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Nitrogen-doped hollow carbon spheres with tunable shell thickness for high-performance supercapacitors†

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Nitrogen-doped hollow carbon spheres (NHCSs) are well prepared by using Cu₂O microspheres as a hard template and 3-aminophenol formaldehyde resin polymer as carbon and nitrogen precursors. The thickness of the carbon shell can be easily controlled in the range of 15–84 nm by simply adjusting the weight ratios of the precursors to Cu₂O microspheres, and the Cu₂O templates can also be further reused. Physicochemical characterization demonstrates that the obtained NHCSs possess a well-developed hollow spherical structure, thin carbon shell and high nitrogen doping content. Due to these characteristics, when further utilized as electrodes for supercapacitors, the NHCSs with the carbon shell thickness of 15 nm show a high capacitance of 263.6 F g⁻¹ at 0.5 A g⁻¹, an outstanding rate performance of 122 F g⁻¹ at 20 A g⁻¹ and an excellent long-term cycling stability with only 9.8% loss after 1000 cycles at 5 A g⁻¹ in 6 M KOH aqueous electrolyte. This finding may push forward the development of carbon materials, exhibiting huge potential for electrochemical energy storage applications.

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1. Introduction

Hollow carbon spheres (HCSs) have been attracting significant scientific attention due to their unique structure and excellent performance, such as very high specific surface, good mechanical flexibility, outstanding chemical stability and high conductivity.^{1–3} These remarkable properties enable them to be widely applied in the field of carriers for drugs, catalysis, fuel batteries and supercapacitors.^{4–8} In particular, HCSs are beneficial in the application of supercapacitors because of their high conductivity and transport lengths for mass and charge transport.^{3,9} In addition, research shows that doping of heteroatoms, such as nitrogen (N) phosphorus (P), boron (B) and sulfur (S), has attracted plenty of attention as an effective way to improve the electrochemical performance by the faradic process generated from the additional electrochemical active sites.^{10–14} Specifically, it is well known that the nitrogen-doped carbon spheres (NHCSs) materials can enhance the capacitance of the material by improving their electronic conductivity and surface wettability.¹⁵ Currently, a great many of research efforts have

been devoted to NHCSs-based electrolyte materials for supercapacitors.

NHCSs have been prepared by means of various synthetic approaches, including the Chemical Vapor Deposition (CVD), electric arc and template methods.^{16–22} Among them, the template method has attracted people's attention due to its advantages such as simple operation and low cost. Template method can be divided into two types: soft-template method and hard-template method. Based on the co-self-assembly of carbon precursors and decomposable amphiphilic molecules, the soft template process is quite easy, while the morphology controllability is relatively unsatisfied.²³ The hard templates are commonly used to synthesize homogeneous NHCSs by coating the spherical template core with carbon precursors, followed by carbonization and subsequent template etching by acids or organic solvents.^{15,24} Qu²⁵ has developed a facile modified template method to synthesize NHCSs in the presence of resorcinol/formaldehyde as carbon precursors and ethylenediamine as both a base catalyst and nitrogen precursor, the obtained NHCSs exhibit a large specific surface area and unique hollow structure. Therefore, the hard template method is simple and effective way to prepare NHCSs. However, the templates currently used are mainly SiO₂ nanoparticles and polymer latex particles,^{26,27} these templates are non-recyclable, which increases the cost of hard template method to synthesize NHCSs in some certain cases. Thence, we propose a method that the template can be used again, *i.e.* Cu₂O used as a hard template. Remarkably, the Cu₂O template can be recycled by the evaporation and crystallization of final waste liquid and followed by reduction for repeated production of NHCSs. It makes

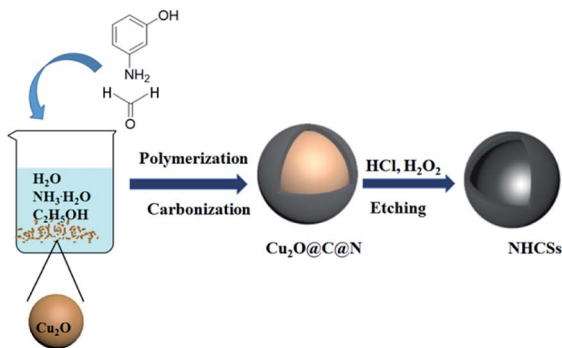
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Scheme 1 Illustration of the synthesis process of N-HCSs.

the process highly cost-effective and environmentally benign for large-scale production, especially while combining with cheap precursors. Moreover, the thickness of carbon shell can be well controlled. It is realized that the synthesis with an adjustable thickness of carbon shell can effectively improve the dielectric loss characteristics and increase matching characteristic impedance of NHCSSs, resulting in enhancing electrochemical performances. However, there are few reports about NHCSS materials with tunable shell thickness in supercapacitors.

