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3-Aminooxetanes: versatile 1,3-amphoteric molecules for intermolecular annulation reactions†

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Despite the versatility of amphoteric molecules, stable and easily accessible ones are still limitedly known. As a result, the discovery of new amphoteric reactivity remains highly desirable. Herein we introduce 3-aminooxetanes as a new family of stable and readily available 1,3-amphoteric molecules and systematically demonstrated their amphoteric reactivity toward polarized π -systems in a diverse range of intermolecular [3 + 2] annulations. These reactions not only enrich the reactivity of oxetanes, but also provide convergent access to valuable heterocycles.

Amphoteric molecules, which bear both nucleophilic and electrophilic sites with orthogonal reactivity, represent an attractive platform for the development of chemoselective transformations.¹ For example, isocyanides are well-established 1,1-amphoteric molecules, with the terminal carbon being both nucleophilic and electrophilic, and this feature has enabled their exceptional reactivity in numerous multi-component reactions.² In the past few decades, substantial effort has been devoted to the search for new amphoteric molecules.^{1–5} Among them, 1,3-amphoteric molecules proved to be versatile. The Yudin and Beauchemin laboratories have independently developed two types of such molecules, α -aziridine aldehydes and amino isocyanates, respectively.^{4,5} With an electrophilic carbon and a nucleophilic nitrogen in relative 1,3-positions, these molecules are particularly useful for the chemoselective synthesis of heterocycles with high bond-forming efficiency without protective groups (Fig. 1). However, such elegant amphoteric systems still remain scarce. Therefore, the development of new stable amphoteric molecules with easy access remains highly desirable.

In this context, herein we introduce 3-aminooxetanes as a new type of 1,3-amphoteric molecules and systematically demonstrate their reactivity in a range of [3 + 2] annulations, providing rapid access to diverse heterocycles. Notably, 3-aminooxetanes are bench-stable and either commercially available or easily accessible. However, their amphoteric reactivity has not been appreciated previously.

Oxetane is a useful functional group in both drug discovery and organic synthesis.^{6–9} Owing to the ring strain, it is prone to nucleophilic ring-opening, in which it serves as an electrophile (Scheme 1A).^{6–8} We envisioned that, if a nucleophilic group is installed in the 3-position (e.g., amino group), such molecules should exhibit 1,3-amphoteric reactivity due to the presence of both nucleophilic and electrophilic sites (Scheme 1B). Importantly, the 1,3-relative position is crucial for inhibiting self-destructive intra- or intermolecular ring-opening (i.e. the 3-nucleophilic site attack on oxetane itself) due to high barriers. Thus, such orthogonality is beneficial to their stability. In contrast, the nucleophilic site is expected to react with an external polarized π bond (e.g., X = Y, Scheme 1B), which enables a better-positioned nucleophile (Y) to attack the oxetane and cyclize. Thus, a formal [3 + 2] annulation should be expected. Unlike the well-known S_N2 reactivity of oxetanes with simple bond formation, this amphoteric reactivity would greatly enrich the chemistry of oxetanes with multiple bond formations and provide expedient access to various heterocycles. In contrast to the conventional approaches that require presynthesis of advanced intermediates (e.g., intramolecular ring-opening),⁸ the exploitation of such amphoteric reactivity in an intermolecular convergent manner from simple substrates

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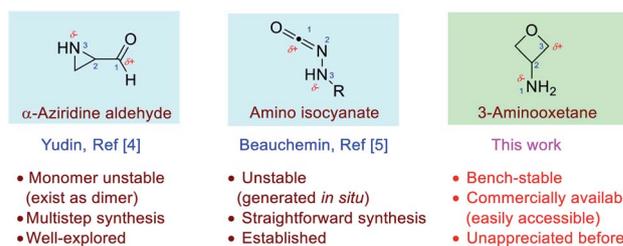
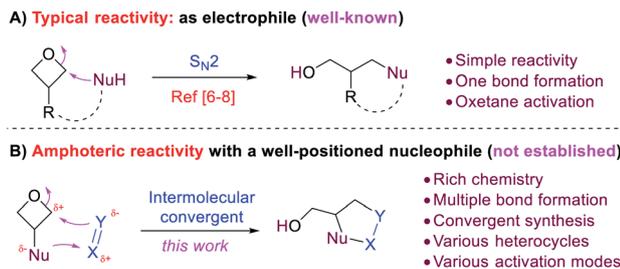


Fig. 1 Representative [1,3]-amphoteric molecules versus 3-aminooxetanes.





