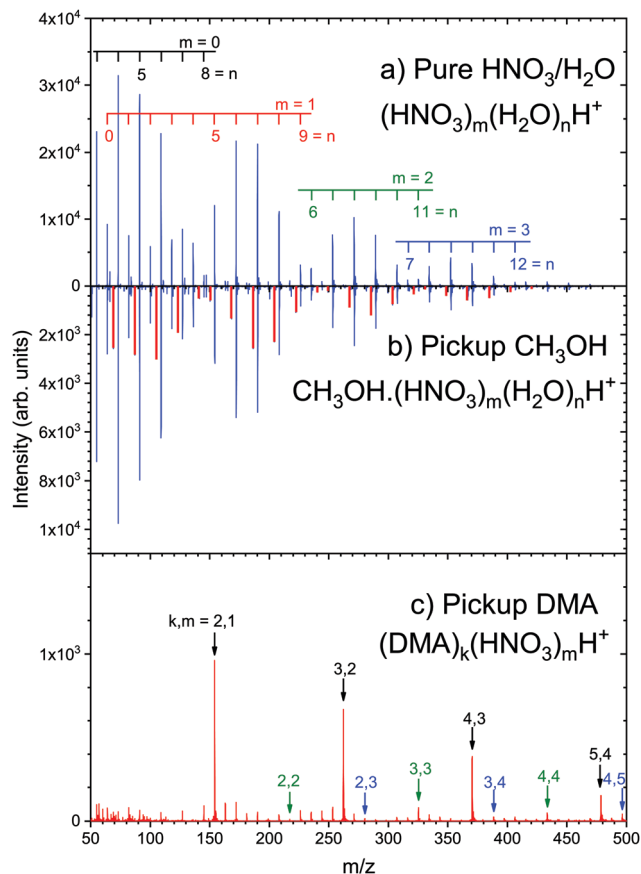


Fig. 4 Mass spectra showing the pickup of methanol (CH_3OH) and dimethylamine ($(\text{CH}_3)_2\text{NH}$ (DMA)) on the mixed nitric acid–water clusters: (a) pure $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ clusters without the pickup; $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ series with different m , n are labeled.^{51,72} (b) The pickup of methanol at a pickup pressure of about 4×10^{-4} mbar results in the additional analogous $\text{CH}_3\text{OH}\cdot(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ series highlighted in red.¹⁵⁸ The spectrum is plotted upside down for better comparison with the spectrum (a). (c) The spectrum after the DMA pickup at a pickup pressure of about 2×10^{-3} mbar. The black arrows indicate the major series $\text{DMA}\cdot(\text{DMA}\cdot\text{HNO}_3)_m\text{H}^+$. Further, less pronounced series $(\text{DMA}\cdot\text{HNO}_3)_m\text{H}^+$ and $(\text{DMA}\cdot\text{HNO}_3)_m\text{H}^-$ are indicated by green and blue arrows, respectively.

methanol $m/z = 32$. By velocity measurements performed with the QMS tuned to different fragment ions, we could prove that $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ fragments corresponded to the clusters which passed through the pickup cell without undergoing sticking collisions with CH_3OH . The velocities measured for the $\text{CH}_3\text{OH}\cdot(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ fragments were significantly lower than those of $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ fragments. Clearly, the former fragments originated from the clusters which picked up the CH_3OH molecule and were slowed down by the momentum transfer. The later velocities of $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ fragments were shifted only very little to lower values with respect to the velocities measured without any methanol gas in the pickup cell. This small shift was attributed to non-sticking collisions of a grazing character. Therefore, we could assign the $\text{CH}_3\text{OH}\cdot(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ and $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ fragments to the clusters which underwent sticking and non-sticking collisions, respectively, and the pickup probability could be determined

from the abundance ratio of the corresponding fragment series. The method was demonstrated for methanol pickup on $(\text{HNO}_3)_m(\text{H}_2\text{O})_n\text{H}^+$ clusters,¹⁵⁸ and several other molecules X discussed below showed qualitatively the same behavior.¹⁵⁹ These experiments represented the proof of concept for measuring the pickup probabilities, which can be useful for atmospheric aerosol modeling.

In atmospheric chemistry, bulk surface properties such as uptake coefficient γ and surface (or mass) accommodation coefficient α_s are generally used in the modeling of heterogeneous processes.^{160–163} These properties are usually determined in the experiments with macroscopic bulk surfaces. In such experiments, adsorption and desorption of the molecules from the surface, diffusion of the molecules from the surface into the bulk and their solubility in the bulk, and also chemical reactions have to be taken into account. In addition, the involved rates often change with time as the gas concentration gradient near the surface, as well as the concentration gradient in the bulk, changes. On the other hand, in our experiments, the clusters move through the pickup gas with extremely diluted concentration under high vacuum (10^{-4} mbar) undergoing just a few collisions within the pickup chamber. Thus, we can neglect any diffusion in the gas due to the gas concentration gradient near the cluster surface. No chemical reactions take place for the molecules discussed below on $(\text{HNO}_3)_m(\text{H}_2\text{O})_n$ clusters, and the adsorbed molecules do not dissolve nor diffuse into the cluster interior due to the small size of the clusters. We are able to determine the fraction of the clusters to which the molecules stick upon collisions; therefore our pickup probability is purely a kinetic parameter disentangling the molecule pickup from its evaporation. Under these conditions the pickup probability corresponds to the surface accommodation coefficient α_s . Thus our experiments provide a unique and direct way to determine experimentally α_s for the investigated molecules on the proxies of ultrafine aerosol particles.

