Current advances in tin cluster chemistry

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This perspective summarizes highlights and most recent advances in tin cluster chemistry, thereby addressing the whole diversity of (mostly) discrete units containing tin atoms. Although being a (semi-)metallic element, tin is in the position to occur both in formally positive or negative oxidation states in these molecules, which causes a broad range of fundamentally different properties of the corresponding compounds. Tin(iv) compounds are not as oxophilic and not as prone to hydrolysis as related Si or Ge compounds, hence allowing for easier handling and potential application. Nevertheless, their reactivity is high due to an overall reduction of bond energies, which makes tin clusters interesting candidates for functional compounds. Beside aspects that point towards bioactivity or even medical applications, materials composed of naked or ligand-protected tin clusters, with or without bridging ligands, show interesting optical, and ion/molecule-trapping properties.

1. Introduction

According to the element tin’s broad range of accessible oxidation states (+IV through –I), compounds of this main group element are multi-variant regarding structures, bonding modes, and reactivities. For organotin compounds, for instance, it is well-known that subtle changes of ligand types and numbers greatly affect bioactivity and toxicity, with monoorganotin species being relatively benign with respect to tin compounds with more than one organic ligand. Inorganic tin compounds, like tin chalcogenides, on the other hand, are semiconductors and significantly less harmful regarding their biological impact. Hence, investigations of tin compounds can be found in essentially all fields of chemical science.

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Currently, one of the most rapidly developing families of tin compounds to be obtained and analyzed in condensed phase are clusters, with their correspondingly wide range of properties. While some of the clusters give answers to fundamental aspects of chemical bonding and inorganic reaction pathways or even mechanisms, others point to new opportunities for nonlinear optics or semiconductor technologies. Clusters with organic ligands furthermore allow for fine-tuning of the biological activity of tin compounds through alterations of the organic moieties.14-16

This perspective article intends to highlight some of the most impressive results reported in this context in recent years and to point towards future developments. For addressing clusters with tin in its highest oxidation states as well as clusters with tin in its lowest oxidation states, the following sections will focus on (i) ligand-free chalcogenido stannate(-metalate) clusters, (ii) organotin(-metal) chalcogenide clusters, and (iii) tin-containing Zintl-type and related clusters. All three types of compounds are prepared from fundamentally different starting materials that typically differ in the oxidation state of the tin atoms. While this choice determines the general type of reaction product, an accurate prediction of reaction products is still almost impossible in cluster chemistry. Yet, upon surveying the field, one recognizes common patterns in structures and bonding, which allows to understand more about these compounds, their formation and their properties. In the meantime, there exist well-established synthesis protocols for reproducible cluster syntheses, indicating that these clearly developed away from purely black-box chemistry.

Without going too much into details of synthesis procedures that can be found elsewhere,12-17 structural, chemical, and physical properties of the selected examples of tin clusters will be named, illustrating the breadth and beauty of this area of research, and also pointing to possible implications of this field for inorganic, organoelement, and materials chemistry.

2. Chalcogenido stannate and metalate clusters

Chalcogenido stannate clusters can be understood as heavier homologues of molecular silicate anions, yet with notable...
structural differences from these as a consequence of larger ionic radii and longer bonds, and with smaller HOMO–LUMO gaps that cause different (photo-)physical properties. The expansion of the clusters’ compositions towards ternary chalcogenido metlate clusters resembles the step from silicates to metalate-silicates, such as zeolites, with similar differences as indicated above, and the additional difference that zeolites never form discrete molecular clusters.

Wet-chemical, solvothermal and flux processes belong to traditional cluster synthesis methods that have been known for several decades.\(^\text{12-17}\) Employed in chalcogenido stannate chemistry, they provide access to various cluster motifs in tetrahedral topologies, commonly known as \(Tn\)-type or \(Pn\)-type clusters.\(^\text{18}\) \(Tn\) \((n=1-6)\) denotes supertetrahedral fragments of the sphalerite-type structure with \(n\) “layers” of simple \((T1\)-type\) tetrahedra \(\{XY_4\}\) being connected via corner-sharing. The smallest member of the penta-tetrahedral \((Pn, n=1, 2)\) family, \(P1\), is based on an “inverse” \(T1\) unit \(\{XY_4\}\), with its faces capped by four “normal” \(T1\) units \(\{XY_4\}\). A vast number of open framework structures in different topologies have been reported, in which supertetrahedral clusters of the same or of different types are linked to form materials with photocatalytic or strong luminescent properties, or ion exchange capabilities.\(^\text{18-20}\) Isolated, ligand-free clusters, in contrast, have remained rare. Two \(T3\)-type clusters \([M_5Sn_2S_{30}]^{10-}\) \((M=Co, Zn; Fig. 1, top left)\), obtained by reactions in water,\(^\text{21}\) and a family of \(P1\)-type clusters \([M_5Sn_2C_{12}]^{10-}\) \((M=Fe, Zn, Co, Mn; Ch=S, Se, Te; Fig. 1, top center)\), obtained from wet-chemical approaches\(^\text{22-28}\) or from \(K_2S\) flux reactions,\(^\text{29}\) show fine-tunable optical properties via alteration of their composition.

