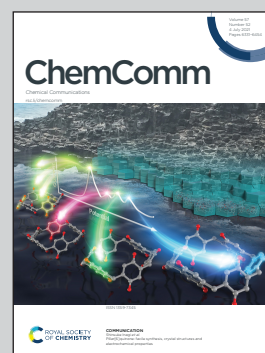


Showcasing research from the laboratories of Andrew Cassidy, using the ASTRID light source at the Department of Physics and Astronomy, Aarhus University, Denmark and Valeria Lauter, Neutron Scattering Division, Oak Ridge National Lab, Oak Ridge, USA.

A mechanism for ageing in a deeply supercooled molecular glass

The temporal decay of spontaneous polarization in a molecular film is coupled with spin-polarised neutron scattering measurements to identify the ageing pathways for a deeply supercooled glassy solid. (Image credit: Oak Ridge National Laboratory/Jill Hemman).

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Measurements of the decay of electric fields, formed spontaneously within vapour-deposited films of *cis*-methyl formate, provide the first direct assessment of the energy barrier to secondary relaxation in a molecular glass. At temperatures far below the glass transition temperature, the mechanism of relaxation is shown to be through hindered molecular rotation. Magnetically-polarised neutron scattering experiments exclude diffusion, which is demonstrated to take place only close to the glass transition temperature.

Ageing or relaxation in supercooled materials may be a two component process in polymeric films,^{1,2} metallic glasses^{3–6} liquid crystals⁷ and, as here, molecular glasses. When heated close to the glass transition temperature, T_g , the two processes, termed the primary and secondary relaxations, merge but in deeply supercooled materials they separate out with different rates and different temperature dependencies. Here, we use measurements of spontaneous dipole orientation in polarised, deeply supercooled molecular solids to follow both processes across a wide temperature range, from $0.5 T_g - T_g$. We observe that the energy barrier for the secondary process drops as the growth temperature of the material is decreased.

The growth of glassy molecular solids by condensation of gas is the main growth mechanism for molecular solids in the interstellar medium,⁸ is frequently used for the production of

molecular electronic components⁹ and, increasingly, for drug delivery solutions.¹⁰ It has recently been discovered that molecular ordering can occur in films formed through physical vapour deposition.^{11–14} The phenomenon appears to be a general consequence of the growth of thin films of molecular materials from the vapour phase. It has been established that a class of polar molecules can condense and harbour spontaneously-generated electric fields *via* molecular dipolar-orientation.^{11,12,15} Such fields are created with no external agency, such as an applied field, but are a natural consequence of deposition of a polar molecule. This so-called ‘spontelectric effect’ relies on molecular orientation and arises in the absence of any free charge.¹¹ Since electric fields arise from collective dipole orientation, migration of dipoles, away from this ordered state, results in depolarisation. Thus, any observed decay of the field tracks molecular rotation. Accordingly, molecular rotations are followed here through the temporal decay of the spontelectric field in films of *cis*-methyl formate, henceforth MF. This is an extremely sensitive means of detecting molecular rotation in solids. Changes of a few hundredths of a degree in average angle of dipole orientation with time are readily observable, allowing for investigations on short time-scales of a few hours.

Fig. 1 shows that the spontelectric phenomenon gives a new and highly sensitive handle on molecular rotation in polarised glasses, at deeply supercooled temperatures, where secondary relaxation is active.¹⁶ The strength of the electric field harboured by a film of MF was inferred from the surface potential measured at the film-vacuum interface atop a film of a known thickness.¹⁷ Throughout, we assume that MF adopts the *cis* conformation as this is energetically preferred, with a 546 meV barrier blocking transformation to *trans*.¹⁸ The correlation between measured electric field strength and deposition temperature, T_d , for MF is shown in Fig. 1. For $T_d = 80$ K to 89 K, Fig. 1A, field strength decreases with time, with a lag period of up to ~ 2000 s before this decay begins. In this temperature regime, the internal electric field also increases with increasing T_d and this has been explained elsewhere.¹⁷