Now, with the method established,¹⁵⁸ we can investigate the uptake of other atmospheric molecules. The volatile organic compounds (VOCs) belong to the key components in the aerosol generation.^{164–168} Therefore we have investigated the pickup of different VOCs and their oxidation products on $(\text{HNO}_3)_m(\text{H}_2\text{O})_n$ clusters.¹⁵⁹ The pickup of isoprene, α -pinene, 2-methyl-3-buten-2-ol, 3-methyl-3-buten-1-ol and verbenone was studied. The experiments delivered the α_s coefficients, which can be used for atmospheric aerosol modeling, and demonstrated that the oxidation increases the surface accommodation of VOCs by more than an order of magnitude. Accompanying theoretical calculations justified the experimental findings by forming of hydrogen bonds between the oxidized compounds and the clusters, whereas the interactions of the parent VOCs with the clusters were weaker and nonspecific.¹⁵⁹

Bases, e.g., ammonia and amines, contribute strongly to the aerosol nucleation as well.^{46,142,169} Therefore we picked up dimethylamine, $(\text{CH}_3)_2\text{NH}$ (DMA), on the hydrated acid clusters. However, a completely different mass spectrum was encountered (Fig. 4c). It is strongly dominated by a single



series, which could be assigned to the $\text{DMA} \cdot (\text{DMA} \cdot \text{HNO}_3)_m \text{H}^+$ ions. This suggests the acid–base reaction between HNO_3 and DMA leading to proton transfer and generating the ion pair $\text{NO}_3^- \cdots \text{DMA} \cdot \text{H}^+$. The energy released in the acid–base reactions leads to water evaporation from the clusters, since the anhydrous $\text{DMA} \cdot (\text{DMA} \cdot \text{HNO}_3)_m \text{H}^+$ ions dominate the mass spectrum. The water evaporation is in line with recent investigation of nanoparticle formation and growth from dimethylamine and nitric acid,¹⁷⁰ where relatively low binding energies of water molecules in these neutral clusters were calculated. There are also weaker mass peaks corresponding to the $(\text{DMA} \cdot \text{HNO}_3)_m \text{H}^+$ and $(\text{DMA} \cdot \text{HNO}_3)_m \text{HNO}_3 \text{H}^+$ series in Fig. 4c indicated by the green and blue arrows, respectively. These series are also consistent with the assumption of the acid–base reaction and $(\text{DMAH})^+ \cdots \text{NO}_3^-$ ion pair generation.

It should be noted that the above assignment is not unambiguous due to the mass coincidences between various $(\text{DMA})_k \cdot (\text{HNO}_3)_m \cdot (\text{H}_2\text{O})_n \text{H}^+$ fragments with different number of DMA, HNO_3 and H_2O molecules. In principle, this could be resolved by high-resolution measurements. We have demonstrated previously⁷² that the resolution of $M/\Delta M \approx 4 \times 10^3$ could be achieved with our RTOF allowing a clear separation of protonated water heptamer $(\text{H}_2\text{O})_7 \text{H}^+$ and nitric acid dimer $(\text{HNO}_3)_2 \text{H}^+$ peaks in the mass spectra. However, the separation of, *e.g.*, protonated water hexamer $(\text{H}_2\text{O})_6 \text{H}^+$ from nitric acid–DMA complex $\text{HNO}_3 \cdot \text{DMA} \cdot \text{H}^+$ would require even much higher mass resolution. Nevertheless, the present assignment is the most straightforward one, based on the acid–base reaction. There are still many open questions, for example: Does the acid–base reaction and subsequent water evaporation happen in the neutral cluster, or only after the ionization process? The presence of mainly anhydrous clusters in the spectrum supports the former scenario, since complete water evaporation is not observed upon ionization of the pure $(\text{HNO}_3)_M \cdot (\text{H}_2\text{O})_N$ clusters (Fig. 4a). However, such questions can be addressed and clarified by theoretical calculations and further experiments which are currently performed.

Similar experiments might be useful for the interpretation of the mass spectra measured in aerosol chambers, such as the Cosmics Leaving Outdoor Droplets (CLOUD) experiment.⁴⁷ A significant increase of the aerosol nucleation was observed in the presence of ammonia^{169,171} and DMA¹⁷² in the CLOUD experiments. In those experiments, all aerosol nucleation precursors are added into the system (acid, base, water, ionizing radiation, *etc.*), and the mass spectra are recorded varying the conditions. In the CLUB experiments, we start with relatively well-defined hydrated acid clusters, on which the amine molecules are picked up. Thus, we can investigate the processes step by step providing complementary information to the CLOUD experiments, where the final results of more complex processes are monitored. For example, we can clearly see the water evaporation upon DMA and HNO_3 reactions in our cluster experiments.

3.3 Clusters for interstellar chemistry

Despite the seemingly hostile conditions for chemistry in space, more than 200 molecules have been identified in the

interstellar medium (ISM), among them some quite large and complex ones such as amino acids and polycyclic aromatic hydrocarbons (PAHs).¹⁷³ How are these molecules formed in the ISM? Low densities of molecules, which in the densest regions of the giant molecular clouds hardly exceed 10^6 cm^{-3} (some 13 orders of magnitude less than the air in the Earth's atmosphere), imply infrequent collisions. Low temperatures, typically $\sim 10\text{--}100 \text{ K}$, often do not allow to surmount reaction barriers. However, there are other chemical pathways prevailing in the ISM, *e.g.*, surface-catalyzed reactions or radical and ion–molecule reactions, which are essentially barrierless.^{39–41,174,175} The ISM contains dust/ice grains acting as “sponges”, where the molecules can be adsorbed and stored for a long time, and eventually interact and react spontaneously or the reactions can be triggered by radiation. To simulate the dust/ice catalyzed chemistry, usually bulk surfaces are employed in the laboratory studies.^{176–178}

The clusters and nanoparticles in molecular beams can provide an alternative approach. They can mimic the dust/ice grains and offer some advantages for understanding the surface catalyzed chemistry. For example, the reaction intermediates and even the final products are sometimes difficult to detect in the bulk system.¹⁷⁷ In the case of clusters, they often leave the finite-size cluster and can be detected and characterized by mass spectrometry or optical spectroscopy. Thus the clusters can offer a detailed insight into the individual steps of complex reactions on dust/ice grains.

The ice nanoparticles investigated in our experiments can represent a model system for the water-ice covered grains in the ISM. Quite common in the ISM are various carbonaceous dust particles.¹⁷⁹ Also PAHs and their clusters have been suggested to be implemented in the ISM molecule synthesis.^{180,181} The PAH clusters have already been investigated experimentally by different methods, *e.g.*, IR and UV spectroscopy¹⁸² mass spectrometry,^{183,184} collisions with energetic ions,^{185–187} photoelectron spectroscopy,^{188,189} and in He-nanodroplets.¹⁹⁰ We have started to investigate the pickup of molecules by the carbonaceous and PAH clusters.

