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## Regioisomeric control of layered crystallinity in solution-processable organic semiconductors†

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The construction and control of 2D layered molecular packing motifs with functionally substituted  $\pi$ -electron cores are crucial for developing organic electronic materials and devices. We investigated a regioisomeric structure–property relationship in high-performance and solution-processable layered organic semiconductors based on *mono*-octyl-substituted benzothieno[3,2-*b*]naphtho[2,3-*b*]thiophene (*mono*-C8-BTNT). We demonstrated that an isomorphous *bilayer-type* layered herringbone packing motif is obtainable in a series of four positional isomers of *mono*-C8-BTNTs whose  $\pi$ -electron core is substituted by an octyl chain at one of the four most peripheral positions with roughly keeping the rod-like molecular shape. These regioisomeric compounds exhibited systematic variations in the solvent solubility and liquid-crystalline phase transitions at elevated temperatures. The analysis of intermolecular interaction energies in the crystals based on dispersion-corrected DFT calculations revealed that the crystals of 2- and 8-*mono*-C8-BTNTs are more stable than those of 3- and 9-*mono*-C8-BTNTs owing to the higher ordering of alkyl chain layers in the crystals. Such differences of the stability in their crystal formation are closely correlated with TFT performances, where the single-crystal devices of the 2- and 8-*mono*-C8-BTNTs substituted at the most peripheral positions exhibit high-performance TFT characteristics with a mobility of approximately  $10 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ .

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## Introduction

Solution-processable organic semiconductors (OSCs) are composed, in most cases, of  $\pi$ -electron cores bonded with functional substituents.<sup>1–5</sup> The substituents are structurally flexible and electrically inert, as observed in typical substituents such as normal and branched alkyl chains,<sup>6–15</sup> silylethynyl,<sup>16–19</sup> fluoroalkyl,<sup>20–24</sup> or others,<sup>25–29</sup> which is in contrast to the rigid nature of  $\pi$ -electron cores that play major roles in the carrier

transport characteristics. However, the role of the substituents is often critical, not only for increasing the solubility characteristics, but also for affecting the whole material characteristics through the variation of intermolecular packing geometries. The important issue is to identify optimal combinations between the  $\pi$ -electron cores and the substituents, including their substituting positions, to develop prolific organic electronic applications.

Among a variety of substituents, normal alkyl-chain substitutions are especially significant for achieving high-performance OSCs as used in thin-film transistors (TFTs).<sup>30–44</sup> In particular, it has been demonstrated recently that self-organized alkyl-chain layers play essential roles in enhancing the layered crystallinity, which is effective and useful for the formation of perfectly aligned semiconductor–insulator interfaces in TFT device structures.<sup>45–48</sup> For example, a simple alkyl-chain substitution on some  $\pi$ -electron cores, such as benzothieno[3,2-*b*]benzothiophene (BTBT), allows the formation of *bilayer-type* layered herringbone (*b*-LHB) packing motifs;<sup>45–52</sup> the alkylated unsymmetric rod-like molecules form unipolarly aligned molecular layers by their side-by-side intermolecular arrangements. Consequently, the resultant unipolar layers form an alternating antiparallel stacking alignment such that the  $\pi$ -core layers and alkyl chain layers are in head-to-head (tail-to-tail) contact within the crystal. High-precision quantum chemical calculations for the intermolecular interaction energies

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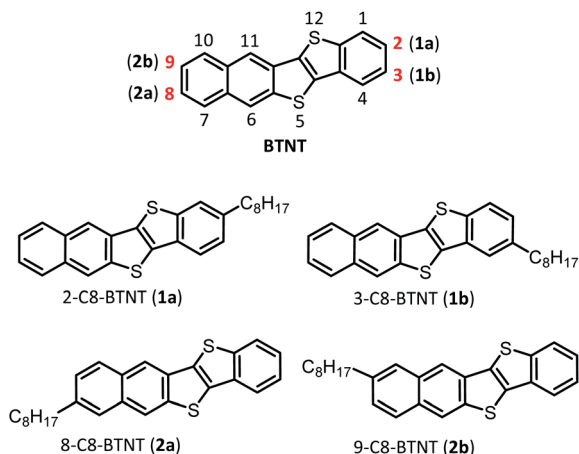


Chart 1 Chemical structures of four positional isomers of *mono*-C8-BTNTs.

revealed that the crystalline stability is considerably enforced due to the aligned alkyl-chain layer formation, which assists the  $\pi$ -core-layer formation.<sup>46,53</sup> Also, it has been demonstrated recently that alkyl-chain substitution is effective in realizing the independent layer formation of  $\pi$ -cores in crystals, even though the unsubstituted  $\pi$ -cores do not exhibit layered molecular packing motifs.<sup>53</sup>

Despite these findings, however, the controllability or designability has not been investigated yet for layered molecular packing by functional substitutions to produce high-performance OSCs. In particular, the effect and role of the substituting positions on OSCs remain unclear, whereas such regioisomerism is generally known as a powerful methodology to modify and develop organic molecules, as demonstrated in broad research fields for fine chemicals<sup>54–59</sup> and/or pharmaceuticals.<sup>60–63</sup> Such an ambiguous nature can be attributed to

the fact that the change of substituting positions may violate the layered crystallinity, which eventually prevents the uniform layer formation of OSCs and thus contributes to deteriorating the semiconducting properties of TFTs.<sup>42,64–67</sup> Nonetheless, the systematic understanding of the regioisomeric effect has not yet been achieved in OSCs.

In this study, we investigated the effect of substituting positions on the structure–property relationships of high-performance and solution-processable layered-crystalline OSCs, with a focus on an extended  $\pi$ -electron core of benzothieno[3,2-*b*]naphtho[2,3-*b*]thiophene (BTNT)<sup>68</sup> substituted simply by a normal octyl chain, which is denoted as *mono*-C8-BTNT. As the BTNT core does not have inversion or mirror symmetry, we could choose four unique (or independent) substituting positions, with roughly keeping the rod-like molecular shape over the whole molecules, as presented in Chart 1; **1a** (2-C8-BTNT) and **2a** (8-C8-BTNT) were substituted at the two most peripheral positions of the rod-like  $\pi$ -core of BTNT, while **1b** (3-C8-BTNT) and **2b** (9-C8-BTNT) were substituted at the two second most peripheral positions. We found that all four regioisomers formed an isomorphous *b*-LHB molecular packing motif that is the most suitable for TFT device structure formation. We present and discuss the crystal structure–property relationship based on the regioisomers in terms of the thermal stability, solubility, and TFT device characteristics. We showed that the *b*-LHB molecular packing motif and the materials' characteristics undergo a systematic change by varying the substituting positions.

## Results and discussion

### Material synthesis

We examined two synthetic routes as illustrated in Chart 2 to obtain *mono*-C8-BTNT regioisomers selectively. **1a** and **1b** were synthesized according to a reported procedure of BTNT

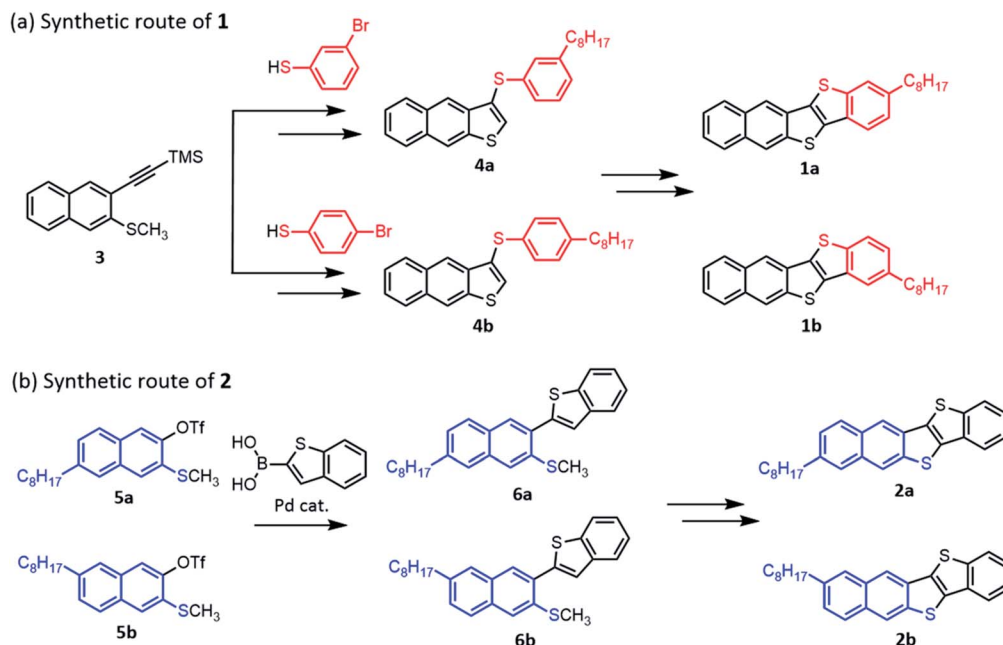


Chart 2 Two synthetic routes to *mono*-C8-BTNT regioisomers.



