



Cite this: DOI: 10.1039/c9sc05507j

All publication charges for this article have been paid for by the Royal Society of Chemistry

Rotaxane Pt^{II}-complexes: mechanical bonding for chemically robust luminophores and stimuli responsive behaviour†

Zhihui Zhang,^a Graham J. Tizzard,^a J. A. Gareth Williams^{id}*^b and Stephen M. Goldup^{id}*^a

We report an approach to rotaxanes in which the metal ion of a cyclometallated Pt^{II} luminophore is embedded in the space created by the mechanical bond. Our results show that the interlocked ligand environment stabilises a normally labile Pt^{II}-triazole bond against displacement by competing ligands and that the crowded environment of the mechanical bond retards oxidation of the Pt^{II} centre, without perturbing the photophysical properties of the complex. When an additional pyridyl binding site is included in the axle, the luminescence of the Pt^{II} centre is quenched, an effect that can be selectively reversed by the binding of Ag^I. Our results suggest that readily available interlocked metal-based phosphors can be designed to be stimuli responsive and have advantages as stabilised triplet harvesting dopants for device applications.

Received 1st November 2019

Accepted 2nd January 2020

DOI: 10.1039/c9sc05507j

rsc.li/chemical-science

Introduction

Many mechanically chelated¹ metal complexes, in which a metal ion bridges mechanically bonded covalent sub-components,² have been synthesised since Sauvage's first report of the Cu^I templated synthesis of catenanes.³ These complexes are typically reported as intermediates in the synthesis of interlocked molecules,⁴ as opposed to being objects of study in their own right.^{5,6} This is despite the potential benefits that embedding the metal ion within the crowded environment of the mechanical bond might bring in terms of the kinetic stabilisation and electronic, magnetic and catalytic properties. Indeed, Sauvage and co-workers highlighted the kinetic stabilisation of a bound Cu^I ion due to a "catenand effect".^{3,7} Recently, we demonstrated that the mechanical bond allows access to complexes of which the non-interlocked equivalents are inaccessible,⁸ including examples with highly distorted coordination geometries reminiscent of the entatic state of enzyme active sites.^{9,10} These reports suggest that interlocked metal complexes in which the mechanical bond alters the stability or properties of the ion are an untapped resource in the development of coordination complexes for a range of

applications, including catalysis, metallo-pharmaceuticals and light harvesting.

Metal-based luminophores have been at the forefront of many recent developments in light-emitting organic electronic materials for various applications.^{10,11} Iridium(III)-based phosphors are now widely incorporated into commercial organic light-emitting devices (OLEDs) to allow the harvesting of triplet excitons that would otherwise be wasted.¹² The strong spin-orbit coupling associated with the heavy metal ion promotes the T₁ → S₀ phosphorescence process, which is strongly forbidden in purely organic materials.¹³ Complexes of other 2nd and 3rd row metal ions are also widely investigated in this context, particularly those of platinum(II).¹⁴

To date, only a limited number of mechanically chelated heavy-metal complexes have been reported. This is perhaps not surprising as most metal-directed passive template syntheses rely on early transition metals,⁵ due to their more favourable ligand exchange kinetics. Furthermore, the small number of passive templates based on Pd^{II},¹⁵ Ru^{II},¹⁶ and Au^I (ref.17) reported are not luminescent.^{18,19} Thus, the limited examples of interlocked molecules containing metal-based luminophores incorporate the metal complex as a substituent of the interlocked scaffold, or as a structural unit. In 2012, Terao and co-workers demonstrated that emissive Pt^{II}-acetylide units embedded in the axle of a rotaxane were insulated from the local environment, leading to no change in emission between solution and solid state.²⁰ More recently, Ma and co-workers demonstrated that the emission of a Pt^{II}-porphyrin complex appended to a rotaxane could be modulated by mechanical motion.²¹

^aChemistry, University of Southampton, Southampton, SO51 5PG, UK. E-mail: s.goldup@soton.ac.uk

^bDepartment of Chemistry, Durham University, Durham, DH1 3LE, UK. E-mail: j.a.g.williams@durham.ac.uk

† Electronic supplementary information (ESI) available: Full experimental details and characterisation data for all novel compounds. CCDC 1961263–1961269. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9sc05507j



Here we describe the synthesis and properties of rotaxane-based Pt^{II} complexes in which the metal ion is embedded within the mechanical bond. We show that the mechanical bond does not significantly perturb the photophysical properties of the metal ion but, pleasingly, the chemical stability of the complexes is greatly enhanced relative to non-interlocked analogues. Furthermore, while studying this stabilising influence, we serendipitously identified a metallorotaxane in which the phosphorescence shows a dramatic and selective enhancement in the presence of Ag^I, through the inhibition of a quenching pathway arising from a pyridyl unit incorporated in the structure.

Results and discussion

Synthesis of mechanically chelated Pt^{II} complexes

Our study was in part inspired by a cyclometallated catenane-based Pd^{II} complex reported in early work by Sauvage and co-workers.²² Rotaxane 1.H (Scheme 1) was synthesised in excellent yield from readily available starting materials²³ using our small macrocycle modification²⁴ of the active template²⁵ Cu-mediated alkyne-azide cycloaddition reaction (AT-CuAAC)²⁶ (see ESI† for details). Treatment of 1.H with PtCl₂(DMSO)₂ under mild conditions failed to give the target cyclometallated complex [Pt(1)]⁺ directly, but led instead to *trans*-Pt(1.H)(DMSO)Cl₂, in which the Pt centre is coordinated only by the triazole of the rotaxane axle (Scheme 1a). Subsequent heating of Pt(1.H)(DMSO)Cl₂ in AcOH led to [Pt(1)]BF₄ in good yield after purification.²⁷ Using these more forcing conditions, it was also possible to obtain this complex directly from 1.H and K₂PtCl₄ in reasonable yield. The formation of the corresponding non-interlocked complex [Pt(2)(3)]BF₄ required a two-step procedure involving initial formation of Pt(2)Cl followed by Ag^I-mediated substitution of the chloride ligand by the triazole (Scheme 1b).²⁸

All novel compounds and intermediates were characterised by NMR spectroscopy and MS (see ESI†). The ¹H NMR spectrum of [Pt(1)]BF₄ clearly shows the expected desymmetrisation of the metallated aromatic ring; individual signals are observed for H_N, H_O and H_P, the first of which is significantly shielded (5.03 ppm). Perhaps counterintuitively, coordination of Pt to the triazole N leads to an upfield shift of H_d as metal coordination interrupts a non-classical H-bonding interaction with the bipyridine Ns. Unusually, whereas H_D and H_E each appear as 4H doublets in the ¹H NMR spectra of 1.H and Pt(1.H)(DMSO)Cl₂,

indicating free rotation of the flanking aromatic ring, H_D and H_E of [Pt(1)]BF₄ each appear as two 1H doublets suggesting this motion is slow on the ¹H NMR timescale, presumably because of the crowded nature of the macrocycle cavity with the large Pt^{II} ion included.