The mixed clusters of PAHs with other molecules were produced by other than pickup techniques in numerous studies. For example benzene–water^{191,192} and naphthalene–water^{193,194} cluster cations were generated by solvation of PAH cations in collision with water. The naphthalene–water clusters were also produced by co-expansion of naphthalene with water vapor in argon carrier gas, and subsequently ionized by vacuum ultraviolet photoionization.¹⁹⁵ In connection with our experiments on adamantane (AD) discussed below, we should mention also the spectroscopy of the ions formed by electron ionization of adamantane,¹⁹⁶ and the adamantane–water cluster cation.¹⁹⁷

Here, we present one of our first investigations, the pickup of methanol, which can be quite abundant in the ISM, on adamantane clusters $(\text{AD})_N$ as a prerequisite for the future investigations of chemical reactions of molecules adsorbed on such carbonaceous clusters. Fig. 5 shows the corresponding mass spectrum. The dominating cluster ion series $(\text{AD})_n^+$ is



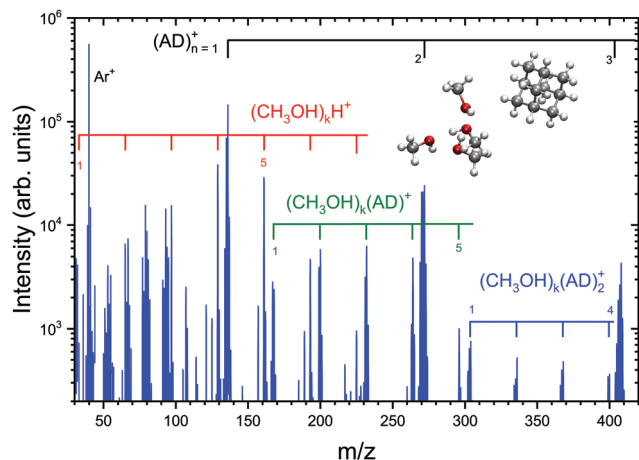


Fig. 5 Mass spectrum of adamantane (AD) $_N$ clusters after the methanol pickup showing the $(\text{AD})_n^+$ cluster fragment series together with the $(\text{CH}_3\text{OH})_k\text{H}^+$, $(\text{CH}_3\text{OH})_k(\text{AD})^+$ and $(\text{CH}_3\text{OH})_k(\text{AD})_2^+$ series. The inset shows a structure of the AD complex with 4 methanol molecules for illustration.

indicated together with clear methanol containing series $(\text{CH}_3\text{OH})_k\text{H}^+$ and $(\text{CH}_3\text{OH})_k(\text{AD})_n^+$, $n = 1, 2$. Each mass peak is accompanied by satellites with different number of H atoms; nevertheless, the major observations can be summarized: adamantane clusters are generated, and several methanol molecules can be adsorbed on $(\text{AD})_N$ in the pickup cell and coagulate to $(\text{CH}_3\text{OH})_K$ clusters with $K \geq 7$. An interesting observation is that the protonated $(\text{CH}_3\text{OH})_k\text{H}^+$ fragments are ejected from the clusters after the ionization. The protonated cluster fragments are typical for the hydrogen bonded clusters such as methanol or water in the gas phase. On the other hand, if the methanol clusters stick to one or two AD molecules after the ionization, the major peaks correspond to $(\text{CH}_3\text{OH})_k(\text{AD})_n^+$ ions, *i.e.*, non-protonated clusters. Theoretical investigation can help to investigate the structure of these clusters.

3.4 Coagulation of molecules on clusters

So far, we have concentrated on the uptake process of the molecule by the cluster. The next logical question is: What happens with the molecules adsorbed on the cluster? Do they migrate on the cluster surface or remain bound to a spot? Do they stay on the surface or submerge into the cluster interior? These questions were addressed previously in some cases.²² For example, the surface mobility of barium atoms and various molecules and generation of their clusters within large argon clusters was observed.^{24,198} In another spectroscopic experiment, methanol clusters were coagulated from individually adsorbed molecules on argon clusters.⁹⁴ In helium nanodroplets, complexes of picked up molecules were generated.^{16–18,21,83,88,89,91,199} However, rare gas clusters, and especially helium, represent rather special cases, and here we concentrate on the molecular clusters such as the ice nanoparticles.

We addressed the questions of migration and coagulation of molecules on/in clusters²⁰⁰ picking up different atmospherically relevant molecules X (X = HCl, CH_4 , CH_3Cl , $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$,

chlorobenzene, and benzene) on Ar_N and $(\text{H}_2\text{O})_N$ clusters with the mean sizes $\bar{N} \approx 330$ and 430, respectively, corresponding to approximately the same geometrical cross sections. The clusters underwent multiple pickup collisions with the molecules. In the case of Ar_N , the electron ionization mass spectra exhibited X_k^+ ion fragments up to $k \approx 10$ for most molecules. For methane, which did not stick to Ar_N easily, and is not prone to clustering, CH_4 complexes up to the trimer were observed.

On the other hand, there was no evidence for cluster formation on $(\text{H}_2\text{O})_N$ for the above molecules. This was not caused by the lack of adsorption of the molecules on $(\text{H}_2\text{O})_N$. The nanoparticles were slowed down by the momentum transfer, and individual molecules X could be seen in the mass spectra arriving with the ice nanoparticles into the mass spectrometer. In some cases, also the photodissociation experiments (discussed in Section 4.1) proved the presence of the molecules on $(\text{H}_2\text{O})_N$.⁶² Thus we could conclude that the above molecules coagulated to X_k clusters on Ar_N , while they remained isolated on $(\text{H}_2\text{O})_N$ during the flight time of about 1 ms from the pickup cell to the RTOF.

