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# Mannich bases derived from lawsone and their metal complexes: synthetic strategies and biological properties

 Abolfazl Olyaei,<sup>a</sup> Mahdieh Sadeghpour<sup>b</sup> and Mehdi Khalaj<sup>c</sup>

Lawsone (2-hydroxynaphthalene-1,4-dione) is a natural product which shows significant biological activity. Aminomethylnaphthoquinone Mannich bases derived from lawsone constitute an interesting class of naphthoquinones and/or their metal complexes have demonstrated a series of important biological properties. So, this review aimed to document the publications concerning the synthesis of aminomethylnaphthoquinone Mannich bases from lawsone, aldehydes and amines and their metal complexes using different conditions, and investigation of their applications.

## 1. Introduction

Molecules with the quinoid structure bearing a hydroxy group on the quinone ring constitute one of the most interesting classes of synthetic compounds in organic chemistry and they are widely distributed in nature. Specifically, naphthoquinone structures receive great attention because of their pharmacological activities such as antibacterial, antiviral and antifungal activities.<sup>1–3</sup> 2-

Hydroxy-1,4-naphthoquinone, or lawsone, or hennotannic acid, has been known for the past 4000 years, is one of the simplest naturally occurring naphthoquinones. It is a red-orange dye, classified as Natural Orange 6 (C.I.75480), present in the leaves of the henna plant (*Lawsonia* spp., family Lythraceae) as well as in the flower of water hyacinth (*Eichhornia crassipes*) and in jewelweed *Impatiens balsamina*.<sup>4,5</sup> It is traditionally used for coloring hair and dyeing nails, skin, wool and cotton.<sup>6–10</sup> Besides it has several other uses including antiaging additive to vulcanized natural rubber,<sup>11</sup> skin protection from ultraviolet radiation,<sup>12,13</sup> oxidation of chlorinated compounds,<sup>14,15</sup> corrosion inhibition for steel,<sup>16,17</sup> and sensitive colorimetric and electrochemical sensor for anions.<sup>18</sup> Extracts of henna are widely used in folk medicine to treat burn

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wounds infected by microorganisms<sup>19</sup> due to its antibacterial properties<sup>20</sup> as well as headaches, lumbago, bronchitis, ophthalmia, syphilis, sores and amenorrhoea.<sup>21</sup> The classical Mannich reaction is one of the most important carbon-carbon bond-forming reactions in organic synthesis because of its atom-economy advantages and application in biologically active molecule syntheses. However, a three-component condensation between structurally diverse substrates containing at least one acidic hydrogen atom, an aldehyde component and an amine reagent leads to Mannich bases. This multicomponent reaction usually occurs under acid catalysis, although catalysis is not mandatory.

3-(Aminomethyl)-2-hydroxy-1,4-naphthoquinones known as aminonaphthoquinone Mannich bases, constitute an interesting class of naphthoquinones and/or their metal complexes have demonstrated a series of important biological properties such as anticancer,<sup>22</sup> antimalarial,<sup>23</sup> antimolluscicidal,<sup>24</sup> antibacterial,<sup>25</sup> cholinesterase inhibitors,<sup>26</sup> leishmanicidal,<sup>27</sup> antiviral,<sup>28</sup> antifungal,<sup>29</sup> anti parasitic<sup>30</sup> activities and they have also been employed as fluorescent compounds and dyes.<sup>31</sup> Because of their prevalent applications, several methods have been reported for the synthesis of these important compounds and their metal complexes. This review will summarize the reported methods for the synthesis of aminonaphthoquinone Mannich base derivatives by multicomponent reactions of lawsone with a non-enolizable aldehyde and a primary or secondary amine *via* the Mannich reaction and their metal complexes using different conditions, and applications of these compounds.

## 2. Synthesis of aminomethylnaphthoquinone Mannich bases

### 2.1. Uncatalyzed reactions

The first synthesis of 2-hydroxy-3-aminomethyl-1,4-naphthoquinones **1** by the Mannich reaction has been reported



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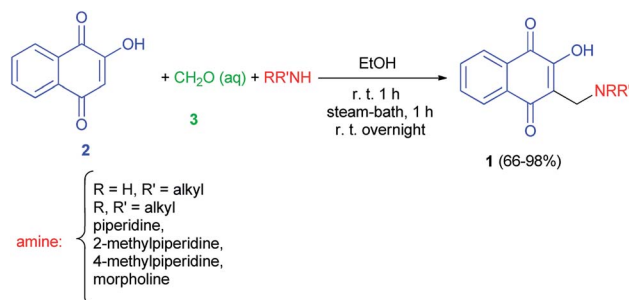
University of Buinzahra Branch from 2014 and now has become to Associate Professor in that University. His research interests include organic and inorganic synthesis, porous coordination polymers based on phosphonates, synthesis, complexation and structural studies of some diimine ligands, nanocomposite green chemistry and so on.

by Leffler and Hathaway in 1948. The reaction mixture of lawsone (**2**), amine and 37% formalin (**3**) in absolute ethanol was allowed to stir at room temperature for one hour was warmed on the steam-bath for an additional hour, and was then left at room temperature overnight gave the corresponding aminonaphthoquinone **1** in 66–98% yields. In this study diethylamine failed to react, on the other hand, primary amines, such as butylamine, gave especially good yields (Scheme 1).<sup>32</sup>

However, two mechanisms may be operating in the formation of the aminomethylnaphthoquinone derivatives (**4**), depending on the reaction conditions: the first mechanism is preformation of the imine *via* the reaction of amine and aldehyde followed by nucleophilic attack of lawsone (**2**) to the activated imine or iminium species, and/or the second mechanism is Knoevenagel-type condensation of lawsone with aldehyde followed by Michael-type 1,4-addition of amine to the alpha, beta-conjugated carbonyl species formed. Apparently, mechanism (1) operates mostly in catalyzed transformations while mechanism (2) will probably be the one operating for uncatalyzed transformations (Scheme 2).

In 1949, Dalglish demonstrated that Mannich bases **5** were prepared by the reaction of lawsone, aldehydes and primary aliphatic amines or 2-aminopyridine in EtOH at room temperature for 3–24 h. The mechanism of the Mannich reaction is still obscure, it being uncertain whether the primary reaction is that of formaldehyde with the amine or with the active hydrogen component (Scheme 3).<sup>33</sup> It has been suggested that the formation of aminonaphthoquinone **5** was occurred through the second mechanism as shown in Scheme 2.

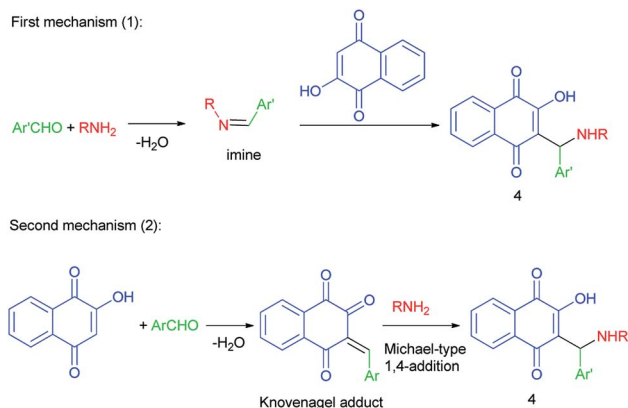
Hydroxynaphthoquinones **1** and **6a–d**<sup>24</sup> were synthesized as previously described methods (Fig. 1).<sup>32,34</sup> They have been submitted to molluscicidal bioassays against the snail *Biomphalaria glabrata*, intermediate host of *Schistosoma mansoni*. Several of the quinones **1** and **6a–d** assayed showed significant molluscicidal activities, and correlation of their activities and electrochemical parameters showed that the first wave reduction potential is an important parameter. The easily reduced quinones ( $>E_{p1c}$ ) were more active against adult snails and against their egg masses, whilst the 3-methylamino-2-hydroxy derivatives presented higher negative reduction potentials and were not active as molluscicides.



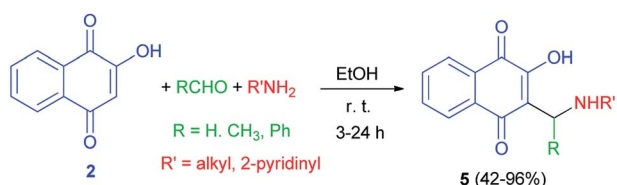
Scheme 1 Synthesis of 2-hydroxy-3-aminomethyl-1,4-naphthoquinones **1**.



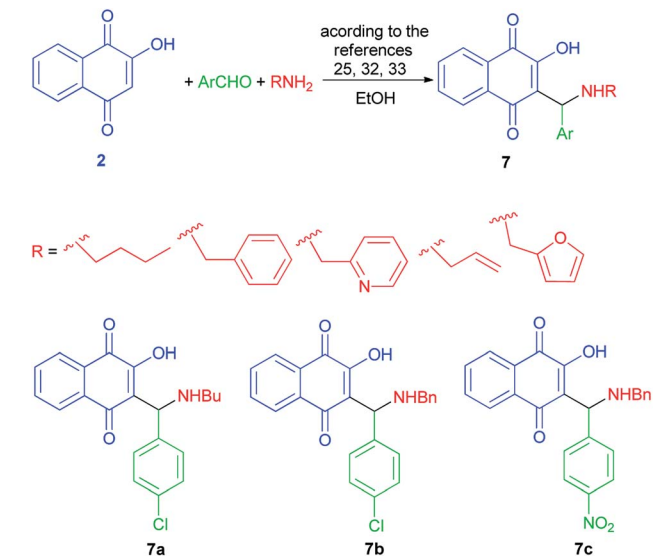
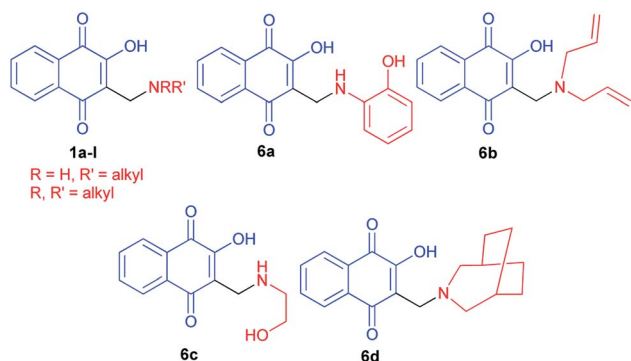
## Review



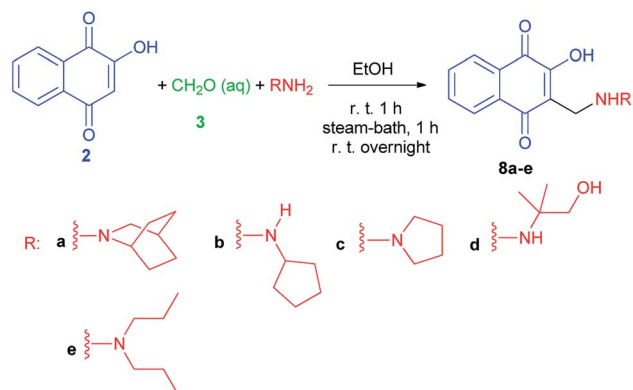
Scheme 2 Proposed mechanisms for the synthesis of aminomethylnaphthoquinones 4.



Scheme 3 Synthesis of aminonaphthoquinines 5.



Scheme 4 Synthesis of aminomethylnaphthoquinone derivatives 7.



Scheme 5 Naphthoquinones 8 obtained by Mannich reaction from lawsone.

Fig. 1 Structures of the synthesized hydroxynaphthoquinones 1 and 6a–d.

