



### Pd-catalyzed C(sp<sup>3</sup>)-C(sp) bond formation in iodocyclobutenes†

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**A synthesis of skipped 1,4-enynes through functionalization of the cyclobutene core with alkynes has been achieved, suggesting an unusual pathway of oxidative addition in tertiary iodoalkanes.**

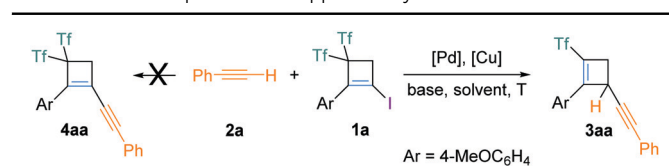
The preparation of alkyl alkynes through irradiation or metal-catalyzed reactions of alkyl halides with alkynes is attractive.<sup>1</sup> Lei and others have recently reported the Sonogashira-type alkynylation with alkanes,<sup>2</sup> while Park's group described the addition of an alkyne to an allenamide.<sup>3</sup> Nonetheless, there is absence of protocol for the direct incorporation of alkynyl groups into existing cyclobutene motifs through C(sp<sup>3</sup>)-C(sp) bond formation.

Having in hand a convenient method for the direct preparation of starting halocyclobutenes **1a-e** from iodoalkynes,<sup>4</sup> we intended to study the feasibility of their palladium-catalyzed reactions with terminal alkynes **2**. Bis(triflyl)iodocyclobutene **1a** was selected as model substrate and was reacted with phenylacetylene **2a** using a Pd-Cu bimetallic catalytic system. After extensive experimentation, the adoption of 5 mol% of Pd(PPh<sub>3</sub>)<sub>2</sub>Cl<sub>2</sub> as the palladium complex and 5 mol% of CuI as the copper salt gave the best performance (Table 1, entry 3). Serendipitously, 4-alkynyl-2-(triflyl)cyclobutene **3aa** was obtained in a reasonable 70% yield instead of the expected Sonogashira product, namely, 2-alkynyl-4,4-bis(triflyl)cyclobutene **4aa**. The screening of different palladium complexes such as Pd(PPh<sub>3</sub>)<sub>4</sub>

and copper salts produced inadequate conversions and less gratifying yields (Table 1, entries 1-5). The screening of solvents and bases revealed that the use of triethylamine as both base and solvent was superior to other combinations (Table 1, entry 8). Temperature studies disclosed that 35 °C was the optimal T (Table 1, entries 6 and 7). The cooperative presence of the palladium complex and the copper salt was required (Table 1, entries 9 and 10).

When we studied the halogen effect it was observed that the nature of the halide affected the efficiency of the reaction. (Triflyl)cyclobutene **3aa** was obtained at 80 °C in poor yield (34%) starting from bis(triflyl)bromocyclobutene **1a-Br**, while no reaction was observed from the chloride counterpart **1a-Cl** (Scheme 1A). In order to see the generality of the process in the presence of different alkynes, bis(triflyl)iodocyclobutene **1a** was reacted with a series of terminal alkynes **2a-g**. As shown in Scheme 1B, aryl-alkynes bearing on the aryl ring both electron-donating (MeO) as well as electron-withdrawing (CF<sub>3</sub>)

**Table 1** Alkylation of phenylacetylene **2a** with bis(triflyl)iodocyclobutene **1a** under modified palladium-copper-catalyzed reaction conditions



Entry	Catalyst <sup>a</sup>	Reaction conditions	Yield <sup>b</sup> (%)
1	Pd(PPh <sub>3</sub> ) <sub>4</sub> , CuI	Et <sub>3</sub> N, 35 °C, 3 h	55
2	Pd(AcO) <sub>2</sub> , CuI	Et <sub>3</sub> N, 35 °C, 3 h	43
3	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub> , CuI	Et <sub>3</sub> N, 35 °C, 1 h	70
4	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub> , CuBr	Et <sub>3</sub> N, 35 °C, 3 h	47
5	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub> , CuOTf	Et <sub>3</sub> N, 35 °C, 4 h	16
6	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub> , CuI	Et <sub>3</sub> N, 20 °C, 14 h	58
7	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub> , CuI	Et <sub>3</sub> N, 60 °C, 2 h	51
8	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub> , CuI	K <sub>2</sub> CO <sub>3</sub> , 35 °C, 3 h <sup>c</sup>	24
9	CuI	Et <sub>3</sub> N, 35 °C, 12 h	—
10	Pd(PPh <sub>3</sub> ) <sub>2</sub> Cl <sub>2</sub>	Et <sub>3</sub> N, 35 °C, 12 h	—

