


Cite this: *RSC Adv.*, 2020, 10, 41703

In situ synthesis of SO₃H supported Fe₃O₄@resorcinol–formaldehyde resin core/shell and its catalytic evaluation towards the synthesis of hexahydroquinoline derivatives in green conditions

Aliyeh Barzkar  and Alireza Salimi Beni*

A novel spherically shaped core@double-shell acidic nanocatalyst (Fe₃O₄@SiO₂@RF–SO₃H) [RF: resorcinol–formaldehyde resin] was prepared *in situ* and completely characterized using X-ray diffraction, Fourier transform infrared spectroscopy, vibrating sample magnetometry, energy dispersive X-ray spectroscopy, thermogravimetric analysis, transmission electron microscopy and field-emission scanning electron microscopy. The concentration of H⁺ loaded on the Fe₃O₄@SiO₂@RF was reported to be 1.3 mmol g^{−1}. The well-defined Fe₃O₄@SiO₂@RF–SO₃H core–shell heterostructures exhibited high stability, efficient recyclability (10 cycles), and promoted catalytic activity for one-pot condensation reaction between the aromatic aldehydes, dimedone, malononitrile, and ammonium acetate for the synthesis of hexahydroquinoline derivatives.

Received 13th August 2020
Accepted 23rd October 2020
DOI: 10.1039/d0ra06972h
rsc.li/rsc-advances

1. Introduction

Recently, quinolines and their derivatives have attracted considerable attention in organic chemistry due to their pharmacological significance. These compounds can be found in many natural products because of their biological properties. They are also benefited from favorable antibacterial, anti-fungal, antioxidant, anti-cancer and anticonvulsant properties.^{1,2} Quinoline and its derivatives are generally synthesized from four compounds using a one-step reaction of Hantzsch *via* a long cyclocondensation of aldehydes, β-keto ester, ammonia, and relevant catalyst. To improve the production efficiency of quinoline and its derivatives, several catalysts, such as Yb trifluoromethanesulfonate (OTf)₃,³ K₇[PW₁₁CoO₄₀],⁴ *p*-toluenesulfonic acid (*p*-TSA),⁵ HClO₄–SiO₂,⁶ (Sc(OTf)₃),⁷ cerium(IV) ammonium nitrate,⁸ and iron(III) fluoride (FeF₃),⁹ have been developed. All these catalysts suffer from several drawbacks, including difficult separation from the reaction medium, high cost, and the use of toxic organic solvents. An environmentally friendly and one-step synthesis approach, which involves the use of four compounds (*i.e.* aryl aldehydes, dimedone, ammonium acetate, and malononitrile), has been recently developed in the presence of a heterogeneous catalyst.^{10–12}

The advancements in magnetic nanoparticles (MNPs)-based heterogeneous catalysts have opened new horizons for the development of modern science because of their unique features such as cost-effectiveness, non-toxicity and convenient

separation with an external magnetic field. MNPs can be recovered and reused many times, which makes them more economical.^{13,14} MNPs, however, have several weaknesses among which low stability in the acidic and alkaline conditions, easy oxidation under humid atmospheric conditions, and agglomeration can be mentioned.^{15,16} These limitations could be overcome through surface modification with stable materials such as silica, carbon and polymer.^{17–20} The MNPs surface modification approaches can be divided into two categories: coating with inorganic materials (*e.g.* silica and metal oxides) and organic coatings containing carbon and polymers (*e.g.* resorcinol–formaldehyde resin).^{18,21} Resorcinol–formaldehyde resin has excellent characteristics, thus making it an appropriate candidate for stabilization. It can serve as a good precursor for carbon source; moreover, it has high chemical stability and hydrophilicity. To date, the use of resorcinol–formaldehyde resin in the synthesis of core–shell and yolk–shell structures has gained considerable interest.^{22,23} In some cases, the Stöber method was used for the sequential hydrolysis and concentration of organo-silanes²⁴ to prepare the silica colloidal spheres as a reactive linking layer on the magnetic substrates.^{25,26}

The magnetite based core/shell nanostructures have been studied for more than two decades due to their use in various fields such as catalysis,^{27,28} drug carriers,^{29,30} lithium batteries,^{31,32} sensors,^{33,34} and energy storage^{35,36}, adsorbents.^{37–39} The core–shell structures containing magnetite as the core and the resorcinol–formaldehyde (RF), glucose, and dopamine as the shell can be used for substrate stabilization. Resorcinol–formaldehyde have attractive properties such as good stability, high surface

Department of Chemistry, Faculty of Science, Yasouj University, Yasouj 75918-74831, Iran. E-mail: salimibeni@yu.ac.ir; alirezasalimi7173291@gmail.com



area, low cost, remarkable electrical conductivity, controllable structure, and outstanding thermal and mechanical properties.^{40–42} Several core-shell catalysts, including Fe_3O_4 @polydopamine,⁴³ $\text{AgBr}@ \text{SiO}_2@ \text{RF}$,⁴⁴ noble metal@RF,²¹ $\text{Fe}_3\text{O}_4@ \text{RF}/\text{Cu}_2\text{O}$,⁴⁵ $\text{Fe}_3\text{O}_4@ \text{CFR}-\text{PdNPs}$, and $\text{Fe}_3\text{O}_4@ \text{FR}@ \text{graphene-oxide}-\text{PdNPs}$,⁴⁶ have been developed in this category. An effective approach to increase the catalytic activity of these nanocomposites is to stabilize their functional groups by grafting and sol-gel methods. Sulfuric acid grafting on the surface of the core shells has several unique advantages, including large surface area, large pore volume, and easy functionalization. In particular, the core/shell structures with two shells is the best strategy to protect the magnetic core against oxidation and acidic conditions. Furthermore, the presence of acidic groups on the surface can improve its catalytic performance; moreover, sulfuric acid is a good Brønsted acid for the catalysis of several reactions (such as aldol condensations, acylations, nucleophilic additions, and hydrolysis).^{20,47,48}

By increasing active sites in catalysts, especially in heterogeneous catalysts, the number of reactive sites and catalytic efficiency will be increased, thus providing additional active sites in the structure of the catalyst and can save the catalyst amount and the time-consumption. Regarding the importance of magnetic nanoparticles, this study is the first report on nanocatalysts modified by silica and resorcinol-aldehyde with high surface area and double-shell core-shell structure ($\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}-\text{SO}_3\text{H}$). In this manner, the MNPs surface was protected against oxidation, thus making them suitable for use as a stabilizing agent for sulfonic acid. In this case, the modification of the surface of nanoparticles with organic shells, in addition to protecting the magnetic properties of these Fe_3O_4 NPs, increases the surface hydrophobicity due to the presence of organic groups. This hydrophobicity increases the application of core-shell-structured organic coated NPs in catalytic processes. This nanostructure was characterized by various analysis methods; it was then employed as a new and useful acidic nanocatalyst for synthesizing hexahydroquinoline derivatives under solvent-free conditions at 40 °C.