Aminomethylnaphthoquinones 7 were synthesized from the Mannich reaction of lawsone (2) with the desired substituted benzaldehydes and primary amines in ethanol as described previously,<sup>32,33</sup> with some modifications (Scheme 4).<sup>25</sup> Although AMNQs 7a–c exhibited similar antiviral activities against BoHV-5, AMNQ 7c was at least 50-fold less cytotoxic than the other two compounds.<sup>3</sup>

2-Hydroxy-3-substituted-aminomethyl naphthoquinones 8a–e,<sup>35</sup> Mannich adducts, were obtained in high quantities by the classic reaction of lawsone (2), using amine and formaldehyde in alcoholic solution (Scheme 5), as previously reported by the Pinto group,<sup>24,36,37</sup> based on the procedure described by Leffler and Hathaway.<sup>32</sup> Compounds 8 exhibit molluscicidal activity<sup>24,36,37</sup> against the snail *Biomphalaria glabrata*, the

causative agent of schistosomiasis. Compound 8e, the most active member of this group, was two times less effective than benznidazole ( $IC_{50}/24\text{ h} = 103.6 \pm 0.6\ \mu\text{M}$ ), the standard anti-*T. cruzi* drug, and 2.6 times more active than crystal violet ( $IC_{50}/24\text{ h} = 536.0 \pm 3.0\ \mu\text{M}$ ), which is recommended by WHO to hemotherapeutic centers in endemic areas for eliminating the parasite in blood used for transfusions.<sup>35</sup>

Krettli *et al.* noted that the naphthoquinones derived from lawsone (compounds 8a and 8e) showed the highest activity against *P. falciparum* chloroquine-resistant blood-stage parasites (clone W2), indicated by their low inhibitory concentration for 50% ( $IC_{50}$ ) of parasite growth. The therapeutic potential of the compounds was evaluated according to the selectivity index, which is a ratio of the cytotoxicity minimum lethal dose which eliminates 50% of cells and the *in vitro*  $IC_{50}$ . Naphthoquinones 8a and 8e, with activities similar to the reference antimalarial chloroquine, were also active against malaria in mice and suppressed parasitaemia by more than 60% in contrast to Noralpachol (9) which was inactive (Fig. 2).<sup>38</sup>



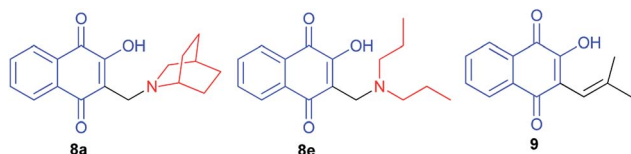
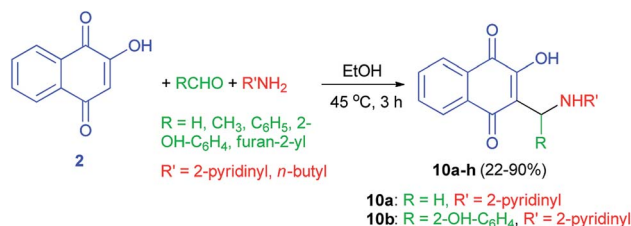


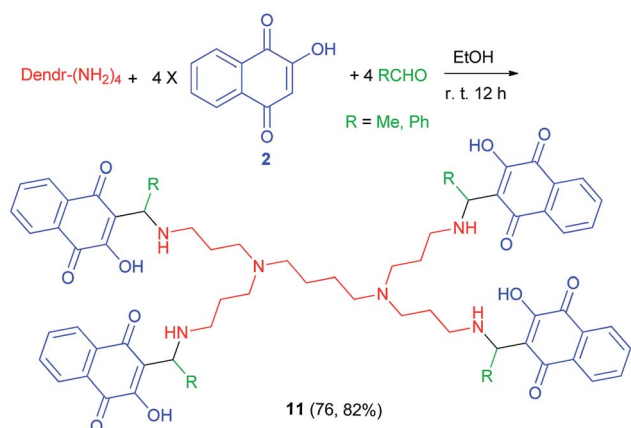
Fig. 2 Structure of the naphthoquinones derived from lawsone (8a, e and 9).



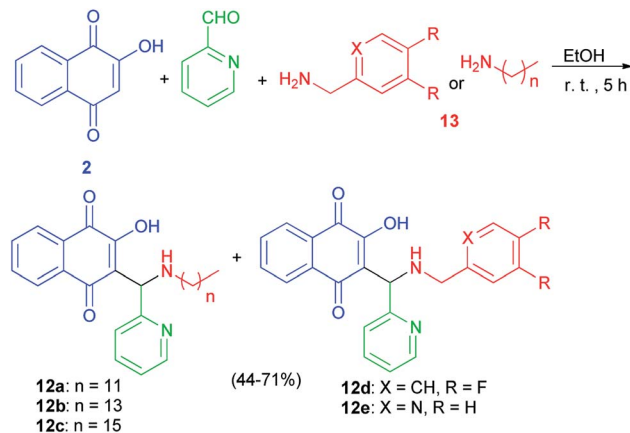
Scheme 6 Synthesis of 2-hydroxy-1,4-naphthoquinones 10.

A series of substituted derivatives of 2-hydroxy-1,4-naphthoquinones **10** were prepared by a Mannich reaction from lawsone (**2**), aryl/alkyl aldehydes and 2-aminopyridine/*n*-butylamine in EtOH at 45 °C for 3 h in 22–90% yields (Scheme 6). Minimum inhibitory concentration (MIC) of the synthesized compounds was measured against *Mycobacterium tuberculosis* (MTB) H37Ra strain. Among them, two compounds **10a,b** showed antibacterial activity within the MIC range of 20–50  $\mu\text{g mL}^{-1}$  while the other compounds did not show antibacterial activity. It was found that compound **10a** showed the high activity anti-tuberculosis among other compounds, although cytotoxicity issues have been observed when evaluated on Vero-cells for further concerns.<sup>39</sup>

The synthesis of two new PPI-G1 dendrimers containing lawsone **11** as pharmacophore is reported by Ndiaye and co-workers. The compounds were synthesized in course of Mannich reaction involving lawsone (**2**), the commercial PPI-G1 (polypropyleneimine-G1) dendrimer with functional amino groups on the exterior and acetaldehyde or benzaldehyde in



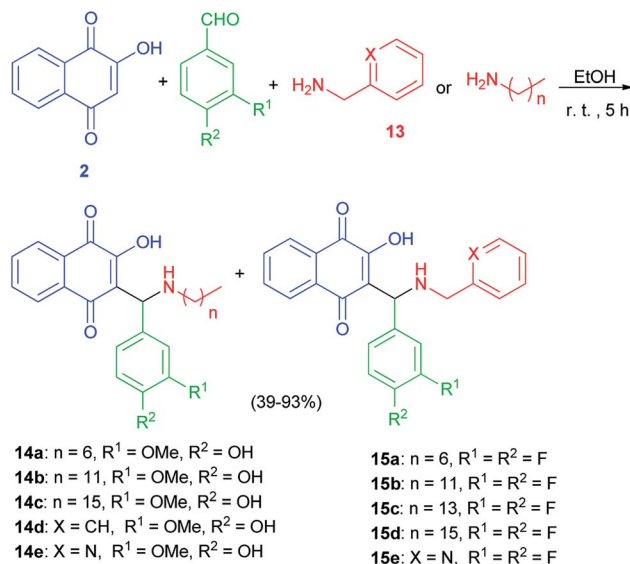
Scheme 7 Synthesis of PPI-G1 dendrimers containing lawsone 11.



Scheme 8 Synthesis of Mannich bases 12a–e.

4 : 1 : 4 molar ratio in absolute EtOH at room temperature for overnight in the dark (Scheme 7). The reaction mechanism could be explained briefly in a few words, *i.e.* condensation of the amino with the acidified aldehyde to give the imine followed by protonation and addition of lawsone anion (mechanism (1)).<sup>40</sup>

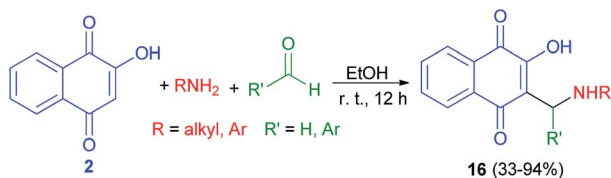
The Mannich bases **12a–e** were prepared in 44–71% yields *via* Mannich reaction of lawsone, 2-pyridylcarboxaldehyde and the corresponding amines **13** in detail, dodecyl amine, tetradecyl amine, hexadecyl amine, 3,4-difluorobenzylamine and 2-pyridylmethylamine in EtOH at room temperature for 5 h (Scheme 8). In addition, closely related vanillyl and 3,4-difluorophenyl derivatives **14a–e** and **15a–e** were prepared. Analogously to the synthesis of compounds **12a–e**, the Mannich reaction of lawsone with vanillins and the corresponding primary amines in EtOH gave the Mannich bases **14a–e** as orange-red solids. In addition, reaction of lawsone with 3,4-difluorobenzaldehyde and various primary amines led to



Scheme 9 Synthesis of Mannich bases 14 and 15.

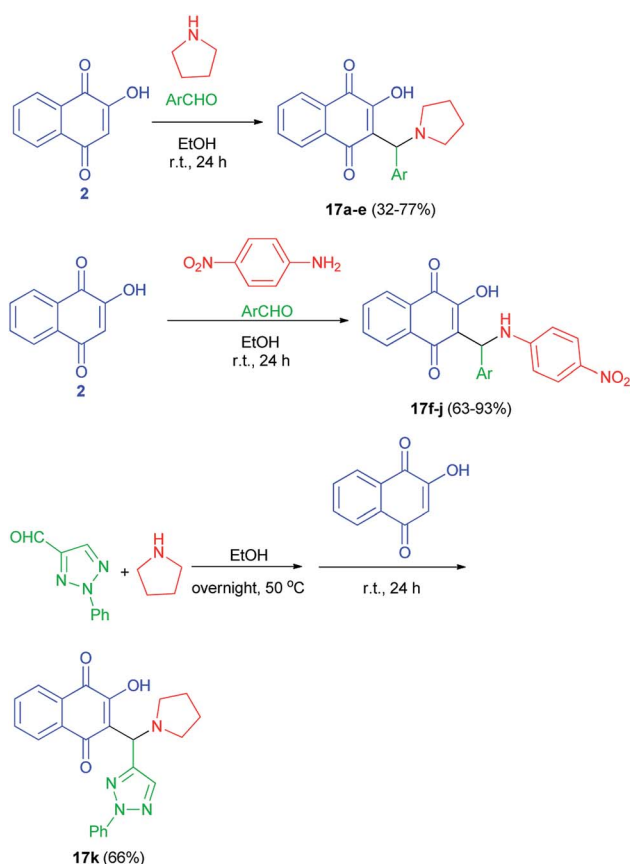




Scheme 10 Synthesis of 3-(aminomethyl)naphthoquinones **16**.

compounds **15a–e** (orange-red solids). These target compounds were obtained in EtOH at room temperature in moderate to high yields (39–93%) (Scheme 9). Substituted lawsonsone Mannich bases **12a–e**, **14a–e** and **15a–e** tested for their biological activities. The new fatty alkyl substituted compounds **12a–c** exhibited strong and selective growth inhibitory activities in the low one-digit micromolar and sub-micromolar range against a panel of human cancer cell lines associated with ROS formation. In addition, compounds **12a–c** revealed sub-micromolar anti-trypanosomal activities against parasitic *Trypanosoma brucei* cells *via* deformation of the microtubule cytoskeleton. The *N*-hexadecyl compound **12c** was also highly active against locally isolated *Entamoeba histolytica* parasite samples exceeding the activity of metronidazole.<sup>30</sup>

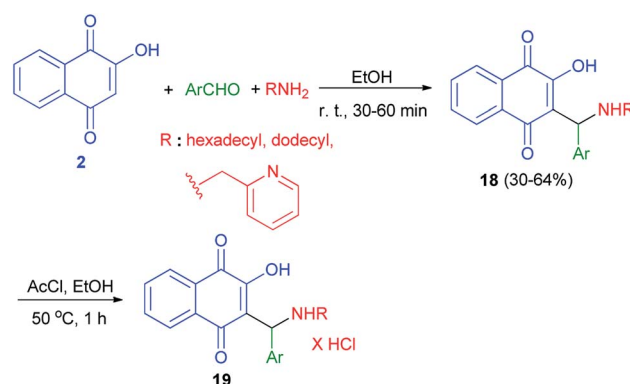
The synthesis of 3-(aminomethyl)-2-hydroxy-1,4-naphthoquinones **16** was carried out following a procedure described previously.<sup>25</sup> The products were isolated as

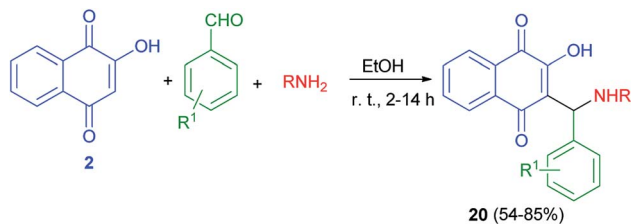
Scheme 11 Synthesis of aminonaphthoquinone Mannich bases **17**.

analytically pure orange powders with yields ranging from 33–94% (Scheme 10). In general, the compounds **16** have shown high to moderate activity against the HL-60 (promyelocytic leukaemia) cell line with  $\text{IC}_{50} = 1.1\text{--}19.2 \mu\text{M}$ . The results suggest that the nature of the aryl moiety introduced in the structure of products by the aldehyde component is crucial to the cytotoxicity, and although the group originated from the primary amine has a lesser influence, aromatic ones were found to suppress the activity.<sup>41</sup>

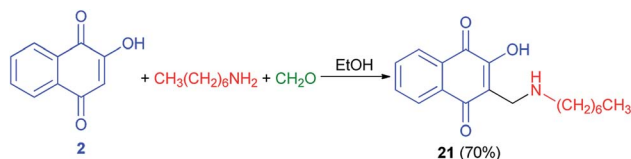
Greco *et al.* succeeded in preparation of amino-naphthoquinone derivatives **17** from lawsonsone (**2**), aldehydes and amines in EtOH at room temperature in the dark for 24 h. Also, for the synthesis of the lawsonsone derivative **17k**, a mixture of 2-phenyl-2H-1,2,3-triazole-4-carbaldehyde and pyrrolidine in ethanol was stirred overnight at 50 °C. After this time, lawsonsone was added and the mixture was allowed to react for 24 h in the dark (Scheme 11). Among the synthetic lawsonsone derivatives, 2-hydroxy-3-((2-hydroxyphenyl)(pyrrolidin-1-yl)methyl)naphthalene-1,4-dione, 2-hydroxy-3-(((4-nitrophenyl)amino)(phenyl)methyl)naphthalene-1,4-dione and 2-hydroxy-3-(((2-hydroxyphenyl)((4-nitrophenyl)amino)methyl)naphthalene-1,4-dione showed high activity against *Candida albicans* ATCC 10231, with minimal inhibitory concentrations (MICs) and minimal fungicidal concentrations (MFCs) ranging from 20 to 330 and from 80 to 330  $\mu\text{g mL}^{-1}$ , respectively. Moreover, they also showed a mechanism of action on exogenous ergosterol.<sup>29</sup>

Biersack *et al.* have been prepared a series of new lawsonsone Mannich bases **18** in 30–64% yields from lawsonsone (**2**), salicylaldehydes or nitrofurfural and dodecylamine or hexadecylamine or 2-aminomethylpyridine in EtOH at room temperature for 30–60 min. Compounds **18** were converted to the hydrochloride salts **19** by reaction with acetyl chloride in ethyl alcohol at 50 °C for 1 h (Scheme 12). These compounds tested for activities against *Leishmania major*, *Toxoplasma gondii*, and *Trypanosoma brucei* parasites. The hydrochloride salts **19** of the Mannich bases **18**, derived from unsubstituted salicylaldehyde and long chained alkyl amines, were selectively and strongly active against *T. gondii* cells and appear to be new promising drug candidates against this parasite. Also, Some of them

Scheme 12 Synthesis of lawsone Mannich bases **18** and their hydrochloride salts **19**.