Scheme 1 Typical oxetane reactivity and the new amphoteric reactivity.

would be more practically useful. Moreover, more activation modes could be envisioned in addition to oxetane activation. In 2015, Kleij and coworkers reported an example of cyclization between 3-aminooxetane and CO₂ in 55% yield, which provided a pioneering precedent.¹⁰ However, a systematic study to fully reveal such amphoteric reactivity in a broad context remains unknown in the literature.

To test our hypothesis, we began with the commercially available 3-aminooxetanes **1a** and **1b** as the model substrates. Phenyl thioisocyanate **2a** and CS₂ were initially employed as reaction partners, as they both have a polarized C=S bond as well as a relatively strong sulfur nucleophilic motif. Moreover, the resulting desired products, iminothiazolidines and mercaptothiazolidines, are both heterocycles with important biological applications (Fig. 2).¹¹ To our delight, simple mixing these two types of reactants in DCM resulted in spontaneous reactions at room temperature without any catalyst. The corresponding [3 + 2] annulation products iminothiazolidine **3a** and mercaptothiazolidine **4a** were both formed with excellent efficiency (Scheme 2). It is worth mentioning that catalyst-free ring-opening of an oxetane ring is rarely known, particularly for intermolecular reactions.⁶⁻⁹ In this case, the high efficiency is likely attributed to the suitable choice and perfect position of the *in situ* generated sulfur nucleophile.

The catalyst-free annulation protocol is general with respect to various 3-aminooxetanes and isothiocyanates. A range of iminothiazolidines and mercaptothiazolidines were synthesized with high efficiency under mild conditions (Scheme 3). Many of them were obtained in quantitative yield. Quaternary carbon centers could also be generated from 3-substituted 3-aminooxetanes (*e.g.*, **3j**). The structure of product **3b** was unambiguously confirmed by X-ray crystallography.

With the initial success of thiocarbonyl partners, we next turned our attention to isocyanates, in which the carbonyl group serves as the [3 + 2] annulation motif. Compared with

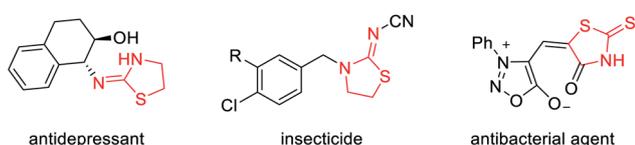
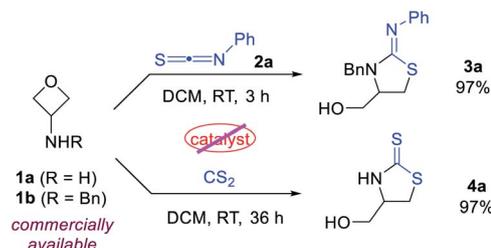
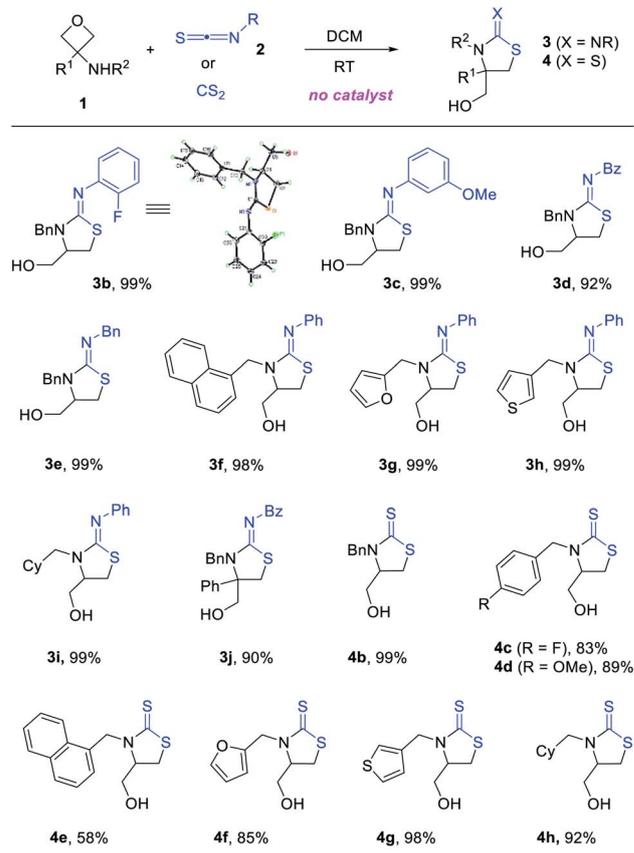


