1. Introduction

The proton exchange membrane fuel cell (PEMFC) is a promising candidate for a clean and sustainable energy source to cope with the growing energy consumption and related environmental concerns. Pt particles are used as a catalyst for the oxygen reduction reaction (ORR) at the PEMFC cathode because of their incomparable catalytic activity and stability in acidic solution, as discussed in recent reviews; we cite here reviews from the last 5 years because many reviews have been published. However, their limited availability on earth and high cost still remain major obstacles for the wide use of PEMFCs. One of the promising methods to solve this problem is the use of core–shell Pt particles (M@Pt) consisting of a Ru13 core and a Pt42 shell with Pt for the core, because Pt content can be tuned by varying the Pt shell thickness; Ru@PtxML (x = 1 and 3) demonstrated that the catalytic activity of Ru@Pt could be tuned by varying the Pt shell thickness; Ru@Pt3ML with two Pt layers was more active than Ru@PtxML (x = 1 and 3). Jackson and coworkers reported that the catalytic activity of Ru@Pt for the ORR exceeded that of a commercial Pt electrode. Jackson and co-workers also interestingly obtained a volcano plot of the catalytic activity against the Pt binding energy, suggesting that both overly strong and overly weak O binding with the Pt surface is not good for ORR catalysts.

Recently, core–shell Ru@Pt particles have been reported as excellent ORR catalysts. For instance, Adzic and co-workers demonstrated that the catalytic activity of Ru@Pt could be tuned by varying the Pt shell thickness; Ru@Pt3ML with two Pt layers was more active than Ru@PtxML (x = 1 and 3). Jackson and coworkers reported that the catalytic activity of Ru@Pt for the ORR exceeded that of a commercial Pt electrode. Jackson and co-workers also interestingly obtained a volcano plot of the catalytic activity against the Pt binding energy, suggesting that both overly strong and overly weak O binding with the Pt surface is not good for ORR catalysts. However, the relation between electronic structure and origin of the O binding energy has been unclear.
Ru13@Pt42 and Pt55. We believe that the theoretical purposes here are to explore O2 adsorption, O protons and enough supply of electrons to the electrode. Our currently, particles is still challenging and its development is needed even ORR. Considering that the theoretical study of nanoscale metal particles is still challenging and its development is needed even currently,46-50 the theoretical study of the O2 reaction on Ru@Pt and related metal particles is indispensable.

In this work, we theoretically investigated dioxygen (O2) adsorption and O–O bond cleavage by the Ru13@Pt42 particle in comparison with the Pt55 particle using DFT computations. In the O–O bond cleavage, two reaction courses are plausible; in one, the O–O bond of the adsorbed O2 molecule is cleaved. In the other reaction course, OOH species are formed on the surface, followed by O–OH bond cleavage, because it is likely that the OOH species are easily formed in the presence of excess protons and enough supply of electrons to the electrode. Our purposes here are to explore O2 adsorption, O–O bond cleavage, OOH formation, and O–OH bond cleavage, compare the reactivity between Ru13@Pt42 and Pt55, find important factor(s) determining the reactivities of Ru13@Pt42 and Pt55, and present a theoretical understanding of the differences between Ru13@Pt42 and Pt55. We believe that the theoretical findings on these issues are valuable for understanding the chemistry of nanometal particles.

2. Computational methods and models

Spin-polarized periodic DFT calculations were carried out using the Vienna Ab initio Simulation Package (VASP)11,52 where plane-wave basis sets were employed with an energy cutoff of 400 eV, and the projector-augmented-wave (PAW) pseudopotentials were used to represent core electrons. The PBE-D3 functional was employed in all calculations,52 where “D3” represents the dispersion correction proposed by Grimme’s group.54,55 Ru@ Pt and Pt particles were placed in a large supercell (25 Å × 25 Å × 25 Å) to ensure enough separation by vacuum. Geometry optimization of equilibrium structure was carried out in gas phase using the energy threshold of 0.0001 eV and force threshold of 0.01 eV Å−1. Optimization of transition state (TS) was carried out using the climbing image nudged elastic band (CI-NEB) method with the VASP transition-state tools (VTST),56,57 in which thresholds for convergence were set to 0.0001 eV for energy and 0.02 eV Å−1 for force. To evaluate solvent effects, an implicit solvation model, which describes the effect of electrostatic interaction between solute and solvent, was employed as implemented in VASPsol,58,59 where optimized geometry in gas phase was used.

Though Pt55 and Ru13@Pt42 particles are not very large compared to real catalysts, we employed these particles here as model nanoparticles because we have to optimize many intermediates and transition states. Also, the use of Pt55 and Ru13@Pt42 particles is not unreasonable, considering that subnanometer-scale metal particles with 55 metal atoms have been employed as model particle in theoretical studies of O2 adsorption and O–O bond cleavage46-50 and other catalytic reactions.64-69,74 The number “55” is a magic number for icosahedral (Ih) and cuboctahedral (Oh) structures. Here, we employed the Ih-like structure because the Ih structure is more stable than Oh in Ru13@Pt42, the relative stabilities of various spin states and comparison between core–shell and non-core–shell structures have been investigated recently.64 As shown in Scheme 1, the Pt42 surface of icosahedral Pt55 and Ru13@Pt42 consists of 20 triangular facets. Each facet has three different types of binding sites: top (t), bridge (b) and hollow (h). Adsorptions at these binding sites are classified as follows: (i) adsorption at the top site is denoted as t1 or t2, in which the adsorbate binds with one Pt atom at the edge or the vertex position. (ii) Adsorption in a bridging manner is denoted as b1 or b2, in which the adsorbate binds with two adjacent Pt atoms at either the edge and vertex positions or two adjacent edge positions. (iii) Adsorption at the hollow site is denoted as h1 or h2, in which the adsorbate binds at either the fcc-like position among three edge-Pt atoms or the hcp-like position among two edge-Pt atoms and one vertex-Pt atom. We explored all these possible adsorption sites.