Scheme 13 Synthesis of aminonaphthoquinone derivatives 20.

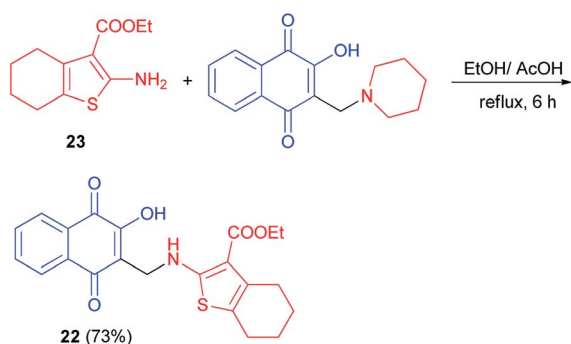


Scheme 14 Synthesis of 3-(N-(1-heptyl)aminomethyl)-2-hydroxynaphthoquinone 21.

proved active against *L. major* promastigotes and up to four times more efficacious against *L. major* amastigotes.<sup>42</sup>

Thakore *et al.* described synthesis of Mannich bases of lawsone 20 by the reaction of lawsone, aromatic aldehydes and aliphatic amines in EtOH at room temperature for 2–14 h in 54–85% yields. It primarily forms the iminium ion by condensation of amine and aldehyde. Subsequently, the nucleophile lawsone attacks iminium ion (analogous to the enolate) to give a Mannich base 20. Overall, the compounds with aliphatic substituent at R position and hydroxyl substituted phenyl ring at R<sup>1</sup> position are associated with the highest activities. The effective IC<sub>50</sub> concentration for compounds are 1.68–4.64 μM. These compounds showed higher anticancer potential against HepG2 cells (Scheme 13).<sup>43</sup>

Moreover, compound 21 was obtained in 70% yield using a direct condensation of *n*-heptylamine on 2-hydroxynaphthoquinone (2) in the presence of formaldehyde in EtOH (Scheme 14). This novel atovaquone derivative 21 was tested for their *in vitro* activity against two apicomplexan parasites of medical importance, *Toxoplasma gondii* and *Plasmodium falciparum*, including resistant strains to atovaquone (*T. gondii*) and chloroquine (*P. falciparum*).<sup>44</sup>

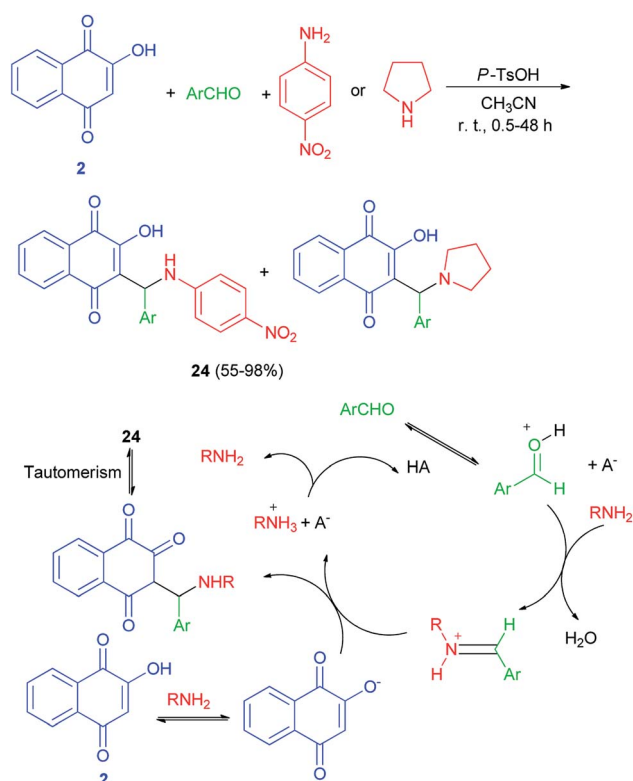


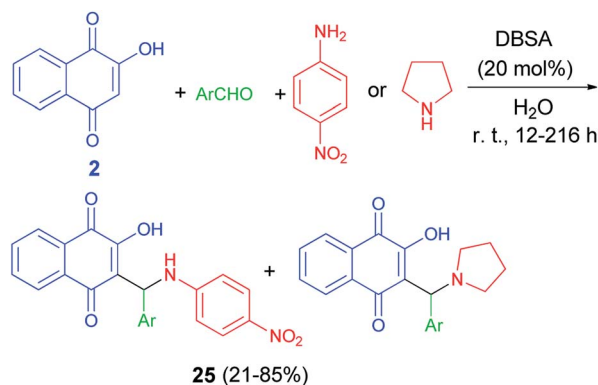
Scheme 15 Synthesis of aminonaphthoquinone 22.

## 2.2. Acid catalyzed reactions

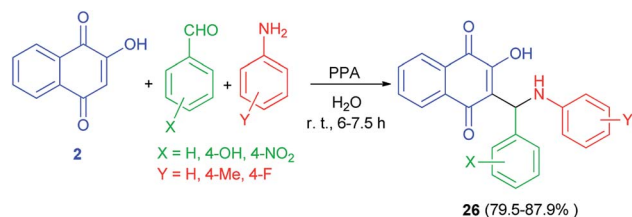
Synthesis of aminonaphthoquinone 22 in 73% yield has been carried out by the reaction of ethyl 2-amino-4,5,6,7-tetrahydrobenzo[*b*]thiophene-3-carboxylate 23 with 2-hydroxy-3-((piperidin-1-yl)methyl)naphthalene-1,4-dione in a mixture ethanol/acetic acid under reflux for 6 h (Scheme 15). It was screened for antioxidant activity. Calculation of the inhibition ratio (%) showed clearly that compound 22 exhibited moderate activity (63.17%) compared with ascorbic acid (89.87%) as a standard antioxidant. Also, it was selected for bleomycin-dependent DNA-damage testing. If the sample to be tested is able to reduce the bleomycin-Fe<sup>3+</sup> to bleomycin-Fe<sup>2+</sup>, DNA degradation in this system will be stimulated, resulting in a positive test for pro-oxidant activity. The results showed that compound 22 has an ability to protect DNA from the induced damage by bleomycin. Thus, it would appear generally that introducing of naphthoquinone moiety enhances the antioxidant properties of 2-aminothiophene compounds.<sup>45</sup>

Greco *et al.* demonstrated that one-pot protocol was developed for the synthesis of a series of new aminonaphthoquinones 24 in 55–98% yields from 2-hydroxy-1,4-naphthoquinone (2), aldehydes and 4-nitroaniline or pyrrolidine by three-component Mannich reaction using catalytic amount of *p*-TsOH in CH<sub>3</sub>CN at room temperature for 0.5–48 h. A reasonable mechanism possibility is shown in Scheme 16. First, the catalyst protonates the carbonyl oxygen of the aldehyde favoring the formation of the iminium ion which reacts

Scheme 16 *p*-TsOH catalyzed synthesis of aminonaphthoquinones 24.



Scheme 17 DBSA catalyzed synthesis of aminonaphthoquinones 25.

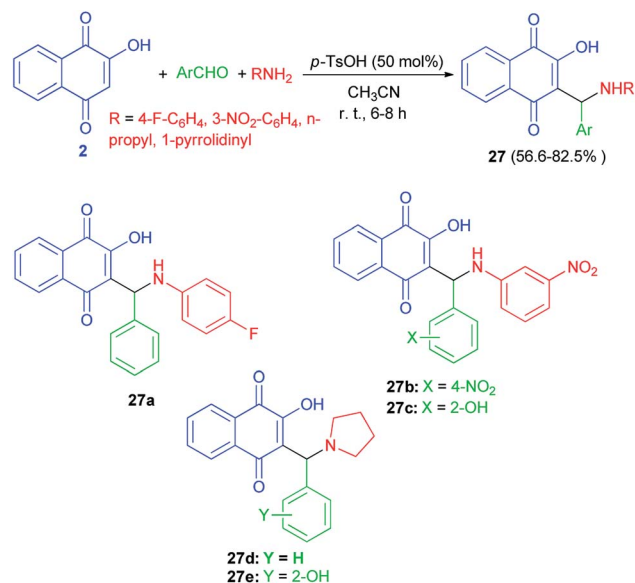
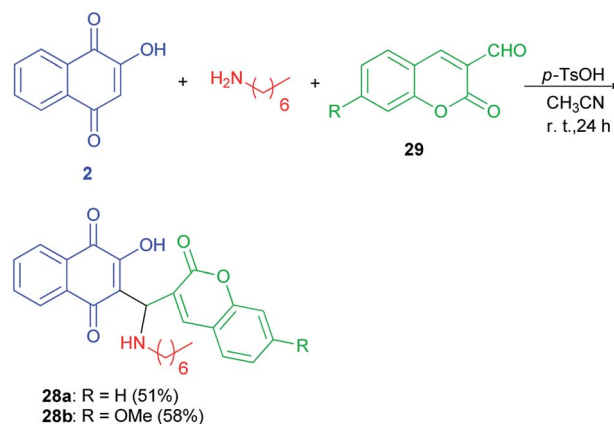


Scheme 18 PPA catalyzed synthesis of hydroxynaphthalene-1,4-dione derivatives 26.

with the nucleophilic species, that may be the ion formed from deprotonation of lawsone by the amine. Finally, the product is obtained after tautomerism and the catalyst is regenerated.<sup>46</sup> Environmentally friendly one-pot protocol was developed for the first synthesis of aminonaphthoquinones 25 in 21–85% yields derived from lawsone (2) *via* a multi-component Mannich reaction in aqueous media using a catalytic amount of dodecyl benzenesulfonic acid (DBSA) (20 mol%), exploring a Brønsted acid-surfactant catalyst (BASC) concept, at room temperature for 12–216 h. Most likely, all of the reaction steps, from the iminium salt formation to the final Mannich adduct, occur inside of the surfactant micelles due to the affinity of the medium for reactants/intermediates (Scheme 17). In this context, DBSA enhances the chemical yield and decreases the reaction time by building a colloidal dispersion system and protonating electrophilic species, the aldehyde and imine, increasing its reactivity towards the nucleophile at the same time.<sup>47</sup>

An efficient one-pot synthesis of lawsone derivatives 26 in 79.5–87.9% yields is accomplished by a three-component reaction of 2-hydroxynaphthalene-1,4-dione (2), aromatic aldehydes and amines carrying electron-donating or electron-withdrawing substituents in water under ambient temperature catalyzed by phenylphosphinic acid for 6–7.5 h (Scheme 18). It should be noted that aliphatic aldehyde reacted poorly under the same conditions. The proposed mechanism for the formation of the compounds 26 is shown in Scheme 2 (mechanism (1)).<sup>48</sup>

A series of Mannich bases of lawsone as 3-alkyl/aryl/heteroaryl substituted aminonaphthoquinones 27 were synthesized in 56.6–82.5% yields from lawsone (2), amines and

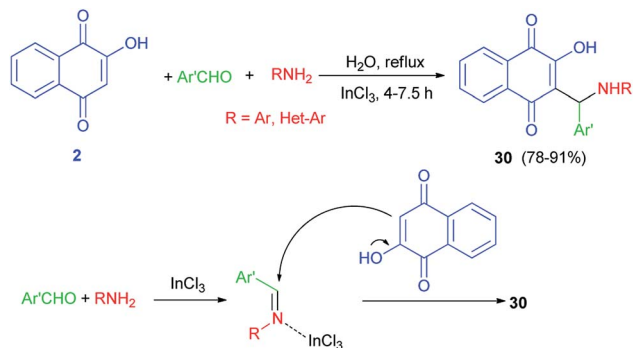
Scheme 19 *p*-TsOH catalyzed Mannich bases of lawsone 27.

