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A facile method for the preparation of black TiO₂ by Al reduction of TiO₂ and their visible light photocatalytic activity

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Black TiO₂ has attracted widespread attention due to its visible light absorption and wide range of applications. However, the currently reported preparation methods for black TiO₂ are not suitable for large-scale production due to its being prepared under high vacuum and over a long time. We have successfully prepared black TiO₂ under normal pressure and short time conditions. The as-prepared black titanium dioxide was characterized by XRD, XPS, TEM, UV-visible absorption spectrum and other characterization methods. The result shows that the as prepared black titanium dioxide had a disordered structure and oxygen vacancy defects on the surface, and exhibits excellent visible and near infrared absorption performance. The black TiO₂ sample was prepared under 650 °C 60 min exhibits excellent visible light photocatalytic performance, and can degrade 56% MO after visible light irradiation for 120 min.

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1. Introduction

TiO₂ has become one of the most studied photocatalyst materials since the Japanese scientists Fujishima and Honda discovered that the TiO₂ electrode photolyzes water.¹ TiO₂ has been widely studied for hydrogen production,^{2,3} dye-sensitized solar cells,^{4,5} CO₂ reduction,^{6–8} environmental pollution removal,⁹ toxic gas oxidation¹⁰ and photocatalytic organic synthesis of chemicals.¹¹ However, TiO₂ can only absorb the energy in the ultraviolet region because the band width of TiO₂ is 3.0 to 3.2 eV. The solar energy utilization rate of TiO₂ is seriously insufficient due to the UV light area only accounting for less than 5% of sunlight. Many methods such as element doping, oxygen vacancies, dye-sensitization and preparation of composite materials have been made to improve light absorption properties of TiO₂ materials.^{12–24} However, the results are not up to the mark.

Black TiO₂ is first produced by Chen *et al.* in 2011,²⁵ which have received widespread attention due to enhanced sunlight absorption by forming oxygen vacancies and Ti³⁺ in TiO₂ or introducing a disordered layer on the surface of highly crystalline TiO₂. Since then, black TiO₂ became an active research area for many applications which include photocatalytic pollutant degradation, hydrogen production through water splitting, photocatalytic CO₂ reduction, dye sensitized solar cells, lithium batteries and other photoelectrochemical applications with great success.

Currently, different strategies have been employed for the synthesis of black TiO₂. The preparation methods and process conditions of common black titanium dioxide are shown in Table 1.

Hydrogenation is the widely used method for synthesizing black TiO₂. Although there are many ways to prepare black titanium dioxide. However, novel green synthesis methods which are easy, low cost, fast and eco-friendly need to be further developed.

The methods for preparing black titanium dioxide described above all require high-pressure hydrogen or inert gas or vacuum conditions. The preparation conditions are relatively harsh, the equipment requirements are high and continuous preparation is very difficult, resulting in higher costs for preparing black titanium dioxide and difficult to prepare on a large scale.

In this paper, the author cleverly designed a molten B₂O₃ covering the surface of the reaction raw materials to effectively isolate the air and B₂O₃ can be recycled repeatedly, as shown in Fig. 1. Black titanium dioxide can be prepared under normal pressure, which does not require inert atmosphere or vacuum conditions. At the same time, this method has low reaction temperature, short reaction time and low equipment requirements. The cost of preparing black titanium dioxide can greatly reduce and easy to realize large-scale preparation. The photocatalytic activity of the synthesized black TiO₂ materials was also investigated by the degradation of methyl orange (MO) under the visible light irradiation.

At present, the preparation methods of black TiO₂ mainly include gas phase atmosphere sintering (high pressure pure hydrogen),^{25,26} chemical reduction (aluminum reduction, magnesium reduction),^{27,28} chemical oxidation²⁹ and

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Table 1 The currently reported preparation process of black titanium dioxide