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† Electronic supplementary information (ESI) available: Demonstrating neutron scattering from trilayer films and X-ray diffraction data from cryo-cooled samples of methyl formate. See DOI: 10.1039/d1cc01639c



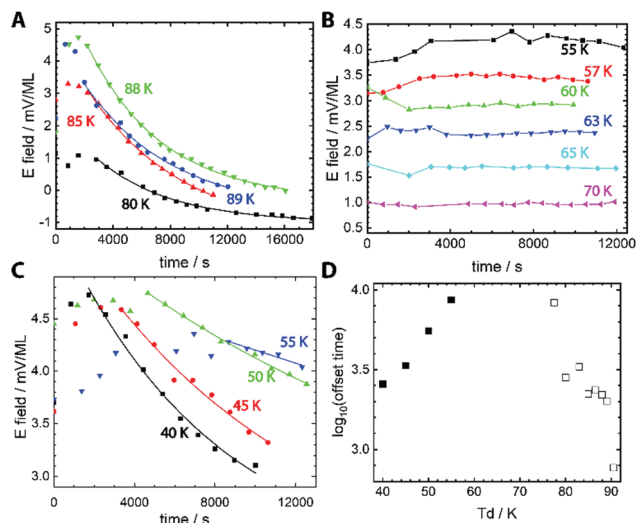


Fig. 1 Isothermal surface potential decay. (A–C) The decay in the strength of electric field measured in films of *cis*-methyl formate is plotted for films prepared across a range of temperatures of deposition (T_d). Measurements were performed during growth of the film, with each data point on a given curve representing a different film thickness, but data are plotted as a function of time. Data are divided into three temperature regimes: (A) high T_d where the electric field decays to a constant value irrespective of T_d ; lines are results from fits (see text); (B) mid-range T_d , where the electric field values are stable. Lines here are added as a guide to the eye and; (C) low T_d , where the electric field does not decay to zero where lines here result from fits. (D) The lag period before field decay begins, for fits presented in (A) and (C) are plotted as a function of T_d . Systematic errors in (E) field measurements arise from systematic errors in the film thickness, $\pm 20\%$.

Data are presented as a function of time, illustrating the rate of field decay. Data in Fig. 1A demonstrate exponential decay of the field to a constant value, c , and the solid lines represent fits resulting from $y = y_0 \exp(-t/\tau)^\beta + c$, where t is the elapsed time, y_0 is the initial electric field before decay begins and τ is the time constant for the decay process.¹⁹ Fits resulted in β values very close to 1, and we set $\beta = 1$ throughout. This value of β may be related to the short timescales over which aging is reported here.²⁰ The residual surface potential measured at extended time delays in Fig. 1A, represented by the constant c , are a measure of the work function of the MF film and are not related to the spontelectric effect.

In the mid-temperature regime, for T_d between 55 K and 70 K, Fig. 1B, the electric field values are constant at any given T_d and this constant value decreases as T_d increases. This is the behaviour expected of spontelectric materials.^{11,17} At low deposition temperatures, below 55 K in Fig. 1C and extending to the lowest temperature of 40 K, there is again a decay in the electric field as a function of time and, again, a lag period before electric field values start to decrease. The lag period decreases as T_d is decreased. In addition, the final value to which the field decays, variable “ c ” above, is lower for lower values of T_d , representing residual dipole alignment. The lag periods before decay, for $T_d \leq 55$ K and $T_d \geq 75$ K are collected in Fig. 1D. For 90 K $> T_d > 75$ K the lag times change by an order of magnitude and decrease with increasing T_d . Thus films