Scheme 20 Synthesis of Mannich bases 28a–b.

aromatic aldehydes in CH<sub>3</sub>CN at room temperature in presence of *p*-TsOH (50 mol%) for 6–8 h (Scheme 19). Among synthesized compounds, five compounds (27a–e) showed good antimalarial activity against both chloroquine (CQ)-sensitive (3D-7) and -resistant (RKL-2) strains of *P. falciparum*, which, however, was considerably less than that of the standard reference drug, CQ. The antimalarial activity of these five compounds was found better against sensitive (IC<sub>50</sub> = 0.411–0.502 μg mL<sup>-1</sup>) strain than the resistant (IC<sub>50</sub> = 1.391–2.394 μg mL<sup>-1</sup>) one.<sup>23</sup>

Mannich bases 28a–b derived from lawsone, heptylamine and 3-formylcoumarins (29) have been synthesized in the presence of *p*-TsOH (20 mol%) in CH<sub>3</sub>CN at room temperature for 24 h in the dark in 51 and 58% yields, respectively (Scheme 20). The cyclic voltammetry data of these compounds showed a charge transfer (CT) process from the coumaryl to the naphthoquinonoid group, in spite of the absence of conjugation between these two fragments, with the nitrogen atom playing an important role.<sup>49</sup>



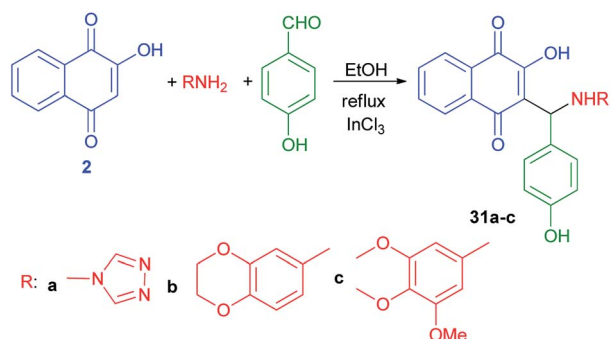
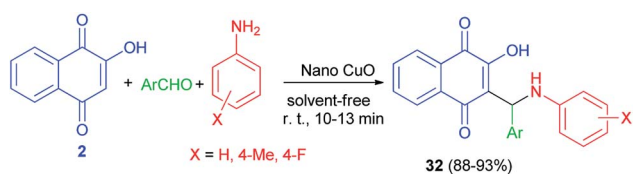
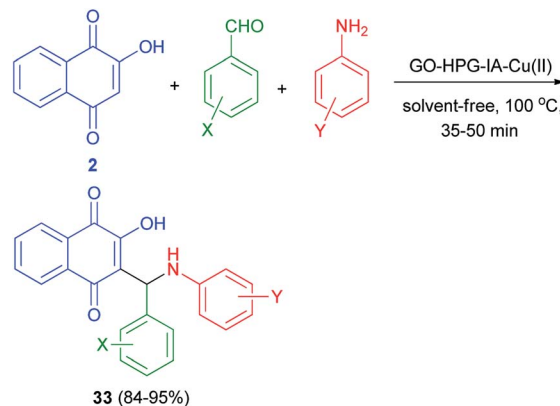
Scheme 21  $\text{InCl}_3$  catalyzed synthesis of aminonaphthoquinones **30**.

### 2.3. Metal catalyzed reactions

An efficient one-pot synthesis of fluorescent hydroxyl naphthalene-1,4-dione derivatives **30** in 78–91% yields by atom-economical and three-component reaction of 2-hydroxynaphthalene-1,4-dione (**2**), aromatic aldehydes and heterocyclic or carbocyclic amines in the presence of a catalytic amount of  $\text{InCl}_3$  in refluxing water for 4–7.5 h is reported. The products were fluorescent in methanol solution emitting at green light (546–560 nm). Possible mechanism for the formation of products **30** as shown in Scheme 21.<sup>31</sup>

Ghods *et al.*<sup>50</sup> have published synthesis of 2-arylamino-naphthoquinone derivatives **31a–c** from lawson (**2**), 4-hydroxybenzaldehyde and aryl amines in the presence of  $\text{InCl}_3$  in refluxing EtOH according to the previously reported procedure (Scheme 22).<sup>31</sup>

Aminonaphthoquinone derivatives **32** have been prepared in 88–93% yields by one-pot three-component condensation

Scheme 22  $\text{InCl}_3$  Catalyzed synthesis of 2-arylamino-naphthoquinones **31a–c**.Scheme 23 Nano CuO catalyzed synthesis of aminonaphthoquinone derivatives **32**.Scheme 24 Synthesis of aminonaphthoquinone derivatives **33** using  $\text{Cu(II)}$  immobilized on graphene oxide.

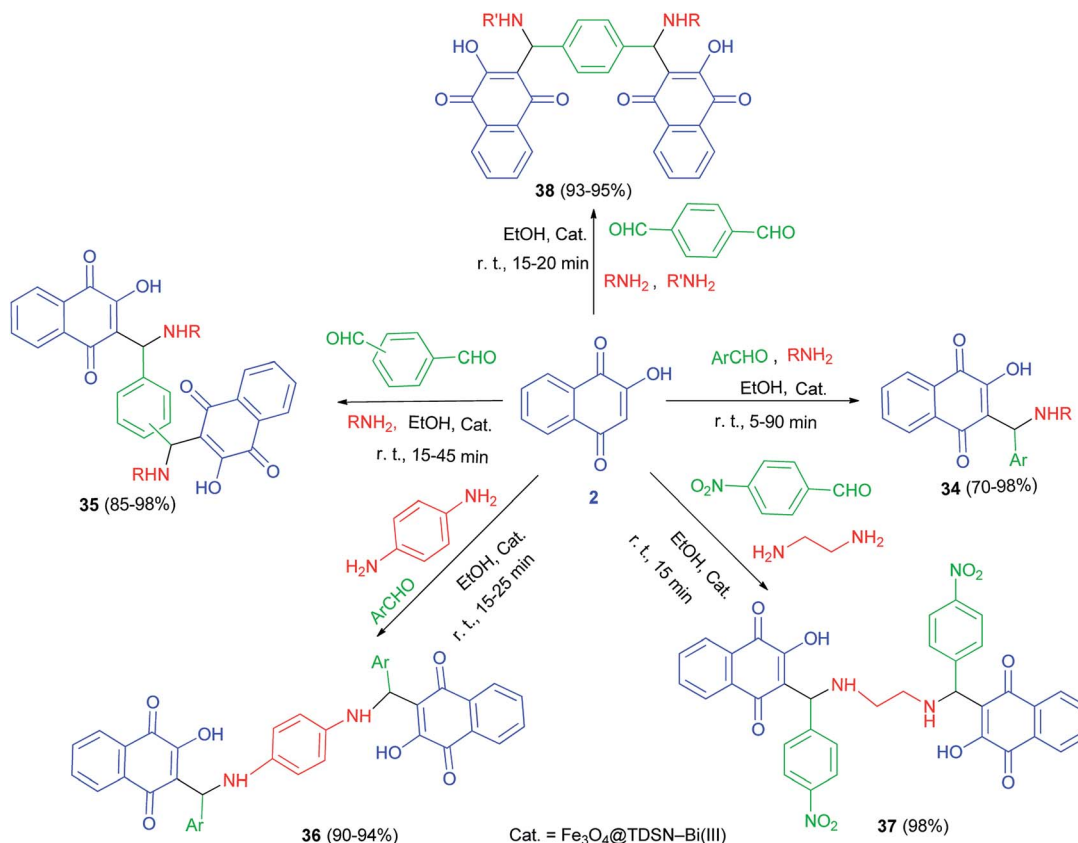
reaction of aromatic aldehydes, aromatic amines, and 2-hydroxynaphthalene-1,4-dione (**2**) in the presence of nano-copper(II) oxide (8 mol%) as catalyst under mild, ambient and solvent-free conditions for 10–13 min. The proposed catalytic mechanism for the preparation of **32** using nano-copper(II) oxide as the catalyst is described in Scheme 23.<sup>51</sup> The proposed mechanism is similar to the first mechanism as shown in Scheme 2.

$\text{Cu(II)}$  immobilized on hyperbranched polyglycerol (HPG) functionalized graphene oxide catalyzed synthesis of aminonaphthoquinones **33** in 84–95% yields *via* one-pot three-component condensation reaction of lawson (**2**), benzaldehyde and aniline derivatives at 100 °C for 35–50 min under solvent-free conditions (Scheme 24). According to the plausible mechanism, firstly, the aldehyde is activated by the GO-HPG-IA- $\text{Cu(II)}$  catalyst which upon nucleophilic attack of amine to form imine. Subsequently, lawson attacks to imine in the presence of the GO-HPG-IA- $\text{Cu(II)}$  catalyst and undergoes tautomeric proton shift to generate the aminonaphthoquinone derivative **33** and releases the catalyst for the next run.<sup>52</sup>

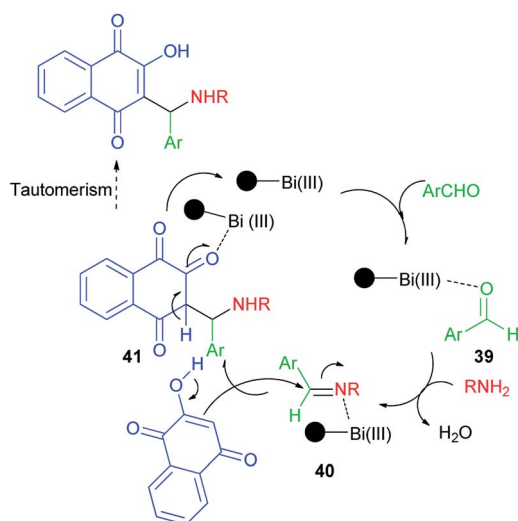
Synthesis of mono-aminonaphthoquinone derivatives **34** catalyzed by  $\text{Bi(III)}$  immobilized triazine dendrimer-stabilized magnetic nanoparticle  $\text{Fe}_3\text{O}_4@\text{TDSN-Bi(III)}$  (3 mol%) *via* a one-pot three-component reaction of lawson, aldehydes containing electron withdrawing or electron-donating groups and amines (anilines, 4-methylaminopyridines, *i*-butylamine, pyrrolidine and piperidine) in EtOH as a green solvent at room temperature for 5–90 min in 70–98% yields. This catalytic system also showed excellent activity in the synthesis of symmetric and unsymmetric bis-aminonaphthoquinones **35–38** from dialdehyde (terephthalaldehyde and isophthalaldehyde) and/or diamines (benzene-1,4-diamine and 1,2-ethylenediamine) in EtOH at room temperature for 15–45 min in high yields (85–98%) and purity *via* an easy work-up procedure (Scheme 25). A plausible mechanism for the  $\text{Fe}_3\text{O}_4@\text{TDSN-Bi(III)}$  catalyzed synthesis of aminonaphthoquinones is depicted in Scheme 26. First, the aldehyde is activated by the catalyst to give **39** which upon nucleophilic attack of amine forms imine **40**. Then, 2-hydroxynaphthalene-1,4-dione attacks imine **40** in







Scheme 25 Synthesis of mono- and bis-aminonaphthoquinone derivatives 34–38.

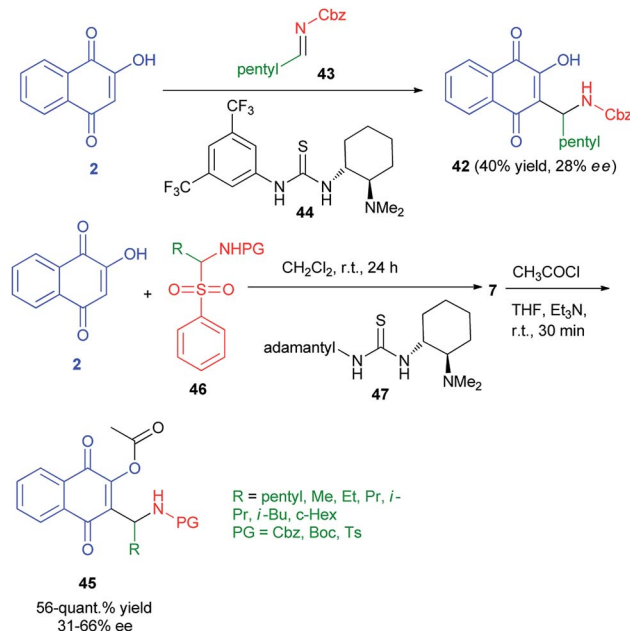
Scheme 26 Plausible mechanism for the  $\text{Fe}_3\text{O}_4\text{@TDSN-Bi(III)}$  catalyzed synthesis of aminonaphthoquinones 34–38.

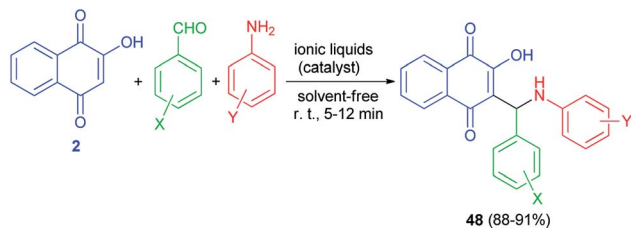
the presence of the catalyst to furnish intermediate **41** desired products and releases the catalyst for the next run.<sup>53</sup>

#### 2.4. Other-types of catalyzed reactions

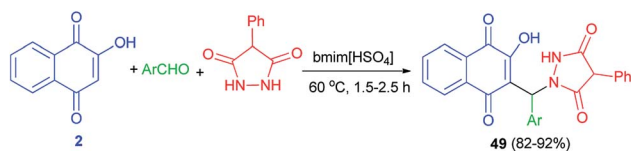
Enantioenriched naphthoquinone Mannich base **42** by organocatalyzed nucleophilic additions of 2-hydroxy-1,4-

naphthoquinone (**2**) to *in situ* formed imines has been reported by Westermann and Ayaz. In this process, lawsone (**2**) to preformed imine **43** using Takemoto's catalyst **44** (10 mol%) afforded the Mannich base **42** in only 40% yield and a very low

Scheme 27 Preparation of Mannich bases **42** and **45**.