Number	Preparation process	Process conditions
1	High/low pressure hydrogen treatment	The black TiO ₂ nanocrystals was synthesized in a 20 bar H ₂ atmosphere at elevated temperatures ranges from 200 to 450 °C for few days ^{25–27}
2	Hydrogen–argon treatment	The black TiO ₂ nanocrystals were obtained by under H ₂ –Ar mixture at 450 °C for 1 h ²⁸
3	Argon–nitrogen treatment	The black TiO ₂ –B nanoparticles were obtained by heating the solid product obtained in the above process at 340 °C in Ar atmosphere for 2 h ²⁹ A core–shell black anatase TiO ₂ with a high concentration of Ti ³⁺ and oxygen vacancy defects by a one-pot synthetic method calcination of colloidal TiO ₂ precursor under N ₂ atmosphere alone ³⁰
4	Hydrogen nitrogen treatment	Black TiO ₂ through hydrogen spill in a H ₂ –N ₂ atmosphere at 200–700 °C ³¹
5	Argon treatment	The black TiO ₂ precursor (Ni doped TiO ₂) was mixed with 2 g of NaBH ₄ and heated at 350 °C under Ar atmosphere for 1 h ³²
6	Plasma treatment	Precursor TiO ₂ was subjected to heat at 350–500 °C for 3 h under the hot filament (2000 °C) resulted in black colored TiO ₂ nanotubes ³³
7	NaBH ₄ reduction	P25 (anatase and rutile) was ground thoroughly with NaBH ₄ and the mixture was heated in a tubular furnace under Ar atmosphere at 300–400 °C for different time intervals up to 1 h ³⁴
8	Metal reduction	Black TiO ₂ nanoparticles could be synthesized through reduction with various metals. Aluminum ³⁵ and magnesium ³⁶ are identified as reducing agents to obtain black TiO ₂ . Wang <i>et al.</i> used Al as a reducing agent in an evacuated two-zone vacuum furnace at 300 to 50 °C. In a typical procedure, pre-annealing of aluminum was done at 800 °C for 6 h and that of pristine TiO ₂ was done at 500 °C for 20 h respectively. Further the post annealing was carried out at 800 and 900 °C for 12 h respectively ³⁷

electrochemical reduction.³⁰ However, these preparation methods are not suitable for large-scale production due to their high cost. Therefore, it is of great significance to seek a simple, efficient and low-cost method for preparing black titanium dioxide.

In this paper, a novel method for preparation of black TiO₂ by aluminothermic reduction had been studied. Through clever design, black TiO₂ materials has been successfully prepared at a lower calcination temperature and a shorter time under normal pressure. The photocatalytic activity of the synthesized black TiO₂ materials was also investigated by the degradation of methyl orange (MO) under the visible light irradiation.

2. Experimental

2.1 Preparation of black TiO₂

Anatase type TiO₂ (typically 0.4–0.6 μm in diameter) were purchased from Sichuan province excellence vanadium and titanium Co., Ltd. Aluminum powder (99.3% purity, 6 μm particle size) were purchased from Jiangsu tian yuan metal powder co., Ltd. Firstly, the anatase TiO₂ powder mixed with Al powder was placed in alumina crucible. Secondly, an aluminum sheet as large as the corundum crucible was placed on the surface of the mixed raw material and then a layer of molten B₂O₃ was covered the surface of the aluminum sheet so as to isolate the raw material from air contact. Thirdly, corundum crucible was roasted in a resistance furnace. Finally, hydrochloric acid pickling method was used to remove excess Al from the as-prepared sample. The prepared sample was mixed with hydrochloric acid in a beaker at a stirring rate of 100 rpm, and it was filtered using a vacuum suction funnel after the reaction was completed. The concentration of hydrochloric acid was

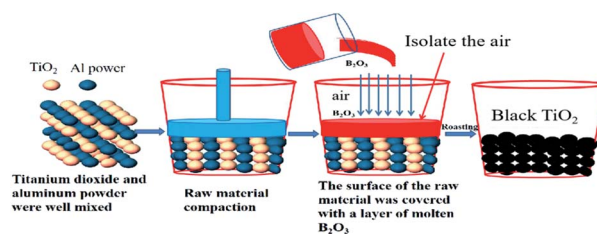
4 mol L^{−1} and the liquid–solid ratio was 1 : 25. The preparation process of black titanium dioxide was shown in Fig. 2.

2.2 Photocatalytic activity test

Photocatalytic activity of the synthesized black TiO₂ powders by the degradation of methylene orange (MO) under a 300 W xenon lamp equipped with a cutoff filter (λ > 420 nm). A quantity of 100 mg of photocatalyst was poised in a 200 mL aqueous MO dye solution (0.5 g L^{−1}). Previous to the light, the suspensions were ultrasonicated for 25 min under dark condition to make sure that the MO dye was adsorbed to saturation level on the surface of catalysts.