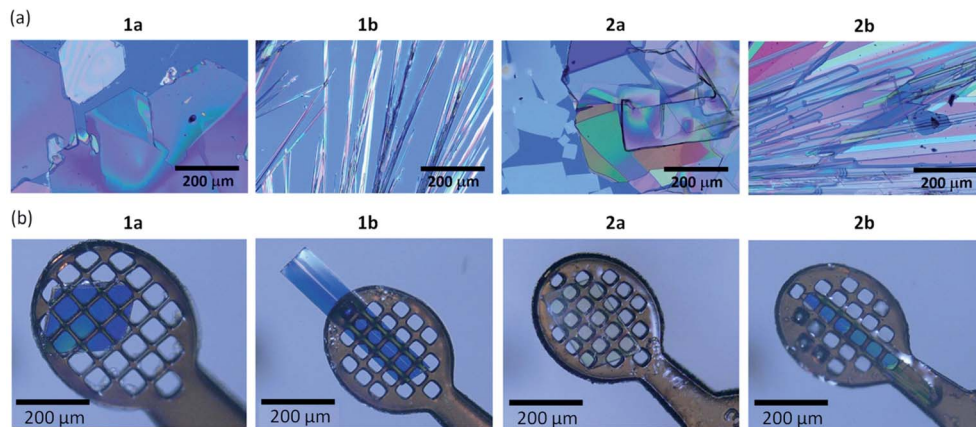


Fig. 1 Micrographs for single crystals of *mono*-C8-BTNT regioisomers. (a) Precipitated crystals grown by drop-casting of anisole solution at room temperature. (b) Single crystals for crystal structure analyses, obtained by recrystallization from anisole (**1a**, **2a**, and **2b**), and from a mixed solution of chloroform and ethanol at room temperature (**1b**).

derivatives<sup>49,68</sup> with several modifications. As shown in Chart 2(a), the key intermediates **4a** and **4b** for obtaining **1a** and **1b** were synthesized by a thiophene-annulation reaction between naphthalene derivative **3** and corresponding bromobenzene-thiols, which is followed by palladium catalyzed cross-coupling with octylboronic ester. Subsequent iodization and palladium catalyzed intramolecular coupling afforded **1a** and **1b**. As shown in Chart 2(b), the key intermediates **6a** and **6b** for obtaining **2a** and **2b** were synthesized by the Suzuki coupling reaction between benzothiopheneboronic acid and corresponding triflates **5a** and **5b**. Subsequent oxidation and intramolecular cyclization with triflic acid afforded **2a** and **2b**. The details of all synthetic procedures are described in the ESI.† The final products were purified by column chromatography and repeated

recrystallization from a mixed solution of chloroform and ethanol.

### Crystalline morphologies of regioisomers

Crystalline morphologies of all the compounds were flake (plate) like, although the obtained crystal size depends on the respective compounds, as presented in Fig. 1. The crystals of **1a** and **2a** are slightly larger and are obtained as more plate-like, while those of **1b** and **2b** are more needle-like.

### Isomorphous packing and regioisomeric effects

Views of crystal structures are shown in Fig. 2 for all four types of regioisomers of *mono*-C8-BTNTs. All the compounds exhibited an isomorphous *b*-LHB molecular packing motif with the

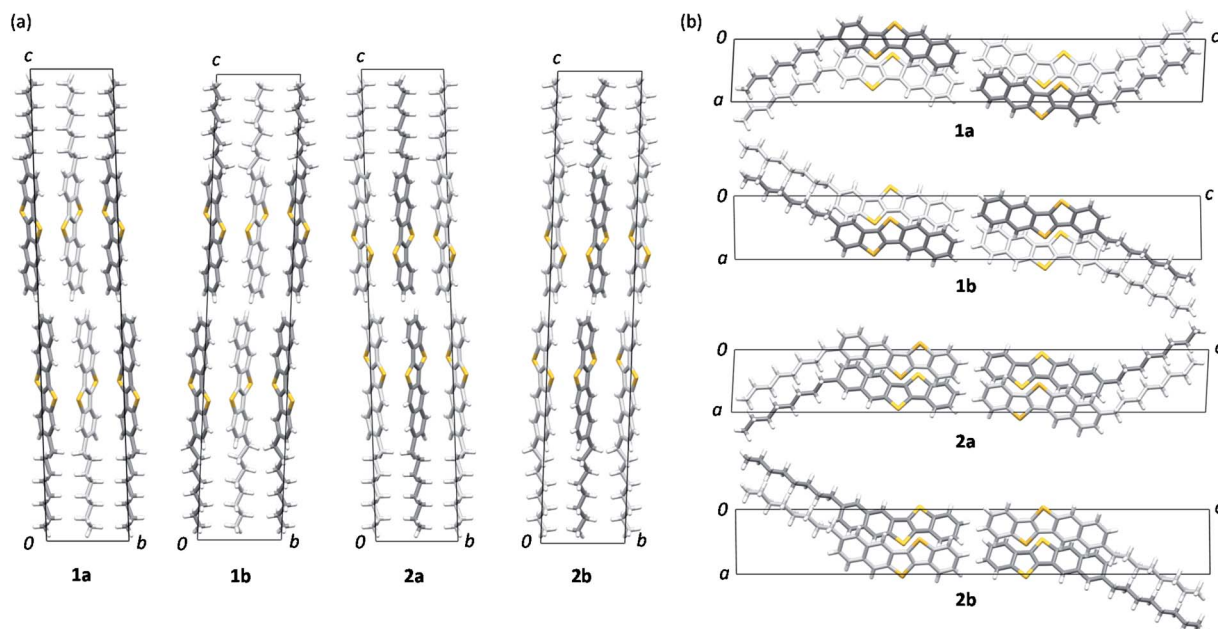
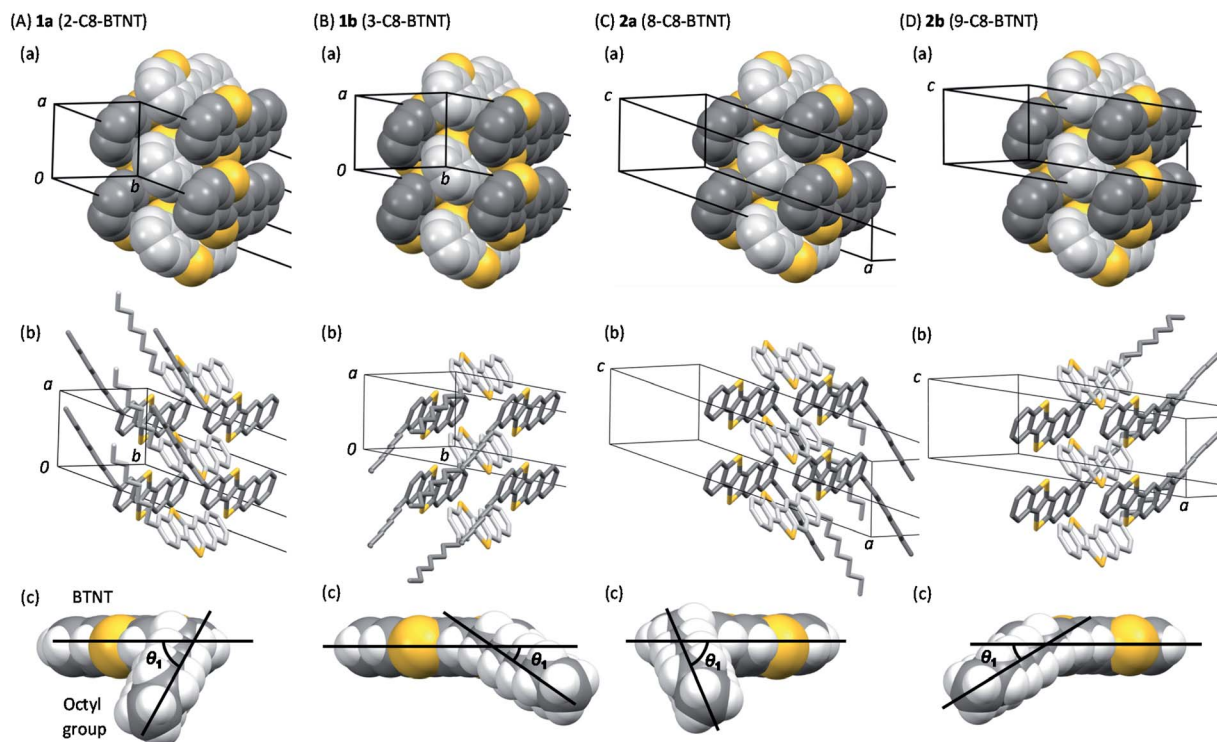


Fig. 2 Unit-cell packing diagrams of *mono*-C8-BTNT regioisomers. (a) Projection of the *b*-*c* plane. (b) Projection of the *a*-*c* plane. Crystallographically independent molecules are shown, respectively, in different colors (white and gray). Sulfur atoms are shown in yellow color.







**Fig. 3** 3D intermolecular arrangements of *mono*-C8-BTNT regioisomers. (a) Space-filling views for intralayer 3D packing of BTNT cores, and also shown by (b) stick models with alkyl chains. Each crystallographically independent molecule is shown in different colors (white and gray) in (a) and (b). (c) Space-filling views along both the BTNT core plane and the all-*trans* octyl chain plane. Definition of the dihedral angle ( $\theta_1$ ) between the planes is also shown. Sulfur atoms are shown in yellow color.

