

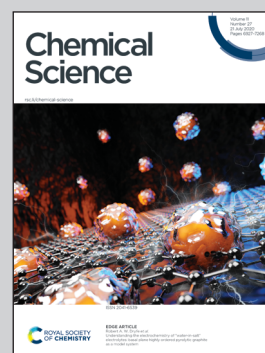


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and TEMPOH catalyzed by the phenolate anion under
visible light

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with strong reducing potentials (-3.16 V vs. SCE). The visible
light activated phenolate anion enabled the reduction of
(hetero)aryl halides to (hetero)aryl radicals through single
electron transfer. A novel and efficient protocol for the
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TEMPOH was developed with the new photocatalyst.

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Intermolecular oxyarylation of olefins with aryl halides and TEMPOH catalyzed by the phenolate anion under visible light†

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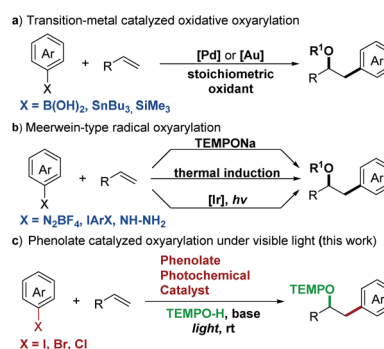
The phenolate anion was discovered as a new photocatalyst with strong reduction potentials. Under visible light irradiation, the phenolate anion enabled the reduction of (hetero)aryl halides (including electron-rich aryl chlorides) to (hetero)aryl radicals through single electron transfer. Based on this new photocatalyst, a novel and efficient photocatalytic protocol for the intermolecular oxyarylation of olefins with aryl halides and TEMPOH was developed. The developed three-component coupling reaction proceeded under redox-neutral reaction conditions with stable and readily available synthons and exhibited broad substrate scope. The utility of this process was further highlighted by the diversified chemical manipulation of the resulting oxyarylation products and the late-stage modification of active pharmaceutical ingredients.

Introduction

Since the chemical reactivity of electronically excited molecules differs fundamentally from that of those in the ground state, the photochemistry induced by visible light provides fresh opportunities to expand the potential of organic chemistry. Various photocatalysts, including Ir or Ru complexes and organic dyes, have been developed and extensively explored.¹ Phenolate anions are a useful model system for photoinduced electron ejection mechanism studies.² Recent research showed that phenolate anions acted as photoreductants to activate perfluoroalkyl iodides in a perfluoroalkylation reaction.³ We also discovered that the visible light-excited vinylphenolate anions enabled the direct reduction of aryl halides to aryl radicals.⁴ However, attempts to use the vinylphenolate anion as a new photocatalyst failed and only afforded the Heck-type arylation products by coupling with aryl radicals. In this manuscript, we discovered that the tri-substituted phenolate anion possessed strong reduction potential and acted as a new photocatalyst with redox-neutral features. To demonstrate abilities of this new photocatalyst, we design an intermolecular oxyarylation of olefins with aryl halides under visible light irradiation.

Intermolecular oxyarylation has attracted great attention since one oxygen substituent and one aryl group add across an

olefin *via* a vicinal difunctionalization process. In the light of these benefits, transition-metal (Pd and Au) catalyzed oxidative oxyarylation has been extensively studied and several protocols for oxyarylation of olefins with arylboronic acids,⁵ arylstannanes,⁶ or arylsilanes⁷ were reported (Scheme 1a). In addition, Meerwein-type radical oxyarylation methods have been developed as an attractive alternative in recent years with aryl diazonium salts,⁸ diaryliodonium salts,^{8e,9} or aryl hydrazines¹⁰ as radical precursors (Scheme 1b). While numerous advances have been made in this area, the ability to access this molecular scaffold array with modular flexibility remains limited. Requirement of the use of toxic, unstable, or unavailable aryl precursors has impeded the further adaptation of these methods in the synthetic community. Furthermore, no general



Scheme 1 Strategies for intermolecular oxyarylation of olefins. (a) Transition-metal catalyzed oxidative oxyarylation. (b) Meerwein-type radical oxyarylation. (c) Phenolate catalyzed oxyarylation under visible light.

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protocol for oxyarylation of olefins with aromatic heterocyclic motifs which are essential in a wide range of pharmaceuticals has yet been reported.

In this context, a catalytic method for the direct oxyarylation of olefins with (hetero)aryl halides which are the cheapest and the most readily available arylation reagents would circumvent these issues and greatly expand the scope of aryl precursors. The single-electron reductive activation of aryl halides by photoredox catalysis has been a powerful strategy for aryl radical generation from stable starting materials.¹¹ However, most of these methods require a large excess of sacrificial electron donors which result in a fast competing hydrogen atom abstraction by radical intermediates to deliver undesired side products and limit the choice of radical coupling partners. Additionally, aryl chlorides without electron withdrawing substituents are still a poorly accessible substrate for currently known visible light photocatalysts due to their large negative reduction potentials.¹² Here, we demonstrated the application of the colored phenolate anion as a new visible light photocatalyst for the catalytic generation of (hetero)aryl radicals from non-activated (hetero)aryl halides and the successful development of an operationally simple and redox-neutral protocol for the intermolecular oxyarylation of a wide range of olefins (Scheme 1c).