2.3 Sample characterization

Powder X-ray diffraction (XRD) profiles were obtained using a D/max22500PC of Japanese Science and Technology with Cu Kα (λ = 0.154 nm) radiation as the incident beam. Transmission electron microscopy (TEM) was performed on a Hitachi H-9000 instrument operating at 300 kV. Scanning electron microscopy (SEM) was performed on JEOL JSM-6060S and JSM-6700F

**Fig. 1** Schematic diagram of roasting stage.

instruments. XPS profiles were obtained using Phi-5000 VersaProbe of America.

3. Results and discussion

3.1 XRD and Raman analysis of samples

Black TiO₂ materials were successfully synthesized using an Al reduction process. Commercial anatase TiO₂ power (50 g) were uniformly mixed with 30 g of Al powder, and then calcined at 650 °C for 60 min, 90 min, 120 min, 150 min. The pure TiO₂ were then obtained after excess Al was removed by pickling with hydrochloric acid and distilled water.

Macro picture of experimental results is depicted in Fig. 3(a). The color of anatase titanium dioxide starts to change from white to black at 650 °C for 60 min. The XRD patterns of the raw material and as-prepared samples are shown in Fig. 2(a). The prepared black TiO₂ has good crystallinity, and its characteristic peaks match well with anatase TiO₂ without any other diffraction peaks. The characteristic peak of the (101) plane of black TiO₂ shifts to the right small diffraction angle and the larger the diffraction angle shifts to the right as the sintering time increases, which may be due to certain lattice defects.³⁸ In addition, the half peak width of as-prepared black TiO₂ is wider comparing with anatase TiO₂, which indicate that the degree of crystallinity of as-prepared black TiO₂ is poor compared with anatase TiO₂ and the crystal size of as-prepared is smaller than the anatase TiO₂.

Fig. 3(b) shows the Raman spectrum of the raw material and as-prepared samples. The strongest intensity Raman active mode at 144 cm⁻¹ and three medium intensity Raman active modes at 394 cm⁻¹, 515 cm⁻¹, and 636 cm⁻¹ were detected in all the samples, which match well with anatase TiO₂. The as-prepared black TiO₂ confirmed the presence of oxygen vacancies defects because the strongest E_g mode at 144 cm⁻¹ broadening and blue shift compared with that of pristine anatase TiO₂.

3.2 X-ray photoelectron spectroscopy (XPS) analysis of samples

Valence band position and useful information of chemical binding on the sample surface can be provided by X-ray photoelectron spectroscopy (XPS). Fig. 4(a–d) shows the Ti2p and O1s core level XPS spectra of the pristine TiO₂ and as-

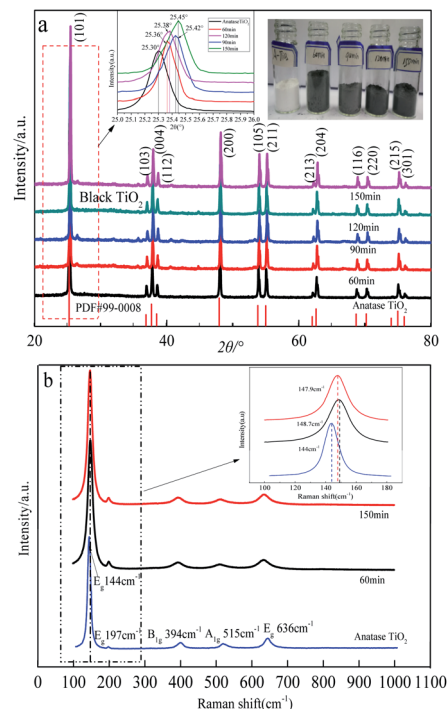


Fig. 3 XRD and Raman analysis of samples. (a) XRD and macro pictures of samples; (b) Raman spectrum of the samples.

prepared black TiO₂. Fig. 3(a) shows the Ti2p XPS spectra of anatase TiO₂ and black TiO₂. The two Ti2p peaks of anatase TiO₂ and black TiO₂ were located at 458.6 eV and 464.4 eV match well with Ti⁴⁺ in TiO₂.³⁹ However, the significant Ti³⁺ signal could not be detected in the black TiO₂. The possible reason is that Ti³⁺ didn't appear in the crystal lattice, but it plays