Some open-framework compounds, obtained via solvothermal reactions, comprise pseudo-\(Tn\)-type clusters and feature a range of interesting (physico-)chemical properties. A network of the largest examples of (pseudo-)\(Pn\)-type clusters, which was also obtained by means of solvothermal reactions, comprises pseudo-\(P2\)-type clusters \([Cu_4In_{10.9}Sn_{9.1}S_{32}]^{10-}\) \((\text{Fig. 1, top right})\), \([Cu_6In_{12.5}Sn_{17}S_{42}]^{10-}\) \((\text{Fig. 1, top center})\), \([Cu_6In_{12.5}Sn_{17}S_{42}]^{10-}\), \([Cu_4In_{10.6}Sn_{9.2}S_{42.5}]^{10-}\) \((\text{Fig. 1, top right})\), \([Cu_4In_{10.6}Sn_{9.2}S_{42.5}]^{10-}\). The compounds show a relatively rapid photocurrent response and high electrocatalytic oxygen reduction activity \((\text{Fig. 1, center})\). A network based on the pseudo-\(T4\)-type \([In_3Sn_{18}O_{10}S_{34}]^{12-}\) \((\text{Fig. 1, bottom left})\), for instance, exhibits efficient heavy metal ion sequestration capabilities.\(^\text{30}\) The pseudo-\(T6\)-type cluster \([Sn_{42}O_{16}Se_{72}]^{8-}\) \((\text{Fig. 1, bottom right})\), composed of four corner-linked \(T3\)-type \((Sn_{10}O_{4}Se_{20})\) units and an \((Sn_{2}Se_{6})\) unit, forms an extremely robust framework with a narrow band gap.\(^\text{31}\)

Both for isolated clusters and for networks of the latter, it was shown that the cluster size and composition, especially regarding the nature of involved transition metal atoms, has a major influence on the molecular electronic excitation energies and the band gaps of the solid materials, which in turn affect the photocatalytic activity.\(^\text{17-19}\)

The most recent development for the synthesis of chalcogenido stannate and metlate clusters is the application of uncommon reaction media, like surfactants,\(^\text{32-34}\) hydrazine,\(^\text{35,36}\) or ionic liquids.\(^\text{37}\) Ionothermal reaction conditions, that is reaction in ionic liquids under slightly elevated temperatures, provide several advantages, the most important one being the large adjustability of the reaction media’s properties. At the same time, the reaction conditions, including slightly elevated temperatures, provoke the formation of non-classical, highly exceptional structures. Products with extended anionic substructures featuring new network topologies as well as new molecular cluster motifs are obtained. One example for the latter is a giant spherical anion with 192 atoms, \([Ge_{24}Sn_{36}Se_{132}]^{24-}\) \((\text{Fig. 2, top left})\).\(^\text{37}\) This cluster, called “Zeoball”, represents the largest main group element polyanion known to date. Salts comprising this spherical anion are capable of activating halogen–halogen bonds. A very uncommon, pseudo-\(P2\)-type cluster comprising an interstitial Cs\(^+\) cation, \([Cs@Sn_4(Ge,Se_{10})]^{7-}\) \((\text{Fig. 2, top right})\), was obtained under ionothermal conditions upon \textit{in situ} reduction of Sn(v) from SnCl\(_4\) to Sn(u) by a low-valent main group cluster in the presence of \(T2\)-type \([Ge_4Se_{10}]^{1-}\) anions.\(^\text{38}\) Very recently, a selective methylation of terminal chalcogenide ligands was observed to take place in
imidazolium-based ionic liquids, with the (non-innocent) ionic liquid itself acting as a relatively benign methylation agent. The resulting clusters $[\text{Sn}_{10} \text{O}_4 \text{S}_{16}(\text{SMe})_4]^{4-}$ (Fig. 2, bottom left) and $[\text{Sn}_n \text{N}_{12} \text{Se}_{13}(\text{SeMe})_4]^{3-}$ (Fig. 2, bottom right) consequently possess smaller anionic charges than the purely inorganic parent compounds.$^{39}$