Results and discussion

The details of our working hypothesis are described in Scheme 2. We envisioned a photocatalytic cycle wherein a suitable phenol catalyst **A** would first be deprotonated under basic conditions to generate the colored phenolate anion **B**, which could reach an electronically excited state (**B***) under visible light irradiation. On the basis of our previous work,⁴ we proposed that a single electron transfer (SET) process between the excited phenolate anion **B*** and aryl halide **D** could occur to produce a phenoxyl radical **C** and an aryl radical **E**. The electrophilic aryl radical **E** would then undergo intermolecular

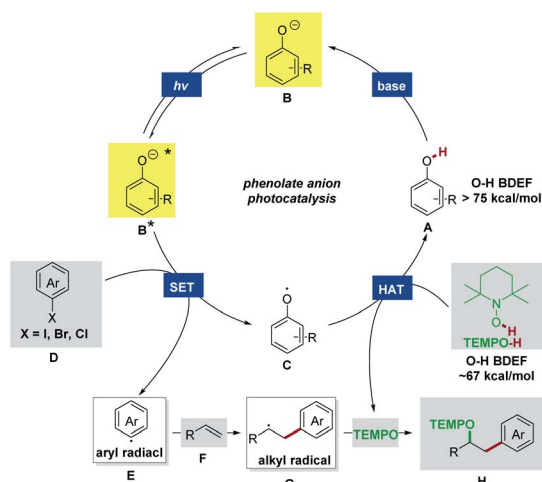
addition to an olefin acceptor **F** to furnish a new C–C bond and an adjacent alkyl radical **G**. In the meantime, the phenoxyl radical **C** could abstract a hydrogen atom from TEMPO–H (the O–H BDFE in TEMPO–H ~67 kcal mol^{−1})¹³ to recycle the phenol **A** (the O–H BDFE in phenol >75 kcal mol^{−1})^{13,14} and afford the persistent TEMPO radical. Finally, the highly selective radical–radical cross coupling of the transient alkyl radical **G** with TEMPO would be kinetically feasible to provide the desired oxyarylation product **H** based on the persistent radical effect.¹⁵

To evaluate these ideas, our initial exploration focused on the oxyarylation of 4-methoxystyrene **2** with iodobenzene **1a** as a model reaction. The 4-phenylphenol derivatives were chosen as catalysts. After irradiation with two 18 W blue light emitting diode (LED) lamps for 16 h, a promising lead result was observed by using 4-phenylphenol (**PC1**) as the catalyst in the presence of TEMPOH and Cs₂CO₃ in DMSO, providing a 40% yield of the desired oxyarylation product **3** at room temperature (Table 1, entry 1). Further screening of catalysts revealed that 2,6-di-*tert*-butyl-4-phenylphenol (**PC3**) was the optimal catalyst for this protocol (entries 2–6), and product **3** was obtained in an 87% yield (entry 3). Interestingly, catalysts with a bis-*t*Bu system (**PC3**–**PC6**) afforded higher yields compared to other tested catalysts (**PC1** and **PC2**). The redox reversibility observed by electrochemical measurements (see ESI Fig. 5†) suggested that the introduction of bis-*t*Bu remarkably prolongs the lifetime of the corresponding phenoxyl radicals. Therefore, we reasoned that the steric hindrance of the bis-*t*Bu system might be good

Table 1 The model reaction and the reaction parameters evaluated

Entry ^a	X	Catalyst	Light sources	Yield ^b (%)
1	I	PC1	Blue LEDs	40
2	I	PC2	Blue LEDs	57
3	I	PC3	Blue LEDs	87
4	I	PC4	Blue LEDs	79
5	I	PC5	Blue LEDs	73
6	I	PC6	Blue LEDs	82
7	I	PC3	Dark	0
8	I	None	Blue LEDs	0
9 ^c	I	PC3	Blue LEDs	0
10	Br	PC3	Blue LEDs	80
11	Cl	PC3	Blue LEDs	42

^a General conditions: halobenzene **1** (0.30 mmol, 1.0 equiv.), 4-methoxystyrene **2** (0.45 mmol, 1.5 equiv.), TEMPOH (0.36 mmol, 1.2 equiv.), Cs₂CO₃ (0.60 mmol, 2.0 equiv.), catalyst (0.03 mmol, 10 mol%), and DMSO (1.5 mL, rigorously degassed by freezing/pumping/thawing). Halobenzene **1** and TEMPOH were dissolved in DMSO and added dropwise over 10 h. ^b Isolated yields by chromatography. ^c Reaction was run in the absence of Cs₂CO₃.