Pt(1.H)(DMSO)Cl₂ (Table S1†) and [Pt(1)]BF₄ (Fig. 1) were further characterised in the solid-state by single-crystal X-ray diffraction (SCXRD). The solid-state structure of [Pt(1)]BF₄ clearly demonstrates the mechanically chelating environment of the metal ion which bridges N³ of the triazole axle and the N^{^N^C} chelate provided by the macrocycle. As is typically observed with small, sterically crowded interlocked molecules,²⁴ several short intercomponent contacts are present in addition to metal-ligand interactions (Fig. 1a), including a CH...Pt contact. Viewing the solid-state structure in space-filling representation (Fig. 1b) underlines the sterically crowded nature of the mechanical bond; the metal ion is embedded in the cavity of the ring and largely isolated from the local environment.

Photophysical properties of metallorotaxane [Pt(1)]BF₄ and [Pt(2)(3)]BF₄

The photophysical properties of the interlocked Pt^{II} complex [Pt(1)]BF₄ and its non-interlocked analogue, [Pt(2)(3)]BF₄, are remarkably similar (Table 1 and Fig. 2), demonstrating that the

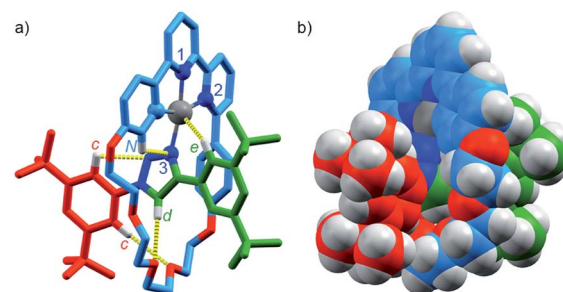
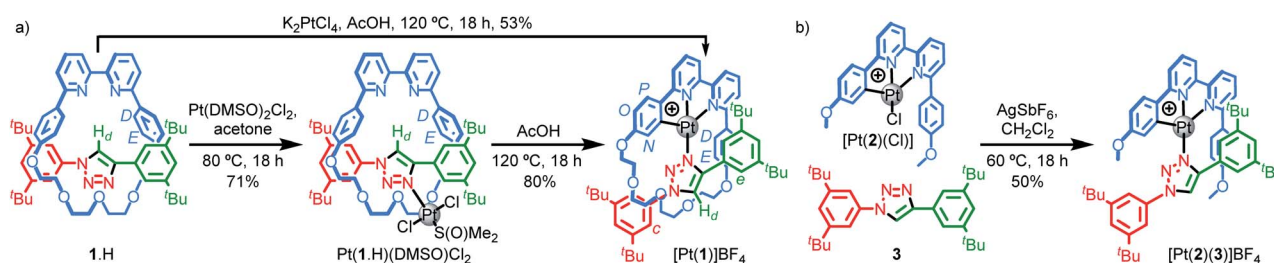


Fig. 1 Solid-state structure of rotaxane [Pt(1)]BF₄ displayed in (a) partial sticks (rotaxane framework)/ball-and-stick (metal and ligand sphere) representation (anions and majority of hydrogen atoms omitted for clarity) and (b) space-filling representation (anions omitted for clarity). Colours: H white, C as in Scheme 1, N dark blue, O red, Pt dark grey. Selected bond lengths (Å) and angles (°): Pt–C 1.99, Pt–N1 1.97, Pt–N2 2.17, Pt–N3 2.05, H_c–N 2.49, H_c–O 2.59, H_d–O 2.53, H_e–Pt 2.74, H_N–N 2.56, C–Pt–N2 160.4, N1–Pt–N3 171.7, C–Pt–N1 81.9, N1–Pt–N2 78.5, N2–Pt–N3 108.4, N3–Pt–C 90.8.



Scheme 1 (a) Synthesis of complexes [Pt(1)]BF₄. (b) Synthesis of complex [Pt(2)(3)]BF₄.



Table 1 Photophysical properties of [Pt(1)]BF₄ and [Pt(2)(3)]BF₄

	[Pt(1)]BF ₄	[Pt(2)(3)]BF ₄
298 K in deoxygenated CH₂Cl₂		
$\lambda_{\text{abs max/nm}}$ ($\epsilon/\text{M}^{-1} \text{cm}^{-1}$)	264 (58 900), 358 (16 400), 441 sh (946)	262 (48 700), 356 (14 700), 448 sh (1120)
$\lambda_{\text{em max/nm}}$	575	582
$\Phi_{\text{lum}} \times 10^{2a}$	2.7	1.9
$\tau/\mu\text{s}$	2.1 [0.46] ^b	2.2 [0.39] ^b
$k_{\text{r}}/10^3 \text{ s}^{-1c}$	13	8.6
$\Sigma k_{\text{nr}}/10^5 \text{ s}^{-1c}$	4.6	4.5
77 K in frozen EPA (2 : 2 : 1 v/v Et₂O-ⁱC₅H₁₂-EtOH)		
$\lambda_{\text{em max/nm}}$	536, 576, 623	521, 559, 603
$\tau/\mu\text{s}$	15.9	19

^a Quantum yield in deoxygenated solution measured using [Ru(bpy)₃]Cl₂(aq) as the standard. ^b Luminescence lifetimes in deoxygenated solution; values in parenthesis refer to air-equilibrated solution. ^c Radiative k_{r} and non-radiative Σk_{nr} rate constants estimated using the approximation that the emissive state is formed with unitary efficiency and thus $k_{\text{r}} = \Phi/\tau$ and $\Sigma k_{\text{nr}} = (1 - \Phi)/\tau$.³²



Fig. 2 Absorption and excitation spectra in CH₂Cl₂ at 298 K (black and dotted purple lines respectively), and emission spectra in CH₂Cl₂ at 298 K and in EPA at 77 K (red and blue lines respectively) for (a) [Pt(1)]BF₄ and (b) [Pt(2)(3)]BF₄. (EPA is defined in Table 1).

mechanical bond does not significantly perturb the pertinent electronic excited states. Both complexes display UV-visible absorption properties typical of cyclometallated Pt^{II} complexes of arylpyridine ligands,²⁹ with intense intra-ligand bands in the far-UV and a somewhat less intense band at lower energy (350–380 nm). Bands in the latter region are typically associated with charge–transfer transitions of mixed $d_{\text{Pt}}|\pi_{\text{Ar}} \rightarrow \pi_{\text{N}^{\wedge}\text{N}}^*$ orbital parentage, that are introduced upon cyclometallation. Notably, there is no such band for the non-cyclometallated Pt(1.H)(DMSO)Cl₂ complex (Fig. S59†).