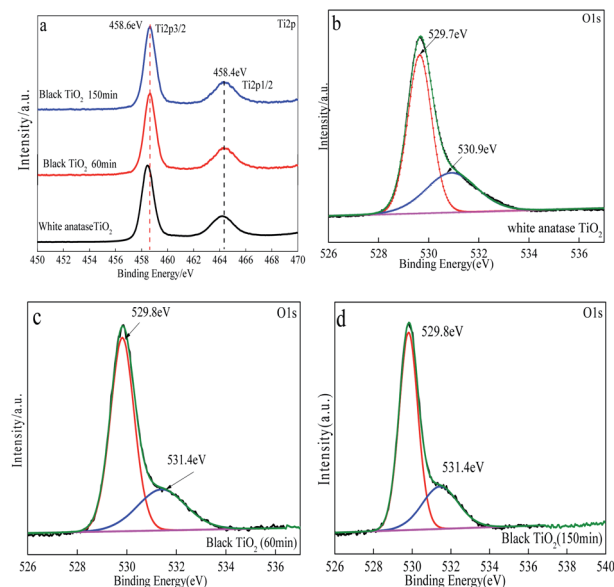


Fig. 4 XPS analysis of the samples. (a) Ti2p core level XPS spectra of the TiO₂ and black TiO₂; (b–d) O1s core level XPS spectra of the TiO₂ and black TiO₂.

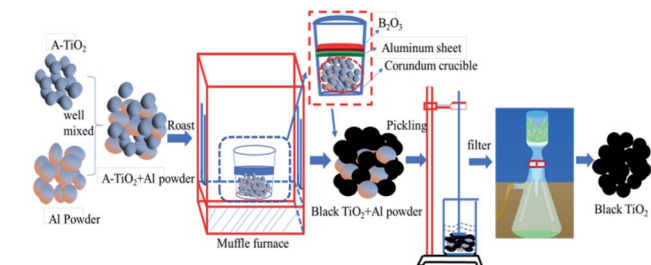


Fig. 2 Schematic diagram of black titanium dioxide was synthesized by aluminothermic reduction and hydrochloric acid pickling.



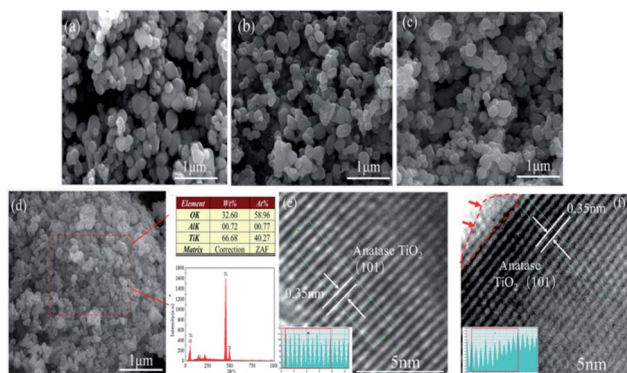


Fig. 5 SEM, EDS and HR-TEM analysis of samples. (a)–(c) SEM images of pristine TiO₂ and as-prepared black TiO₂ (650 °C, 60 min and 650 °C, 150 min); (d) EDS analysis of the black TiO₂ was pickled by hydrochloric acid; (e) and (f) HR-TEM image of the TiO₂ and black TiO₂.

an important role in color change.⁴⁰ The O1s XPS spectra of the pristine TiO₂ and black TiO₂ are depicted in Fig. 3(c and d). The peak for O1s could be deconvoluted into two peaks at about 529.7 eV and 530.9 eV. The peaks at 529.7 eV in anatase TiO₂ and black TiO₂ were belong to the crystal lattice of the Ti–O bond, while the peaks at 530.9 eV or 531.4 eV were the external –OH group or H₂O molecules absorbed on the anatase TiO₂ and black TiO₂ surface.⁴¹ It was clearly seen that –OH group absorption intensity of black TiO₂ was significantly higher than that of anatase TiO₂ due to the presence of oxygen vacancies in the black TiO₂.

3.3 SEM, EDS and HR-TEM analysis of samples

Fig. 5(a–c) shows SEM images of pristine TiO₂ and black TiO₂, respectively. Pristine TiO₂ with a particle size of 200–300 nm, and the morphology of black TiO₂ had not obvious sintering growth compared to pristine anatase TiO₂. The EDS analysis (Fig. 4(d)) shows only trace amounts of Al in black TiO₂.