Herein, we present a simple and effective strategy to synthesize NHCSSs with the well-preserved spherical morphology, carbon shell thickness-controllable and high-nitrogen content *via* a hard template method (Scheme 1). The fabrication process uses 3-aminophenol and formaldehyde polymer resins as carbon and nitrogen precursor, and Cu₂O solid spheres as hard template. Moreover, the shell thicknesses of NHCSSs can be adjusted by changing the ratio of 3-aminophenol to Cu₂O. Meanwhile the obtained materials reveal fully composed of spherical particles with the diameter of 600–700 nm and shell thickness can be easily controlled in the range of 15–84 nm. In particular, NHCSSs with the thinnest carbon shell as supercapacitor electrodes exhibit a high specific capacitance of 263.6 F g⁻¹ at a current density of 0.5 A g⁻¹, and long-cycling experiments present good cycle stability in 6 M KOH aqueous solution.

2. Experimental

2.1. Synthesis of Cu₂O template

Cu₂O solid spheres with the diameter of 600–700 nm were synthesized *via* a solution-phase method. For synthesis of the spheres, 50 ml of 0.4 M CuCl₂·2H₂O solution was mixed with 100 ml of 1 M NaOH and 0.1 g of cetyltrimethylammonium bromide (CTAB) and stirred at room temperature to form a uniform solution. Then, 60 ml of 0.5 M glucose solution and 15 ml of 1 M N₂H₄ solution were then added successively to the above solution by strong stirring. During this process, the color of the solution changed gradually into brick-red indicating the formation of the yellow precipitate of cuprous oxide (Cu₂O), the precipitate was collected by centrifugation and washed sequentially with distilled water and ethanol several time, then dried at 60 °C for 5 h under vacuum.

2.2. Synthesis of N-doped hollow carbon spheres (NHCSSs)

Nitrogen-doped hollow carbon spheres (NHCSSs) were synthesized by using Cu₂O microspheres as hard template, 3-aminophenol formaldehyde resin polymer as carbon and nitrogen precursors. The synthesis mechanism is illustrated in Scheme 1. Firstly, 1.5 g of Cu₂O microspheres were dispersed in a stock solution, which contained 80 ml of water, 32 ml of absolute ethanol, and 0.4 ml of ammonia, then the mixed solution was stirred at room temperature for more than 1 h. Subsequently, 0.1 g of 3-aminophenol was added into the solution. The mixture was stirred at room temperature for another 1 h before adding the corresponding amount of formaldehyde solution (38 wt%) to start the polymerization for 24 h. Then the previous mixture solution was transferred into an autoclave and sealed to heat at 180 °C for 6 h. The obtained brown black particles were filtered and washed with deionized water and ethanol to remove residues, and evaporated at 60 °C under vacuum. The carbonization was carried out in a tubular furnace under nitrogen atmosphere in two steps: the sample was heated to 300 °C for 1 h (heating rate: 1 °C min⁻¹) and 1000 °C for 2 h (heating rate: 5 °C min⁻¹). The resulted composites (named Cu₂O@C@N) were immersed in concentrated 30 ml of hydrochloric acid (37 wt%), and 30 ml of hydrogen peroxide solution (10 wt%) added dropwise into the solution, then stirred for 8 h at room temperature to dissolve the Cu₂O templates, this process was repeated at least three times until the copper ions were completely removed. The NHCSSs materials were collected by centrifugation, washed with distilled water and dried at 60 °C for 5 h under vacuum. Finally, the obtained samples were denoted as S1. For comparison, we prepared three additional sets of samples with different carbon shell thicknesses. (denoted as S2, S3 and S4, respectively. ESI, Table S1†).

2.3. Characterization

The morphology of samples were investigated by field emission scanning electron microscope (FESEM, Quanta 200 FEG) and transmission electron microscope (TEM, JEM2100). The crystallographic structures of the samples were analyzed by X-ray diffraction (XRD, X Pert PRO) equipped with Cu-K α radiation ($\lambda = 0.154$ nm) source at scanning speed of 10° min⁻¹ in the range of 10–80° and Raman spectrum (DXR-Raman) under $\lambda_{exc} = 514$ nm laser excitation. The pore structure of the materials was examined by N₂ adsorption–desorption at 77 K (Micrometrics, ASAP 2020). The BET surface area (S_{BET}) of samples was calculated according to the BET (Brunauer–Emmett–Teller). The chemical bonding state of carbon, oxygen and nitrogen elements in NHCSSs were studied by X-ray photoelectron spectroscopy (XPS, Thermo Scientific ESCALAB250Xi).

