



Cite this: *Chem. Commun.*, 2021, 57, 4979

Correction: Tethered CAAC–CAAC dimers: oxidation to persistent radical cations and bridging-unit dependent reactivity/stability of the dications

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DOI: 10.1039/d1cc90161c

rsc.li/chemcomm

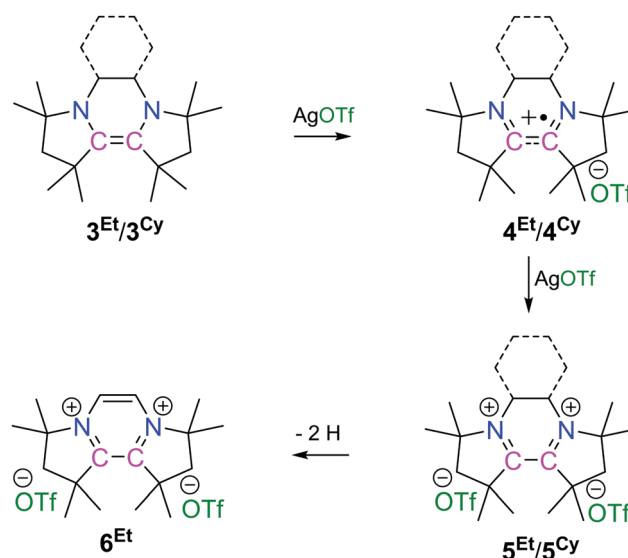
Correction for 'Tethered CAAC–CAAC dimers: oxidation to persistent radical cations and bridging-unit dependent reactivity/stability of the dications' by Mithilesh Kumar Nayak *et al.*, *Chem. Commun.*, 2021, 57, 1210–1213, DOI: 10.1039/D0CC07385G.

The authors regret that some data in the original article was mis-interpreted.

In case of 2-electron oxidation, the ethylene bridged CAAC–CAAC dimer leads to the dication **6^{Et}** with the concurrent elimination of two hydrogens. For the *trans*-1,2-cyclohexylene bridged CAAC–CAAC dimer, however, it leads to the dication **5^{Cy}** instead of **6^{Cy}** as reported in the original article. Consequently, sections of the text in the manuscript should be adjusted according to this change, and these are detailed below.

The CCDC number, 1971758, refers to compound **5^{Cy}** rather than compound **6^{Cy}** as originally reported.

Scheme 3 has been revised to reflect this change



Scheme 3 Oxidation of **3^{Et}** and **3^{Cy}**.

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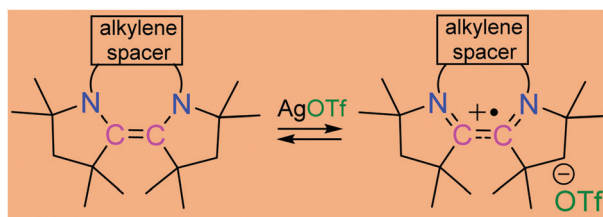
The caption of Fig. 3 should read “Fig. 3 Solid-state structures of 4^{Cy} (left) and 5^{Cy} (right). Hydrogen atoms are omitted for clarity.”

The paragraph on page 1212 beginning “The 1 : 2 reactions of 3^{Et} and 3^{Cy} with AgOTf did not lead to the isolation of the expected dications 5^{Et} and 5^{Cy} ...” should be correctly given as “The 1 : 2 reactions of 3^{Et} and 3^{Cy} with AgOTf led to the isolation of 6^{Et} (72%) and 5^{Cy} (80%), respectively (Scheme 3). The elimination of two hydrogen atoms for 3^{Et} is most likely driven forward by the gain of aromaticity. The theoretical calculation at the B3LYP/def2-TZVP level of theory for the ethylene bridged analogue as a representative example suggests that the reaction is exergonic by $\Delta G = -0.92 \text{ kcal mol}^{-1}$.¹⁸ Crystallographic analysis showed that the central C_4N_2 -six-membered ring is planar for 6^{Et} and non-planar for 5^{Cy} with a N1–C4–C3–N2 dihedral angle of 33.72° (Fig. S67, ESI,† and Fig. 3).”

The paragraph on page 1212 beginning “Surprisingly, the results obtained from electrochemistry and spectroelectrochemistry are in contrast to the chemical reduction of $6^{\text{Et}}/6^{\text{Cy}}$...” should be correctly given as “The results obtained from electrochemistry and spectroelectrochemistry are in line with the chemical reduction of 6^{Et} (ESI,†).”

In the sentence on page 1212 beginning “In contrast to 3^{Et} and 3^{Cy} , the oxidation of 3^{Pr} ...”, 3^{Cy} should be removed.

The corrected table of contents entry is shown below:



The electronic supplementary information and crystal structure data have been updated to reflect this correction.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