Fig. 2 Selected bioactive molecules containing iminothiazolidine and mercaptothiazolidine motifs.



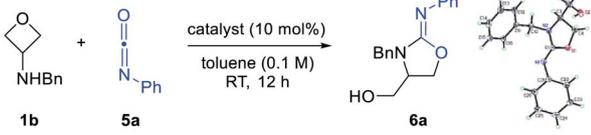
Scheme 2 Initial results between 3-aminooxetanes and thiocarbonyl compounds.



Scheme 3 Formal [3 + 2] annulation with isothiocyanates and CS₂. Reaction conditions: **1** (0.3–0.4 mmol), **2** (1.1 equiv.) or CS₂ (1.5 equiv.), DCM (2 mL), RT, 3 h for **3** and 36 h for **4**. Yields are for the isolated products.

sulfur as the nucleophilic site in the above cases, the oxygen atom is less nucleophilic. As expected, initial tests of the reactivity by mixing **1b** and **5a** resulted in no desired annulation product **6a** in the absence of a catalyst (Table 1, entry 1). Next, Brønsted acids, including TsOH and the super acid HNTf₂, were examined as catalysts, but with no success (entries 2 and 3). We then resorted to various Lewis acids, particularly those oxophilic ones, in hope of activating the oxetane unit. Unfortunately, many of them still remained ineffective (*e.g.*, ZnCl₂, AuCl, and FeCl₃). However, to our delight, further screening of stronger Lewis acids helped identify Sc(OTf)₃, Zn(OTf)₂, and



Table 1 Reaction conditions for annulation with isocyanates^a


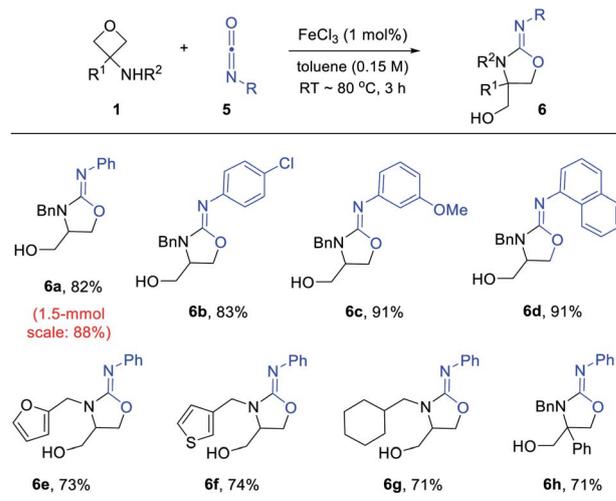
Entry	Catalyst	Yield ^b (%)
1	—	0
2	TsOH·H ₂ O	0
3	HNTf ₂	0
4	ZnCl ₂	0
5	AuCl	0
6	FeCl ₃	0
7	Sc(OTf) ₃	74
8	Zn(OTf) ₂	78
9	In(OTf) ₃	90
10	FeCl ₃ ^c	61
11	TsOH·H ₂ O ^c	0
12	FeCl ₃ ^c (1 mol%)	89(84) ^d
13	FeCl ₃ ^c (0.5 mol%)	85

^a Reaction scale: **1b** (0.1 mmol), **5a** (0.1 mmol), catalyst (10 mol%), toluene (1 mL). ^b Yield based on analysis of the ¹H NMR spectra of the crude reaction mixture using trichloroethylene as an internal reference. For all the entries, the urea product from simple amine addition to isocyanate **5a** accounts for the mass balance. ^c Run at 80 °C. ^d Isolated yield.