**Table 1** Crystallographic parameters and values of the intramolecular dihedral angle between the BTNT core plane and all-*trans* chain plane ( $\theta_1$ ), core–core herringbone angle ( $\theta_2$ ) and inter-core distance ( $d_2$ ), chain–chain herringbone angle ( $\theta_3$ ) and inter-chain distance ( $d_3$ )

Crystallographic parameters							Dihedral angles	Packing diagrams of BTNT cores		Alignment of octyl chains		
<i>a</i> (Å)	<i>b</i> (Å)	<i>c</i> (Å)	$\alpha$ (deg)	$\beta$ (deg)	$\gamma$ (deg)	<i>V</i> (Å <sup>3</sup> )	$\theta_1$ (degree)	$\theta_2$ (degree)	<i>d</i> <sub>2</sub> (Å)	$\theta_3$ (degree)	<i>d</i> <sub>3</sub> (Å)	
<b>1a</b>	6.01908 (11)	7.87883 (16)	45.2168 (12)	91.9646 (19)	92.9466 (18)	90.3409 (16)	2140.17 (8)	63.36	53.95	2.71	80.45	3.30
<b>1b</b>	6.0576 (3)	8.0009 (4)	44.519 (3)	87.629 (5)	88.944 (4)	89.750 (4)	2155.4 (2)	29.85	54.90	2.76	111.58	4.26
<b>2a</b>	6.00978 (14)	7.87052 (18)	45.2747 (12)	91.737 (2)	92.907 (2)	90.2170 (18)	2137.73 (9)	65.38	53.76	2.71	84.93	3.51
<b>2b</b>	6.1144 (3)	7.9480 (5)	44.660 (2)	87.923 (5)	89.367 (4)	89.762 (5)	2168.8 (2)	34.57	53.24	2.71	114.45	4.51

same space group symmetry of  $P\bar{1}$ . In the crystals, the BTNT cores formed head-to-head alignment between the unipolar layers, irrespective of the substituting positions. Thus, head-to-head contacts were observed between the naphthalene-ring sides of the BTNT cores in **1a** and **1b**, and between the benzene-ring sides in **2a** and **2b**. The details of interlayer molecular packings are shown in Fig. S1–S5.†

Fig. 3 presents the space-filling views of the intermolecular arrangements between the BTNT cores for all the regioisomer crystals. It was found that the *b*-LHB arrangements between the BTNT cores were almost the same for all the regioisomers, despite the change in the substituting positions from which the alkyl chains were sprouted out in the crystals. Meanwhile, the

substituted alkyl chains formed all-*trans* chain planes (see Fig. 3) arranged to form ordered alkyl-chain layers through the *T*-shaped and slipped parallel contacts between the chain planes. The study demonstrated that the characteristic angles formed by the core planes and chain planes present a systematic variation depending on the substituting positions, as presented in Table 1. All the inter-core herringbone angles  $\theta_2$  were roughly the same for all the compounds. In contrast, the interchain herringbone angle  $\theta_3$  and core-chain dihedral angle  $\theta_1$  differed significantly between the two groups; those of **1a** and **2a** and those of **1b** and **2b** were similar to each other, respectively, but the former was much different from the latter. The differences between the two groups are associated with the



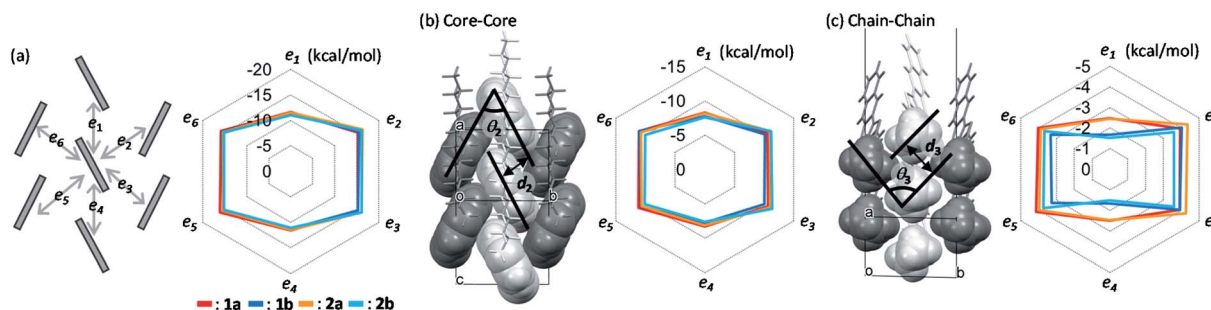


Fig. 4 Calculated intermolecular interaction energies of *mono*-C8-BTNT regioisomers. (a) Total intermolecular interaction energies between adjacent molecules at the *T*-shaped ( $e_2$ ,  $e_3$ ,  $e_5$ , and  $e_6$ ) and slipped parallel ( $e_1$  and  $e_4$ ) contacts. Scheme for the interlayer herringbone packing motif is shown on the left. (b) Contributions from the core–core interactions. Space-filling view for the intralayer core–core arrangement with the definition of the herringbone angle ( $\theta_2$ ) and inter-core distance ( $d_2$ ) is also shown on the left. (c) Contributions from the chain–chain interactions. Space-filling view for the intralayer chain–chain arrangement with the definition of the herringbone angle ( $\theta_3$ ) and inter-chain distance ( $d_3$ ) is also shown on the left.

feature of the substituting positions; those of **1a** and **2a** were located at the most peripheral positions, while those of **1b** and **2b** were located at the second most peripheral positions.

### Intermolecular interaction energies

The calculated intermolecular interaction energies between neighboring molecules in the crystals are summarized in Fig. 4. The interaction energies between  $\pi$ -electron cores are roughly the same regardless of the substituting position by the alkyl chains (Fig. 4b), while those between alkyl chains depend clearly on the substituting positions (Fig. 4c). Additionally, the size ratios of the total intermolecular interaction energies between the *T*-shaped and slipped parallel contacts were kept at approximately 3 : 2 for all the compounds, the features of which were also observed in many other LHB compounds.<sup>46,53</sup> It is most probable that the size ratio should remain the same to stabilize the LHB packing motifs.

A distinctive difference in the interchain interaction energies was observed between the two groups, as shown in Fig. 4c; **1a** and **2a** show higher interaction energies along the slipped parallel contacts than **1b** and **2b**. The result indicates that the formation of alkyl chain layers is energetically more beneficial in **1a** and **2a** than in **1b** and **2b**. Total cohesive energies,  $E_{\text{total}}$ ,

are summarized in Table 2. Because of the higher interchain ordering in **1a** and **2a** than in **1b** and **2b**, cohesive energy is larger in the former than the latter, despite the same *b*-LHB packing. Actually, the volume of **1a** and **2a** in crystal cells is smaller than that of **1b** and **2b**. The results also imply that the alkyl-chain layer formation has more degrees of freedom than the core layer formation, which is the reason why the isomorphous *b*-LHB packing is kept for all four regioisomeric compounds. The alkyl chains are known to form a number of polymorphs in terms of the interchain ordering.<sup>69–71</sup>

### Intermolecular transfer integral

Fig. 5 shows the calculated transfer integrals. The values were estimated to be 60 meV between the slipped parallel contacts for all the regioisomers. In contrast, those between the *T*-shaped contacts varied between 20–80 meV, depending on the pairs of adjacent molecules, as well as on the different regioisomers. The feature is first ascribed to the unsymmetric nature of the BTNT core that does not have inversion or mirror symmetry, leading to nonequivalent relationships on the *T*-shaped contacts ( $t_2$ ,  $t_3$ ,  $t_5$ , and  $t_6$ ) within the herringbone packing.

To investigate the origin of the above variation, we estimated the contributions of respective atomic pairs to the intermolecular transfer integrals, whose results are summarized in Fig. S6 and Table S2–S5.† Major contributions for both *T*-shaped and slipped parallel contacts are anti-phase overlaps between the atomic orbitals of sulfur and in-phase overlaps between the  $sp^2$ -orbitals of carbon located around the center of the thienothio-phenylene ring. It implies that the transfer integral is based on the sum of anti-phase and in-phase overlaps, whose balance is sensitive to the slight change of the packing motif. This feature makes us expect the discovery of OSCs, showing a much higher performance by chance in the near future.

### Solubility & thermal characteristics

All the *mono*-C8-BTNT regioisomers were soluble in some aromatic solvents such as toluene, xylene, anisole, and tetralin. However, the solubility was lower in nonaromatic polar solvents such as *N,N*-dimethylformamide, acetonitrile, and ethanol.

Table 2 Lattice energies calculated from the crystal structures of *mono*-C8-BTNT regioisomers

	Lattice energy per molecules (kcal mol <sup>−1</sup> )		
	From intermolecular interaction energy		From QE
	$E_{\text{total}}^a$	$E_{\text{core-core}}^b$	$E_{\text{chain-chain}}^c$
<b>1a</b>	−42.88	−30.02	−10.50
<b>1b</b>	−42.19	−28.76	−8.96
<b>2a</b>	−43.10	−29.95	−10.70
<b>2b</b>	−42.29	−28.97	−8.85

<sup>a</sup> Lattice energies estimated from calculated interaction energies with six neighboring molecules in crystals. <sup>b</sup> Contributions of interactions between aromatic cores. <sup>c</sup> Contributions of interactions between alkyl chains.



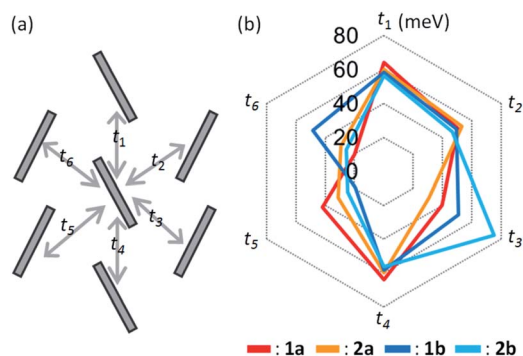


Fig. 5 Calculated intermolecular transfer integrals of *mono*-C8-BTNT regioisomers. Scheme for the intralayer herringbone packing motif is shown on the left.