3. Organotin chalcogenide clusters and related compounds

In contrast to the chalcoenogen stannate clusters described in the previous section, organotin chalcogenide clusters are usually neutral, hybrid inorganic/organic compounds that consist of an inorganic core comprising tin and chalcogen atoms, which is protected by organic substituents at the tin atoms. Organotin compounds have been used as biocidal agents, for antifouling treatment,* fungicides, acaricides, and antimicrobial agents,$^9$ which has been one of the motivations of extending the research in this field, besides the large variety of structures and the physical properties of such compounds.

Historically, this chemistry was dominated by organotin oxide clusters.$^{41}-^{49}$ Such compounds were obtained by different procedures, including reactions of organotin chlorides with ($^7$Bu$_2$SnO)$_3$ or oxide source, which afforded compounds like the ladder-type cluster $[\text{SnCl(CH}_2)_3 \text{SnCl(CH}_2)_3 \text{SnClO}_3]_4$ (R = CH$_3$SiMe$_3$; Fig. 3, left).$^{40}$ Hydrolysis reactions led to a variety of organotin oxide clusters like $[^7$Bu$_2$(OH)Sn(O)SnPh(CH$_2$Me$_2$SiO)$_2$Sn$^7$Bu$_2$O]_2$ or $[\text{Sn(OH)}_3]_6$ (R = 2,6-(Me$_2$NCH$_2$)$_2$C$_6$H$_3$)$_2$. A rather unusual way of preparing organotin oxide clusters is the reaction of organotin trichlorides with Na$_2$O in liquid ammonia, yielding the adamantane-type cluster $[\text{Sn}_n \text{O}_4]_4$ (R = C(SiMe$_3$)$_3$).$^{49}$ Recent highlights in this field include

$$\text{[(}^{10}\text{BuSn})_2 \text{O}_4 \text{]/(O}^\text{O}_2\text{)]_4}[\text{O}_2\text{C}_6\text{H}_4\text{N}^\text{H}_2]_2,$$

which exhibits significant cytostatic activity towards lung cancer cells,$^{50}$ or the gold complex-decorated $[[\text{MeSn}(\text{dppba})(\text{AuCl})_2]_2]$; (dppba = p-diphenylphosphino benzolate, Ph$_2$P[4-(COO $^\text{C}_6$H$_4$)]) which exhibits multi-color photoluminescence upon incorporation of different aniline derivatives into the voids of its crystal structure.$^{51}$

Organotin chalcogenide clusters, hence containing the heavier homologues of oxygen, are typically obtained by reactions of organotin trichlorides with a chalcogenide source (Ch = S, Se, Te). The first known cluster of this type, $[\text{MeSn}_2 \text{Se}_4]$, was initially prepared in 1903,$^{52}$ structurally characterized in 1968,$^{53}$ and found to possess an adamantane-type inorganic scaffold. A second, “double-decker”-type isomer was found as late as in 2007 for $[[\text{S}_0\text{Me}_2 \text{NCH}_2 \text{CH}_2 \text{Sn}_4 \text{Se}_4]_2]$; the isomeric $[\text{Sn}_4 \text{Se}_4]$ cluster core is stabilized by coordination of $\sigma$-donor ligands to the tin atoms, and a corresponding increase of the coordination number from four to five.$^{54}$ Stepwise condensation reactions of $\text{RSnCl}_3$ and $[\text{Me}_2 \text{Si}_2]_2$ Ch$^{55-56}$ allowed for the isolation of several reaction intermediates, including semicube-type clusters $[[\text{R}_4 \text{SnCl}]_4]$ (R = $\text{C}_6$H$_2$CO(Me); Fig. 2, center), features tin atoms in the rare +III oxidation state. However, some mixed-valence clusters have also been reported, in which Sn(+IV) and Sn(+II) atoms coexist.$^{59}$ Additionally, one example of a mixed-valence cluster, $[[\text{R}_4 \text{Sn}]_6 \text{Sn}_2 \text{Se}_6]_2$ (R = $\text{C}_6$H$_2$CO(Me); Fig. 3, center), features tin atoms in the rare +III oxidation state.$^{60,61}$

