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Oxygen vacancy engineering in spinel-structured nanosheet wrapped hollow polyhedra for electrochemical nitrogen fixation under ambient conditions†

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Electrochemical nitrogen-to-ammonia conversion by the nitrogen reduction reaction (NRR) under ambient conditions is regarded as a potential approach to tackle the energy-intensive Haber–Bosch process with excessive CO₂ emission. However, the NRR is still restricted by low faradaic efficiency and NH₃ yield, which is due to the chemical inertness of N-related groups for efficient adsorption/activation on the electrocatalysts. Here, a series of spinel-structured nanosheet wrapped hollow nitrogen-doped carbon polyhedra with abundant oxygen vacancies are constructed successfully. From theoretical aspects, these materials show increased charge density on their surface for enhanced capture and activation of N₂ molecules. As a result, oxygen vacancy-rich NiCo₂O₄ on hollow N-carbon polyhedra (V_O-rich NiCo₂O₄@HNCP) shows outstanding electrocatalytic NRR performance with high production yield (NH₃: 4.1 μg h^{−1} cm^{−2}/17.8 μg h^{−1} mg^{−1}; faradaic efficiency: 5.3%) and high stability under ambient conditions and is superior to the counterpart oxygen vacancy-poor electrocatalysts. Oxygen vacancy engineering introduces a new concept for rational design of advanced NRR catalysts for energy conversion systems.

hydrogen density^{1,2} and is widely used in producing numerous chemicals, such as nitrogen fertilizers or pharmaceuticals, to satisfy the increasing demand from a booming world population.^{3–5} Up to now, over 1% of the world's energy supply is used to produce ammonia by the traditional Haber–Bosch process which is responsible for 1.6% of global CO₂ emissions and requires harsh operating conditions (400–600 °C and 20–40 MPa) due to the chemical inertness of N₂ with a high bond energy of 940.95 kJ mol^{−1}.^{6–9} In contrast, the electrochemical N₂ reduction reaction (NRR, N₂ + 6H⁺ + 6e[−] → NH₃) occurs at ambient temperature and pressure and is regarded as a more sustainable and energy-efficient process for ammonia generation.^{10–14} Although some electrocatalysts have recently been investigated for the NRR, including noble metal electrocatalysts (such as Ru, Au, and Rh),^{13–16} non-noble metal-based electrocatalysts (such as Fe-, Co-, and Mo-based materials),^{7,12,17,18} and conducting polymers (such as polyaniline and polypyrrole),^{19,20} their electrocatalytic properties are still limited by the low faradaic efficiency due to the confined electro-active sites in the corresponding bulk.^{10,11,21} Therefore, constructing multilevel nanostructures (such as one-dimensional nanofibers/nanotubes,^{22,23} two-dimensional ultrathin nanosheets,^{24,25} and three-dimensional hollow nanostructures^{26,27}) has become an efficient solution to obtain outstanding NRR electrocatalysts with a high specific surface area, thus resulting to an increased number of active sites for efficient reactant adsorption.

Apart from this, the impeded NRR process is deeply rooted in the chemical inertness of N₂ molecules. Then, the question on how to capture and activate N₂ gas much more efficiently using electrocatalysts arises. Atomic surface engineering (such as defect engineering^{28,29} and doping engineering³⁰) is the best choice to regulate the electronic structure and alter the charge density distribution. Among various types of defects, the vacancy-type defect (such as oxygen vacancies) is regarded as a kind of point defect, which is caused by independent atoms escaping from the atomic surface.^{31–33} Up to now, several possible strategies have been discovered for generating oxygen vacancies in metal oxides, including chemical reduction,³⁴ annealing in an oxygen poor atmosphere,^{35,36} and

Introduction

Ammonia (NH₃) is one of the most important carbon-free energy intermediates with low liquefying pressure and high

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heterogeneous atom substitution.³⁷ Furthermore, the oxygen vacancies created on the metal oxide surface could serve as trapping sites to capture and activate inert gas molecules of N_2 . However, to our knowledge, the proper control of the oxygen vacancy content in the electrocatalytic N_2 reduction reaction is still in the rudimentary state.

To figure out the contribution of the oxygen vacancies in NRR applications, we conducted density functional theory (DFT) calculations by using $NiCo_2O_4$ with a perfect surface and oxygen vacancy-introduced surface as two models. As shown in Fig. 1, the introduction of oxygen vacancies in $NiCo_2O_4$ is beneficial for enhancing its stability towards NNH groups due to the presence of fewer coordination sites of Ni/Co cations around the vacancies. As further proved by the differential charge density of $NiCo_2O_4$ with oxygen vacancies, the bonded electrons can be delocalized to electron donors (Ni and Co elements) near the oxygen vacancies and be activated dramatically. As a result, the $NiCo_2O_4$ with oxygen vacancies exhibits a lower stabilization energy of 0.30 eV for NNH groups than $NiCo_2O_4$ with a perfect surface (0.61 eV), indicating the positive role of oxygen vacancies in the following nitrogen reduction reactions. In addition, the delocalized electrons on the surface of $NiCo_2O_4$ with oxygen-vacancies are easier to excite to the conduction band, which is favorable not only for enhanced conductivity^{35,38} but also for activation of N-related groups, by transferring its electrons into their antibonding orbitals. At the same time, the hollow carbon polyhedron is beneficial not only for electron transfer from buried carbon to the oxygen vacancy-rich $NiCo_2O_4$ surface for the fast NRR process, but also for increasing the number of active sites for the capture of NNH groups.

In this study, we design a series of highly efficient oxygen vacancy-rich spinel-structured nanosheets on hollow N-carbon polyhedra to facilitate the electrocatalytic N_2 reduction reaction. Taking the oxygen vacancy-rich $NiCo_2O_4$ on hollow N-carbon polyhedra (V_o -rich $NiCo_2O_4@HNCP$) as an example, it exhibited a high production yield (NH_3 : $4.1 \mu g h^{-1} cm^{-2}/17.8 \mu g h^{-1} mg^{-1}$ and faradaic efficiency: 5.3%), good selectivity at $-0.25 V$ versus RHE, and high stability during the NRR process and is much more outstanding than the corresponding oxygen vacancy-poor $NiCo_2O_4$ on hollow N-carbon polyhedra (V_o -poor

$NiCo_2O_4@HNCP$). Furthermore, this oxygen vacancy engineering can also be applied in other spinel-structured nanosheets (such as $ZnCo_2O_4$ and Co_3O_4), which leads to a general solution for the design of advanced NRR electrocatalysts.

