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Nucleophilic transformations of azido-containing carbonyl compounds *via* protection of the azido group[†]

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Nucleophilic transformations of azido-containing carbonyl compounds are discussed. The phosphazide formation from azides and di(tert-butyl)(4-(dimethylamino)phenyl)phosphine (Amphos) enabled transformations of carbonyl groups with nucleophiles such as lithium aluminum hydride and organometallic reagents. The good stability of the phosphazide moiety allowed us to perform consecutive transformations of a diazide through triazole formation and the Grignard reaction.

Azides are of great significance in a broad range of research fields including synthetic organic chemistry, chemical biology, and materials chemistry.¹ Diverse reliable transformations of azides such as copper-catalyzed azide–alkyne cycloaddition (CuAAC),² strain-promoted azide–alkyne cycloaddition (SPAAC),³ the Staudinger reaction,⁴ and reduction into amines⁵ have been used in the synthesis of various organonitrogen compounds (Fig. 1A). One drawback of azide chemistry is that an azido group is easily damaged with various nucleophiles due to the high electrophilicity.^{1,6,7} We herein disclose selective nucleophilic transformations of a carbonyl group through the protection of the azido group (Fig. 1B and C).

In the course of our studies on azide^{4d,e,8,9} and organophosphorus chemistry,^{9,10} we recently found a protection method for the azido group from CuAAC and SPAAC reactions by phosphazide formation with di(tert-butyl)(4-(dimethylamino)phenyl)phosphine (Amphos) (Fig. 1B).⁹ The phosphazide formation took place smoothly without forming azaylides through

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a denitrogenation process. The deprotection was easily realized by treating with elemental sulfur to afford azides in high yields. Considering the significance of azides in synthetic organic chemistry, we focused on the stability of phosphazides under nucleophilic or radical conditions for transformations of azides, which were still unclear due to the possible equilibrium liberation of azides from phosphazides under various conditions.⁹

We firstly examined denitrogenative reductions of azide **1a** and phosphazide **4** (Fig. 2). As a result, aniline **2** was obtained

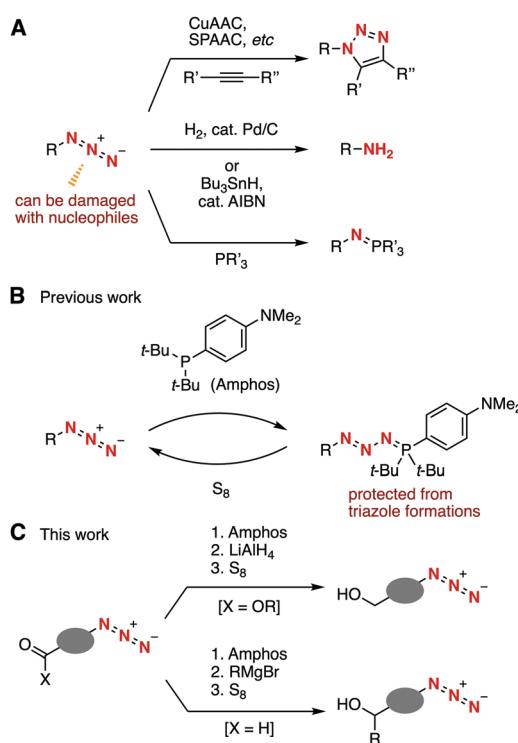


Fig. 1 Backgrounds and plan of this study. (A) Conventional transformations of azides. (B) Our previous work. (C) This work.



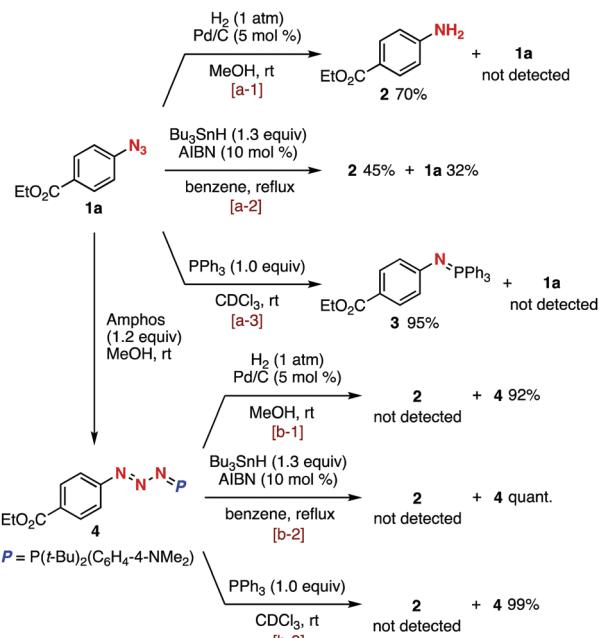


Fig. 2 Transformations of azide **1a** and phosphazide **4** under reductive conditions.