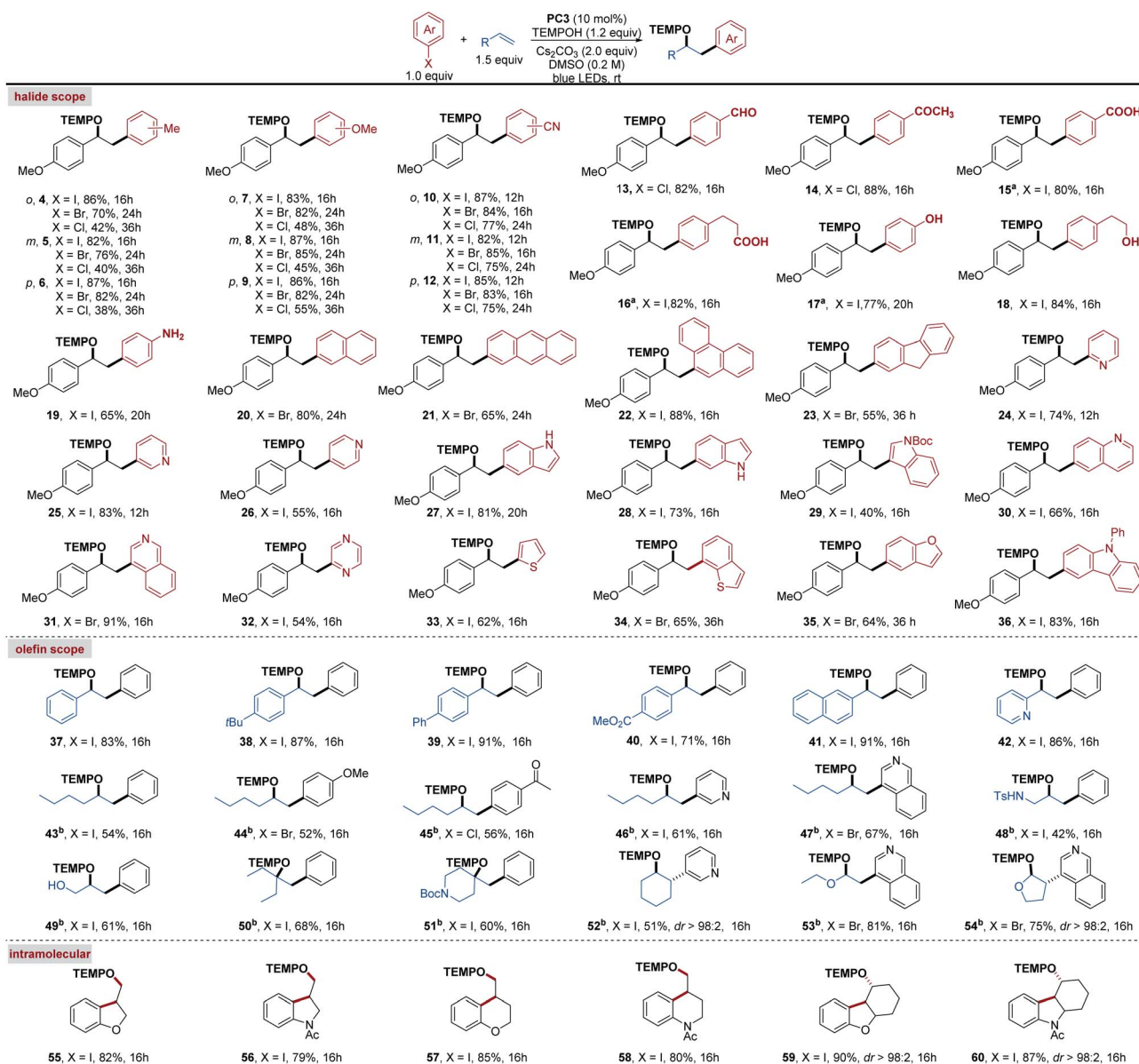


Scheme 2 Proposed photocatalytic mechanism for intermolecular radical oxyarylation.



for the catalyst regeneration by improving the kinetic stability of the phenoxyl radicals, thus leading to high efficiency. In an attempt to seek better reaction conditions, a survey of other bases (KHCO_3 , Na_2CO_3 , K_2CO_3 , K_3PO_4 , DBU, and TMG) and solvents (DMF, DMA, acetone, and CH_3CN) indicated that they were not comparable to the combination of Cs_2CO_3 and DMSO (see ESI Table 1†). The control experiments revealed that the light, catalyst, and base were essential for this reaction (entries 7–9). Besides the bromobenzene **1b**, the chlorobenzene **1c** was also efficiently activated for the oxyarylation, albeit a moderate yield was achieved (entries 10 and 11). These results showed that the excited state phenolate anion was capable enough for reducing inert chlorobenzene.