In(OTf)₃ to be effective at room temperature, leading to the desired iminoxazolidine product **6a** in good yield (entries 7–9). Its structure was confirmed by X-ray crystallography. Nevertheless, aiming to search for a cheaper catalyst, we continued to optimize this reaction at a higher temperature using previous ineffective catalysts. Indeed, FeCl₃ was found to be effective at 80 °C (61% yield, entry 10), while Brønsted acid TsOH remained ineffective at this temperature (entry 11). Notably, decreasing the loading of FeCl₃ to 1 mol% led to a higher yield (89% yield, entry 12). However, further decreasing to 0.5 mol% resulted in slightly diminished efficiency (entry 13).

While there are multiple effective catalysts, FeCl₃ was selected for the scope study in view of its low price. Various substituted 3-aminooxetanes and isocyanates were subjected to this annulation protocol (Scheme 4). The corresponding iminoxazolidine products were all obtained in good to excellent yields. Isocyanates containing an electron-donating or electron-withdrawing group were both suitable reaction partners. Remarkably, a 1.5 mmol scale reaction of **6a** also worked efficiently.

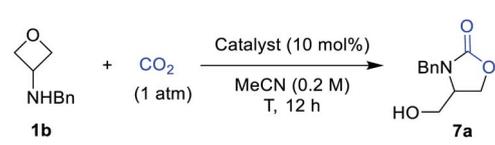
Although (thio)isocyanates and CS₂ have been successfully utilized in the formal [3 + 2] annulation with 3-aminooxetanes, these partners are relatively reactive. We were curious about whether the C=O bond in relatively inert molecules could react in a similar manner. For example, the C=O bond in CO₂ is both thermodynamically and kinetically inert relative to typical organic carbonyl groups. However, as a cheap, abundant and green one-carbon source, CO₂ has been a subject of persistent investigations owing to its versatility in various transformations leading to valuable materials.¹² Specifically, if CO₂ could be employed as



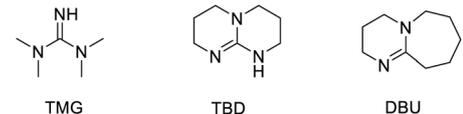
Scheme 4 Formal [3 + 2] annulation between 3-aminooxetanes and isocyanates. Reaction scale: **1** (0.3 mmol), **5** (0.3 mmol), FeCl₃ (1 mol%), toluene (2 mL).

a partner for the [3 + 2] annulation with 3-aminooxetanes, it would represent an attractive synthesis of oxazolidinones, a well-known heterocycle with applications in both organic synthesis and medicinal chemistry.¹³ In this context, we next studied the possibility of utilizing CO₂ in our annulation.

As expected, the reaction between **1b** and CO₂ at 1 atmospheric pressure did not proceed without a catalyst (Table 2, entry 1). Next, we examined representative Lewis acids, such as

Table 2 Reaction conditions for annulation with CO₂^a


Entry	Catalyst	T	Conv. (%)	Yield (%)
1	—	RT	0	0
2	Sc(OTf) ₃	RT	48	22
3	In(OTf) ₃	RT	33	9
4	Zn(OTf) ₂	RT	7	0
5	Sc(OTf) ₃	60 °C	100	61
6	Sc(OTf) ₃	80 °C	100	65
7	Et ₃ N	80 °C	0	0
8	DABCO	80 °C	5	0
9	TMG	80 °C	72	54
10	TBD	80 °C	100	88
11	DBU	80 °C	100	89