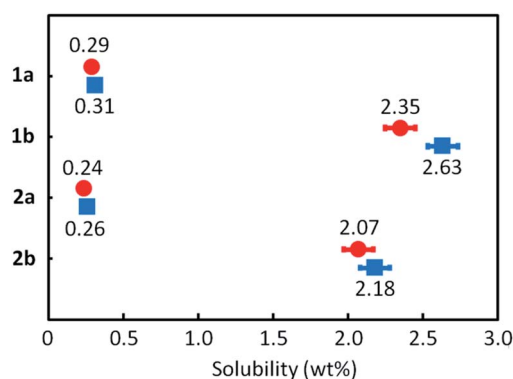


Fig. 6 Solubility of *mono*-C8-BTNT regioisomers in *o*-xylene at 20 °C determined by the concentration of saturated solution. Red filled circles show the concentration estimated by the dissolution process, and blue filled squares show the concentration estimated by the precipitation process.

Fig. 6 shows their solubilities in *o*-xylene. The solubility of **1b** (2.35 wt%) and **2b** (2.07 wt%) is much higher than that of **1a** (0.29 wt%) and **2a** (0.24 wt%). The results are consistent with the calculated intermolecular interactions, which show that the crystals of **1a** and **2a** are more stable than those of **1b** and **2b**. We also found that the solubility estimated by the different processes (dissolution and precipitation) showed a slight difference in **1b** and **2b**, but not in **1a** and **2a**. We consider that the result also indicates the lower crystallinity of **1b** and **2b** as compared to **1a** and **2a**.

DSC charts for all the regioisomers are shown in Fig. 7. Both **1a** and **2a** exhibited two exothermic and endothermic peaks, respectively, associated with the liquid-crystal phase transition at 180–190 °C and the melting transition at 221–222 °C. Notably, the results indicate that the thermal durability of both **1a** and **2a** should be high enough for use in printed electronics. In contrast, both **1b** and **2b** showed similar but lower transition temperature (melting at 160–169 °C) than **1a** and **2a**. Fig. S7† shows the total enthalpies ( $\Delta H_{\text{total}}$ ) estimated from the respective peaks of the DSC charts. We found that the  $\Delta H_{\text{total}}$  values of **1a** and **2a** are approximately the same in the heating and

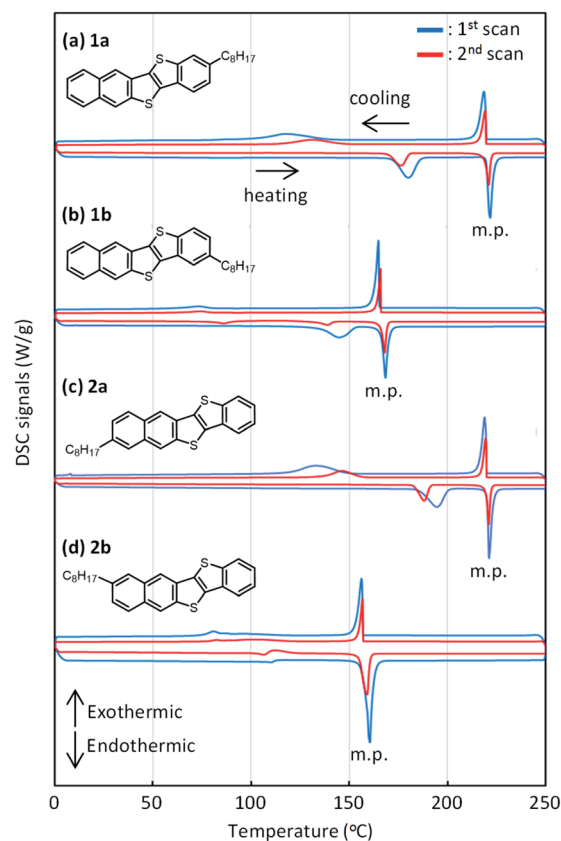


Fig. 7 DSC curves of *mono*-C8-BTNT regioisomers. Blue curves show the results for scans at 10 K min<sup>−1</sup>, and red curves show the results for 2nd scans at 5 K min<sup>−1</sup>.

cooling processes. In contrast, those of **1b** and **2b** estimated in the heating process are much larger than those estimated in the cooling process by ca. 7–10 kJ mol<sup>−1</sup>. We consider that **1b** and **2b** should show lower crystallinity or require a longer structural relaxation time for the crystallization, as compared to **1a** and **2a**.

We note here that two types of unsymmetrically substituted regioisomers were reported for Ph-BTNT-C10,<sup>49</sup> in which the phenyl group and the alkyl chain were substituted at the most peripheral (2-alkyl-8-phenyl) and at the second most peripheral (3-alkyl-9-phenyl) positions. In the compounds, a similar trend was observed in the solubility and thermal characteristics, where the latter showed much higher solubility and lower melting points than the former.

### Polycrystalline thin films & TFT characteristics

Fig. 8 shows the out-of-plane X-ray diffraction profiles of the spin-coated polycrystalline thin films treated by annealing at high temperature. We found that two types of diffraction peaks were observed for all the regioisomers, depending on the annealing conditions. The as-coated films and the films annealed at 80 °C exhibited distinctive diffraction peaks due to the *d*-spacing of 2.3 nm, which is approximately half of the lattice constant *c* (or molecular bilayer thickness), as shown in Table 1. By annealing at elevated temperatures, diffraction peaks due to the *d*-spacing of 4.5 nm, corresponding to the



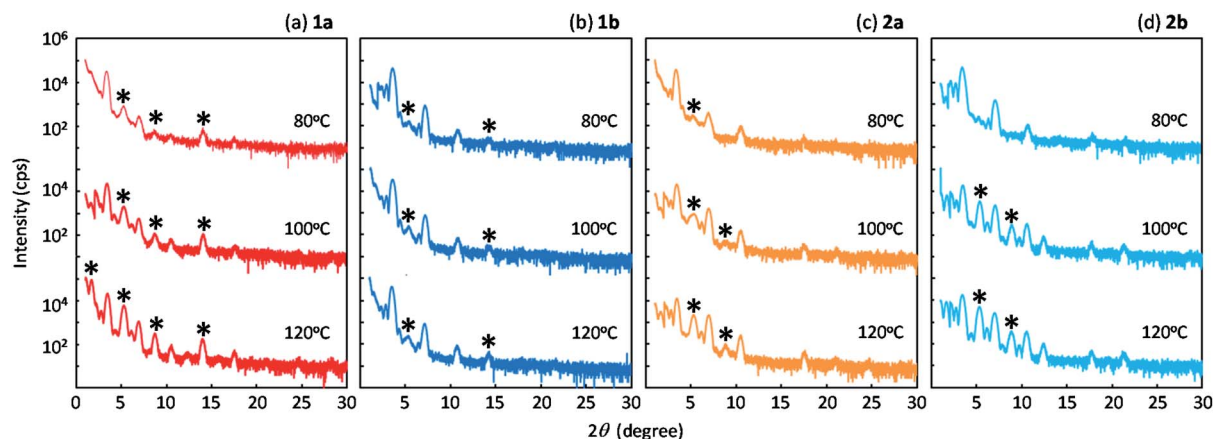


Fig. 8 Out-of-plane X-ray diffraction profiles of spin-coated films of *mono*-C8-BTNT regioisomers with thermal treatment at 80 °C, 100 °C, and 120 °C for 5 min. Diffraction peaks due to the *d*-spacing of 4.5 nm, corresponding to the molecular bilayer thickness, are shown by the asterisk.

molecular bilayer thickness, appeared and grew, as shown by the asterisks. The *d*-spacing of 2.3 nm indicates the formation of a monomolecular layer, which should be grown by rapid solidification by the spin-coating process. The growth of diffraction peaks due to annealing was most clearly observed in the films of **1a**. It is likely that the structural relaxation from the initial quenched phase to the stable *b*-LHB packing phase was promoted by thermal annealing conditions, depending on the thermal characteristics of each regioisomer. Similar annealing effects were also observed in polycrystalline thin films of Ph-BTBT-C10 and Ph-BTNT-C10.<sup>49</sup> We consider that the rapid formation of the monomolecular layer and the structural relaxation of molecular bilayer films at elevated temperatures should be the common characteristics for the molecular bilayer organic semiconductors.

Fig. 9 presents the device characteristics of TFTs based on spin-coated polycrystalline films of **1a** (the results for the other regioisomers are summarized in Fig. S8 and S9†). The field-effect mobility for both **1a** and **2a** was estimated to be more than  $1 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$  with low hysteresis and low turn-on voltage by annealing at elevated temperatures, although the measurements were conducted under ambient conditions. Table 3 summarizes the estimated values of polycrystalline TFT characteristics for all the regioisomers. Fig. 10 also show the dependence of TFT characteristics on the annealing temperature. We found that an improvement of the device mobility was observed in **1a** and **2a** by annealing at elevated temperature, but not in **1b** and **2b**. The results indicate the high thermal durability of **1a** and **2a** for practical use. In contrast, the mobility for **1b** decreased by the increase of the annealing temperature. The