Scheme 28 Ionic liquid catalyzed synthesis of amino-naphthoquinones 48.

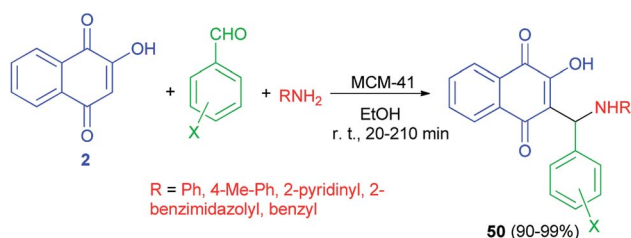


Scheme 29 Task-specific ionic liquid catalyzed synthesis of naphthoquinone-urazole hybrids 49.

enantioselectivity of 28% ee. Also, enantioenriched naphthoquinone Mannich bases 45 have been synthesized by the reaction of 2 with  $\alpha$ -amid sulfone 46 in the presence of organocatalyst 47 in CH<sub>2</sub>Cl<sub>2</sub> at room temperature for 24 h. Then, the reaction of title product with acetyl chloride in the presence of Et<sub>3</sub>N in THF at room temperature for 30 min gave the corresponding products 45 in 56-quant.% yield and 31-66% ee (Scheme 27).<sup>54</sup>

Condensation of 2-hydroxynaphthalene-1,4-dione (2), anilines, and aromatic aldehydes in the presence of the mild basic ionic liquids, (A) DBU [CH<sub>3</sub>COO] (10 mol%), (B) Pyr [CH<sub>3</sub>COO] (15 mol%), (C) Pyr[HCOO] (15 mol%), (D) Pip [CH<sub>3</sub>COO] (10 mol%), (E) Pip[HCOO] (10 mol%), (F) [Hmim] [HCOO] (15 mol%) and (G) 3-HPAA (15 mol%) as the most efficient catalysts with respect to the reaction time and yields under solvent-free and ambient conditions for 5-12 min afforded the corresponding aminonaphthoquinones 48 in high to excellent yields (Scheme 28).<sup>55</sup> The proposed mechanism for the preparation of compounds 48 is similar to the first mechanism as indicated in Scheme 2.

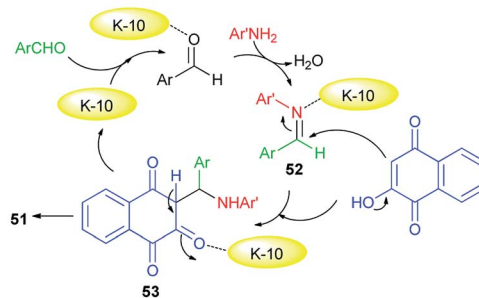
Khurana *et al.*<sup>56</sup> described synthesis of naphthoquinone-urazole hybrids 49 in high yields (82-92%) *via* one-pot condensation of lawsone (2), various aromatic aldehydes having electron withdrawing and electron-releasing groups and 4-phenylurazole using task specific ionic liquid (bmim[HSO<sub>4</sub>])



Scheme 30 MCM-41 catalyzed synthesis of amino-methyl-naphthoquinones 50.



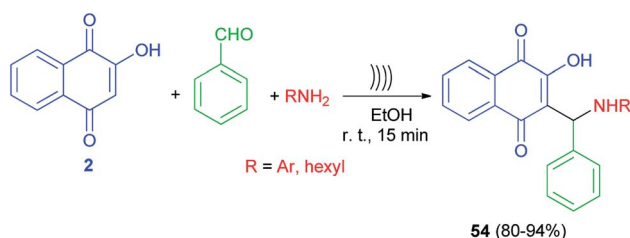
Scheme 31 Montmorillonite K-10 catalyzed synthesis of amino-naphthoquinones 51.



(10 mol%) at 60 °C for 1.5-2.5 h (Scheme 29). A mechanistic rationale portraying the probable sequence of steps is similar to the second mechanism as shown in Scheme 2. All newly synthesized compounds were screened for *in vitro* antioxidant and anticancer activities against human breast (T47D), colon (HCT-15), lung (NCI-H522), liver (HepG-2) and ovary (PA-1) cancer cell lines. The *in vitro* antioxidant and anticancer activity of these compounds revealed that all of the synthesized naphthoquinone-urazole hybrids 49 have a significant activity. 3-Bromophenyl, 2-methylphenyl, 2-naphthyl and 3-methylphenyl derivatives were more active among all the naphthoquinone-urazole hybrids.

Naimi-Jamal *et al.* succeeded in preparation of 2-hydroxy-1,4-naphthoquinone derivatives 50 from 2-hydroxynaphthalene-1,4-dione (2), aromatic aldehydes and heterocyclic or carbocyclic amines using MCM-41 nanoporous as catalyst in EtOH at room temperature for 20-210 min. The aromatic aldehydes carrying both electron-withdrawing and electron-releasing substituents were converted to their corresponding products in high yields (90-99%) (Scheme 30).<sup>57</sup>

An efficient one-pot protocol for the synthesis of 2-((substituted amino)(4-phenyl)methyl)-3-hydroxy-naphthalene-1,4-diones 51 has been developed by the three-component reaction of 2-hydroxynaphthalene-1,4-dione (2), aromatic



Scheme 32 Synthesis of Mannich bases 54 derived from lawsone under ultrasonic irradiation.



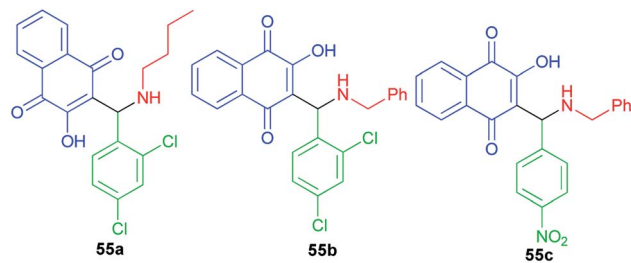


Fig. 3 Structures of 3-aminomethylnaphthoquinones **55a–c** active against HSV-1.

aldehydes and anilines/heterocyclic amines using montmorillonite K-10 (10 mol%) as a catalyst in EtOH at room temperature for 8–10 h. Benzaldehydes and aromatic amines carrying both electron-donating and electron-withdrawing substituents and heterocyclic amines displayed high reactivity under this optimized conditions to afford the desired products in 81–93% yields. A speculative mechanistic explanation for this reaction is provided in Scheme 31. In the first step, aromatic amines attacked the activated benzaldehydes to form activated imines **52**. The latter compounds reacted with lawsone, to give an intermediate **53** that underwent tautomerization to yield the products. Montmorillonite K-10 is likely to enhance the rate of Mannich reaction.<sup>58</sup>

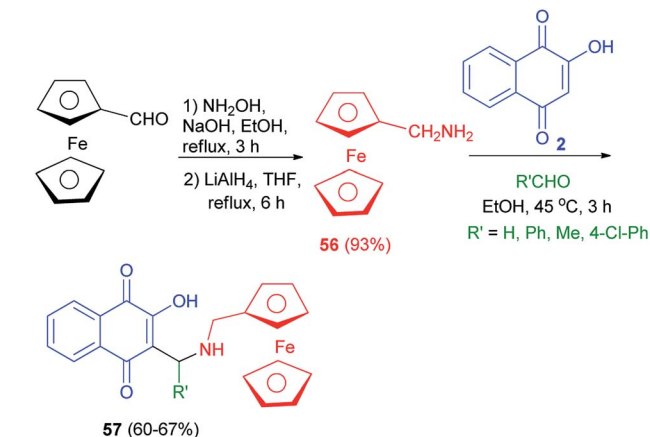
Lopez-Lopez *et al.* developed a non-catalytic effective procedure to prepare Mannich base lawsone derivatives **54** in 80–94% yields using one-pot three-component reaction of lawsone (**2**), amines and benzaldehyde promoted by ultrasound irradiation at room temperature for 15 min (Scheme 32).<sup>59</sup> A possible mechanism for the formation of **54** is proposed according to the first mechanism as indicated in Scheme 2.

A small series of 3-aminomethylnaphthoquinones **55a–c** active against HSV-1 with low *in vitro* and *in silico* toxicity profile has been evaluated by Paixao and co-workers. In spite of the highest antiviral activity against HSV-1 of compound **55b**, the most promising molecule is probably compound **55a** due to its greater SI value ( $SI = CC_{50}/EC_{50}$ ). All compounds exhibited L-phase of lytic replication. Compounds **55a–b** also inhibited IE and E phases to different degrees and they affect gD protein expression (L-phase). Structural features, such the nature of the substituent on the nitrogen atom (benzyl *versus* *n*-butyl) the conformation and LUMO distribution and energy profiles seem to modulate the antiviral activity of these compounds (Fig. 3).<sup>28</sup>

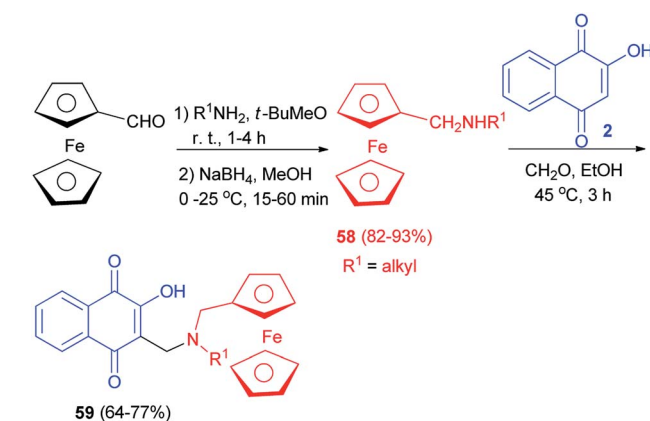
### 3. Synthesis of aminomethylnaphthoquinone Mannich base–metal complexes

#### 3.1. Complexes derived from ferrocenylalkanamines

Ferrocenylmethylamine **56** was first synthesized from ferrocene carboxaldehyde by condensation of hydroxylamine on the aldehyde function followed by a reduction with  $LiAlH_4$ . The amine **56** was obtained in 93% global yield. Ferrocenyl amino-hydroxynaphthoquinones **57**, analogues of atovaquone, were



Scheme 33 Synthesis of ferrocenyl aminohydroxynaphthoquinones **57**.



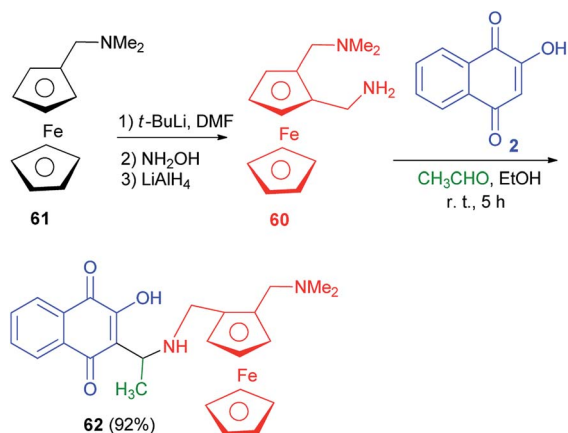
Scheme 34 Synthesis of ferrocenyl aminohydroxynaphthoquinones **59**.

synthesized from 2-hydroxyquinone, ferrocenyl amines **56** and aldehydes in EtOH at 45 °C for 3 h in 60–67% yields (Scheme 33).<sup>60</sup>

Condensation of commercial amines on ferrocene carboxaldehyde followed by reduction with  $NaBH_4$  led to secondary ferrocenyl amines **58** in high global yield (82–93%). The ferrocenyl derivatives **59** were then prepared by condensation of lawsone (**2**) with the formaldehyde in the presence of the corresponding ferrocenylmethylamines **59** in EtOH at 45 °C for 3 h. Compounds **59** were obtained in 64–77% yields (Scheme 34).<sup>60</sup> Ferrocenic atovaquone derivative **59** composed of the hydroxynaphthoquinone core plus an amino-ferrocenic group and an aliphatic chain with 6–8 carbon atoms were found to be significantly active against *T. gondii*. Moreover, this novel compound were also effective against the atovaquone-resistant strain of *T. gondii* (AtoR).<sup>44</sup>

Following the procedure described in the literature,<sup>60</sup> the diamine **60** was prepared from commercially available *N,N*-dimethylaminomethylferrocene (**61**). The treatment of compound **60** with 2-hydroxynaphthoquinone (**2**) in the presence of acetaldehyde in EtOH at room temperature for 5 h afforded the compound **62** as a mixture of diastereomers in 55/

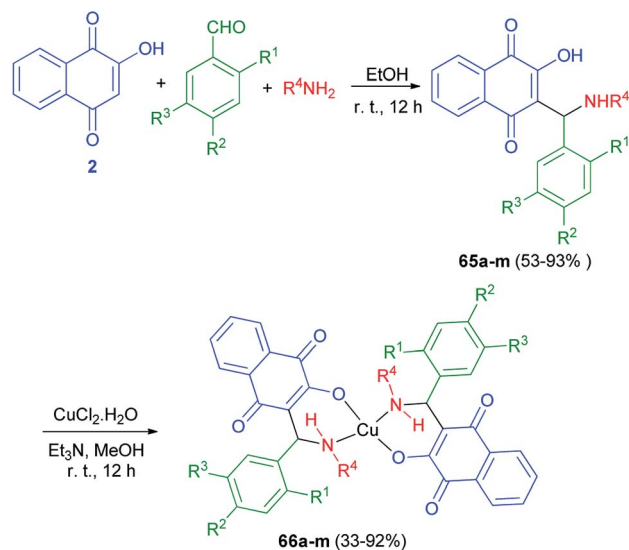




Scheme 35 Synthesis of 3{N-(2-(N,N-dimethylaminomethyl)ferrocenylmethyl)-1-aminoethyl}-2-hydroxynaphthoquinone **62**.