2. Experimental section

2.1. Materials and methods

Chemicals such as tetraethyl orthosilicate (TEOS), resorcinol, formaldehyde, ammonia solution (25–28%), $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$, $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$, ethanol, HCl, malononitrile, dimedone, and all applied aldehydes were purchased from Merck, Fluka, and Aldrich. All solvents were dried and purified before application in the reactions. Purification of reaction products was performed *via* TLC on silica gel polygram SILG/UV 254 plates. The melting points were determined by a Barnstead Electrothermal (BI 9300) apparatus. FTIR spectra were obtained using an FT-IR JASCO-680 spectrometer. NMR spectra were obtained with a Bruker 400 MHz Ultrashield spectrometer at 400 MHz (^1H) and 100 MHz (^{13}C) using CDCl_3 or $\text{DMSO}-d_6$ as the solvent with TMS as the internal standard. Filed-emission scanning electron microscopy (FESEM) analysis was conducted by a Philips, XL30 emission electron microscope. Thermogravimetric analysis

(TGA) was performed by NETZSCH STA 409 PC/PG from room temperature to 800 °C.

2.2. Preparation of Fe_3O_4

Magnetic nanoparticles were prepared as follows: $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ 2 g and $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ 5.2 g was mixed in 25 mL of HCl (1 N). Then, 250 mL of NaOH solution was dropwise added for 20 min under N_2 purging at 80 °C. Finally, the black precipitates were separated by an external magnet, washed several times with distilled water and dried at 40 °C.⁴⁹

2.3. Preparation of Fe_3O_4 core@ SiO_2 shell

To coat a SiO_2 layer around the Fe_3O_4 nanoparticle, the Stöber method was used. In a typical procedure, 0.15 g of as-prepared Fe_3O_4 NPs was monodispersed in 60 mL of distilled water/ethanol (1 : 2) for 30 min. Then, 10 mL of NH_4OH and 2 mL of TEOS was added to the mixture in a round-bottom flask and stirred at room temperature for 6 h. Finally, the crude product was separated by an external magnet and washed several times with water and ethanol and dried at 60 °C for 12 h.⁵⁰

2.4. Preparation of $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}$

Polymeric resorcinol-formaldehyde shell was coated on the $\text{Fe}_3\text{O}_4@ \text{SiO}_2$ surface through the following steps: 0.3 g of $\text{Fe}_3\text{O}_4@ \text{SiO}_2$ nanostructure, 0.9 g of formaldehyde, and 28 mL of distilled water were transferred in a round-bottom flask and fully dispersed under ultrasound waves for 90 min. Then, 1.4 g of resorcinol, 120 mL of ethanol and 0.4 mL of ammonium hydroxide were added to the mixture. To complete the reaction, after 30 min of stirring at 35 °C, 2 mL of formaldehyde solution was added and the mixture was stirred for additional 6 h at room temperature. The reaction mixture was aged at room temperature overnight without complete polymerization. Finally, the product was collected by an external magnet and washed several times by water and ethanol and dried in an oven at 60 °C for 12 h.⁵¹

2.5. Preparation of $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}-\text{SO}_3\text{H}$

The $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}-\text{SO}_3\text{H}$ catalyst preparation involved the replacement of the hydroxyl groups of $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}$ with sulfuric acid groups. For this purpose, 1.0 g of $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}$ core-double shell nanostructure was dispersed by ultrasonic waves into 40 mL of chloroform in a two-neck round-bottom flask. Then, 1 mL of chlorosulfonic acid was dropwise added to the reaction mixture and stirred for 3 h at room temperature under an argon atmosphere. Finally, the magnetic product was separated with an external magnet and washed several times with chloroform and ethanol and dried at 40 °C. All stages for the production of $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}-\text{SO}_3\text{H}$ are presented in Fig. 1.

2.6. Determination of acidity of the $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}-\text{SO}_3\text{H}$ nanocatalyst

The amount of sulfuric acid loaded on the nanostructure surface was calculated by back titration by 0.5 N HCl. To this end, 0.3 g of $\text{Fe}_3\text{O}_4@ \text{SiO}_2@ \text{RF}-\text{SO}_3\text{H}$ nanocatalyst was added into 15 mL of



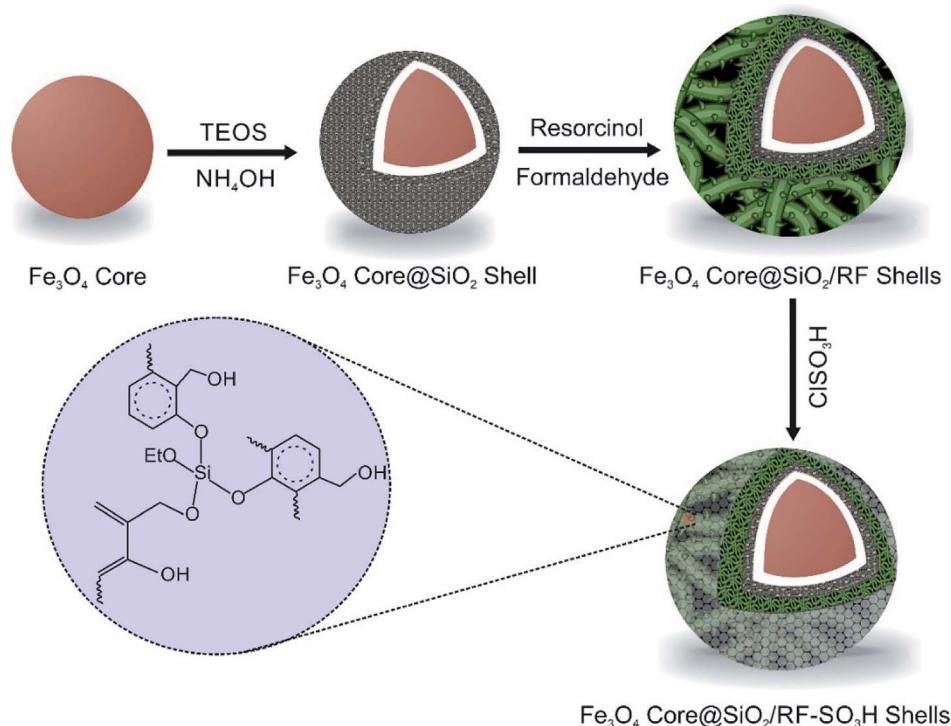


Fig. 1 Illustration for the synthetic steps of $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@RF-SO}_3\text{H}$.

NaOH solution (0.2 N) and stirred at room temperature for 24 h; then, the extracted NaOH was titrated by HCl. According to this experiment, the loading of H^+ was 1.3 mmol g^{-1} .