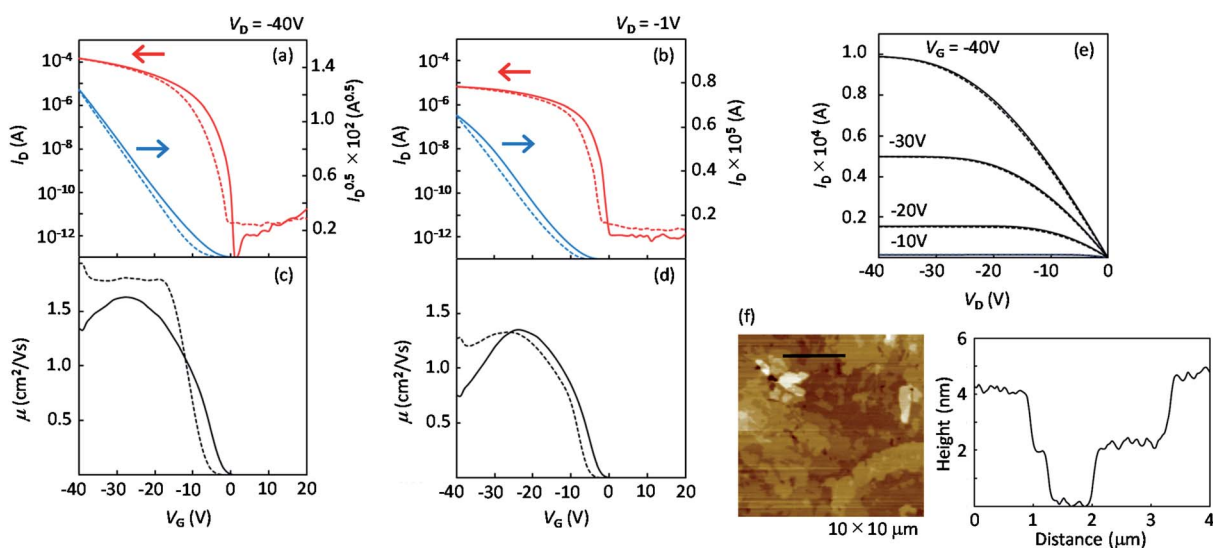


Fig. 9 Typical device characteristics of spin-coated polycrystalline TFTs of **1a**. (a) Transfer characteristics in the saturation regime and (b) in the linear regime. (c) and (d) show the plots for mobility as a function of  $V_G$ . (e) Output characteristics. Solid lines and dashed lines are forward and backward scans, respectively. (f) AFM image for the spin-coated polycrystalline-thin film of **1a** after annealing at 120 °C for 5 min.





**Table 3** Parameters of TFT characteristics after thermal treatment for spin-coated polycrystalline thin films of *mono*-C8-BTNT regioisomers. All the values were obtained as an average from more than 6 devices

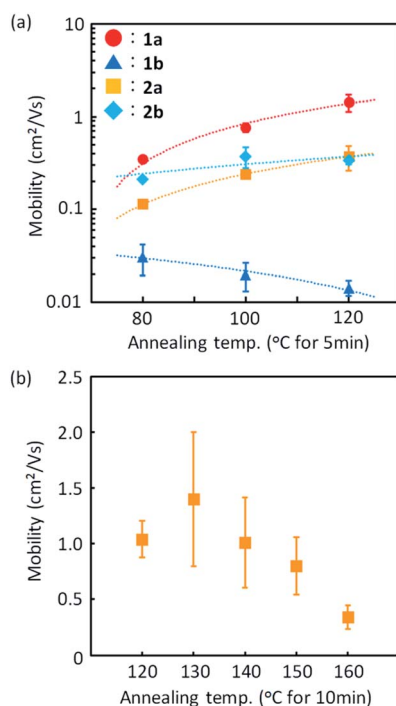
	Annealing temp. (°C)	Annealing time (min)	Mobility ( $\text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ )/standard deviation (%)		$V_{\text{th}}$ (V)		$I_{\text{on}}/I_{\text{off}}$	
			$V_{\text{D}} = -40 \text{ V}$	$V_{\text{D}} = -1 \text{ V}$	$V_{\text{D}} = -40 \text{ V}$	$V_{\text{D}} = -1 \text{ V}$	$V_{\text{D}} = -40 \text{ V}$	$V_{\text{D}} = -1 \text{ V}$
<b>1a</b>	80	5	0.39/24	0.38/26	$0.2 \pm 3.9$	$-2.9 \pm 3.2$	$10^6$	$10^6$
	100		0.77/11	0.72/9	$-4.9 \pm 1.5$	$-8.2 \pm 1.0$	$10^6$	$10^7$
	120		1.43/21	1.26/20	$-10.7 \pm 3.5$	$-11.5 \pm 2.8$	$10^7$	$10^7$
<b>1b</b>	80	5	$3.1 \times 10^{-2}/37$	$2.1 \times 10^{-2}/54$	$-1.1 \pm 0.6$	$-10.4 \pm 1.2$	$10^5$	$10^5$
	100		$2.0 \times 10^{-2}/34$	$1.0 \times 10^{-3}/20$	$-6.6 \pm 5.0$	$-13.7 \pm 0.9$	$10^5$	$10^4$
	120		$1.4 \times 10^{-2}/19$	$9.1 \times 10^{-3}/28$	$-2.2 \pm 0.4$	$-9.7 \pm 0.6$	$10^5$	$10^4$
<b>2a</b>	80	5	0.11/9	0.11/15	$-2.6 \pm 0.4$	$-5.3 \pm 0.6$	$10^5$	$10^5$
	100		0.24/11	0.22/9	$-3.5 \pm 0.6$	$-6.1 \pm 0.5$	$10^6$	$10^5$
	120		0.37/29	0.30/25	$-4.2 \pm 2.0$	$-6.8 \pm 1.3$	$10^6$	$10^6$
		10	1.04/16	0.67/8	$-12.6 \pm 1.3$	$-13.7 \pm 2.6$	$10^6$	$10^6$
		20	1.54/11	1.02/5	$-9.7 \pm 1.7$	$-8.1 \pm 1.4$	$10^6$	$10^6$
<b>2b</b>	80	5	0.21/6	0.13/18	$-0.8 \pm 2.9$	$-6.1 \pm 0.8$	$10^6$	$10^6$
	100		0.37/25	0.33/22	$-9.0 \pm 2.6$	$-12.2 \pm 1.3$	$10^6$	$10^6$
	120		0.34/10	0.33/18	$-9.2 \pm 1.3$	$-12.4 \pm 0.5$	$10^7$	$10^6$

result is consistent with the fact that the thin-film diffraction peaks, due to the *b*-LHB packing, were not observed in **1b**. It is most probable that the quenching of the polycrystalline thin film from the initial drying state to the *b*-LHB crystalline state did not proceed effectively by the annealing treatment at temperatures higher than 80 °C for **1b**. We here note that these mobility values are affected not only by the inclusion of quenched phases but also by the grain boundaries in polycrystalline thin films.

### Single-crystalline thin films & intrinsic mobility

In order to compare the intrinsic device performance between the regioisomeric compounds, it is necessary to use single-crystalline thin-film devices. By using very slow blade-coating, we successfully obtained crystalline thin films with large single-crystal domains for all the regioisomers. Optical microscope images and AFM profiles for all the films are shown in Fig. S9 and S10,† respectively. It was feasible to obtain large single-crystal domains in **1a** and **2a**, while it was much more difficult in **1b** and **2b**. As a result, the obtained domain size of **1b** and **2b** was limited and much smaller than that of **1a** and **2a**. It is clear that such a difference in the domain size should be associated with the degree of layered crystallinity in each regioisomers; **1a** and **2a** are more effectively promoted to have higher layered crystallinity by the alkyl chains than **1b** and **2b**. Note that the inter- $\pi$ -core arrangements responsible for the carrier transport are retained in the single crystals of all the regioisomers, as discussed in the former section.

The typical transfer and output characteristics of single-crystal TFTs are shown in Fig. 11 for **1a**, and in Fig. S11† for all the regioisomers. The device performance of the single-crystal TFTs are much higher than that of polycrystalline TFTs for all the *mono*-C8-BTNT regioisomers. Table 4 summarizes the estimated mobility, threshold voltage, and on/off current ratio. The intrinsic field-effect mobility of **1a** and **2a** in the saturation regime reaches as high as approximately  $10 \text{ cm}^2/\text{Vs}$ . In contrast, **1b** and **2b** show relatively lower mobility at approximately  $2.5 \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$  and  $5 \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$ , respectively, showing a large negative threshold voltage. These results imply that single-crystal thin films of **1b** and **2b** should involve a larger number of deep trap sites for carriers, associated with the lower degree of layered crystallinity, as stated above. We conclude that the closer *b*-LHB packing achieved by the substitution at the most peripheral positions is more favorable both for the crystalline stability and the semiconducting properties of *mono*-alkylated BTNTs.



**Fig. 10** Effect of thermal treatment on the TFT mobility of spin-coated films of **1a** (●), **1b** (▲), **2a** (■), and **2b** (◆) measured at a  $V_{\text{D}}$  of  $-40 \text{ V}$ . (a) Changes in TFT mobilities for the four *mono*-C8-BTNT regioisomers. (b) Changes in TFT mobilities for **2a**.





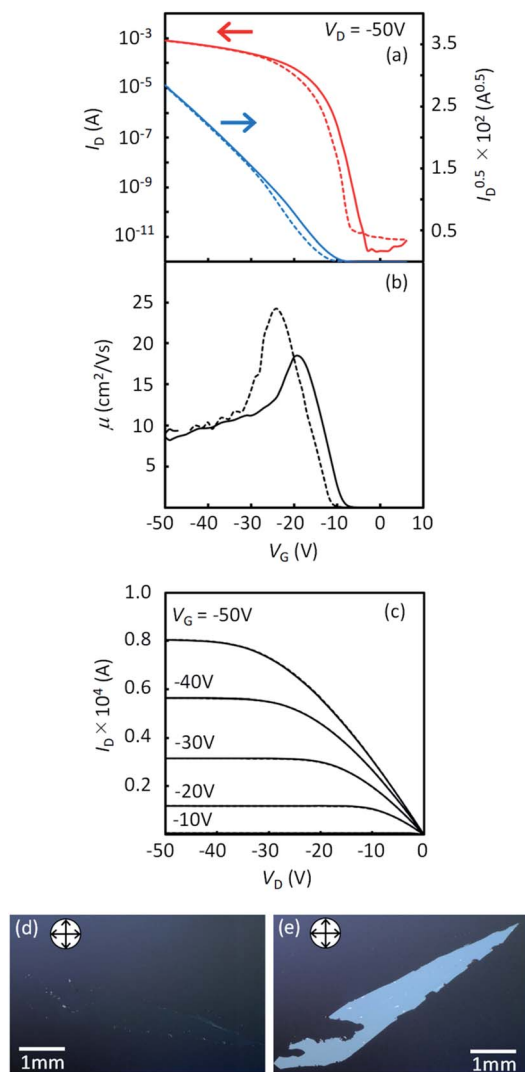


Fig. 11 Device characteristics of single-crystal TFTs with **1a**. (a) Transfer characteristics. (b) Plot of mobility as a function of  $V_G$ . (c) Output characteristics. (d), (e) Crossed-Nicols polarized micrographs of a single-crystal thin film of **1a**.