Excitation of both cyclometallated complexes at 360 nm in degassed solution leads to an emission band centred around 575 nm. The corresponding luminescence lifetimes of around 2 μs in deoxygenated solution are indicative of a formally spin-forbidden phosphorescence process, promoted by the spin-orbit coupling associated with the metal ion. Such phosphorescence is typical of Pt(II) complexes of 6-phenylbipyridine-based ligands, of the form Pt(N[^]N[^]C)X or [Pt(N[^]N[^]C)L]⁺, and assigned to a $^3[d_{\text{Pt}}|\pi_{\text{Ar}} \rightarrow \pi_{\text{N}^{\wedge}\text{N}}^*]$ state.³⁰ The measured quantum yields of around 2–3% under these conditions are also quite typical of such complexes. The phosphorescence is modestly quenched by dissolved molecular oxygen, as evident from the shorter lifetimes observed in air-equilibrated solution. At 77 K,

vibrational structure becomes clearly resolved, with a progression of around 1300 cm^{-1} , quite typical of complexes featuring aromatic ligands. The complexes are also luminescent in the solid state, with similar quantum yields and lifetimes (Table S8†).³¹

Chemical stability of interlocked Pt^{II} complexes

Having confirmed that metallorotaxane [Pt(1)]BF₄ retains the emissive properties typical of Pt(N[^]N[^]C) complexes, we turned to the effect of the mechanical bond on the stability of the metal complex. Triazole ligands are only weakly coordinating³³ and indeed, treatment of the non-interlocked complex [Pt(2)(3)]BF₄ with NBu₄Cl led to decomposition to produce [Pt(2)Cl] and axle 3 through exchange of the monodentate ligand (Scheme 2). In contrast, the interlocked analogue [Pt(1)]BF₄ is stable over at least 5 days under the same conditions. Similarly, treatment of [Pt(2)(3)]BF₄ with ^tBuO₂H led to rapid decomposition to a complex mixture of products (Scheme 2). The same experiment with [Pt(1)]BF₄ led to a slow reaction to give an oxidised Pt^{IV} complex as the major product, which was identified by ESI-MS as [Pt(1)Cl₂]BF₄. The Cl ligands in this compound are presumably provided through oxidative decomposition of CH₂Cl₂ over the extended reaction time (Scheme 2).

These results suggest that the mechanical bond significantly retards displacement of the weakly coordinating triazole ligand, presumably due to steric repulsion of the incoming Cl nucleophile and destabilisation of the expected 5-coordinate intermediate of the presumed associative pathway. The same steric hindrance would also thermodynamically destabilise the product of the reaction with [Pt(1)]BF₄ as the macrocycle cannot readily accommodate both a Pt–Cl moiety and the uncoordinated axle. The mechanical bond also appears to kinetically disfavour oxidation of the Pt^{II} centre, presumably for similar reasons, as the change from Pt^{II} to Pt^{IV} requires the inclusion of two additional ligands within the metal primary coordination sphere.

To examine if the latter effect could be overcome by providing an additional donor atom in the rotaxane framework, we synthesised complex [Pt(κ^4 -4)]BF₄ (see ESI†) and examined





its stability to oxidation. As predicted, when $[\text{Pt}^{\text{II}}(\kappa^4\text{-}4)]\text{BF}_4$ was treated with ^tBuO₂H, ¹H NMR analysis (Fig. S55[†]) revealed that a rapid reaction (<2 h) took place to give a mixture in which the species assigned as $[\text{Pt}^{\text{IV}}(\kappa^5\text{-}4)(\text{OH})](\text{BF}_4)_2$ ($m/z = 616.3$ corresponding to M^{2+}) was the major product. Replacing ^tBuO₂H with H₂O₂ led to a cleaner reaction (Fig. S56[†]) in which the species assigned as $[\text{Pt}^{\text{IV}}(\kappa^5\text{-}4)(\text{OH})](\text{BF}_4)_2$ was the only significant product. This assignment was supported by SCXRD analysis of crystals grown from the crude product mixture which, although of low quality, confirmed the expected connectivity (Table S5[†]).

Serendipitously (*vide infra*), when $[\text{Pt}(\kappa^4\text{-}4)]\text{BF}_4$ was treated with AgSbF₆ without protection from light (Scheme 2c), higher quality crystals were produced and found to correspond to $[\text{Pt}(\kappa^5\text{-}4)(\text{OH})](\text{SbF}_6)_2$ in which the same coordination environment was also observed (Fig. 3b). The structures of $[\text{Pt}(\kappa^4\text{-}4)]\text{BF}_4$ and $[\text{Pt}(\kappa^5\text{-}4)(\text{OH})](\text{SbF}_6)_2$ (Fig. 3) clearly demonstrate the change from a square-planar Pt^{II} unit to a pseudo-octahedral Pt^{IV} complex in which the rotaxane framework provides five of the six donor atoms required to complete the Pt^{IV} coordination sphere (three from the macrocycle and now *two* from the axle).

The corresponding non-interlocked complex, $[\text{Pt}(2)(5)]\text{BF}_4$, also reacted rapidly with ^tBuO₂H to give a mixture of products (Fig. S57[†]), including uncoordinated axle 5, and

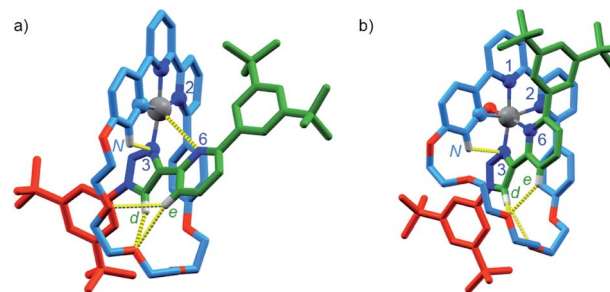


Fig. 3 Solid-state structures of (a) $[\text{Pt}(\kappa^4\text{-}4)]\text{BF}_4$ and (b) $[\text{Pt}(\kappa^5\text{-}4)(\text{OH})](\text{SbF}_6)_2$ in partial sticks (rotaxane framework)/ball-and-stick (metals and ligand sphere) representation (anions and majority of hydrogen atoms omitted for clarity). Colours: H white, C as in Scheme 2, N dark blue, O red, Pt dark grey. $[\text{Pt}(\kappa^4\text{-}4)]\text{BF}_4$ selected distances (Å) and angles (°): Pt–C 1.97, Pt–N1 1.95, Pt–N2 2.17, Pt–N3 2.03, Pt–N6 3.39, H_d–O 2.25, H_e–O 2.79, H_e–O 2.80, H_N–N 2.55, C–Pt–N2 160.7, N1–Pt–N3 172.0, C–Pt–N1 82.1, N1–Pt–N2 78.8, N2–Pt–N3 107.3, N3–Pt–C 91.6. $[\text{Pt}(\kappa^4\text{-}4)(\text{OH})](\text{SbF}_6)_2$ selected distances (Å) and angles (°): Pt–C 2.01, Pt–N1 1.99, Pt–N2 2.22, Pt–N3 2.01, Pt–N6 2.13, Pt–O 1.97, H_d–O 2.49, H_d–O 2.76, H_e–O 2.55, H_N–N3 2.54, C–Pt–N2 160.3, N1–Pt–N3 172.6, N6–Pt–O 171.6, C–Pt–N1 82.4, N1–Pt–N2 77.9, N2–Pt–N3 108.2, N3–Pt–C 91.4.

a species that was assigned as Pt^{IV} complex by ESI-MS analysis ($m/z = 551.2$ corresponding to M^{2+}). Replacing ^tBuO₂H with H₂O₂ simplified the product mixture at short reaction times (Fig. S58[†]). However, whereas $[\text{Pt}^{\text{IV}}(\kappa^5\text{-}4)(\text{OH})](\text{BF}_4)_2$ proved stable over extended reaction times, treatment of $[\text{Pt}(2)(5)]\text{BF}_4$ with H₂O₂ for 24 h led to decomposition to produce significant quantities of non-interlocked axle (Fig. S58[†]), suggesting that even in the Pt^{IV} oxidation state, the Pt–triazole bond remains labile.