Results and discussion

According to the theoretical work, oxygen vacancy-rich $NiCo_2O_4$ on hollow N-carbon polyhedra (V_o -rich $NiCo_2O_4@HNCP$) was synthesized (Fig. S1†). As shown in Fig. 2a, the V_o -rich $NiCo_2O_4@HNCP$ displays a regular geometrical shape with a maximum length of ~ 850 nm. A blurry interface between HNCP and V_o - $NiCo_2O_4$ nanosheets can be observed in the corresponding HRTEM image (Fig. 2b), with mixed phases of amorphous carbon and crystalline $NiCo_2O_4$, which is beneficial for the charge accumulation and transfer. As a result, the two parts of HNCP and V_o -rich $NiCo_2O_4$ nanosheets can be well connected with a shared electron transfer region. Meanwhile, the fringe spacing of V_o -rich $NiCo_2O_4$ is revealed to be ~ 0.23 nm in its HRTEM image (corresponding to the (311) plane), which demonstrates a negative role of oxygen vacancies in the crystalline $NiCo_2O_4$. As shown in the X-ray diffraction (XRD) pattern of V_o -rich $NiCo_2O_4@HNCP$ (Fig. 2c), three typical peaks at $2\theta = 36.9^\circ$, 43.0° , and 62.7° can be well indexed to the (311), (400) and (440) planes of spinel-structured $NiCo_2O_4$ (JCPDS card no. 73-1702), demonstrating its successful loading on the hollow nitrogen-doped carbon polyhedra. The nine-fold higher specific surface area ($82.1 m^2 g^{-1}$) of V_o -rich $NiCo_2O_4@HNCP$ compared to the V_o -rich $NiCo_2O_4$ bulk ($9.0 m^2 g^{-1}$) (Fig. 2d) indicates its higher number of active sites, leading to shortened diffusion paths for reagents during the NRR process.

Further, for comparison, oxygen vacancy-poor $NiCo_2O_4$ on hollow N-carbon polyhedra (V_o -poor $NiCo_2O_4@HNCP$) was obtained by calcining the corresponding nickel-cobalt layered double hydroxide nanosheet wrapped hollow nitrogen-doped carbon polyhedra ($Ni-Co$ LDH@HNCP) in an oxygen atmosphere, and a series of characterization experiments was conducted (Fig. S2–S8†). In the N 1s X-ray photoelectron spectroscopy (XPS) spectrum (Fig. 2e), the peaks at 400.6 eV, 399.7 eV, and 398.7 eV corresponded to the graphitic N, pyrrolic N, and pyridinic N of the HNCP template, which is beneficial for the efficient connection of both V_o -rich and V_o -poor $NiCo_2O_4$ nanosheets to form a charge-rich region. In addition, the nitrogen content was calculated to be 11.6 wt% (Fig. S9†). The binding energies of Co 2p_{3/2} (780 eV) and Co 2p_{1/2} (795 eV) for V_o -rich $NiCo_2O_4@HNCP$ show negative shifts compared to those for V_o -poor $NiCo_2O_4@HNCP$ (Fig. 2f). This indicates that some Co^{3+} ions in the $NiCo_2O_4$ nanosheets are reduced to Co^{2+} after introduction of oxygen vacancies,^{39–42} due to the charge transfer from the V_o to the $Co(Ni)^{3+}$ with a decreased number of $Co(Ni)-O$ bonds at the surfaces (Fig. S10†). In the O 1s XPS spectra (Fig. 2g), the peaks at 529.5, 531.1 and 532.5 eV are associated with $Co(Ni)-O$ bonds (O1), oxygen vacancies or defects with lower oxygen coordination (O2), and surface adsorbed water molecules (O3), respectively.^{43–45} The proportion of the integral-area of the O2 peak in V_o -rich $NiCo_2O_4@HNCP$ is much larger than that in V_o -poor

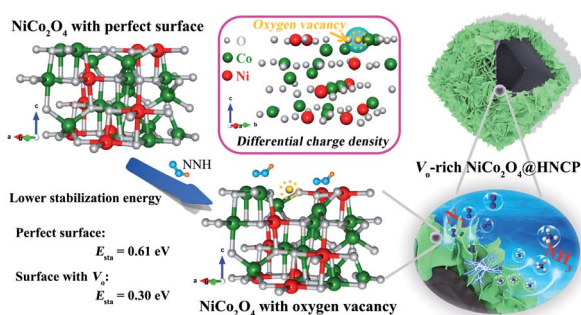


Fig. 1 Schematic illustration of the stabilization of NNH groups on the $NiCo_2O_4$ with perfect and oxygen vacancy-introduced surfaces, as well as the partial charge density of V_o -rich $NiCo_2O_4$.