Table 4 Parameters of TFT characteristics for single crystal thin films of *mono*-C8-BTNT regioisomers

VD (V)	Mobility ( $\text{cm}^2 \text{V}^{-1} \text{s}^{-1}$ )		$V_{th}$ (V)		$I_{on}/I_{off}$	
	-50	-5	-50	-5	-50	-5
<b>1a</b>	10.3	8.3	-9.6	-12.7	$10^8$	$10^7$
<b>1b</b>	2.5	1.8	-26.2	-27.1	$10^7$	$10^7$
<b>2a</b>	10.3	2.8	-6.0	-4.7	$10^7$	$10^6$
<b>2b</b>	5.1	3.4	-35.5	-38.1	$10^7$	$10^7$

## Conclusions

We investigated the regioisomeric structure–property relationship in layered organic semiconductors based on alkyl-substituted BTNTs. We found that a series of four positional

isomers whose BTNT core is substituted by the octyl chain at either of the four most peripheral positions afford isomorphous *b*-LHB molecular packing. The intralayer herringbone arrangements of the BTNT cores were almost the same for all the regioisomers, while the all-*trans* chain planes of the substituents formed two groups of interchain ordering depending on the substituting positions within the crystals: **1a** and **2a**, substituted at the most peripheral positions, and **1b** and **2b**, substituted at the second most peripheral positions. The thermal and solubility characteristics were determined by the difference in crystalline stability between the groups of these regioisomers. The results were consistent with the calculated intermolecular interaction energies in the crystals, where the crystals of **1a** and **2a**, substituted at the most peripheral positions, are more stable than those of **1b** and **2b**. It was also found that the higher crystalline stability of **1a** and **2a** is more favorable for providing larger single-crystal domains as well as higher semiconducting properties. Especially, **1a** and **2a** exhibit higher intrinsic field-effect mobility as high as approximately  $10 \text{ cm}^2 \text{V}^{-1} \text{s}^{-1}$ . The results indicate that an additional factor of “crystallinity”, which remains chemically unexplored, is crucial for the further development of organic semiconductors. These findings also demonstrate that the regioisomeric control allows the design of practical layered organic semiconductor materials with sufficient solvent solubility and thermal stability in addition to excellent carrier transport characteristics through the enhancement of the layered crystallinity. We envision that these findings should be also applicable for various aspects of molecular engineering in broad fields of chemical science.

## Experimental

### Structural analysis

Single crystals for structural analyses were obtained by recrystallization from saturated solution in anisole at room temperature for **1a**, **2a**, and **2b**. Those for **1b** were obtained by recrystallization from a solution in a mixed solvent of chloroform and ethanol through solvent evaporation at room temperature. The obtained crystals were carefully mounted on mounting apparatuses for structural analyses (LithoLoops, Molecular Dimensions Ltd.). Single-crystal X-ray diffraction measurements were conducted by using a Rigaku AFC10 four-circle diffractometer equipped with a Pilatus 200 K hybrid pixel detector for **2a**, and a Rigaku VariMax Dual four-circle diffractometer equipped with a Pilatus 200 K for **1a**, **1b**, and **2b**. All the structural analysis calculations were conducted using the CrysAlisPro software package<sup>72</sup> and Crystal-Structure software package<sup>73</sup> (Rigaku Co., Ltd.). The structures were solved by a direct method using the SIR92 (ref. 74) and the SIR2004 program.<sup>75</sup> In the analyses, sulfur and carbon atoms were refined anisotropically, while hydrogen atoms were refined with the riding model using SHELXL97.<sup>76</sup> All the crystallographic parameters are listed in Table S1.†

### Theoretical calculation

Density functional theory (DFT) calculations for the intermolecular interaction energies between adjacent molecules were



performed using the Gaussian 16 program package,<sup>77</sup> based on the molecular packing geometries as obtained by the crystal structure analyses. The intermolecular interaction energies were calculated at the B3LYP/6-311G\*\* level with Grimme's D3 dispersion correction.<sup>78,79</sup> The basis set superposition error (BSSE)<sup>80</sup> was corrected by the counterpoise method.<sup>81</sup> The interaction energies between BTNT cores and those between alkyl chains were calculated, respectively, using the component structures fragmented in the crystals. The dangling bonds formed by the fragmentations were capped by hydrogen atoms in the calculations. The contributions of the interactions with six neighboring molecules to the lattice energy per molecule ( $E$ ) was calculated as half the sum of the interaction energies ( $E_i$ ) with the neighboring molecules as:

$$E = \frac{1}{2} \sum E_i.$$

The lattice energy was also calculated using another electronic-structure calculation package of Quantum Espresso (QE),<sup>82,83</sup> in which the crystal structures with a unit cell composed of four molecules were locally optimized at the PBE<sup>84</sup> level with Grimme's D3 dispersion correction. The lattice energies  $E$  estimated by the QE were obtained by using the following equation:<sup>85</sup>

$$E = E_{\text{cryst}} - 4E_{\text{mono}},$$

where  $E_{\text{cryst}}$  is the energy of crystal structures with a unit cell composed of four molecules, and  $E_{\text{mono}}$  is the energy of the isolated molecule.

Intermolecular electron transfer integrals were also calculated by DFT calculations using the Gaussian 09 program package<sup>86</sup> at the B3LYP/6-31G\* level, based on the molecular packing geometries as obtained by the crystal structure analyses. The transfer integral  $t_{AB}$  between the molecular orbitals  $|A\rangle$  and  $|B\rangle$  was calculated by using the following equation:<sup>87</sup>

$$t_{AB} = \frac{\langle A|F|B\rangle - \frac{1}{2}(\langle A|F|A\rangle + \langle B|F|B\rangle)\langle A|S|B\rangle}{1 - \langle A|S|B\rangle^2}.$$

where  $F$  is the Fock matrix, and  $S$  is the overlap matrix. Considering that the molecular orbitals can be represented as the linear combination of atomic orbitals  $|i\rangle, |j\rangle$  as:

$$|A\rangle = \sum_i a_i |i\rangle, \quad |B\rangle = \sum_j b_j |j\rangle,$$

the transfer integral  $t_{AB}$  was decomposed into the contributions of all atomic orbital pairs  $t_{ij}$  as:

$$t_{AB} = \sum_{i,j} a_i^* b_j \frac{\langle i|F|j\rangle - \frac{1}{2}(\langle i|F|i\rangle + \langle j|F|j\rangle)\langle i|S|j\rangle}{1 - \langle i|S|j\rangle^2} \equiv \sum_{i,j} t_{ij}.$$

### Thermal characteristics

The powdered samples were thermally analyzed by differential scanning calorimetry (DSC; DSC7000X, Hitachi High-Tech Science Co.) at a scan rate of 5–10 K min<sup>−1</sup>. The measured temperature was calibrated using the melting point of indium

(429.8 K), and synthetic sapphire was used as the standard for determining the heat capacity. For the DSC measurements, the powdered sample was heated and subsequently cooled at rates of 10 K min<sup>−1</sup> (first scan) and 5 K min<sup>−1</sup> (second scan). The melting points of all the compounds were determined by visual inspection of the changes in powdered samples on a hot plate. Transition enthalpies ( $\Delta H$ ) were estimated by the integration of respective DSC peaks. All the results are summarized in Tables S6 and S7.†

### Solubility

The solubility of the respective materials was estimated by the following two methods. The first one is the estimation of the solubility by inspecting the dissolution process at room temperature. Organic solvents were added to a powdered sample (3–5 mg) in increments of 20  $\mu$ L using a micropipette, and the mixture was stirred to dissolve the sample completely at 20 °C. The complete dissolutions in organic solvents were confirmed by visual inspection. The second method is the estimation of the solubility by inspecting the precipitation process with decreasing the temperature: a number of solution vials with different concentrations were prepared at 80 °C, and subsequently cooled down to 20 °C with continuous stirring for 12 h. The solubility was calculated from the sample weight and the total volume of the solvent.

### Thin-film fabrication and characterization

We prepared two types of thin films of *mono*-C8-BTNT; polycrystalline thin films were fabricated by spin-coating, and single-crystal thin films were fabricated by blade coating. In the fabrication of all the spin-coated films, 0.4 wt% semiconductor solutions in *o*-xylene were used. The solution membranes were obtained by short-term spin-coating (1000 rpm/3 s), and subsequently annealed to produce polycrystalline semiconductor layers by the post-evaporation of solvents. The thickness of the film was *ca.* 15–20 nm, as was estimated by AFM measurements. In the fabrication of single-crystal films, 0.1 wt% semiconductor solutions in chlorobenzene for **1a**, **2a**, and **2b** and 0.15 wt% of anisole for **1b** were used for blade-coating at sweep rates of 3.0  $\mu$ m s<sup>−1</sup> for **1a** and **2a** and 1.5  $\mu$ m s<sup>−1</sup> for **1b** and **2b** at room temperature (23 °C). Optical microscope images were collected by using a digital microscope (VHX-6000; Keyence Co., Ltd.).