2.4. Electrochemical measurement

The working electrodes were prepared by mixing the NHCSSs with 10 wt% of polytetrafluoroethylene (PTFE) and acetylene black (the mass ratio of 8 : 1 : 1) in *N*-methyl-2-pyrrolidone to form a homogenous slurry. The slurry was coated on nickel foam (1 cm × 2 cm) with a surface area of 1 cm², the electrodes



were dried in vacuum at 120 °C for 12 h and pressed under a pressure of 10 MPa for 30 s. The loading active material was in the range of 3.0–6.0 mg cm⁻². All electrochemical behaviour of the working electrodes were investigated in a three-electrode system with 6 M KOH aqueous solution as the electrolyte. The saturated calomel electrode (SCE) and platinum were served as reference electrode and counter electrode, respectively. Cyclic voltammetry (CV), galvanostatic charge–discharge (GCD) and life cycle test were conducted on a CHI660E electrochemical workstation (Shanghai Chenhua, China). The range of CV test varied from –1.0 to 0 V with the different scan rates of 10–100 mV s⁻¹ and the current density of GCD were from 0.5 to 20 A g⁻¹. The specific capacitance was calculated from GCD according to equation:²⁸

$$C_g = I\Delta t / \Delta V m;$$

where C_g is the gravimetric-specific capacitance, I is the discharge current (A); Δt is the discharge time (s), ΔV is the potential window (V) and m is the mass of the active material on the working electrode (g), respectively.

3. Result and discussion

The morphology and microstructure of the prepared samples were characterized by scanning electron microscope (SEM) and transmission electron microscope (TEM). As can be seen in Fig. 1a, the Cu₂O microspheres present fully composed of spherical particles with the diameter of 600–700 nm, and these microspheres have rough surfaces. After self-polymerization of 3-aminophenol formaldehyde resin, followed by carbonization and etching template, all obtained NHCSs mainly inherit the spherical structure with smooth surfaces, and the diameters of the carbon spheres are almost unchanged compared with the Cu₂O templates. Fig. S1a–d (ESI†) clearly show the morphology

evolution of NHCSs with different carbon content. Some cracked NHCSs observed suggest their hollow structure (as shown in white rectangle of Fig. 1b). It should be mentioned that using the prepared Cu₂O microspheres as cores, carbon coating can be achieved on these microspheres through *in situ* polymerization of 3-aminophenol formaldehyde resin following by a carbonization process to produce composites with hollow structures. When the amount of 3-aminophenol gradually increases, the thickness of NHCSs display monotonic increase with more 3-aminophenol and formaldehyde applied in the experiments, and the hollow structure of the NHCS with the larger thickness is also more and more difficult to be destroyed during the etching template process. (Fig. S1, ESI†). According to the above results, we can conclude that the content of 3-aminophenol and formaldehyde in synthesis system plays a key role on the morphology and the thickness of the NHCSs. Fig. 1c–f show the TEM images of all the samples, those images further reveal the hollow spherical morphology of the NHCSs. In addition, these spherical NHCSs went through a slight deformation, this might be attributed to the fact that cuprous oxide loses one oxygen atom after being reduced by carbon, resulting in a change in the spherical structure of the hard tin plating agent, thereby slightly deforming the spherical morphology of the NHCSs. Furthermore, it is obviously that the variation of the content of carbon source affects the uniformity and thickness of the NHCSs remarkably. From these TEM images, we can see that those hollow structured NHCSs can be fabricated with the diameter of >600 nm and the thicknesses of carbon shells for S1, S2, S3 and S4 are about 15 nm, 32 nm, 63 nm and 84 nm, respectively, while the carbon shell of S1 collapses after removing the Cu₂O cores due to its thinnest carbon shell. These above results further indicate that the carbon shell can be successfully prepared on the cores by properly controlling the experimental conditions and the relative ratio of 3-aminophenol to Cu₂O microspheres, and more importantly, the shell thickness can be adjusted at the nanoscale.

