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One-step-synthesized D-gluconic acetal-based supramolecular organogelators with effective phase-selective gelation†

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Two effective and one-step-synthesized organogelators based on D-gluconic acetal derivatives have been developed to show phase-selective gelation behaviours towards aromatic solvents from their biphasic mixtures with water. The dominant factors that drive gelation have been studied using FT-IR and temperature-dependent ¹H NMR spectroscopy. Particularly, gelator **GAA-2** in powder form could selectively congeal toluene, benzene and o-xylene at room temperature under mild stirring. Additionally, **GAA-2** could gelate the aromatic solvents within 10 min and the recovery rate of the aromatic solvents could reach about 82% under a certain condition. The benefits of wide source availability, being easy to synthesize, and recyclable performance of the gelator make **GAA-2** ideal for real-world remediation of aromatic solvents.

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Introduction

With the rapid development of industry and the increasing demand for substances in people's daily lives, the pollution of freshwater arising from organic solvent pollutants such as grease, aliphatic hydrocarbons, and aromatic hydrocarbons has become a serious environmental and economic problem, prompting people to develop advanced functional materials to deal with this situation.^{1–4} In recent years, the phase-selective supramolecular gelators (PSSGs) have attracted much attention because they could selectively transform organic solvent phases into gels from the organic solvent/water two-phase systems. The obtained gel would realize the reversible phase transition of gel–sol under certain external stimuli (such as heating, mechanical shearing, *etc.*), which was beneficial to the recycling of gelators and organic solvent.^{5–7} Therefore, many scientists have been committed to finding suitable PSSGs for application research in organic solvent pollution control.^{8–10} The current self-assembly process of PSSGs needed to be achieved by heating–cooling, ultrasonic and mechanical stirring methods.^{11–14} It was obvious that heating–cooling and ultrasonic operation were not feasible in practical situations from

the perspective of technical feasibility. In addition, the co-solvent could also be used to dissolve PSSGs in advance, and then evenly dispersed into the organic phase to form a gel.^{15–19} However, this method always introduces toxic water-soluble solvents (such as tetrahydrofuran, acetonitrile), which may cause secondary pollution and limit the application of PSSGs in practice. Therefore, it is necessary to formulate a more effective strategy to solve the above problems.

An ideal alternative is to spray solid powder PSSGs directly on the surface of the organic phase, and then form a room-temperature gel, which has the advantages of simple operation, safety and environmental protection.^{20–24} For example, K. M. Sureshan *et al.*²⁵ synthesized a series of glucose-based PSSGs that could be used to congeal crude oils from the crude oil/water mixed system in powder form. J. R. Li *et al.*²⁶ prepared a phase-selective organogelator based on diurea, and they claimed that the solid powder of this PSSG could selectively transform gasoline into gels from the gasoline/water mixing system. However, the above-mentioned PSSGs exhibited quite slow gelation time (3 h and 5 h, respectively), which limited their wide application in the actual pollution control. Recently, H. Q. Zeng *et al.*²⁷ reported a wetting powdered PSSGs which could effectively decrease the gelation time, but co-solvent acetonitrile might still leak into the water body if they were employed on a large scale. H. T. Yu²⁸ and J. Song²⁹ found that the formation of the gel could be effectively accelerated by increasing the amount of powdered PSSGs, but the cost was relatively high. Therefore, the development of the more economical and environmentally friendly PSOGs that can be directly applied in solid form and has excellent performance still remains a major challenge and is in great requirement.

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Previously, we have designed a simple phase-selective gelator based on a sorbitol acetal compound that exhibited excellent gelation ability in nonpolar solvents and could selectively congeal the gasoline layer in gasoline–water mixture by adding the co-solvent THF. However, it still failed to gel the crude oils in powder form. Based on this, we herein used inexpensive gluconic acid as raw material to synthesize two gluconic acid acetal compounds in one step, which combined aromatic ring structures with different substituents and polyhydroxy groups. They are expected to be effective PSSGs for the treatment of aromatics solvent contaminants in water body.