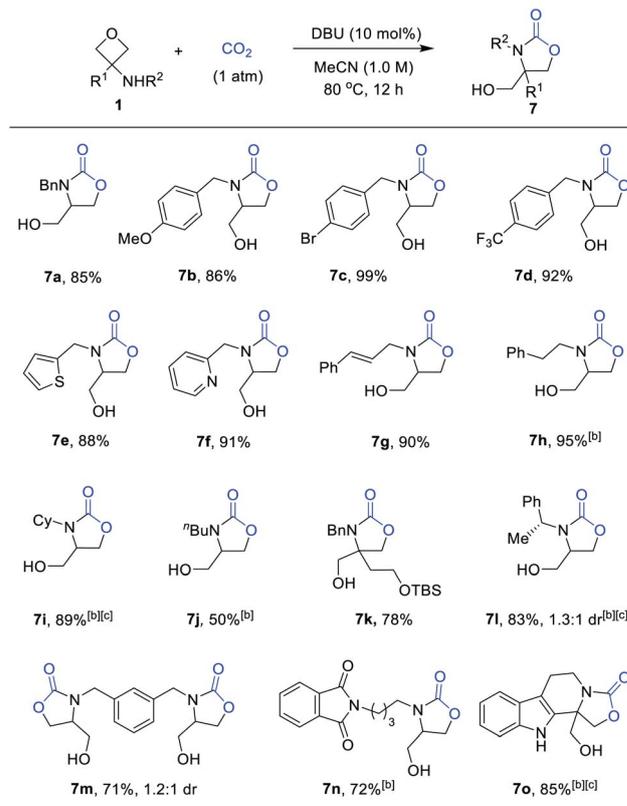
^a Reaction scale: **1b** (0.1 mmol), CO₂ (1 atm), solvent (0.5 mL). Yields based on analysis of the ¹H NMR spectra of the crude reaction mixture using CH₂Br₂ as an internal standard.



Sc(OTf)₃, In(OTf)₃ and FeCl₃. Among them, Sc(OTf)₃ exhibited the highest catalytic activity at room temperature (22% yield, entry 2). The reaction efficiency could be improved at 80 °C (65% yield, entry 6), but no further improvement could be made at a higher temperature or with other solvents. Next, we resorted to organic nitrogen bases, as they were known as effective activators of CO₂.¹⁴ While Et₃N and DABCO were completely ineffective for the reaction in MeCN at 80 °C, fortunately, TMG, TBD, and DBU were competent for the desired process (entries 7–11). Among them, DBU exhibited the best performance, leading to the desired product **7a** in 89% yield (entry 11). It is worth noting that the polar solvent MeCN was found to be crucial for the base-catalyzed reactivity. Less polar solvents, such as toluene, DCE or THF, completely shut down the reaction. We believe that effective stabilization of certain polar intermediates involved here is critically beneficial to decreasing the reaction barrier. Finally, unlike the previous Lewis acid-catalyzed annulation with isocyanates, this base-catalyzed [3 + 2] annulation with CO₂ proceeds *via* a different activation mode (*i.e.*, to activate CO₂ rather than oxetane). We believe that expansion of possible activation modes in this type of amphoteric reactivity will enrich the chemistry of oxetanes.

We next examined the scope of this CO₂-fixation process. Unfortunately, at a larger scale (0.5 mmol), the same condition (entry 11, Table 2) could not lead to complete conversion within 12 h. Therefore, further optimization aiming to accelerate the reaction was performed. Indeed, a higher concentration (1.0 M) resulted in a higher rate without affecting the yield. As shown in Scheme 5, a wide variety of 3-aminooxetanes were smoothly converted to the corresponding oxazolidinones in high yields. Both electron-donating and electron-withdrawing substituents on the *N*-benzyl group did not affect the efficiency. Heterocycle-based *N*-benzyl or *N*-allylic substituents are all suitable substrates. However, for regular alkyl substituents, such as homobenzyl (**7h**) or *n*-butyl (**7j**), the stronger base catalyst TBD was needed to achieve good efficiency. Furthermore, this reaction can tolerate steric hindrance in the 3-position of the oxetane (**7k**), where a quaternary carbon center could be incorporated. However, increasing the size of the *N*-substituent, such as the secondary alkyl groups in **7i** and **7l**, did influence the reactivity, thus requiring a higher temperature (100 °C). This process exhibited good compatibility with diverse functional groups, such as ethers, pyridines, aryl halides, olefins, silyl-protected alcohols, and phthalimides. Finally, this protocol is also capable of generating various oxazolidinones embedded in a different structural context, such as chiral oxazaolindinone **7l**, bis(oxazolidinone) **7m**, and polyheterocycle-fused oxazolidinone **7o**.