The treatment of low valent tin species, such as $R_2$Sn (R = 2,6-(Me$_2$NCH$_2$)$_2$C$_6$H$_3$), bearing sterically demanding organic substituents with elemental chalcogens yields compounds like $[[\text{R}_4 \text{Sn}]_2 \text{SnSe}]_2$ (R = 2,6-(Me$_2$NCH$_2$)$_2$C$_6$H$_3$; Fig. 3, right),$^{62}$ or the mixed-chalcogen compounds $[[\text{RSnCl}]_4]$ (R = 2,6-(Me$_2$NCH$_2$)$_2$C$_6$H$_3$; Ch/Ch’ = S/Se; S/Se; Se/Te).$^{63}$ Clusters of the composition $[[\text{RR’Sn}_2 \text{SnTe}]_2]_2$ (R = N[2,6-(Pr)$_2$C$_6$H$_4$]SiMe$_3$; R’ = 2,4,6-(CH$_2$SiMe$_3$)$_2$C$_6$H$_3$) were also obtained by reaction of RR’SnLi$_2$ with TeCl$_4$. Another synthetic approach starts out from the organotin hydride $\text{RSnH}_3$ (R = 2,6-(2,4,6-(Pr)$_2$C$_6$H$_4$)$_2$C$_6$H$_3$), which is reacted with elemental chalcogens to yield clusters with trichalcogenide bridges.$^{64}$

Organotin chalcogenide clusters can serve as starting ground for follow-up reactions, either by extension of their inorganic core, or by derivatization of their organic substituents. Modifications of the inorganic core are usually achieved by reactions
with transition metal complexes. The first cluster with a respective ternary inorganic core, [(PhSn)₂(CuPhMe₂)₃Se₆], was obtained from [(PhSn)₄S₆], Na₂S and [Cu(CuPhMe₂)Cl].

Some of the structurally most exceptional examples are mixed-valent clusters like [(RSn)₄(Sn₈Cl₂(MPPPh₃)₂)₂] (R = CMe₂CH₂CO(Me); M/Ch = Cu,Ag/S,55 and Cu,Ag/Se67,68), but also large clusters like the isomeric species [(RSn)₁₂Ag₁₄Se₂₅] with R = CMe₂CH₂CO(Me); M/Ch = {CMe₂CH₂C(Me)NNHCO}₂(C₂H₄)₄, {CMe₂CH₂C(Me)NNH}₂−

4. Zintl and related low-valent clusters containing tin atoms

While all clusters that were discussed in the preceding sections contain tin atoms in positive oxidation states, this section will address clusters, in which tin atoms are (formally) negatively charged. Anionic main group element clusters were first mentioned in the literature in 1891 by Joannis, and famously investigated in a systematic fashion by Zintl in the 1930s.
Milestone improvements to their synthesis and characterization followed with the introduction of ethylenediamine as a solvent and cation sequestering agents, like crypt-222. 94–96

Historically, the first stannide clusters were investigated alongside the corresponding lead and germanium clusters. Two structure types are dominant for the polytetrelide cluster anions, namely Tt₄⁻ tetrahedral and Tt₅⁻ cages. 12,97 The latter are fluxional in shape, often adopting capped square antiprismatic or tricapped trigonal prismatic geometries. 98,99 Investigations and follow-up chemistry of the Tt₄⁻ anions are restricted to liquid ammonia as a solvent due to the high charge per atom in these anions and Sn₄⁺ anions still have not been employed to date. 100–103 Room temperature reactivity studies have therefore focused on the reactivity of the larger Tt₅⁻ anions, most commonly accessed through dissolution of the respective Zintl phases A₄Tt₉ (A: alkali metal, most often K; Tt = Ge, Sn, Pb). 104–106 In recent times, new developments included oxidative coupling of such Zintl anions, 107–109 the formation of organic derivatives, 110–112 and the extension and transformation of the inorganic core by formation of intermetalloid or heterometallic clusters that combine main group (semi-)metal atoms with d- or f-block metal atoms M. 12–14,113 Regarding clusters that comprise tin atoms, most results were obtained in the latter area of research, which will therefore be addressed herein.