Table 1 Parameters describing the rate of field decay in films of MF

T_d (K)	Lag period (s)	Rate constant, k , for decay of $\langle \mu_z \rangle / \mu$ (s^{-1})	Activation energy, E_{lag}^* (eV)	Time constant, τ , for exp decay of $\langle \mu_z \rangle / \mu$ (s)
40	2564	3.900×10^{-4}	0.112	5622
45	3348	2.987×10^{-4}	0.127	11 185
50	5520	1.812×10^{-4}	0.143	17 230
55	8672	1.153×10^{-4}	0.160	102 894
77.5	8336	1.200×10^{-4}	0.225	$2.35 \times 10^{+5}$
80	2816	3.551×10^{-4}	0.225	5180
83	3298	3.032×10^{-4}	0.234	5420
85	2222	4.500×10^{-4}	0.237	6763
86.5	2350	4.255×10^{-4}	0.241	4485
88	2200	4.545×10^{-4}	0.245	5291
89	2010	4.975×10^{-4}	0.247	6196
90.5	770	1.299×10^{-3}	0.244	1031

age faster as the growth temperature is increased, as expected. Conversely, lag times for $T_d \leq 55$ K decrease as T_d drops lower, meaning films age faster as growth temperature drops.

If the lag period represents the time taken before the first molecules begin to rotate away from a dipole-oriented conformation, then the inverse lag time can be used to establish a rate for this rotation process. The degree of dipole orientation is defined as $\langle \mu_z \rangle / \mu$, where $\langle \mu_z \rangle$ indicates the component of the molecular dipole, μ , pointing in the direction of the electric field, that is, normal to the plane of the film. The process of the decay of $\langle \mu_z \rangle / \mu$ has been modelled as first order and the corresponding rate coefficients and activation energies are shown in columns 2 and 3 of Table 1.

The time constants used to produce the exponential fits in Fig. 1 represent the rate at which the decay of $\langle \mu_z \rangle / \mu$ propagates throughout the film. The interpretation of this and an understanding of the dynamics of molecular rotations, in the context of the spontelectric field permeating the films, will be considered elsewhere.

We now show that molecular diffusion does not occur below ~ 80 K and thus that aging between 40 K and 55 K is not caused by diffusion. Magnetically-contrasted neutron scattering was used to measure the self-diffusion of molecules of MF on molecular length scales as a function of sample temperature. Alternate dosing of D- and H-MF was used to prepare a pentalayered-sandwich of MF, in which the isotopic substitution does not affect the chemical, kinetic and thermodynamic properties of the MF film. Reflectivity plots, as a function of scattering vector magnitude Q , from two distinct, pentalayered-films, grown at $T_d = 45$ K and $T_d = 65$ K, and recorded repeatedly as both films were annealed to progressively higher temperatures, and data at $T_d = 45$ K are shown in Fig. 2. Full datasets and data for trilayer-films with better temperature resolution are included as ESI.†

The reflectivity profiles in Fig. 2A result from the depth-sensitive interaction of the polarised neutron beam with the various isotopically labelled interfaces created within the structure. Data were collected for approximately 1 hour at each temperature point, yielding the time dependence of physical changes. A model that allowed for loss of definition at



