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Synthesis of oxygen-doped graphitic carbon nitride and its application for the degradation of organic pollutants *via* dark Fenton-like reactions†

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Graphitic carbon nitride (g-C₃N₄) is a promising photocatalyst for environmental protection but its development is greatly limited for its application in dark Fenton-like reactions due to its extremely low specific surface area and lack of suitable active sites. Herein, for the first time, graphitic carbon nitride with large surface area and abundant defect sites was developed by tailoring oxygen *via* a simple and green method without any templates, namely, the calcination–hydrothermal–calcination successive treatment of melamine. The structure of the catalyst was characterized using several technologies, including XRD, SEM, TEM, N₂-physisorption, FT-IR, Raman spectroscopy and XPS. The results revealed that it possessed a large specific surface area (ca. 236 m² g^{−1}), while changes in its structural properties such as the formation of new defect sites and change in the content of nitrogen atoms were observed. These properties were beneficial for the *in situ* activation of H₂O₂ toward reactive oxygen species, as confirmed by the reactive oxygen species capturing experiments. Furthermore, various influencing factors were systemically investigated. The results clearly showed that the oxygen-doped g-C₃N₄ was light-independent and metal-free Fenton-like catalyst for the enhanced degradation of organic pollutants in wastewater. Compared to the pristine g-C₃N₄, the oxygen-doped g-C₃N₄ showed superior performance under various conditions such as broad pH range and excellent stability. Thus, this study provides a novel pathway for the treatment of organic pollutants in water.

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1. Introduction

Organic compounds are widely used in many applications of manufactured products, such as plastics, paper, and textiles.¹ However, some dye pollutants, such as rhodamine B (RhB) and methylene blue (MB), are not easily degraded in aqueous solution, and thus severely threaten the environment. Therefore, the development of an effective, inexpensive, and benign technology for dealing with these pollutants is an important issue. Several oxidation processes such as Fenton-like reactions have been employed to remove organic compounds from wastewater.^{2–5} However, there are still many problems, for instance,

the abundant Fe- and Cu-based catalysts employed exhibit strong toxicity and limited pH range functionality. Thus, there is a considerable urgency to explore green Fenton-like catalysts with both non-metal participants and strong resistance to a broad range of reaction conditions. Accordingly, g-C₃N₄ appears to be an ideal candidate material, having many merits, including safety, low cost, and good stability. Currently, this compound exhibits catalytic properties for the degradation of various organic pollutants, which are also strongly dependent on light illumination.^{1,6} Cui *et al.*¹ first reported a g-C₃N₄ catalyst activated by H₂O₂ to degrade RhB in solution under visible light illumination. However, it hardly exhibited any catalytic activity in the dark, which apparently hindered large-scale development in industry. Thus, to solve this problem, much effort has been devoted to achieving excellent performances in the dark for better practical use.^{7–10} For example, Ge *et al.*⁷ prepared a MgO/g-C₃N₄ catalyst by employing high-temperature calcination. Zhu *et al.*⁸ synthesized Cu-g-C₃N₄ composites *via* the pyrolysis of [H₂mela]₂[CuCl₅]Cl, and Ding *et al.*⁹ developed the g-C₃N₄-supported iron oxide composite. Their results demonstrated that these catalysts were capable of activating H₂O₂ to degrade organic compounds in the dark. However, these catalysts still contained metal elements, and their photocatalytic performance and stability need to be improved. Consequently, the

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development of effective, stable, and non-metal g-C₃N₄-based catalysts remains pursued with the dark Fenton-like reactions.

Oxygen (O)-doped g-C₃N₄ catalysts have been recognized as promising due to their non-toxicity, abundance, and low cost, and thus have been widely employed in several fields,^{11–17} including pollutant removal,¹¹ bioimaging probes,¹³ and singlet O generation.¹⁴ However, to the best of our knowledge, there is no report on the application of O-doped g-C₃N₄ in dark Fenton-like reactions to date. In addition, in the past few years, it has been reported that the synthesis of O-doped g-C₃N₄ can be carried out in a liquid environment,^{11–17} mainly including hydrothermal treatment by H₂O₂,¹¹ HNO₃,¹² H₂SO₄/KMnO₄,¹³ and HNO₃/H₂SO₄.¹⁴ Clearly, many toxic chemical reactants are required, which can lead to serious environmental pollution. Also, in most cases, the specific surface area of the catalyst is not greater than 100 m² g^{−1}, and thus not adequate for generating the desired activity. Consequently, the further development of O-doped g-C₃N₄ is of great interest and necessary. Creating defects in a catalyst is a useful way to regulate its structure and catalytic properties. This can increase the specific surface area of the catalyst *via* the introduction pore structure and/or altered morphology.^{18,19} Importantly, it may directly activate oxidants *in situ* to generate reactive oxygen species (ROS), which are usually observed with metal-based catalysts.^{20,21} For example, Liu *et al.*²⁰ demonstrated that the vacancy sites on nickel(II) oxide were crucial for activating peroxydisulfate through a non-radical pathway. Gao *et al.*²¹ reported that the O defect sites on Fe₂O₃/CeO₂ could activate H₂O₂ *in situ* to produce ·OH species under an O₂ atmosphere. However, the study of non-metal-based catalysts with defect sites in Fenton-like reactions has rarely been reported. As mentioned above, increasing the specific surface area and introducing defects in a catalyst are two effective strategies for strengthening its performance for the degradation of organic pollutants. It will be beneficial for boosting the catalytic performance in dark Fenton-like reactions if the above targets are concurrently achieved in one g-C₃N₄-based material, and thus organic pollutants can be effectively degraded under light-independent conditions. To date, there are no reports on g-C₃N₄-based materials with a large specific surface area and many defects to allow dark Fenton-like reactions. Based on the above studies and problems associated with both the synthesis and application of O-doped g-C₃N₄, it is highly desirable to explore new methods for the synthesis of O-doped g-C₃N₄ possessing a large surface area and abundant defect sites to meet these requirements. With the aim of introducing O through the hydrothermal treatment of g-C₃N₄ and defect generation *via* thermal treatment, this study explores the use of combined hydrothermal and calcination technologies to increase the specific surface area and introduce defect sites in g-C₃N₄ without the addition of templates. Accordingly, to the best of our knowledge, this practical approach has not been reported to date.

In this study, O-doped g-C₃N₄ catalysts with high specific surface areas and many defect sites were synthesized in a multistep post-formation treatment for the degradation of organic pollutants in dark Fenton-like reactions. These catalysts showed superior performances under the optimal conditions.

In addition, the adsorption and active catalytic sites and ROS were identified and examined based on the degradation efficiencies for organic pollutants in water.

2. Experimental

2.1. Materials and reagents

All chemical reagents were of analytical grade and used as received without further purification. Deionized water (resistivity 18.2 M) was filtered using a Millipore Milli-Q water purification system.

2.2. Preparation of g-C₃N₄-based catalysts

The target catalyst was prepared *via* a three-step successive treatment from melamine (Scheme 1). Briefly, 10 g of melamine powder was placed in a 50 mL crucible with an aluminum foil cover and then heated to 550 °C for 4 h at a rate of 5 °C min^{−1} under a static air atmosphere. The obtained yellow sample was denoted as g-C₃N₄. 0.3 g of the obtained powder was added to 25 mL of deionized water, transferred to a 50 mL Teflon-sealed autoclave, and heated at 180 °C for 4 h. After cooling to room temperature, the sample was washed and dried at 120 °C, which was denoted as g-C₃N₄(HY), and then the sample was calcined at 550 °C for 4 h at a rate of 5 °C min^{−1} under a static air atmosphere. The resultant sample was then denoted as g-C₃N₄(HY-HT). Similarly, g-C₃N₄ was again calcinated at 550 °C for 4 h at a rate of 5 °C min^{−1} under a static air atmosphere. The resultant sample was denoted as g-C₃N₄(HT). 0.3 g of g-C₃N₄(HT) powder was then hydrothermally treated under the same conditions as that for g-C₃N₄(HY) and denoted as g-C₃N₄(HT-HY).