Photophysical properties of $[\text{Pt}^{\text{II}}(\kappa^4\text{-}4)]\text{BF}_4$ – response to Ag^I

The pyridine appended rotaxane complex $[\text{Pt}(\kappa^4\text{-}4)]\text{BF}_4$ displays a similar phosphorescence spectrum to that of $[\text{Pt}(1)]\text{BF}_4$. However, its emission quantum yield in solution is dramatically reduced by around 20-fold, and the phosphorescence lifetime is an order of magnitude shorter (Table 2 vs. Table 1). A similar difference was observed between the corresponding non-interlocked complexes $[\text{Pt}(2)(5)]\text{BF}_4$ and $[\text{Pt}(2)(3)]\text{BF}_4$. Estimation of the radiative and non-radiative rate constants, k_r and Σk_{nr} respectively, from the quantum yields and lifetimes suggests that the effect is due primarily to the introduction of an additional non-radiative decay pathway that increases k_{nr} by an order of magnitude when the additional pyridine ring is present. Interestingly, the quenching effect is apparently eliminated in the solid state (Table S9[†]), under which conditions the quantum yields and lifetimes of all of the complexes are similar to one another, suggesting that the quenching process in the pyridine-functionalised structures requires some degree of intramolecular reorganisation that is inhibited in the rigid environment of a solid matrix.³⁴

Examining the solid state structure of $[\text{Pt}(\kappa^4\text{-}4)]\text{BF}_4$, a striking feature is that the lone pair of the axle pyridine moiety is projected near to the Pt centre (3.39 Å) in the ground state. Thus,



Table 2 Photophysical properties of [Pt(κ^4 -4)]BF₄ and [Pt(2)(5)]BF₄

	[Pt(κ^4)]BF ₄	[Pt(2)(5)]BF ₄
298 K in deoxygenated CH₂Cl₂		
$\lambda_{\text{abs max}}/\text{nm}$ ($\epsilon/\text{M}^{-1} \text{cm}^{-1}$)	268 (48 500), 359 (12 900), 445 sh (698)	266 (47 000), 352 (12 900), 443 (1120)
$\lambda_{\text{em max}}/\text{nm}$	591	595
$\Phi_{\text{lum}} \times 10^{2a}$	0.12	0.22
$\tau/\mu\text{s}^b$	0.22 [0.18]	0.10 [0.07]
$k_r/10^3 \text{ s}^{-1c}$	5.5	22
$\Sigma k_{\text{nr}}/10^5 \text{ s}^{-1c}$	45	99
298 K in deoxygenated CH₂Cl₂ with Ag^I (3 equiv.)		
$\lambda_{\text{em max}}/\text{nm}$	581	578
$\Phi_{\text{lum}} \times 10^{2a}$	10	6.0
$\tau/\mu\text{s}^b$	4.3 [0.95]	5.4 [1.0]
$k_r/10^3 \text{ s}^{-1c}$	23	11
$\Sigma k_{\text{nr}}/10^5 \text{ s}^{-1c}$	2.1	1.7
77 K in frozen EPA (2 : 2 : 1 v/v Et₂O-ⁱC₅H₁₂-EtOH)		
$\lambda_{\text{em max}}/\text{nm}$	534, 574, 622	529, 555
$\tau/\mu\text{s}$	19	18

^a Quantum yield in deoxygenated solution measured using [Ru(bpy)₃]Cl_{2(aq)} as the standard. ^b Luminescence lifetimes in deoxygenated solution; values in parenthesis refer to air-equilibrated solution. ^c Radiative k_r and non-radiative Σk_{nr} rate constants estimated using the approximation that the emissive state is formed with unitary efficiency and thus $k_r = \Phi/\tau$ and $\Sigma k_{\text{nr}} = (1 - \Phi)/\tau$.

one explanation for the quenching of the emission is through exciplex formation, where a conformational rearrangement takes place to bring the pyridine lone pair closer to the Pt^{II} centre, which becomes more electrophilic in the ³MLCT excited state. Indeed, it has been noted that Pt^{II} complexes with “dangling” nucleophilic groups are often non-emissive (or only weakly so), ostensibly due to such axial interactions, whilst similar reasoning has been invoked to explain the frequently lower quantum yields of Pt^{II} complexes in solvents incorporating potential donor atoms like DMF.³⁵ McMillin and co-workers have highlighted the correlation between the Gutmann donor number³⁶ and the rate constant for intermolecular quenching of the luminescence of Pt^{II} complexes by donors such as pyridine.³⁷

To probe the role of the axle pyridine in [Pt(κ^4 -4)]BF₄, we attempted to protonate it, in order to determine if its elimination as a potential donor group would enhance the luminescence. However, no change was observed by ¹H NMR spectroscopy (Fig. S61†) when [Pt(κ^4 -4)]BF₄ was treated with trifluoroacetic acid. The apparent lack of basic behaviour may be due to charge–charge repulsion between the formally cationic Pt^{II} centre and the protonated pyridine moiety that destabilises the protonated state. We therefore examined instead the binding of the soft, monovalent cations Cu^I, Ag^I, and Au^I in the hope that this charge–charge repulsion might be compensated for by attractive metal–metal interactions. Pleasingly, when [Pt(κ^4 -4)]BF₄ was titrated with AgSbF₆ or [Cu(MeCN)₄]PF₆, significant changes were observed by ¹H NMR (Fig. 4a), suggesting that Ag^I or Cu^I may bind to the pyridine ring of the axle. No significant changes were observed upon addition of [AuCl(SMe₂)]. Similarly, divalent cations Zn^{II} and Cd^{II} failed to elicit any observable change by ¹H NMR (Fig. S61†).

Titration of [Pt(κ^4 -4)]BF₄ with Ag^I and Cu^I monitored by UV-vis or luminescence spectroscopy revealed significant changes with incremental addition of metal salt (Fig. S61–S70†). Most strikingly, a switch-on of the emission was observed as a result of Ag^I binding but not in the case of Cu^I; in the presence of 3 equiv. of Ag^I, the emission intensity increased by ~25-fold at λ_{max} upon excitation at 400 nm (Fig. 4b). The quantum yield and lifetime of the emission increased by two orders of magnitude, and the corresponding estimated rate constants (Table 2) suggest that the dramatic effect of silver is due to the combined effect of suppressed non-radiative decay and an increase in k_r . The consequent increase in luminescence is strikingly visible to the naked eye (Fig. 4b inset).

SCXRD analysis of crystals obtained from a solution of [Pt(κ^4 -4)]BF₄ and AgSbF₆ in the dark confirmed that the isolated product incorporates a Ag^I ion coordinated to the pyridine ring of the axle (Fig. 4c). This heterodinuclear complex appears to be stabilised both by a metal–metal interaction between the Pt^{II} and Ag^I centres (2.83 Å) and by an η^2 π -interaction between the Ag^I centre and the bipyridine unit (2.54 Å), although it should be noted that in crystals grown under different conditions, only the Ag–Pt contact is maintained (Table S7†). Thus, we tentatively assign the observed increase in the luminescence quantum yield of [Pt(κ^4 -4)]BF₄ in the presence of Ag^I to inhibition of the proposed quenching mechanism involving the pyridine-N \cdots Pt interaction in the excited state, thus decreasing Σk_{nr} . The increase in k_r , albeit a modest one, may reflect more efficient spin–orbit coupling pathways in the dinuclear system, in line with other recent reports of hetero-dinuclear³⁸ and polynuclear³⁹ complexes incorporating Pt^{II}, although these pathways remain poorly understood.