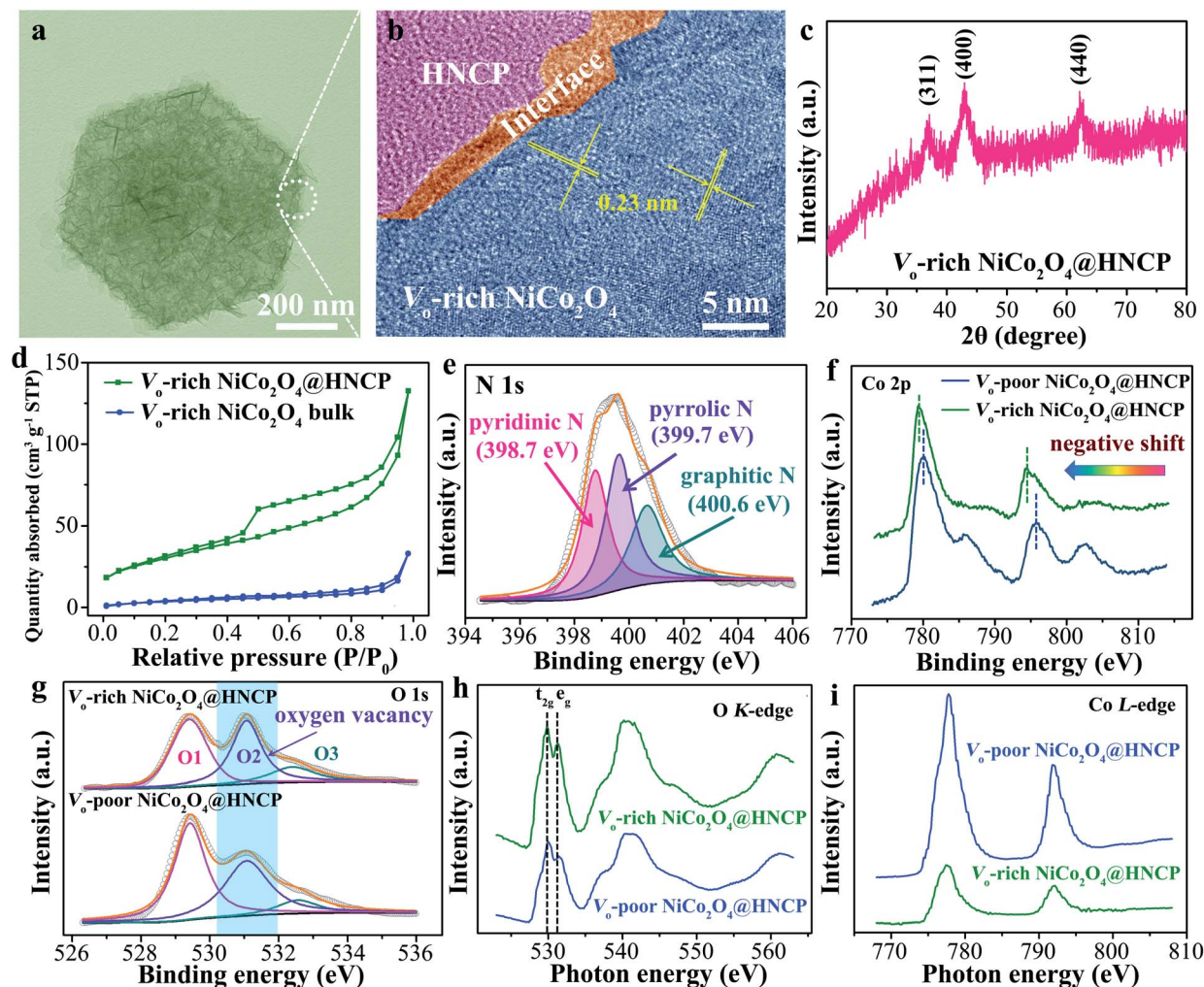


Fig. 2 (a) TEM and (b) HRTEM images of V_o -rich $NiCo_2O_4@HNCP$. (c) XRD pattern of V_o -rich $NiCo_2O_4@HNCP$. (d) N_2 adsorption isotherms of V_o -rich $NiCo_2O_4@HNCP$ and the V_o -rich $NiCo_2O_4$ bulk. (e) N 1s spectrum of V_o -rich $NiCo_2O_4@HNCP$. (f) Co 2p and (g) O 1s spectra of V_o -rich $NiCo_2O_4@HNCP$ and V_o -poor $NiCo_2O_4@HNCP$. (h) O K-edge and (i) Co L-edge XANES spectra of V_o -rich $NiCo_2O_4@HNCP$ and V_o -poor $NiCo_2O_4@HNCP$.

$NiCo_2O_4@HNCP$, indicating the successful introduction of abundant oxygen vacancies in V_o -rich $NiCo_2O_4@HNCP$.^{35,39} In addition, the V_o (O2) proportion in V_o -rich $NiCo_2O_4@HNCP$ increases from 33.0% to 68.2% as the annealing temperature increased from 200 to 300 °C (Fig. S11†), which means that the incorporated V_o amount can be well controlled by adjusting the annealing temperature. This finding could be further confirmed using the X-ray absorption near edge structure (XANES) of the O K-edge and Co L-edge. Fig. 2h shows the O K-edge XANES spectra of V_o -rich and V_o -poor $NiCo_2O_4@HNCP$ with two main folds. In the first fold at lower energy, another two sub-folds are centered at about 529.7 and 531.4 eV, which are denoted as t_{2g} and e_g by the ligand field, respectively. A relatively high intense ratio of e_g to t_{2g} can be observed for V_o -rich $NiCo_2O_4@HNCP$, indicating the successful generation of abundant V_o , which can be further proved in the magnified version of t_{2g} and e_g at the O K-edge (Fig. S12–S13†). Meanwhile, attributed to the Co 2p to 3d transition and spin–orbit coupling, the Co L-edge spectra can be split into two parts, L3-edge and L2-edge (Fig. 2i).^{46,47} The lower

peak intensity for both the Co L3-edge and L2-edge in V_o -rich $NiCo_2O_4@HNCP$ indicates that more electrons would occupy the Co 3d orbitals, providing more crucial evidence for the presence of abundant V_o on the surface of V_o -rich $NiCo_2O_4@HNCP$. A similar phenomenon is observed in the Ni L-edge spectrum (Fig. S14†). The generation of oxygen vacancies is due to the inadequate oxidation of $NiCo_2O_4$ in an oxygen-deficient atmosphere as compared to an oxygen atmosphere, leading to abundant partially unoxidized metal domains in the $NiCo_2O_4$ structure.^{48,49}

To reveal the role of oxygen vacancies in tuning the electronic structures of the $NiCo_2O_4$ nanosheets, DFT calculations were performed for different types of oxygen vacancies. Considering that there are two types of oxygen (e.g. surface and sub-surface oxygens) in the perfect crystallographic lattice of $NiCo_2O_4$ (Fig. 3a), we separately studied each situation regarding their density of states, formation possibility and adsorption ability towards the key NRR intermediates of NNH groups. The surface oxygen was removed first for building the