2.7. General procedure for the preparation of hexahydroquinoline

Aldehydes (1 mmol), dimedons (1 mmol), ammonium acetate (1 mmol), malononitrile (1 mmol), and $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@RF-SO}_3\text{H}$

nanocatalyst (0.9 mol%) were thoroughly mixed in a round-bottom flask and heated to 40°C under continuous stirring. Reaction completion was confirmed by TLC and then the products were dissolved in hot ethanol (8 mL) and $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@RF-SO}_3\text{H}$ was separated by an external magnet. The solvent was then evaporated and crude products were recrystallized and purified in ethanol, followed by IR, ^1H NMR, and ^{13}C NMR characterization.

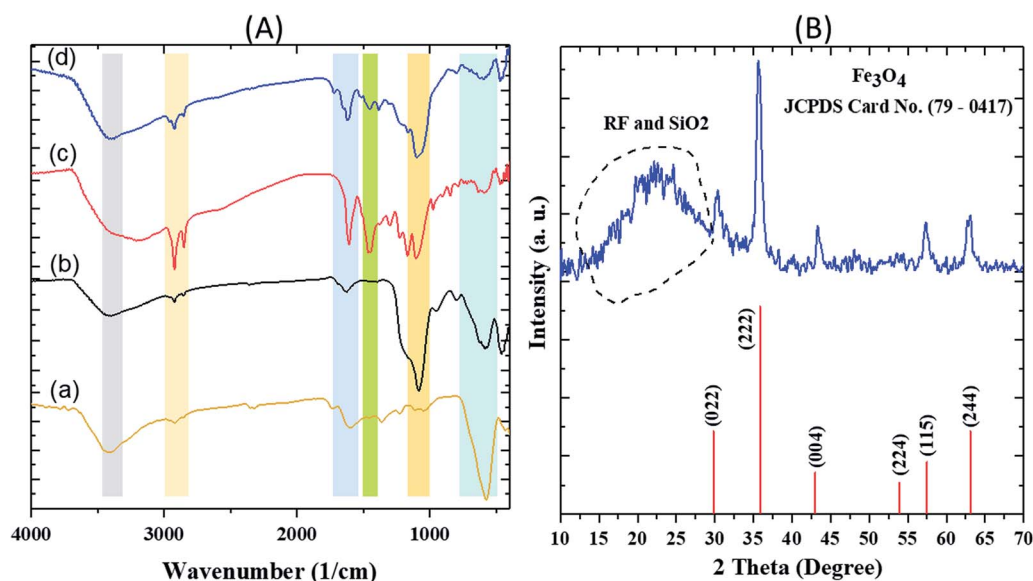


Fig. 2 FT-IR spectra of (A) (a) Fe_3O_4 (b) $\text{Fe}_3\text{O}_4\text{@SiO}_2$ (c) $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@RF}$ and (d) $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@RF-SO}_3\text{H}$ and (B) XRD pattern of $\text{Fe}_3\text{O}_4\text{@SiO}_2\text{@RF-SO}_3\text{H}$ sample.



3. Results and discussions

3.1. Materials characterization

The FT-IR spectra of various steps involved during the synthesis of the nanocatalyst are first described. In the FTIR spectrum of Fe_3O_4 , a strong and evident peak at 585 cm^{-1} can be related to the Fe–O vibration. The vibrations of the hydroxyl group at 1630 cm^{-1} (bending vibration) and 3405 cm^{-1} (anti-symmetric stretching vibration) confirmed the successful preparation of Fe_3O_4 . For $\text{Fe}_3\text{O}_4@\text{SiO}_2$, a wide absorption band at 1083 cm^{-1} is related to the Si–O–Si symmetric stretching vibrations, Si–O bands could be observed in the range of $460\text{--}789\text{ cm}^{-1}$. The vibrations of hydroxyl group related to the SiO_2 at 3434 cm^{-1} are easily observable. In the spectrum of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}$, intensity of the Si–O–Si and Si–O peaks decreased due to the presence of RF polymer. The absorption bands at 1420 cm^{-1} and 1635 cm^{-1} correspond to the aromatic rings in the RF. The

successful stabilization of sulfuric acid groups on the surface of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}$ can be confirmed based on the vibrations at $1000\text{--}1300\text{ cm}^{-1}$ (Fig. 2A).

The XRD pattern of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ is depicted in Fig. 2B. The diffraction peaks at $2\theta = 30.0^\circ, 35.5^\circ, 43.2^\circ, 53.8^\circ, 57.2^\circ$, and 62.6° can be attributed to 022, 222, 004, 224, 114 and 244 crystal planes, respectively, confirming the crystalline structure of the magnetic Fe_3O_4 .⁵² The broadband peak observed in the range of $15^\circ\text{--}30^\circ$ could be assigned to SiO_2 and RF amorphous layers.^{53,54}

FESEM (Fig. 3a) images of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ show the nano spherical shape of the samples and confirm the formation of a quite rough structure. TEM (Fig. 3b) and HRTEM (Fig. 3c) was used to visualize the formation of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ core-shell and $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ core-double-shell nanostructure, respectively. Fig. 3b shows the nanoparticle-assembled cluster core of $\sim 30\text{ nm}$ and uniform carbon shell of $\sim 20\text{ nm}$ around the