Most syntheses of such mixed metallic M/Sn clusters from Zintl anions made use of redox-inert d¹⁰ transition metal compounds or such that provide stable complex fragments, prone to simple coordination chemistry. By reacting K₄Sn₉ in en with Cu(Me)₃ the simplest intermetalloid clusters possible, [Cu@Sn₉]³⁻, can be obtained, in which the Cu⁺ ion is incorporated into the [Sn₉]⁻ cage without affecting its structure. 144 Similarly, reactions with [M(CO)₅Me] (M = Cr, Mo, W) afford the heterometallic cluster [Sn₉M(CO)₅]³⁻ with an intact Sn₉⁻ cage acting as a 6 electron ligand to an M(CO)₅ fragment. 115–117 However, Tt₅⁻ cages do not remain intact in all reactions; in fact, they are far more likely to rearrange and form new cluster ions that are only stable in the presence of the respective d- or f-block metal atoms. A process similar to the transformation of Pb₅⁻ into the first intermetalloid cluster, [Pt@Pb₁₂]²⁻, 118 was observed for Sn₉⁻, affording the isoelectronic anion [Ir@Sn₁₂]³⁻. 119 All of these clusters can be interpreted in terms of Wade–Mings' rules. 120–123 Zintl chemists quickly explored this large reaction space, making use of a vast variety of available transition metal compounds and by reintroducing liquid ammonia as a solvent.

Despite the beauty and aesthetics of deltahedral Wade clusters, it should be noted that their formation is expected for group 14 metallide clusters, as Sn⁻ (or its homologues) is isoelectronic to a BH group. Hence, group 14 (semi-)metal anions that deviate from these geometries are more interesting in terms of their electronic structure. Two of the first examples that featured drastically different geometries were the pentagonal prismatic clusters [M@Ge₁₀]³⁻ (M = Fe, Co). 124,125 Later, the structural variety in 10, 12, and 14-vertex intermetalloid clusters of group 14 elements was investigated systematically in a series of quantum chemical studies. 126–128 These studies were sparked by the synthesis and crystallization of the isoelectronic stannide cluster [Fe@Sn₁₀]³⁻, 126 which exhibits yet another cluster geometry that can be considered as an intermediate between the two extremes of a deltahedral and a pentagonal prismatic cage. A very flat potential energy hypersurface was found for this anion that led to severe disorder in the solid state and very poor single crystal data.

Another exceptional cluster geometry is observed in the anion [Ti@Sn₉Ti₄Cp₂]⁴⁻/⁵⁻ (Fig. 6, top left) that was obtained from reactions of K₄Sn₉ with [TiCp₂Cl₂] in liquid ammonia. 119 This study is interesting for a number of reasons. First, this cluster was the first polystannide to comprise hard, Lewis-acidic transition metal cations, which is rare in Zintl cluster chemistry in general; most intermetalloid and heterometallic clusters contain softer, electron rich metal atoms. Second, the cluster comprises a naked, endohedral Ti atom besides {TiCp} and {TiCp₂} fragments, thereby blurring the boundaries between intermetalloid clusters (with the M atoms inside the stannide cage) and heterometallic clusters (with the M atoms being integrated in a joint cluster shell).

Lastly, some intermediates isolated alongside the formation of the product cluster provided real, experimental glimpses into potential cluster formation pathways. Indeed, the mechanisms of cluster growth remain a mostly speculative aspect in this chemistry due to short reaction times and a lack of spectroscopic handles. 130 Very recently, another report was published that again provides some indirect evidence for a possible cluster formation pathway. Reactions of K₄Sn₉ with [(coe)₂Rh(Cl)][coe = cyclooctene] afford a series of Rh-centered polystannide clusters: [Rh@Sn₁₀]³⁻, [Rh@Sn₁₂]³⁻, [Rh₂@Sn₁₇]⁴⁻ and [Rh₂@Sn₂₄]⁵⁻ (Fig. 6, top right). 131 The first three of these anions represent well-known intermetalloid cluster types, even though (drastic) distortions away from idealized geometries are observed. This is most notable in [Rh@Sn₁₀]³⁻, and these...
effects can be attributed to electron back-donation from the endohedral rhodium atom to the outer cage. \([\text{Rh}_3@\text{Sn}_{10}]^{5-}\) represents the largest polystannide cluster known to date. It was obtained via two different synthetic routes, both of which involve heating solutions in which \([\text{Rh}@\text{Sn}_{10}]^{3-}\) anions are present. A closer look at the structure of \([\text{Rh}_3@\text{Sn}_{12}]^{5-}\) shows that it can be viewed as a fusion product of three \([\text{Rh}@\text{Sn}_{10}]^{3-}\) clusters, suggesting that this is at least a plausible reaction pathway. Two further, remarkable structures that were reported some years ago, should also be mentioned: the onion type cluster \([\text{Sn}@\text{Cu}_3@\text{Sn}_{30}]^{12-}\) (Fig. 6, bottom left), which occurs in an intermetallic phase and was referred to as a molecular “bronze”, \(^{132}\) and the largest single-cage deltahedral cluster \([\text{Pd}_2@\text{Sn}_{18}]^{4-}\) (Fig. 6, bottom right).\(^{133,134}\)

