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Selective extraction and determination of beryllium in real samples using amino-5,8-dihydroxy-1,4-naphthoquinone functionalized magnetic MIL-53 as a novel nanoadsorbent†

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Herein, a magnetic MOF for preconcentration of Be(II) was synthesized. The material is obtained from magnetite (Fe₃O₄) nanoparticles that were modified with 2-amino-5,8-dihydroxy-1,4-naphthoquinone (ADHNQ) and then reacted with terephthalic acid and iron(III) chloride to form a metal–organic framework of the type MIL-53(Fe) capable of extracting Be(II). The extraction parameters were optimized by employing design of experiments methodology. Beryllium concentration was determined by flame atomic absorption spectrometry (FAAS). The detection limit was as low as 0.07 ng mL⁻¹ and the quantification limit is 0.2 ng mL⁻¹. Linearity extends from 0.2 to 100 ng mL⁻¹, and the precision of the method (intra and inter-day assay) is <9.5%. The recoveries of real sample analysis were in the range of 89–101%. The method was successfully applied to the analysis of various real water samples and an alloy sample.

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1 Introduction

Beryllium (Be) as one of the lightest (three times lighter than aluminium) and strongest metals (six times harder than steel) is a strategic metal because of its vital importance and applications in several industries.^{1,2} Pure Be and its alloys with some metals, including Ni, Cu, Mg and Al, have been applied in aerospace components, electronic instrumentation, electrical equipment, the automotive industry, the nuclear industry, jewelry, ceramics, dental amalgams and prostheses.^{1,3} Beryllium and its compounds can cause acute toxicity with the lungs as the target organ in human beings. As early as the 1930s, in beryllium extraction and processing industries, the cases of acute beryllium disease were reported as an inflammation of the respiratory tract.¹ Beryllium exposure takes place through inhalation and ingestion.¹ However, the mechanism of chronic beryllium disease is not thoroughly known. Probably Be combines with special proteins such as ferritin, then enters into the lungs and is taken up by macrophages of the lungs, subsequently causing lesions in the lungs.³ The water supplies and environment are mainly contaminated due to the discharges of related industries and also due to burning coal, while Be exists in the earth crust in the range of ~2.8–5.0 µg g⁻¹.⁴ It has been reported that many serious and fatal problems can occur due to the contamination of Be and its compounds,

below 4 µg L⁻¹ in waters. Based on various national guidelines, the concentration of Be(II) in tap and surface water should not exceed a limit of 0.1 µg L⁻¹ and 0.2 µg L⁻¹, respectively.⁵ Generally, in natural water, the average concentration of Be(II) is in the range of ng L⁻¹.⁶ As International Agency for Research on Cancer (IARC) states, Be and its compounds are classified as group 1 due to their carcinogenic effect to human and animal.⁷ Thus, it is necessary to develop and apply highly accurate and sensitive methods for determination of Be(II) at very low concentrations. The techniques applied for Be(II) determination are as follows: atomic absorption spectrometry (AAS),⁶ spectrofluorimetry,⁸ spectrophotometric method,¹⁶ inductively coupled plasma mass spectrometry,⁹ inductively coupled plasma optical emission spectrometric detection¹⁰ and gas chromatography.¹¹

Solid phase extraction (SPE) has been considered as one of the principal sample preparation methods due to its merits, including easy configuration, high enrichment factors and recoveries, low cost and facile on-line hyphenation.⁵ Magnetic solid-phase extraction (MSPE) is a type of SPE in dispersion mode utilizing magnetic adsorbents which has triggered interest in sample pretreatment strategies due to its unique magnetic behavior, facile implementation and low cost. It is necessary to consider the nature, features and behavior of the magnetic adsorbent in the MSPE adsorbent development.^{12,13} Thus, diverse adsorbent have been introduced for the extraction of various analytes from miscellaneous matrices such as Fe₃-O₄@metal–organic frameworks (MOFs) composites, Fe₃O₄@covalent organic frameworks (COFs), magnetic molecularly imprinted and magnetic carbon materials.¹³

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MOFs are classified as multifunctional microporous materials formed from metal ions and organic ligands through coordinate bonds. Tailorable chemistry, large specific surface areas, tunable cavities and good stability are the special features of MOFs, which has made this new type of scaffolds suitable for the incorporation of magnetic nanoparticles, providing versatile and efficient MSPE adsorbents.¹³ Some adsorbents that provide selective separation and concentration of trace amounts of Be in natural water resources are silica gel,¹⁴ activated carbon,¹⁵ anion exchange resin,¹⁶ salicylate chelating resin¹⁷ and mixed micelle mediated extraction.¹⁸

Herein, a magnetic MOF for preconcentration of Be(II) was synthesized. The material is obtained from magnetite (Fe₃O₄) nanoparticles that were modified with 2-amino-5,8-dihydroxy-1,4-naphthoquinone (ADHNQ) and then reacted with terephthalic acid and iron(III) chloride to form the metal-organic framework of the type MIL-53(Fe) capable of extracting Be(II).