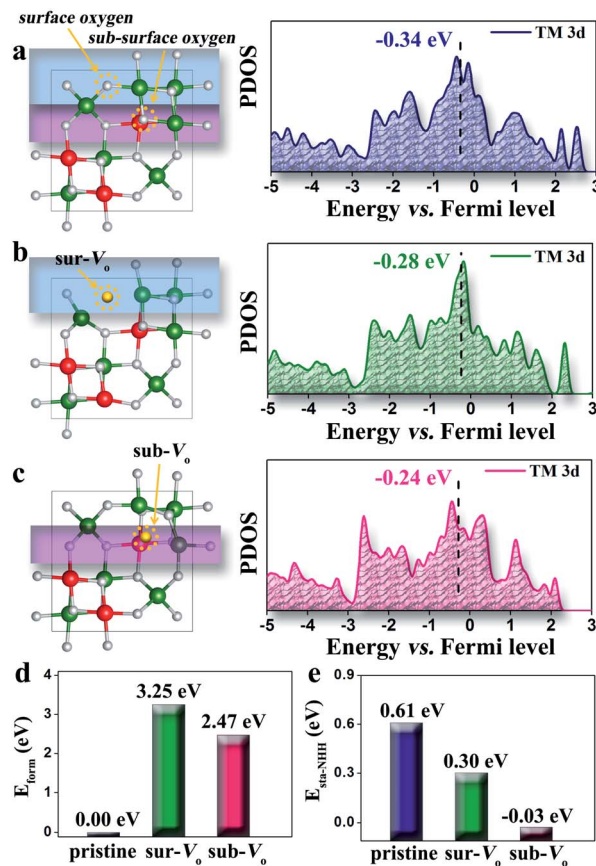


Fig. 3 (a) Pristine NiCo₂O₄ structure and projected density of states (PDOS) of the surface transition metal. (b) The defect model with removal of a surface oxygen vacancy (sur-V_o) and its PDOS. (c) The defect model with removal of a sub-surface oxygen vacancy (sub-V_o) and its PDOS. (d) The formation energy of oxygen vacancies for the three catalyst models. (e) Adsorption energy of NNH groups in the three catalyst models.

surface oxygen vacancy (sur-V_o) model, as shown in Fig. 3b. We found that the d-band center of the sur-V_o model is -0.28 eV, much closer to the Fermi level compared with that of the pristine NiCo₂O₄ model (-0.34 eV). This result indicates that oxygen vacancies can induce more activity of the metal atoms, that is, the adsorption strength of molecules and radicals over the metal atoms will be enhanced. As shown in Fig. 3c, we tested another type of oxygen vacancy, the sub-surface oxygen vacancy (sub-V_o) model, which has the closest d-band center (-0.24 eV) to the Fermi level among the three models. In the subsequent calculation, we found that the possibility of the formation of these two types of oxygen vacancies is not the same (Fig. 3d). As a more positive formation energy denotes that a vacancy is more difficult to form, the results show that sub-V_o (2.47 eV) has a higher stability than sur-V_o (3.25 eV). Finally, we calculated the adsorption strength of the three models by using NNH as a probe and found that the sub-V_o model presents the most favorable energy (-0.03 eV). The adsorption energy of the sur-V_o (0.30 eV) is also more negative than that of the pristine model (0.61 eV), indicating its enhanced adsorption ability, which is consistent with the d-band center results. According to the

trend of formation energy, we can deduce that the sub-V_o is preferentially formed when the vacancy density is relatively low. Therefore, the adsorption enhancement will be significantly dependent on the vacancy density in the region of low vacancy density.

To illustrate the role of oxygen vacancies in the NRR, a two-compartment cell was assembled by employing V_o-rich NiCo₂O₄@HNCP as the cathode catalyst with continuous N₂ bubbling in a 0.1 M Na₂SO₄ electrolyte. The highest average yields and corresponding faradaic efficiencies of V_o-rich NiCo₂O₄@HNCP are achieved when the negative potential increases to -0.25 V versus the reversible hydrogen electrode (RHE) and are calculated to be about 4.1 μg h⁻¹ cm⁻²/17.8 μg h⁻¹ mg⁻¹ and 5.3%, respectively (Fig. 4a and S15[†]). This production yield at -0.25 V is also confirmed by the nuclear magnetic resonance (NMR) method as shown in Fig. S16–S18 and Table S1.[†] Beyond this negative potential, the NH₃ yields and faradaic efficiencies decrease significantly because of the overwhelming competition from the hydrogen evolution reaction (HER). The NRR performance of the V_o-rich NiCo₂O₄@HNCP catalyst is much more outstanding than that of the V_o-poor NiCo₂O₄@HNCP catalyst (1.6 μg h⁻¹ cm⁻²/6.9 μg h⁻¹ mg⁻¹ and 1.8% at -0.25 V), as well as other results under ambient conditions or at high temperatures and pressures (Table S2[†]). Note that no hydrazine is detected in the electrolyte (Fig. S19[†]). In addition, the charge amounts of the experimentally quantified NH₃ and H₂ gases and the calculated one (Fig. S20[†]) are roughly in agreement (S2 ≈ S3). These results indicate the high selectivity of the V_o-rich NiCo₂O₄@HNCP catalyst for NH₃ generation except for H₂ gas. For practical use, stability is another critical criterion to evaluate the NRR performance of a catalyst. As shown in Fig. 4b and S21–26,[†] the NH₃ yield, faradaic efficiency, and current density of the V_o-rich NiCo₂O₄@HNCP catalyst are all stable without obvious fluctuation. As shown in Fig. S27,[†] the V_o-rich NiCo₂O₄@HNCP catalyst displays ultra-stable properties in both NH₃ yield and faradaic efficiency even after working for 100 h and is comparable to the recently reported catalysts but with a much longer practical life.⁵⁰ By using XRD and XPS analyses (Fig. S28[†]), the V_o-rich NiCo₂O₄@HNCP catalyst is found to exhibit an unchanged crystal structure and only slightly decreased oxygen vacancy content after working for 100 h, which demonstrates the relatively stable structure of V_o-rich NiCo₂O₄@HNCP as an NRR catalyst. Due to the similar results of NH₃ yield and faradaic efficiency after varying the nitrogen flow rate at -0.25 V (Fig. 4c), the N₂ diffusion process becomes a non-rate-determining step as it is an independent gas-solid interface. Moreover, by varying the reaction temperature of the NRR, both the NH₃ yield and faradaic efficiency of the V_o-rich NiCo₂O₄@HNCP catalyst increase simultaneously (Fig. 4d). For instance, the NRR yield is about 2.5 times higher at 60 °C than at 0 °C, indicating that mass transfer plays a key role in enhancing the reaction rate of the V_o-rich NiCo₂O₄@HNCP catalyst. In accordance with the Arrhenius equation and Arrhenius plot (Fig. S29[†]), the apparent activation energy of V_o-rich NiCo₂O₄@HNCP for the NRR is calculated to be 11.4 kJ mol⁻¹. The UV/Vis absorption spectra of various samples (Fig. 4e) and photographs of NH₄⁺-containing solutions before