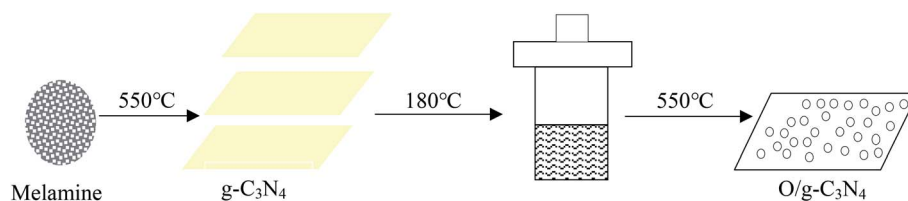
2.3. Characterization

X-ray diffraction (XRD) spectroscopy was carried out using a D8-Advance X-ray diffractometer (Bruker Corp., Billerica, MA, USA); N₂-physisorption was conducted on an Autosorb-1 instrument (Quantachrome Instruments Corp., Boynton Beach, FL, USA), at liquid-N₂ temperature; scanning electron microscopy (SEM) was carried out on a JSM 6700 F instrument (JEOL Ltd., Tokyo, Japan), operating at an accelerating voltage of 10 kV; transmission electron microscopy (TEM) was conducted on a TALOS F200 X instrument (Thermo Fisher Scientific Inc., Waltham, MA, USA), operating at an accelerating voltage of 200 kV; Fourier transform-infrared spectroscopy (FT-IR) was recorded on a Varian 3100 spectrometer (Varian, Inc., Palo Alto, CA, USA), equipped with a DTGS detector; Raman spectroscopy was recorded on an InviaTM (Renishaw Co., Inc., Gloucestershire, UK), using a 785 nm laser; and X-ray photoelectron spectroscopy (XPS) was carried out on a Thermo Fisher Scientific K-Alpha (Thermo Fisher Scientific Inc.), using an Al-K α source.

2.4. The degradation of organic pollutants in the dark and light illumination

The degradation reaction was conducted in a flask at room temperature and atmospheric pressure. The flask contained a certain amount of catalyst, H₂O₂ solution (30 wt%), and





Scheme 1 Synthetic process to obtain O/g-C₃N₄.

organic pollutants. The above solution was then stirred rapidly in the dark, with samples taken at fixed time and filtered. The filtrates were collected and measured using a UV-Vis spectrophotometer to quantitate the concentration of residual organic pollutants. Except for LED illumination, the reaction was performed in the dark.

3. Results and discussion

3.1. Characterization of the catalysts

The XRD patterns of the g-C₃N₄, g-C₃N₄(HY), and g-C₃N₄(HY-HT) catalysts showed diffraction peaks at $2\theta = 13.0^\circ$ and 27.5° from g-C₃N₄, which were assigned to the responses of the repeating tri-s-triazine units in the g-C₃N₄ unit layer (100) and inter-planar stacking of the unit layers (002), respectively (Fig. 1).¹ Besides the characteristic peaks of g-C₃N₄, both the g-C₃N₄(HY) and g-C₃N₄(HY-HT) catalysts exhibited no other diffraction peaks, which indicated that the original structure of g-C₃N₄ was retained after the treatment. It was also observed that the intensities of the diffraction peaks followed the order of g-C₃N₄(HY) > g-C₃N₄ >> g-C₃N₄(HY-HT). On the one hand, these results indicated that the degree of crystallinity of g-C₃N₄(HY) slightly increased after the hydrothermal treatment, which was consistent with the literature.²² This was most probably due to the flat layered structure being exposed and/or surface impurities being removed from the catalyst surface during the hydrothermal treatment. Wang *et al.*²³ considered repairing the defects of bulk g-C₃N₄ via the self-assembly of the internal

hydrogen bonds of cyano and amino groups. Ming *et al.*¹⁵ considered the formation of contained O in g-C₃N₄ through hydrolysis and oxidation during hydrothermal processing, and Wang *et al.*²⁴ suggested that the structure of g-C₃N₄ nanosheets was rearranged. On the other hand, the structure of g-C₃N₄(HY-HT) was greatly altered through delamination and depolymerization due to the subsequent calcination. The main reasons for this result are as follows: (i) the contained oxygen groups were generated on the surface and interlayers of the g-C₃N₄ nanosheets;¹¹ (ii) the accumulated g-C₃N₄ particles and their inter-crystalline structures were seriously destroyed,¹³ which may have generated a great number of pores or defects; and (iii) the periodic structure of the g-C₃N₄ nanosheets was reduced.¹⁴ The above analysis accounted for the degree of crystallinity of g-C₃N₄(HY) being relatively higher than that of g-C₃N₄(HY-HT), which was clearly lower than that of the bulk g-C₃N₄. In addition, the strongest diffraction peaks from the g-C₃N₄(HY-HT) catalyst were seen to shift toward higher angles, specifically, $2\theta = 27.8^\circ$, relative to that of g-C₃N₄, resulting in a decrease in the interlayer spacing of the layered structure.¹¹ This was mainly attributed to the doping effect of the heteroatoms and distortion of the graphite structure. In the framework of g-C₃N₄, nitrogen (N) atoms were replaced by O atoms with higher electronegativity during the preparation process, which created stronger interactions between the nanosheets.²⁵ Therefore, the interlayer spacing of the g-C₃N₄ nanosheets became smaller. Similarly, doping with O atoms disturbed the structure of g-C₃N₄. The literature demonstrated that amine functional groups were replaced by O atoms.^{26,27} One electron in an O atom interacted with an adjacent carbon (C) atom to form one C–O group and another electron departed its original location and moved to an *in situ* conjugate triazine ring structure, leading to an electron-rich state for the targeted product.

The SEM micrographs of the g-C₃N₄, g-C₃N₄(HY), and g-C₃N₄(HY-HT) catalysts showed that the g-C₃N₄ catalyst possessed agglomerated particles consisting of nanosheets with relatively rough surfaces (Fig. 2). However, after hydrothermal treatment, the surfaces became smooth, which may have been due to the partial delamination and depolymerization. Also, a notable change appeared in the morphology of g-C₃N₄(HY-HT) relative to that of g-C₃N₄, which may have been because the hydrothermal-calcination treatment led to more delamination and depolymerization, and thus separated more g-C₃N₄ nanosheets, even forming single g-C₃N₄ nanosheets. Meanwhile, the evolution of the nanosheets during high temperature calcination generated some cavities due to delamination and depolymerization. This difference was further confirmed by the TEM

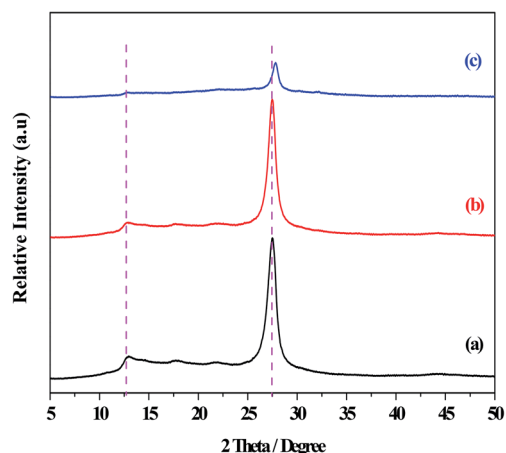


Fig. 1 XRD patterns of the series of catalysts. (a) g-C₃N₄, (b) g-C₃N₄(HY), and (c) g-C₃N₄(HY-HT).



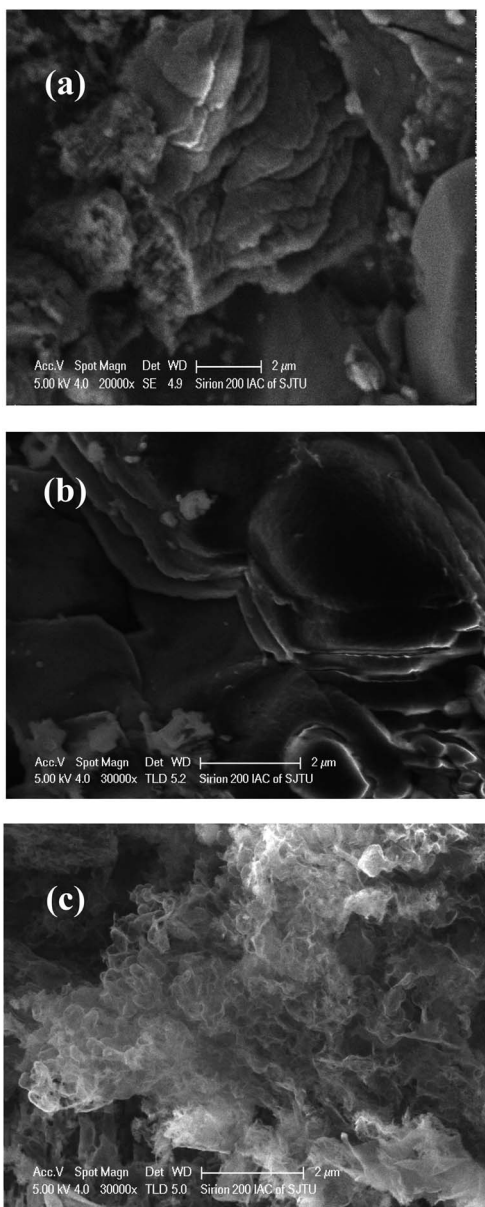


Fig. 2 SEM images of the series of catalysts. (a) $\text{g-C}_3\text{N}_4$, (b) $\text{g-C}_3\text{N}_4(\text{HY})$, and (c) $\text{g-C}_3\text{N}_4(\text{HY-HT})$.

results (Fig. S1, ESI[†]). This deduction was further confirmed by examining the main elemental content in $\text{g-C}_3\text{N}_4$ and $\text{g-C}_3\text{N}_4(\text{HY-HT})$ via EDS characterization (Table S1, ESI[†]). The N content was observed to decrease, but the atomic ratio of C and N atoms increased in the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst relative to that of $\text{g-C}_3\text{N}_4$. This indicated that the process of delamination and depolymerization indeed occurred and N atoms were lost, consequently forming some N defect sites. For $\text{g-C}_3\text{N}_4$, the content of O mainly increased, which was derived from physically adsorbed water. The content of O atoms in the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst was about twice that of $\text{g-C}_3\text{N}_4$, which indicated that more O atoms were doped in the framework structure of $\text{g-C}_3\text{N}_4$ during the preparation process. Accordingly, the SEM results offered additional evidence for the delamination and

depolymerization of the $\text{g-C}_3\text{N}_4$ nanosheets, as deduced from the XRD characterization.