2 Experimental

2.1 Reagents and instrumentation

Benzen-1,4-dicarboxylic acid (H₂BDC), (3-chloropropyl)-triethoxysilane (3-CPTS), 2-amino-5,8-dihydroxy-1,4-naphthoquinone (ADHNQ), toluene, triethylamine, ammonium hydroxide (28% w/v), tetraethyl orthosilicate (TEOS), methanol, ethanol, *N,N*-dimethylformamide (DMF), acetone, FeCl₃, (NH₄)₂-Fe(SO₄)₂·6H₂O, NaOH, HCl, H₂SO₄ and HNO₃, were supplied by Merck (Darmstadt, Germany) or Fluka (Seelze, Germany). Stock solution (1000 mg L⁻¹) of Be(II) was purchased from Merck. All the other solutions were daily prepared using distilled water.

2.2 Instrumentation

Be(II) determination was performed using a FAAS instrument model AA-680 Shimadzu (Japan) consisting an acetylene/nitrous oxide flame and a Be hollow cathode lamp with wavelength of 234.9 nm. Manufacturer's manual was used to set up the instrument. A digital 827 WTW Metrohm pH-meter composed of a combined glass-calomel electrode (Herisau, Switzerland) was employed for the pH measurements. Fourier transform infra-red (FT-IR) analysis carried out on a Bruker spectrophotometer model IFS-66. Elemental analysis of the nanoadsorbent was conducted using an EA112 flash elemental analyzer (Thermo Finnigan, Okehampton, UK). Scanning electron microscopy (SEM) study was performed on a KYKY-3200 instrument (Beijing, China). Magnetic features of the nanoadsorbents were recorded on a vibrating sample magnetometer (VSM) (AGFM/VSM 117 3886 Kashan, Iran). A Bahr STA-503 instrument (Behrthermo, Germany) was used for thermogravimetric analysis (TGA). Surface analysis was explored by nitrogen adsorption-desorption method employing a Micromeritics ASAP 2010 instrument. X-ray diffraction (XRD) analysis was conducted on a Philips-PW 12C diffractometer (Amsterdam, The Netherlands) using Cu K α radiation.

2.3 Synthesis of the magnetic nanoadsorbent

MNPs@SiO₂ nanoparticles (NPs) was prepared according to the literatures¹⁹⁻²¹ and the detailed information were provided in

ESI.† For functionalization, in brief 1.0 g silica-coated magnetite NPs (Fe₃O₄@SiO₂) was added to 75 mL dried toluene, and then 1.0 g 3-CPTES was added to the mixture and the reaction continued at 130 °C for 18 h under the reflux condition. Thereafter, the fabricated solid product was magnetically isolated from the reaction medium and it was washed with 50 mL ethanol three times and then dried at vacuum oven at 50 °C (Fig. 1a). For fabricating ADHNQ-functionalized MNPs, 1.0 g Fe₃O₄@SiO₂@3-CPTES was suspended in 150 mL 1:1 v/v methanol-TEA and then 0.25 g ADHNQ was added to the mixture. Afterwards, the mixture was heated under the reflux condition for 24 h (Fig. 1a). Ultimately, the final product was magnetically separated and washed with methanol and acetone several times and then dried in vacuum oven at 50 °C.

Fe₃O₄@SiO₂@ADHNQ@MIL-53(Fe) was fabricated *via* solvothermal method.²² Briefly, 3.0 mmol H₂BDC was dissolved in 10 mL DMF containing 0.25 g Fe₃O₄@SiO₂@ADHNQ and the mixture was sonicated for 20 min. Thereafter, 3.0 mmol ferric chloride was dissolved in another vial that contains 5.0 mL DMF *via* sonication. Afterwards, the two mixture were mixed in a 50 mL Teflon lined stainless-steel autoclave, and the reaction was performed at 150 °C during 2 h. After the completion of the reaction, the temperature of mixture was allowed to decrease down to the room temperature. The Fe₃O₄@SiO₂@ADHNQ/MIL-53(Fe) was magnetically separated and washed with 3 × 30 mL deionized (DI) water and then 3 × 25 mL ethanol, and finally it was dried. The synthesis of Fe₃O₄@SiO₂@ADHNQ/MIL-53(Fe) was confirmed by FT-IR spectroscopy, SEM, elemental analysis, XRD, and VSM techniques. The same process was applied for the synthesis of MIL-53(Fe) in the absence of MNPs.

2.4 Extraction procedure

In brief, pH of samples were adjusted at 6.4 utilizing 0.25 mol L⁻¹ NH₃ and HCl solutions. Thereafter, 32 mg Fe₃O₄@SiO₂@ADHNQ/MIL-53(Fe) nanoadsorbent was dispersed to the sample solution and it was stirred for 6.5 minutes. Afterwards, the nanosorbent was isolated magnetically and the supernatant solution was subjected to FAAS to determine the sorption%. For elution, the nanoadsorbent was dispersed in 1.0 mL of a solution containing 0.6 mol L⁻¹ HNO₃ and the resultant mixture was stirred for 13.5 minutes. Finally, the mixture was separated using magnet and the supernatant was subjected to FAAS to determine recovery%.