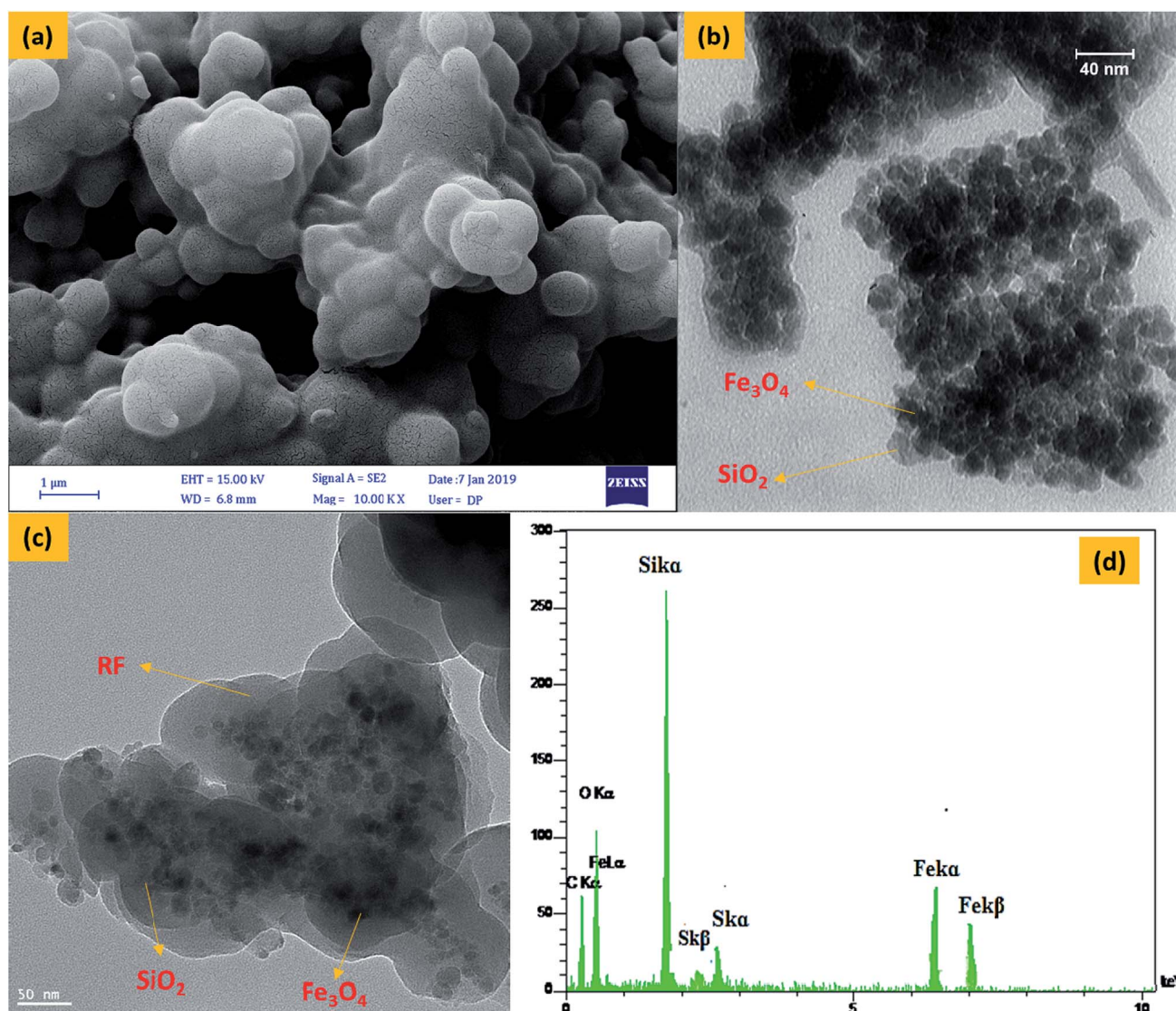


Fig. 3 FESEM image of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ (a), TEM image of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ (b) and $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ (c) and EDS spectra of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ (d).



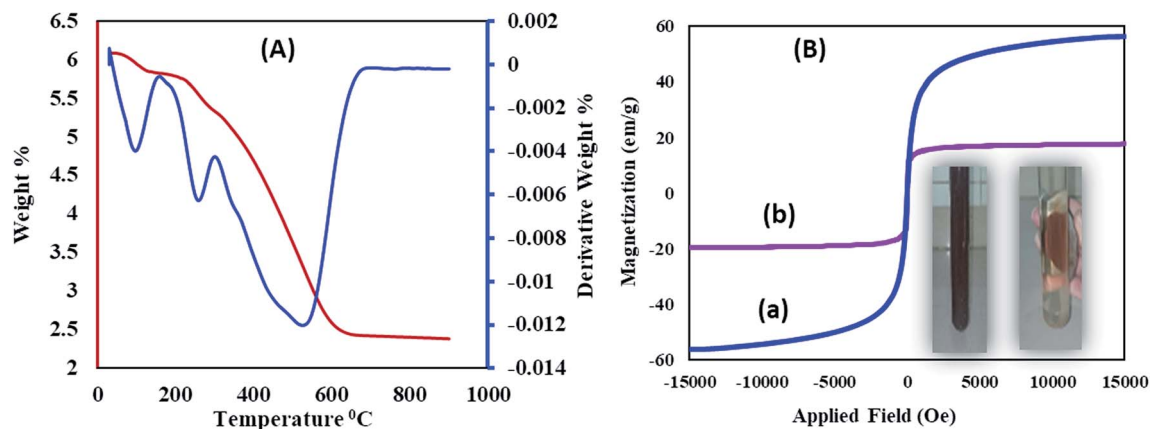


Fig. 4 (A) Thermal gravimetric analysis of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ and (B) VSM analysis of (a) Fe_3O_4 and (b) $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$.

Fe_3O_4 NPs. Subjected to the successive precipitation polymerization through the binding of RF, the magnetic composite nanoparticles were densely encapsulated by the RF of ~ 70 nm (Fig. 3c). A close inspection of the outer shell in Fig. 3c indicates a rough surface with discontinuous coverage on the $\text{Fe}_3\text{O}_4@\text{SiO}_2$ shell. Energy-dispersive X-ray (EDAX) spectra (Fig. 3d) confirms the presence of Fe, Si, O, C and S, thus proving the existence of silica and polymeric layers on the surface of magnetite nanoparticles and the stabilization of acidic groups of sulfuric acid on the core-shell surface.

Thermogravimetric and differential thermal analyses of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ (Fig. 4A) confirm the thermal stability

of the as-prepared sample. The presence of fixed groups can be interpreted as follows: the initial weight loss below 200°C is related to the removal of adsorbed solvents and hydroxyl groups; however, the second weight loss is observed in the temperature range of 200 to 300°C , which reflects the removal of sulfonic acid groups. The last weight loss can be assigned to the thermal decomposition of the nanostructure at 350 – 700°C , thus confirming the stability of the synthesized nanostructure.

Vibrating sample magnetometer (VSM) curve of (Fig. 4B) revealed that the saturation magnetization of Fe_3O_4 nanoparticles was approximately 60 emu g^{-1} , which decreased to 17.05 emu g^{-1} after $\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ coating. This decrease in

Table 1 Optimization conditions of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ catalyst amount, solvent type and temperature on blank reaction^a

Entry	Solvent	Catalyst loading (mol%)	<i>T</i> (°C)	<i>t</i> [min]	Yield ^b [%]
1	—	—	40	240	—
2	—	0.18	40	35	>90%
3	—	0.54	40	25	>95%
4	—	0.9	40	15	>98%
5	—	1.8	40	10	>98%
6	—	0.9	RT	50	60%
7	—	0.9	30	45	90
8	—	0.9	60	10	>98%
9	H ₂ O	0.9	40	25	90
10	EtOH	0.9	40	25	93%
12	Toluene	0.9	40	35	50%
13	CH ₃ CN	0.9	40	40	58%
14	—	$\text{Fe}_3\text{O}_4@\text{SiO}_2^c$	40	60	—
15	—	$\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}^d$	40	60	—

^a Reaction conditions: benzaldehyde (1 mmol), dimedone (1 mmol), malononitrile (1 mmol), ammonium acetate (1 mmol) and $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst (0.9 mol%). ^b Isolated yields. ^c 0.02 g of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ was used as a catalyst. ^d 0.02 g of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}$ was used as a catalyst.



the saturation magnetization could be due to the formation of silicone and polymer coatings on the surface of Fe_3O_4 nanoparticles, which confirms the successful stabilization of the acidic groups on the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}$ surface. An image for the collection of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ catalyst by an external magnet is in the inset of Fig. 4B. Magnetic property is one of the best and most effective advantages of this novel nanocatalyst, which contributes in its easy separation and cost-effectiveness.