Interestingly, it appeared that the morphology of the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst was similar to that of graphene, specifically, graphene-like carbon nitride may have been generated. Actually, graphene-like carbon nitride have been reported by researchers in the past few years.^{19,28} For example, Niu *et al.*¹⁹ reported the thermal oxidation etching of bulk $\text{g-C}_3\text{N}_4$ using dicyandiamide as a precursor. Similarly, Meng *et al.*²⁸ obtained graphene-like $\text{g-C}_3\text{N}_4$ nanosheets using melamine. In these cases, the specific surface area was generally $>100 \text{ m}^2 \text{ g}^{-1}$. In terms of specific surface area, the values obtained in this study are much higher than that in most recent reports (see below).

The N_2 adsorption-desorption isotherms for the $\text{g-C}_3\text{N}_4$, $\text{g-C}_3\text{N}_4(\text{HY})$, and $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalysts were shown in Fig. S2, ESI[†] and the corresponding pore size-distribution profiles using the Barrett-Joyner-Halenda method, were shown in Fig. S3, ESI[†]. The isotherms for the $\text{g-C}_3\text{N}_4$ and $\text{g-C}_3\text{N}_4(\text{HY})$ catalysts were relatively the same over the whole range of relative pressures, which were characteristic of small pore structures. The isotherm for the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst displayed only a low uptake of N_2 at $p/p_0 < 0.1$, which was related to limited microporosity, typical of $\text{g-C}_3\text{N}_4$.^{11–13} This isotherm also contained an H4-type hysteresis loop, which indicated the presence of mesopores with an irregular and slit-like shape in the catalyst.^{11–13} Mesopores were most probably generated as inter-crystal voids through the aggregation of the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ nanosheets and/or as intracrystal cavities through the delamination and depolymerization of the bulk $\text{g-C}_3\text{N}_4$ during the subsequent hydrothermal-calcination treatment, but the mesoporous structures were generally limited. The pore sizes for the $\text{g-C}_3\text{N}_4$ and $\text{g-C}_3\text{N}_4(\text{HY})$ catalysts were observed to be broadly distributed, covering a range from micropore size up to 10 nm. In contrast, that for the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst was relatively narrow, with its value falling predominantly in the range of 4–7 nm. This result was because the $\text{g-C}_3\text{N}_4$ series catalysts contained nanosheets, and accordingly, the aggregation of the nanosheets in $\text{g-C}_3\text{N}_4$ catalyst was incomplete, thus generating pores with a broad size distribution, compared to the latter. In addition, an appreciably greater amount of pores with sizes $<5 \text{ nm}$ were observed in the $\text{g-C}_3\text{N}_4(\text{HY})$ catalyst compared to that of the bulk $\text{g-C}_3\text{N}_4$, which was most probably due to the generation of additional micropores and mesopores through the delamination and depolymerization after hydrothermal treatment. For the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst, its pore size distribution was broader than that of the $\text{g-C}_3\text{N}_4$ and $\text{g-C}_3\text{N}_4(\text{HY})$ catalysts, which was due to the fact that, on the one hand, a proportion of $\text{g-C}_3\text{N}_4$ nanosheets was exposed through the successive hydrothermal-calcination treatment, as evidenced by the SEM and TEM characterizations, and thus the pore size distribution generated via the aggregation of the $\text{g-C}_3\text{N}_4$ nanosheets became broader. On the other hand, additional mesopores may have been generated within the $\text{g-C}_3\text{N}_4$ nanosheets through delamination and depolymerization during hydrothermal-calcination treatment, which also contributed to the broadening of the pore size distribution. Nevertheless, the pore size distribution was not clearly observed in all the catalysts.



Generally, a large pore size is favorable for mass transfer in pore channels, and therefore enhances catalytic activity.

The textural properties of the $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4(\text{HY})$, and $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalysts were determined via N_2 adsorption-desorption experiments (Table 1). The results showed that, from $g\text{-C}_3\text{N}_4$ to $g\text{-C}_3\text{N}_4(\text{HY})$ and then $g\text{-C}_3\text{N}_4(\text{HY-HT})$, the specific surface area (S_{BET}), first slightly increased and then significantly increased by more than 20-fold. Simultaneously, the total pore volume (V_{total}) first decreased and then remarkably increased by ~ 10 -fold. In addition, the average pore size (D_p) in all three samples decreased. It is well known that $g\text{-C}_3\text{N}_4$ can be easily obtained from various precursors through high-temperature calcination, but the resulting S_{BET} value is often $<10 \text{ m}^2 \text{ g}^{-1}$. This is because the interactions between $g\text{-C}_3\text{N}_4$ layers are very strong, arising from van der Waals forces and/or hydrogen bonds in the material, leading to serious particle aggregation. After hydrothermal treatment, these interactions became weaker due to the attack of water molecules at high temperature and pressure, thus releasing these layers. Subsequently, these exposed nanosheets were further attacked by O in the high-temperature calcination process, and therefore a new morphology appeared, and consequently, the BET greatly increased. Combined with the SEM, TEM, and pore size distribution results, the observed increased BET was concluded to be mainly attributed to the huge morphology changes. Generally, a large S_{BET} is an important factor for strengthening the catalytic performance of materials, which is ascribed to the presence of more exposed adsorption and catalytic active sites.

The FT-IR spectra of the $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4(\text{HY})$, and $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalysts showed that generally, the absorption bands for $g\text{-C}_3\text{N}_4$ were mainly located in three different regions of $2900\text{--}3200 \text{ cm}^{-1}$, $1700\text{--}1000 \text{ cm}^{-1}$, and 807 cm^{-1} (Fig. 3). These bands were assigned to the characteristic peaks of the terminal amino ($-\text{NH}$) and hydroxyl ($-\text{OH}$) groups, typical C–N heterocycles, and tri-*s*-triazine units, respectively. Similar absorption bands were observed for all the catalysts, indicating that the original graphitic structure was basically maintained after the treatment, which was consistent with the XRD results. Moreover, the absorption bands for $g\text{-C}_3\text{N}_4(\text{HY-HT})$ were sharper than that of $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4(\text{HY})$. Combined with the XRD analysis, this further confirmed that tri-*s*-triazine units with less order in a plane appeared in the $g\text{-C}_3\text{N}_4(\text{HY-HT})$ sample. The absorption band at 807 cm^{-1} for $g\text{-C}_3\text{N}_4$ shifted by 1 and 6 cm^{-1} toward higher frequencies in $g\text{-C}_3\text{N}_4(\text{HY})$ and $g\text{-C}_3\text{N}_4(\text{HY-HT})$, respectively. This implied that the interactions between the $g\text{-C}_3\text{N}_4$ nanosheets decreased to some degree, which was due to

the partial substitution of N atoms by O atoms. Fang *et al.*²² suggested that the appearance of this phenomenon could be mainly attributed to an increase in the crystallinity of $g\text{-C}_3\text{N}_4$ and O-doping effects from the hydrothermal treatment. The absorption bands at $3500\text{--}2900 \text{ cm}^{-1}$, corresponding to $-\text{NH}$ and $-\text{OH}$ groups, were almost the same in all three catalysts, which originated from the catalyst and/or physically adsorbed water, respectively. This indicated that $-\text{NH}$ groups still existed in the $g\text{-C}_3\text{N}_4(\text{HY})$ and $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalysts after the hydrothermal and calcination treatments. In the range of $900\text{--}1700 \text{ cm}^{-1}$, the absorption bands for $g\text{-C}_3\text{N}_4$ appeared at 1639, 1571, 1459, 1406, 1318, and 1237 cm^{-1} , which belonged to characteristic peaks of C–N heterocycles, such as C–C and C–N groups. However, due to the near force constant, it was extremely difficult to distinguish the O-containing groups, such as C–O, by FT-IR spectroscopy.^{25,27} Most of the absorption bands in the $g\text{-C}_3\text{N}_4(\text{HY})$ and $g\text{-C}_3\text{N}_4(\text{HY-HT})$ samples shifted toward higher wavenumbers compared to that of the $g\text{-C}_3\text{N}_4$ sample. This confirmed that the C–N and C=N bonds were enhanced because of the addition of O after the hydrothermal and calcination treatments. For example, the absorption bands at 1406 and 1237 cm^{-1} for the $g\text{-C}_3\text{N}_4$ sample were assigned to the $=\text{C}(\text{sp}^2)$ bending and N– sp^3 stretching vibrations, respectively.²⁵ After treatment, these peaks shifted by 2–3 and 3–5 cm^{-1} toward higher wavenumbers, respectively. This was most probably due to the electrophilic effects resulting from the O atoms being located near C–N.²⁷ Also, two new absorption bands appeared at 1082 and 1011 cm^{-1} for the $g\text{-C}_3\text{N}_4(\text{HY-HT})$ sample, which are assigned to the C–O species from the C–O–C and C–OH groups, respectively. Wang *et al.*²⁴ demonstrated by DFT that the N atoms in two-coordinated positions were preferentially displaced by O atoms due to their remarkably low reaction/formation energy. Accordingly, the present functionalization with different O-containing groups was concluded to have been successfully achieved for the $g\text{-C}_3\text{N}_4$ nanosheets after successive hydrothermal and calcination treatment.