2.5 Real sample and certified reference material preparation

The new method was utilized to a certified reference sample (NIST-SRM 1640), Cu-Be alloy and three different water samples (seawater river water and well water). The seawater (Caspian Sea, North of Iran), river water (Karaj, Iran) and well water (Karaj, Iran) samples were filtered into cleaned polyethylene bottles and their pH was adjusted to 6.4 before the extraction procedure. Acid digestion procedure was conducted to digest Cu-Be alloy sample. In this regard, 0.4 g of the sample was transferred into a 50 mL beaker and 5 mL 50% v/v HNO₃ was added to dissolve the alloy and finally it was diluted up to 100 mL using a volumetric flask and the pH of resulting solution was adjusted to 6.4 before extraction process.



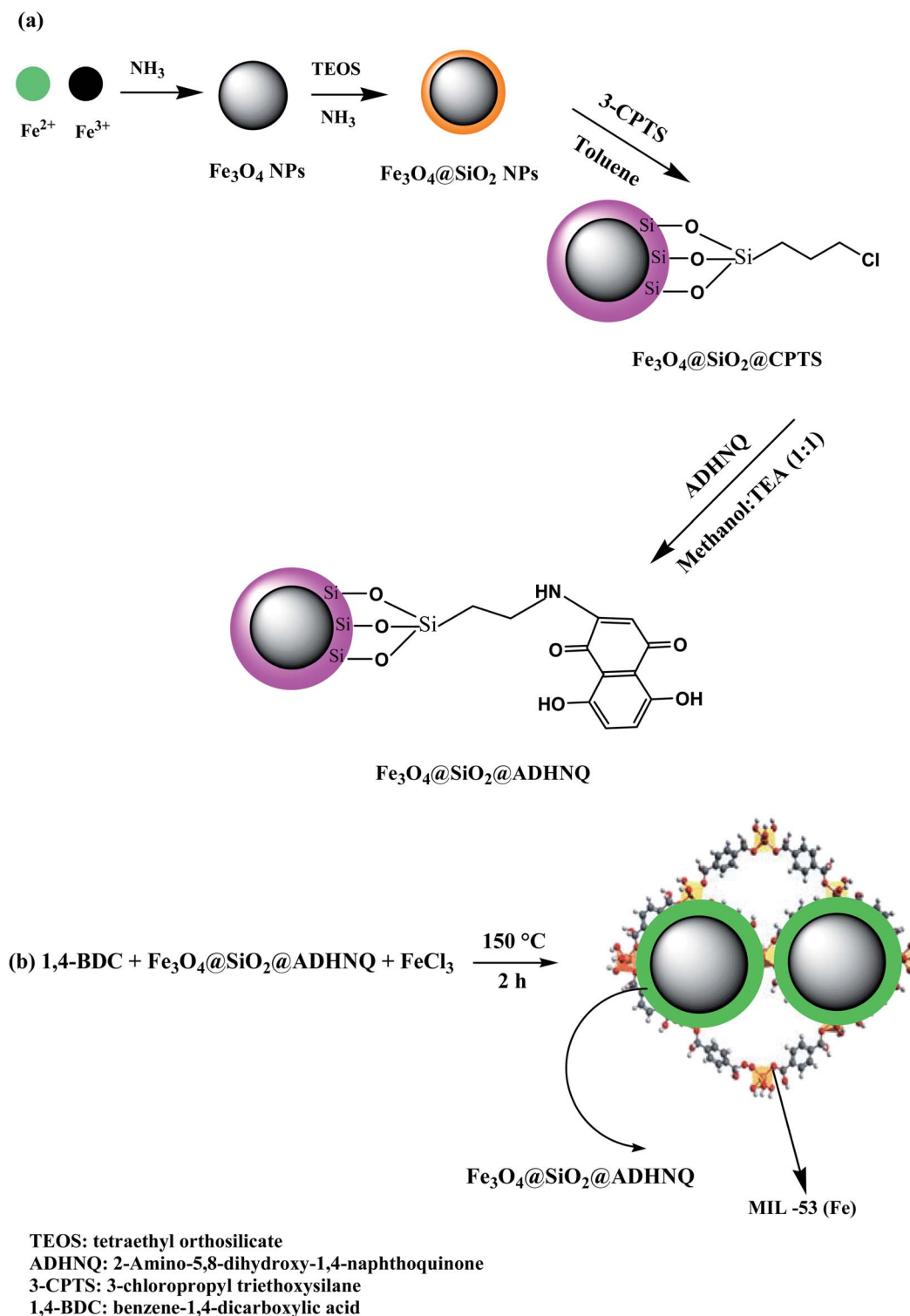


Fig. 1 (a) A scheme for the synthesis of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ and (b) $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ nanocomposite.