3.2. Optimization of hexahydroquinoline synthesis using $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$

The catalytic reaction for the hexahydroquinoline preparation was performed in the presence of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$. Before the synthesis of these derivatives, the ideal conditions were optimized in the blank reaction. To identify the best reaction conditions, the reactions of benzaldehyde with malononitrile, dimedons, and ammonium acetate were considered as blank reactions. The effects of various parameters such as catalyst loading, solvent type, and temperature were then optimized (Table 1). The reaction took a long time (4 h) in the absence of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$, which confirms the importance of the presence of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$. A increase in the amount of catalyst from 0.18 to 0.90 mol% led to an enhance in the product yield; however, its additional increase from 0.9 to 1.8 mol% did not change the yield (Table 1, entries 1–5). Therefore, 0.9 mol% of the catalyst was selected as the optimum catalyst amount (Table 1, entry 4). Several solvents were used to identify the optimal solvent type (Table 1, entries 9–13). It was observed that nonpolar solvents such as toluene and acetonitrile resulted in lower yields compared to polar and protic solvents (e.g. water and ethanol). In the solvent-free condition, 98% yield was observed (Table 1, entries 4). The reaction temperature was studied for selecting the appropriate temperature conditions (room temperature, 30, 40, and 60 °C). A comparison of obtained yields showed that the best yield can be achieved at 40 °C; therefore, this temperature was selected as the optimum temperature conditions (Table 1, entries 4–8). To show the neat effect of the supported sulfuric acid sites during this catalytic process, the catalytic activity of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ and $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}$ nanomaterials in the model reaction was studied; the results were compared with those of the present catalyst. Interestingly, the results showed that in the presence of both sulfuric acid and free nanomaterials, no yield of the desired product was obtained under the same conditions and time as that of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$, thus confirming that the synthesis of hexahydroquinoline process is primarily catalyzed by the supported $-\text{SO}_3\text{H}$ moieties.

3.3. Synthesis of hexahydroquinoline derivatives at the presence of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$

After optimizing the blank reaction conditions, certain benzaldehyde derivatives with electron-donating groups and electron-accepting groups were studied. The results indicated a short reaction time and high efficiency for both types of groups. This indicates that the position and type of substitution did not have a significant impact on the reaction process, thus confirming

Table 2 Synthesis of hexahydroquinoline derivatives in the presence of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst at 40 °C under solvent-free condition^a

Entry	R	Time	Yield ^b (%)	MP	Ref.
1	H	15	95	285–286	286 (ref. 12)
2	4-CH ₃	18	92	290–292	294–295 (ref. 12)
3	4-OCH ₃	20	91	290–291	289–293 (ref. 12)
4	2-OCH ₃	21	91	288–289	—
5	2-OH	22	90	290–291	—
6	4-Isopropyl	20	93	289–290	—
7	4-NO ₂	12	98	280–292	290–293 (ref. 12)
8	3-NO ₂	13	97	282–284	282–283 (ref. 12)
9	4-CN	13	97	287–290	—
10	4-Cl	14	92	289–291	290–291 (ref. 10)
11	2,4-Cl	13	94	290–293	—
12	4-Br	13	92	296–298	295–296 (ref. 10)
13	3-Br	12	91	292–294	293–294 (ref. 12)

^a Reaction conditions: benzaldehyde (1 mmol), dimedone (1 mmol), malononitrile (1 mmol), ammonium acetate (1 mmol) and $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst (0.9 mol%). ^b Isolated yields.

the effective performance of the new $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst for synthesizing hexahydroquinoline derivatives (Table 2). FTIR, ¹H NMR, and ¹³C NMR spectral data were used to identify the structure of the product.

3.4. Reusability and recovery of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$

The recovery and reusability test is a simple and easy approach to assess the heterogeneity and performance of a catalyst. To study the recycling ability and reusability of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst, the reaction between benzaldehyde, dimedone, malononitrile, and ammonium acetate was performed under optimum conditions. After completion of the reaction (detected by TLC), the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ catalyst was collected by an external magnet and washed with hot ethanol. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ was again used in the same reaction under the same conditions for 10 cycles and did not show any efficiency variations (see Fig. 5).

3.5. Comparisons with literature

The performance and efficiency of the synthesized acidic $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst was compared with the other

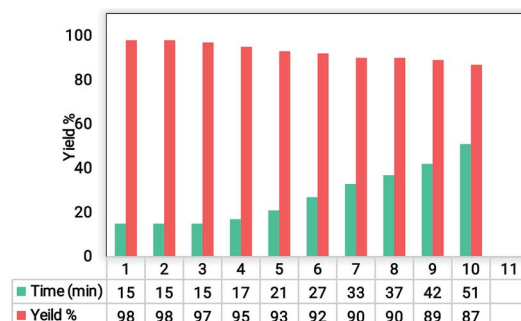


Fig. 5 Recoverability and reusability results of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst.



Table 3 Comparison of catalytic activity of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ nanocatalyst with several known catalysts in the synthesis of hexahydroquinoline derivatives

Entry	Catalyst	Solvent	Time (min)	Reaction cycles	Yield%	Ref.
1	Melamine trisulfonic acid (MTSA)	—	180	4	94	55
2	$n\text{-Fe}_3\text{O}_4@\text{TDI}@\text{TiO}_2$	—	20	6	92	12
3	Nano- $\text{ZrO}_2-\text{SO}_3\text{H}$	—	16	5	94	10
4	Nano- Fe_3O_4	EtOH	12	—	91	56
5	$\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$	—	15	10	>98	This work