Titration of [Pt(2)(5)]BF₄ with Ag^I revealed an increase in luminescence efficiency, albeit with a lower switch-on effect



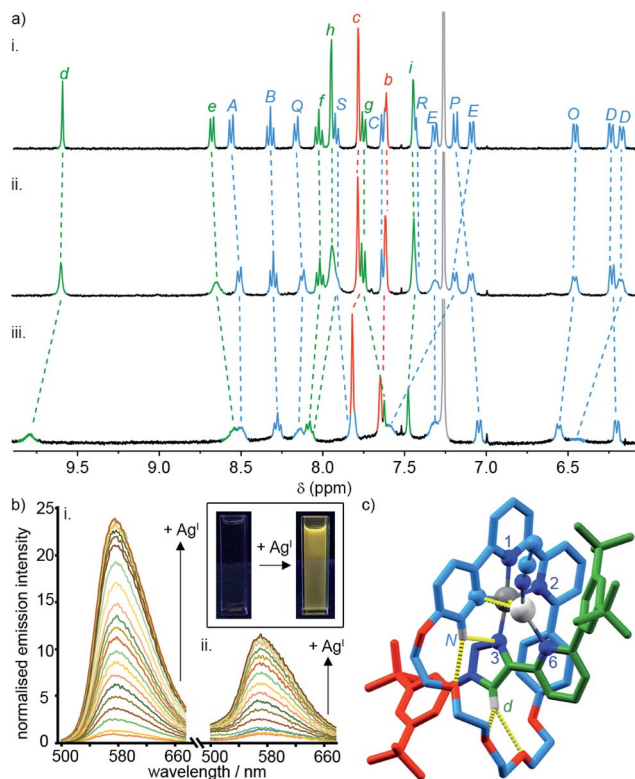


Fig. 4 (a) Partial ^1H NMR of (i) $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4$, (ii) $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4 + 1$ equiv. AgSbF_6 , (iii) $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4 + 1$ equiv. of $[\text{Cu}(\text{MeCN})_4]\text{PF}_6$. (b) Phosphorescence response of (i) $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4$ and (ii) $[\text{Pt}(2/5)]\text{BF}_4$ to the portion-wise addition of AgSbF_6 (inset: visual demonstration of the luminescent switch-on of $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4$ in the presence of 10 equiv. of Ag^+). (c) Solid-state structure of $[\text{Pt}(\kappa^4\text{-4})\text{Ag}(\text{MeCN})]\text{BF}_4$ in partial sticks (rotaxane framework)/ball-and-stick (metals and ligand sphere) representation (anions and majority of Hs omitted for clarity). Colours: H white, C as in Scheme 2, N dark blue, O red, Pt dark grey, Ag light grey. Selected distances (Å) and angles ($^\circ$): Pt–C 1.98, Pt–N1 1.95, Pt–N2 2.17, Pt–N3 2.02, Ag–Pt 2.83, Ag–C 2.54, $\text{H}_d\text{-O}$ 2.28, $\text{H}_d\text{-O}$ 2.60, $\text{H}_N\text{-N3}$ 2.51, $\text{H}_N\text{-O}$ 2.59, C–Pt–N2 161.3, N1–Pt–N3 172.1, C–Pt–N1 82.2, N1–Pt–N2 79.1, N2–Pt–N3 108.4, N3–Pt–C 90.3.

(~12 fold). The difference in behaviour between $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4$ and the corresponding non-interlocked complex may reflect the lower rigidity of the non-interlocked system; the pyridine moiety is sterically constrained in $[\text{Pt}(\kappa^4\text{-4})]\text{BF}_4$ to project towards the Pt^{II} centre, whereas the non-interlocked complex has more conformational freedom which may entropically disfavour the $\text{Pt}^{\text{II}}\cdots\text{N}$ interaction and thus lower the efficiency of the proposed quenching mechanism.

Conclusions

In conclusion, we have demonstrated that interlocked cyclometallated Pt^{II} complexes – which are readily synthesised using an active template strategy – retain the photophysical properties of the parent non-interlocked complexes. The mechanical bond sterically stabilises the metal centre towards ligand displacement and – when no additional donors are present in the structure – towards oxidation, suggesting that this approach may have significant benefits for the construction of more

stable emitters (e.g., as required in applications such as OLEDs, where stability and longevity are crucial). In particular, it allows one or more weakly coordinating ligands to be considered for the optimisation of the luminescence properties, since the chemical integrity of the luminophore is maintained by the mechanical bond rather than relying on the *intrinsic* strength of the metal–ligand interaction.³³ In addition, we serendipitously identified a luminescent sensory system for Ag^+ with >20 fold switch-on, in which the preorganisation provided by the mechanical bond appears to lead to enhanced performance compared with the non-interlocked analogue.

More generally, we have demonstrated that interlocked, mechanically chelated late transition metal complexes can be stabilised by taking advantage of the catenand effect, something that has previously, to our knowledge, only been observed in the case of Cu^{I} .⁷ Just taking the example of platinum alone, such complexes are not only of interest for their photophysical properties but also in cancer chemotherapy – where both Pt^{II} and Pt^{IV} species are known to be effective^{40,41} – and in catalysis.⁴² Our results indicate that it may be possible to use the mechanical bond to augment the stability of key reaction intermediates or even divert the system down an alternative pathway⁴³ in such chemical applications.

Finally, it is also noteworthy that all of the interlocked Pt complexes reported here are examples of mechanically planar chiral rotaxanes – systems in which the mechanical bond acts as a stereogenic unit.⁴⁴ Indeed, ^1H NMR analysis of $[\text{Pt}(\mathbf{1})]\text{BF}_4$ in the presence of the chiral anion “trisphat” revealed the appearance of two sets of signals corresponding to the diastereomeric ion pairs of $R_{\text{mp}}\text{-}[\text{Pt}(\mathbf{1})]\Delta\text{-trisphat}$ and $S_{\text{mp}}\text{-}[\text{Pt}(\mathbf{1})]\Delta\text{-trisphat}$ (Fig. S75†). $[\text{Pt}(\kappa^5\text{-4})(\text{OH})](\text{BF}_4)_2$ also contains a stereogenic Pt^{IV} centre, the configuration of which is dictated by the mechanical stereochemistry of the precursor. Although all of these compounds were formed as racemic mixtures in the present study, we have recently developed methodology to stereoselectively access mechanically planar chiral rotaxanes⁴⁵ and related topologically chiral catenanes.⁴⁶ These results add further perspectives to the potential applications of interlocked heavy metal complexes in, for example, asymmetric catalysis⁴⁷ and circularly polarised luminescence.⁴⁸

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported financially by the China Scholarship Council (studentship to ZZ) and the European Research Council (Consolidator Grant to SMG, agreement no. 724987). The authors thank Dr J. E. M. Lewis (Imperial College) for assistance with SCXRD and helpful discussions.