With the optimized conditions in hand, we then explored the scope of the (hetero)aryl halides (Scheme 3, top). Iodides, bromides and chlorides (regardless of the electronic properties of the starting substrates) were successfully coupled to 4-methoxystyrene **2** with good to moderate yields (**4**–**12**). These results further emphasized the strong reducing properties of the excited phenolate anion of **PC3**. Moreover, the *o*- (**4**, **7**, and **10**), *m*- (**5**, **8**, and **11**) and *p*- (**6**, **9**, and **12**) substituted aryl halides readily participated in this transformation demonstrating that the site of the substitution on the aromatic ring had no detrimental effects on our method. It is worth highlighting that the mild conditions employed by this reaction were proved to be compatible with aryl halides containing a wide range of



Scheme 3 Reaction scope. General conditions: (hetero)aryl halide (0.30 mmol, 1.0 equiv.), olefin (0.45 mmol, 1.5 equiv.), TEMPOH (0.36 mmol, 1.2 equiv.), Cs_2CO_3 (0.60 mmol, 2.0 equiv.), **PC3** (0.03 mmol, 10 mol%), and DMSO (1.5 mL, rigorously degassed by freezing/pumping/thawing). (Hetero)aryl halides and TEMPOH were dissolved in DMSO and added dropwise over 10 h. Isolated yields were reported. ^aWith 3 equiv. of Cs_2CO_3 . ^bWith 3 equiv. of the olefin.



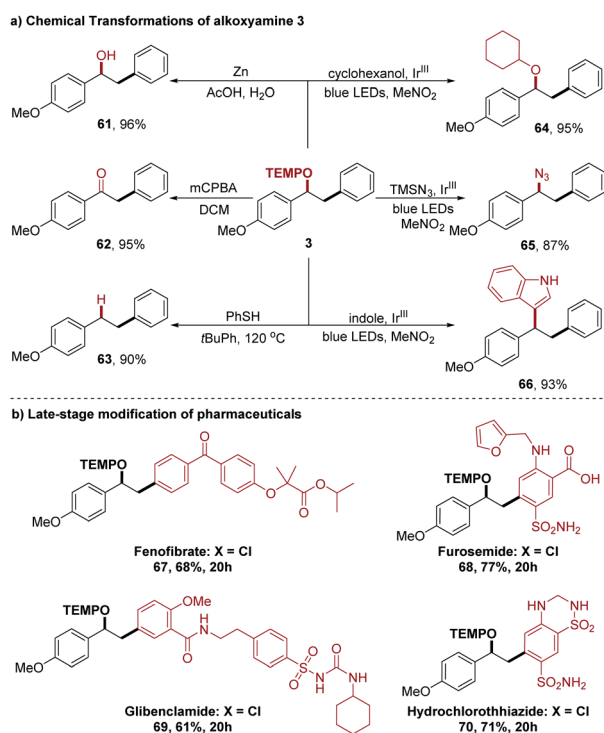
functional groups, including unprotected $-\text{CHO}$ (**13**), $-\text{COCH}_3$ (**14**), $-\text{COOH}$ (**15** and **16**), $-\text{OH}$ (**17** and **18**), and $-\text{NH}_2$ (**19**), thereby eliminating the need for inefficient protection and deprotection procedures often required by traditional methodologies. Pleasingly, polycyclic aromatic hydrocarbon halides, such as naphthalene (**20**), anthracene (**21**), phenanthrene (**22**), and fluorene (**23**) can also be used as radical precursors under these conditions. Lastly, we investigated this photocatalytic transformation with a variety of privileged heterocyclic motifs, which were not involved in previously known methods. Remarkably, heterocycles such as pyridine (**24–26**), indole (**27–29**), quinoline (**30**), isoquinoline (**31**), pyrazine (**32**), thiophene (**33**) thianaphthene (**34**), benzofuran (**35**) and carbazole (**36**) were found to be efficient coupling partners in our reaction system.

We next surveyed the scope of olefin acceptors that could be used for the current transformation (Scheme 3, middle). Using iodobenzene **1a** as a model aryl halide, various styrene derivatives were effectively converted to the corresponding oxyarylation products **37–42** in good to excellent yields under the standard conditions. Furthermore, this oxyarylation method was not only limited to aryl olefins; non-activated aliphatic olefins were also found to be compatible with this reaction system (**43–52**). We observed that a series of representative (hetero)aryl halides were successfully coupled to 1-hexene to give the desired products **43–47** with synthetically useful yields. To our delight, allylic sulfonamide (**48**) and alcohol derivatives (**49**) were also readily accommodated under these reaction conditions. In addition, the 1,1- or 1,2-disubstituted olefins

were found to be competent coupling partners and provided the desired products with acceptable yields (**50–52**). Further studies revealed that enol ethers were also suitable substrates and yielded the oxyarylation products **53** and **54** in good yields. Subsequently, we found that intramolecular variants of this transformation were successful, and a variety of aryl iodides cyclized under our photocatalytic system to deliver a range of five- and six-membered heterocyclic products **55–60** (Scheme 3, bottom).