Accompanying theoretical MD simulations justified our findings: the molecules coagulated on Ar_N while they mostly remained bound where they landed on ice nanoparticles. The simulations also suggested that the molecules partly submerged in the cluster surface on Ar_N , despite the low temperature of the cluster of about 30–40 K and thus its solid-like nature.¹¹⁵ In the case of the ice nanoparticles the molecules remained bound to the surface. In another experiment, we adsorbed Freon CF_2Cl_2 molecules on Ar_N and ice nanoparticles,⁶³ and we observed the same pattern: coagulation to $(\text{CF}_2\text{Cl}_2)_K$ clusters on Ar_N and pickup but no coagulation on $(\text{H}_2\text{O})_N$.

Yet another interesting example of the difference between the coagulation of molecules picked up on Ar_N and $(\text{H}_2\text{O})_N$ can be found in our recent study of the hydrogen peroxide (H_2O_2) pickup and reactions on these clusters.⁶⁶ The top panel of Fig. 6a shows the electron ionization mass spectrum of the

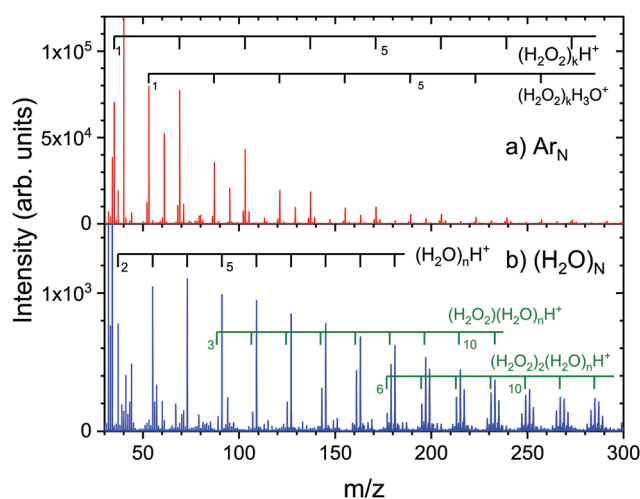


Fig. 6 Mass spectra of H_2O_2 molecules picked up on (a) Ar_N , $\bar{N} \approx 160$, and (b) $(\text{H}_2\text{O})_N$, $\bar{N} \approx 120$. The major fragment ion series are indicated and discussed in the text.



clusters generated after the pickup on Ar_N , $\bar{N} \approx 160$. The spectrum exhibits several fragment ion series containing H_2O_2 clusters. We label the major series corresponding to the protonated hydrogen peroxide $(\text{H}_2\text{O}_2)_k\text{H}^+$ (further series and their discussion can be found in our recent publication⁶⁶). Fragments with up to $k = 20$ H_2O_2 molecules could be identified in the spectrum upon detailed analysis.

Fig. 6b shows the spectrum after the pickup of H_2O_2 on $(\text{H}_2\text{O})_N$, $\bar{N} \approx 120$. Here, the spectrum is dominated by the protonated water $(\text{H}_2\text{O})_n\text{H}^+$; nevertheless, it contains also the series with H_2O_2 molecules, $(\text{H}_2\text{O}_2)_k(\text{H}_2\text{O})_n\text{H}^+$, with k up to 8 ($k = 1$ and 2 are labeled in Fig. 6b). Clearly, multiple H_2O_2 molecules could be adsorbed on the ice nanoparticles; however, there are no anhydrous cluster ion fragments $(\text{H}_2\text{O}_2)_k\text{H}^+$. For the $(\text{H}_2\text{O}_2)_k(\text{H}_2\text{O})_n\text{H}^+$ series, we find that each k number of H_2O_2 molecules is accompanied by $n > k$ water molecules, and the maximum of each series corresponds to the fragment composition with many more water molecules than H_2O_2 molecules, $n \gg k$. This suggests that the H_2O_2 molecules do not coagulate to $(\text{H}_2\text{O}_2)_K$ clusters on $(\text{H}_2\text{O})_N$, just that the electron ionization generates the protonated water fragments and some H_2O_2 molecules remain attached. However, this hypothesis needs further experimental and theoretical verification.

Apparently, all the picked up molecules, discussed so far, coagulated on Ar_N , but were immobilized on $(\text{H}_2\text{O})_N$. Is this a general pattern? It turns out that we had evidence for molecules coagulating on the ice nanoparticles as well. The first example was represented by 2-hydroxypyridine (HP).⁷⁴ These molecules generated large clusters on Ar_N , but we found the evidence for generation of $(\text{HP})_K$ clusters with K up to 8 also on $(\text{H}_2\text{O})_N$. It is worth noting that these molecules showed a strong propensity for dimer formation on $(\text{H}_2\text{O})_N$. Accompanying calculations showed that the dimers exhibited hydrogen bonds analogous to the base pairing in DNA. Thus this result can contribute to

the discussion of the biomolecule synthesis on the ice grains in the space.^{40,174,201–203} Similarly, we observed coagulation of uracil and 5-bromouracil on $(\text{H}_2\text{O})_N$.²⁰⁴ Thus the observed propensity for aggregation on ice nanoparticles seems to be a more general trend for molecules forming strong hydrogen bonds among themselves, *e.g.*, biomolecular analogues. Recently, we have observed also pyruvic and valeric acid coagulating on $(\text{H}_2\text{O})_N$.^{205,206} Table 1 provides the overview of molecules picked up on different clusters within our investigations, and indicates whether the molecules coagulated to clusters or not by labels *c* or *x*, respectively. Theoretical investigations could elucidate the interplay of forces acting between the molecules on nanoparticles, which lead to their mobility and coagulation.

4 Reactions triggered by photons and electrons

So far, we have focused on the pickup of molecules by the clusters. Now, we will discuss reactions between the adsorbed molecules. The reactions between neutral ground-state molecules were observed in some special cases previously, *e.g.*, the reactions of Ba atoms with different molecules mentioned in the introduction.^{23–29} We provide some evidence for the acid–base reaction between HNO_3 and DMA in Section 3.2.2. However, in this section, we concentrate on the reactions triggered by photons or electrons. In Table 1 we indicate for all the picked up molecules whether we have observed reactions after the electron ionization i^+ , or photon excitation-ionization $h\nu$, or slow electron attachment e^- .