Notes and references

- 1 By this we mean complexes in which a metal ion bridges donor atoms provided by two or more interlocked covalent



- subcomponents: J. E. M. Lewis, M. Galli and S. M. Goldup, *Chem. Commun.*, 2017, **53**, 298–312.
- 2 It is an open question whether an interlocked molecule in which the components are bridged in this manner are true rotaxanes or if the metal–ligand interactions constitute a covalent link, rendering them entangled but not mechanically bonded, strictly speaking. Sauvage proposed the term “catenate” to denote such complexes in the context of catenanes. The equivalent noun “rotaxanate” is more commonly used as a verb meaning “to make a rotaxane”. Here we use the term “metallorotaxane”, which has seen some use to denote the mechanically chelated complex.^{5a}
- 3 C. O. Dietrich-Buchecker, J. P. Sauvage and J. P. Kintzinger, *Tetrahedron Lett.*, 1983, **24**, 5095–5098.
- 4 Reviews: (a) J. E. Beves, B. A. Blight, C. J. Campbell, D. A. Leigh and R. T. McBurney, *Angew. Chem., Int. Ed.*, 2011, **50**, 9260–9327; (b) J. E. M. Lewis, P. D. Beer, S. J. Loeb and S. M. Goldup, *Chem. Soc. Rev.*, 2017, **46**, 2577–2591.
- 5 Reviews: (a) C. J. Bruns and J. F. Stoddart, *The Nature of the Mechanical Bond: From Molecules to Machines*, Wiley, 2016; (b) N. H. Evans and P. D. Beer, *Chem. Soc. Rev.*, 2014, **43**, 4658–4683; (c) S. Erbas-Cakmak, D. A. Leigh, C. T. McTernan and A. L. Nussbaumer, *Chem. Rev.*, 2015, **115**, 10081–10206.
- 6 Selected examples: (a) D. A. Leigh, P. J. Lusby, A. M. Z. Slawin and D. B. Walker, *Angew. Chem., Int. Ed.*, 2005, **44**, 4557–4564; (b) D. A. Leigh, P. J. Lusby, R. T. McBurney, A. Morelli, A. M. Z. Slawin, A. R. Thomson and D. B. Walker, *J. Am. Chem. Soc.*, 2009, **131**, 3762–3771; (c) A. Noor, W. K. C. C. Lo, S. C. Moratti and J. D. Crowley, *Chem. Commun.*, 2014, **50**, 7044–7047; (d) G. Baggi and S. J. Loeb, *Angew. Chem., Int. Ed.*, 2016, **55**, 12533–12537; (e) G. Baggi and S. J. Loeb, *Chem.–Eur. J.*, 2017, **23**, 14163–14166.
- 7 A. M. Albrecht-Gary, Z. Saad, C. O. Dietrich-Buchecker and J. P. Sauvage, *J. Am. Chem. Soc.*, 1985, **107**, 3205–3209.
- 8 M. Cirulli, A. Kaur, J. E. M. Lewis, Z. Zhang, J. A. Kitchen, S. M. Goldup and M. M. Roessler, *J. Am. Chem. Soc.*, 2019, **141**, 879–889.
- 9 (a) R. J. P. Williams, *J. Mol. Catal.*, 1985, **30**, 1–26; (b) R. J. P. Williams, *Eur. J. Biochem.*, 1995, **234**, 363–381.
- 10 For selected examples and reviews of synthetic complexes designed to display “entatic” behaviour see: (a) P. Comba, *Coord. Chem. Rev.*, 2000, **200–202**, 217–245; (b) G. Chaka, J. L. Sonnenberg, H. B. Schlegel, M. J. Heeg, G. Jaeger, T. J. Nelson, L. A. Ochrymowycz and D. B. Rorabacher, *J. Am. Chem. Soc.*, 2007, **129**, 5217–5227; (c) A. Hoffmann, S. Binder, A. Jesser, R. Haase, U. Flörke, M. Gnida, M. Salomone Stagni, W. Meyer-Klaucke, B. Lebsanft, L. E. Grünig, S. Schneider, M. Hashemi, A. Goos, A. Wetzels, M. Rübhausen and S. Herres-Pawlis, *Angew. Chem., Int. Ed.*, 2014, **53**, 299–304; (d) L. Garcia, F. Cisnetti, N. Gillet, R. Guillot, M. Aumont-Nicaise, J. P. Piquemal, M. Desmadril, F. Lambert and C. Policar, *J. Am. Chem. Soc.*, 2015, **137**, 1141–1146; (e) E. W. Dahl and N. K. Szymczak, *Angew. Chem., Int. Ed.*, 2016, **55**, 3101–3105; (f) B. Dicke, A. Hoffmann, J. Stanek, M. S. Rampp, B. Grimm-Lebsanft, F. Biebl, D. Rukser, B. Maerz, D. Göries, M. Naumova, M. Biednov, G. Neuber, A. Wetzels, S. M. Hofmann, P. Roedig, A. Meents, J. Bielecki, J. Andreasson, K. R. Beyerlein, H. N. Chapman, C. Bressler, W. Zinth, M. Rübhausen and S. Herres-Pawlis, *Nat. Chem.*, 2018, **10**, 355–362.
- 11 (a) Q. Zhao, F. Li and C. Huang, *Chem. Soc. Rev.*, 2010, **39**, 3007–3030; (b) Q. Zhao, C. Huang and F. Li, *Chem. Soc. Rev.*, 2011, **40**, 2508–2524; (c) E. Baggaley, J. A. Weinstein and J. A. G. Williams, *Coord. Chem. Rev.*, 2012, **256**, 1762–1785.
- 12 (a) *Highly Efficient OLEDs with Phosphorescent Materials*, ed. H. Yersin, Wiley, 2007; (b) Y. Chi and P. T. Chou, *Chem. Soc. Rev.*, 2010, **39**, 638–655; (c) L. F. Gildea and J. A. G. Williams, in *Organic Light-Emitting Diodes (OLEDs): Materials, devices and applications*, Woodhead, 2013, pp. 77–113.
- 13 H. Yersin, A. F. Rausch, R. Czerwieńiec, T. Hofbeck and T. Fischer, *Coord. Chem. Rev.*, 2011, **255**, 2622–2652.
- 14 (a) W. Y. Wong and C. L. Ho, *Chem. Soc. Rev.*, 2009, **253**, 1709; (b) J. Kalinowski, V. Fattori, M. Cocchi and J. A. G. Williams, *Coord. Chem. Rev.*, 2011, **255**, 2401–2425; (c) C. M. Che, C. C. Kwok, S. W. Lai, A. F. Rausch, W. J. Finkenzeller, N. Y. Zhu and H. Yersin, *Chem.–Eur. J.*, 2010, **16**, 233; (d) C. Cebrian, M. Mauro, D. Kourkoulos, P. Mercandelli, D. Hertel, K. Meerholz, C. A. Strassert and L. De Cola, *Adv. Mater.*, 2013, **25**, 437–442; (e) S. Q. Huo, J. Carroll and D. A. K. Vezzu, *Asian J. Org. Chem.*, 2015, **4**, 1210; (f) M. C. Tang, A. K. W. Chan, M. Y. Chan and V. W. W. Yam, *Top. Curr. Chem.*, 2016, **374**, 46.
- 15 Selected examples: (a) A. M. Fuller, D. A. Leigh, P. J. Lusby, I. D. H. Oswald, S. Parsons and D. B. Walker, *Angew. Chem., Int. Ed.*, 2004, **43**, 3914–3918; (b) Y. Furusho, T. Matsuyama, T. Takata, T. Moriuchi and T. Hirao, *Tetrahedron Lett.*, 2004, **45**, 9593–9597; (c) A. M. L. Fuller, D. A. Leigh, P. J. Lusby, A. M. Z. Slawin and D. B. Walker, *J. Am. Chem. Soc.*, 2005, **127**, 12612–12619; (d) A.-M. L. Fuller, D. A. Leigh and P. J. Lusby, *Angew. Chem., Int. Ed.*, 2007, **46**, 5015–5019; (e) A.-M. L. Fuller, D. A. Leigh and P. J. Lusby, *J. Am. Chem. Soc.*, 2010, **132**, 4954–4959.
- 16 (a) J. P. Sauvage and M. Ward, *Inorg. Chem.*, 1991, **30**, 3869–3874; (b) P. Mobian, J. M. Kern and J. P. Sauvage, *J. Am. Chem. Soc.*, 2003, **125**, 2016–2017; (c) P. Mobian, J. M. Kern and J. P. Sauvage, *Angew. Chem., Int. Ed.*, 2004, **43**, 2392–2395.
- 17 (a) C. P. McArdle, J. J. Vittal and R. J. Puddephatt, *Angew. Chem., Int. Ed.*, 2000, **39**, 3819–3822; (b) C. P. McArdle, S. Van, M. C. Jennings and R. J. Puddephatt, *J. Am. Chem. Soc.*, 2002, **124**, 3959–3965; (c) S. M. Goldup, D. a. Leigh, P. J. Lusby, R. T. McBurney and A. M. Z. Slawin, *Angew. Chem., Int. Ed.*, 2008, **47**, 6999–7003; (d) R. J. Puddephatt, *J. Organomet. Chem.*, 2015, **792**, 13–24.
- 18 For an example of a cyclometallated pseudorotaxane-based Pt complex see: D. Sooksawat, S. J. Pike, A. M. Z. Slawin and P. J. Lusby, *Chem. Commun.*, 2013, **49**, 11077–11079.