45 ratio (92% yield) which cannot be separated by silica gel column chromatography (Scheme 35).<sup>60</sup>

The new Mannich bases **63a–c** were prepared in 57–98% yields *via* Mannich reaction of lawsone (**2**), aryl carboxaldehyde (2-pyridylcarboxaldehyde) for **63a**, 4-pyridylcarboxaldehyde for **63b**, and 3,4-difluorobenzaldehyde for **63c** and ferrocene-1-ylmethylamine (**64**) in EtOH at room temperature for 5 h. Analogously, compound **63d** was obtained in 25% yield from the reaction of **2** with ferrocene-1-yl carboxaldehyde and heptylamine in EtOH at room temperature for 5 h (Scheme 36). The 2-pyridyl derivative **63a** was distinctly more active than its analogs **63b–d** in breast, prostate and pancreatic cancer cells. Compound **63a** also exhibited greater antiproliferative effects in the androgen-receptor negative PC-3 prostate and Pgp-expressing KB-V1/Vbl cervix carcinoma cell lines. Compound **63a** reached sub-micromolar activities in these aggressive

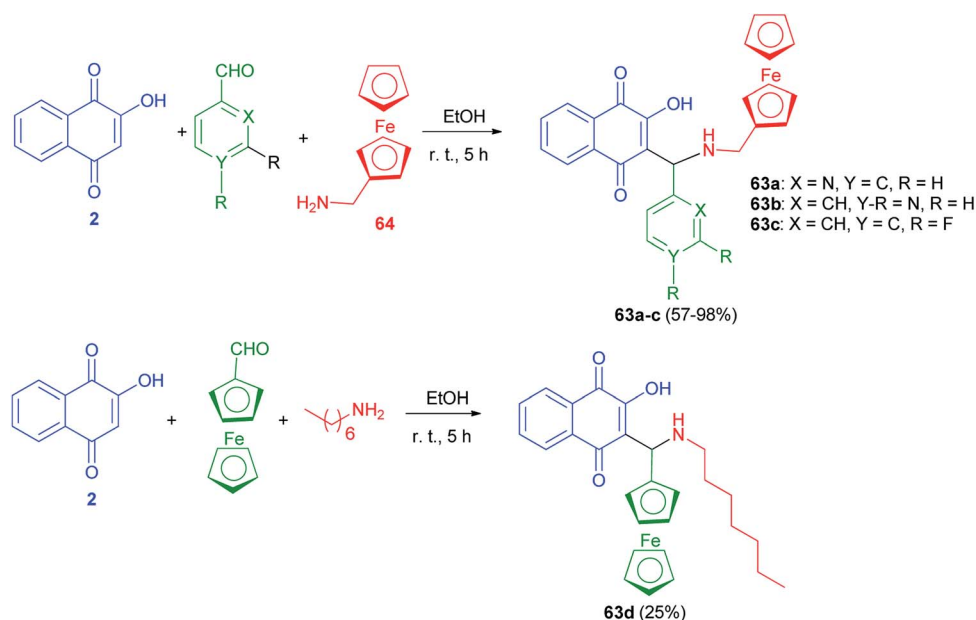


Scheme 37 Synthesis of Mannich bases of **65** and their copper complexes **66**.

cancer cells and, thus, features a promising drug candidate for the efficient treatment of hormone- or multidrug-resistant cancer types.<sup>22</sup>

### 3.2. Copper complexes

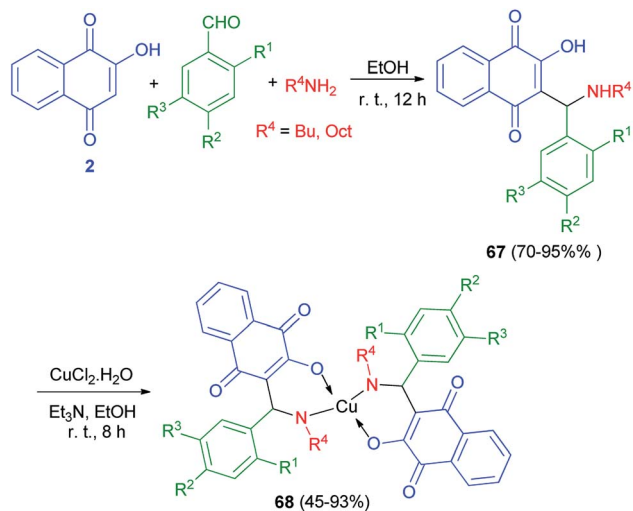
A series of novel Mannich bases **65a–m** derived have been synthesized in 53–93% yields from 2-hydroxy-1,4-naphthoquinone (**2**), substituted benzaldehydes and various primary amines ( $\text{NH}_2\text{R}^4$ ,  $\text{R}^4 = n$ -butyl, benzyl, allyl, 2-furfuryl) in EtOH at room temperature for 12 h in the dark. Followed, their  $\text{Cu}^{2+}$  complexes **66a–m** have been prepared in 33–92% yields from the reaction of the compounds **65a–m** with  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  in



Scheme 36 Ferrocene modified lawsone Mannich bases **63**.





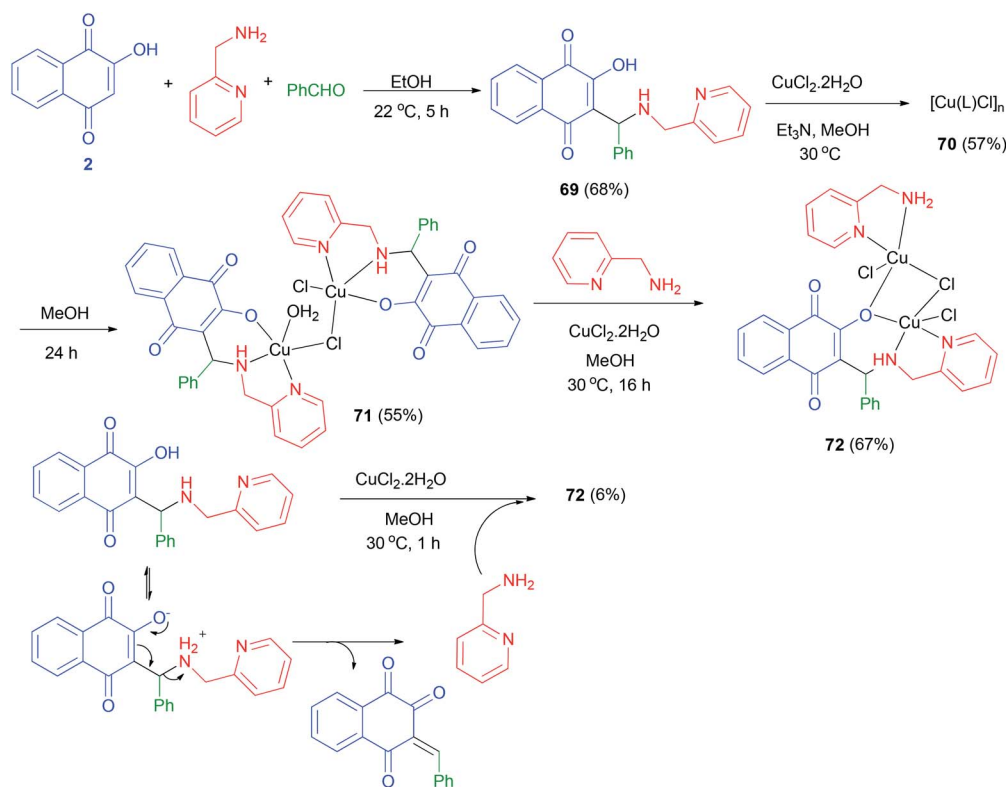


Scheme 38 Synthesis of Mannich bases **67** and their copper complexes **68**.

MeOH in the presence of  $\text{Et}_3\text{N}$  in the dark for 12 h at room temperature (Scheme 37). The antimicrobial activity of all compounds has been tested. In general, Mannich bases were more active than complexes, **65k** ( $\text{R}^1 = \text{OH}$ ;  $\text{R}^2 = \text{H}$ ;  $\text{R}^3 = \text{Me}$ ;  $\text{R}^4 = \text{Bn}$ ) and **65m** ( $\text{R}^1 = \text{OH}$ ;  $\text{R}^2 = \text{H}$ ;  $\text{R}^3 = \text{Br}$ ;  $\text{R}^4 = \text{Bn}$ ) being the most potent inhibitors. The MIC for the most active compound **65k** against *S. Coli* was  $20 \mu\text{mol L}^{-1}$  ( $8 \mu\text{g mL}^{-1}$ ), better than chloramphenicol ( $90 \mu\text{mol L}^{-1}$ ) and well below most values reported for other naphthoquinones.<sup>25</sup>

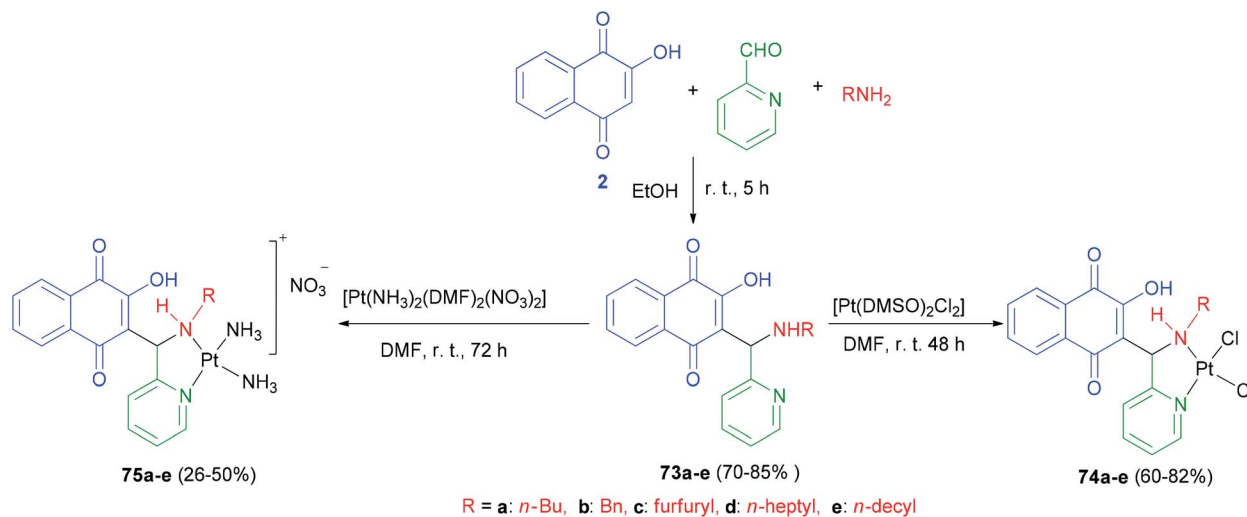
A new series of Mannich bases **67** derived from 2-hydroxy-1,4-naphthoquinone (**2**), substituted benzaldehydes and two primary amines, and their  $\text{Cu}^{2+}$  complexes **68** were synthesized by Cardoso and co-workers. At the first stage, 2-hydroxy-1,4-naphthoquinone (**2**) undergo an amino alkylation with butylamine or octylamine and benzaldehydes in EtOH at room temperature for 12 h, which led to lawsone derivatives **67** in high yields 70–95%. Complexes **68** were obtained by addition of triethylamine to an ethanolic suspension of compounds **68** and  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  under stirring at room temperature for 8 h, with yields varying from 45 to 93% (Scheme 38). Compounds **68** evaluated for their potential as selective cholinesterase inhibitors (ChEIs). Eight copper complexes were identified and characterized as potent reversible and selective ChEIs with inhibitory potencies ( $\text{IC}_{50}$ ) and constants of inhibition ( $K_i$ ) ranging from 1.24 to  $11.5 \mu\text{mol L}^{-1}$ . One of the compounds was particularly promising, showing  $\text{IC}_{50}$  and  $K_i$  values of  $1.24 \pm 0.01$  and  $1.06 \pm 0.01 \mu\text{mol L}^{-1}$ , respectively, for huAChE.<sup>26</sup>

A novel versatile tridentate 3-(aminomethyl)naphthoquinone proligand, 3-[N-(2-pyridylmethyl)aminobenzyl]-2-hydroxy-1,4-naphthoquinone (**69**), was obtained from the Mannich reaction of 2-hydroxy-1,4-naphthoquinone (**2**) with 2-aminomethylpyridine and yielded a polymeric compound **70** in 57% yield. When the reaction mixture was left stirring under the same conditions for over 5 h benzaldehyde in EtOH at  $22^\circ\text{C}$  for 5 h in 68% yield. The reactions of **69** with  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  in the presence of  $\text{Et}_3\text{N}$  in MeOH at  $30^\circ\text{C}$  yielded a polymeric compound **70** in 57% yield. When the reaction mixture was left



Scheme 39 Synthesis of Mannich base **69** and their complexes **71** and **72**.