The morphology and structure of pristine TiO₂ and as-prepared black TiO₂ were further examined by TEM. Fig. 5(e and f) shows HR-TEM analysis of pristine TiO₂ and black TiO₂. A lattice space of as-prepared black TiO₂ and pristine TiO₂ were 0.35 nm, which were well-resolved (101) lattice plane of typical anatase TiO₂ plane. Both HR-TEM and XRD analysis results indicate that the as-prepared black TiO₂ contains a pure anatase phase. The area was marked by the red arrow in Fig. 4(f) could be seen that the as-prepared black TiO₂ contained an outer disordered structure on the surface compared to pristine anatase TiO₂.

3.4 UV-Vis diffuse reflectance spectroscopy analysis of samples

UV-Vis diffuse reflectance spectroscopy are used to evaluate the light absorption characteristics of the sample. The UV-Vis diffuse reflectance spectra is depicted in Fig. 6(a) shows the as-prepared black TiO₂ exhibits excellent visible and near infrared absorption performance, and anatase TiO₂ mainly absorb ultraviolet light. It can be clearly seen that the UV-visible

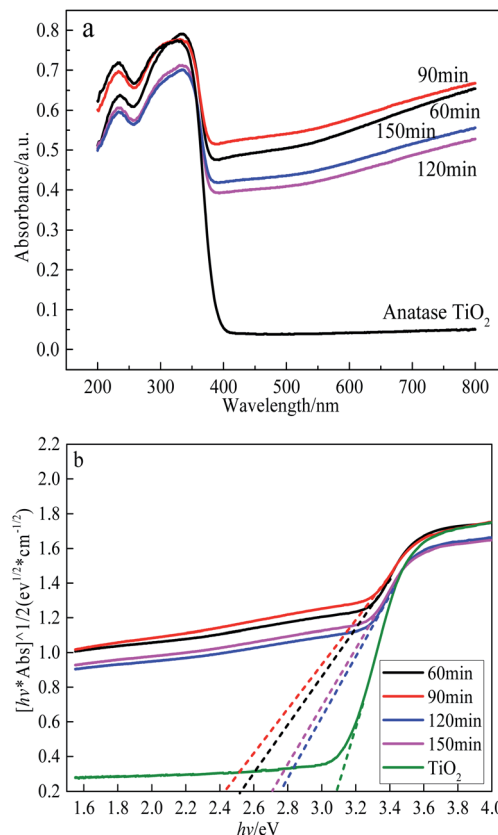


Fig. 6 UV-Vis diffuse reflectance spectroscopy analysis. (a) UV-Vis diffuse reflectance spectroscopy of the TiO₂ and black TiO₂; (b) plots of $(\alpha h\nu)^{1/2}$ versus photon energy ($h\nu$).

spectrum of as-prepared black TiO₂ occurred a red shift comparing with white anatase TiO₂. Kubelka Munk function formula ($\alpha h\nu = C(h\nu - E_g)^{1/2}$) could be used to calculate the band gap of semiconductor materials based on UV-Vis DRS, as shown in Fig. 6(b). It can be seen that the band gap of the prepared black TiO₂ was 2.4–2.8 eV, which was significantly narrowed compared with white TiO₂ (3.1 eV). The band gap of black TiO₂ was narrowed and the absorption spectrum expand to the visible light scope may be caused by Ti dangling bonds was related with oxygen vacancy defects.⁴²

3.5 Photocatalytic activity analysis of samples

The photocatalytic activities of black TiO₂ samples were evaluated by the decomposition of MO under visible light irradiation (>420 nm), as shown in Fig. 7(a). For comparison, pure anatase TiO₂ was also tested under the same conditions, which can only degrade 5% MO within 120 min, respectively. The as-prepared black TiO₂ sample (650 °C, 60 min) could degrade 56% MO after reaction for 120 min under visible light. However, the prepared black TiO₂ (650 °C, 150 min) only degradation degrade 25% MO, which proves that extending the roasting time couldn't promote the visible light photocatalytic performance.