The N₂ adsorption/desorption characterization was further carried out to evaluate the pore structure of NHCSs. The adsorption/desorption isotherm and the corresponding pore size distribution are shown in the ESI (Fig. S2a and S2b),† respectively. All samples exhibit typical type IV isotherms according to the IUPAC classification, indicating the assembly of micro and mesoporous.²⁹ The pore size distribution was analyzed using the Barrette–Joyner–Halenda (BJH) method and the distribution curves of the S1, S2, S3 and S4 show that their average pore sizes are around 1.3 nm (Fig. S2b, ESI†). A summary of the surface area, average pore size and total volume is recorded in Table 1, The BET surface areas of S1, S2, S3 and S4 are 112.36, 74.05, 25.65, and 18.93 m² g⁻¹ and the pore volumes are 0.34, 0.24, 0.08 and 0.07 cm³ g⁻¹, respectively, which is basically negatively correlated with the thickness of carbon shell of NHCSs. Notably, the microporous and mesoporous assembly structures of NHCSs could provide a huge amount of “ion-adsorption sites” for the formation of electric double layer, showing a promise to explore the synthesized NHCSs as electrode materials for advanced supercapacitor.

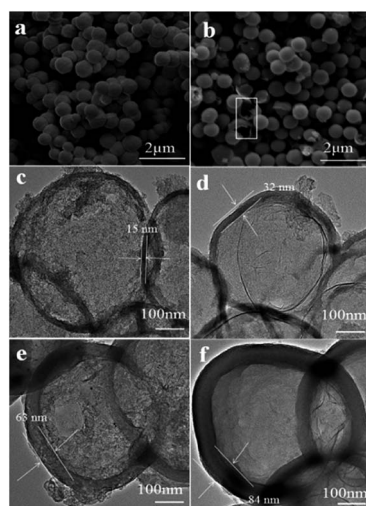


Fig. 1 SEM images of (a) Cu₂O microspheres and nitrogen-doped hollow carbon spheres (b) S1; TEM images of NHCSs with different thickness of carbon shell (c) S1, 25 nm (d) S2, 32 nm (e) S3, 63 nm and (f) S4, 84 nm.



Table 1 Texture properties of NHCSs

Sample	S_{BET} ($\text{m}^2 \text{g}^{-1}$)	V_{p}^a ($\text{cm}^3 \text{g}^{-1}$)	Pore size ^b (nm)	C ^c (wt%)	O ^c (wt%)	N ^c (wt%)
S1	112.36	0.34	1.21	86.98	7.93	5.09
S2	74.05	0.24	1.34	86.72	8.27	5.01
S3	25.65	0.08	1.27	87.69	7.34	4.97
S4	18.93	0.07	1.49	87.26	7.79	4.95

^a Total pore volume from N_2 adsorbed at $p/p_0 = 0.995$. ^b Maximum of pore size distribution. ^c C, N and O, the content of carbon, oxide and nitrogen tested by XPS.

The carbon state in NHCSs was identified by the X-ray diffraction (XRD) and Raman spectroscopy. As shown in Fig. 2a, all samples exhibit two typical broadened diffraction peaks at about $2\theta = 26^\circ$ and 43° , which can be approximately indexed as (002) and (100) reflection of standard graphite, respectively.^{25,30,31} The pronounced intensity of the wide (002) diffraction peak attributing to stacks of parallel layer planes indicates certain graphitizing degree of the as fabricated NHCSs. With the decrease of the thickness of carbon shell, the (002) peak gradually has a slight upward trend, indicating that the NHCSs has the tendency of graphitization. By comparing the XRD pattern of Cu_2O and $\text{Cu}_2\text{O}@C@N$ (as shown in Fig. S2, ESI†), there is no peak which is ascribed to copper (or cuprous oxide) appeared in the pattern, implying the complete removal of copper and cuprous oxide species during the templates etching. The Raman spectrum of the NHCSs depicted in Fig. 2b exhibits two peaks at about 1350 cm^{-1} (D band) and 1580 cm^{-1} (G band). The D band is considered to the crystal defects and disordered structures of carbon materials, while the G band is attributed to the phonon mode with E_{2g} symmetry of graphite.^{31,32} By calculating the relative intensity of the D band and the G band ($I_{\text{D}}/I_{\text{G}}$), the graphitic degree of the carbon materials can be determined.³³ The $I_{\text{D}}/I_{\text{G}}$ intensity ratios of S1, S2, S3 and S4 are 0.98, 0.99, 1.01 and 1.02, respectively, indicating the high degree of defects for obtained NHCSs, which is good agreement with the above XRD result.