3. Results and discussion

In this work, we firstly discuss O2 adsorption to Pt55 and Ru13@Pt42, followed by O–O bond cleavage, OOH formation from the adsorbed O2 molecule, and O–OH bond cleavage on the Pt55 and Ru13@Pt42 surfaces. Next, we show the differences in reactivity between Pt55 and Ru13@Pt42 in these reactions and elucidate the reasons for the differences. Lastly, we unveil the characteristic feature(s) of Ru13@Pt42 in comparison with Pt55.

3.1 Dioxygen (O2) adsorption

The binding energy $E_b(O_2)$ of dioxygen molecule (O2) increases (becomes more negative) following the order $O_2^{-1/2}h1 < O_2^{-1/2}t1 < O_2^{-1/2}t2 < O_2^{-1/2}b2 < O_2^{-1/2}h2 < O_2^{-1/2}h1 < O_2^{-1/2}b1$ for pure Pt55 and $O_2^{-1/2}h1 < O_2^{-1/2}b2 < O_2^{-1/2}h2 < O_2^{-1/2}t1 < O_2^{-1/2}h1 < O_2^{-1/2}h1 <$
O$_2$-η$_1^*/t2$ < O$_2$-μ$_3$/b1 for Ru$_{13}@Pt_{42}$, as shown in Fig. S1 in the ESI, where h1, t1 etc., represent the adsorption site shown in Scheme 1 and “A-μ$_3$/x” represents the interaction of adsorbate A with Pt at the x binding site in a μ$_3$ manner, hereinafter. Obviously, O$_2$ is preferentially adsorbed at the b1 site of both Pt$_{55}$ and Ru$_{13}@Pt_{42}$ in a μ$_3$-side-on manner (O$_2$-μ$_3$/b1). This is the most stable O$_2$ adsorption structure. The coordination number of the surface Pt atom is one of the important factors for stabilization of O$_2$ adsorption: because the coordination number of the vertex Pt atom is 6 but that of the edge Pt is 7, the O$_2$ molecule tends to interact with the vertex Pt atom compared to the edge Pt atom. However, the O$_2$ molecule cannot interact with two vertex Pt atoms because the vertex Pt is far from the neighboring vertex Pt. Thus, O$_2$ interacts with one vertex Pt and one edge Pt in a bridging manner, as seen by the O$_2$-μ$_3$/b1 structure.

In the most stable O$_2$-μ$_3$/b1-binding species (2), the O–O distance of Pt$_{55}$(O$_2$) 2$_{\text{Pt}}$ is moderately longer than that of Ru$_{13}@Pt_{42}$(O$_2$) 2$_{\text{RuPt}}$, and the Pt–O distance of 2$_{\text{Pt}}$ is moderately shorter than that of 2$_{\text{RuPt}}$, as shown in Fig. 1. The $E_a$(O$_2$-μ$_3$/b1) value is −1.85 (−1.70) eV for Pt$_{55}$ and −1.07 (−0.95) eV for Ru$_{13}@Pt_{42}$, as shown in Fig. 2, where the figures in parentheses represent the binding energy in gas phase. The larger binding energy of the O$_2$ molecule with Pt$_{55}$ than with Ru$_{13}@Pt_{42}$ is consistent with such geometrical features as the shorter Pt–O and longer O–O distances in 2$_{\text{Pt}}$ than in 2$_{\text{RuPt}}$. The reasons for the stronger O$_2$ adsorption with Pt$_{55}$ than with Ru$_{13}@Pt_{42}$ are discussed below. It is also noted that solvation by water enhances O$_2$ binding with these metal particles.

### 3.2 O–O bond cleavage

Starting from 2$_{\text{Pt}}$ and 2$_{\text{RuPt}}$, O–O bond cleavage occurs via transition states TS2/3a$_{\text{Pt}}$ and TS2/3a$_{\text{RuPt}}$ to afford Pt$_{55}$(O)$_2$ 3$_{\text{aPt}}$ and Ru$_{13}@Pt_{42}$(O)$_2$ 3$_{\text{aRuPt}}$, respectively, as shown in Fig. 1. In TS2/3a$_{\text{Pt}}$, the O–O distance is considerably elongated to 1.835 Å, by 0.451 Å, and the Pt–O distances become shorter to 1.825 Å and 1.845 Å, by 0.144 and 0.152 Å (Fig. 1). In TS2/3a$_{\text{RuPt}}$, the O–O distance is more elongated to 1.939 Å (by 0.565 Å) than in TS2/3a$_{\text{Pt}}$, indicating that TS2/3a$_{\text{Pt}}$ is more reactant-like than TS2/3a$_{\text{RuPt}}$. Consistent with the O–O bond elongation, the Pt–O distances become shorter in TS2/3a$_{\text{Pt}}$. Though they are almost the same between TS2/3a$_{\text{Pt}}$ and TS2/3a$_{\text{RuPt}}$, the average of Pt–O distances is moderately shorter in 3$_{\text{aPt}}$ than in 3$_{\text{aRuPt}}$, suggesting the stronger binding energy of the O atom with Pt$_{55}$ than with Ru$_{13}@Pt_{42}$, as discussed below. The activation barrier ($E_a$) relative to 2 is 0.31 (0.33) eV and 0.35 (0.49) eV for TS2/3a$_{\text{Pt}}$ and TS2/3a$_{\text{RuPt}}$, respectively, and the reaction energy ($\Delta E$) relative to 2 is −0.90 (−0.83) eV and −0.95 (−0.85) eV for 3$_{\text{aPt}}$ and 3$_{\text{aRuPt}}$, as shown in Fig. 2, where a negative $\Delta E$ value represents exothermicity. The smaller $E_a$ for the O–O bond cleavage on Pt$_{55}$ than on Ru$_{13}@Pt_{42}$ is consistent with the more reactant-like TS2/3a$_{\text{Pt}}$ than TS2/3a$_{\text{RuPt}}$. The moderately smaller $\Delta E$ in the Pt$_{55}$ case than in the Ru$_{13}@Pt_{42}$ case is seemingly inconsistent with the smaller $E_a$ of the former case than in the latter. But, this is not unreasonable because the O$_2$ binding energy with Pt$_{55}$ is overly larger than that with Ru$_{13}@Pt_{42}$. It is noted that the $E_a$ is smaller and $\Delta E$ is more negative in water than in gas phase, because CT is generally enhanced by polar solvents.