3 Results and discussion

3.1 Characterization of nanoadsorbent

The fabricated $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ sorbent at first was characterized by FT-IR spectroscopy (Fig. 2). The absorption bands at 583 cm^{-1} , 1014 cm^{-1} , 1168 cm^{-1} ,

3370 cm^{-1} , 1389 cm^{-1} , 1503 cm^{-1} , 1598 cm^{-1} , and 1669 cm^{-1} were attributed to (Fe-O), (Si-O-Si), (C-N), (O-H), (C-OH), (C-O), (C=C) and (C=O), respectively, confirmed the presence of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ in MIL-53(Fe) structure and also proved the synthesis of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$. Elemental analysis of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ was conducted to



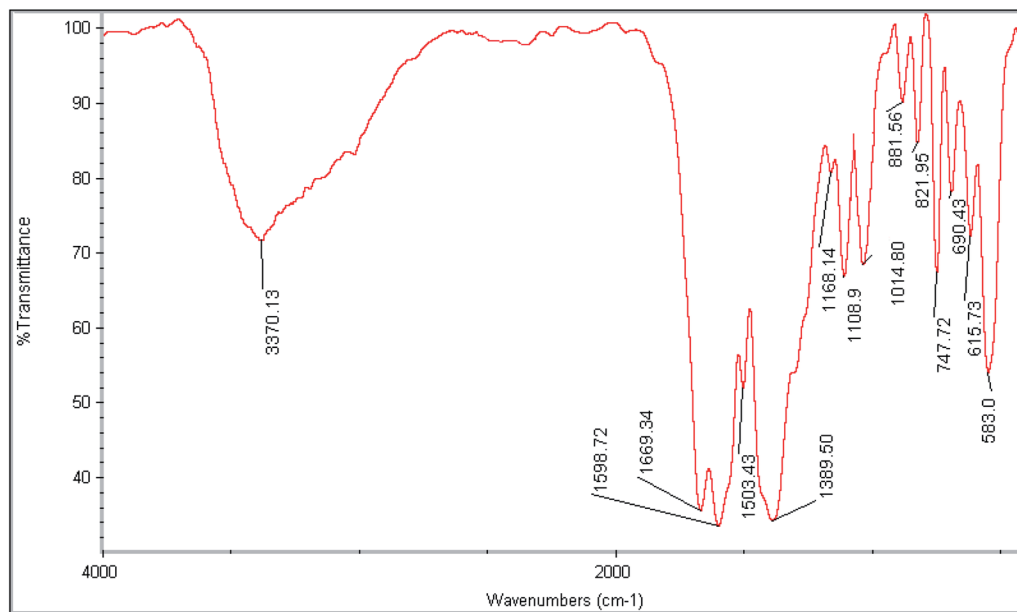


Fig. 2 FT-IR spectrum of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ nanoadsorbent.

determine the elements of nanocomposite. The results revealed 37.5% C, 1.7% H, and 3.5% N in the structure of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ and confirmed the successful synthesis of nanoadsorbent.

To explore the surface morphology of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ and $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$, SEM analysis was conducted. As depicted in Fig. 3a, $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ NPs are spherical with a mean diameter of about 50 nm. After functionalization, as can be seen in Fig. 3b, an increase in the particle size of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ was observed (200 nm) which can be attributed to the coating of these particles with a MIL-53(Fe) layer.

The magnetic features of magnetite, $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ and $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ were explored at ambient temperature with a VSM system and the obtained curves are exhibited in Fig. 4. The saturation magnetization for magnetite, $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ and $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ was 65, 35.6 and 19.7 emu g^{-1} , respectively. Thus, this nanoadsorbent can be easily separated from the solution during DMSPE process with an external magnet due to their superparamagnetic and high saturation magnetization.²²

The surface of the as-fabricated and reused $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ were analyzed using BET N_2 adsorption-desorption method to obtain the specific surface area and pore volume of this sorbents. Accordingly, the surface and pore volume of as-fabricated $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ was 21 $\text{m}^2 \text{g}^{-1}$ and 0.022 $\text{cm}^3 \text{g}^{-1}$, respectively. The surface and pore volume of five times used $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ was 18 $\text{m}^2 \text{g}^{-1}$ and 0.016 $\text{cm}^3 \text{g}^{-1}$, respectively. The decrease in the surface and pore volume of the nanoadsorbent can be related to the possible damage of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ after reusing.

The structure of as-fabricated and reused $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ were analyzed by XRD technique (Fig. 5).

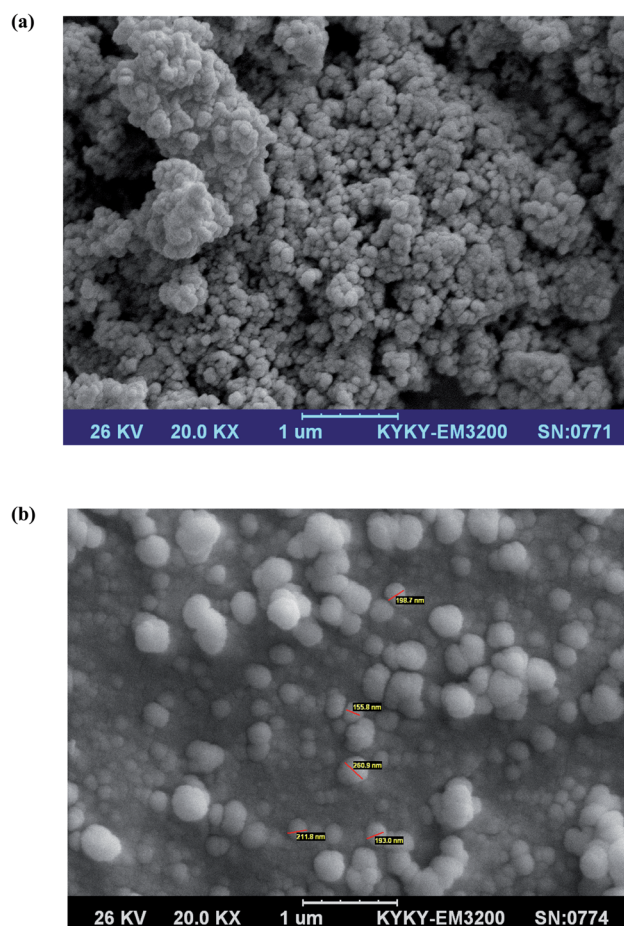


Fig. 3 SEM image of (a) $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ NPs and (b) $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$.