4.1 Photodissociation of molecules on clusters

The photodissociation of a molecule is sometimes referred to as the half-reaction, and in that sense it can be regarded as one

Table 1 Overview of different molecules picked up on different clusters and processes observed in this review. Investigated processes: *c* – coagulation of molecules on/in the cluster, *x* – non-coagulation, σ_{pu} – pickup cross section, α_S – surface accommodation coefficient, i^+ – reactions after electron ionization, $h\nu$ – reactions after UV photon excitation/ionization, e^- – electron attachment and negative ion reactions

Molecules/cluster	Ar_N	$(\text{H}_2\text{O})_N$	$(\text{HNO}_3)_M(\text{H}_2\text{O})_N$	$(\text{NH}_3)_N$	$(\text{AD})_N$	Ref.
Ar	σ_{pu}	—	—	—	—	67
H_2O	σ_{pu}	σ_{pu}	—	—	—	67 and 68
HCl, HBr	$\sigma_{\text{pu}}, c, h\nu$	$\sigma_{\text{pu}}, x, h\nu$	—	—	—	62, 67, 68 and 207
NO, NO_2	—	σ_{pu}	—	—	—	68
Methanol (CH_3OH)	σ_{pu}, c	σ_{pu}	$\sigma_{\text{pu}}, \alpha_S$	c, i^+	c, i^+	67, 68, 158, 159, 208 and 209
Methane (CH_4)	<i>c</i>	σ_{pu}, x	—	—	—	68 and 200
Ethanol ($\text{CH}_3\text{CH}_2\text{OH}$)	—	σ_{pu}	—	—	c, i^+	68
$\text{CH}_3\text{Cl}, \text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}, \text{C}_6\text{H}_5\text{Cl}, \text{C}_6\text{H}_6, \text{CF}_2\text{Cl}_2$	<i>c</i>	<i>x</i>	—	—	—	63 and 200
Isoprene, α -pinene, verbenone, 2-methyl-3-buten-2-ol, 3-methyl-3-buten-1-ol	—	—	α_S	—	—	159
Hydrogen peroxide (H_2O_2)	c, i^+, e^-	x, i^+, e^-	—	—	—	66
Formic acid (HCOOH)	$c, i^+, h\nu$	—	—	—	—	209
Amylene (2-methyl-2-butene) ($\text{H}_3\text{C}-\text{C}(\text{CH}_3)=\text{CHCH}_3$)	$c, i^+, h\nu$	—	—	—	—	75
Pyruvic acid ($\text{H}_3\text{C}-\text{C}(\text{O})-\text{C}(\text{O})-\text{OH}$)	—	$c, i^+, h\nu$	—	—	—	205
Valeric acid ($\text{H}_3\text{C}-\text{CH}_2\text{CH}_2-\text{C}(\text{O})-\text{OH}$)	$c, i^+, h\nu$	$c, i^+, h\nu$	—	—	—	206
2-Hydroxypyridine	$c, i^+, h\nu$	$c, i^+, h\nu$	—	—	—	74
Uracil, 5-Br-uracil	—	c, i^+	—	—	—	204
Dimethylamine ($(\text{CH}_3)_2\text{NH}$)	—	—	c, i^+	—	—	—
Methylamine (CH_3NH)	—	—	—	—	c, i^+	—



of the simplest chemical reactions.^{210,211} The effect of a solvent on the photodissociation is of prime importance for understanding the chemistry in bulk systems irradiated by UV photons.^{212,213} With this motivation the photodissociation of molecules in rare gas clusters was investigated.^{22,36,37,214}

Although the rare gas clusters represented a benchmark system emulating a weakly interacting environment, let us concentrate on the atmospherically more relevant systems such as the large water clusters $(\text{H}_2\text{O})_N$. They can mimic the ice particles in the polar stratospheric clouds (PSCs). Type II PSCs are composed essentially of pure water ice.¹⁵⁵ The PSC particles represent the key players in the chemistry leading to the Antarctic ozone hole.^{155,215–218} The heterogeneous chemistry on their surfaces converts so called reservoir species (HCl, ClONO₂) to the active ones (Cl₂, HOCl), which can be readily photolyzed by the sun radiation yielding the ozone destroying Cl• and other radicals. The reservoir species originate from the photolysis of the Cl-containing molecules in the stratosphere such as the infamous Freon CF₂Cl₂, and subsequent Cl• reactions with other atmospheric molecules, e.g., reactions with CH₄ produce HCl.

We have investigated the photodissociation of hydrogen halides HX (X = Cl, Br) on $(\text{H}_2\text{O})_N$ implementing also Ar_N for comparison.⁶² The molecules adsorbed on the Ar_N clusters formed (HX)_K clusters, as revealed above, and the H fragments underwent caging after the photodissociation yielding the fragments with essentially zero kinetic energy. The heavy Cl• fragments were also caged; however, some of them escaped the cluster without losing their kinetic energy. On $(\text{H}_2\text{O})_N$ clusters, the HX molecules did not coagulate and remained isolated, and their (photo)chemistry was quite different. First, the acidic dissociation occurred on $(\text{H}_2\text{O})_N$ already in the ground state generating the H₃O⁺ ··· X⁻ ion pair. This ion pair was then excited by the UV photon yielding a neutral hydronium radical H₃O•.^{219–223} H₃O• was metastable and quickly decayed yielding the observed H• fragments with low kinetic energy. This reaction pathway was observed in previous photodissociation studies,^{207,224,225} and also confirmed theoretically.^{219,220,226} The corresponding Cl• fragment was trapped in the ice nanoparticles. Possible relevance of such photodissociation of HCl on the PSC particles was discussed.^{62,224} The acidic dissociation enhances the HCl photolysis rate in the 200–300 nm region by about four orders of magnitude.²²⁴ On the other hand, this enhancement of the Cl• radical yield from HCl on the ice could be counterbalanced by the trapping of the Cl• in the particles.⁶²