Scheme 40 Synthesis of 3-(aminomethyl)naphthoquinones **73** and their Pt-complexes **74** and **75**.

stirring under the same conditions for over 5 h darkening was observed. Dark green solid **71** isolated after 24 h in 55% yield. The reaction of complex **71** with 2-aminomethylpyridine in using  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  in MeOH at 30 °C for 16 h in afforded complex **72** in 67% yield. Complex **72** could also be obtained in 6% yield from decomposition of **69** in methanol using  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  at 30 °C for 1 h (Scheme 39).<sup>61</sup>

### 3.3. Platinum complexes

The first examples of platinum(II) complexes of 3-(aminomethyl)naphthoquinone Mannich bases **73** have been synthesized by Vargas and coworkers. Proligands 3-[( $\text{R}^1$ -amino)(pyridin-2-yl)methyl]-2-hydroxy-1,4-naphthoquinones ( $\text{R} = n\text{-Bu}$ , **73a**; Bn, **73b**; furfuryl, **73c**; *n*-heptyl, **73d** and *n*-decyl, **73e**) were synthesized in 70–85% yields from the Mannich reactions of lawsone (**2**) with the respective primary amines and 2-pyridinecarboxaldehyde, in ethanol, under stirring, at room temperature for 5 h. The reactions of equimolar amounts of **73a–e** and *cis*-[Pt(DMSO)<sub>2</sub>Cl<sub>2</sub>] in dimethylformamide (DMF) at room temperature for 48 h in the dark yielded the chlorido complexes *cis*-[Pt(**74a–e**)Cl<sub>2</sub>] **74a–e** in 60–82% yields as yellow solids (Scheme 40).

The amino complexes *cis*-[Pt(**75a–e**)(NH<sub>3</sub>)<sub>2</sub>]NO<sub>3</sub> **75a–e** were obtained as orange powders in 26–50% yields from the reactions of the solid proligands **73a–e** with a solution of *cis*-[Pt(NH<sub>3</sub>)<sub>2</sub>(DMF)<sub>2</sub>](NO<sub>3</sub>)<sub>2</sub> in DMF at room temperature in the dark for 72 h (Scheme 40). The cytotoxic activities of all compounds have been tested for different cancer cell lines. Proligands **73d** and **73e** have exhibited high activity against the tested cell lines, although they were only moderately active against the PC-3 cell line (IC<sub>50</sub> = 29.9 and 15.6 mmol L<sup>-1</sup>, respectively). In general the compounds with the longest carbon chains ( $\text{R} = n\text{-heptyl}$  and *n*-decyl) have exhibited the highest activities.<sup>62</sup>

2-Hydroxy-3-(aminomethyl)-1,4-naphthoquinones **73a,d,e** derivatives, chlorido **74a,d,e** and amino **75a,d,e** Pt<sup>2+</sup> complexes

were synthesized as described in the literature.<sup>44</sup> The DMF solutions of compounds [Pt(**73a,d,e**)(DMF)<sub>2</sub>](NO<sub>3</sub>)<sub>2</sub> were obtained by reacting a solution of AgNO<sub>3</sub> with **73a,d,e** in DMF at 50 °C for 24 h. [Pt(**76a,d,e**)(H<sub>2</sub>O)<sub>2</sub>](NO<sub>3</sub>)<sub>2</sub> **76a,d,e** were obtained by dilution of the DMF solutions with water to a final concentration of 95 : 5 (H<sub>2</sub>O : DMF). The water solution of [Pt(**68a**)(H<sub>2</sub>O)<sub>2</sub>](NO<sub>3</sub>)<sub>2</sub> **76a** was also prepared using the same procedure above, except for the use of H<sub>2</sub>O instead of DMF, for the gel mobility shift assays (Fig. 4). Several chlorido and amino Pt<sup>2+</sup> complexes of 2-hydroxy-3-(aminomethyl)-1,4-naphthoquinone Mannich bases **73a,d,e** exhibiting moderate to high cytotoxicity against cancer cell lines were studied in order to investigate their modes of DNA binding, *in vitro* DNA strand breaks, mechanism of topoisomerase (Topo I) inhibition and cellular accumulation. DNA model base studies have shown that complex **74a** was capable of binding covalently to 9-ethyl-guanine (9-EtG) and 5'-GMP. The chlorido Pt<sup>2+</sup> complexes **74a,d,e** highly accumulate in prostate (PC-3) and melanoma (MDA-MB-435) cell lines, being able to induce DNA strand

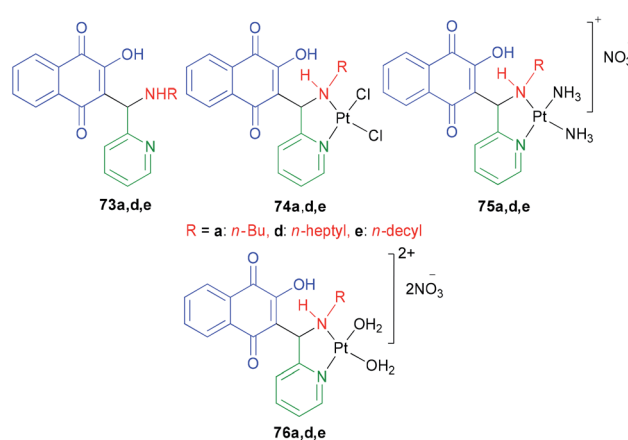
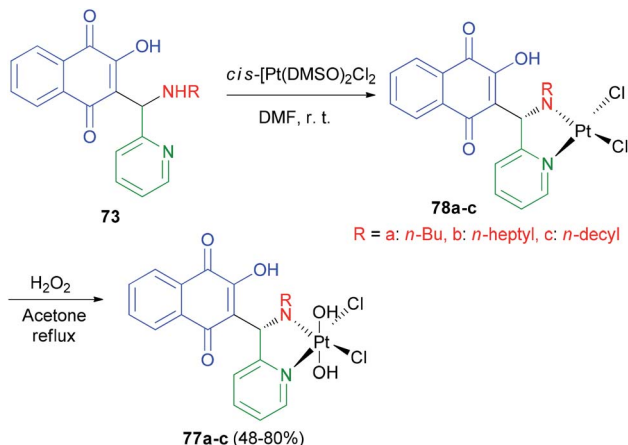
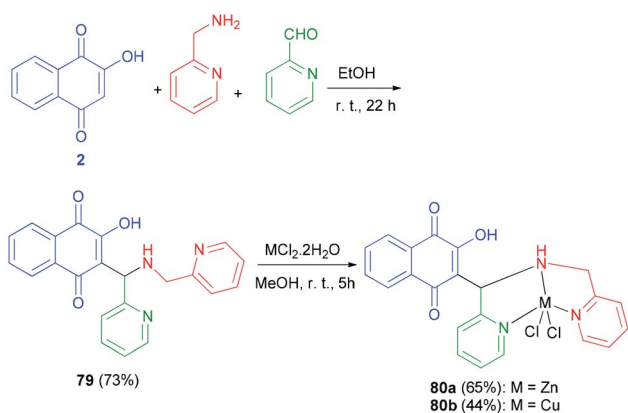


Fig. 4 Structures of 2-hydroxy-3-(aminomethyl)-1,4-naphthoquinones **73** and their Pt<sup>2+</sup> complexes **74**, **75**, **76a,d,e**.





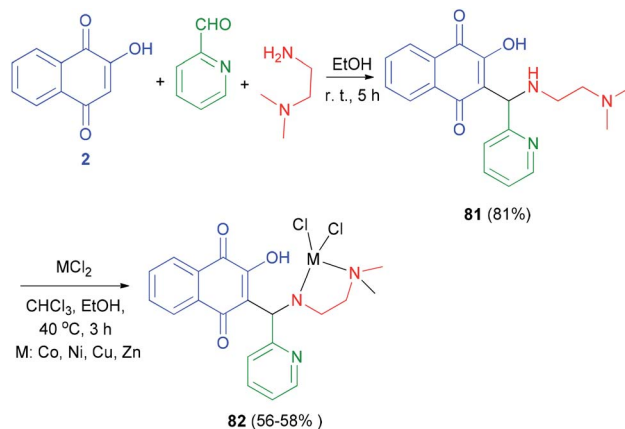
Scheme 41 Synthesis of Mannich base-platinum(IV) complexes 77a-c.



Scheme 42 Synthesis of Mannich base 79 and complexes 80a-b.

breaks *in vitro* and inhibit Topo I by a catalytic mode. On the other hand, the free 2-hydroxy-3-(aminomethyl)-1,4-naphthoquinones **73a,d,e** and the amino Pt<sup>2+</sup> complexes **75a,d,e** neither cause DNA strand breakage nor exhibit strong DNA interaction, nevertheless the latter were also found to be catalytic inhibitors of Topo I at 100 μM. Thus, coordination of the Mannich bases **73a,d,e** to the “PtCl<sub>2</sub>” fragment substantially affects the chemical and biophysical properties of the proligands, leading to an improvement of their DNA binding properties and generating compounds that cleave DNA and catalytically inhibit Topo I. Finally, the high cytotoxicity exhibited by the free (uncomplexed) 2-hydroxy-3-(aminomethyl)-1,4-naphthoquinones **73a,d,e** might be associated with their decomposition in solution, which is not observed for the Pt<sup>2+</sup> complexes.<sup>63</sup>

Vargas *et al.* described synthesis of three novel platinum(IV) complexes *cis,cis,trans*-[Pt(**77a-c**)Cl<sub>2</sub>(OH)<sub>2</sub>] **77a-c** (77 = 2-hydroxy-3-[(R-amino)(pyridin-2-yl)methyl]-1,4-naphthoquinone, R = *n*-butyl, **77a**; *n*-heptyl, **77b** and *n*-decyl, **77c**). In this process, the yellow-orange complexes **77a-c** were obtained by selective oxidation of **78a-c** with H<sub>2</sub>O<sub>2</sub> (30%) in refluxing acetone for 7 h in 48–80% yields (Scheme 41). The cytotoxicity studies against

Scheme 43 Metal(II) complexes of Mannich base derived from lawsone **82**.

four human cancer cell lines have shown that in general the platinum(IV) and platinum(II) derivatives exhibit the same cytotoxic profile and are all more active than cisplatin. The lowest *in vitro* IC<sub>50</sub> values have been observed for **77b-c**, which bear ligands with the largest R groups being the most lipophilic. Furthermore similar IC<sub>50</sub> values for platinum(II) and platinum(IV) complexes of the same ligands have been associated with rapid *in vitro* reduction of the latter complexes to afford **77a-c**.<sup>64</sup>

### 3.4. Other-types metal complexes

Tridentate Mannich base **79** derived from lawsone, pyridine carboxyaldehyde and 2-aminomethylpyridine, have been synthesized in EtOH at room temperature in the dark for 22 h in 73% yield. Dichloro{3-[*N*-(2-pyridylmethyl)aminomethyl-2-pyridyl]-2-hydroxy-1,4-naphthoquinone}zinc(II) **80a** was obtained from the reaction of **79** with 1 equiv. ZnCl<sub>2</sub>·2H<sub>2</sub>O in methanol at room temperature in the dark for 5 h in 65% yield. The analogous dichloro 3-[*N*-(2-pyridylmethyl)aminomethyl-2-pyridyl]-2-hydroxy-1,4-naphthoquinone copper(II) complex **80b** was obtained from equimolar amounts of **79** and CuCl<sub>2</sub>·2H<sub>2</sub>O in methanol at room temperature for 5 h in 44% yield (Scheme 42).<sup>65</sup>

The Mannich base ligand **81** has been synthesized in 81% from lawsone, pyridine-2-carbaldehyde and *N,N*-dimethylethylenediamine in EtOH at room temperature for 5 h. Ligand **81** was dissolved in chloroform and then corresponding metal chloride (Co(II), Cu(II), Ni(II) and Zn(II)) dissolve in ethanol was added drop wise and stirred at 40 °C for 3 h afforded metal(II) complexes of Mannich base **82** in 56–58% yields (Scheme 43).<sup>66</sup>

## 4. Conclusions

Aminomethylnaphthoquinone Mannich bases derived from lawsone and their metal complexes belong to an important class of organic/organometallic compounds and exhibit a wide range of biological properties and due to their potent activities, thus the synthesis of Mannich bases derived from lawsone is an area



of current interest. Several multi-component reactions for the synthesis of aminomethylnaphthoquinones have been reported from lawsone, aldehydes and amines using green solvents, Bronsted and Lewis acid catalysts, ionic liquids, ultrasound irradiation under room temperature, reflux and solvent-free conditions. Also, Metal complexes of 3-(aminomethyl)naphthoquinone Mannich bases have been synthesized from lawsone, aldehydes and organometallic amines or the reaction of Mannich bases with metal cations under different conditions. This review summarizes the most multi-component strategies in the synthesis of aminomethylnaphthoquinone Mannich bases derived from lawsone and their metal complexes and applications.