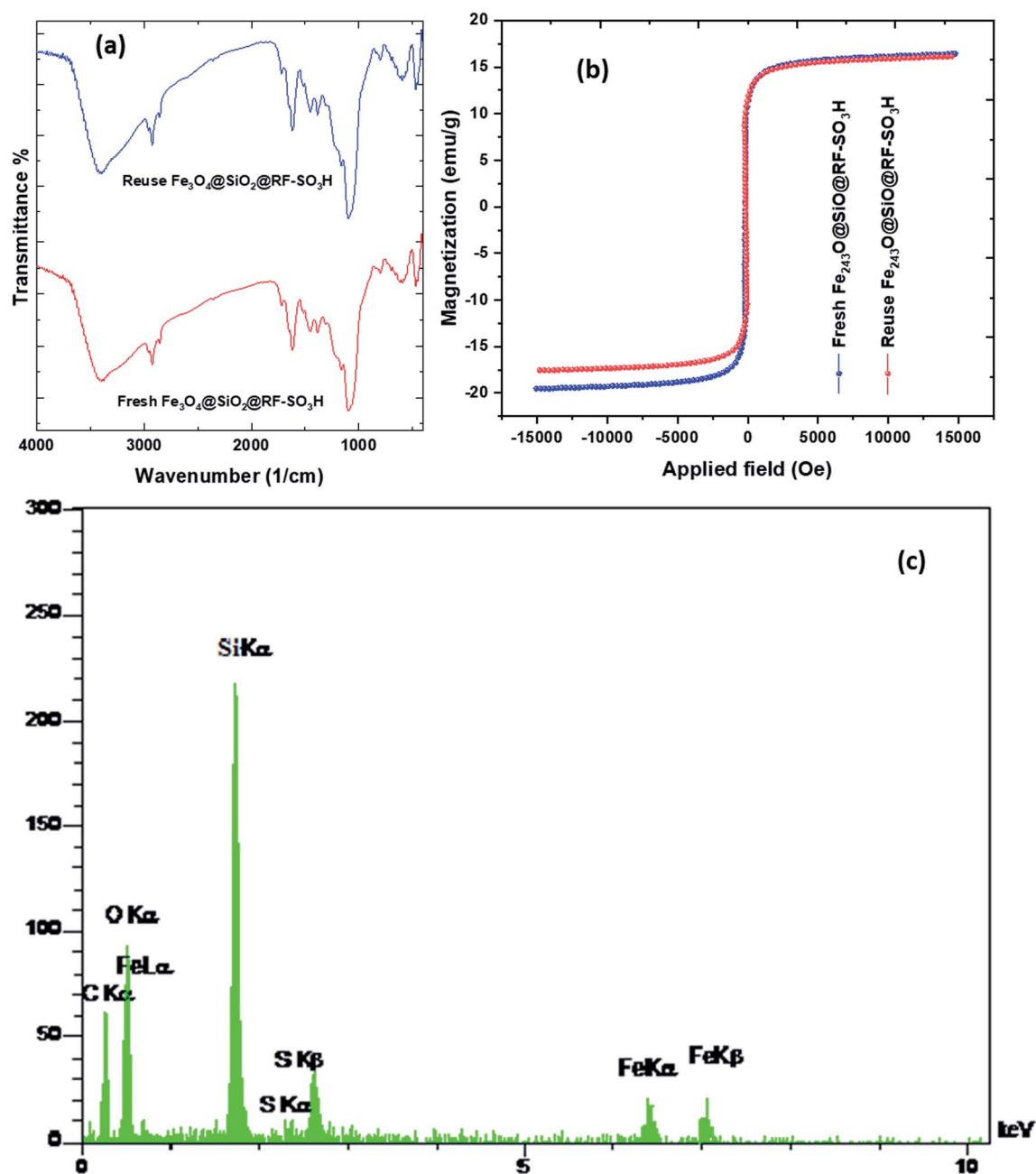


Fig. 6 FTIR spectra (a) and VSM analysis (b) of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ after and before used as the catalyst and EDS spectrum of the recovered $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}-\text{SO}_3\text{H}$ after seventh reaction cycle and (c).



synthesized catalysts in Table 3. Accordingly, all previously reported catalysts were among the components of our blank reaction. All of them required higher temperatures, took longer time to react, and had lower yields and reaction cycles than $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$. The superior performance of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ nanocatalyst, (*e.g.* its shorter reaction time, solvent-free conditions, high efficiency, and higher response cycle) can be assigned to its ease of recyclability, structural features (*e.g.* high surface area, surface uniformity), and high stability of functionalized acid groups.

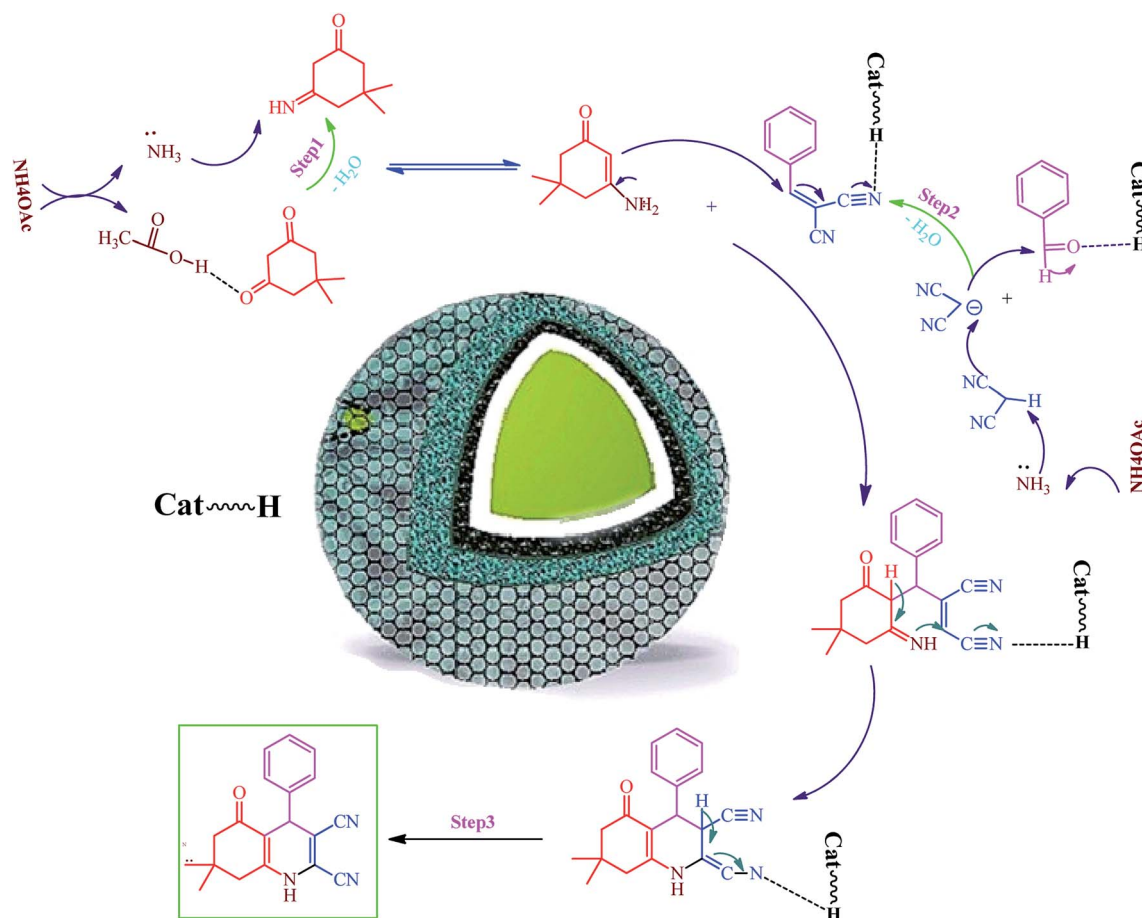
3.6. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ stability