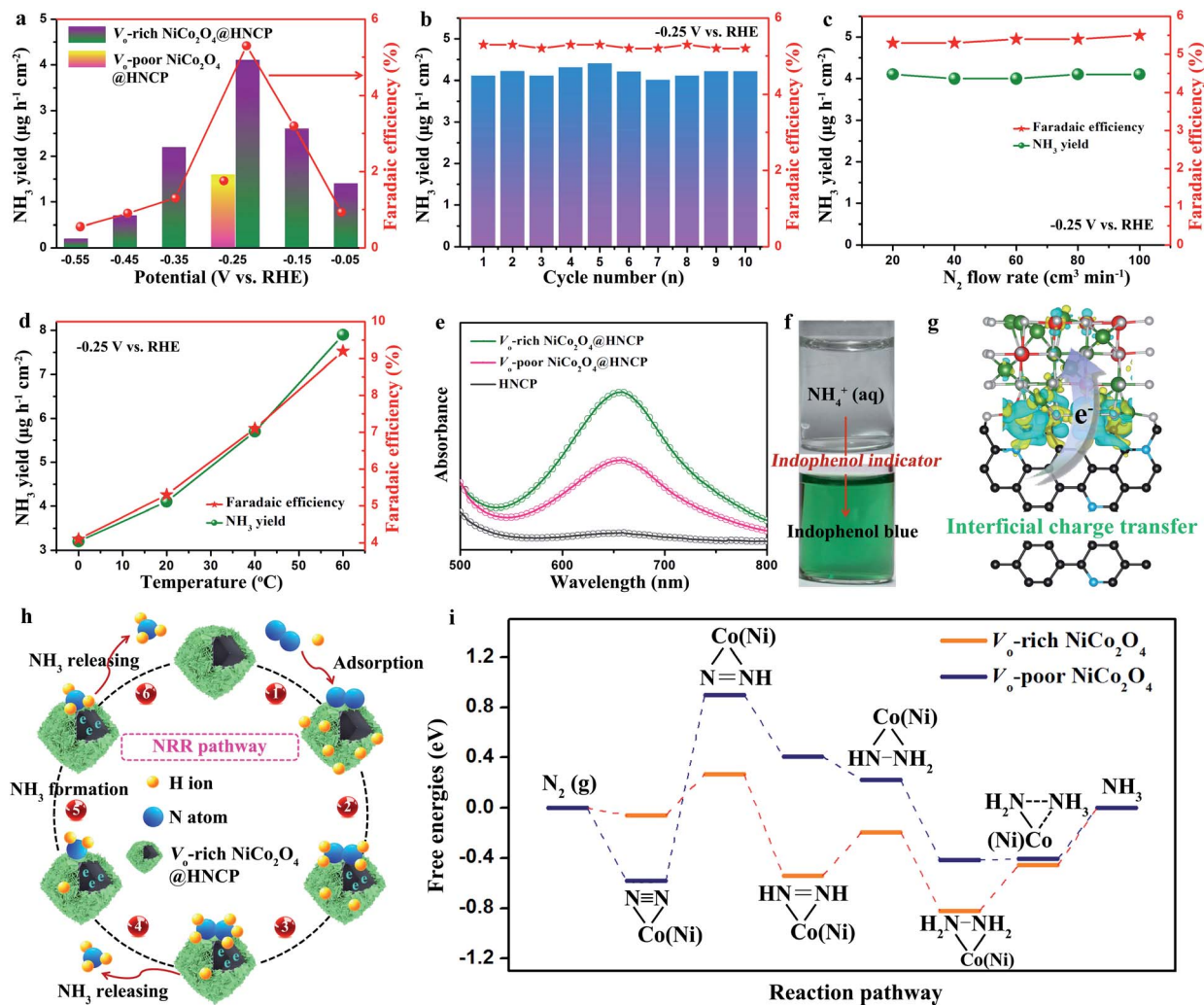


Fig. 4 Electrocatalytic NRR of V_o -rich and V_o -poor $NiCo_2O_4@HNCP$ in an aqueous solution of 0.1 M Na_2SO_4 under ambient conditions. (a) Yield of NH_3 (bar graph) and faradaic efficiency (red) at each given potential. (b) Recycling test of V_o -rich $NiCo_2O_4@HNCP$ at a potential of -0.25 V. Yield rate of NH_3 (green) and faradaic efficiency at different (c) N_2 flow rates and (d) temperatures at -0.25 V. (e) UV-Vis absorption spectra of various samples after chronoamperometry testing at -0.25 V versus RHE. (f) NH_4^+ -containing solution stained with indophenol indicator. (g) Charge density distribution of V_o -rich $NiCo_2O_4@HNCP$ around the interface and oxygen vacancy. (h) Schematic illustration of the proposed electrochemical NRR pathway by using the V_o -rich $NiCo_2O_4@HNCP$ catalyst. (i) Free energy diagram of the electrochemical reduction of N_2 to NH_3 in an alternating pathway on the surface of the V_o -rich $NiCo_2O_4@HNCP$ catalyst.

and after staining with indophenol indicator (Fig. 4f) are exhibited. After comparing the EIS spectra (Fig. S30†) and the performance of the V_o -rich $NiCo_2O_4$ bulk ($1.4\ \mu g\ h^{-1}\ cm^{-2}/6.1\ \mu g\ h^{-1}\ mg^{-1}$ and 1.6% at -0.25 V) in Fig. S31,† it could be seen that the hollow nitrogen-doped carbon polyhedron template in V_o -rich $NiCo_2O_4@HNCP$ is extremely vital for its enhanced NRR performance, due to the increased specific surface area (Fig. 2d) and well-formed interface between HNCP and V_o -rich $NiCo_2O_4$ nanosheets for charge accumulation/transfer. For the verification of the source of ammonia, a ^{15}N isotope labeling experiment using NMR (600 MHz) was performed. As shown in Fig. S32,† the 1H NMR signals of $^{14}NH_4^+$ produced ^{14}N triplets in the region of 6.8–7.1 ppm. Importantly, the 1H NMR signal of $^{15}NH_4^+$ only produced an ^{15}N doublet without the appearance of ^{14}N triplets in the same region, which demonstrates that the

doped nitrogen atoms in the HNCP template are stable and do not escape from the structures. Therefore, the source of ammonia is the feed N_2 gas rather than the electrocatalyst.