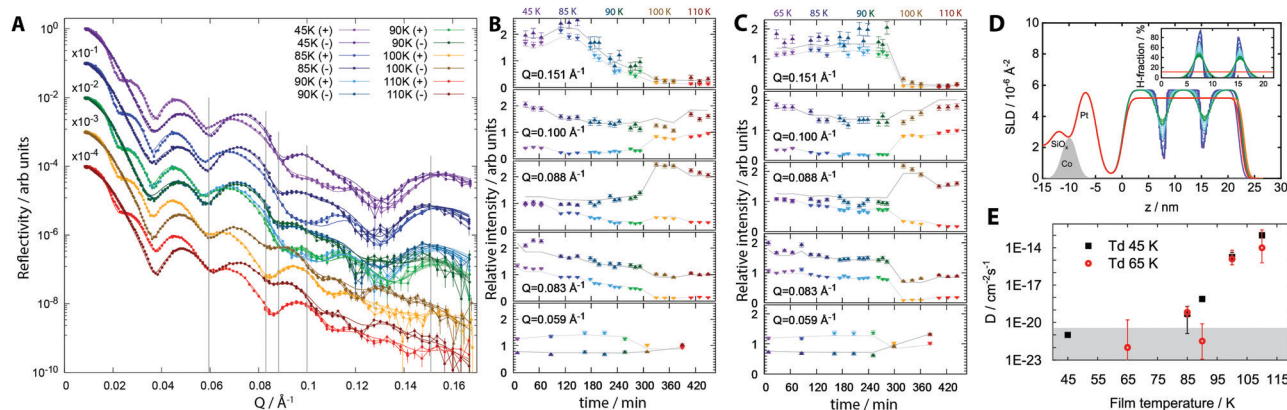


Fig. 2 Neutron scattering data from distinct penta-layered films of methyl formate. (A) Reflectivity plot, as a function of scattering vector magnitude Q , for a film prepared at $T_d = 45$ K and gradually annealed to higher temperatures. (B) Relative intensity for data in (A), $T_d = 45$ K; fits at specific Q values were extracted and plotted as a function of time, where 0 min indicates the beginning of the first measurement. There are several time values for each annealing temperature (see main text and ESI†) and colour has been used to distinguish between annealing points. Errors arise from signal to noise limits. Lines represent fits resulting from the self-diffusion model. (C) Relative intensity at specific Q values for $T_d = 65$ K. (D) Scattering length density profiles, $T_d = 45$ K; inset shows H-fraction. (E) Self-diffusion coefficients for $T_d = 45$ K and $T_d = 65$ K measured as films were annealed to progressively higher temperatures. The grey area indicates the lower limit of detection and diffusion coefficient values in this regime are too low to be modelled reliably.

isotopically labelled interfaces *via* diffusion, according to Fick's Law, was fit to the reflectivity profiles. The resulting fits in Fig. 2A allow us to produce; (i) time resolved data, Fig. 2B and C, where relative intensity values have been extracted at specific Q values as a function of both time and temperature for films with $T_d = 45$ K and 65 K, and (ii) a real space visualisation of the interface diffusion as scattering length density (SLD) profiles, Fig. 2D (see ESI†). High values of Q are most sensitive to the layered structure of a film and changes in these values indicate mixing between layers. Neither the $T_d = 45$ K nor the $T_d = 65$ K films show any mixing upon growth and very little mixing after annealing to 85 K. Above 85 K and the 120 min mark, relative intensities at $Q = 0.151 \text{ \AA}^{-1}$ in the $T_d = 45$ K film begin to change gradually. Both $T_d = 45$ K and $T_d = 65$ K films show a step change above 90 K in Q data. Above 90 K, the composite film is no longer identifiable as distinct D- and H-layers, for either the $T_d = 45$ K or the $T_d = 65$ K films; above 100 K, the MF isotopologues have completely mixed homogeneously throughout the entire film (see ESI†). This mixing results from primary structural reorganisation of the films, and demonstrates that both films have now passed the glass transition temperature. Methyl formate remains adsorbed on the substrate up to 135 K.²¹ Mixing for the $T_d = 45$ K film begins at a lower temperature, 90 K compared to 100 K for the $T_d = 65$ K film, indicating a lower T_g , and a less stable material when $T_d = 45$ K.

Self-diffusion coefficients, D , can be plotted as a function of anneal temperature for both penta-layered films, Fig. 2E, and trilayered films, ESI.† Following growth, films have an upper limit of $D = 1 \times 10^{-20} \text{ cm}^2 \text{ s}^{-1}$, the lowest diffusion coefficient that can be reliably inferred from this analysis, and there is no measurable diffusion at a resolution of 0.5 nm. Upon annealing to 85 K, $D = 4.8 \times 10^{-20} \text{ cm}^2 \text{ s}^{-1}$ for $T_d = 45$ K and $6.6 \times 10^{-20} \text{ cm}^2 \text{ s}^{-1}$ for $T_d = 65$ K. Thus, both systems demonstrate very small, but measureable molecular translational motions. On annealing to 100 K, the coefficient increases

to $1.77 \times 10^{-15} \text{ cm}^2 \text{ s}^{-1}$ for $T_d = 45$ K and $1.24 \times 10^{-15} \text{ cm}^2 \text{ s}^{-1}$ for $T_d = 65$ K. Both values are representative of viscous glassy materials demonstrating liquid-like behaviour above the glass transition temperature.²² From 90 K and above, diffusion constants for the $T_d = 45$ K film are consistently higher than the values for the $T_d = 65$ K film, confirming the film with $T_d = 65$ K to be the more stable glass.²³ In a separate experiment, films with $T_d = 45$ K subjected to slow thermal anneal produce a gradual reduction in surface potential as the temperature is increased, while films with $T_d = 65$ K demonstrate stable surface potentials under the same conditions; see ESI.†