- 19 Examples in which late transition metals are structural components of non-emissive interlocked molecules: (a) M. Fujita, F. Ibukuro, K. Yamaguchi and K. Ogura, *J. Am. Chem. Soc.*, 1995, **117**, 4175–4176; (b) C. P. McArdle, J. J. Vittal and R. J. Puddephatt, *Angew. Chem., Int. Ed.*, 2000, **39**, 3819–3822; (c) A. Hori, A. Akasaka, K. Biradha, S. Sakamoto, K. Yamaguchi and M. Fujita, *Angew. Chem., Int. Ed.*, 2002, **41**, 3269–3272; (d) A. Hori, K. I. Yamashita and M. Fujita, *Angew. Chem., Int. Ed.*, 2004, **43**, 5016–5019; (e) T. J. Burchell, D. J. Eisler and R. J. Puddephatt, *Dalton Trans.*, 2005, 268–272; (f) N. Weisbach, Z. Baranová, S. Gauthier, J. H. Reibenspies and J. A. Gladysz, *Chem. Commun.*, 2012, **48**, 7562; (g) Z. Baranová, H. Amini, N. Bhuvanesh and J. A. Gladysz, *Organometallics*, 2014, **33**, 6746–6749; (h) W. Wang, L.-J. Chen, X.-Q. Wang, B. Sun, X. Li, Y. Zhang, J. Shi, Y. Yu, L. Zhang, M. Liu and H.-B. Yang, *Proc. Natl. Acad. Sci. U. S. A.*, 2015, **112**, 5597–5601; (i) Y. X. Wang, Q. F. Zhou, L. J. Chen, L. Xu, C. H. Wang, X. Li and H. B. Yang, *Chem. Commun.*, 2018, **54**, 2224–2227; (j) X.-Q. Wang, W. Wang, W.-J. Li, L.-J. Chen, R. Yao, G.-Q. Yin, Y.-X. Wang, Y. Zhang, J. Huang, H. Tan, Y. Yu, X. Li, L. Xu and H.-B. Yang, *Nat. Commun.*, 2018, **9**, 3190.
- 20 (a) H. Masai, J. Terao, S. Makuta, Y. Tachibana, T. Fujihara and Y. Tsuji, *J. Am. Chem. Soc.*, 2014, **136**, 14714–14717; (b) H. Masai and J. Terao, *Bull. Chem. Soc. Jpn.*, 2019, 529–539.
- 21 X. Ma, J. Zhang, J. Cao, X. Yao, T. Cao, Y. Gong, C. Zhao and H. Tian, *Chem. Sci.*, 2016, **7**, 4582–4588.
- 22 A. J. Blake, C. O. Dietrich-Buchecker, T. I. Hyde, J. P. Sauvage and M. Schröder, *J. Chem. Soc., Chem. Commun.*, 1989, 1663–1665.
- 23 J. E. M. Lewis, R. J. Bordoli, M. Denis, C. J. Fletcher, M. Galli, E. A. Neal, E. M. Rochette and S. M. Goldup, *Chem. Sci.*, 2016, **7**, 3154–3161.
- 24 (a) H. Lahlali, K. Jobe, M. Watkinson and S. M. Goldup, *Angew. Chem., Int. Ed.*, 2011, **50**, 4151–4155; (b) E. A. Neal and S. M. Goldup, *Chem. Sci.*, 2015, **6**, 2398–2404; (c) E. A. Neal and S. M. Goldup, *Angew. Chem., Int. Ed.*, 2016, **55**, 12488–12493; (d) J. E. M. Lewis, F. Modicom and S. M. Goldup, *J. Am. Chem. Soc.*, 2018, **140**, 4787–4791.
- 25 Reviews: (a) J. D. Crowley, S. M. Goldup, A.-L. Lee, D. A. Leigh and R. T. McBurney, *Chem. Soc. Rev.*, 2009, **38**, 1530–1541; (b) M. Denis and S. M. Goldup, *Nat. Rev. Chem.*, 2017, **1**, 0061.
- 26 (a) V. Aucagne, K. D. Hänni, D. A. Leigh, P. J. Lusby and D. B. Walker, *J. Am. Chem. Soc.*, 2006, **128**, 2186–2187; (b) V. Aucagne, J. Berna, J. D. Crowley, S. M. Goldup, K. D. Hänni, D. A. Leigh, P. J. Lusby, V. E. Ronaldson, A. M. Z. Slawin, A. Viterisi and D. B. Walker, *J. Am. Chem. Soc.*, 2007, **129**, 11950–11963.
- 27 H. Yersin, A. F. Rausch, R. Czerwieńiec, T. Hofbeck and T. Fischer, *Coord. Chem. Rev.*, 2011, **255**, 2622–2652.
- 28 J. H. K. Yip, Suwarno and J. J. Vittal, *Inorg. Chem.*, 2000, **39**, 3537–3543.
- 29 (a) J. Brooks, Y. Babayan, S. Lamansky, P. I. Djurovich, I. Tsyba, R. Bau and M. E. Thompson, *Inorg. Chem.*, 2002, **41**, 3055–3066; (b) S. Fernández, J. Forniés, B. Gil, J. Gómez and E. Lalinde, *Dalton Trans.*, 2003, **2**, 822–830; (c) F. Niedermair, O. Kwon, K. Zojer, S. Kappaun, G. Trimmel, K. Mereiter and C. Slugovc, *J. Chem. Soc., Dalton Trans.*, 2008, 4006–4014; (d) M. Spencer, A. Santoro, G. R. Freeman, Á. Diez, P. R. Murray, J. Torroba, A. C. Whitwood, L. J. Yellowlees, J. A. G. Williams and D. W. Bruce, *Dalton Trans.*, 2012, **41**, 14244–14256; (e) H. Uesugi, T. Tsukuda, K. Takao and T. Tsubomura, *Dalton Trans.*, 2013, **42**, 7396–7403; (f) J. Moussa, G. R. Freeman, J. A. G. Williams, L. M. Chamoreau, P. Herson and H. Amouri, *Eur. J. Inorg. Chem.*, 2016, **2016**, 761–767.
- 30 (a) S. Lai, M. C. Chan, T. Cheung, S. Peng and C. M. Che, *Inorg. Chem.*, 1999, **38**, 4046–4055; (b) T. Cheung and C. Che, *J. Chem. Soc., Dalton Trans.*, 1996, **2**, 1645–1651.
- 31 Non-cyclometallated complex Pt(1.H)(DMSO)Cl₂ displays weak ligand-centred fluorescence and no detectable phosphorescence at room temperature, but phosphoresces in the red region at 77 K (Fig. S59[†]). The long lifetime of 140 μs is consistent with a more ligand-centred formulation to the emissive excited state, with little participation of the metal, reflecting the absence of the Pt–C covalent bond.
- 32 J. R. Lakowicz, *Principles of Fluorescence Spectroscopy*, Springer, 3rd edn, 2006, p. 9.
- 33 (a) G. Aromí, L. A. Barrios, O. Roubeau and P. Gamez, *Coord. Chem. Rev.*, 2011, **255**, 485–546; (b) R. A. S. Vasdev, D. Preston and J. D. Crowley, *Dalton Trans.*, 2017, **46**, 2402–2414.
- 34 For related examples see: (a) M. L. Clarke, *Polyhedron*, 2001, **20**, 151–164; (b) M. Ghedini, A. Golemme, I. Aiello, N. Godbert, R. Termine, A. Crispini, M. La Deda, F. Lelj, M. Amati and S. Belviso, *J. Mater. Chem.*, 2011, **21**, 13434–13444; (c) S. Liu, H. Sun, Y. Ma, S. Ye, X. Liu, X. Zhou, X. Mou, L. Wang, Q. Zhao and W. Huang, *J. Mater. Chem.*, 2012, **22**, 22167–22173.
- 35 S. Chatterjee, J. A. Krause, K. Madduma-Liyanage and W. B. Connick, *Inorg. Chem.*, 2012, **51**, 4572–4587.
- 36 V. Gutmann, *Chimia*, 1977, **31**, 1–7.
- 37 D. K. Crites Tears and D. R. McMillin, *Coord. Chem. Rev.*, 2001, **211**, 195–205.
- 38 G. Turnbull, J. A. G. Williams and V. N. Kozhevnikov, *Chem. Commun.*, 2017, **53**, 2729–2732.
- 39 (a) E. V. Puttock, M. T. Walden and J. A. G. Williams, *Coord. Chem. Rev.*, 2018, **367**, 127–162; (b) S. Culham, P.-H. Lanoë, V. L. Whittle, M. C. Durrant, J. A. G. Williams and V. N. Kozhevnikov, *Inorg. Chem.*, 2013, **52**, 10992–11003; (c) M. Z. Shafikov, R. Daniels, P. Pander, F. B. Dias, J. A. G. Williams and V. N. Kozhevnikov, *ACS Appl. Mater. Interfaces*, 2019, **11**, 8182–8193.
- 40 T. C. Johnstone, K. Suntharalingam and S. J. Lippard, *Chem. Rev.*, 2016, **116**, 3436–3486.
- 41 (a) M. D. Hall and T. W. Hambley, *Coord. Chem. Rev.*, 2002, **232**, 49–67; (b) S. J. Berners-Price, *Angew. Chem., Int. Ed.*, 2011, **50**, 804–805.
- 42 (a) M. L. Clarke, *Polyhedron*, 2001, **20**, 151–164; (b) M. L. Clarke, *Polyhedron*, 2001, **20**, 151–164.
- 43 F. Modicom, E. M. G. Jamieson, E. Rochette and S. M. Goldup, *Angew. Chem., Int. Ed.*, 2019, **58**, 3875–3879.