The measurements of out-of-plane X-ray diffraction profiles were carried out with a thin-film diffractometer (SmartLab; Rigaku Co., Ltd.) by using a monochromatized synchrotron radiated X-ray beam with an energy of 9.0 keV at the beamlines BL-7C of Photon Factory (PF) in KEK. Diffraction intensity was recorded using a scintillation counter. The film thickness and height profiles of the films were measured by tapping mode AFM (VN-8010; Keyence Co., Ltd. and Cypher; Asylum Research Inc., USA).

### Device fabrication and characterization

Two types of organic TFTs based on the *mono*-C8-BTNT thin film were fabricated; one was bottom-gate, top-contact TFTs



composed of spin-coated polycrystalline thin films, and the other was bottom-gate, top-contact TFTs composed of blade-coated single-crystal thin films. We used highly doped (p<sup>+</sup>)-Si wafers with 100 nm-thick silicon dioxide layers as substrates for all the devices. Source/drain electrodes were fabricated by vacuum deposition of Au through a shadow mask to define the length (*L*) and width (*W*) of all the TFT channels, respectively, at 100 and 500 μm for polycrystalline thin films, at 200 and 500 μm for single-crystal thin films of **1a** and **2a**, and at 100 and 300 μm for single-crystal thin films of **1b** and **2b**. A micromanipulator (AxisPro; Systems Engineering Inc.) was used to trim away the films outside the channels for the proper evaluation of the device mobility.

The OTFT characteristics were measured using a semiconductor parametric analyzer (E5270A; Agilent Technologies Co. Ltd., B2912A; Keysight Technologies Inc.) under ambient conditions. The field-effect mobility ( $\mu$ ) was defined as the derivative of the transfer curve according to the following equations:

$$\text{Linear regime } \mu_{\text{linear}} = \frac{L}{WC_i V_D} \left( \frac{\partial I_D}{\partial V_G} \right)$$

$$\text{Saturation regime } \mu_{\text{saturation}} = \frac{2L}{WC_i} \left( \frac{\partial \sqrt{I_D}}{\partial V_G} \right)^2$$

where  $I_D$ ,  $C_i$ ,  $V_D$ , and  $V_G$  are the drain current, gate capacitance per unit area, drain voltage, and gate voltage, respectively. The measurements were conducted at a  $V_D$  of −1 V in the linear regime and at a  $V_D$  of −40 V in the saturation regime for polycrystalline TFTs and at a  $V_D$  of −5 V in the linear regime and at a  $V_D$  of −50 V in the saturation regime for single-crystal TFTs.

## Conflicts of interest

There are no conflicts to declare.

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## Notes and references

- J. Mei, Y. Diao, A. L. Appleton, L. Fang and Z. Bao, *J. Am. Chem. Soc.*, 2013, **135**, 6724–6746.
- H. Sirringhaus, *Adv. Mater.*, 2014, **26**, 1319–1335.

- B. Kang, F. Ge, L. Qiu and K. Cho, *Adv. Electron. Mater.*, 2017, **3**, 1600240.
- Q. Yue, W. Liu and X. Zhu, *J. Am. Chem. Soc.*, 2020, **142**, 11613–11628.
- L. Dou, J. You, Z. Hong, Z. Xu, G. Li, R. A. Street and Y. Yang, *Adv. Mater.*, 2013, **25**, 6642–6671.
- J. Mei and Z. Bao, *Chem. Mater.*, 2014, **26**, 604–615.
- T. Lei, J. Wang and J. Pei, *Chem. Mater.*, 2014, **26**, 594–603.
- V. Savikhin, M. Babics, M. Neophytou, S. Liu, S. D. Oosterhout, H. Yan, X. Gu, P. M. Beaujuge and M. F. Toney, *Chem. Mater.*, 2018, **30**, 7872–7884.
- M. Suzuki, K. Terai, C. Quinton, H. Hayashi, N. Aratani and H. Yamada, *Chem. Sci.*, 2020, **11**, 1825–1831.
- H. Yu, H. N. Kim, I. Song, Y. H. Ha, H. Ahn, J. H. Oh and Y.-H. Kim, *J. Mater. Chem. C*, 2017, **5**, 3616–3622.
- T. He, P. Leowanawat, C. Burschka, V. Stepanenko, M. Stole and F. Würthner, *Adv. Mater.*, 2018, **30**, 1804032.
- F. Wang, K. Nakano, H. Segawa and K. Tajima, *Chem. Mater.*, 2019, **31**, 2097–2105.
- T.-T. Do, Y. Takeda, S. Manzhos, J. Bell, S. Tokito and P. Sonar, *J. Mater. Chem. C*, 2018, **6**, 3774–3786.
- F. Zhang, Y. Hu, T. Schuettfort, C. Di, X. Gao, C. R. McNeill, L. Thomsen, S. C. B. Mannsfeld, W. Yuan, H. Sirringhaus and D. Zhu, *J. Am. Chem. Soc.*, 2013, **135**, 2338–2349.
- M. Sawamoto, M. J. Kang, E. Miyazaki, H. Sugino, I. Osaka and K. Takimiya, *ACS Appl. Mater. Interfaces*, 2016, **8**, 3810–3824.
- A. J. Petty, Q. Ai, J. C. Sorli, H. F. Haneef, G. E. Purdum, A. Boehm, D. B. Granger, K. Gu, C. P. L. Rubinger, S. R. Parkin, K. R. Graham, O. D. Jurchescu, Y.-L. Loo, C. Risko and J. E. Anthony, *Chem. Sci.*, 2019, **10**, 10543–10549.
- L. Zhang, A. Fonari, Y. Liu, A. L. M. Hoyt, H. Lee, D. Granger, S. Parkin, T. P. Russell, J. E. Anthony, J. E. J. L. Brédas, V. Coropceanu and A. L. Briseno, *J. Am. Chem. Soc.*, 2014, **136**, 9248–9251.
- J. Smith, W. Zhang, R. Sougrat, K. Zhao, R. Li, D. Cha, A. Amassian, M. Heeney, I. McCulloch and T. D. Anthopoulos, *Adv. Mater.*, 2012, **24**, 2441–2446.
- G. Giri, S. Park, M. Vosgueritchian, M. M. Shulaker and Z. Bao, *Adv. Mater.*, 2014, **26**, 487–493.
- M. H. Lee, M. Kang, H. G. Jeong, J. J. Park, K. Hwang and D. Y. Kim, *Macromol. Rapid Commun.*, 2018, **39**, 1800431.
- J. Ma, K. Hashimoto, T. Koganezawa and K. Tajima, *J. Am. Chem. Soc.*, 2013, **135**, 9644–9647.
- M. Chikamatsu, A. Itakura, Y. Yoshida, R. Azumi and K. Yase, *Chem. Mater.*, 2008, **20**, 7365–7367.
- M. Mamada, H. Shima, Y. Yoneda, T. Shimano, N. Yamada, K. Kakita, T. Machida, Y. Tanaka, S. Aotsuka, D. Kumaki and S. Tokito, *Chem. Mater.*, 2015, **27**, 141–147.
- B. J. Jung, N. J. Tremblay, M. L. Yeh and H. E. Katz, *Chem. Mater.*, 2011, **23**, 568–582.
- Y. Wang, E. Zeglio, H. Liao, J. Xu, F. Liu, Z. Li, I. P. Maria, D. Mawad, A. Herland, I. McCulloch and W. Yue, *Chem. Mater.*, 2019, **31**, 9797–9806.