In summary, 3-aminooxetanes have been systematically demonstrated, for the first time, as versatile 1,3-amphoteric molecules. They are a new addition to the limited family of amphoteric molecules. Though previously unappreciated, these molecules exhibited various advantages over the related known 1,3-amphoteric molecules (*e.g.*, α -aziridine aldehydes and amino isocyanates), including easy access and extraordinary stability. The perfect position of the nucleophilic nitrogen together with the orthogonal electrophilic carbon allowed them to participate in a diverse range of intermolecular formal [3 + 2] annulations



Scheme 5 Formal [3 + 2] annulation between 3-aminooxetanes and CO₂. ^aReaction scale: **1** (0.5 mmol), CO₂ (1 atm), DBU (10 mol%), MeCN (0.5 mL). Isolated yield. ^bRun with TBD as the catalyst. ^cRun with DMF as solvent at 100 °C.

with polarized π -systems, leading to rapid access to various valuable nitrogen heterocycles. Different types of polarized double bonds, from reactive (thio)isocyanates to inert CO₂, all participated efficiently in these highly selective annulations with or without a suitable catalyst. Furthermore, the involvement of more functional groups in such amphoteric reactivity allowed manifold activation modes, thereby greatly enriching the reactivity of the already versatile oxetane unit to a new dimension. These reactions, proceeding in an intermolecular convergent manner from readily available substrates, provide expedient access to various valuable nitrogen heterocycles, thus being complementary to those traditional methods that either required multiple steps or less available substrates. More studies on the 1,3-amphoteric reactivity of 3-oxetanes, particularly those with other partners as well as their asymmetric variants, are ongoing in our laboratory.

Conflicts of interest

There are no conflicts to declare.