For further detailed insights into the structure–activity relationship between the V_o -rich $NiCo_2O_4@HNCP$ catalyst and its enhanced NRR properties, DFT calculation was carried out from the perspective of thermodynamic and kinetic acceleration. First, from the thermodynamic point of view, the electrons that previously occupied the O 2p orbital would partially delocalized to the neighboring Ni/Co cations after introduction of oxygen vacancies, which suggests that a surface with oxygen vacancies is much more activated. Meanwhile, a rational model for the V_o -rich $NiCo_2O_4@HNCP$ catalyst was constructed by aligning optimized V_o -rich $NiCo_2O_4$ with highly conductive nitrogen-doped carbon (Fig. 4g and S33†). After analysis in detail,



abundant charges accumulated on the blurry interface between V_o -rich NiCo_2O_4 and nitrogen-doped carbon, leading to the successful generation of highly active regions for the electrocatalytic process. As shown in Fig. S34 and S35,[†] the electrons are prone to transfer from nitrogen-doped carbon to V_o -rich NiCo_2O_4 due to their inconsistent charge distribution. In turn, the Ni/Co cations show increased electronic states of the d-orbital around the Fermi level, which is beneficial for activation of N_2 and formation of an N-catalyst bond. From the kinetic point of view, the hollow nitrogen-doped carbon polyhedron template endows the V_o -rich NiCo_2O_4 nanosheets with a higher specific surface area with more exposed active sites for adsorption/activation of N-related species. Attributed to the structural merits of the V_o -rich NiCo_2O_4 @HNCP catalyst, a possible NRR mechanism is depicted in Fig. 4h and i. Firstly, the N_2 gas can be easily adsorbed on the surface of the V_o -rich NiCo_2O_4 @HNCP catalyst to form chemisorbed $\text{Co}(\text{Ni})\text{-N}_2$ bonds, which can be denoted as $\text{N}_2(\text{g}) \rightarrow \text{N}_2^*$ (here, the asterisk * denotes an adsorption site). After full structural relaxation, two energetically favorable configurations were found: end-on and side-on. On the V_o -poor NiCo_2O_4 surface, an N_2 molecule was more likely to anchor in the end-on configuration with a Co–N bond length of 1.84 Å (Fig. S36a[†]), whereas, on the V_o -rich NiCo_2O_4 surface, two Co–N bonds (1.20 Å in length) in the side-on configuration are formed with an elongated N–N bond length from 1.12 Å to 1.16 Å (Fig. S36b[†]). When taking ΔE_{ZPE} and entropy into consideration, the ΔG values for N_2 adsorption are -0.05 eV for the V_o -rich NiCo_2O_4 surface and -0.57 eV for the V_o -poor NiCo_2O_4 surface, respectively. Secondly, six consecutive protonation and reduction processes on the V_o -rich(poor) NiCo_2O_4 surfaces were further proposed with atomic configurations at various states of each elementary step (Fig. S37[†]). As discussed in Fig. 4i and Table S3,[†] the free energies of all these states were slightly downhill for the V_o -rich NiCo_2O_4 surface compared to the V_o -poor NiCo_2O_4 surface, which also proves that the successful introduction of oxygen vacancies in NiCo_2O_4 is beneficial for the nitrogen reduction reaction process theoretically.

We further studied the NRR activity of other spinel-structured nanosheet wrapped HNCP, such as ZnCo_2O_4 @HNCP and Co_3O_4 @HNCP, in which oxygen vacancies were deliberately introduced. It should be noted that the XPS O 1s spectra (Fig. S38[†]) reveal that abundant oxygen vacancies were successfully introduced into the surface of both the V_o -rich ZnCo @HNCP and V_o -rich Co @HNCP catalysts. As shown in Fig. 5a, all the catalysts with abundant oxygen vacancies showed a higher NH_3 yield rate and faradaic efficiency than their V_o -poor counterparts in 0.1 M Na_2SO_4 . Taking V_o -rich ZnCo @HNCP and its V_o -poor counterpart as examples, the NH_3 yield rate and faradaic efficiency for the V_o -rich ZnCo @HNCP catalyst were $3.7 \mu\text{g h}^{-1} \text{cm}^{-2}$ / $16.0 \mu\text{g h}^{-1} \text{mg}^{-1}$ and 3.6%, respectively, which are much more outstanding than those of the V_o -poor ZnCo @HNCP catalyst ($1.6 \mu\text{g h}^{-1} \text{cm}^{-2}$ / $6.9 \mu\text{g h}^{-1} \text{mg}^{-1}$ and 1.7% at -0.25 V). Besides, the time-dependent curves of various catalysts at -0.25 V (Fig. 5b) exhibited excellent stability, indicating the stable vacancy structure in both the V_o -rich ZnCo @HNCP and V_o -rich Co @HNCP catalysts during the NRR process.

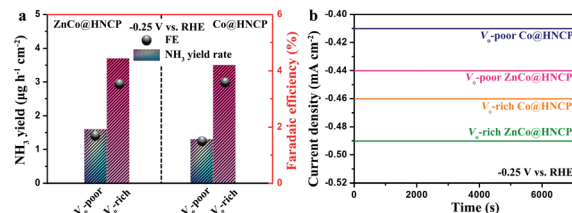


Fig. 5 (a) Average NH_3 yields and faradaic efficiency of other spinel-structured nanosheet wrapped HNCP (ZnCo@HNCP and Co@HNCP) with and without oxygen vacancies after chronoamperometry testing at -0.25 V versus RHE. (b) Time-dependent current density curves of various catalysts for the NRR at -0.25 V versus RHE.