by hydrogenation of azide **1a** catalyzed by Pd/C under a hydrogen atmosphere (Fig. 2, a-1) or radical-mediated reduction with tributylstannane and 2,2'-azodiisobutyronitrile (AIBN) (Fig. 2, a-2). The Staudinger reaction of azide **1a** with triphenylphosphine furnished azaylide **3** in high yield (Fig. 2, a-3). In contrast, we found that phosphazide **4** prepared from azide **1a** with Amphos was stable under the reductive conditions using Pd/C and H₂ (Fig. 2, b-1), tributylstannane and AIBN (Fig. 2, b-2), or triphenylphosphine (Fig. 2, b-3). These results obviously indicated that azides can be protected by phosphazide formation under various reduction conditions, although azides can form in equilibrium from phosphazides.⁹

Ester-selective lithium aluminum hydride (LAH) reduction was achieved through the phosphazide formation (Fig. 3).⁶ While the treatment of azide **1a** with LAH in tetrahydrofuran (THF) at $-20\text{ }^{\circ}\text{C}$ provided 4-aminobenzyl alcohol (**5**) and 4-(ethoxycarbonyl)aniline (**2**) *via* reduction of the azido group (Fig. 3A, upper), pretreatment of azide **1a** with Amphos at room temperature followed by the addition of LAH at $-20\text{ }^{\circ}\text{C}$ and subsequent deprotection using elemental sulfur successfully afforded 4-azidobenzyl alcohol (**6a**) by selective reduction of the ethoxycarbonyl group leaving the azido group untouched (Fig. 3A, lower).¹¹ It is worth noting that the phosphazide moiety tolerated highly nucleophilic LAH. We also succeeded in the ester-selective LAH reduction of a broad range of azido-substituted esters **1** *via* the formation of phosphazides (Fig. 3B). Indeed, phosphazide formation, LAH reduction, and deprotection with elemental sulfur took place efficiently using esters **1** to afford the corresponding alcohols **6b**–**6f** in moderate to good yields *via* the protection of the heteroaromatic and primary, secondary, and tertiary alkyl azido groups. These results clearly show that a wide range of azido groups were protected using

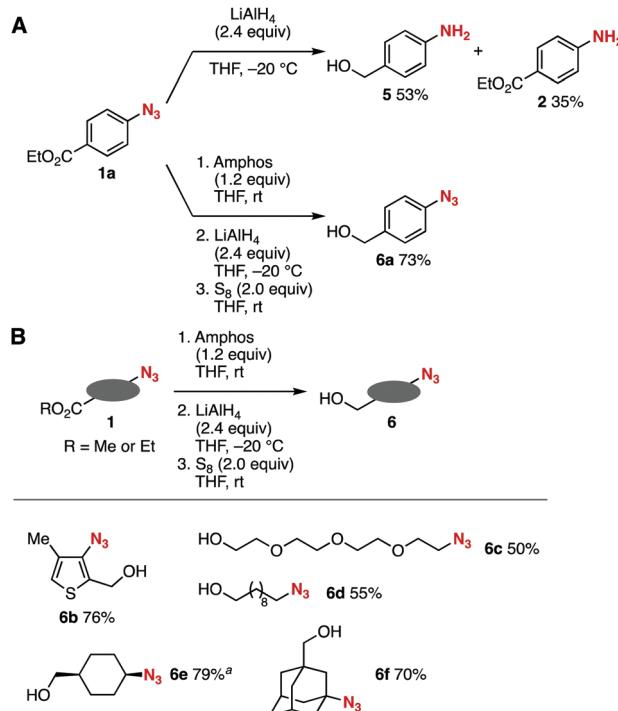


Fig. 3 Reduction of azido-substituted esters **1** with LiAlH₄. (A) Reductions with or without the protection of azide **1a**. (B) Scope of synthesized azido-substituted alcohols **6**. See ESI† for the structures of **1**.

Amphos even in the presence of highly nucleophilic reagents such as LAH.