The short Pt–Pt distance of the surface has been discussed as one important factor for high catalytic activity.70-73 The surface

Fig. 1  Geometry changes in O$_2$ adsorption followed by O–O bond cleavage on Pt$_{55}$ and Ru$_{13}@Pt_{42}$. Values represent bond distance in angstrom.
Pt–Pt distance becomes longer by the O$_2$ adsorption and the O–O bond cleavage in both Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$ (Fig. 1), suggesting that the short Pt–Pt distance of the surface is not beneficial to these processes. The other important factor is flexibility of the Pt surface. Actually, the energy destabilization of Pt$_{55}$ is much smaller than that of Ru$_{13}@$Pt$_{42}$ when surface Pt–Pt distance is elongated; it is 1.0 kcal mol$^{-1}$ in Pt$_{55}$ and 7.5 kcal mol$^{-1}$ in Ru$_{13}@$Pt$_{42}$ for the Pt–Pt elongation by 0.4 Å, where we employed rather arbitrarily the elongation of 0.4 Å because the Pt–Pt distance is elongated by about 0.3–0.4 Å at TS2/3 of the O–O bond cleavage on Pt$_{55}$. These results suggest that the longer Pt–Pt distance and the larger flexibility of Pt$_{55}$ than those of Ru$_{13}@$Pt$_{42}$ are favorable for O$_2$ adsorption and O–O bond cleavage; in other words, the discussion that short Pt–Pt distance is good for high catalytic activity is not useful when these processes are rate-determining. In addition, it should be noted that the flexibility of the Pt surface is a crucially important factor besides Pt–Pt distance.

### 3.3 OOH formation followed by O–OH bond cleavage

In ORR, it is likely that the proton is adsorbed easily to the cathode surface because protons exist in excess in solution. Also, electrons are always supplied to the cathode. These features suggest that H species is formed on the cathode surface. Actually, the reactions of adsorbed H atoms with oxygen-containing species were discussed in recent works, which indicate they correspond to the Langmuir–Hinshelwood pathway. First, we investigated the formation of OOH species from the adsorbed O$_2$ molecule and H atom on the surface; the adsorption sites of OOH and H are shown in Fig. S3, Tables S2 and S3 in the ESL.†

The H$^+$/e$^-$ addition occurs with significant exothermicity in 2$^\text{Pt}$ and 2$^\text{RuPt}$ to afford Pt$_{55}$(O$_2$)(H)(H$^+$) 3$^\text{bPt}$ and Ru$_{13}@$Pt$_{42}$(O$_2$)(H$^+$) 3$^\text{bRuPt}$, as shown in Fig. 2, where one-half of $\Delta E$ of the eq. 2H$^+$ + 2e$^-$ → H$_2$, was taken as the energy of H$^+$/e$^-$; it is noted that this step can be tuned experimentally by the cell voltage. In 3$^\text{bPt}$ and 3$^\text{bRuPt}$, H$^+$ takes the position bridging two Pt atoms, as shown in Fig. 3 and 4. Starting from 3$^\text{bPt}$ and 3$^\text{bRuPt}$, the H$^+$ reacts with the adsorbed O$_2$ via transition states TS3/4$^\text{bPt}$ and TS3/4$^\text{bRuPt}$ to afford an OOH species adsorbed on the Pt surface Pt$_{55}$(OOH) 4$^\text{bPt}$ and Ru$_{13}@$Pt$_{42}$(OOH) 4$^\text{bRuPt}$, respectively (Fig. 3 and 4). In the transition state, the H$^+$ is approaching the O$_2$, keeping a bonding interaction with one Pt, and simultaneously, the O$^-$ is leaving the Pt$_{42}$ surface. Though the Pt–O$^-$ distance is very long in TS3/4$^\text{bPt}$, the O$^-$–H$^+$ distance is still long (1.41 Å), and the O$^-$–O$^-$ distance is moderately elongated. In TS3/4$^\text{bRuPt}$, the O$^-$–H$^+$ distance is much longer (2.006 Å) than that of TS3/4$^\text{bPt}$, suggesting that TS3/4$^\text{bRuPt}$ is more reactant-like than TS3/4$^\text{bPt}$, in contrast to the more product-like TS2/3$^\text{aPt}$ than TS2/3$^\text{aRuPt}$. This contrast is reasonable according to the Hammond rule because the O–O bond cleavage occurs with smaller $E_a$ (relative to 2$^\text{Pt}$), but the OOH formation occurs with larger $E_a$ on Pt$_{55}$ than on Ru$_{13}@$Pt$_{42}$. In 4$^\text{bPt}$ and 4$^\text{bRuPt}$, the O$^-$–O$^-$ distance is 1.451 Å and 1.456 Å, respectively, which is moderately shorter than that (1.471 Å, the PBE-D3-optimized value) of free HOOH. The surface Pt–Pt$^\text{a}$ distance becomes longer in this reaction, suggesting that the short Pt–Pt distance of the Pt surface is not favorable for this step, either.