<sup>a</sup> The reactions were run using **1a** (0.1 mmol) and catalyst (5 mol%).  
<sup>b</sup> Yield of pure isolated product. <sup>c</sup> 1,4-Dioxane was used as solvent.

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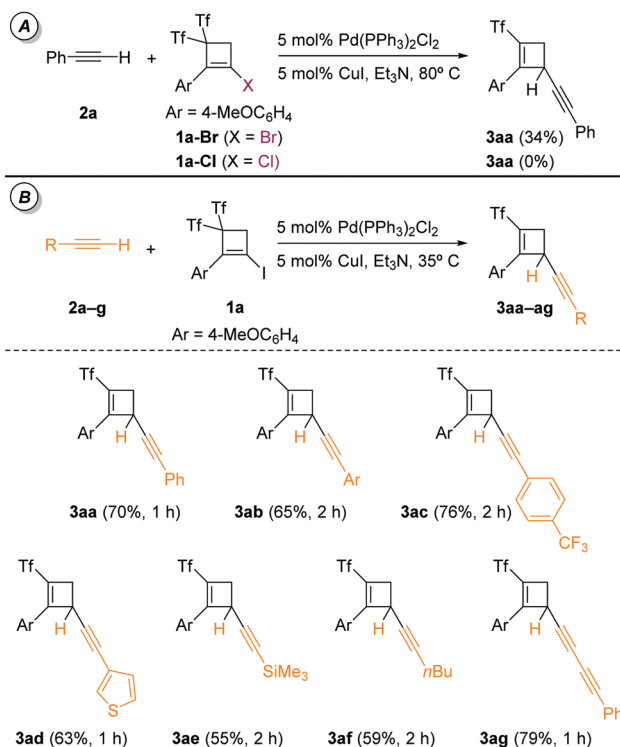
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† Electronic supplementary information (ESI) available: Computational details, experimental procedures, characterization data of new compounds, and copies of NMR spectra for all new compounds. See DOI: 10.1039/d1cc03087f



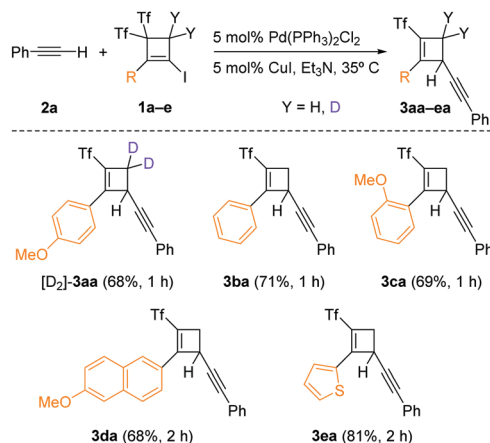


**Scheme 1** Palladium–copper-catalyzed alkylation of terminal alkynes **2a–g** with bis(triflyl)iodocyclobutene **1a**.

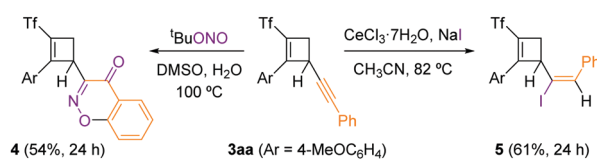
substituents were well accommodated. Heterocyclic (thienyl) and heteroatomic (trimethylsilyl) functionalities were well tolerated in the alkyne. Likewise, alkyl and alkynyl-substituted alkynes were also successful coupling partners. The resulting (triflyl)cyclobutenes **3aa–ag** were smoothly obtained in good isolated yields after flash chromatography in silica gel (Scheme 1B).