The Raman spectra for the $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalysts showed some characteristic peaks at ~ 1613 , 1465, 1345, 1308, 1230, 759, 707, 558, and 477 cm^{-1} for bulk $g\text{-C}_3\text{N}_4$, belonging to the typical characteristic modes of CN heterocycles (Fig. S4†).²⁹ For instance, the peaks at ~ 1613 and 1230 cm^{-1} were assigned to the C=N stretching and $=\text{C}(\text{sp}^2)$ bending vibrations, respectively. Compared to $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4(\text{HY-HT})$ possessed similar characteristic peaks, implying that the successive hydrothermal and calcination treatment for $g\text{-C}_3\text{N}_4$ did not cause any remarkable destruction of the $g\text{-C}_3\text{N}_4$ framework. This was consistent with the

Table 1 The results of the pore structure of the various catalysts^a

Catalyst	$S_{\text{micro}} (\text{m}^2 \text{ g}^{-1})$	$S_{\text{ext}} (\text{m}^2 \text{ g}^{-1})$	$S_{\text{BET}} (\text{m}^2 \text{ g}^{-1})$	$V_{\text{micro}} (\text{cm}^3 \text{ g}^{-1})$	$V_{\text{meso}} (\text{cm}^3 \text{ g}^{-1})$	$V_{\text{total}} (\text{cm}^3 \text{ g}^{-1})$	$D_A (\text{nm})$
$g\text{-C}_3\text{N}_4$	0	11.5	11.5	0.00352	0.06565	0.06927	24.13
$g\text{-C}_3\text{N}_4(\text{HY})$	0	17.02	17.02	0.00474	0.05522	0.05996	14.09
$g\text{-C}_3\text{N}_4(\text{HY-HT})$	0	236.4	236.4	0.06221	0.57219	0.6344	10.74

^a S_{micro} , S_{ext} and S_{BET} respectively represented the microporous surface area, external surface area and specific surface area; V_{micro} , V_{meso} and V_{total} represented the microporous pore volume, mesoporous pore volume and total pore volume, respectively; and D_A represented the average pore diameter.



XRD and FT-IR analyses. Furthermore, the peaks at 1613, 1345, 1230, 759, and 477 cm^{-1} slightly shifted toward higher frequencies, and four new peaks appeared at 1567, 1489, 1150, and 1116 cm^{-1} , suggesting that the O-doping effects led to chemical bond changes near the C atoms.

The binding energies of C-1s and N-1s from the XPS characterization of the $\text{g-C}_3\text{N}_4$, $\text{g-C}_3\text{N}_4(\text{HY})$, and $\text{g-C}_3\text{N}_4(\text{HY-HT})$ samples showed that in the bulk $\text{g-C}_3\text{N}_4$ sample, there were two peaks at 284.9 and 288.4 eV for the C-1s spectra (Fig. 4, Table S2†). These were attributed to the reference C and sp^2 -hybridized C (N-C=N), respectively. Four different peaks at 398.9, 399.8, 401.5, and 404.6 eV appeared in the N-1s spectra, corresponding to the sp^2 -hybridized N (N-C=N , denoted as $\text{sp}^2(\text{N})$), sp^3 -hybridized N (N-[C]_3 , denoted as $\text{sp}^3(\text{N})$), N-H bonds, and π bands in the localization of heterocycles, respectively (Table S3†).^{10,11,14–18} In terms of the $\text{g-C}_3\text{N}_4(\text{HY})$ and $\text{g-C}_3\text{N}_4(\text{HY-HT})$ samples, in their C-1s spectra, besides the two peaks at ~ 284.8 and 288.4 eV, a new peak appeared at ~ 289 eV, which was attributed to the C-O bonds.^{10,11,14–18} Notably, the content of C-O species remarkably increased and their binding energy shifted toward lower values in the $\text{g-C}_3\text{N}_4(\text{HY})$ sample relative to that of $\text{g-C}_3\text{N}_4(\text{HY-HT})$. This indicated that partial oxidation continued to occur after the hydrothermal and high-

temperature oxidation calcination. Nevertheless, as observed in the earlier FT-IR analysis, the C-O bonds may have originated from the formation of C-OH and/or C-O-C species. In the N-1s spectra, the $\text{g-C}_3\text{N}_4(\text{HY})$ and $\text{g-C}_3\text{N}_4(\text{HY-HT})$ samples possessed peaks similar to that of $\text{g-C}_3\text{N}_4$. In addition, the concentration of $\text{sp}^2(\text{N})$ and ratio of $\text{sp}^2(\text{N})$ and $\text{sp}^3(\text{N})$ were observed to first decrease and then increase, whereas the content of the N-H band and $\text{sp}^3(\text{N})$ initially increased and then decreased for $\text{g-C}_3\text{N}_4(\text{HY})$ and $\text{g-C}_3\text{N}_4(\text{HY-HT})$ compared to $\text{g-C}_3\text{N}_4$. This accounted for some $\text{sp}^2(\text{N})$ and $\text{sp}^3(\text{N})$ atoms being substituted for O atoms during the treatment, specifically, a decrease in $\text{sp}^2(\text{N})$ atoms meant the formation of C-OH and/or C-O-C by replacing the N atoms in C-NH and C=N-C. Meanwhile, the decrease in $\text{sp}^3(\text{N})$ atoms indicates the significant decomposition of the entire $\text{g-C}_3\text{N}_4$ nanosheets. Actually, the mass of $\text{g-C}_3\text{N}_4(\text{HY-HT})$ was confirmed to be remarkably reduced relative to that of $\text{g-C}_3\text{N}_4$, which illustrated that N loss preferentially occurred, and thus some N defect sites were formed. In addition, the content of π bands in the heterocycles increased after the treatment, which implied the formation of many more unpaired electrons.

Based on the above characterization results, the mechanism for the preparation of the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst was proposed.