Conclusions and outlook

In summary, we uncovered the positive role of oxygen vacancies in spinel-structured nanosheets on hollow N-carbon polyhedra (e.g. V_o -rich NiCo_2O_4 @HNCP, V_o -rich ZnCo_2O_4 @HNCP, and V_o -rich Co_3O_4 @HNCP) towards the electrocatalysis of the nitrogen reduction reaction. For example, the V_o -rich NiCo_2O_4 @HNCP catalyst showed higher production yield (NH_3 : $4.1 \mu\text{g h}^{-1} \text{cm}^{-2}$ / $17.8 \mu\text{g h}^{-1} \text{mg}^{-1}$ and faradaic efficiency: 5.3%), good selectivity and high stability when compared with its V_o -poor counterpart (NH_3 : $1.6 \mu\text{g h}^{-1} \text{cm}^{-2}$ / $6.9 \mu\text{g h}^{-1} \text{mg}^{-1}$ and faradaic efficiency: 1.8%). As revealed by DFT calculations, the oxygen vacancies enhance the reactivity of the active sites, leading to reduced stabilization energy of NNH groups. Meanwhile, the HNCP template increased the number of active sites for facilitating the reaction on the surface of V_o -rich NiCo_2O_4 nanosheets and improved the conductivity for interfacial electron transfer between V_o -rich NiCo_2O_4 and HNCP. Systematic study from both theoretical and experimental aspects further confirmed the superior kinetics for NH_3 production using V_o -rich NiCo_2O_4 @HNCP. Therefore, this work provides a guideline for the rational design of novel and highly efficient catalysts towards N_2 electrochemical reduction by increasing the surface area and introducing surface oxygen vacancies in spinel-structured nanosheets simultaneously.

Experimental section

Details of the synthetic procedures, characterization and theoretical calculation methods can be found in the ESI[†].

Conflicts of interest

The authors declare no conflict of interest.

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References

- 1 T. Spatzal, K. A. Perez, O. Einsle, J. B. Howard and D. C. Rees, *Science*, 2014, **345**, 1620–1623.
- 2 R. Lan, J. Irvine and S. W. Tao, *Int. J. Hydrogen Energy*, 2012, **37**, 1482–1492.
- 3 K. A. Brown, D. F. Harris, M. B. Wilker, A. Rasmussen, N. Khadka, H. Hamby, S. Keable, G. Dukovic, J. W. Peters, L. C. Seefeldt and P. W. King, *Science*, 2016, **352**, 448–450.
- 4 B. M. Hoffman, D. Lukoyanov, Z. Y. Yang, D. R. Dean and L. C. Seefeldt, *Chem. Rev.*, 2014, **114**, 4041–4062.
- 5 V. Rosca, M. Duca, M. T. de Groot and M. Koper, *Chem. Rev.*, 2009, **109**, 2209–2244.
- 6 M. Kitano, Y. Inoue, Y. Yamazaki, F. Hayashi, S. Kanbara, S. Matsuiishi, T. Yokoyama, S. W. Kim, M. Hara and H. Hosono, *Nat. Chem.*, 2012, **4**, 934–940.
- 7 S. Kuriyama, K. Arashiba, K. Nakajima, Y. Matsuo, H. Tanaka, K. Ishii, K. Yoshizawa and Y. Nishibayashi, *Nat. Commun.*, 2016, **7**, 12181.
- 8 M. Ali, F. L. Zhou, K. Chen, C. Kotzur, C. L. Xiao, L. Bourgeois, X. Y. Zhang and D. R. MacFarlane, *Nat. Commun.*, 2016, **7**, 11335.
- 9 Y. H. Lu, Y. Yang, T. F. Zhang, Z. Ge, H. C. Chang, P. S. Xiao, Y. Y. Xie, L. Hua, Q. Y. Li, H. Y. Li, B. Ma, N. J. Guan, Y. F. Ma and Y. S. Chen, *ACS Nano*, 2016, **10**, 10507–10515.
- 10 C. X. Guo, J. R. Ran, A. Vasileff and S. Z. Qiao, *Energy Environ. Sci.*, 2018, **11**, 45–56.
- 11 G. F. Chen, X. R. Cao, S. Q. Wu, X. Y. Zeng, L. X. Ding, M. Zhu and H. H. Wang, *J. Am. Chem. Soc.*, 2017, **139**, 9771–9774.
- 12 Y. Nishibayashi, *Inorg. Chem.*, 2015, **54**, 9234–9247.
- 13 S. J. Li, D. Bao, M. M. Shi, B. R. Wulan, J. M. Yan and Q. Jiang, *Adv. Mater.*, 2017, **29**, 170000133.
- 14 M. M. Shi, D. Bao, B. R. Wulan, Y. H. Li, Y. F. Zhang, J. M. Yan and Q. Jiang, *Adv. Mater.*, 2017, **29**, 160655017.
- 15 V. Kordali, G. Kyriacou and C. Lambrou, *Chem. Commun.*, 2000, 1673–1674.
- 16 H. M. Liu, S. H. Han, Y. Zhao, Y. Y. Zhu, X. L. Tian, J. H. Zeng, J. X. Jiang, B. Y. Xia and Y. Chen, *J. Mater. Chem. A*, 2018, **6**, 3211–3217.
- 17 K. Arashiba, E. Kinoshita, S. Kuriyama, A. Eizawa, K. Nakajima, H. Tanaka, K. Yoshizawa and Y. Nishibayashi, *J. Am. Chem. Soc.*, 2015, **137**, 5666–5669.
- 18 S. Kuriyama, K. Arashiba, H. Tanaka, Y. Matsuo, K. Nakajima, K. Yoshizawa and Y. Nishibayashi, *Angew. Chem. Int. Ed.*, 2016, **55**, 14291–14295; *Angew. Chem.*, 2016, **128**, 14503–14507.
- 19 F. Koleli and D. B. Kayan, *J. Electroanal. Chem.*, 2010, **638**, 119–122.
- 20 F. Koleli and T. Ropke, *Appl. Catal., B*, 2006, **62**, 306–310.
- 21 R. Lan, J. Irvine and S. W. Tao, *Sci. Rep.*, 2013, **3**, 1145.
- 22 F. L. Lai, D. Y. Yong, X. L. Ning, B. C. Pan, Y. E. Miao and T. X. Liu, *Small*, 2017, **13**, 16028667.
- 23 F. X. Ma, L. Yu, C. Y. Xu and X. W. Lou, *Energy Environ. Sci.*, 2016, **9**, 862–866.
- 24 S. C. Chen, Z. X. Kang, X. Hu, X. D. Zhang, H. Wang, J. F. Xie, X. S. Zheng, W. S. Yan, B. C. Pan and Y. Xie, *Adv. Mater.*, 2017, **29**, 170168730.
- 25 J. Q. Xu, X. D. Li, W. Liu, Y. F. Sun, Z. Y. Ju, T. Yao, C. M. Wang, H. X. Ju, J. F. Zhu, S. Q. Wei and Y. Xie, *Angew. Chem. Int. Ed.*, 2017, **56**, 9121–9125; *Angew. Chem.*, 2017, **129**, 9249–9253.
- 26 B. Y. Guan, S. L. Zhang and X. W. Lou, *Angew. Chem. Int. Ed.*, 2018, **57**, 6176–6180; *Angew. Chem.*, 2018, **130**, 6284–6288.
- 27 L. Yu, J. F. Yang, B. Y. Guan, Y. Lu and X. W. Lou, *Angew. Chem. Int. Ed.*, 2018, **57**, 172–176; *Angew. Chem.*, 2018, **130**, 178–182.
- 28 Y. Tong, Y. Q. Guo, K. J. Mu, H. Shan, J. Dai, Y. Liu, Z. Sun, A. D. Zhao, X. C. Zeng, C. Z. Wu and Y. Xie, *Adv. Mater.*, 2017, **29**, 1703123.
- 29 J. F. Xie and Y. Xie, *ChemCatChem*, 2015, **7**, 2568–2580.
- 30 Y. Shi, Y. Zhou, D. R. Yang, W. X. Xu, C. Wang, F. B. Wang, J. J. Xu, X. H. Xia and H. Y. Chen, *J. Am. Chem. Soc.*, 2017, **139**, 15479–15485.
- 31 A. R. Singh, B. A. Rohr, J. A. Schwalbe, M. Cargnello, K. Chan, T. F. Jaramillo, I. Chorkendorff and J. K. Nørskov, *ACS Catal.*, 2017, **7**, 706–709.
- 32 C. van der Ham, M. Koper and D. Hettterscheid, *Chem. Soc. Rev.*, 2014, **43**, 5183–5191.
- 33 H. K. Lee, C. Koh, Y. H. Lee, C. Liu, I. Y. Phang, X. M. Han, C. K. Tsung and X. Y. Ling, *Sci. Adv.*, 2018, **4**, eaar32083.
- 34 S. J. Peng, F. Gong, L. L. Li, D. S. Yu, D. X. Ji, T. R. Zhang, Z. Hu, Z. Q. Zhang, S. L. Chou, Y. H. Du and S. Ramakrishna, *J. Am. Chem. Soc.*, 2018, **140**, 13646–13653.
- 35 J. Bao, X. D. Zhang, B. Fan, J. J. Zhang, M. Zhou, W. L. Yang, X. Hu, H. Wang, B. C. Pan and Y. Xie, *Angew. Chem. Int. Ed.*, 2015, **54**, 7399–7404; *Angew. Chem.*, 2015, **127**, 7507–7512.
- 36 H. Yuan, J. T. Li, W. Yang, Z. C. Zhuang, Y. Zhao, L. He, L. Xu, X. B. Liao, R. Q. Zhu and L. Q. Mai, *ACS Appl. Mater. Interfaces*, 2018, **10**, 16410–16417.
- 37 Y. X. Zeng, Z. Z. Lai, Y. Han, H. Z. Zhang, S. L. Xie and X. H. Lu, *Adv. Mater.*, 2018, **30**, 1802396.
- 38 Y. C. Wang, T. Zhou, K. Jiang, P. M. Da, Z. Peng, J. Tang, B. A. Kong, W. B. Cai, Z. Q. Yang and G. F. Zheng, *Adv. Energy Mater.*, 2014, **4**, 140069616.
- 39 W. J. Xu, F. L. Lyu, Y. C. Bai, A. Q. Gao, J. Feng, Z. X. Cai and Y. D. Yin, *Nano Energy*, 2018, **43**, 110–116.
- 40 S. Zhang, P. Kang and T. J. Meyer, *J. Am. Chem. Soc.*, 2014, **136**, 1734–1737.
- 41 L. Xu, Q. Q. Jiang, Z. H. Xiao, X. Y. Li, J. Huo, S. Y. Huang and L. M. Dai, *Angew. Chem. Int. Ed.*, 2016, **55**, 5277–5281; *Angew. Chem.*, 2016, **128**, 5363–5367.
- 42 J. J. Wu, R. M. Yadav, M. J. Liu, P. P. Sharma, C. S. Tiwary, L. L. Ma, X. L. Zou, X. D. Zhou, B. I. Yakobson, J. Lou and P. M. Ajayan, *ACS Nano*, 2015, **9**, 5364–5371.
- 43 Y. L. Ge, K. Kan, Y. Yang, L. Zhou, L. Q. Jing, P. K. Shen, L. Li and K. Y. Shi, *J. Mater. Chem. A*, 2014, **2**, 4961–4969.
- 44 V. M. Jimenez, A. Fernandez, J. P. Espinos and A. R. Gonzalez-Elipio, *J. Electron Spectrosc. Relat. Phenom.*, 1995, **71**, 61–71.