Result and discussion

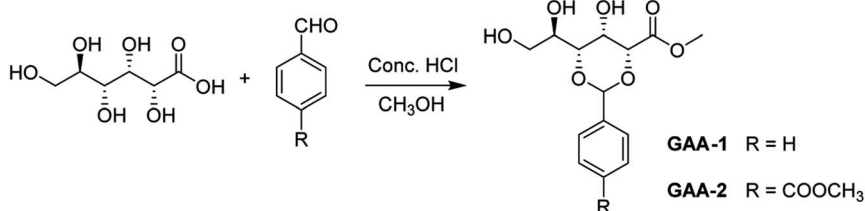
Synthesis and gelation behaviours of the gelators

Firstly, the target gelators could be easily synthesized in one step outlined in Scheme 1. It was worth noting that D-gluconic acid as a raw material (commercially available) reacted with 1 equivalent of the corresponding aldehyde to obtain acetals **GAA-1** and **GAA-2**, respectively. The gelation behaviors of compounds **1** in various solvents was tested by heating–cooling method (Table 1). Obviously, **GAA-2** showed gelation ability in eight organic liquids, including various alcohols (*n*-butanol and *n*-octanol), aromatic solvents (chlorobenzene, nitrobenzene,

toluene, *o*-xylene and benzene) and acetonitrile, but **GAA-1** displayed gelation behaviors just in five of them. In addition, the critical gel concentration (CGC) values for toluene and *o*-xylene were lower than 0.1 wt%, elevating gelator **GAA-2** to the category of supergelator.¹⁵ Moreover, the gelation ability would not change even after multiple repetitions of the sol–gel cycles, indicating that these gels had complete thermal reversibility. Then all the obtained gels could remain stable at room temperature and no obvious changes were observed when stored at room temperature for 2 months, demonstrating the stability of the gels.

Characterization of the gels

To shed light on the self-assembly mechanism of the gelators, we took the gelator **GAA-2** as an example for FT-IR investigations (Fig. 1a). In the chloroform solution, the stretching band of OH was observed at 3489 cm^{−1}, whereas it shifted to 3419 cm^{−1} in the acetonitrile xerogel, indicating that the OH groups were involved in the hydrogen bonding. The stretching vibration peaks of C=C and C–O–C from the methyl ester group appeared at 1724, 1147 cm^{−1}, the corresponding bands of the xerogel appeared at 1726, 1147 cm^{−1}, respectively. These data indicated that the methyl ester groups were involved in the hydrogen bonding. In order to further study the interaction force between



Scheme 1 The synthetic route of gelators **GAA-1** and **GAA-2**.

Table 1 Gelation ability results of gelator **1** in various solvents^a

Solvents	GAA-1			GAA-2		
	State	CGC (wt%)	<i>T</i> _{gel} (°C)	State	CGC (wt%)	<i>T</i> _{gel} (°C)
Hexane	I	—	—	I	—	—
Cyclohexane	I	—	—	I	—	—
<i>n</i> -Octane	I	—	—	I	—	—
Kerosene	I	—	—	I	—	—
Gasoline	I	—	—	I	—	—
Acetonitrile	P	—	—	OG	0.19	52
Chlorobenzene	TG	0.11	78	TG	0.10	75
Nitrobenzene	S	—	—	TG	0.15	81
Toluene	TG	0.16	86	TG	0.09	89
<i>o</i> -Xylene	TG	0.11	91	TG	0.07	99
<i>p</i> -Xylene	TG	0.09	94	TG	0.06	102
<i>m</i> -Xylene	TG	0.13	89	TG	0.09	95
Benzene	TG	0.19	66	TG	0.16	72

^a These tests were performed by adding 30 mg of **1** into 1 mL of a known liquid and then heating the tube until the gelators were dissolved or could not be dissolved ever. A “stable to inversion test tube method” was adopted. I: insoluble, OG: opaque gel, TG: transparent gel. Each experiment was done in duplicate.