The scope of the C(sp<sup>3</sup>)–C(sp) bond formation in cyclobutenes was investigated with respect to the bis(triflyl)iodocyclobutene component. As depicted in Scheme 2, a range of substituents are amenable for this Pd-catalyzed conditions. Both, aryl (phenyl and naphthyl) and hetaryl (thienyl) moieties are tolerated. Besides, the bulky naphthyl ring provided the corresponding (triflyl)cyclobutene **3da** in good yield. Noteworthy, the presence of deuterium atoms in the cyclobutene framework is compatible with our protocol, which allows a convenient access to *D*-labelled adduct [D<sub>2</sub>]-**3aa**. The synthetic utility of alkynyl-cyclobutenes **3** was demonstrated by the preparation of chromanone-like compound **4** and alkenyl iodide **5** (Scheme 3).

The addition of TEMPO to the reaction between **1a** and phenylacetylene **2a** resulted in insignificant differences in the obtained yields of the final product **3aa**, which exclude a radical mechanism. At least, two possible mechanisms may be postulated for this Pd-catalyzed alkylation reaction of terminal alkynes (Scheme 4). A Sonogashira coupling followed by an allylic substitution with water acting as a hydrogen source<sup>5</sup> may be postulated in the first possible route (Scheme 4, path a).



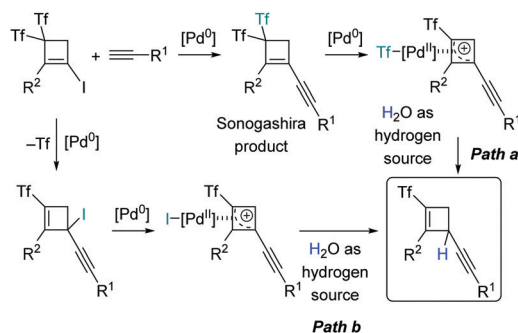
**Scheme 2** Palladium–copper-catalyzed alkylation of phenylacetylene **2a** with bis(triflyl)iodocyclobutenes [D<sub>2</sub>]-**1a** and **1b–e**.



**Scheme 3** Transformations of alkynyl–cyclobutene **3aa**.

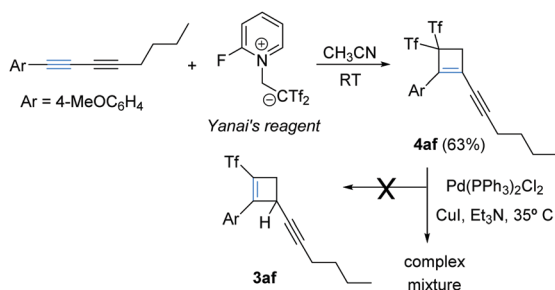
Alternative incorporation of H from Et<sub>3</sub>N by β-elimination and reductive elimination can be ruled out since the formation of **3aa** in the presence of *d*<sub>15</sub>-Et<sub>3</sub>N gave no deuteration. The second plausible pathway (Scheme 4, path b) may involve as initial step an allylic substitution with the participation of the alkyne as the nucleophile, and a subsequent oxidative addition of the alkyl iodide. To assess the above proposed pathways, we used the cyclization reaction of 1-methoxy-4-(octa-1,3-diyne-1-yl)benzene with 2-(2-fluoropyridin-1-ium-1-yl)-1,1-bis[(trifluoromethyl)sulfonyl]ethan-1-ide (Yanai's reagent)<sup>6</sup> to prepare the formal Sonogashira adduct **4af** which may be considered as an intermediate. Under the standard reaction conditions bis(triflyl)-1,3-enyne **4af** could not be converted into (triflyl)-1,4-enyne **3af** (Scheme 5), which should point to Path a as the less likely pathway.