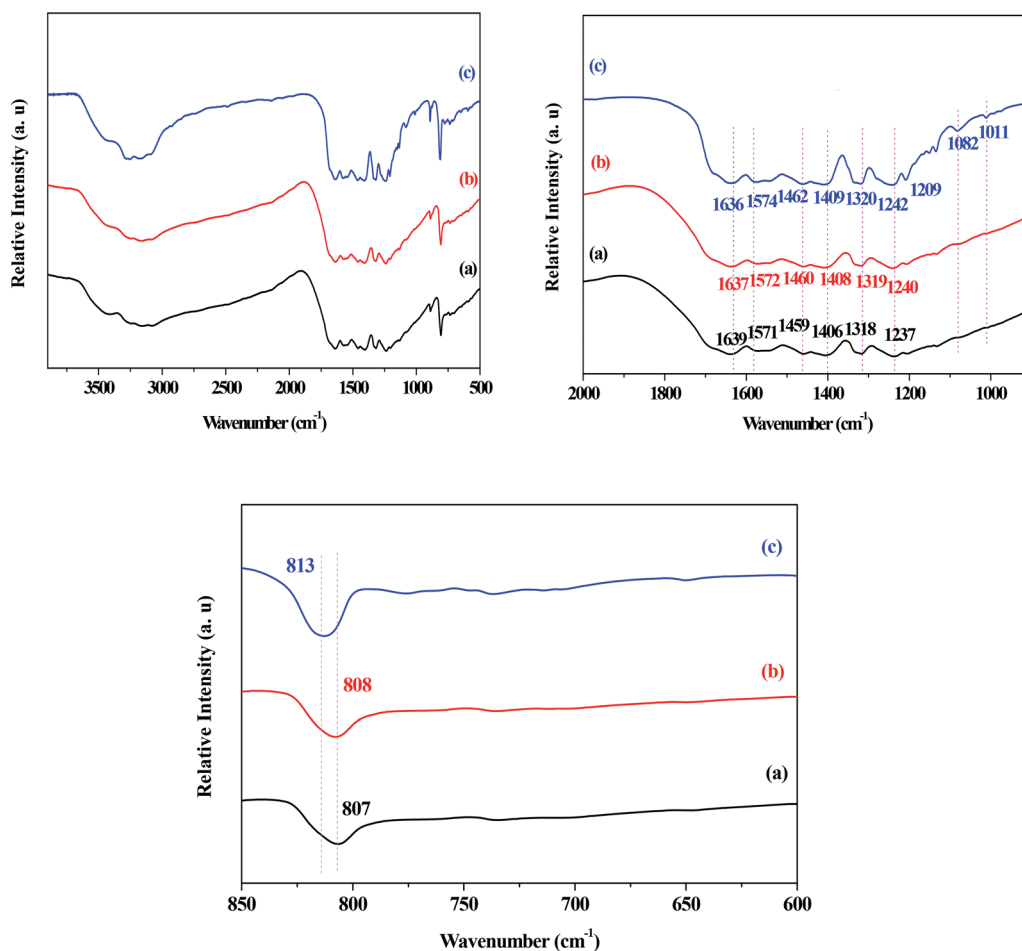


Fig. 3 FT-IR spectra of the series of catalysts. (a) $\text{g-C}_3\text{N}_4$, (b) $\text{g-C}_3\text{N}_4(\text{HY})$, and (c) $\text{g-C}_3\text{N}_4(\text{HY-HT})$.



Melamine is one of the promising precursors for the synthesis of $g\text{-C}_3\text{N}_4$ and only through high temperature calcination can bulk $g\text{-C}_3\text{N}_4$ be easily generated. Here, once the bulk was generated, a portion of the bulk structure was altered by hydrothermal treatment. During this process, $g\text{-C}_3\text{N}_4$ was induced to generate some new O-containing groups, such as $(\text{C}_3\text{H}_3\text{N}_3)_3\text{-OH}$, which were mainly attributed to H_2O attack during the delamination and depolymerization.¹⁷ With an increase in the hydrothermal time and/or temperature, more $(\text{C}_3\text{H}_3\text{N}_3)_3\text{-NH}_2$ was transformed into $(\text{C}_3\text{H}_3\text{N}_3)_3\text{-OH}$. Also,

a weak alkaline solution, arising from NH_3 , was detected by pH monitoring. Finally, the $g\text{-C}_3\text{N}_4(\text{HY})$ catalyst was again treated at high temperature. On the one hand, the $g\text{-C}_3\text{N}_4(\text{HY})$ nanosheets were exfoliated at high temperature, and thus fewer layers of nanosheets were obtained. On the other hand, delamination and depolymerization continued to occur through interactions with O_2 in the air atmosphere. In this case, the material was vulnerable to decomposition into small molecular gas products, such as CO_2 , NH_3 , and NO_x . These strong reactions drove these products to cross the layers, specifically, they acted as a gas

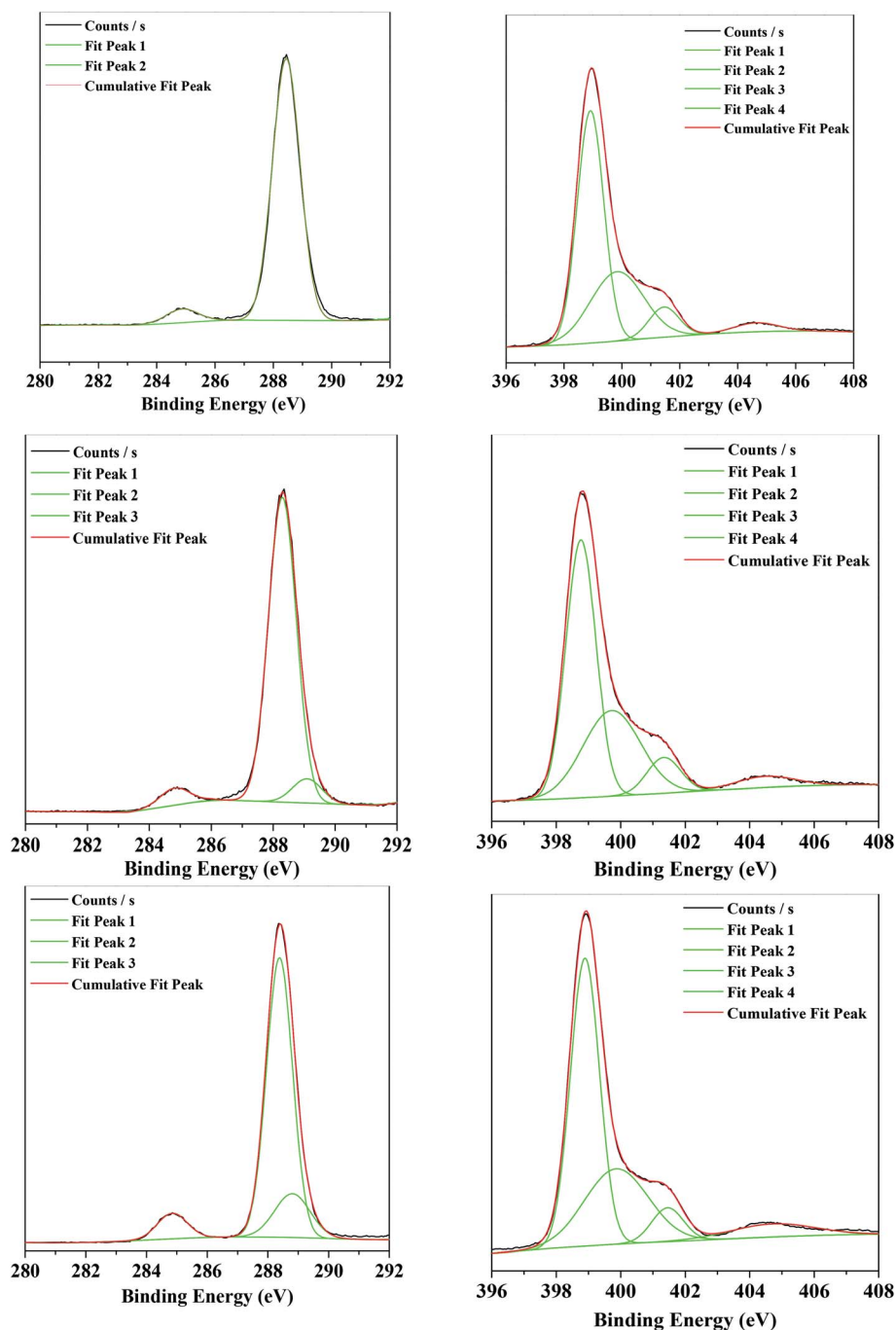


Fig. 4 XPS results of C1s (left) and N1s (right) for the various catalysts. (Top) $g\text{-C}_3\text{N}_4$, (middle) $g\text{-C}_3\text{N}_4(\text{HY})$, and (Bottom) $g\text{-C}_3\text{N}_4(\text{HY-HT})$.

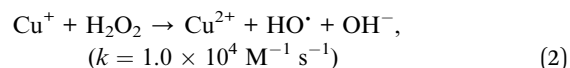
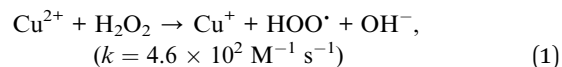


template. Accordingly, a cotton morphology in the g-C₃N₄ nanosheets was formed, and concurrently, many defects were generated. For example, some N atoms were preferentially removed from the framework of g-C₃N₄(HY) under high-temperature calcination, and thus N vacancy defects were formed. Consequently, a higher specific surface area and rich defect population were generated in the g-C₃N₄(HY-HT) catalyst relative to g-C₃N₄(HY). Actually, the color of g-C₃N₄ turned from light yellow to white after calcination, which meant that the degree of oxidation was extremely serious. At a lower calcination temperature was used, the color of the treated catalyst hardly changed from than that of g-C₃N₄. However, the higher the calcination temperature, the whiter the color of the treated catalyst was obtained. However, it completely disappeared at >600 °C, which was attributed to the structural damage to the catalyst, and therefore, the material was consumed at higher temperature. Notably, the amount of g-C₃N₄(HY) during calcination was a key factor for the formation of the final products. Under identical conditions, high doses of g-C₃N₄(HY) barely generated the target products because of insufficient O₂ to react with the g-C₃N₄(HY) nanosheets. In contrast, no product was obtained if the dose of g-C₃N₄(HY) was too small due to complete oxidation. Accordingly, a suitable dosage of g-C₃N₄(HY) during the second high temperature treatment was quite important and necessary.