Decay in the electric field for films with $T_d \leq 55$ K, Fig. 1C, is not associated with structural reorganisation caused by diffusion and therefore cannot be ascribed to the primary relaxation process. The depolarisation mechanism in MF films is directly linked to the secondary relaxation mechanism for glassy solids and electric field decay can act as a direct probe of molecular dynamics. For films of MF we see that this occurs over laboratory time scales, through molecular rotation. Films with $T_d \leq 55$ K relax *via* this secondary relaxation pathway, on laboratory timescales, while films with $T_d > 55$ K do not. We note that our observations are in accord with the formation of 'ultra-stable' glasses, as distinct from ordinary glasses. These have been observed to form if T_d falls within a critical temperature window, when grown, as here, *via* physical vapour deposition. This window has been reported to lie between 0.8 to 0.91 T_g .²⁴ Here we find that the critical window for ultra-stable MF glasses lies between 60 K and 75 K, that is, 0.66 T_g to 0.83 T_g given $T_g = 90$ K.

These real-time measurements of molecular motions are complimentary to other measurements of relaxation, such as solid-state NMR or dielectric spectroscopies, that rely on interpretation of data in the frequency domain.^{25,26} Solid-state NMR spectroscopy has revealed that molecular mobility may originate from collective oscillations of individual molecules within



a tightly confined volume, with energy barriers limiting the degree of oscillation possible.^{25,26} Such motions may not always be evident from dielectric spectroscopy measurements, leading to some confusion in the literature about the generality of the secondary relaxation pathway for species which form molecular glasses.²⁶ Other measurements of dipole ordering *via* ellipsometry or X-ray scattering, for example, have not been used to probe secondary relaxations.^{12,27}

The spontelectric effect offers a new and very sensitive means of studying relaxation phenomena in glasses far below the glass transition temperature. In a separate paper, a mean field model of the spontelectric effect is used to consider the energetics of the molecular rotations presented here. Neutron scattering data, allow us to exclude diffusion in this temperature regime. Recent work with metallic glasses has suggested that on short time scales, as reported here, it may be useful to separate out secondary relaxation processes into enthalpy and entropy changes.²⁸ Electric field measurements, similar to those reported here, may also prove useful in studying memory effects in the rejuvenation of glasses successively annealed to different temperatures on short time scales, which have been reported for a variety of glassy materials.^{29,30}

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Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 D. Cangialosi, V. M. Boucher, A. Alegría and J. Colmenero, *Phys. Rev. Lett.*, 2013, **111**, 095701.