The photodissociation of Freon CF₂Cl₂ was investigated in pure $(\text{CF}_2\text{Cl}_2)_N$ clusters and also when adsorbed on Ar_N and $(\text{H}_2\text{O})_N$ clusters,⁶³ as well as a single CF₂Cl₂ molecule embedded inside Ar and Xe clusters.²²⁷ The general feature of the photodissociation process was the caging of Cl• observed in all kinds of clusters. However, the Cl• fragments from Ar_N clusters were detected, which meant that the fragments were slowed down to near-zero kinetic energies, but they still exit the cluster or the cluster decayed upon the photodissociation process. On the other hand, no Cl• fragments from $(\text{H}_2\text{O})_N$

were detected, despite the clear evidence for the uptake of CF₂Cl₂ on $(\text{H}_2\text{O})_N$. The loss of Cl• can be atmospherically important. The Cl• fragments trapped in the ice nanoparticle after the photodissociation can be washed out from the atmosphere. The accompanying theoretical calculations rationalized this observation by the formation of halogen bonds, where the Cl atoms of CF₂Cl₂ pointed to the oxygen of the water molecules in the nanoparticle. MD simulations of the dissociation process showed that the Cl• entered the cluster after the photodissociation, and remained trapped.⁶³

4.2 Reactions after excitation and ionization

The clusters represent a very special environment for reactions of molecules in the excited or ionized states, and the reaction pathways can differ from the reactions in both the gas and the condensed phases. In clusters, the reactants are in direct contact already before they are excited or ionized. Thus the collision process, preceding a reaction in the gas phase, is absent in the cluster, and the reaction can readily proceed at a kinetic rate. The interaction potentials can be altered by the cluster environment, and reaction pathways different from the gas phase can be followed. The cluster can lead to the product caging as discussed above for the photodissociation processes. The cluster can serve as an efficient heat bath with many closely spaced degrees of freedom, where excited states can be efficiently quenched. On the other hand, in comparison with the condensed phase, the clusters provide a much simpler environment allowing molecular-level understanding of the investigated reactions. The energy released in the reaction leads to the cluster decay and the reaction products, which are usually trapped in the bulk, can be released from the cluster and detected. Thus, reaction intermediates can be observed, such as activated complexes and radicals, which usually react immediately in the bulk. All these aspects are well known and have been discussed in numerous studies.^{3,4,22,43,53,55,104,113,228}

We have investigated quite a few systems where the molecules were picked up on different clusters and reactions were subsequently triggered by interaction with electrons (typically of 70 eV) and/or photons (usually at 193 nm, 6.4 eV). For example, the reactions of amylene (2-methyl-2-butene, H₃C-C(CH₃)=CH-CH₃) molecules picked up on Ar_N have been observed recently.⁷⁵ The mass spectra after the electron ionization and photoionization exhibited striking differences pointing to the excited state reactions, which occurred only during the photoionization process. From the perspective of atmospheric chemistry, small carboxylic acids play an important role in the secondary organic aerosol formation.^{46,229–231} We have investigated valeric acid (H₃C-(CH₂)₃-(C=O)-OH) and pyruvic acid (H₃C-(C=O)-(C=O)-OH) in clusters and picked up on $(\text{H}_2\text{O})_N$ and Ar_N clusters.^{205,206} From the perspective of the interstellar chemistry, ammonia and methanol together with water represent the primary hydrogen-bonding molecules detected in the interstellar clouds, ices and comets in high abundance.^{232–236} We have studied the uptake and ionization of methanol on large ammonia (NH₃)_N, $\bar{N} \approx 230$, clusters.²⁰⁸ However, we are not discussing these experiments in more



detail and the interested reader is referred to the above references.

Our ultimate goal was to observe a reaction between two different molecules adsorbed on a cluster. This has recently been achieved for methanol (CH_3OH) and formic acid (HCOOH) adsorbed on Ar_N nanoparticles.²⁰⁹ The Ar_N nanoparticles picked up the molecules in two sequential pickup cells in SC and PC chambers (see Fig. 2). We probed the clusters by photon and electron ionization yielding very similar spectra in this case. When only one kind of molecules was adsorbed on Ar_N , the spectra revealed protonated clusters of either $(\text{CH}_3\text{OH})_m\text{H}^+$ or $(\text{HCOOH})_m\text{H}^+$. When both molecules were adsorbed on the nanoparticles, mixed clusters appeared as well. We used isotopically labeled molecules to prove the reactions between CH_3OH and HCOOH . Perdeuterated methanol (CD_3OD) demonstrated the deuteron transfer reactions to formic acid resulting in $(\text{HCOOH})_m\text{D}^+$ cluster ions, as well as the proton transfer from formic acid resulting in $(\text{CD}_3\text{OD})_m\text{H}^+$. The latter process was expected for the acidic molecule, while the former reaction was, perhaps, less intuitive. The experiments with the partially deuterated methanol CD_3OH and CH_3OD further revealed the proton transfer from both methyl and hydroxyl groups. These findings were supported by *ab initio* calculations.²⁰⁹ Both molecules CH_3OH and HCOOH are abundant in the ISM, and the methoxy radical $\text{CH}_3\text{O}^\bullet$ and the hydromethoxy radical $\text{CH}_2\text{OH}^\bullet$ were implemented in the formation of complex organic molecules in astrochemistry.^{237–240} Our study shows that such radicals might be generated in the proton transfer processes on the grains and released from them. More realistic proxies of interstellar grains will be used in future experiments, *e.g.*, the carbonaceous clusters discussed in Section 3.3.

4.3 Electron attachment

Another area of reactions on free clusters is the electron attachment and negative ion reactions. The dissociative electron attachment (DEA) of isolated molecules is important in many areas of physics and chemistry, as reviewed recently by Fabrikant *et al.*²⁴¹ Indeed, the question, how does a more complex environment of clusters influence the DEA process, has been ever-present already in the early studies.^{242,243} A recent review covered several fundamental effects of the environment on the DEA process in clusters, which can lead to an enhancement as well as to a suppression of the electron attachment cross section.²⁴⁴