- 45 M. Ma, K. Zhang, P. Li, M. S. Jun, M. J. Jeong and J. H. Park, *Angew. Chem. Int. Ed.*, 2016, **55**, 11819–11823; *Angew. Chem.*, 2016, **128**, 11998–12002.
- 46 N. Kornienko, J. Resasco, N. Becknell, C. M. Jian, Y. S. Liu, K. Q. Nie, X. H. Sun, J. H. Guo, S. R. Leone and P. D. Yang, *J. Am. Chem. Soc.*, 2015, **137**, 7448–7455.
- 47 P. Z. Chen, T. P. Zhou, M. L. Chen, Y. Tong, N. Zhang, X. Peng, W. S. Chu, X. J. Wu, C. Z. Wu and Y. Xie, *ACS Catal.*, 2017, **7**, 7405–7411.
- 48 X. R. Li, J. L. Wei, Q. Li, S. S. Zheng, Y. X. Xu, P. Du, C. Y. Chen, J. Y. Zhao, H. G. Xue, Q. Xu and H. Pang, *Adv. Funct. Mater.*, 2018, **28**, 1800886.
- 49 L. Hua, Z. Y. Hui, Y. Sun, X. Zhao, H. Xu, Y. J. Gong, R. Y. Chen, C. Y. Yu, J. Y. Zhou, G. Z. Sun and W. Huang, *Nanoscale*, 2018, **10**, 21006–21012.
- 50 H. Cheng, L. X. Ding, G. F. Chen, L. L. Zhang, J. Xue and H. H. Wang, *Adv. Mater.*, 2018, **30**, 1803694.