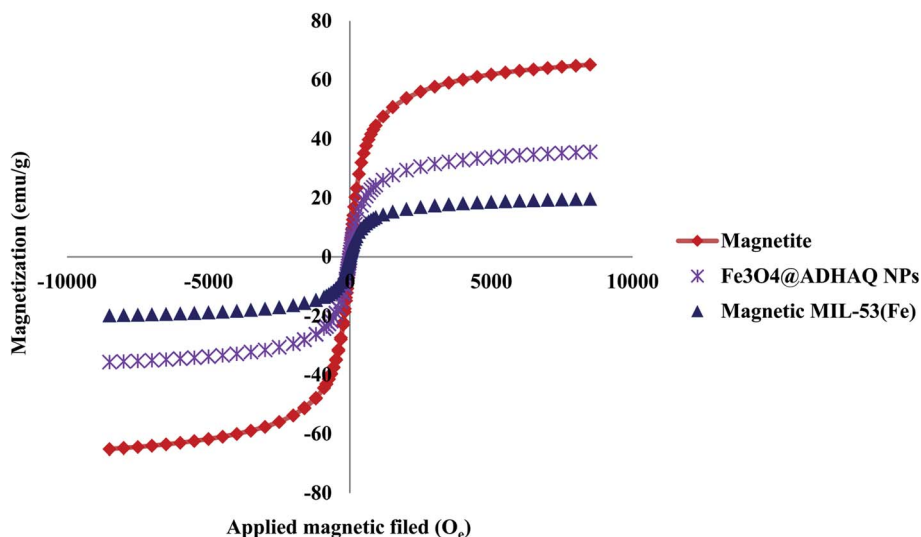


Fig. 4 VSM curves of the magnetite, $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ$ NPs and $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ nanocomposite.

The diffraction peaks at $2\theta = 9.2^\circ, 12.6^\circ, 17.6^\circ, 18.9^\circ,$ and 25.4° are related to the MIL-53(Fe) backbone.²² The diffraction peaks at $2\theta = 30.3^\circ, 35.6^\circ, 57.3^\circ,$ and 62.0° are corresponded to MNPs and confirm the structure of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$. The same XRD pattern was observed for 5-times used nano-adsorbent with a slight decrease in the peak's intensity which can be related to the partial damage of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ after reusing.

The thermal behavior of as-fabricated and reused $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ were explored by TGA technique (Fig. 6). The as-fabricated and reused nano-adsorbents exhibited the same thermal behavior. A sharp losing weight between 270°C and 420°C was observed and the total weight loss up to 800°C for as-fabricated and reused $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ was 86.5% and 91.5%, respectively. This difference can be attributed to the partial damage of $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ after reusing.

3.2 Optimization

The experimental design methodology enables the analyst to select the best optimum conditions with minimum numbers of

experiment, minimum of time and cost consumption. In order to achieve the best extraction performance, the main parameters affecting the extraction of the suggested DMSPE method were evaluated and optimized through the central composite design (CCD). CCD is composed of a two level factorial design (2^f), star points ($2f$) and center points that are located at the center of the experimental region (C_0). Center points are performed to establish the experimental error and the star points are located at $\pm\alpha$ from the center of experimental domain and provide wider extremes for the low and high settings of factors.²³ Response surface methodology (RSM) has been extensively employed for the effective modeling and specify of main factors involved the DMSPE. Desirability function is the predominant and most approach applied to access the best agreement between the two responses based on the mathematical models constructed in response surface methodology. In this regards, the effect of following variable were explored and optimized: (a) pH; (b) $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ dosage; (c) adsorption time; (d) eluent type; (e) eluent concentration; (f) eluent volume; (g) elution time. Based on the number of experiments for a CCD assay ($N = 2^f + 2f + C_0$, N : number of experiments; f : number of variables and C_0 : number of

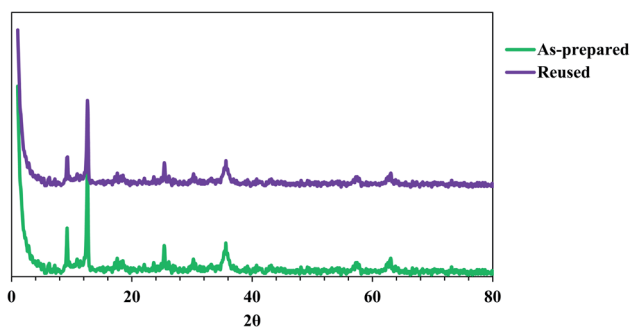


Fig. 5 XRD patterns of as-prepared $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ and reused $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$.