Product **3aa** did not undergo H/D exchange in presence of D<sub>2</sub>O under the reaction conditions. Next, we decided to test the reaction



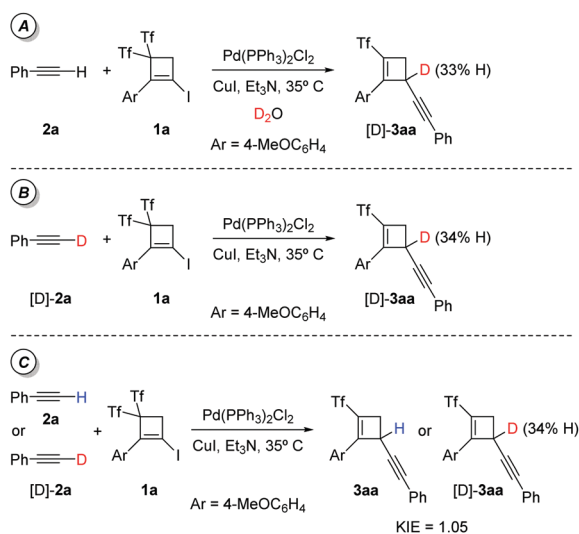
**Scheme 4** Possible pathways of the C(sp<sup>3</sup>)–C(sp) bond formation reactions.





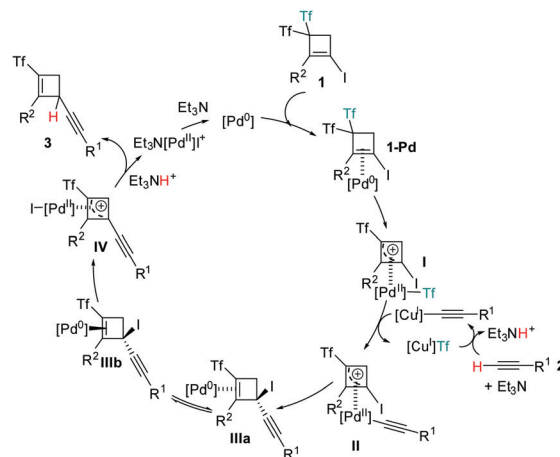
Scheme 5 Control experiment.

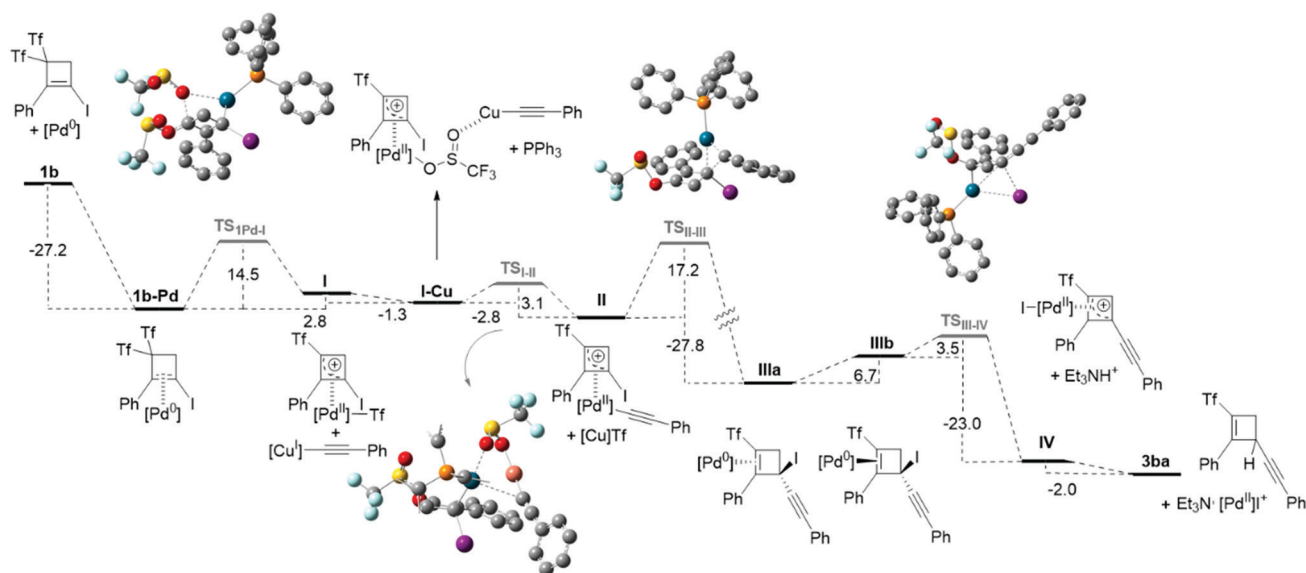
between **1a** and **2a** under the standard reaction conditions but with the inclusion of 2 equiv. of  $D_2O$  (Scheme 6A). Intriguingly, the above experiment did not form exclusively (triflyl)cyclobutene **3aa**. Instead, labelled (triflyl)cyclobutene  $[D]$ -**3aa** was formed as very major (67% deuterium content) adduct. Unexpectedly, the combination of labelled alkyne  $[D]$ -**2a** and bis(triflyl)iodocyclobutene **1a** using the optimized Pd–Cu bimetallic catalytic system was capable of promoting the formation of partially deuterated (triflyl)cyclobutene  $[D]$ -**3aa** (66% deuterium content) (Scheme 6B). This should be attributed to the fact that the new hydrogen atom at the skipped 1,4-enyne moiety may arise from the terminal alkyne, and also may point to a competence in the final protonation step when external water (*e.g.*  $D_2O$ , Scheme 6A) is added to the reaction. A kinetic isotopic effect (KIE) was not clearly detected in the standard reaction when deuterated phenylacetylene  $[D]$ -**2a** was used as alkyne in the standard reaction with **1a** in parallel experiments (Scheme 6C), which may suggest that the breakage of the C(sp)–H bond is not implicated in the rate-determining step. Consequently, the above control experiment should call into question Path b (Scheme 4) as a plausible pathway.