Another branch of Zintl chemistry makes use of small, binary anions as building blocks for larger intermetalloid and heterometallic clusters. The high charge of \(\text{Sn}^{4-}\) tetrahedra can be reduced by formal replacement of Sn with group 15 elements, yielding isoelectronic \((\text{T}_{2}\text{PN})^{3-}\) anions. These are soluble in ethylenediamine and can be used in follow-up reactions, just like the homoatomic anions discussed previously. Such reactions afford ternary clusters, as first demonstrated with the synthesis of \([\text{Zn}@\text{Zn}_3\text{Bi}_4]^{6-}\).\(^{135}\) The use of these binary precursors also allowed for the first incorporation of f-elements into polyatomic cages, as observed in \([\text{Eu}@\text{Sn}_6\text{Bi}_8]^{4-}\) (Fig. 7, left), which still has not been achieved with homoatomic tetrélide precursors.\(^{136}\) Since then, this chemistry has successfully been expanded to other elemental combinations, also including \((\text{TrBI})_3^{2-}\) anions \((\text{Tr} = \text{Ga, In, Tl})\).\(^{137-140}\) Reactions with transition metal complexes indicated drastic differences in reactivity and product formation with respect to homoatomic polystannides: reactions of binary \((\text{Sn}_2\text{Bi})^{3-}\) with \([\text{K}(\text{thf})_{0.13}\) \([\text{Co(cod)}]_2\)] (cod = 1,5-cyclooctadiene) and \([\text{[nacnac}][\text{Cu(NCMe)}] \) (nacnac = \([\text{N}((\text{C}_6\text{H}_3\text{tBu})\text{Me})\text{CH}])\) afforded the ternary clusters \([\text{Co}@\text{Sn}_3\text{Bi}_4]^{3-}\) (Fig. 7, right), and \([\text{[CuSn}_9\text{Bi}_4]^{4-}\)\(^{141,142}\)

As the electronic situation in such anions is often very complex, their characterization is bearing new challenges. The mentioned cluster, \([\text{[CuSn}_9\text{Bi}_4]^{4-}\) (Fig. 8, top), is a recent example for such a study.\(^{143}\) The anion is to be viewed as a dimer of two \{\text{CuSn}_5\text{Bi}_3\}^{3-} 9-atom cages with weak contacts between them. Quantum chemical analyses suggested the absence of any significant bonding interactions between the copper atoms, despite a relatively short Cu–Cu distance. Detailed studies of the electronic situations addressed (i) the hypothetic halves of the cluster as well as (ii) the final product upon hypothetic dimerization of the cluster halves versus (iii) the respective combination of two homoatomic 9-atomic tin clusters (Fig. 8, center and bottom).

These indicated large differences between the homoatomic and the heteroatomic system, respectively, and thereby helped to understand the ternary cluster. Thus, studies of binary Zintl anions allow for a significant expansion of polyanionic main group element cluster chemistry and additionally provide further insights into aspects of cluster bonding and their formation.