To expand the synthetic value of this method, we investigated the further chemical manipulation of these TEMPO-derived oxyarylation products (Scheme 4a). For instance, cleavage of the N–O bond in alkoxyamine **3** with Zn in acetic acid afforded the alcohol **61**, oxidation of alkoxyamine **3** smoothly proceeded to provide ketone **62** with *m*-chloroperoxybenzoic acid (MCPBA) in DCM, and radical deoxygenation occurred upon heating alkoxyamine **3** in the presence of thiophenol to deliver compound **63**.^{8a} In addition, alkoxyamine **3** is also an ideal carbocation precursor in the Ir-catalyzed alkoxyamine radical cation mesolytic cleavage reaction, and the resulting carbocation intermediate can be trapped by different nucleophiles such as cyclohexanol, TMSN₃, and indole to afford a series of useful molecular architectures **64–66**.¹⁶ Finally, we evaluated our method for late-stage functionalization applications (Scheme 4b). Pharmaceutical ingredients (fenofibrate, furosemide, glibenclamide, and hydrochlorothiazide) that contained an aryl chloride were effectively converted into corresponding oxyarylation products **67–70** in 61–77% yield, further highlighting the utility of our method in a complex setting.

As a means of illuminating the mechanism of this proposed photocatalytic transformation, a number of UV-vis experiments were conducted (Fig. 1a). In analogy with our previous studies,⁴ we found that the colorless solution of PC3 (orange line in Fig. 1a) immediately turned primrose yellow upon addition of



Scheme 4 Synthetic applications of this photocatalytic oxyarylation reaction. (a) Chemical transformations of alkoxyamine **3**. (b) Late-stage modification of pharmaceuticals.

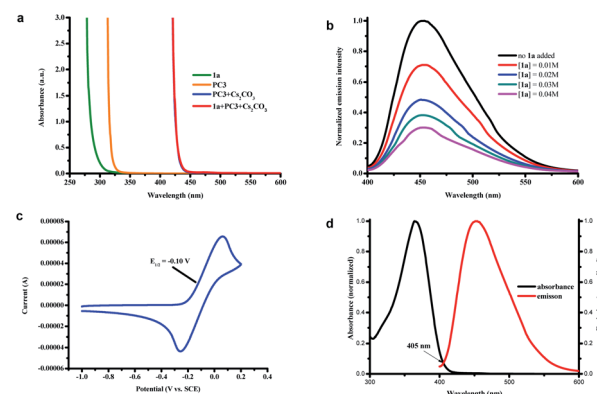
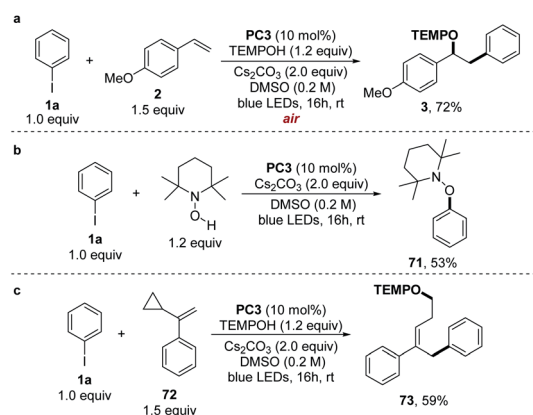
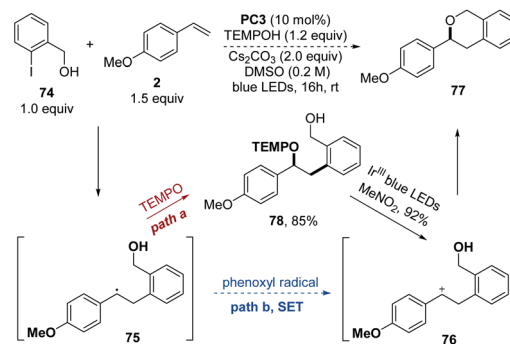


Fig. 1 (a) UV/Vis absorption spectra of the DMSO solutions (0.10 M) of **1a**, PC3, mixture of PC3 and Cs₂CO₃, and mixture of **1a**, PC3 and Cs₂CO₃. (b) Quenching of the phenolate of PC3 emission (5 × 10⁻⁵ M in DMSO) in the presence of increasing amounts of **1a**. (c) The cyclic voltammogram of the phenolate anion of PC3 vs. SCE in DMSO at 0.1 V s⁻¹. (d) Normalized absorption and emission spectra of the phenolate anion of PC3 in dry DMSO (5 × 10⁻⁵ M), and the intersection wavelength was considered to be 405 nm.