3.2. Catalytic performance for the degradation of RhB under various reaction conditions

3.2.1. Comparison of various catalysts for the degradation of RhB. Different types of solid catalysts were compared for their dark Fenton-like degradation of RhB (Fig. 5). The tested TiO₂, Fe₂O₃ and CeO₂ materials were found to all exhibit extremely low catalytic activities for the degradation of RhB even though H₂O₂ was added to the reaction solutions, which implied that this was mainly a simple adsorption process. The CuO and MgO

catalysts displayed moderate activities under the same reaction conditions, which was due to the fact that CuO contained different copper ions, such as Cu²⁺ and Cu⁺. These ions had the ability to directly react with H₂O₂ to produce reactive O species (ROS),³⁰ which then easily attacked and removed organic pollutants (eqn (1) and (2)).



Consequently, with an increase in the reaction time, RhB degradation continued. Similarly, MgO had many alkaline active sites, and therefore was capable of the *in situ* activation of H₂O₂ for removing RhB. Wu *et al.*³¹ proposed that active O in MgO₂ was obtained *via* the reaction between MgO and H₂O₂, acting as active sites, and then eliminating MB in solution. Actually, the above results were not ideal because of the limited activating ability of CuO and MgO with H₂O₂ for producing ROS. In contrast, the g-C₃N₄(HY-HT) catalyst showed superior activity in this reaction and ~81% of RhB was decolorized within only 60 min. These results illustrated that, here, the introduction of O in g-C₃N₄ by post-treatments was beneficial for yielding a high specific surface area, which exposed more active sites. In this case, more ROS were generated from H₂O₂, thus degrading more RhB in water, and accordingly, remarkably increased the degradation efficiency. Usually, metal ions are extremely difficult to treat in aqueous solution, and thus the use of metal-based catalysts induces secondary pollution. In contrast, here, the catalytic process of g-C₃N₄(HY-HT) does not release toxic metallic ions since it is a metal-free strategy for the green degradation of organic pollutants.

The catalytic activity of g-C₃N₄ after the post-treatments for the degradation of RhB in dark Fenton-like reactions was

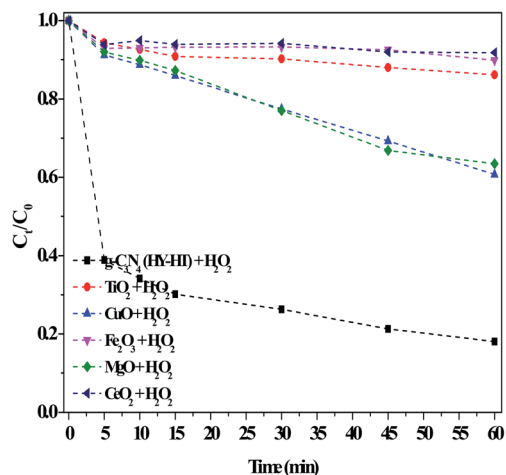


Fig. 5 Comparison of the performance of various heterogeneous catalysts for the degradation of RhB in the dark. Conditions: [RhB]₀ = 10 mg L⁻¹, [H₂O₂] = 36 mM, [Catal.] = 1.0 g L⁻¹, initial pH no adjustment, 25 °C and 60 min.

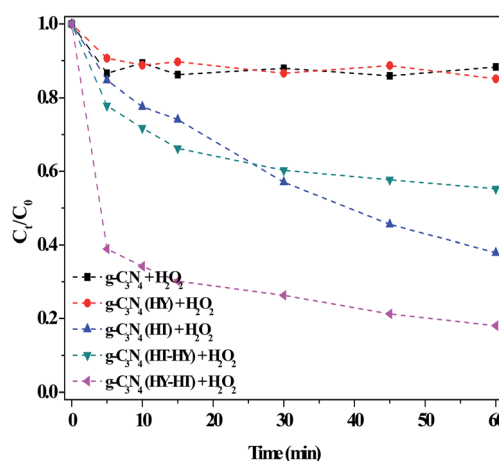


Fig. 6 Effect of the preparation method on the degradation of RhB in the dark. Conditions: [RhB]₀ = 10 mg L⁻¹, [H₂O₂] = 36 mM, [Catal.] = 1.0 g L⁻¹, initial pH no adjustment, 25 °C and 60 min.

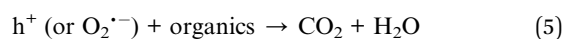
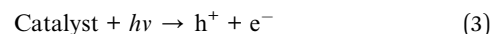


carefully studied (Fig. 6). The degradation efficiency of RhB over $g\text{-C}_3\text{N}_4(\text{HY})$ was only $\sim 14\%$ at 60 min in the presence of H_2O_2 , which was almost the same as that over the bulk $g\text{-C}_3\text{N}_4$. This indicated that this treatment had no effect on the RhB degradation ability, although the former possessed a higher specific surface area than the latter. After the second high-temperature treatment of $g\text{-C}_3\text{N}_4$ (denoted as $g\text{-C}_3\text{N}_4(\text{HT})$), the RhB degradation evidently increased, which showed that catalytic activity was enhanced by the second calcination. However, no useful products were obtained if a third high-temperature calcination was applied. Thus, based on these results for $g\text{-C}_3\text{N}_4(\text{HT})$, we speculated about the enhanced catalytic activity observed by introducing hydrothermal treatment (denoted as $g\text{-C}_3\text{N}_4(\text{HT-HY})$). Surprisingly, this resulted in a decrease in the degradation of RhB (*ca.* 40%) during a 60 min reaction, which indicated that the treatment had a negative impact on RhB degradation after the hydrothermal treatment of the $g\text{-C}_3\text{N}_4(\text{HT})$ catalyst. One of the reasons for this may have been that the crystallinity of $g\text{-C}_3\text{N}_4(\text{HT})$ increased due to the reduction in defect sites after the hydrothermal treatment, similar to that for the hydrothermal treatment of $g\text{-C}_3\text{N}_4$, which was confirmed by the XRD results. This meant that defects in the catalyst were a necessary factor for enhancing its catalytic activity. Further exploration of how to intensify this effect was underway *via* sequential hydrothermal and high-temperature treatment for the production of $g\text{-C}_3\text{N}_4$ catalysts. Encouragingly, the degradation of RhB prominently increased in the dark. The synthetic process of $g\text{-C}_3\text{N}_4(\text{HT})$ was noted to be similar with that of graphene-like carbon nitride with a high specific area,¹⁹ where the latter had not been reported thus far regarding dark Fenton-like reactions. However, these results showed that its catalytic activity was still lower than that of $g\text{-C}_3\text{N}_4(\text{HY-HT})$, which illustrated that $g\text{-C}_3\text{N}_4(\text{HY-HT})$ possessed a higher degradation efficiency than the current reported non-metal catalysts for the degradation of organic pollutants *via* dark Fenton-like reactions.

To gain further insight into this phenomenon, the impact of various catalysts on the degradation of RhB in the dark was systematically studied (Fig. S5†). With an increase in the reaction time, the trend of RhB degradation efficiency initially decreased and then remained unchanged when H_2O_2 and the $g\text{-C}_3\text{N}_4$ -based catalysts were respectively added to the reaction solution. All the degradation efficiencies were no more than 15% at 60 min. With only H_2O_2 , RhB degradation barely occurred due to the limited oxidation. For the $g\text{-C}_3\text{N}_4$ -based catalysts, their activity was attributed to adsorption. After the maximum adsorption capacity of catalyst was reached, the degradation process stopped. Also, the degradation significantly increased on the $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalyst relative to $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4(\text{HY})$, which was mainly attributed to its high specific surface area. To alter this situation, a useful strategy was exploited. After the introduction of H_2O_2 in the solution, a large difference was notably observed, with the highest RhB degradation efficiency over $g\text{-C}_3\text{N}_4(\text{HY-HT})$ among them. This was due to the combined action between $g\text{-C}_3\text{N}_4(\text{HY-HT})$ and H_2O_2 . As a comparison, the $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4(\text{HY})$ catalysts in the presence of H_2O_2 exhibited slight degradation in the dark,

which was consistent with the literature.¹ The principal reason for these results was attributed to the synergistic effect between the increased specific surface area and abundant defects in the $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalyst compared to that of $g\text{-C}_3\text{N}_4$ and $g\text{-C}_3\text{N}_4(\text{HY})$. This effect was due to that large surface area of the catalyst, which was beneficial for improving reactant absorption capacity and also offered more catalytic active sites. Meanwhile, the defect sites in the $g\text{-C}_3\text{N}_4(\text{HY-HT})$ catalyst activated H_2O_2 *in situ* to create more ROS, although its activating ability was not high in this reaction. Actually, RhB was completely degraded after several hours. In summary, the best activity was achieved in the $g\text{-C}_3\text{N}_4(\text{HY-HT})$ and H_2O_2 mixed system. Here, the catalytic activity of $g\text{-C}_3\text{N}_4(\text{HY-HT})$ was dramatically strengthened by H_2O_2 activation without any light irradiation. This is a useful finding regarding Fenton-like reactions since, to the best of our knowledge, there are few reports on the effective degradation of organic pollutants in the dark involving a $g\text{-C}_3\text{N}_4$ -based catalyst without any added metal.