- 44 (a) N. H. Evans, *Chem.–Eur. J.*, 2017, 1–13; (b) N. Pairault and J. Niemeyer, *Synlett*, 2018, 29, 689–698; (c) E. M. G. Jamieson, F. Modicom and S. M. Goldup, *Chem. Soc. Rev.*, 2018, 47, 5266–5311.
- 45 (a) R. J. Bordoli and S. M. Goldup, *J. Am. Chem. Soc.*, 2014, 136, 4817–4820; (b) M. A. Jinks, A. de Juan, M. Denis, C. J. Fletcher, M. Galli, E. M. G. Jamieson, F. Modicom, Z. Zhang and S. M. Goldup, *Angew. Chem., Int. Ed.*, 2018, 57, 14806–14810.
- 46 M. Denis, J. E. M. Lewis, F. Modicom and S. M. Goldup, *Chem*, 2019, 5, 1512–1520.
- 47 (a) H. Huo, X. Shen, C. Wang, L. Zhang, P. Röse, L. A. Chen, K. Harms, M. Marsch, G. Hilt and E. Meggers, *Nature*, 2014, 515, 100–103; (b) C. Wang, L. A. Chen, H. Huo, X. Shen, K. Harms, L. Gong and E. Meggers, *Chem. Sci.*, 2015, 6, 1094–1100; (c) Y. Zheng, Y. Tan, K. Harms, M. Marsch, R. Riedel, L. Zhang and E. Meggers, *J. Am. Chem. Soc.*, 2017, 139, 4322–4325; (d) N. Hu, H. Jung, Y. Zheng, J. Lee, L. Zhang, Z. Ullah, X. Xie, K. Harms, M. H. Baik and E. Meggers, *Angew. Chem., Int. Ed.*, 2018, 57, 6242–6246; (e) C. Zhang, S. Chen, C.-X. Ye, K. Harms, L. Zhang, K. N. Houk and E. Meggers, *Angew. Chem., Int. Ed.*, 2019, 58, 14462–14466.
- 48 (a) C. Shen, E. Anger, M. Srebro, N. Vanthuyne, K. K. Deol, T. D. Jefferson, G. Muller, J. A. G. Williams, L. Toupet, C. Roussel, J. Autschbach, R. Réau and J. Crassous, *Chem. Sci.*, 2014, 5, 1915–1927; (b) J. R. Brandt, X. Wang, Y. Yang, A. J. Campbell and M. J. Fuchter, *J. Am. Chem. Soc.*, 2016, 138, 9743–9746.