We explored here another OOH formation pathway in which adsorbed O$_2$ molecule undergoes protonation followed by one-electron reduction, as proposed by several works. This reaction corresponds to the Eley–Rideal pathway. We compared the energy change in this pathway with that of the Langmuir–Hinshelwood pathway, as shown in Scheme 2. In the Pt$_{55}$ case, the energy changes differ little between these two pathways, suggesting that the OOH formation occurs via both pathways. In Ru$_{13}@$Pt$_{42}$, the H$^+$/e$^-$ addition to the adsorbed O$_2$ molecule is more exothermic than that to the Pt$_{42}$ surface (Scheme 2), indicating that the Eley–Rideal pathway is more favorable than the Langmuir–Hinshelwood pathway from the viewpoint of reaction energy. Also, Scheme 2 strongly suggests that the OOH formation occurs more easily on Ru$_{13}@$Pt$_{42}$ via the Eley–Rideal pathway than that on Pt$_{55}$ via the Langmuir–Hinshelwood and Eley–Rideal pathways. Here, we need to mention that the reaction pathway significantly depends on the coverage of Pt surface by O$_2$ molecules; the Langmuir–Hinshelwood pathway preferentially occurs at low coverage. On the other hand, the Eley–Rideal pathway preferentially occurs at high coverage from the viewpoint of possibility, while the Eley–Rideal pathway becomes less easy at high coverage than at low coverage, from
the viewpoint of reactivity of adsorbed O₂ molecule, because the adsorbed O₂ molecule becomes less negatively charged at high coverage. The mechanism of OOH formation significantly depends on reaction conditions, which must be investigated carefully in the near future. In both pathways, it is reasonably concluded that OOH formation is an easy process on Pt₅₅ and Ru₁₃@Pt₄₂ particles.

Fig. 3 Geometry changes in OOH formation through the reaction between adsorbed O₂ molecule and H atom, followed by O–OH bond cleavage on Pt₅₅ particle. Numbers represent bond distance in angstrom.

Fig. 4 Geometry changes in OOH formation through the reaction between adsorbed O₂ molecule and H atom, followed by O–OH bond cleavage on Ru₁₃@Pt₄₂ particle. Numbers represent bond distance in angstrom.
The OOH formation from adsorbed O₂ and H species occurs with a smaller $E_a$ of 0.49 (0.50) eV on Ru₁₃@Pt₄₁ than on Pt₅₅ ($E_a$ = 0.82 (0.80) eV), as shown in Fig. 2. The smaller $E_a$ in the reaction on Ru₁₃@Pt₄₂ than on Pt₅₅ is consistent with the more reactant-like TS₃/4₅RuPt than TS₃/4₅Pt. The O–OH bond cleavage occurs with a very small $E_a$ on both Pt₅₅ ($E_a$ = 0.12 (0.12) eV) and Ru₁₃@Pt₄₁ ($E_a$ = 0.17 (0.0) eV; Fig. 2). The very small $E_a$ for the O–OH bond cleavage is consistent with the reactant-like transition states TS₄/5₅Pt and TS₄/5₅RuPt. Though the $E_a$ is moderately smaller in the reaction on Pt₅₅ than on Ru₁₃@Pt₄₂, the difference is small, and therefore, the geometry of TS₄/5₅Pt differs little from that of TS₄/5₅RuPt.

It should be noted that the O–OH bond cleavage occurs with a smaller $E_a$ than the O–O bond cleavage of the adsorbed O₂ molecule on both Pt₅₅ and Ru₁₃@Pt₄₂. This is not surprising because the O–O bond of the adsorbed O₂ molecule is weaker than the original O–O double bond of free O₂ molecule, as shown by the elongated O–O bond, but still stronger than the O–O single bond of the OOH species. Another result to be noted is that the $E_a$ of the OOH formation from adsorbed O₂ and H species is larger in the Pt₅₅ case than in the Ru₁₃@Pt₄₂ case, but the $E_a$ for O–O bond cleavage is smaller in the Pt₅₅ case than in the Ru₁₃@Pt₄₂ case. These findings are discussed below in more detail on the basis of electronic structure.

### 3.4 Electronic process in O₂ adsorption, O–O bond cleavage, OOH formation, and O–OH bond cleavage

As shown in Table 1, the O₂ moiety is negatively charged in Pt₅₅(O₂) $2^\text{Pt}$ and Ru₁₃@Pt₄₂(O₂) $2^\text{RuPt}$, because O₂ adsorption occurs with charge-transfer (CT) from metal particle to O₂ molecule. This CT is enhanced by the polar solvent (water), as expected. Notably, the O₂ molecule is more negatively charged in $2^\text{RuPt}$ than in $2^\text{Pt}$. Seemingly, this result is inconsistent with the larger adsorption energy of the O₂ molecule with Pt₅₅ than with Ru₁₃@Pt₄₂, which is discussed below. In the O–O bond cleavage ($2 \rightarrow 3^\alpha$), both O¹ and O² atoms become much more negatively charged, as well known. A moderately larger CT