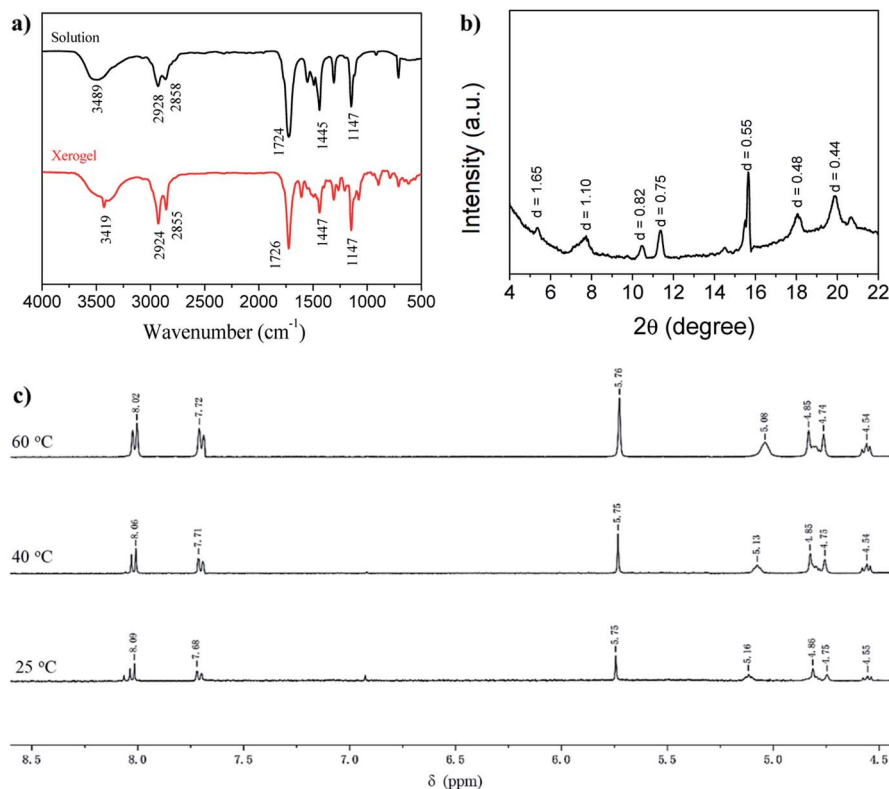


Fig. 1 (a) FTIR spectra of solution (0.1 mM) and xerogel (1.0% w/v) of **GAA-2** in acetonitrile. (b) WAXRD spectra of **GAA-2** xerogels made from acetonitrile. (c) Change in ^1H NMR chemical shift of hydroxyl and aromatic protons of gelator **GAA-2** in CD_3CN at different temperatures.

the gelators, a temperature-dependent ^1H NMR spectroscopy study was performed in acetonitrile (Fig. 1c), and it was found that as the temperature increased from 20 °C to 60 °C, the chemical shift of the hydroxyl proton moved to a high field, which further proves that hydroxyl protons participate in intermolecular hydrogen bonding in the gel phase.⁴ At the same time, the signals of the aromatic protons at 8.09 (Ha) and 7.68 ppm shifted slightly to the high field, to 8.02 and 7.72 ppm, respectively, indicating that the π - π stacking effect is slightly changed because the compound is still in the gel state.²⁴ The

proposed π - π interactions (J-aggregation) can lead to the red shift in the absorption features of the gelator **GAA-2** in solutions with high gelator concentrations as compared with those with low concentrations (Fig. S5†). The XRD pattern of acetonitrile xerogel (Fig. 1b) shows that the d -spacing values are 1.65, 0.82 and 0.55 nm, and the ratio is 1 : 1/2 : 1/3, indicating a layered arrangement.²¹ In addition, scanning electron microscope (SEM) revealed the morphologies of the xerogels prepared from different solvents of **GAA-2**, and well-defined three-dimensional (3D) networks intertwined by fibrous structure can be observed (Fig. 2). These results indicate that the synergistic effect of hydrogen bonding and π - π stacking was the driving forces for the self-assembly of **GAA-2**.

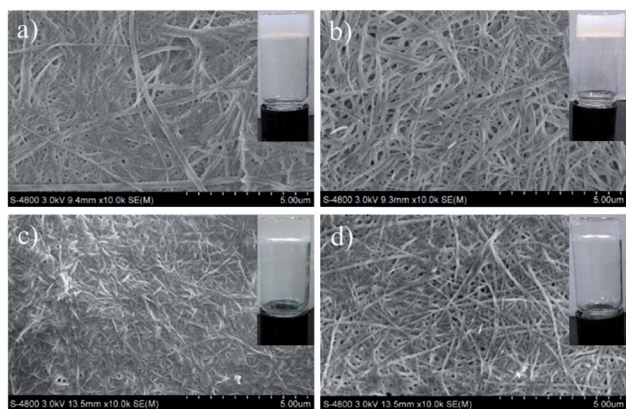


Fig. 2 SEM images of **GAA-2** xerogels obtained from (a) *n*-butanol, (b) acetonitrile, (c) toluene, (d) *o*-xylene.