Cs_2CO_3 (blue line in Fig. 1a, the absorption of the phenolate anion of **PC3**) and did not observe any color change when iodobenzene **1a** was added to the solution of the phenolate anion of **PC3** (red line in Fig. 1a, perfectly overlapping with the absorption of the phenolate anion of **PC3**). These results excluded the formation of a ground-state EDA complex between the phenolate anion and aryl halide and indicated that the excited phenolate anion was responsible for triggering the formation of the aryl radical from its halide.^{3,17} Further support for this conclusion was obtained from Stern–Volmer quenching studies (Fig. 1b), in which the excited state of the phenolate anion of **PC3** was effectively quenched by iodobenzene **1a**. Moreover, a decline in the excited state lifetime also was observed in the time-resolved quenching studies, and the Stern–Volmer analysis revealed a linear correlation indicating a dynamic quenching of the excited phenolate anion of **PC3** by **1a** (see ESI note 5†).^{12b} In addition, the thermodynamic feasibility of the photoinduced SET was analyzed from the oxidation–reduction potentials. The ground-state potential $E_{1/2}$ (**PC3***/**PC3**[−]) was determined to be -0.10 V vs. SCE (Fig. 1c) and the excited-state energy E_{0-0} (**PC3***/**PC3**[−]) was read from the intersection of the normalized absorbance and emission spectra at 405 nm as 3.06 eV (Fig. 1d).¹⁸ Therefore, the redox potential of the excited phenolate anion of **PC3** was calculated to be -3.16 V vs. SCE ($E_{1/2}$ (**PC3***/**PC3**[−]) = $E_{1/2}$ (**PC3***/**PC3**[−]) − E_{0-0} (**PC3***/**PC3**[−])), which indicated that the reduction of iodobenzene **1a** (-2.24 V vs. SCE), bromobenzene **1b** (-2.44 V vs. SCE), and even chlorobenzene **1c** (-2.78 V vs. SCE) by the excited phenolate anion of **PC3** is thermodynamically feasible.¹⁹ To further probe whether the singlet excited state of the phenolate anion of **PC3** could be responsible for the catalysis in our system, an oxygen tolerance experiment was performed open to air (Scheme 5a). Under these conditions, where triplet pathways should be inhibited by oxygen (a potent triplet quencher),²⁰ a 72% yield was still observed after 16 h. These results indicated that the singlet state may be the primary mode of catalysis, but the combination catalysis of singlet and triplet states in our system cannot be completely ruled out at this time.^{20b}



Scheme 5 (a) Oxygen tolerance experiments. (b) Control experiments without 4-methoxystyrene. (c) Radical clock experiments.



Scheme 6 Intramolecular carbocation trapping experiments.

To support the radical mechanism suggested in Scheme 2, we performed several mechanism experiments. In the absence of 4-methoxystyrene, the reaction provided TEMPO-Ph (**71**) in a 53% yield (Scheme 5b), proving that a phenyl radical along with a TEMPO was formed under reaction conditions. We then carried out the radical clock experiments with α -cyclopropylstyrene (**72**) and found that the ring-opened adduct **73** was obtained in a 59% yield (Scheme 5c). These results strongly suggested the formation of a benzyl radical as the intermediate.²¹

Besides the proposed pathway of coupling between the alkyl radical and TEMPO as shown in Scheme 2, it is also possible that the alkyl radical is first oxidized by the phenoxyl radical in the SET pathway to give the carbocation and then is trapped by the TEMPOH alcohol to afford the oxyarylation product. In order to exclude the possible SET pathway between the benzyl radical and the phenoxyl radical, 2-iodophenylmethanol **74** was chosen as a substrate to trap the carbocation intermediate by the intramolecular alcohol (Scheme 6). If the carbocation **76** was generated by the SET process with the phenoxyl radical (path b), an intermolecular cyclization will follow to give product **77**. However, only the intermolecular oxyarylation product **78** was obtained in a 85% yield without any formation of intramolecular cyclization product **77**. Instead, the alkoxyamine **78** was efficiently cyclized to provide **77** in a 92% yield through a carbocation intermediate in the Ir-catalyzed alkoxyamine radical cation mesolytic cleavage reaction,¹⁶ confirming that the cyclization could easily take place if the carbocation **76** was generated. These results implied that the radical–radical cross coupling of the benzyl radical with the HAT product TEMPO to provide the desired oxyarylation product would be more plausible.