To study the structure and confirm the stability of the recovered catalyst, FTIR, EDS, and VSM results of the recycled $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ were investigated. The FTIR spectra results of reused catalyst indicated that the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ structure remained intact during the hexahydroquinolines preparation reaction. The VSM analysis of recovered catalyst did not show a significant change in the saturation magnetization as both the fresh and recovered catalysts had the saturation magnetization of $\sim 17 \text{ emu g}^{-1}$ (Fig. 6b). By comparing the EDS results of fresh and recovered catalysts, it can be observed that all elements in the fresh catalyst (such as Fe, Si, O, C, and S) are present in the recovered catalyst, thus

indicating the stability of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ structure during the reaction (Fig. 6c). These results reveal the stability of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ structure and indicate the high activity of the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ catalyst after recovery.

3.7. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ -catalyzed hexahydroquinolines synthesis mechanism

Hexahydroquinoline preparation reaction occurs in three stages. (i) The addition of ammonium acetate to the dimedone and the formation of enaminone (enaminone possesses the nucleophilic characters of enamine and enone).⁵⁷ At this stage, ammonium acetate serves as a source of hydrogen and activates the carbonyl group from dimedone and malononitrile to form enaminone.⁵⁷ (ii) A Knoevenagel reaction between the benzaldehyde and the malononitrile gives rise to the formation of arylidenemalononitrile (the ammonium acetate also acts as a base source and converts malononitrile into a nucleophile by removing hydrogen from the malononitrile).⁵⁸ (iii) A Michael addition reaction involving the intramolecular cyclization between the enaminone and arylidenemalononitrile; water elimination under this condition results in the formation of the final product (Scheme 1). Owing to its acidic hydrogen, carbonyl, and cyanide groups, $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ can be



Scheme 1 Proposed mechanism of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF-SO}_3\text{H}$ catalyzed hexahydroquinolines synthesis.



activated in three different steps and enhance the activity of intermediates in each step (Scheme 1).

4. Conclusions

In this study, a new core-double shell nanostructure of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}@\text{SO}_3\text{H}$ nanocatalyst was successfully synthesized and characterized using FT-IR, VSM, SEM, TEM, EDAX, TGA, and XRD analyses. FT-IR, EDAX, and TGA showed successful stabilization of sulfuric acid acidic groups on the $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}$ surface. FESEM and TEM analyses confirmed the spherical core-double-shell structure of the products. This new nanocatalyst was successfully applied in a one-step reaction to synthesize the hexahydroquinoline derivatives with high purity and yield. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}@\text{SO}_3\text{H}$ nanocatalyst was simply recycled and reused 10 times without significant reduction in its efficiency. Moreover, the efficiency of this new nanocatalyst was compared with the previously reported catalysts, which indicated shorter reaction time, higher yields, solvent-free conditions, and higher reaction cycle of the proposed nanocatalyst. Finally, EDAX and VSM results of the recycled $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{RF}@\text{SO}_3\text{H}$ confirmed its high stability.

Conflicts of interest

There is no conflicts of interest.