Phase-selective gelation performances in gelator powder form

Then, we explored the applications of gelators in the fields of room-temperature phase-selective gelation. To our surprise, gelator **GAA-2** has been found to be able to congeal selectively benzene, toluene and xylene, from their water mixtures (0.5 mL/2 mL) with a gentle agitation (Table S1†), whereas **GAA-1** fails to have such behaviours. Furthermore, we carried out the experiment to simulate aromatic solvents recovery in large scale: when 500 mg of **GAA-2** powder was sprinkled into the toluene/water mixture (20 mL/200 mL), after being gently stirred at room temperature for 65 min, a toluene gel was formed which could be easily skimmed off with a net spoon to achieve separation (Fig. 3). Next, toluene could be separated from the gel *via* simple



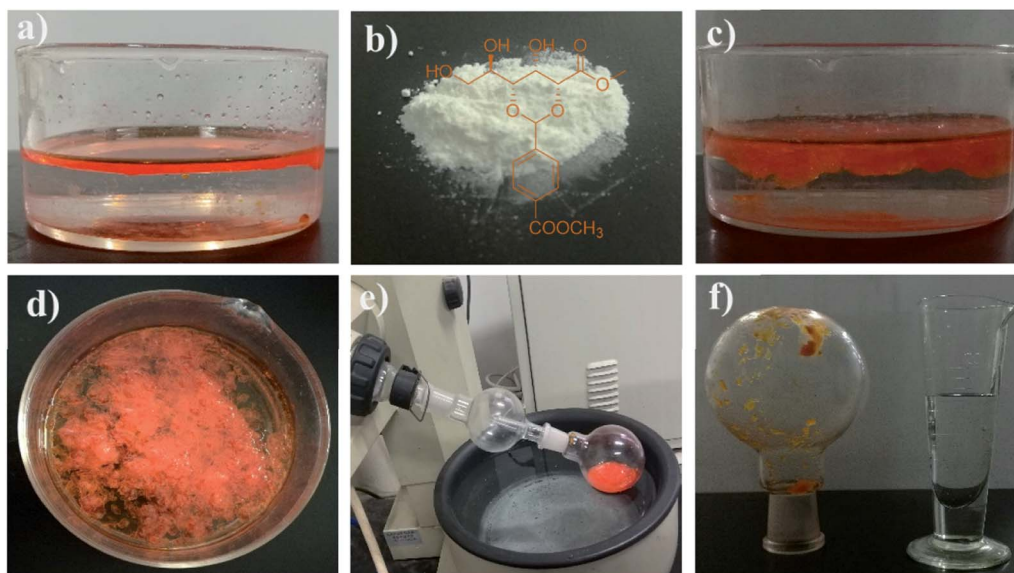


Fig. 3 Demonstration of aromatic solvents recovery. (a) Biphasic system was prepared by adding 20 mL of toluene (dyed with Methyl Red) into 200 mL of water. (b) Powders of GAA-2. (c) 500 mg of gelator powders were dispersed over the solvent surface after being gently stirred at room temperature for 65 min. (d) Formation of a stable gel. (e) Distillation of toluene extracted from the gel. (f) Toluene was recycled from water by GAA-2.

distillation with a recovery of 75% by weight, and the gelator was left in the flask with a recovery of 85% which could also be reused. The phase-selective gelation ability could slow down the diffusion of solvents in water, indicating its potential application in the recovery of actual aromatic solvents.

Given that the amount of the gelator would directly affect its gelation behavior, it was necessary to study the effect of the dosage of the gelator on the gelation time and the recovery rate of the aromatic solvents. In the experiment, the gelation time means the time for the solution turns into a semi-solid that can