Conclusions

In conclusion, the phenolate anion was discovered as a new photocatalyst under visible light and enabled the intermolecular oxyarylation of olefins with aryl and heteroaryl halides. This process utilized stable and readily available halogenated arenes as radical precursors, avoided the use of sacrificial reductants, and exhibited broad functional group tolerance. Moreover, the developed three-component coupling reaction was further



applied to the late-stage modification of active pharmaceutical ingredients, and the resulting TEMPO-derived oxyarylation products are ideal precursors for diversified chemical manipulations to afford a series of useful molecular architectures. The highly reducing character of this photocatalytic system together with the mild reaction conditions will open new avenues for the development of aryl radical-based transformations.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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Notes and references

- For recent reviews, see: (a) C. K. Prier, D. A. Rankic and D. W. C. MacMillan, *Chem. Rev.*, 2013, **113**, 5322–5363; (b) N. A. Romero and D. A. Nicewicz, *Chem. Rev.*, 2016, **116**, 10075–10166; (c) M. D. Kärkäs, J. A. Porco and C. R. J. Stephenson, *Chem. Rev.*, 2016, **116**, 9683–9747; (d) M. Silvi and P. Melchiorre, *Nature*, 2018, **554**, 41; (e) J.-R. Chen, X.-Q. Hu, L.-Q. Lu and W.-J. Xiao, *Acc. Chem. Res.*, 2016, **49**, 1911–1923; (f) M. N. Hopkinson, A. Tlahuext-Aca and F. Glorius, *Acc. Chem. Res.*, 2016, **49**, 2261–2272; (g) D. Staveness, I. Bosque and C. R. J. Stephenson, *Acc. Chem. Res.*, 2016, **49**, 2295–2306; (h) X. Huang and E. Meggers, *Acc. Chem. Res.*, 2019, **52**, 833–847; (i) X. Lang, J. Zhao and X. Chen, *Chem. Soc. Rev.*, 2016, **45**, 3026–3038; (j) S. Protti, S. Garbarino, D. Ravelli and A. Basso, *Angew. Chem., Int. Ed.*, 2016, **55**, 15476–15484; (k) J. Xie, H. Jin and A. S. K. Hashmi, *Chem. Soc. Rev.*, 2017, **46**, 5193–5203.
- (a) X. Chen, D. S. Larsen, S. E. Bradforth and I. H. M. van Stokkum, *J. Phys. Chem. A*, 2011, **115**, 3807–3819; (b) T. Ichino and R. W. Fessenden, *J. Phys. Chem. A*, 2003, **107**, 9257–9268; (c) R. W. Fessenden and N. C. Verma, *J. Am. Chem. Soc.*, 1976, **98**, 243–244; (d) G. N. Lewis and D. Lipkin, *J. Am. Chem. Soc.*, 1942, **64**, 2801–2808.
- G. Filippini, M. Nappi and P. Melchiorre, *Tetrahedron*, 2015, **71**, 4535–4542.
- K. Liang, T. Li, N. Li, Y. Zhang, L. Shen, Z. Ma and C. Xia, *Chem. Sci.*, 2020, **11**, 2130–2135.
- (a) S. Kirchberg, R. Fröhlich and A. Studer, *Angew. Chem., Int. Ed.*, 2010, **49**, 6877–6880; (b) A. D. Melhado, W. E. Brenzovich, A. D. Lackner and F. D. Toste, *J. Am. Chem. Soc.*, 2010, **132**, 8885–8887.
- A. D. Satterfield, A. Kubota and M. S. Sanford, *Org. Lett.*, 2011, **13**, 1076–1079.
- (a) L. T. Ball, M. Green, G. C. Lloyd-Jones and C. A. Russell, *Org. Lett.*, 2010, **12**, 4724–4727; (b) W. E. Brenzovich, J.-F. Brazeau and F. D. Toste, *Org. Lett.*, 2010, **12**, 4728–4731; (c) L. T. Ball, G. C. Lloyd-Jones and C. A. Russell, *Chem.-Eur. J.*, 2012, **18**, 2931–2937.
- (a) M. Hartmann, Y. Li and A. Studer, *J. Am. Chem. Soc.*, 2012, **134**, 16516–16519; (b) C.-J. Yao, Q. Sun, N. Rastogi and B. König, *ACS Catal.*, 2015, **5**, 2935–2938; (c) S. Kindt, K. Wicht and M. R. Heinrich, *Angew. Chem., Int. Ed.*, 2016, **55**, 8744–8747; (d) M. R. Heinrich, A. Wetzel and M. Kirschstein, *Org. Lett.*, 2007, **9**, 3833–3835; (e) M. N. Hopkinson, B. Sahoo and F. Glorius, *Adv. Synth. Catal.*, 2014, **356**, 2794–2800; (f) M. Bu, T. F. Niu and C. Cai, *Catal. Sci. Technol.*, 2015, **5**, 830–834; (g) R. Govindarajan, J. Ahmed, A. K. Swain and S. K. Mandal, *J. Org. Chem.*, 2019, **84**, 13490–13502.
- (a) G. Fumagalli, S. Boyd and M. F. Greaney, *Org. Lett.*, 2013, **15**, 4398–4401; (b) M. Hartmann, Y. Li, C. Mück-Lichtenfeld and A. Studer, *Chem.-Eur. J.*, 2016, **22**, 3485–3490.
- (a) Y. Su, X. Sun, G. Wu and N. Jiao, *Angew. Chem., Int. Ed.*, 2013, **52**, 9808–9812; (b) T. Taniguchi, H. Zaimoku and H. Ishibashi, *Chem.-Eur. J.*, 2011, **17**, 4307–4312; (c) S. Kindt, H. Jasch and M. R. Heinrich, *Chem.-Eur. J.*, 2014, **20**, 6251–6255; (d) Y.-H. Chen, M. Lee, Y.-Z. Lin and D. Leow, *Chem.-Asian J.*, 2015, **10**, 1618–1621.
- For selected reviews and examples, see: (a) I. Ghosh, L. Marzo, A. Das, R. Shaikh and B. König, *Acc. Chem. Res.*, 2016, **49**, 1566–1577; (b) G. Qiu, Y. Li and J. Wu, *Org. Chem. Front.*, 2016, **3**, 1011–1027; (c) R. Lekkala, R. Lekkala, B. Moku, K. P. Rakesh and H.-L. Qin, *Eur. J. Org. Chem.*, 2019, **2019**, 2769–2806; (d) J. D. Nguyen, E. M. D'Amato, J. M. R. Narayanam and C. R. J. Stephenson, *Nat. Chem.*, 2012, **4**, 854–859; (e) I. Ghosh, T. Ghosh, J. I. Bardagi and B. König, *Science*, 2014, **346**, 725–728; (f) A. J. Boyington, C. P. Seath, A. M. Zearfoss, Z. Xu and N. T. Jui, *J. Am. Chem. Soc.*, 2019, **141**, 4147–4153.
- (a) A. U. Meyer, T. Slanina, A. Heckel and B. König, *Chem.-Eur. J.*, 2017, **23**, 7900–7904; (b) M. Schmalzbauer, I. Ghosh and B. König, *Faraday Discuss.*, 2019, **215**, 364–378; (c) H. Kim, H. Kim, T. H. Lambert and S. Lin, *J. Am. Chem. Soc.*, 2020, **142**, 2087–2092; (d) I. A. MacKenzie, L. Wang, N. P. R. Onuska, O. F. Williams, K. Begam, A. M. Moran, B. D. Dunietz and D. A. Nicewicz, *Nature*, 2020, **580**, 76–80; (e) N. G. W. Cowper, C. P. Chernowsky, O. P. Williams and Z. K. Wickens, *J. Am. Chem. Soc.*, 2020, **142**, 2093–2099.
- J. J. Warren, T. A. Tronic and J. M. Mayer, *Chem. Rev.*, 2010, **110**, 6961–7001.
- G. Brigati, M. Lucarini, V. Mugnaini and G. F. Pedulli, *J. Org. Chem.*, 2002, **67**, 4828–4832.
- (a) H. Fischer, *Chem. Rev.*, 2001, **101**, 3581–3610; (b) A. Studer, *Chem.-Eur. J.*, 2001, **7**, 1159–1164; (c) A. Studer, *Chem. Soc. Rev.*, 2004, **033**, 267–273; (d) D. Leifert and A. Studer, *Angew. Chem., Int. Ed.*, 2020, **59**, 74–108.
- Q. Zhu, E. C. Gentry and R. R. Knowles, *Angew. Chem., Int. Ed.*, 2016, **55**, 9969–9973.