- 2 V. M. Boucher, D. Cangialosi, A. Alegría and J. Colmenero, *J. Chem. Phys.*, 2017, **146**, 203312.
- 3 L. Song, W. Xu, J. Huo, J.-Q. Wang, X. Wang and R. Li, *Intermetallics*, 2018, **93**, 101–105.
- 4 D. P. B. Aji and G. P. Johari, *J. Chem. Phys.*, 2015, **142**, 214501.
- 5 Q. Yang, S.-X. Peng, Z. Wang and H.-B. Yu, *Natl. Sci. Rev.*, 2020, **7**, 1896–1905.
- 6 M. Gao and J. H. Perepezko, *Phys. Rev. Mater.*, 2020, **4**, 025602.
- 7 J. Fan, E. I. Cooper and C. A. Angell, *J. Phys. Chem.*, 1994, **98**, 9345–9349.
- 8 E. F. Van Dishoeck, E. Herbst and D. A. Neufeld, *Chem. Rev.*, 2013, **113**, 9043–9085.
- 9 W. Brütting, S. Berleb and A. G. Mückl, *Org. Electron.*, 2001, **2**, 1–36.
- 10 O. Shalev, S. Raghavan, J. M. Mazzara, N. Senabulya, P. D. Sinko, E. Fleck, C. Rockwell, N. Simopoulos, C. M. Jones, A. Schwendeman, G. Mehta, R. Clarke, G. E. Amidon and M. Shtein, *Nat. Commun.*, 2017, **8**, 711.
- 11 D. Field, O. Plekan, A. Cassidy, R. Balog, N. C. Jones and J. Dunger, *Int. Rev. Phys. Chem.*, 2013, **32**, 345–392.
- 12 D. Yokohama, Y. Setoguchi, A. Sakaguchi, M. Suzuki and C. Adachi, *Adv. Funt. Mater.*, 2010, **20**, 386–391.
- 13 H.-W. Lin, C.-L. Lin, H.-H. Chang, Y.-T. Lin, C.-C. Wu, Y.-M. Chen, R.-T. Chen, Y.-Y. Chien and K.-T. Wong, *J. Appl. Phys.*, 2004, **95**, 881–886.
- 14 D. Yokoyama, A. Sakaguchi, M. Suzuki and C. Adachi, *Appl. Phys. Lett.*, 2008, **93**, 173302.
- 15 I. K. Gavra, A. N. Pilidi and A. A. Tsekouras, *J. Chem. Phys.*, 2017, **146**, 104701.
- 16 G. P. Johari and M. Goldstein, *J. Chem. Phys.*, 1970, **53**, 2372–2388.
- 17 O. Plekan, A. Cassidy, R. Balog, N. C. Jones and D. Field, *Phys. Chem. Chem. Phys.*, 2012, **14**, 9972–9976.
- 18 M. Roman, A. Dunn, S. Taj, Z. G. Keolopile, A. Rosu-Finsen, M. Gutowski, M. R. S. S. McCoustra, A. M. Cassidy and D. Field, *Phys. Chem. Chem. Phys.*, 2018, **20**, 29038–29044.
- 19 A. Amir, Y. Oreg and Y. Imry, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 1850–1855.
- 20 Z. W. Wu, W. Kob, W.-H. Wang and L. Xu, *Nat. Commun.*, 2018, **9**, 5334.
- 21 D. J. Burke, F. Puletti, W. A. Brown, P. M. Woods, S. Viti and B. Slater, *Mon. Not. R. Astron. Soc.*, 2015, **447**, 1444–1451.
- 22 S. F. Swallen, K. Traynor, R. J. McMahon, M. D. Ediger and T. E. Mates, *J. Phys. Chem. B*, 2009, **113**, 4600–4608.
- 23 S. F. Swallen, K. L. Kearns, M. K. Mapes, Y. S. Kim, R. J. McMahon, M. D. Ediger, T. Wu, L. Yu and S. Satija, *Science*, 2007, **315**, 353–356.
- 24 K. L. Ngai, L. M. Wang and H. Bin Yu, *J. Phys. Chem. Lett.*, 2017, **8**, 2739–2744.
- 25 M. Vogel and E. Rössler, *J. Chem. Phys.*, 2001, **115**, 10883–10891.
- 26 M. Vogel, C. Tschirwitz, G. Schneider, C. Koplin, P. Medick and E. Rössler, *J. Non-Cryst. Solids*, 2002, **307–310**, 326–335.
- 27 T. Komino, H. Nomura, M. Yahiro and C. Adachi, *J. Phys. Chem. C*, 2012, **116**, 11584–11588.
- 28 L. Song, W. Xu, J. Huo, F. Li, L.-M. Wang, M. D. Ediger and J.-Q. Wang, *Phys. Rev. Lett.*, 2020, **125**, 135501.
- 29 H. Yardimci and R. L. Leheny, *EPL*, 2003, **62**, 203.
- 30 C. Scalliet and L. Berthier, *Phys. Rev. Lett.*, 2019, **122**, 255502.