## Conflicts of interest

There are no conflicts to declare.

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## Notes and references

- 1 Y. Kumagai, Y. Shinkai, T. Miura and A. K. Cho, *Annu. Rev. Pharmacol. Toxicol.*, 2012, **52**, 221–247.
- 2 L. Mafioletti, I. F. Silva Junior, E. M. Colodel, A. Flach and D. T. O. Martins, *J. Ethnopharmacol.*, 2013, **150**, 576–582.
- 3 A. M. Pinto, J. P. Leite, A. P. Neves, G. B. da Silva, M. D. Vargas and I. C. Paixao, *Arch. Virol.*, 2014, **159**, 1827–1833.
- 4 A. C. Dweck, *Int. J. Cosmet. Sci.*, 2002, **24**, 287–302.
- 5 V. A. Motz, C. P. Bowers, L. M. Young and D. H. Kinder, *J. Ethnopharmacol.*, 2012, **143**, 314–318.
- 6 D. C. Mcmillan, S. D. Sarvate, J. E. Oatis Jr and D. J. Jollow, *Toxicol. Sci.*, 2004, **82**, 647–655.
- 7 J. Araldi and S. S. Guterres, *Infarma*, 2005, **17**, 78–83.
- 8 A. E. Musa, B. Madhan, W. Madhulatha, J. R. Rao, G. A. Gasmelseed and S. Sadulla, *J. Am. Leather Chem. Assoc.*, 2009, **104**, 183–190.
- 9 M. Dogru, Z. Baysal and C. Aytekin, *Prep. Biochem. Biotechnol.*, 2006, **36**, 215–221.
- 10 S. Ali, T. Hussain and R. J. Nawaz, *J. Clean. Prod.*, 2009, **17**, 61–66.
- 11 A. Kosmalska, M. Zaborski and A. Masek, *Przem. Chem.*, 2010, **89**, 420–424.
- 12 A. C. Watkinson, K. R. Brian, K. A. Walters and J. Hadgraft, *Int. J. Cosmet. Sci.*, 1992, **14**, 265–275.
- 13 R. R. Korac and K. M. Khambholja, *Pharmacogn. Rev.*, 2011, **5**, 164–173.
- 14 R. Andoong and H. Chihchiang, *Environ. Sci. Technol.*, 2005, **39**, 7460–7468.
- 15 R. A. Doong and Y. J. Lai, *Water Res.*, 2005, **39**, 2309–2318.
- 16 A. Ostovari, S. M. Hoseinie, M. Peikari, S. R. Shadizadeh and S. J. Hashemi, *Corros. Sci.*, 2009, **51**, 1935–1949.
- 17 A. Y. El-Etre, M. Abdallah and Z. E. El-Tantawy, *Corros. Sci.*, 2005, **47**, 385–395.
- 18 Y. M. Hijji, B. Barare and Y. Zhang, *Sens. Actuators, B*, 2012, **169**, 106–112.
- 19 M. N. Rahmoun, M. Benabdallah, D. Villemin, K. Boucherit, B. Mostefa-Kara, C. Ziani-Cherif and N. Choukchou-Braham, *Der Pharma Chem.*, 2010, **2**, 320–326.
- 20 H. S. Muhammad and S. Muhammad, *Afr. J. Biotechnol.*, 2005, **4**, 934–937.
- 21 D. Anju, S. Kavita, G. Jugnu, G. Munish and S. Asha, *J. Pharm. Sci. Innovation*, 2012, **1**, 17–20.
- 22 A. Ahmad, K. Mahal, S. Padhye, F. H. Sarkar, R. Schobert and B. Biersack, *J. Saudi Chem. Soc.*, 2017, **21**, 105–110.
- 23 M. Arundhati, C. Dipak and R. Mithun, *Indian J. Pharm. Educ. Res.*, 2018, **52**, 472–479.
- 24 N. M. F. Lima, C. S. Correia, P. A. L. Ferraz, A. V. Pinto, M. do C. R. F. Pinto, A. E. G. Santana and M. O. F. Goulart, *J. Braz. Chem. Soc.*, 2002, **13**, 822–829.
- 25 A. P. Neves, C. C. Barbosa, S. J. Greco, M. D. Vargas, L. C. Visentin, C. B. Pinheiro, A. S. Mangrich, J. P. Barbosa and G. L. da Costae, *J. Braz. Chem. Soc.*, 2009, **20**, 712–727.
- 26 A. F. L. Vilela, B. M. Frugeri, A. L. F. Sarria, R. O. S. Kitamura, J. B. Fernandes, M. F. G. F. Silva, Q. B. Cass and C. L. Cardoso, *J. Braz. Chem. Soc.*, 2016, **27**, 535–545.
- 27 T. T. Guimaraes, F. R. Maria do Carmo, J. S. Lanza, M. N. Melo, L. Rubens, I. M. de Melo, E. B. Diogo, V. F. Ferreira, C. A. Camara, W. O. Valenca, R. N. de Oliveira, F. Frezard and E. N. da Silva Junior, *Eur. J. Med. Chem.*, 2013, **63**, 523–530.
- 28 C. P. P. de Mello, N. S. Sardoux, L. Terra, L. C. Amorim, M. D. Vargas, G. B. da Silva, H. C. Castro, V. A. Giongo, L. F. Madeira and I. C. N. P. Paixao, *Antiviral Ther.*, 2016, **21**, 507–515.
- 29 J. F. Allochio Filho, L. L. Roldi, M. Delarmelina, R. G. Fiorot, J. T. Andrade, Á. A. Aleixo, R. S. Carvalho, M. G. F. Araújo, J. M. S. Ferreira, A. G. Taranto, W. Romao and S. J. Greco, *J. Braz. Chem. Soc.*, 2016, **27**, 2127–2140.
- 30 K. Mahal, A. Ahmad, F. Schmitt, J. Lockhauserbäumer, K. Starz, R. Pradhan, S. Padhye, F. H. Sarkar, W. S. Koko, R. Schobert, K. Ersfeld and B. Biersack, *Eur. J. Med. Chem.*, 2017, **126**, 421–431.
- 31 M. Dabiri, Z. Noroozi Tisseh and A. Bazgir, *Dyes Pigm.*, 2011, **89**, 63–69.
- 32 M. Leffler and R. Hathaway, *J. Am. Chem. Soc.*, 1948, **70**, 3222–3223.
- 33 C. E. Dalglish, *J. Am. Chem. Soc.*, 1949, **71**, 1697–1702.
- 34 S. C. Hooker, *J. Am. Chem. Soc.*, 1936, **58**, 1190–1197.
- 35 E. N. da Silva Júnior, I. M. M. de Melo, E. B. T. Diogo, V. A. Costa, J. D. de Souza Filho, W. O. Valenca, C. A. Camara, R. N. de Oliveira, A. S. de Araujo, F. S. Emery, M. R. dos Santos, C. A. de Simone, R. F. S. Menna-Barreto and S. L. de Castro, *Eur. J. Med. Chem.*, 2012, **52**, 304–312.
- 36 M. H. C. Lagrota, M. D. Wigg, M. G. M. do Santos, A. V. Pinto and M. C. F. R. Pinto, *Rev. Microbiol.*, 1988, **19**, 338–342.





- 37 A. F. dos Santos, P. A. L. Ferraz, A. V. Pinto, M. C. F. R. Pinto, M. O. F. Goulart and A. E. G. SantAna, *Int. J. Parasitol.*, 2000, **30**, 1199–1202.
- 38 N. B. de Souza, I. M. de Andradel, P. F. Carneiro, G. A. M. Jardim, I. M. M. de Melo, E. N. da Silva Júnior and A. U. Krettli, *Mem. Inst. Oswaldo Cruz*, 2014, **109**, 546–552.
- 39 W. Paengsri, V. S. Lee, W. L. Chong, H. A. Wahab and A. Baramée, *Int. J. Biol. Chem.*, 2012, **6**, 69–88.
- 40 T. Ndior, C. M. Ndiaye, F. D. Faye and A. Samb, *Asian J. Chem.*, 2015, **27**, 4063–4066.
- 41 G. B. da Silva, A. P. Neves, M. D. Vargas, J. D. B. Marinho-Filho and L. V. Costa-Lotufo, *Bioorg. Med. Chem. Lett.*, 2016, **26**, 3537–3543.
- 42 I. Al Nasr, J. Jentzsch, I. Winter, R. Schobert, K. Ersfeld, W. S. Koko, A. A. H. Mujawah, T. A. Khan and B. Biersack, *Arch. Pharm.*, 2019, **352**, 1900128.
- 43 P. Nariya, F. Shukla, H. Vyas, R. Devkar and S. Thakore, *Synth. Commun.*, 2020, **50**, 1724–1735.
- 44 C. Biot, G. Glorian, L. Maciejewski, J. Brocard, O. Domarle, G. Blampain, P. Millet, A. J. Georges, H. Abessolo, D. Dive and J. Lebib, *J. Med. Chem.*, 1997, **40**, 3715–3718.
- 45 M. A. Gouda, H. F. Eldien, M. M. Girges and M. A. Berghot, *Med. Chem.*, 2013, **3**, 2228–2232.
- 46 R. G. Fiorot, A. J. F. Filho, T. M. C. Pereira, V. Lacerda Jr, R. B. dos Santos, W. Romao and S. J. Greco, *Tetrahedron Lett.*, 2014, **55**, 4373–4377.
- 47 J. F. Allochio Filho, R. G. Fiorot, V. Lacerda Jr, R. B. dos Santos, G. Vanini, W. Romao and S. J. Greco, *Colloids Interface Sci. Commun.*, 2015, **4**, 14–18.
- 48 D. Liu, S. Zhou and J. Gao, *Synth. Commun.*, 2014, **44**, 1286–1290.
- 49 M. O. B. Sousa, M. D. Vargas and F. S. Miranda, *J. Mol. Struct.*, 2018, **1164**, 260–272.
- 50 S. Golmakaniyoon, V. R. Askari, K. Abnous, A. Zarghi and R. Ghodsi, *Iran. J. Pharm. Res.*, 2019, **18**, 16–29.
- 51 H. R. Shaterian and M. Mohammadnia, *J. Mol. Liq.*, 2013, **177**, 353–360.
- 52 H. Naeimi and M. Farahnak Zarabi, *Tetrahedron*, 2018, **74**, 2314–2323.
- 53 B. Asadi, I. Mohammadpoor-Baltork, S. Tangestaninejad, M. Moghadam, V. Mirkhani and A. Landarani-Isfahani, *New J. Chem.*, 2016, **40**, 6171–6184.
- 54 M. Ayaz and B. Westermann, *Synlett*, 2010, **10**, 1489–1492.
- 55 H. R. Shaterian and F. Moradi, *Res. Chem. Intermed.*, 2015, **41**, 291–297.
- 56 P. Saluja, J. M. Khurana, K. Nikhil and P. Roy, *RSC Adv.*, 2014, **4**, 34594–34603.
- 57 S. Shaabani, M. R. Naimi-Jamal and A. Maleki, *Dyes Pigm.*, 2015, **122**, 46–49.
- 58 S. Jayashree and K. Shivashankar, *Synth. Commun.*, 2018, **48**, 1805–1815.
- 59 L. I. Lopez-Lopez, S. D. Nery-Flores, A. Saenz-Galindo and D. de Loera, *Synth. Commun.*, 2017, **47**, 2247–2253.
- 60 A. Baramée, A. Coppin, M. Mortuaire, L. Pelinski, S. Tomavo and J. Brocard, *Bioorg. Med. Chem.*, 2006, **14**, 1294–1302.
- 61 A. P. Neves, K. C. B. Maia, M. D. Vargas, L. C. Visentin, A. 3 Casellato, M. A. Novak and A. S. Mangrich, *Polyhedron*, 2010, **29**, 2884–2891.
- 62 A. P. Neves, G. B. da Silva, M. D. Vargas, C. B. Pinheiro, C. Visentin Ldo, J. D. B. M. Filho, A. J. Araujo, L. V. Costa-Lotufo, C. Pessoa and M. O. de Moraes, *Dalton Trans.*, 2010, **39**, 10203–10216.
- 63 A. P. Neves, M. X. G. E. J. Pereira Peterson, R. Kipping, M. D. Vargas, F. P. Silva Jr, J. W. M. Carneiro and N. P. Farrell, *J. Inorg. Biochem.*, 2013, **119**, 54–64.
- 64 G. B. da Silva, A. P. Neves, M. D. Vargas, W. A. Alves, J. D. B. Marinho-Filho, C. Pessoa, M. O. Moraes and L. V. Costa-Lotufo, *J. Braz. Chem. Soc.*, 2013, **24**, 675–684.
- 65 A. P. Neves, M. D. Vargas, C. A. Tellez Soto, J. M. Ramos, C. Visentin Ldo, C. B. Pinheiro, A. S. Mangrich and E. I. P. de Rezende, *Spectrochim. Acta, Part A*, 2012, **94**, 152–163.
- 66 M. Suganthi and K. P. Elango, *J. Biomol. Struct. Dyn.*, 2019, **37**, 1136–1145.