### Table 1 Changes of the Bader charges in O₂ adsorption and O–O bond cleavage

<table>
<thead>
<tr>
<th></th>
<th>2</th>
<th>TS2/3a</th>
<th>3a</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pt₅₅</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O¹</td>
<td>−0.305 (−0.273)</td>
<td>−0.496 (−0.434)</td>
<td>−0.610 (−0.509)</td>
</tr>
<tr>
<td>O²</td>
<td>−0.300 (−0.259)</td>
<td>−0.480 (−0.406)</td>
<td>−0.706 (−0.684)</td>
</tr>
<tr>
<td>Pt₅₅</td>
<td>+0.605 (+0.532)</td>
<td>+0.976 (+0.840)</td>
<td>+1.316 (1.193)</td>
</tr>
<tr>
<td>Δ(CT)</td>
<td>0.0 (0.0)</td>
<td>0.371 (0.308)</td>
<td>0.711 (0.661)</td>
</tr>
<tr>
<td>Ru₁₃@Pt₄₂</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O¹</td>
<td>−0.326 (−0.289)</td>
<td>−0.519 (−0.445)</td>
<td>−0.640 (−0.536)</td>
</tr>
<tr>
<td>O²</td>
<td>−0.320 (−0.268)</td>
<td>−0.506 (−0.427)</td>
<td>−0.719 (−0.689)</td>
</tr>
<tr>
<td>Ru₁₃@Pt₄₂</td>
<td>+0.646 (+0.557)</td>
<td>+1.025 (0.872)</td>
<td>+1.359 (+1.225)</td>
</tr>
<tr>
<td>Δ(CT)</td>
<td>0.0 (0.0)</td>
<td>0.379 (0.315)</td>
<td>0.713 (0.668)</td>
</tr>
</tbody>
</table>

* A positive value represents positive atomic charge, and *vice versa.*
* Values in parentheses represent the Bader charge in gas phase.

### Table 2 Changes of the Bader charges in OOH formation from adsorbed O₂ and H species and O–OH bond cleavage

<table>
<thead>
<tr>
<th></th>
<th>3b</th>
<th>TS3/4b</th>
<th>4b</th>
<th>TS4/5b</th>
<th>5b</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pt₅₅</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O¹</td>
<td>−0.295 (−0.263)</td>
<td>−0.246 (−0.218)</td>
<td>−0.371 (−0.339)</td>
<td>−0.669 (−0.630)</td>
<td>−1.016 (−0.935)</td>
</tr>
<tr>
<td>O²</td>
<td>−0.289 (−0.250)</td>
<td>−0.252 (−0.197)</td>
<td>−0.624 (−0.575)</td>
<td>−0.384 (−0.349)</td>
<td>−0.615 (−0.511)</td>
</tr>
<tr>
<td>H</td>
<td>−0.035 (−0.037)</td>
<td>0.198 (0.204)</td>
<td>0.683 (0.609)</td>
<td>0.727 (0.653)</td>
<td>0.676 (0.610)</td>
</tr>
<tr>
<td>Pt₅₅</td>
<td>0.619 (0.550)</td>
<td>0.300 (0.211)</td>
<td>0.312 (0.305)</td>
<td>0.326 (0.326)</td>
<td>0.955 (0.836)</td>
</tr>
<tr>
<td>Ru₁₃@Pt₄₂</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O¹</td>
<td>−0.312 (−0.261)</td>
<td>−0.260 (−0.235)</td>
<td>−0.401 (−0.367)</td>
<td>−0.945 (−0.671)</td>
<td>−1.182 (−1.138)</td>
</tr>
<tr>
<td>O²</td>
<td>−0.278 (−0.245)</td>
<td>−0.210 (−0.140)</td>
<td>−0.800 (−0.728)</td>
<td>−0.401 (−0.367)</td>
<td>−0.708 (−0.680)</td>
</tr>
<tr>
<td>H</td>
<td>0.008 (0.007)</td>
<td>0.069 (0.083)</td>
<td>0.872 (0.782)</td>
<td>1.000 (0.675)</td>
<td>0.822 (0.789)</td>
</tr>
<tr>
<td>Ru₁₃@Pt₄₂</td>
<td>0.582 (0.499)</td>
<td>0.401 (0.292)</td>
<td>0.329 (0.313)</td>
<td>0.346 (0.363)</td>
<td>1.068 (1.029)</td>
</tr>
</tbody>
</table>

* Values in parentheses represent the Bader charge in gas phase.
occurs unexpectedly in Ru$_{13}@$Pt$_{42}$ than in Pt$_{55}$, despite the larger $E_a$ in Ru$_{13}@$Pt$_{42}$ than in Pt$_{55}$.

In the OOH formation via the reaction between the adsorbed O$_2$ molecule and H atom ($3b \rightarrow TS3/4b \rightarrow 4b$), the H atom becomes more positively charged, the O$^-$ and O$^2-$ atoms become more negatively charged, and the positive charges of Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$ decrease. However, these population changes are not simple. The positive charges of Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$ decrease when going from $3b$ to $TS3/4b$ but change little after $TS3/4b$, suggesting that the CT from O$_2$ and H to the metal particle mostly occurs in step $3b \rightarrow TS3/4b$ but little after $TS3/4b$. In this $3b \rightarrow TS3/4b$ step, the H atomic charge becomes considerably positive, but the O$^-$ and O$^2-$ atomic charges moderately change, suggesting that the H atom mainly participates in the CT to the metal particle. As it goes from $TS3/4b$ to $4b$, the O$^-$ becomes more negatively charged, the O$^2-$ is much more negatively charged, and the H atom becomes much more positively charged. Because the CT occurs little to metal particle in this step ($TS3/4b \rightarrow 4b$), as discussed above, the change in electron distribution mainly occurs in the O$^-$H moiety, suggesting that the O$^-$H$^+$ polarization becomes strong in this step. It is noted that the positive charges of Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$ change to a lesser extent in this OOH formation than in the O$_2$ adsorption, O–O bond cleavage, and O–OH bond cleavage, as shown in Tables 1 and 2. These features suggest that not only CT but also some other factors play important roles in this OOH formation, as discussed below.