- 26 Z. P. Fan, X. Y. Li, X. E. Luo, X. Fei, B. Sun, L. C. Chen, Z. F. Shi, C. L. Sun, X. Shao and H. L. Zhang, *Adv. Funct. Mater.*, 2017, **27**, 1702318.
- 27 Y. Suzuki, E. Miyazaki and K. Takimiya, *J. Am. Chem. Soc.*, 2010, **132**, 10453–10466.
- 28 K. Kawabata, S. Usui and K. Takimiya, *J. Org. Chem.*, 2020, **85**, 195–206.
- 29 C. Wang, D. Hashizume, M. Nakano, T. Ogaki, H. Takenaka, K. Kawabata and K. Takimiya, *Chem. Sci.*, 2020, **11**, 1573–1580.
- 30 H. Ebata, T. Izawa, E. Miyazaki, K. Takimiya, M. Ikeda, H. Kuwabara and T. Yui, *J. Am. Chem. Soc.*, 2007, **129**, 15732–15733.
- 31 T. Izawa, E. Miyazaki and K. Takimiya, *Adv. Mater.*, 2008, **20**, 3388–3392.
- 32 A. Y. Amin, A. Khassanov, K. Reuter, T. Meyer-Friedrichsen and M. Halik, *J. Am. Chem. Soc.*, 2012, **134**, 16548–16550.
- 33 K. Nagayama, Y. Hirose, J. Soeda, M. Yoshizumi, T. Uemura, M. Uno, W. Li, M. J. Kang, M. Yamagishi, Y. Okada, E. Miyazaki, Y. Nakazawa, A. Nakao, K. Takimiya and J. Takeya, *Adv. Mater.*, 2011, **23**, 1626–1629.
- 34 H. Minemawari, T. Yamada, H. Matsui, J. Tsutsumi, S. Haas, R. Chiba, R. Kumai and T. Hasegawa, *Nature*, 2011, **475**, 364–367.
- 35 R. Janneck, N. Pilet, S. P. Bommanaboyena, B. Watts, P. Heremans, J. Genoe and C. Rolin, *Adv. Mater.*, 2017, **29**, 1703864.
- 36 T. Hamai, S. Arai, H. Minemawari, S. Inoue, R. Kumai and T. Hasegawa, *Phys. Rev. Appl.*, 2017, **8**, 054011.
- 37 H. Y. Chen, G. Schweicher, M. Planells, S. M. Ryno, K. Broch, A. J. P. White, D. Simatos, M. Little, C. Jellet, S. J. Cryer, A. Marks, M. Hurhangee, J. L. Brédas, H. Sirringhaus and I. McCulloch, *Chem. Mater.*, 2018, **30**, 7587–7592.
- 38 T. Oyama, Y. S. Yang, K. Matsuo and T. Yasuda, *Chem. Commun.*, 2017, **53**, 3814–3817.
- 39 S. Arai, S. Inoue, T. Hamai, R. Kumai and T. Hasegawa, *Adv. Mater.*, 2018, **30**, 1707256.
- 40 H. Iino, T. Usui and J. Hanna, *Nat. Commun.*, 2015, **6**, 6828.
- 41 T. Okamoto, C. P. Yu, C. Mitsui, M. Yamagishi, H. Ishii and J. Takeya, *J. Am. Chem. Soc.*, 2020, **142**, 9083–9096.
- 42 Y. Tsutsui, G. Schweicher, B. Chattopadhyay, T. Sakurai, J. B. Arlin, C. Ruzié, A. Aliev, A. Ciesielski, S. Colella, A. R. Kennedy, V. Lemaure, Y. Olivier, R. Hadji, L. Sanguinet, F. Castet, S. Osella, D. Dudenko, D. Beljonne, J. Cornil, P. Samori and Y. H. Geerts, *Adv. Mater.*, 2016, **28**, 7106–7114.
- 43 Y. Ogawa, K. Yamamoto, C. Miura, S. Tamura, M. Saito, M. Mamada, D. Kumaki, S. Tokito and H. Katagiri, *ACS Appl. Mater. Interfaces*, 2017, **9**, 9902–9909.
- 44 K. He, S. Zhou, W. Li, H. Tian, Q. Tang, J. Zhang, D. Yan, Y. Geng and F. Wang, *J. Mater. Chem. C*, 2019, **7**, 3656–3664.
- 45 S. Inoue, H. Minemawari, J. Tsutsumi, M. Chikamatsu, T. Yamada, S. Horiuchi, M. Tanaka, R. Kumai, M. Yoneya and T. Hasegawa, *Chem. Mater.*, 2015, **27**, 3809–3812.
- 46 H. Minemawari, M. Yanaka, S. Tsuzuki, S. Inoue, T. Yamada, R. Kumai, Y. Shimoi and T. Hasegawa, *Chem. Mater.*, 2017, **29**, 1245–1254.
- 47 S. Arai, K. Morita, J. Tsutsumi, S. Inoue, M. Tanaka and T. Hasegawa, *Adv. Funct. Mater.*, 2019, 1906406.
- 48 M. Yoneya, H. Minemawari, T. Yamada and T. Hasegawa, *J. Phys. Chem. C*, 2017, **121**, 8796–8803.
- 49 S. Inoue, S. Shinamura, Y. Sadamitsu, S. Arai, S. Horiuchi, M. Yoneya, K. Takimiya and T. Hasegawa, *Chem. Mater.*, 2018, **30**, 5050–5060.
- 50 T. Higashino, S. Inoue, Y. Sadamitsu, S. Arai, S. Horiuchi and T. Hasegawa, *Chem. Lett.*, 2019, **48**, 453–456.
- 51 T. Hamai, S. Inoue, S. Arai and T. Hasegawa, *Phys. Rev. Mater.*, 2020, **4**, 074601.
- 52 S. Guo, Y. He, I. Murtaza, J. Tan, J. Pan, Y. Guo, Y. Zhu, Y. He and H. Meng, *Org. Electron.*, 2018, **56**, 68–75.
- 53 T. Higashino, S. Arai, S. Inoue, S. Tsuzuki, Y. Shimoi, S. Horiuchi, T. Hasegawa and R. Azumi, *CrystEngComm*, 2020, **22**, 3618–3626.
- 54 K. Burgess and A. Loudet, *Chem. Rev.*, 2007, **107**, 4891–4932.
- 55 M. Romain, D. Tondelier, J. C. Vanel, B. Geffroy, O. Jeannin, J. Rault-Berthelot, R. Métivier and C. Poriol, *Angew. Chem., Int. Ed.*, 2013, **52**, 14147–14151.
- 56 L. Sicard, C. Quinton, J. D. Peltier, D. Tondelier, B. Geffroy, U. Biapo, R. Métivier, O. Jeannin, J. Rault-Berthelot and C. Poriol, *Chem.–Eur. J.*, 2017, **23**, 7719–7727.
- 57 M. G. Freire, C. M. S. S. Neves, K. Shimizu, C. E. S. Bernardes, I. M. Marrucho, J. A. P. Coutinho, J. N. C. Lopes and L. P. N. Rebelo, *J. Phys. Chem. B*, 2010, **114**, 15925–15934.
- 58 Q. Lu, G. K. Kole, A. Friedrich, K. Müller-Büschbaum, Z. Liu, X. Yu and T. B. Marder, *J. Org. Chem.*, 2020, **85**, 4256–4266.
- 59 C. Poriol, L. Sicard and J. Rault-Berthelot, *Chem. Commun.*, 2019, **55**, 14238–14254.
- 60 J. S. Carey, D. Laffan, C. Thomson and M. T. Williams, *Org. Biomol. Chem.*, 2006, **4**, 2337–2347.
- 61 T. Nishida, H. Ida, Y. Kuninobu and M. Kanai, *Nat. Commun.*, 2013, **5**, 3387.
- 62 L. N. Le, J. M. R. Moore, J. Ouyang, X. Chen, M. D. H. Nguyen and W. J. Galush, *Anal. Chem.*, 2012, **84**, 7479–7486.
- 63 S. Hagihara, L. Gremaud, G. Bollot, J. Mareda and S. Matile, *J. Am. Chem. Soc.*, 2008, **130**, 4347–4351.
- 64 M. J. Kang, T. Yamamoto, S. Shinamura, E. Miyazaki and K. Takimiya, *Chem. Sci.*, 2010, **1**, 179–183.
- 65 M. J. Kang, E. Miyazaki, I. Osaka and K. Takimiya, *Jpn. J. Appl. Phys.*, 2012, **51**, 11PD04.
- 66 J. T. E. Quinn, C. Guo, F. Haider, H. Patel, D. A. Khan and Y. Li, *J. Mater. Chem. C*, 2017, **5**, 5902–5909.
- 67 H. S. Ryu, M. J. Kim, Y. W. Lee, S. H. Lee, T. J. Shin, J. H. Cho and H. Y. Woo, *ACS Appl. Mater. Interfaces*, 2019, **11**, 47170–47181.
- 68 T. Mori, T. Nishimura, T. Yamamoto, I. Doi, E. Miyazaki, I. Osaka and K. Takimiya, *J. Am. Chem. Soc.*, 2013, **135**, 13900–13913.
- 69 F. Kaneko, K. Yamazaki, K. Kitagawa, T. Kikyo and M. Kobayashi, *J. Phys. Chem. B*, 1997, **101**, 1803–1809.
- 70 P. Espeau and R. Céolin, *J. Phys. Chem. B*, 2008, **112**, 2063–2069.
- 71 L. Bayés-García, T. Calvet, M. Á. Cuevas-Diarte, S. Ueno and K. Sato, *CrystEngComm*, 2013, **15**, 302–314.





- 72 CrysAlisPro: Data Collection and Processing Software; Rigaku Corp., 2015.
- 73 Crystalstructure 4.3: Crystal Structure Analysis Package; Rigaku Corp., 2000–2018.
- 74 A. Altomare, G. Cascarano, C. Giacovazzo and A. Guagliardi, *J. Appl. Crystallogr.*, 1993, **26**, 343–350.
- 75 M. C. Burla, R. Caliendo, M. Camalli, B. Carrozzini, G. L. Cascarano, L. De Caro, C. Giacovazzo, G. Polidori and R. Spagna, *J. Appl. Crystallogr.*, 2005, **38**, 381–388.
- 76 G. M. A. Sheldrick, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 2008, **64**, 112–122.
- 77 M. J. Frisch, *et al.*, *Gaussian 16*, Gaussian, Inc., Wallingford, CT, 2016.
- 78 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 79 A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648–5652.
- 80 B. J. Ransil, *J. Chem. Phys.*, 1961, **34**, 2109–2118.
- 81 S. F. Boys and F. Bernardi, *Mol. Phys.*, 1970, **19**, 553–556.
- 82 P. Giannozzi, *et al.*, *J. Phys. Condens. Matter*, 2009, **21**, 395502.
- 83 P. Giannozzi, *et al.*, *J. Phys. Condens. Matter*, 2017, **29**, 465901.
- 84 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 85 The structures of crystals and isolated molecules were optimized with a kinetic energy cutoff of 36 Ry. The energies of crystals and isolated molecules were calculated with a kinetic energy cutoff of 49 Ry.
- 86 M. J. Frisch, *et al.*, *Gaussian 09 Revision A. 02*, Gaussian, Inc., Wallingford, CT, 2016.
- 87 V. Coropceanu, J. Cornil, D. A. da Silva Filho, Y. Olivier, R. Silbey and J. L. Brédas, *Chem. Rev.*, 2007, **107**, 926–952.