In the past few years a novel direction of Zintl cluster chemistry, based on (element-)organic derivatives of \(\text{Ge}_9^{4-}\) anions, most often \((\text{Ge}_9\text{R}_3)^{-}\) units, was developed.\(^{144}\) A rich coordination chemistry, based on these derivatives, has evolved since their first preparation.\(^{144-146}\) However, a transfer of this work to the less stable \(\text{Sn}_9^{4-}\) cages has been done with limited success so far.\(^{147}\) The only reported compound that was obtained this way is \([\text{[Hyp]}][\text{Au}][\text{Sn}_9\text{Hyp}_3][\text{Au}][\text{Sn}_9\text{Hyp}_3][\text{Au}][\text{Hyp}]\] (Hyp = hypersilyl), bearing a cluster-based chain, which was most likely formed upon dissociation of \([\text{Sn}_9\text{Hyp}_3]^{2-}\) in the presence of \([\text{[PPh}_3][\text{Au}][\text{SHyp}]]\).\(^{147}\) However, several related compounds were synthesized starting out from low valent

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Fig. 7 From left: molecular structures of the cluster anions \([\text{Eu}@\text{Sn}_{18}\text{Bi}_4]^{4-}\)\(^{156}\) and \([\text{Co}@\text{Sn}_3\text{Bi}_4]^{3-}\)\(^{141}\) The clusters show positional disorder of the main group elements in the cluster shell. The atom distributions depicted here are based on quantum chemical calculations and represent the energetically preferred distribution.

Fig. 8 Top: molecular structure of the cluster anion \([\text{[CuSn}_9\text{Sb}_4]^{4-}\)\(^{142}\) Center and bottom: difference of electron densities of dimer and monomers upon hypothetic dimerization reactions to form \([\text{[CuSn}_9\text{Sb}_4]^{4-}\) or the non-existent \((\text{Sn}_9^{2-})_2\) analogue. Contours are drawn at 0.0016 a u. For black areas, electron density in the dimer is higher than in the monomers, for light areas it is lower (reproduced from ref. 142 with permission from Wiley-VCH).
organotin compounds or low valent tin halides. These clusters typically contain tin atoms in oxidation states 0 and +I (besides a few atoms with oxidation states +II or −I). Consequently, the compounds feature drastically altered properties, including other solubilities and reactivities, as compared to the naked Sn0 anions. Examples are [Sn13R6] (R = N[(2,6-Pr₂C₆H₄)₂SiMe₃]),[146] [Bu₄MeSn₄],[147] [[tripSn]₄]1/2, [[tripSn]₄] (trip = 2,4,6-trisopropylphenyl),[148] [[Snₓ(SiMe₃)₅]₄]⁻,⁻,⁻,⁻ (Fig. 9, left), [[(Me₃Si)₂CH]₂Sn₉(NHC)]⁻ (Fig. 9, right), and [Mes₁₀Sn₁₀].[152]

Conclusions and outlook

Tin clusters belong to the most contemporary family of tin compounds and exhibit a high variability, owing to the large range of potential oxidation states to be adopted by this group (semi)-metal, and due to its potential to bind to most other atoms.

Very obviously, the oxidation state is correlated with a preference for certain coordination environments and consequently structural features. Tin atoms in higher/highest oxidation states prefer tetrahedral, trigonal bipyramidal, or even octahedral coordination, whereas (slightly) negatively charged or formally neutral tin atoms behave more like metals, and are thus more flexible regarding coordination numbers and geometries.

In its extreme oxidation state, +IV, tin is not as oxophilic and not as sensitive to hydrolysis as the lighter homologues silicon and germanium, which is due to the smaller Sn–O bond energy (414 kJ mol⁻¹) as compared to Ge–O (473 kJ mol⁻¹) and Si–O (598 kJ mol⁻¹),[149] which makes handling of its compounds under ambient conditions less problematic. At the same time, all other bond energies to its own substituents are also generally smaller than for the lighter group 14 elements, which allows for a significant reactivity in according reaction mixtures – especially in the less common oxidation states of tin. For this, tin clusters, with their finely tunable electronic properties, should be ideal candidates for bioactive compounds, for catalysts, or for precursors to innovative materials. Despite the high activity in the field, such expansions are yet in their infancy.

Hence, further developments are to be expected – in each of the discussed areas of research, and also in new directions. Maybe, future work will even end up in the co-existence of the extreme versions of Sn(+IV) and Sn(−I) clusters in the same compound, similar to the tetrelide–tetrelate compounds of the lighter homologues silicon and germanium, such as Cs₁₀[Si₅O₉] and Rb₁₄[Ge₄][Si₆O₁₇].[154]

Conflicts of interest

There are no conflicts to declare.

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