In the O$^1$–O$^2$H bond cleavage ($4b \rightarrow TS4/5b \rightarrow 5b$), the O$^2$H group becomes considerably positive at TS4/5b and then returns to moderately positive at $5b$, while the negative charge of the O$^-$ atom and the positive charges of Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$ increase when going from $4b$ to $5b$. These population changes indicate that this step occurs with CT from the metal particle to the OOH moiety. These population changes resemble those by the oxidative addition in organometallic chemistry. This is reasonable because the σ-bond cleavage needs CT from the metal to the σ*-antibonding orbital. Because the CT deeply relates to the electronic structure of the metal particle, the next task is to elucidate the electronic structures of Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$.

### 3.5 Electronic structures of Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$, M–X bond energy ($M = $ Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$; X = H, O, OH, and OOH), and their relation to O$_2$ activation

The 5d-valence band-top and d-band center of the Pt$_{42}$ shell are calculated at higher energy in Pt$_{55}$ than in Ru$_{13}@$Pt$_{42}$, but the d-conduction band-bottom of the Pt$_{42}$ shell is calculated at lower energy in Pt$_{55}$ than in Ru$_{13}@$Pt$_{42}$ (Fig. 5A and B), where the d-band center was calculated using d-valence bands and the DOS energy was corrected according to Baldereschi and coworkers. We checked if the box size for periodic calculation

![Fig. 5](image-url)  
Fig. 5 Density of states (DOS), partial density of states (PDOS) of total d of whole particle and 5d of the Pt$_{42}$ shell in (A) Pt$_{55}$, (B) Ru$_{13}@$Pt$_{42}$, (C) Pt$_{55}$(O$_2$), and (D) Ru$_{13}@$Pt$_{42}$(O$_2$).
influences little the Fermi level after the correction, as shown in Table S4 in the ESI (page S10†), and also we wish to note the 5d-valence band-top energy differs moderately from that of our previous work,47 because of the different computation method, as explained in the ESI (pages S11 to S12†). The O₂ adsorption decreases the density of the d-valence band-top and that of the d-conduction band-bottom in both Pt₅₅ and Ru₁₃@Pt₄₂, as shown by Fig. 5A, B and C, D, indicating that the d-valence band-top mainly participates in the CT from the Pt₄₂ shell to the O₂, and the d-conduction band-bottom mainly participates in the reverse CT from the O₂ to the Pt₄₂ shell. The higher energy d-valence band-top and the lower energy d-conduction band-bottom induce stronger CT from Pt₅₅ to O₂ than that from Ru₁₃@Pt₄₂ to O₂ and, also, stronger CT from O₂ to Pt₅₅ than that from O₂ to Ru₁₃@Pt₄₂, respectively. Consequently, the binding energy of O₂ molecule with Pt₅₅ is larger than that with Ru₁₃@Pt₄₂, as discussed above, but the O₂ moiety is less negatively charged in Pt₅₅(O₂) than in Ru₁₃@Pt₄₂(O₂); see Table 1. As shown in Fig. 5C and D, the d-valence band-top and d-band center exist at higher energy in Pt₅₅(O₂) than in Ru₁₃@Pt₄₂(O₂). The higher energy d-valence band-top and d-band center in Pt₅₅(O₂) than in Ru₁₃@Pt₄₂(O₂) are the origin of the smaller Eₐ value of the O–O bond cleavage on Pt₅₅ than on Ru₁₃@Pt₄₂. Considering these results, we mainly employ the d-valence band-top and the conduction band-bottom for discussion; these DOSs correspond to HOMO and LUMO in the frontier orbital theory in molecular theory. Also, we used the d-band center for discussion because it is an important property representing the electronic structure of metal particles.

In OOH formation, on the other hand, the charge distribution changes to a lesser extent than in the O–O bond cleavage, as mentioned above, but the Eₐ value is considerably different between Pt₅₅ and Ru₁₃@Pt₄₂. This result suggests that some different factor plays an important role in this reaction. One plausible factor is bond dissociation energy (BDE). In OOH formation, M–O₂ and M–H bonds are broken and M–(OOH) and O–H bonds are formed, where M represents Pt₅₅ and Ru₁₃@Pt₄₂. Because the O–H bond formation is common in both Pt₅₅ and Ru₁₃@Pt₄₂ cases, we focus here on M–O₂, M–H, and M–(OOH) bonds. As shown in Scheme 3, the Pt₅₅–O₂ and Pt₅₅–H bonds are stronger than the Ru₁₃@Pt₄₂–O₂ and Ru₁₃@Pt₄₂–H bonds, respectively, in the reactant side, while the Pt₅₅–(OOH) bond is stronger than the Ru₁₃@Pt₄₂–(OOH) bond in the product side. Therefore, two strong Pt₅₅–(O₂) and Pt₅₅–H bonds (the sum of BDEs = 4.48 eV) are converted to one strong Pt₅₅–(OOH) bond (BDE = 1.96 eV) in the OOH formation on Pt₅₅, where the energy loss is 2.52 eV. In the OOH formation on Ru₁₃@Pt₄₂, on the other hand, two weaker Ru₁₃@Pt₄₂–O₂ and Ru₁₃@Pt₄₂–H bonds (the sum of BDEs = 3.52 eV) are converted to one weaker Ru₁₃@Pt₄₂–(OOH) bond (BDE = 1.55 eV), where the energy loss is 1.71 eV. Apparently, the reaction occurs more easily on Ru₁₃@Pt₄₂ than on Pt₅₅ because of the smaller energy loss in the reaction by the former than by the latter. These results lead us to the conclusion that the stronger Pt₅₅–(O₂) and Pt₅₅–H bonds than Ru₁₃@Pt₄₂–(O₂) and Ru₁₃@Pt₄₂–H bonds, respectively, are reasons why OOH formation from adsorbed O₂ and H needs a larger Eₐ on Pt₅₅ than on Ru₁₃@Pt₄₂.