Scheme 6 Labelling experiments and determination of the KIE for the alkyne partner. Reactions were run using **1a** (0.1 mmol) and catalyst (5 mol%).

Based on the results of the above-mentioned control experiments, a reasonable mechanistic proposal for the generation of cyclobutene-embedded skipped 1,4-enynes is depicted in Scheme 7. In this catalytic cycle, activation of the substrate takes place after Pd(0) coordination to the cyclobutene, by C–O allylic oxidative addition involving one of the triflyl groups to give complex **I**.<sup>7</sup> Then, following the steps of a usual Sonogashira coupling, transmetalation with Cu-acetylide affords **II** which undergoes subsequent reductive elimination to furnish species **IIIa**. A facile equilibration by decoordination–coordination of Pd(0) to cyclobutene would render **IIIb** from **IIIa**. Complex **IIIb**, would evolve by C–I oxidative addition to form  $\pi$ -allyl palladium species **IV**. In turn, the early released  $Et_3NH^+$  (in the formation of the Cu acetylide) plays an essential role in promoting an electrophilic cleavage of the Pd–C bond of intermediate **IV**, to give functionalized products **3** and a Pd(II) salt.<sup>8</sup> Finally, the active palladium(0) catalyst is regenerated by reductive elimination of palladium(II) assisted by triethylamine, closing the catalytic cycle.

The feasibility of this proposal was explored by DFT calculations (Scheme 8 and ESI†) to corroborate whether the proposed steps were energetically preferred over other secondary pathways.<sup>9</sup> Bis(triflyl)iodocyclobutene **1b** and phenylacetylene **2a**, which react to give product **3ba**, were used as model substrates, and  $(PPh_3)_2Pd(0)$  was considered as the catalytically active species. The energy profile for the transformation (Scheme 8) shows that the overall process exhibits a downhill energetic trend and all the computed energies are compatible with a temperature of 36 °C. Although the first oxidative addition is an endergonic step ( $\Delta G = 2.8 \text{ kcal mol}^{-1}$ ), it occurs through  $TS_{I-Pd-I}$  with a relatively low activation energy of  $14.5 \text{ kcal mol}^{-1}$ . The coordination of the copper from the copper acetylide to one of the oxygen atoms of the triflyl group bonded to Pd decreases the energy by  $1.3 \text{ kcal mol}^{-1}$  and generates the association complex **I-Cu**, which facilitates a fast transmetalation step *via*  $TS_{I-II}$  ( $\Delta G = -2.8 \text{ kcal mol}^{-1}$ ,  $\Delta G^\ddagger = 3.1 \text{ kcal mol}^{-1}$ ). The consequent reductive elimination step is highly exergonic ( $\Delta G = -27.8 \text{ kcal mol}^{-1}$ ) and has a  $\Delta G^\ddagger = 17.2 \text{ kcal mol}^{-1}$ .