The kinetic curves for the MO photodegradation were calculated according to the equation ($\ln(C_0/C) = kt$) and then plotted and displayed in Fig. 7(b). The order of photocatalytic



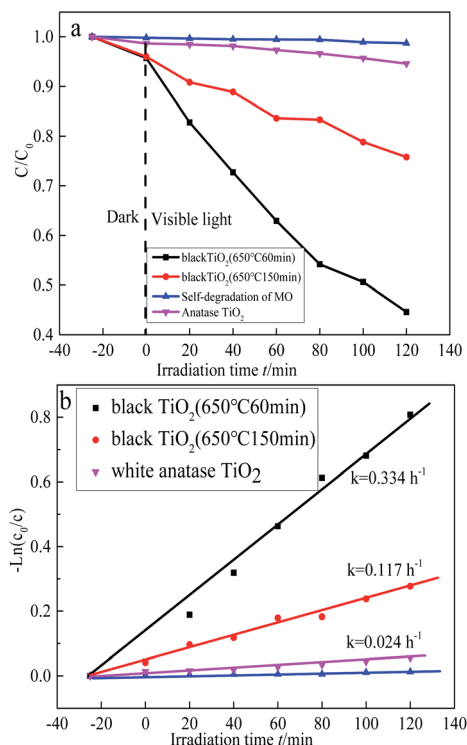


Fig. 7 Photocatalytic activity analysis of samples. (a) Photocatalytic activity of TiO₂ and black TiO₂ for the MO degradation under visible light irradiation; (b) kinetic curves for the MO photodegradation over black TiO₂ and TiO₂.

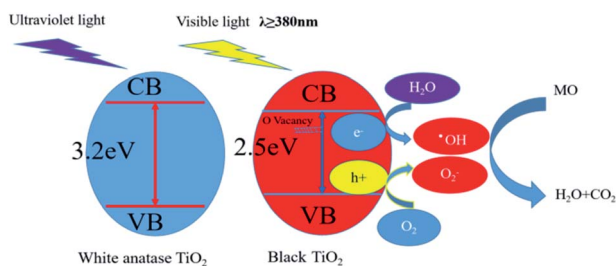


Fig. 8 Proposed mechanism for the MO degradation on the as prepared black TiO₂ photocatalyst under visible light irradiation.

performance of as-prepared photocatalysts is black TiO₂ (650 °C, 60 min) > black TiO₂ (650 °C, 150 min) > pure anatase TiO₂, based on the *k* values.

3.6 Mechanism analysis and discussion

High absorption of visible-light and pollutants, and a lower electron/hole recombination rate can remarkably enhance the photocatalysis. A schematic illustration of the mechanism of enhanced photocatalytic performance of as-prepared black TiO₂ is shown in Fig. 8.

XRD, XPS, Raman, UV-Vis characterizations show that the prepared black titanium dioxide has formed a large number of oxygen vacancies. The presence of oxygen vacancies can reduce the recombination of electron-hole pairs and enhance visible

light absorption by moving the top of VB upward and the bottom of CB downward, respectively, resulting in black TiO₂ having excellent visible light photocatalytic activity.⁴³

Although the prepared black TiO₂ enhances visible light absorption, its photocatalytic activity under visible light is still unsatisfactory due to the short lifetime of electrons and holes excited by visible light. In order to overcome this bottleneck, as-prepared black TiO₂ with a series of non-metal dopants (H, S, N and I) to increase photocatalytic activities under both UV and visible-light irradiations.⁴⁴

4. Conclusion

In this study, the black TiO₂ had been successfully synthesized by aluminothermic reduction and hydrochloric acid pickling. Compared with other preparation methods, this preparation method was suitable for large-scale production because it can be prepared at a lower calcination temperature and shorter calcination time under normal pressure conditions.

The surface of as prepared black titanium dioxide had disorder structure and oxygen vacancy defects. The particles of prepared black titanium dioxide had no obvious sintering growth, which maintains the original appearance of raw titanium dioxide. The band gap of the prepared black TiO₂ are 2.4–2.8 eV, which exhibits excellent visible and near infrared absorption performance. The black TiO₂ sample was prepared under 650 °C, 60 min demonstrates the highest photocatalytic activity, which can degrade 56% MO after visible light irradiation for 120 min. The products show excellent visible photocatalytic activity for dye degradation, which shows the good application value in water cleaning.

Conflicts of interest

There are no conflicts to declare.

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