Examination of the effects of the dark and LED illumination on the degradation of RhB showed that $g\text{-C}_3\text{N}_4(\text{HY-HT})$ with the help of H_2O_2 exhibited significantly better catalytic activity under light illumination relative to that in the dark (Fig. 7). This was because some new reactions simultaneously occurred based on the dark Fenton-like reactions, and they were further capable of promoting the degradation of RhB, which was a typical photo-Fenton-like reaction process (eqn (3)–(5)).



For example, under the light irradiation, a photo-generated hole directly reacted with RhB. In addition, this special electron could have interacted with dissolved oxygen in aqueous

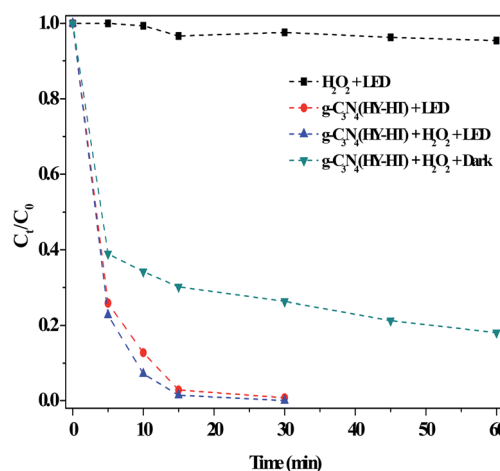


Fig. 7 Effect of the dark and light illumination on the degradation of RhB. Conditions: $[\text{RhB}]_0 = 10 \text{ mg L}^{-1}$, $[\text{H}_2\text{O}_2] = 36 \text{ mM}$, $[\text{Catal.}] = 1.0 \text{ g L}^{-1}$, initial pH no adjustment, 25°C and 60 min.



solution to form $O_2^{\cdot-}$ species, and thus the degradation of RhB clearly increased. In previous studies,^{7–9} for metal- $g-C_3N_4$ -based materials, the degradation of organic pollutants was slightly different in both the dark and under light illumination. This was mainly attributed to be the existence of metal–N bonds. Notably, the present $g-C_3N_4$ (HY-HT) catalyst contained abundant defects but still possessed excellent catalytic activity under LED illumination. This result indicated that the defects in the $g-C_3N_4$ (HY-HT) catalyst played a positive role in the degradation of RhB. In general, these defects can be divided into surface and bulk defects. The former always promotes photocatalytic activity, which is ascribed to its acceleration of the separation of photogenerated charges and role as adsorption sites. The latter inhibits reactions because it becomes the recombination center of photogenerated charges. Therefore, here, $g-C_3N_4$ (HY-HT) with surface defects exhibited increased photocatalytic activity. The literature^{11,15} demonstrated that O-containing groups in $g-C_3N_4$ remarkably enhanced its photocatalytic activity in many reactions relative to that of the parent catalyst. To examine this hypothesis, $g-C_3N_4$ oxidized by H_2O_2 hydrothermal treatment was employed, which indicated that the degradation efficiency over this catalyst was higher than that of the bulk $g-C_3N_4$ (Fig. S6†). Specifically, the defects and O-containing groups in $g-C_3N_4$ (HY-HT) both played important roles for the degradation of RhB under visible light illumination.

3.2.2. Optimal reaction conditions over $g-C_3N_4$ (HY-HT) for the degradation of RhB. The effects of the dose of $g-C_3N_4$ (HY-HT) on the RhB degradation efficiency in the dark showed that the degradation efficiency decreased with an increase in the dose, where the optimum value was 1.0 g L^{-1} , and above this dose, it began to decrease (Fig. 8). Thus, the degradation of RhB was closely related to the catalyst dose, which was because a higher dose provided more adsorption sites, and thus the removal of RhB increased (Fig. S7†). However, an excessive catalyst dose resulted in the agglomeration and increased

counterforces between the catalyst particles, and thus, the RhB removal efficiency decreased. When the catalyst amount was low, there were insufficient catalytic active sites to activate H_2O_2 to produce ROS, and therefore, the degradation efficiency decreased. In contrast, too much catalyst generated surplus ROS, which could be easily captured by H_2O_2 , similarly resulting in reduced ROS availability, and consequently, a decrease in RhB degradation. Relatively, this negative effect was mainly attributed to the decrease in adsorption by the catalyst in the above Fenton-like reactions.

The influence of the concentration of H_2O_2 on the degradation of RhB in the dark was examined, and the degradation efficiency was found to increase with an increase in the H_2O_2 concentration, reaching a maximum at a $C_0(H_2O_2)$ of 36 mM, and a further increase in the concentration of H_2O_2 resulted in a decrease in degradation (Fig. S8†). When the H_2O_2 concentration was low, insufficient ROS were formed. Accordingly, the increase in H_2O_2 concentration increased RhB degradation. However, too much ROS had some detrimental effects on the degradation of RhB, which was due to reactions by ROS, such as with $\cdot OH$ and H_2O_2 . In this case, the possibility of $\cdot OH$ species attacking RhB became lower, and therefore the oxidation kinetics were much slower. Cui *et al.*¹ proposed that many side reactions occur, which were derived from the self-effects of radical species.

Examination of the effects of initial pH on the RhB degradation efficiency in the dark Fenton-like reactions showed that the degradation efficiency was $\sim 81\%$ in 60 min with no adjustment of the solution pH (Fig. 9). After regulating the pH using diluted H_2SO_4 solution (0.1 M) to strongly acidic (pH = 2), the degradation efficiency remained $>40\%$. This result was satisfactory in principle compared to that of the currently reported results.^{1,7–9} For instance, Ge *et al.*⁷ found that methyl orange degradation over $g-C_3N_4/MgO$ hardly occurred at pH < 2 and under visible light irradiation. RhB was found here to be more quickly degraded when the initial pH was adjusted to pH

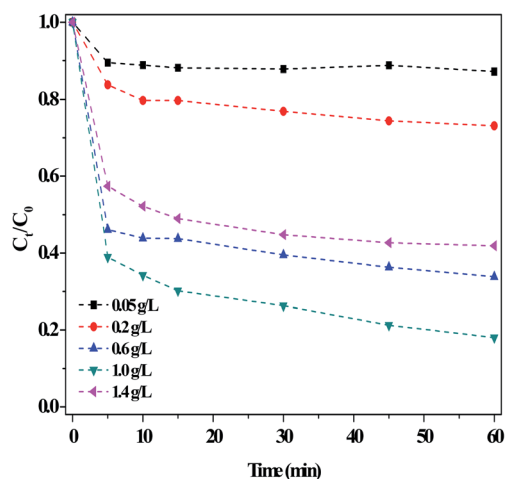


Fig. 8 Effect of the dosage of the $g-C_3N_4$ (HY-HT) catalyst on the degradation of RhB in the dark. Conditions: $[RhB]_0 = 10\text{ mg L}^{-1}$, $[H_2O_2] = 36\text{ mM}$, initial pH no adjustment, 25°C and 60 min.

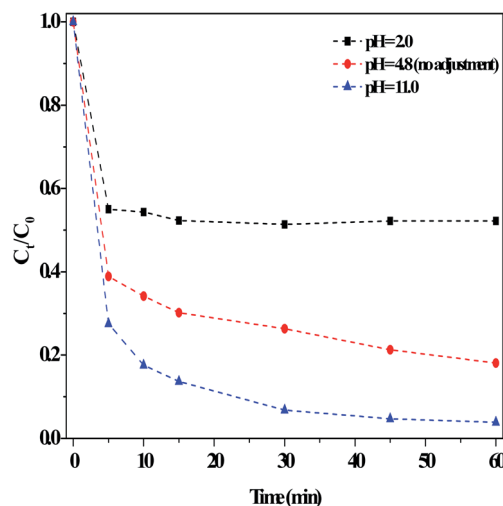


Fig. 9 Effect of adjusting the initial pH using H_2SO_4 or NaOH on the degradation of RhB in the dark. Conditions: $[RhB]_0 = 10\text{ mg L}^{-1}$, $[Catal.] = 1.0\text{ g L}^{-1}$, $[H_2O_2] = 36\text{ mM}$, 25°C and 60 min.