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References

- Z. He, A. Zajdlik and A. K. Yudin, *Acc. Chem. Res.*, 2014, **47**, 1029–1040.
- A. Dömling and I. Ugi, *Angew. Chem., Int. Ed.*, 2000, **39**, 3168–3210.
- (a) J. D. St. Dennis, Z. He and A. K. Yudin, *ACS Catal.*, 2015, **5**, 5373–5379; (b) P. Trinchera, V. B. Corless and A. K. Yudin, *Angew. Chem., Int. Ed.*, 2015, **54**, 9038–9041; (c) D. B. Diaz, C. C. G. Scully, S. K. Liew, S. Adachi, P. Trinchera, J. D. St. Denis and A. K. Yudin, *Angew. Chem., Int. Ed.*, 2016, **55**, 12659–12663.
- (a) R. Hili and A. K. Yudin, *J. Am. Chem. Soc.*, 2006, **128**, 14772–14773; (b) R. Hili and A. K. Yudin, *Angew. Chem., Int. Ed.*, 2008, **47**, 4188–44191; (c) R. Hili and A. K. Yudin, *J. Am. Chem. Soc.*, 2009, **131**, 16404–16406; (d) R. Hili, V. Rai and A. K. Yudin, *J. Am. Chem. Soc.*, 2010, **132**, 2889–2891; (e) B. H. Rotstein, V. Rai, R. Hili and A. K. Yudin, *Nat. Protoc.*, 2010, **5**, 1813–1822; (f) S. Baktharaman, N. A. Afagh, A. Vandersteen and A. K. Yudin, *Org. Lett.*, 2010, **12**, 240–243; (g) L. L. W. Cheung, Z. He, S. M. Decker and A. K. Yudin, *Angew. Chem., Int. Ed.*, 2011, **50**, 11798–11802.
- (a) J.-G. Roveda, C. Clavette, A. D. Hunt, S. I. Gorelsky, C. J. Whipp and A. M. Beauchemin, *J. Am. Chem. Soc.*, 2009, **131**, 8740–8741; (b) J.-G. Roveda, C. Clavette, W. Gan, A. Bongers, T. Markiewicz, A. B. Toderian, S. I. Gorelsky and A. M. Beauchemin, *J. Am. Chem. Soc.*, 2012, **134**, 16111–16114; (c) C. Clavette, J.-F. Vincent Rocan and A. M. Beauchemin, *Angew. Chem., Int. Ed.*, 2013, **52**, 12705–12708; (d) J.-F. Vincent-Rocan, C. Clavetter, K. Leckett and A. M. Beauchemin, *Chem.–Eur. J.*, 2015, **21**, 3886–3890.
- For leading reviews on oxetane chemistry, see: (a) J. A. Burkhard, G. Wuitschik, M. Rogers-Evans, K. Müller and E. M. Carreira, *Angew. Chem., Int. Ed.*, 2010, **49**, 9052–9067; (b) D. J. Mack and J. T. Njardarson, *ACS Catal.*, 2013, **3**, 272–286; (c) Z. Wang, Z. Chen and J. Sun, *Org. Biomol. Chem.*, 2014, **12**, 6028–6032; (d) C. A. Malapit and A. R. Howell, *J. Org. Chem.*, 2015, **80**, 8489–8495; (e) S. Li and J. Xu, *Prog. Chem.*, 2016, **28**, 1798–1810; (f) J. A. Bull, R. A. Croft, O. A. Davis, R. Doran and K. F. Morgan, *Chem. Rev.*, 2016, **116**, 12150–12233.
- Selected examples of intermolecular nucleophilic opening of oxetanes: (a) Y. Wang, H. Bekolo and A. R. Howell, *Tetrahedron*, 2002, **58**, 7101–7107; (b) H. Yoshida, Y. Asatsu, Y. Mimura, Y. Ito, J. Ohshita and K. Takaki, *Angew. Chem., Int. Ed.*, 2011, **50**, 9676–9679; (c) N. Ishida, Y. Nakanishi and M. Murakami, *Angew. Chem., Int. Ed.*, 2013, **52**, 11875–11878; (d) Z. Wang, Z. Chen and J. Sun, *Angew. Chem., Int. Ed.*, 2013, **52**, 6685–6688; (e) W. Yang, Z. Wang and J. Sun, *Angew. Chem., Int. Ed.*, 2016, **55**, 6954–6958.