Because the O–OH bond cleavage needs CT from the metal particle to the OOH moiety, the higher energy d-valence band-top and d-band center of Pt₅₅ than those of Ru₁₃@Pt₄₂ are the origin of the smaller Eₐ on Pt₅₅ than on Ru₁₃@Pt₄₂. In addition, the bond energies relating to this O–OH bond cleavage provide clear understanding of the larger reactivity of Pt₅₅ than that of Ru₁₃@Pt₄₂, as follows: the Pt₅₅–(OOH) bond is stronger than the Ru₁₃@Pt₄₂–(OOH) bond, as shown in Scheme 3, and the Pt₅₅–O and Pt₅₅–(OH) bonds are stronger than the Ru₁₃@Pt₄₂–O and Ru₁₃@Pt₄₂–(OH) bonds, respectively. This means that one stronger bond is broken, but two stronger bonds are formed in the O–OH bond cleavage by Pt₅₅ than by Ru₁₃@Pt₄₂. Thus, Pt₅₅ is more reactive for this reaction than Ru₁₃@Pt₄₂.

As discussed above, the stronger Pt₅₅–(O₂) and Pt₅₅–X bonds (X = H, O₂, OH, and OOH) than the Ru₁₃@Pt₄₂–(O₂) and Ru₁₃@Pt₄₂–X bonds, respectively, are responsible for the reactivity difference in OOH formation and O–OH bond cleavage between Pt₅₅ and Ru₁₃@Pt₄₂. Also, it has been supposed that the overly strong binding energy of oxygen-containing species with the Pt electrode is unfavorable for ORR activity.7 Thus, it is of considerable importance to discuss the Pt₅₅–X and Ru₁₃@Pt₄₂–X bond energies and determining factor of these bond energies. Because the Pt₅₅–(O₂) and Ru₁₃@Pt₄₂–(O₂) bond energies were discussed above in terms of the d-valence band-top and d-band center energies, we focus here on Pt₅₅–X and Ru₁₃@Pt₄₂–X bond energies. These bonds are neither pure ionic nor pure covalent, but they are understood to be strongly polarized covalent bonds. Polarized covalent bond energy E_{cov} (A–B) is approximately represented by eqn (1) on the basis of simple Hückel MO theory:82

\[ E_{cov}(A - B) = \sqrt{(\epsilon_A - \epsilon_B)^2 + 4\beta^2}, \]

where \( \epsilon_A \) and \( \epsilon_B \) are the valence orbital energies of A and B, respectively, and \( \beta \) is a resonance integral. This eqn (1) has been employed to discuss the relative bond strengths of various M–R bonds (M = transition metal element and R = alkyl, silyl, etc.).81–87 Because the M–R bond is understood to be a polarized covalent bond, it is likely that this eqn (1) can be applied to the present discussion. Actually, this equation has recently been applied to the discussion of binding energies of H, O, and OH species with Pd₅₅ and Rh₃₅ particles.84 Eqn (1) indicates that \( E_{cov}(A - B) \) becomes larger as the energy difference in valence orbital \( (\epsilon_A - \epsilon_B) \) increases when the \( \beta \) does not differ. Here, the \( \beta \) value is almost the same between Pt₅₅–X and Ru₁₃@Pt₄₂–X because X is bound with the Pt₄₂ shell in both particles. Thus, the valence

![Scheme 3](image-url)  
Scheme 3 Bond energy changes (in eV) in OOH formation followed by O–OH bond cleavage on Pt₅₅ and Ru₁₃@Pt₄₂.

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scheme of bond cleavages on Pt$_{55}$ than on Ru$_{13}@$Pt$_{42}$. Thus, one of the
bond is stronger than the Ru$_{13}@$Pt$_{42}$
d valence band-top of Pt$_{55}$ than that of Ru$_{13}@$Pt$_{42}$. Consequently, the Pt$_{55}$
apparently shown in Scheme 4; for simplicity, O and OH valence
orbital energies are omitted in this Scheme. Due to their larger electronegativities, the energy difference
between the d-valence band-top of the Pt$_{42}$ shell and the valence orbital of H are at lower energy than these d-valence band-tops
due to their larger electronegativities, the energy difference
between the d-valence band-top of the Pt$_{42}$ shell and the valence orbital of X species is larger in Pt$_{55}$ than in Ru$_{13}@$Pt$_{42}$, as
apparently shown in Scheme 4; for simplicity, O and OH valence
orbital energies are omitted in this Scheme. Consequently, the Pt$_{55}$–X
bond is stronger than the Ru$_{13}@$Pt$_{42}$–X bond.