- 17 (a) M. Silvi, E. Arceo, I. D. Jurberg, C. Cassani and P. Melchiorre, *J. Am. Chem. Soc.*, 2015, **137**, 6120–6123; (b) A. Bahamonde and P. Melchiorre, *J. Am. Chem. Soc.*, 2016, **138**, 8019–8030.
- 18 M. Liang, Z.-Y. Wang, L. Zhang, H.-Y. Han, Z. Sun and S. Xue, *Renewable Energy*, 2011, **36**, 2711–2716.
- 19 L. Pause, M. Robert and J.-M. Savéant, *J. Am. Chem. Soc.*, 1999, **121**, 7158–7159.
- 20 (a) C. D. McTiernan, S. P. Pitre and J. C. Scaiano, *ACS Catal.*, 2014, **4**, 4034–4039; (b) E. H. Discekici, N. J. Treat, S. O. Poelma, K. M. Mattson, Z. M. Hudson, Y. Luo, C. J. Hawker and J. R. de Alaniz, *Chem. Commun.*, 2015, **51**, 11705–11708.
- 21 (a) H. Zhang, W. Pu, T. Xiong, Y. Li, X. Zhou, K. Sun, Q. Liu and Q. Zhang, *Angew. Chem., Int. Ed.*, 2013, **52**, 2529–2533; (b) L. Guo, H.-Y. Tu, S. Zhu and L. Chu, *Org. Lett.*, 2019, **21**, 4771–4776.