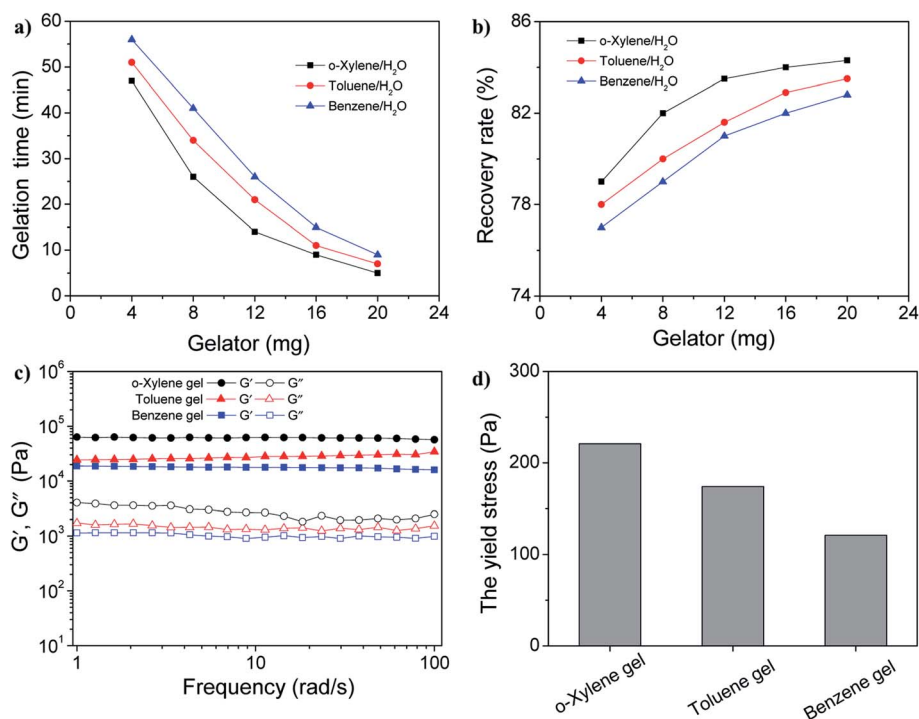


Fig. 4 (a) The gelation time and (b) the recovery rates of gelator GAA-2 in aromatics solvents/water ratios of 1 : 100 (0.1 mL : 10 mL) upon the addition of different amounts of GAA-2. Rheological studies of gels formed by spreading 30 mg of gelator GAA-2 powders over the crude oil surface (0.1 mL) in 10 mL of water. (c) Frequency sweep. G' = storage modulus, G'' = loss modulus. (d) The yield stress of gelator GAA-2 in different aromatics solvents.

be easily removed from water. As shown in Fig. 4a, with the increase in the gelator amount, the gelation time gradually decreases. Moreover, upon the addition of 20 mg of gelator **GAA-2**, the gelation time could be reduced to less than 10 min. In order to calculate the recovery rate, the separated gel was subjected to distillation to recycle the aromatic solvents. It has been found that the recovery rate undergo an enhancement with the increase in the gelator amount, as shown in Fig. 4b. In addition, upon the addition of 20 mg of gelator **GAA-2**, the recovery rates of benzene, toluene and *o*-xylene could reach about 82%. More interestingly, these gels with such a dosage of gelator had excellent mechanical strength, which was implied by its rheological data. For all the gel samples, the storage modulus (G') exhibited much higher values than the loss modulus (G'') in the entire range of frequency sweep, suggesting their gel nature (Fig. 4c). In addition, the high G' values between 2×10^4 – 7.5×10^4 Pa explained the excellent mechanical strength. More importantly, the yield stress values of all the gels were more than 100 Pa, which favours the ability of the gels to withstand high pressure (Fig. 4d and S6†). We then explored the influence of water surface waves on the stability of the thin gel films. The cylindrical gel samples with the same size (10 cm in diameter, 0.2 cm in height) was deposited on the surface of a water bath and the shaking amplitude was slowly increased at a constant frequency until the gel layer broke (Table S2†). For the 4 wt% **GAA-2** in the benzene gel, the amplitude of the waves at which the layer was broken was 0.86 mm at 10 Hz, 0.49 mm at 20 Hz and 0.07 mm at 30 Hz, respectively. The breaking amplitude was increased for toluene gel and *o*-xylene gel, indicating that these gels had higher resistance to water waves compared with benzene gel, which was consistent with the above results of mechanical strength. Once again, no droplets were observed at the breaking amplitude, and the gel film was irreversibly broken into large fragments, which remained on the water surface and could continue to be salvaged. Moreover, the temperature-consecutive frequency spectra for benzene gel, toluene gel and *o*-xylene gel with 3 wt% **GAA-2** were provided in Fig. S7–S9† and demonstrated that an increase in temperature promoted a systematic reduction in G' and G'' over the entire ω range examined. These spectra were then shifted along the frequency axis until they aligned to generate master curves. As shown in the right panels of Fig. S7–S9,† the well-behaved master curves mean that these gels exhibit linear viscoelasticity over a wide frequency range.^{30,31} Therefore, the above results indicated that gelator **GAA-2** could be an efficient and sustainable material for the recovery of aromatic solvents from their water mixtures.

Conclusions

In conclusion, we report two one-step-synthesized organogelators based on the D-gluconic acetal derivatives, which exhibited phase-selective gelation for aromatics solvents. Particularly, **GAA-2** can congeal benzene, toluene and *o*-xylene in powder form at room temperature under mild stirring. Systematic experiments indicate that the gelation time can be realized within 10 min and the recovery rate of the aromatic solvents can reach about 82% under a certain condition.

Besides, the discovery of the PSSGs with the features of wide source, easy to be synthetic, and recyclable performance, lead us to propose gelator **GAA-2** to the real practical use of powdery room-temperature gelators for the treatment of aromatics solvents recovery.

Conflicts of interest

There are no conflicts to declare.

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