- Selected examples of intramolecular nucleophilic opening of oxetanes: (a) T. Bach and J. Shröder, *Tetrahedron Lett.*, 1997, **38**, 3707–3710; (b) T. Bach, K. Kather and O. Krämer, *J. Org. Chem.*, 1998, **63**, 1910–1918; (c) T. Bach and J. Shröder, *J. Org. Chem.*, 1999, **64**, 1265–1273; (d) T. Bach, *Synlett*, 2000, 1699–1707; (e) R. N. Loy and E. N. Jacobsen, *J. Am. Chem. Soc.*, 2009, **131**, 2786–2787; (f) J. A. Burkhard, B. H. Tchitchanov and E. M. Carreira, *Angew. Chem., Int. Ed.*, 2011, **50**, 5379–5382; (g) W. Zhao, Z. Wang and J. Sun, *Angew. Chem., Int. Ed.*, 2012, **51**, 6209–6213; (h) Z. Chen, B. Wang, Z. Wang, G. Zhu and J. Sun, *Angew. Chem., Int. Ed.*, 2013, **52**, 2027–2031; (i) S. A. Ruider, S. Müller and E. M. Carreira, *Angew. Chem., Int. Ed.*, 2013, **52**, 11908–11911; (j) R. A. Croft, J. J. Mousseau, C. Choi and J. A. Bull, *Chem.–Eur. J.*, 2016, **22**, 16271–16276; (k) W. Yang and J. Sun, *Angew. Chem., Int. Ed.*, 2016, **55**, 1868–1871; (l) A. R. White, R. A. Kozłowski, S.-C. Tsai and C. D. Vanderwal, *Angew. Chem., Int. Ed.*, 2017, **56**, 10525–10529; (m) D. Xu, H. Wei, Y. Zhen, Y.-Q. Gao, R. Li, X. Li, Y. He, Z. Zhang and W. Xie, *Org. Chem. Front.*, 2019, **6**, 1681–1685; (n) L. G. DeRatt, E. C. Lawson, C.-Y. Wang and S. D. Kuduk, *Org. Lett.*, 2019, **21**, 9642–9645; (o) R. Zhang, W. Guo, M. Duan, K. N. Houk and J. Sun, *Angew. Chem., Int. Ed.*, 2019, **58**, 18055–18060; (p) H. Huang, W. Yang, Z. Chen, Z. Lai and J. Sun, *Chem. Sci.*, 2019, **10**, 9586–9590.
- Selected recent examples of oxetane ring expansion: (a) D. Rix, R. Ballesteros-Garrido, W. Zeghida, C. Besnard and J. Lacour, *Angew. Chem., Int. Ed.*, 2011, **50**, 7308–7311; (b) C. Gronnier, S. Kramer, Y. Odabachian and F. Gagosz, *J. Am. Chem. Soc.*, 2012, **134**, 828–831; (c) B. Guo, G. Schwarzwald and J. T. Njardarson, *Angew. Chem., Int. Ed.*, 2012, **51**, 5675–5678; (d) Q. Yin and S.-L. You, *Org. Lett.*, 2014, **16**, 1810–1813; (e) Y.-N. Wang, L.-C. Yang, Z.-Q. Rong, T.-L. Liu, R. Liu and Y. Zhao, *Angew. Chem., Int. Ed.*, 2018, **57**, 1596–1600; (f) S. Jana, Z. Yang, C. P. X. Xu and R. M. Koenigs, *Chem. Sci.*, 2019, **10**, 10129–10134.
- J. Rintjema, W. Guo, E. Martin, E. C. Escudero-Adan and A. W. Kleij, *Chem.–Eur. J.*, 2015, **21**, 10754–10762.
- (a) M.-H. Shih, Y.-Y. Xu, Y.-S. Yang and G.-L. Lin, *Molecules*, 2015, **20**, 6520–6532; (b) U. K. Shukla, R. Singh, J. M. Khanna, A. K. Saxena, H. K. Singh, R. N. Sur, B. N. Dhawan and N. Anand, *Collect. Czech. Chem. Commun.*, 1992, **57**, 415–424; (c) A. Hense, R. Fischer, E.-R. Gesing, S. Herrmann, K. Kather, S. Lehr, K. Voigt, H.-J. Riebel, P. Jeschke, C. Erdelen, W. Andersch, P. Lösel and U. Reckmann, *PCT Int. Appl.*, 2002, 96872; (d) M. D'hooghe and N. De Kimpe, *Tetrahedron*, 2006, **62**, 513–535.
- (a) T. Sakakura, J.-C. Choi and H. Yasuda, *Chem. Rev.*, 2007, **107**, 2365–2387; (b) Y. Cao, X. He, N. Wang, H.-R. Li and L.-N. He, *Chin. J. Chem.*, 2018, **36**, 644–659.
- (a) M. R. Barbachyn and C. W. Ford, *Angew. Chem., Int. Ed.*, 2003, **42**, 2010–2023; (b) T. A. Mukhtar and G. D. Wright, *Chem. Rev.*, 2005, **105**, 529–542; (c) K. Michalska, I. Karpiuk, M. Krul and S. Tyski, *Bioorg. Med. Chem.*, 2013, **21**, 577–591; (d) M. M. Heravi and V. Zadsirjan, *Tetrahedron: Asymmetry*, 2013, **24**, 1149–1188.
- DBU and other organic bases have been known to activate CO₂: (a) S. Wang and C. Xi, *Chem. Soc. Rev.*, 2019, **48**, 382–404; (b) C. Villiers, J.-P. Dognon, R. Pollet, P. Thuéry and M. Ephritikhine, *Angew. Chem., Int. Ed.*, 2010, **49**, 3465–3468.