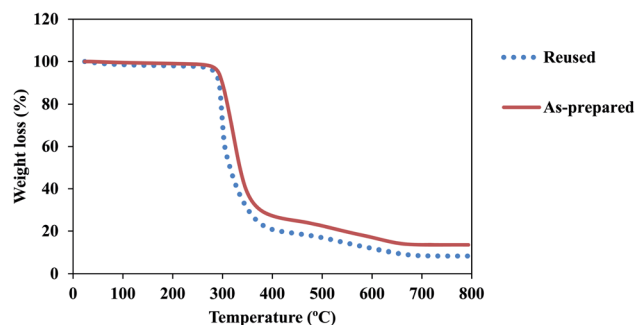


Fig. 6 TGA curves of as-prepared $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$ and reused $\text{Fe}_3\text{O}_4@SiO_2@ADHNQ/MIL-53(Fe)$.



replicates in center point), f and C_0 were set at 3 and 5, respectively and hence 19 tests were designed for each step (sorption and elution). Two replicates were conducted for each test and the average was considered as a suitable response. The results of the experimental design for each step along with their corresponding response were presented in Fig. 7 and 8, Table 1S.† The points with high desirability were depicted with red color and low desirability with green colors. From evaluating the result, the best extraction conditions were: (a) optimum pH: 6.4; (b) $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ dosage: 32 mg; (c) adsorption time: 6.5 min; (d) eluent type: HNO_3 ; (e) eluent concentration: 0.6 mol L^{-1} ; (f) eluent volume: 1.0 mL; (g) desorption time: 13.5 min.

3.3 Effect of interfering

The effect of potentially interfering ions on the recoveries of the $\text{Be}(\text{II})$ ions was also investigated. The obtained results were summarized in Table 2S.† The coexisting ions did not show

considerable effect on the recovery of $\text{Be}(\text{II})$. The results revealed that the nanoadsorbent possessed high selectivity toward $\text{Be}(\text{II})$ ions in the presence of other interfering ions. The presence hydroxyl moieties on the surface of nanoadsorbent impart the selective interaction to $\text{Be}(\text{II})$ ions. According to the hard and soft acid–base theory, hard acids such as $\text{Be}(\text{II})$ ions react more efficiently with the ligands consisting oxygen moieties (such as naphthoquinone) as the hard bases.¹⁸

3.4 Sorption capacity and effect of sample volume

Under the optimized condition, sorption capacity of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ towards $\text{Be}(\text{II})$ was evaluated. To this aim, the adsorption capacity of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ toward $\text{Be}(\text{II})$ was explored utilizing 100 mL standard solution and calculated by using equation:

$$SC = \frac{(C_0 - C_e)V}{q}$$

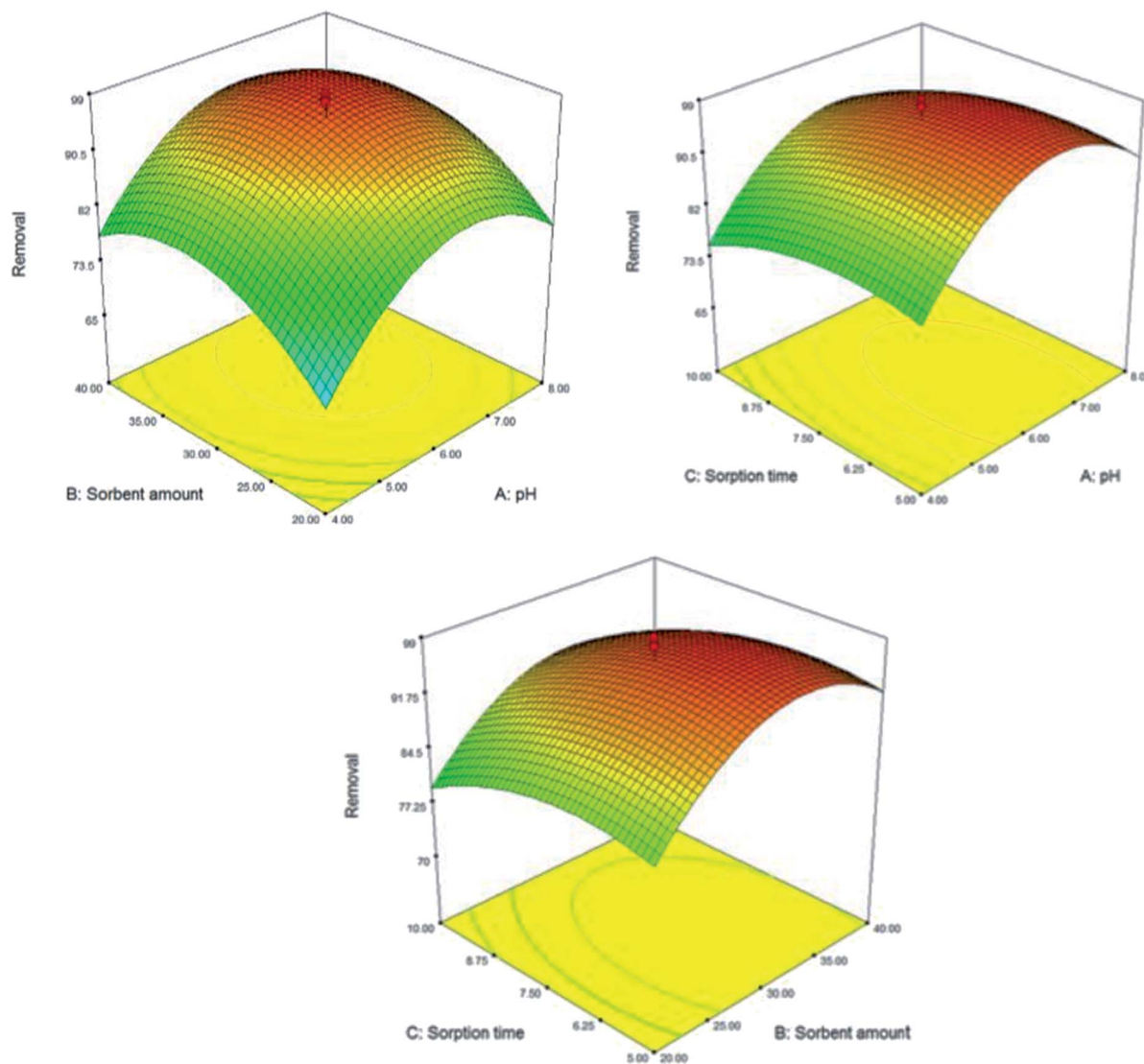


Fig. 7 3D response surface plots of the sorption step.



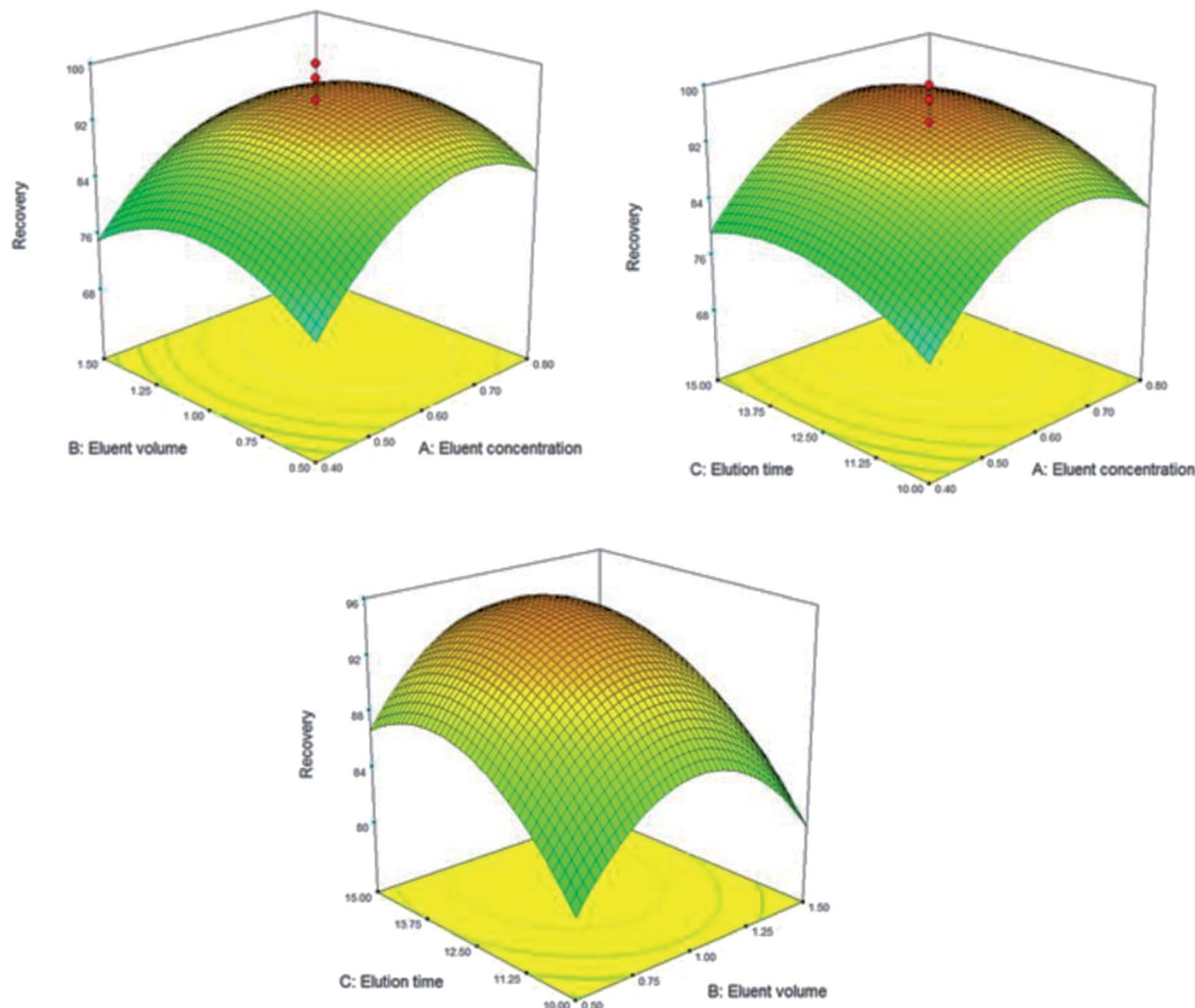


Fig. 8 3D response surface plots of the elution step.

where, SC represents the sorption capacity per gram of nano-adsorbent, C_0 and C_e are the initial and equilibrium concentration of Be(II) (mg L^{-1}), q is the weight of nanoadsorbent (mg) and V is the volume of Be(II) solution (L). Based on the achieved result, the maximum sorption capacity was calculated as 284 mg g^{-1} . This high SC can be attributed to the porous structure of $\text{Fe}_3\text{O}_4@/\text{SiO}_2@/\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ increasing its adsorption capacity.