The good stability of phosphazides toward nucleophiles allowed for alcohol synthesis from 4-azidobenzaldehyde (**7a**) with organomagnesium or organolithium reagents (Fig. 4).^{7,12} The reaction between aldehyde **7a** and ethylmagnesium bromide furnished aniline **9** in moderate yield along with a trace amount of alcohol **8a** due to the denitrogenation reduction of the azido group (Fig. 4A, upper). In sharp contrast, we accomplished an efficient synthesis of alcohol **8a** from aldehyde **7a** and ethylmagnesium bromide without damaging the highly electrophilic aromatic azido group (Fig. 4A, lower). Indeed, the phosphazide formation of azide **7a** with Amphos proceeded smoothly leaving the formyl group. Then, the addition of ethylmagnesium bromide to the resulting mixture followed by the deprotection with elemental sulfur furnished alcohol **8a** in high yield keeping the azido group unreacted. The reaction performed at a 1 mmol scale also proceeded efficiently to afford alcohol **8a** in good yield. Thus, this simple single-purification procedure enabled the efficient Grignard reaction of the formyl group without damaging the azido group.

A wide variety of carbanions successfully participated in the azido-substituted alcohol synthesis from aldehydes **7** (Fig. 4B and C). For example, phenylation, 4-chloro- and 4-methoxyphenylation, and 2-thienylation of aldehyde **7a** efficiently proceeded to provide alcohols **8b**–**8e** possessing an azido group in good yields. Benzylation and allylation of aldehyde **7a** also took place smoothly to afford azido-substituted alcohols **8f** and **8g** in moderate to high yields. Bulky alkyl Grignard reagents such as isopropylmagnesium



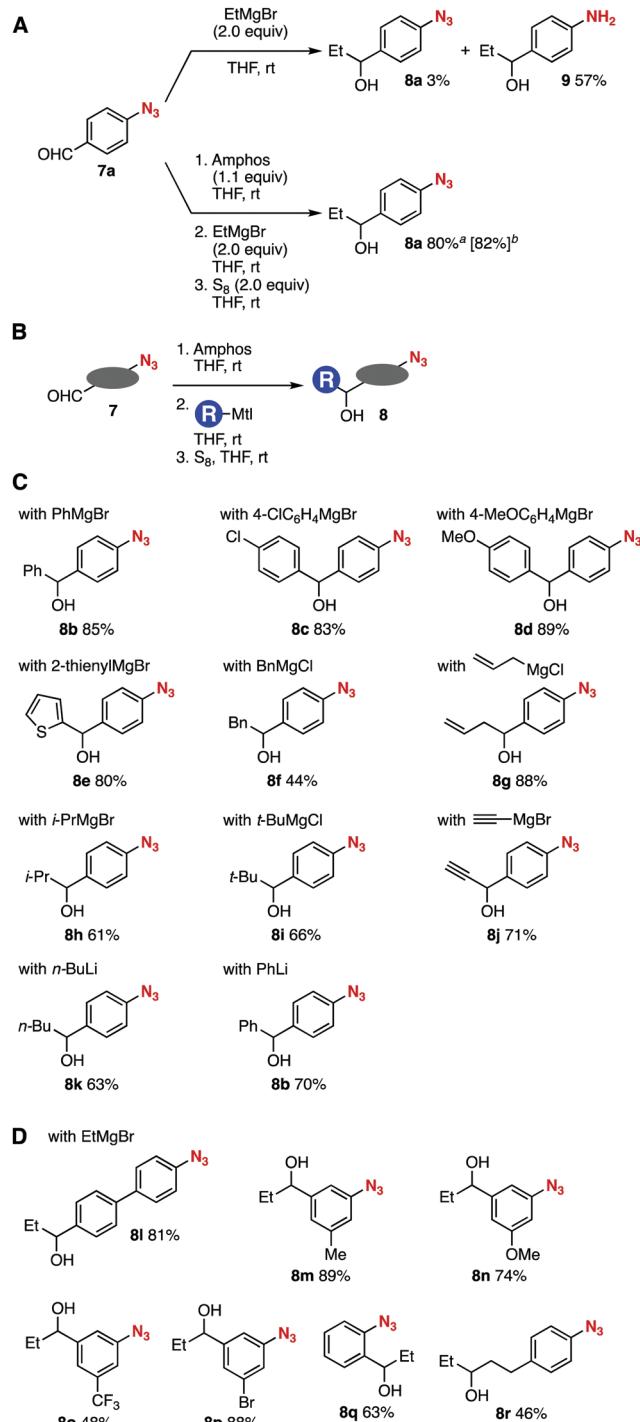


Fig. 4 Azido-substituted alcohol synthesis with organomagnesium or organolithium reagents. (A) Grignard reaction with or without the protection of azide **7a**. (B) General scheme. (C) Scope of synthesized azido-substituted alcohols **8** from **7a**. (D) Scope of synthesized azido-substituted alcohols **8** from various azido-substituted aldehydes **7**. See ESI† for the structures of **7**. ^aThe isolated yield for the reaction conducted on a 0.2 mmol scale. ^bThe isolated yield for the reaction conducted on a 1 mmol scale.