It should be clearly concluded that the higher energy d valence band-top of Pt$_{55}$ than that of Ru$_{13}@$Pt$_{42}$ is the origin of the
stronger Pt$_{55}$–X bond than the Ru$_{13}@$Pt$_{42}$–X bond. The higher energy d valence band-top of Pt$_{55}$ than that of Ru$_{13}@$Pt$_{42}$ is also responsible for the larger O$_2$ adsorption energy to Pt$_{55}$
than to Ru$_{13}@$Pt$_{42}$ and smaller $E_a$ values of the O–O and O–OH
bond cleavages on Pt$_{55}$ than on Ru$_{13}@$Pt$_{42}$. Thus, one of the important characteristic features of Ru$_{13}@$Pt$_{42}$ is the presence of the
valence band-top of the Pt$_{42}$ shell at lower energy than that of Pt$_{55}$; here we wish to mention that higher energy d-valence band-top relates to higher energy d-band center in many cases, indicating that the d-band center is also useful for discussion.

4 Conclusions

O$_2$ adsorption followed by O–O bond cleavage and OOH formation followed by O–OH bond cleavage on Pt$_{55}$ and
Ru$_{13}@$Pt$_{42}$ particles were investigated using DFT computations,
and comparisons were made between Pt$_{55}$ and Ru$_{13}@$Pt$_{42}$.
Several important findings are summarized as follows: (i) O$_2$ is preferentially adsorbed to the vertex Pt and the neighboring edge Pt atoms in a bridging μ$_2$-side-on manner. (ii) The O$_2$
adsorption energy with the Pt$_{42}$ shell is larger in Pt$_{55}$ than in
Ru$_{13}@$Pt$_{42}$. (iii) The O–O bond cleavage occurs with a smaller $E_a$
on Pt$_{55}$ than on Ru$_{13}@$Pt$_{42}$. (iv) The OOH formation from the
adsorbed O$_2$ molecule and H atom occurs with a smaller $E_a$ on
Ru$_{13}@$Pt$_{42}$ than on Pt$_{55}$. The CT occurs much less in this reaction
than in the O–O bond cleavage. The stronger Pt$_{55}$–(O$_2$) and
Pt$_{55}$–H bonds than the Ru$_{13}@$Pt$_{42}$–(O$_2$) and Ru$_{13}@$Pt$_{42}$–H bonds, respectively, are the origin of the larger $E_a$ on Pt$_{55}$ than on Ru$_{13}@$Pt$_{42}$. (v) The OOH formation via H$^+$/e$^-$ addition to the
adsorbed O$_2$ molecule also occurs easily in Pt$_{55}$ similarly to the reaction between adsorbed O$_2$ molecule and H atom, but more
easily in Ru$_{13}@$Pt$_{42}$. And, (vi) the O–OH bond cleavage occurs more easily with much smaller $E_a$ than the O–O bond cleavage of the adsorbed O$_2$ molecule.

The abovementioned differences between Pt$_{55}$ and
Ru$_{13}@$Pt$_{42}$ are understood on the basis of the PDOS of these metal particles. The d-valence band-top and d-band center of the Pt$_{42}$ shell are calculated at higher energy in Pt$_{55}$ than in Ru$_{13}@$Pt$_{42}$, but the d-conduction band-bottom of the Pt$_{42}$ shell is at lower energy in Pt$_{55}$ than in Ru$_{13}@$Pt$_{42}$. Accordingly, the O$_2$ molecule is adsorbed to Pt$_{55}$ more strongly than to Ru$_{13}@$Pt$_{42}$, because the CT from the Pt$_{42}$ shell to O$_2$ and the reverse CT from the O$_2$ to the Pt$_{42}$ shell are more strongly formed with Pt$_{55}$ than with Ru$_{13}@$Pt$_{42}$. Because the O–O bond cleavage needs CT from the metal particle to the O$_2$ moiety, the presence of d-valence band-top at high energy is favorable. Consequently, Pt$_{55}$ is more reactive than Ru$_{13}@$Pt$_{42}$. On the other hand, the reactivity for OOH formation from adsorbed O$_2$ and H depends on the M–O$_2$ and M–H bond energies, as follows: because the Pt$_{55}$–(O$_2$) and Pt$_{55}$–H bonds are stronger than the Ru$_{13}@$Pt$_{42}$–(O$_2$) and
Ru$_{13}@$Pt$_{42}$–H bonds, respectively, OOH formation on Pt$_{55}$ needs a larger $E_a$ than that on Ru$_{13}@$Pt$_{42}$.

The binding energy of oxygen-containing species with Pt-based electrode has been discussed as an important factor for ORR activity. Also, the above discussion suggests that the bond energy is an important property for understanding reactions on the Pt-based electrode. We explored the Pt$_{55}$–X (X = H, O, OH, and OOH) and Ru$_{13}@$Pt$_{42}$–X bond energies and found that the Ru$_{13}@$Pt$_{42}$–X bond is weaker than the Pt$_{55}$–X, and the lower energy d-valence band-top of Ru$_{13}@$Pt$_{42}$ than that of Pt$_{55}$ is the origin of the weaker Ru$_{13}@$Pt$_{42}$–X bond than the Pt$_{55}$–X bond. It is clearly concluded that the lower energy of the d-valence band-top of Ru$_{13}@$Pt$_{42}$ than that of Pt$_{55}$ is one of the important characteristic features of Ru$_{13}@$Pt$_{42}$ in comparison to Pt$_{55}$.

Conflicts of interest

The authors declare no competing financial interest.

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Notes and references