11 using diluted NaOH solution (0.1 M). This result was remarkably better than that under other conditions, where the alkaline conditions reinforced the catalytic activity. However, Ge *et al.*⁷ reported that ~5% of RhB degradation efficiency was lost when the initial pH was increased from 8 to 10. The initial pH had a different influence on the catalyst adsorption capacity (Fig. S9†). It has been reported that the zero potential point (pK_a) of g-C₃N₄ obtained from urea was about 3,²² which meant that the surfaces of g-C₃N₄-based materials became negatively charged if the pH_{solution} was above pK_a , and similarly, their surface was protonated at $pH_{\text{solution}} < pK_a$. Accordingly, RhB developed a pseudo-positive charge at low pH_{solution} , thus leading to a surface repulsion effect with the catalyst. Once the pH_{solution} increased, the effect between catalyst surface and RhB molecule decreased, resulting in an increase in absorption. Clearly, the present catalyst adsorbed less RhB at lower pH, whereas pH had little effect on the degradation of RhB at higher pH. In addition, at pH = 2, H⁺ covered the catalyst defect sites, and thus generated less ROS in the above dark Fenton-like reactions. Thus, the decrease in RhB degradation at pH = 2 was mainly attributed to the reduced adsorption by catalyst, whereas the increase in degradation at the high pH of 11 was because the electrode potential value of OH[•]/OH decreased with a higher OH[•] concentration.¹ In this case, more OH species were generated, thus resulting the greater degradation of RhB. Specifically, this catalyst, assisted by H₂O₂, exhibited a good catalytic performance over a broad pH range. Also, the degradation of MB was examined using this catalyst under the above reaction conditions, and the result clearly indicated that MB was effectively removed in a short time (Fig. S10†).

The stability of g-C₃N₄(HY-HT) was investigated in cyclic experiments for the degradation of RhB under the same reaction conditions (Fig. S11†). After a reaction period, the catalyst was washed with ultrapure water and ethanol, and then used for the next run. The catalyst activity decreased slightly after five cycles, demonstrating its good stability and reusability (Fig. S11†). This result was further confirmed by XRD and FT-IR characterizations (Fig. S12 and S13,† respectively). According to the results, the structure of the used catalyst was concluded to be fundamentally same as that of the fresh catalyst.

Additionally, the degradation of various organic pollutants in wastewater over the g-C₃N₄-based catalysts in dark Fenton-like reactions was examined (Table S4†). As is known, g-C₃N₄ has low photocatalytic activity under visible light illumination,¹ and a low RhB degradation efficiency (<10%) is often observed in the dark, even in the presence of H₂O₂. The main drawbacks associated with this catalyst are its extremely low specific surface area and lack active sites for H₂O₂. As the substitute for g-C₃N₄, one current strategy is to use metal-modified g-C₃N₄. For instance, MgO/g-C₃N₄,⁷ Cu/g-C₃N₄,^{8,32} and FeO_x/g-C₃N₄ (ref. 33) have been employed to degrade organic pollutants in the dark, but they have many issues, including long reaction time, activity at low pH, high cost, and poor stability. Furthermore, metal ions need to be employed, which easily cause environmental contamination. Thus, these factors constitute the main challenges for their industrial application. To overcome these limitations, an N-g-C₃N₄ catalyst was employed for the

degradation of MB,¹⁰ achieving a non-metal-containing catalyst. Nevertheless, the required H₂O₂ concentration is clearly higher than with metal/g-C₃N₄ catalysts. In the present study, g-C₃N₄(HY-HT) was employed as a catalyst for the degradation of organic pollutants in water. Compared to the above catalysts, g-C₃N₄(HY-HT) is green, non-toxic, and easy to prepare. For this catalyst, a large specific surface area was obtained, and when assisted by H₂O₂, it exhibited excellent performances over a wide pH range in the dark. More importantly, the activity of this catalyst was further and significantly enhanced under LED illumination. The above merits indicate that the novel method for synthesis of g-C₃N₄(HY-HT) and its application developed in this study exhibit great potential for commercial application.

3.2.3. Identification of adsorption and catalytic active sites and the mechanism of H₂O₂ activation

(1) *Identification of adsorption and catalytic active sites.* In advanced oxidation processes, O groups, N species, and defects usually act as catalytic sites for activating oxidation to generate ROS. To confirm the role of O-containing groups in the catalyst, a simulation experiment was conducted, and the g-C₃N₄ catalyst was hydrothermally treated with H₂O₂,¹¹ which was one of most common methods for the preparation of O-doped g-C₃N₄. Many O-containing groups were formed, including C=O, C-O-C-, and -NO_x, and the results showed that the degradation of RhB by this catalyst barely increased relative to that by pristine g-C₃N₄ (Fig. S14†), which indicated that these groups in g-C₃N₄ were not main active sites in this reaction. In the g-C₃N₄(HY-HT) catalyst, the O-containing groups were mainly C-O-C and/or C-OH, as confirmed by the FT-IR and XPS results. For example, the C-O-C group was observed in the g-C₃N₄(HY) catalyst, but catalytic degradation results were almost the same for the g-C₃N₄ and g-C₃N₄(HY) catalysts. This indicated that this group was not a catalytically active site. Also, C-OH groups cannot directly activate H₂O₂. The above results demonstrated that O-containing groups were not responsible for catalyzing H₂O₂ to generate ROS for the degradation of organic pollutants in solution. As previously mentioned, g-C₃N₄(HY-HT) possessed abundant defects, and its enhanced RhB degradation confirmed this. It has been reported that defects not only increased reactant adsorption, but also prolonged the bond length of H₂O₂ and reduced the reaction free energy.³⁴ To further confirm these deductions, an experiment was performed in which g-C₃N₄(HY-HT) was first treated with NaF to exchange its -OH groups.¹¹ The results indicated that the removal of RhB was not obvious, which implied that defect sites, besides the -OH groups, were the main adsorption sites in this reaction (Fig. S15†). Specifically, defects served as both adsorption and catalytic active sites in this reaction, with these defects dynamically filled with reactants and/or products during reactions. Accordingly, this process resulted in a reduction of the defect site concentration, and therefore decreased the generation of ROS. With an increase in the reaction time, the RhB degradation rate became much slower. As proof of this effect, RhB decolorization in aqueous solution required several hours (Fig. S16†). After washing, the catalyst was easily regenerated, and its initial catalytic activity was recovered.



(2) *Identification of ROS*. It is well known that ROS generation in H_2O_2 -based Fenton-like reactions generally involve $\cdot\text{OH}$, $\text{O}_2^{\cdot-}$, and $^1\text{O}_2$. Thus, to reveal the activation mechanism, a series of quenching experiments were conducted to identify the dominant ROS in the “ $\text{g-C}_3\text{N}_4(\text{HY-HT}) + \text{H}_2\text{O}_2$ ” system. Different scavengers, including isopropanol (2-PrOH), potassium iodide (KI), 1,4-benzoquinone (BQ), and furfuryl alcohol (FFA), were added separately to this reaction system (Fig. S17†). The RhB degradation efficiency was observed to decrease by up to 65% after 60 min with 30 mM 2-PrOH compared to the control experiments, which showed that the surface-bound $\cdot\text{OH}$ radicals were an important active species here (Fig. S17a†). Meanwhile, the addition of KI appeared to have little effect on the degradation, which indicated that the $\cdot\text{OH}$ in the aqueous solution did not play a role in the catalytic reaction because of its extremely short lifetime. With the addition of BQ, the RhB degradation efficiency reached >80% after 60 min of reaction, which indicated that $\text{O}_2^{\cdot-}$ radicals were also not the active species here (Fig. S17b†). In addition, the addition of FFA, as a capture agent, hardly had any influence on the RhB degradation efficiency (Fig. S17c†), which indicated that singlet oxygen ($^1\text{O}_2$) was also not an active ROS here. The above results clearly showed that the $\cdot\text{OH}$ radicals were the main ROS in this reaction. Thus, it was concluded that the degradation pathway in this reaction was a simultaneous adsorption and surface-reaction process. Specifically, adsorption through a non-radical pathway occurred, while the surface-reaction step involved the typical radical route. The good adsorption observed mainly originated from the large catalyst surface area, while the defect sites were main contributor for activating H_2O_2 for subsequent damage to the molecular structure of organic pollutants, as evidenced by the UV-Vis results (Fig. S18†) and also the reaction pathway shown in Fig. S19.†

4. Conclusions

In this work, a novel O/ $\text{g-C}_3\text{N}_4$ catalyst possessing a large surface area and defect sites for the efficient degradation of organic pollutants was successfully developed through a three-step post-treatment method. The $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst possessed the highest specific surface area and more defects relative to that of the $\text{g-C}_3\text{N}_4$ and $\text{g-C}_3\text{N}_4(\text{HY})$ catalysts, which was used for treating organic pollutants *via* H_2O_2 activation. The $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst exhibited the best catalytic performance among various catalysts. Furthermore, the influencing factors were optimized, *e.g.*, catalyst dosage, concentration of H_2O_2 and initial pH. Besides, this catalyst retained good stability after use in five cycles. The huge surface area and defects in the $\text{g-C}_3\text{N}_4(\text{HY-HT})$ catalyst played a key role in the degradation of organic pollutants in the dark Fenton-like reaction system. The degradation pathway was dominated by the synergistic effect between adsorption and chemical oxidation. This work offers a new catalyst with the ability to *in situ* activate H_2O_2 for the degradation of organic pollutants *via* dark Fenton-like reactions.

Conflicts of interest

There are no conflicts to declare.

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