bromide and *tert*-butylmagnesium chloride reacted with aldehyde **7a** avoiding the reaction at the azido group. Of note, we achieved the introduction of a terminal alkyne moiety by the reaction using

ethynylmagnesium bromide, furnishing azido-substituted propargyl alcohol **8j** in good yield, which will serve in click chemistry. Also, organolithium reagents including *n*-butyllithium and phenyllithium were applicable to the addition reaction without damaging the phosphazide moiety.

The Grignard reactions of various aldehydes **7** bearing an azido group were accomplished to provide alcohols **8l–8r** in moderate to good yields (Fig. 4B and D). A broad range of aldehydes **7** efficiently reacted with ethylmagnesium bromide at the formyl group without damaging the azido group *via* phosphazide formation to yield alcohols **8l–8p** leaving methoxy, trifluoromethyl, and bromo groups intact. The azide protection followed by the Grignard reaction and deprotection unevenly took place when using *o*-azidobenzaldehyde to provide **8q** in good yield. Synthesis of secondary alcohol **8r** was also achieved from an aliphatic aldehyde (*3*-(4-azidophenyl)propanal) and ethylmagnesium bromide *via* the azide protection with Amphos. These results clearly show that nucleophilic additions of formyl-substituted azides with a broad range of organomagnesium and organolithium reagents through the robust phosphazides realized the preparation of diverse alcohols having an azido group.

A synthetic utility of the Grignard reaction of aldehydes having azido groups was showcased by the consecutive reactions between diazide **10**, cycloalkyne **12**, and 4-methoxyphenylmagnesium bromide (Fig. 5). Firstly, the treatment of diazide **10** with an equimolar amount of Amphos resulted in the selective protection of the aromatic azido group by virtue of the stabilization by the benzene ring without phosphazide formation at the benzyl azido moiety. Secondly, the SPAAC reaction of phosphazide **11** with cycloalkyne **12** at the remaining benzylic azido group proceeded smoothly to afford cycloadduct **13**. Thirdly, the Grignard reaction of aldehyde **13** with 4-methoxyphenylmagnesium followed by the deprotection with elemental sulfur provided azide **14** in high yield. Since azides are of great importance in not only synthetic chemistry but also chemical biology, this four-step two-pot method will allow us to construct a vast chemical library of a wide variety of triazoles having an aromatic azido moiety from formyl-substituted diazide **10**, alkynes, and organometallic reagents through the phosphazide formation.

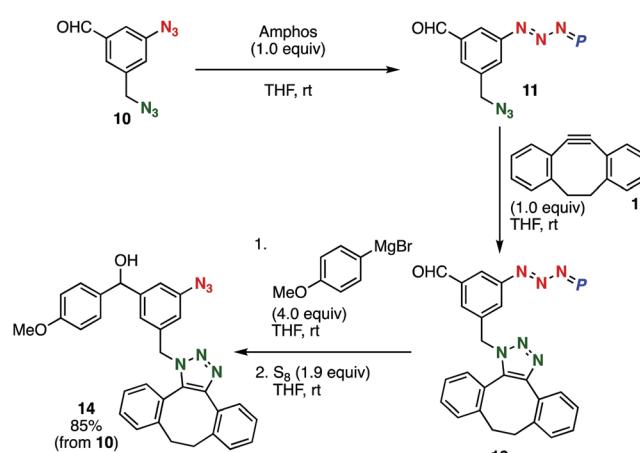


Fig. 5 Synthesis of azide **14** from diazide **10**.

In summary, we have accomplished nucleophilic transformations of diverse azides through phosphazide formation. Good stability of the phosphazide intermediates realized reduction of esters by LAH and addition reaction of aldehydes with organometallic reagents to provide a wide variety of azido-substituted alcohols. Further studies to examine detailed properties of phosphazides, especially equilibrium liberation of azides, and to expand the applicability of the azide protection method are underway in our group.

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Conflicts of interest

There are no conflicts to declare.

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