Scheme 7 Mechanistic proposal for the Pd-catalyzed formation of **3**.



**Scheme 8** Energy profile for the Pd-catalyzed formation of 4-alkynyl-2-(triflyl)cyclobutenes **3**. Values of reaction free energies and activation energies ( $\text{kcal mol}^{-1}$ ) are calculated M06/6-311+G\*\* (C,H,O,S,F,P,N), SDD (Pd,Cu,I), SMD(triethylamine)/ B3LYP/6-31G(d) (C,H,O,S,F,P,N), LANL2DZ (Pd,Cu,I) with [Pd] =  $(\text{PPh}_3)\text{Pd}$  and [Cu] =  $(\text{PPh}_3)\text{Cu}$ . Phenyl substituents in the  $\text{PPh}_3$  ligand are omitted in the ball-and-stick model of  $\text{TS}_{\text{I-II}}$  for clarity.

This step constitutes the driving force of the reaction, since it falls into a highly stable energy minimum **IIIa**. As commented before, species **IIIa** and **IIIb** are related through a low energy equilibrium, from which **IIIb**, would react *via* an oxidative addition though  $\text{TS}_{\text{III-IV}}$ . This step has a low activation energy ( $\Delta G^\ddagger = 3.5 \text{ kcal mol}^{-1}$ ) and is energetically very favored ( $\Delta G = -23.0 \text{ kcal mol}^{-1}$ ). It is worth noting that this oxidative addition takes place in a concerted manner, through a three-membered transition state, as it usually occurs for  $\text{C}(\text{sp}^2)\text{-I}$  activations of aryl- and alkenyl-iodides. Although alkyl halides have been proposed to react with Pd(0) from the backside, following a  $\text{S}_{\text{N}}2$  type mechanism,<sup>10</sup> this tertiary halide undergoes a facile oxidative addition by the metal atom bound to the same face of the cyclobutene. Interestingly, previous calculations showed that this activation mode has a lower activation energy compared to the  $\text{S}_{\text{N}}2$  type reaction.<sup>11</sup> The experimental results in the presence of radical inhibitors allow disregarding radical activation pathways. Finally,  $\text{Et}_3\text{NH}^+$  would act as a proton source to depalladate **IV** and furnish the reaction product **3ba**.<sup>5b</sup>

In conclusion, we have developed a metal-catalyzed direct alkylation reaction of terminal alkynes with bis(triflyl)iodocyclobutenes for the generation of (triflyl)cyclobutene-embedded skipped 1,4-enynes. A pathway which contrasts to the widespread activation of the  $\text{C}(\text{sp}^2)\text{-I}$  bond by the Pd catalyst has been suggested with the aid of control and labelled experiments as well as DFT computations.

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## Conflicts of interest

There are no conflicts of interest to declare.

## Notes and references

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- Perhaps, final protonation at the  $\alpha$  position of alkyne is competitive with protonation at the  $\alpha$  position of Tf group, but it could be possible that the isomerization of the double bond leads to a single product, favoured by the basic reaction conditions and the natural acidity of the protons in  $\alpha$  to sulfones.
- Other feasible reaction pathways such as the initial oxidative addition of the  $\text{C-I}$  bond instead of the  $\text{C-Tf}$  bond, have been explored and disregarded in view of the higher energies (see ESI† for details).
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