In relevance to the atmospheric aerosols discussed above, we have concentrated on the electron attachment to the mixed acid–water complexes, namely $\text{HNO}_3/\text{H}_2\text{O}$ ^{156,157} and $\text{H}_2\text{SO}_4/\text{H}_2\text{O}$.⁵⁷ The former system was a representative example of the role of the environment in the DEA process, where the solvent molecules influence the reaction pathways and change the products. The DEA of an isolated HNO_3 molecule yielded almost exclusively NO_2^- ions (96.5%) and the minor product OH^- (3.4%). The nitrate anion NO_3^- had an abundance of less than 0.1%. This picture changed completely in the mixed $(\text{HNO}_3)_M(\text{H}_2\text{O})_N$ clusters, where the negatively charged cluster

ion fragments with NO_3^- dominated the mass spectra with 57% abundance. It was interpreted by the acidic dissociation of HNO_3 molecules in the water containing clusters, and by cascade intracluster ion–molecule reactions terminated by NO_3^- . The cluster ions containing HNO_3^- were also detected in significant amounts (8%). This ion was not formed in the gas phase DEA, and its generation in the clusters was attributed to the stabilization of HNO_3^- by caging. The measured energy dependence of the DEA¹⁵⁷ revealed that the degree to which the major gas phase product NO_2^- is converted to NO_3^- in the clusters depends strongly on the electron energy: namely NO_2^- prevails at low electron energies below approximately 3.5 eV, while NO_3^- is the major product at higher energies.

The atmospheric relevance of these experiments is underlined by the fact that NO_2^- and NO_3^- are among the most abundant anions in the atmosphere having a strong influence on the aerosol generation and other processes.^{245–247} In addition, the proposed negative ion reactions in $(\text{HNO}_3)_M(\text{H}_2\text{O})_N$ clusters yield the OH^\bullet radical and HONO as the neutral products, which can be evaporated from the particles to the atmosphere. They represent the key species in atmospheric chemistry, namely the hydroxyl radical is the major oxidation agent.⁴⁵ The photolysis of nitrous acid serves as the source of OH^\bullet in urban areas; yet the detailed mechanism of its formation is still discussed.

The experiments in M. Beyer's group, in which the negatively charged water clusters $(\text{H}_2\text{O})_n^-$ are generated and subsequently react with molecules,^{56,248–252} can represent a complementary approach to our investigations of the electron attachment in water clusters. The $(\text{H}_2\text{O})_n^-$ clusters represent the solvated electron e_{aq}^- in the gas phase.^{56,252} This way, the reaction of CF_2Cl_2 with the solvated electron was investigated.²⁵⁰ The reaction yielded the $(\text{H}_2\text{O})_n\text{Cl}^-$ product, and the reaction enthalpy was determined. These results demonstrated that CF_2Cl_2 can undergo the dissociative electron transfer in condensed aqueous environments if thermalized hydrated electrons are present. It can be relevant to the atmospheric chemistry of ozone depletion, since the earlier studies^{253,254} revealed huge enhancement of the bond cleavage in CF_2Cl_2 and other chloro-fluoro-carbon compounds due to the dissociative electron transfer on ice surfaces. Later on, this process was proposed as the mechanism enhancing the Cl^\bullet radical yield in PSCs;²⁵⁵ however, the actual contribution of this process to ozone depletion remains questionable.^{256–260} The study revealed the details and energetics of the dissociative electron transfer process at the molecular level, but its relevance for the stratospheric ozone depletion could not be established by laboratory experiments.²⁵⁰

Hydrogen peroxide, H_2O_2 , plays an important role in many areas of atmospheric and interstellar chemistry as a source of HO_x radicals^{261,262} and oxidation agent contributing to the aerosol nucleation.^{45,46,151,153} We have investigated the electron attachment to hydrogen peroxide (H_2O_2) adsorbed on Ar_N and ice $(\text{H}_2\text{O})_N$ nanoparticles.⁶⁶ The hydrogen peroxide coagulates to $(\text{H}_2\text{O}_2)_M$ clusters on Ar_N , as illustrated by the mass spectra in Fig. 6. The electron attachment to the $(\text{H}_2\text{O}_2)_M$ clusters on Ar_N



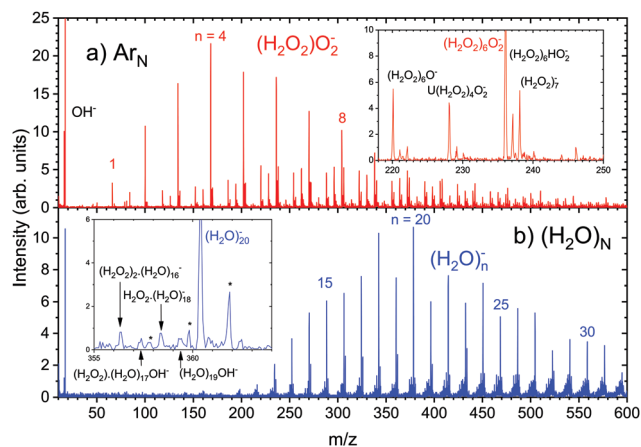


Fig. 7 Negative ion mass spectra of H_2O_2 molecules picked up on (a) Ar_N , $N \approx 160$, and (b) $(\text{H}_2\text{O})_N$, $N \approx 120$. The spectra recorded at the electron energies between 0 and 5 eV were integrated. The spectrum (a) is dominated by the $(\text{H}_2\text{O}_2)_m\text{O}_2^-$ ion series as indicated, and further ion series are shown in the inset. All ions have been discussed in ref. 66. The spectrum (b) is dominated by the water $(\text{H}_2\text{O})_n^-$ series, and further minor ions containing H_2O_2 are shown in the inset (the star indicates metastable species).

yielded $(\text{H}_2\text{O}_2)_m\text{O}_2^-$ ions as the major products, as shown in Fig. 7a. There are further weaker O_2^- containing series, and also HO_2^- and other ions (see the inset of Fig. 7a and ref. 66 for discussion). The presence of O_2^- ions was surprising, since they could not be generated in the DEA of the isolated H_2O_2 molecule. Nandi *et al.*²⁶³ showed that the DEA of H_2O_2 yielded only OH^- (75%) and O^- (25%). We have proposed possible reaction pathways leading from the nascent OH^- and O^- ions to the observed O_2^- , HO_2^- and the other ions in the spectra in ref. 66.