References

- 1 S. Kumar, P. Sharma, K. K. Kapoor and M. S. Hundal, *Tetrahedron*, 2008, **64**, 536–542.
- 2 D. Schade, M. Lanier, E. Willems, K. Okolotowicz, P. Bushway, C. Wahlquist, C. Gilley, M. Mercola and J. R. Cashman, *J. Med. Chem.*, 2012, **55**, 9946–9957.
- 3 L.-M. Wang, J. Sheng, L. Zhang, J.-W. Han, Z.-Y. Fan, H. Tian and C.-T. Qian, *Tetrahedron*, 2005, **61**, 1539–1543.
- 4 M. M. Heravi, K. Bakhtiari, N. M. Javadi, F. F. Bamoharram, M. Saeedi and H. A. Oskooie, *J. Mol. Catal. A: Chem.*, 2007, **264**, 50–52.
- 5 S. R. Cherkupally and R. Mekala, *Chem. Pharm. Bull.*, 2008, **56**, 1002–1004.
- 6 M. Maheswara, V. Siddaiah, G. L. V. Damu and C. V. Rao, *ARKIVOC*, 2006, **2**, 201–206.
- 7 J. L. Donelson, R. A. Gibbs and S. K. De, *J. Mol. Catal. A: Chem.*, 2006, **256**, 309–311.
- 8 C. S. Reddy and M. Raghu, *Chin. Chem. Lett.*, 2008, **19**, 775–779.
- 9 R. Surasani, D. Kalita, A. D. Rao, K. Yarbaki and K. Chandrasekhar, *J. Fluorine Chem.*, 2012, **135**, 91–96.
- 10 A. Amoozadeh, S. Rahmani, M. Bitaraf, F. B. Abadi and E. Tabrizian, *New J. Chem.*, 2016, **40**, 770–780.
- 11 J. Safaei-Ghomi, R. Aghagholi and H. Shahbazi-Alavi, *Z. Naturforsch. B*, 2018, **73**, 269–274.
- 12 E. Tabrizian and A. Amoozadeh, *Catal. Sci. Technol.*, 2016, **6**, 6267–6276.
- 13 A. H. Lu, E. e. L. Salabas and F. Schüth, *Angew. Chem., Int. Ed.*, 2007, **46**, 1222–1244.
- 14 J. Gao, H. Gu and B. Xu, *Acc. Chem. Res.*, 2009, **42**, 1097–1107.
- 15 Y. Kang, L. Zhou, X. Li and J. Yuan, *J. Mater. Chem.*, 2011, **21**, 3704–3710.
- 16 Y.-w. Jun, Y.-M. Huh, J.-s. Choi, J.-H. Lee, H.-T. Song, S. Kim, S. Kim, S. Yoon, K.-S. Kim and J.-S. Shin, *J. Am. Chem. Soc.*, 2005, **127**, 5732–5733.
- 17 C. Jin, Y. Wang, H. Tang, H. Wei, X. Liu and J. Wang, *J. Physic. Chem. C*, 2014, **118**, 25110–25117.
- 18 C. R. Vestal and Z. J. Zhang, *Nano Lett.*, 2003, **3**, 1739–1743.
- 19 G. Huang, C. Zhang, S. Li, C. Khemtong, S.-G. Yang, R. Tian, J. D. Minna, K. C. Brown and J. Gao, *J. Mater. Chem.*, 2009, **19**, 6367–6372.
- 20 S. Taheri, H. Veisi and M. Hekmati, *New J. Chem.*, 2017, **41**, 5075–5081.
- 21 P. Yang, Y. Xu, L. Chen, X. Wang and Q. Zhang, *Langmuir*, 2015, **31**, 11701–11708.
- 22 X. Liu, S. Li, J. Mei, W.-M. Lau, R. Mi, Y. Li, H. Liu and L. Liu, *J. Mater. Chem. A*, 2014, **2**, 14429–14438.
- 23 Y. Shao, L. Zhou, C. Bao, Q. Wu, W. Wu and M. Liu, *New J. Chem.*, 2016, **40**, 9684–9693.
- 24 J. Choma, D. Jamiola, K. Augustynek, M. Marszewski and M. Jaroniec, *Chem. Commun.*, 2012, **48**, 3972–3974.
- 25 S.-H. Wu, C.-Y. Mou and H.-P. Lin, *Chem. Soc. Rev.*, 2013, **42**, 3862–3875.
- 26 J. Liu, S. Z. Qiao, H. Liu, J. Chen, A. Orpe, D. Zhao and G. Q. Lu, *Angew. Chem.*, 2011, **123**, 6069–6073.
- 27 Q. Zhang, I. Lee, J. B. Joo, F. Zaera and Y. Yin, *Acc. Chem. Res.*, 2012, **46**, 1816–1824.
- 28 M. Zhao, K. Deng, L. He, Y. Liu, G. Li, H. Zhao and Z. Tang, *J. Am. Chem. Soc.*, 2014, **136**, 1738–1741.
- 29 Y. Li, J. Jin, D. Wang, J. Lv, K. Hou, Y. Liu, C. Chen and Z. Tang, *Nano Res.*, 2018, **11**, 3294–3305.
- 30 J. Lu, W. Zhou, L. Wang, J. Jia, Y. Ke, L. Yang, K. Zhou, X. Liu, Z. Tang and L. Li, *ACS Catal.*, 2016, **6**, 1045–1053.
- 31 X. Wang, L. Fan, D. Gong, J. Zhu, Q. Zhang and B. Lu, *Adv. Funct. Mater.*, 2016, **26**, 1104–1111.
- 32 Z. Zhang, F. Wang, Q. An, W. Li and P. Wu, *J. Mater. Chem. A*, 2015, **3**, 7036–7043.
- 33 M. Xu, D. Chen, P. Huang, Z. Wan, Y. Zhou and Z. Ji, *J. Mater. Chem. C*, 2016, **4**, 6516–6524.
- 34 S. M. Majhi, P. Rai and Y.-T. Yu, *ACS Appl. Mater. Interfaces*, 2015, **7**, 9462–9468.
- 35 X. Y. Yu, L. Yu and X. W. Lou, *Adv. Energy Mater.*, 2016, **6**, 1501333.
- 36 R. Zou, M. F. Yuen, L. Yu, J. Hu, C.-S. Lee and W. Zhang, *Sci. Rep.*, 2016, **6**, 20264.
- 37 P. Tan, Y. Jiang, X.-Q. Liu, D.-Y. Zhang and L.-B. Sun, *ACS Sustainable Chem. Eng.*, 2016, **4**, 2223–2231.
- 38 T. Wu, Y. Liu, X. Zeng, T. Cui, Y. Zhao, Y. Li and G. Tong, *ACS Appl. Mater. Interfaces*, 2016, **8**, 7370–7380.
- 39 R. Purbia and S. Paria, *Nanoscale*, 2015, **7**, 19789–19873.
- 40 X. Fang, S. Liu, J. Zang, C. Xu, M.-S. Zheng, Q.-F. Dong, D. Sun and N. Zheng, *Nanoscale*, 2013, **5**, 6908–6916.
- 41 R. Liu and R. D. Priestley, *J. Mater. Chem. A*, 2016, **4**, 6680–6692.
- 42 Y. Liu, W. Wang, Q. Chen, C. Xu, D. Cai and H. Zhan, *Inorg. Chem.*, 2019, **58**(2), 1330–1338.



- 43 R. Liu, Y. Guo, G. Odusote, F. Qu and R. D. Priestley, *ACS Appl. Mater. Interfaces*, 2013, **5**, 9167–9171.
- 44 R. Liu, Y.-W. Yeh, V. H. Tam, F. Qu, N. Yao and R. D. Priestley, *Chem. Commun.*, 2014, **50**, 9056–9059.
- 45 M. Wang, Y. Ni and A. Liu, *ACS Omega*, 2017, **2**, 1505–1512.
- 46 Y. Zhang, Y. Yang, H. Duan and C. Lü, *ACS Appl. Mater. Interfaces*, 2018, **10**(39), 12469–12475.
- 47 A. Amoozadeh, S. Golian and S. Rahmani, *RSC Adv.*, 2015, **5**, 45974–45982.
- 48 Z. Wu, C. Chen, L. Wang, H. Wan and G. Guan, *Ind. Eng. Chem. Res.*, 2016, **55**, 1833–1842.
- 49 H. Veisi, S. Taheri and S. Hemmati, *Green Chem.*, 2016, **18**, 6337–6348.
- 50 R. Yang, Y. Liu, X. Yan, S. Liu and H. Zheng, *J. Mater. Chem. A*, 2016, **4**, 9807–9815.
- 51 Y.-X. Wang, J. Yang, S.-L. Chou, H. K. Liu, W.-x. Zhang, D. Zhao and S. X. Dou, *Nat. Commun.*, 2015, **6**, 8689.
- 52 J. Pan, H. Sun, X. Yan, W. Zhong, W. Shen, Y. Zhang and X. Cheng, *Ceram. Int.*, 2020, **520**(1), 146335.
- 53 X. Xie, L. Chen, X. Pan and S. Wang, *J. Chromatogr. A*, 2015, **1405**, 32–39.
- 54 Y. Zhong, Y. Ni, S. Li and M. Wang, *RSC Adv.*, 2016, **6**, 15831–15837.
- 55 K. Aswin, K. Logaiya, P. N. Sudhan and S. S. Mansoor, *J. Taibah Univ. Sci.*, 2012, **6**, 1–9.
- 56 B. Amirheidari, M. Seifi and M. Abaszadeh, *Res. Chem. Intermed.*, 2016, **42**, 3413–3423.
- 57 D. Patil, D. Chandam, A. Mulik, S. Jagdale, P. Patil and M. Deshmukh, *J. Saudi Chem. Soc.*, 2017, **21**, S329–S338.
- 58 J. van Schijndel, D. Molendijk, H. Spakman, E. Knaven, L. A. Canalle and J. Meuldijk, *Green Chem. Lett. Rev.*, 2019, **12**, 323–331.