The effect of the volume of sample on the recovery of Be(II) ions was also explored in the range of 50–500 mL. The results illustrated that the increase in the sample volume up to 350 mL

leads to an enhancement in the recovery, while for the sample volumes more higher 350 mL the recovery decreases. Hence, $\text{Fe}_3\text{O}_4@/\text{SiO}_2@/\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ nanoadsorbent provides a preconcentration factor of 350 regarding to the sample (350 mL) and eluent (1 mL) volumes.

3.5 Analytical features

Under the optimized conditions, analytical features of the DMSPE-combined FAAS method such as limit of detection, dynamic linear range, accuracy and precision were evaluated. An excellent linear relationship between the absorption and

Table 1 Determination of beryllium in certified reference material

| CRM | Concentration | | Recovery (%) | Error (%) | |
|---------------|---------------|--------------------------------|----------------|-----------|-------|
| | Analyte | Certified | | | Found |
| NIST SRM 1640 | Beryllium | 34.94 ($\mu\text{g L}^{-1}$) | 34.3 \pm 2.8 | 98.2 | −1.8 |



Table 2 Determination of beryllium in different real samples (mean \pm SD)^a

| Sample | Real value | Added value | Found value ^b | Recovery (%) |
|-------------------|------------------------------|-------------|------------------------------|--------------|
| Well water | N.D. ^d | 5.0 | 4.9 \pm 0.32 ^c | 98 |
| Caspian sea water | 1.7 \pm 0.05 ^c | 5.0 | 6.75 \pm 0.48 ^c | 101 |
| River water | N.D. ^d | 5.0 | 4.45 \pm 0.10 ^c | 89 |
| Cu–Be alloy | 0.64 ^b \pm 0.04 | 1.0 | 1.58 \pm 0.10 | 94 |

^a Standard deviation ($n = 3$). ^b w/w%. ^c For water samples concentration are based on $\mu\text{g L}^{-1}$. ^d N.D.: not detected.

Be(II) concentration was achieved from 0.2–100 ng mL^{-1} with r^2 value equal to 0.9965. The limit of detection (LOD) was calculated to be 0.07 ng mL^{-1} based on $3S_b/m$, where S_b is the standard deviation of 10 replicates blank signal, m = slop of calibration plot. The obtained LOD is much lower than the maximum permitted level of Be(II) in water. The relative standard deviation (RSD) for 5 replicate extraction and determinations of 30 ng mL^{-1} Be(II) was investigated and found to be lower than 5.8, indicating the good precision of the proposed method. In the next step, the accuracy of the developed method was explored by estimation Be(II) level in NIST SRM 1640 certified reference material. The obtained amount was in well agreement with the certified value and the results confirm the excellent accuracy of the method, as summarized in Table 1.

3.6 Analytical applicability

In the next step, the analytical applicability of the DMSPE combined with FAAS method was evaluated for extraction and determination of Be(II) in seawater, well water, river water and Cu–Be alloy samples. The relative recoveries and standard deviations were presented in Table 2. As summarized in Table 2, the recoveries for the Be(II) were ranged from 89 to 101% with standard deviation lower than 10% for three replicate tests. These result indicated the method could be very promising for determination of Be(II) in real samples without significant matrix effects.

4 Conclusions

Herein, for the first time $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ nanoadsorbent was successfully fabricated and then was DMSPE adsorbent for the extraction and preconcentration of Be(II) at trace levels. $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ exhibited high selectivity to Be(II) due to its functionalization with ADHNQ molecules. Moreover, the porous structure of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{ADHNQ}/\text{MIL-53}(\text{Fe})$ increases its adsorption capacity. The proposed DMSPE method is quick, easy, cost-effective, and has low detection limit (0.07 ng mL^{-1}) and excellent dynamic linear range (0.2–100 ng mL^{-1}) compared to other reports (Table 3S, ESI[†]). A high maximum sorption capacity equal to 284 mg g^{-1} was obtained. The accuracy of the developed method was explored by estimation Be(II) level in NIST SRM 1640 certified reference material. The obtained amount ($34.3 \pm 2.8 \mu\text{g L}^{-1}$) was in well agreement with the certified value ($34.94 \mu\text{g L}^{-1}$) and the results confirm the

excellent accuracy of the method. Ultimately, to explore the feasibility of the method, it was applied for quick extraction and determination of Be(II) in three different water samples and an Cu–Be alloy and satisfactory result was obtained. These results proved that the DMSPE combined FAAS can be considered as an alternative adsorbent for extraction and determination of Be(II) from real samples.

Conflicts of interest

The authors declare that they have no conflict of interest.

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