On the other hand, O_2^- is a well known product of the reaction of solvated electron with H_2O_2 in aqueous solutions in radiation chemistry.^{264–266} Therefore, we have investigated also H_2O_2 adsorbed on the $(\text{H}_2\text{O})_N$ clusters, where the electron attachment yields the solvated electron $(\text{H}_2\text{O})_n^-$. The uptake of H_2O_2 molecules on $(\text{H}_2\text{O})_N$ was proved by the positive spectrum in Fig. 6b. Nevertheless, the negative ion spectrum is dominated by $(\text{H}_2\text{O})_n^-$ and no O_2^- containing products were detected, as shown in Fig. 7b. The $(\text{H}_2\text{O})_n^-$ series was also observed in the electron attachment to the pure water clusters.²⁶⁷ There are other negative ion series of minor abundance, pointing to the presence of H_2O_2 , illustrated in the inset of Fig. 7b. It is worth noting that a relatively strong OH^- peak at $m/z = 17$ is observed in both spectra in Fig. 6. The DEA of H_2O_2 yields OH^- , which leaves the clusters. Although the analysis of the spectrum in Fig. 7b might be still tentative, it is clear that none of the strong peaks could be assigned to O_2^- containing ions. Thus, the reaction of e_{aq}^- with H_2O_2 yielding O_2^- was not observed in the case of $(\text{H}_2\text{O})_N$ clusters. The cluster environment changed the chemistry in a way, which could not be predicted based on the free DEA in the gas phase, nor based on the known chemistry of H_2O_2 in the bulk aqueous solutions and the reactions with the solvated electrons. It was even

different from the chemistry of the electron attachment to the H_2O_2 clusters on Ar_N .

5 Conclusions, challenges and perspectives of the pickup experiments

We have reviewed a few examples of the experiments with the pickup and reactions of molecules on clusters from the perspective of atmospheric and interstellar chemistry. The uptake cross sections of the ice nanoparticles for various molecules were determined. The cross section for the pickup of water molecules turned out to be substantially larger than the geometrical cross section, which is important for the nucleation of water aerosols.

The pickup experiments with the smaller hydrated nitric acid clusters enabled determination of purely kinetic surface accommodation coefficient for the proxies of ultrafine aerosol particles. We have established a promising method providing important parameters for atmospheric aerosol modeling; nevertheless, the method is not universal. When the character of the mass spectra with and without pickup is different, the pickup probability cannot be evaluated as the ratio of the integrated spectra. Such was the case for the uptake of molecules on hydrated sulfuric acid clusters. The investigation of sulfuric acid clusters is thus challenging, yet highly desirable, and will be followed in the future.

Some questions were touched in our pickup studies; yet future work can answer them more comprehensively. For example, the coagulation of molecules on nanoparticles could be modeled in more detail to understand their mobility and interplay of the bonds. Also the question of molecules submerging into the cluster or remaining on its surface should be addressed in more detail.

A clear perspective to follow in the future was shown by the experiments with the pickup of dimethylamine on the hydrated nitric acid clusters, which pointed to the acid–base reaction in the neutral clusters. From a purely fundamental point of view, it can provide molecular-level insight into the basic chemistry. In the atmosphere, such systems play the key role in the new particle formation.

This will be one of our major directions in future studies. On our way towards the clusters of atmospheric relevance, we have moved from rare gas model systems to water clusters, and recently to mixed binary systems such as nitric acid–water clusters. Evidence from aerosol field measurements and model simulations suggests that the new particle formation in the atmosphere is most likely dominated by ternary nucleation of H_2SO_4 – H_2O – NH_3 and subsequent condensation of semivolatile organic compounds (SVOCs).²⁶⁸ Clusters mimicking the composition of the real atmospheric aerosols should be generated. The present experiments with HNO_3 – H_2O clusters and DMA pickup represent the first steps in this direction in our experiments. We will continue generating clusters of hydrated inorganic acids (nitric acid, sulfuric acid, methanesulfonic acid) with bases (ammonia, amines). The pickup of other



atmospheric molecules, such as SVOCs, by these clusters will be studied, providing uptake coefficients for atmospheric modeling. Another challenging project represents incorporating a radical source into the pickup region of our CLUB apparatus and embedding radicals, such as OH[•], into the clusters. This will open the doorway to radical chemistry investigations in these clusters. The reactions of radicals, positive and negative ions and excited species in the clusters can contribute to the understanding of the atmospheric aerosol chemistry.

In relevance to the polar stratospheric clouds and ozone depletion we have reviewed the photodissociation of hydrogen halides and CF₂Cl₂ on the ice nanoparticles. In the case of HX, our experiments revealed their acidic dissociation to the H₃O⁺ ··· X⁻ ion pair in the ground state, analogous to the above acid–base reaction of DMA. However, it is interesting to note that the acid–base reaction in the DMA case was revealed by mass spectrometry, while in the HX case a completely different photochemistry experiment in synergy with theory provided the evidence. Further extension of this work to the photodissociation of molecules on the mixed nitric acid–water clusters and to the clusters with sulfuric acid is planned also with other molecules involved in the ozone depletion, e.g., ClONO₂, CH₃Cl, Cl₂, HOCl.

We have discussed reactions between molecules deposited on the nanoparticles triggered by electrons or photons. The first reactions observed between two different molecules on Ar_N were the proton transfer processes between methanol and formic acid. Here, a vast field is opened for simulating atmospheric and interstellar chemistry, e.g., reactions of different molecules on carbonaceous and PAH clusters, and ice nanoparticles including also other than water nanoices.

Finally, reactions triggered by slow electron attachment in molecules deposited on clusters were revealed. Generally, the reactions of negatively charged ions have been investigated less extensively compared to the positive ion chemistry, and thus the negatively charged clusters represent a large relatively unexplored field.

In terms of the experiment, the following challenges may be defined: (1) new ways of cluster generation with improved control of cluster composition and size; (2) implementing the action spectroscopy to provide information about the bonding motifs and structure of the fragments; (3) detect and analyze the neutral fragments in addition to the charged ones.

Author contributions

MF: conceptualization, methodology, writing-original draft, review and editing, visualization, supervision, funding acquisition, project administration; JF: methodology, writing-review and editing; JK: investigation, writing-review and editing; JL: investigation, writing-review and editing; EP: software, writing-review and editing; VP: investigation, writing-review and editing; AP: investigation, formal analysis, visualization.

Conflicts of interest

There are no conflicts of interest to declare.

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