The chemical composition and electronic state of the materials were analyzed by X-ray photoelectron spectroscopy (XPS). As is shown in the ESI (Fig. S3),† the XPS spectrum of the obtained NHCSs shows the presence of typical C 1s, N 1s and O 1s peaks, which clearly indicates that the N atoms were successfully doped into the carbon structure in NHCSs. The surface chemistry of NHCSs is dominated by carbon, oxygen and

nitrogen, and the content of nitrogen in materials are 5.09, 5.01, 4.97 and 4.95 wt%, listed in Table 1. Fig. 3a presents the C 1s spectrum of NHCSs, which can be decomposed into the following four Gaussian peaks: 284.6 eV (sp^2 -hybridized graphitic C, C-I), 285.9 eV (sp^2 C-N and C-O in phenols and ethers, C-II), 286.8 eV (C=O, C-III), and 288.7 eV (O-C=O, C-IV).^{33,34} Gaussian Fitting of N 1s spectra indicates pyridinic-N, pyrrolic-N, quaternary-N, attributable to binding energies of 398.1 eV, 399.5 eV and 400.7 eV,^{35,36} respectively (Fig. 3b). As reported previously, pyridinic-N and pyrrolic-N play a key role in increasing the specific capacitance of N-doped materials due to their pseudo-capacitive contribution, whereas quaternary-N can increase the conductivity of the material and thus improve the capacitance performance.^{37,38} In addition, the suitable nitrogen and oxygen doping can increase the surface wettability with aqueous electrolyte and thus reduce the resistance of electrode materials.^{34,39,40}

To evaluate the electrochemical properties of the prepared N-doped carbon materials for supercapacitor electrodes, the samples were characterized by cyclic voltammetry (CV) and galvanostatic charge-discharge (GCD) measurements within the potential window of -1.0 to 0 V in 6 M KOH aqueous solution. Fig. 4 shows the CV curves of NHCSs electrodes at scan rates of 10 – 200 mV s^{-1} . All the CV curves display approximately rectangular shapes, indicating the typical characteristic of electrical double layer capacitor (EDLC).³⁴ From the results of the curves, we can see that the S1 gives the largest quasi-rectangular area at the same scan rate, namely, the highest specific capacitance, which could be ascribed to the ultrafast electron transfer and electrolyte transportation rates, which is caused by the specific surface area and ultra-thin carbon layer of S1.

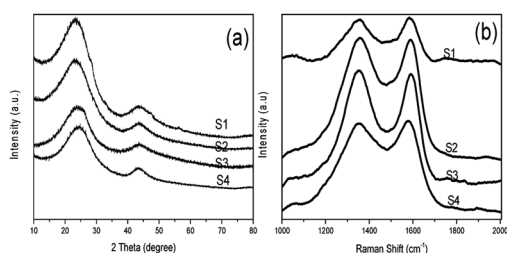


Fig. 2 (a) XRD pattern and (b) Raman spectra of NHCSs (S1, S2, S3 and S4).

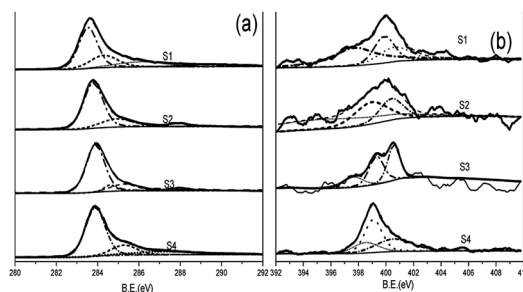


Fig. 3 XPS spectrum of (a) C 1s spectrum and (b) N 1s spectrum of NHCSs.



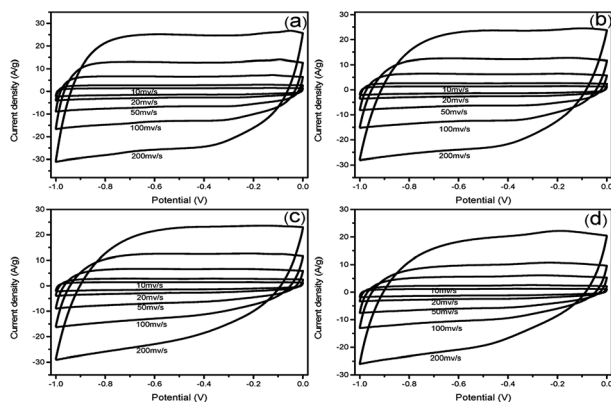


Fig. 4 CV curves of NHCSs with different carbon shell thickness (a) S1, (b) S2, (c) S3 and (d) S4 at different scan rates range from 10 to 200 mV s^{-1} in 6 M KOH electrolyte.

The GCD curves of at various current densities (0.5, 1, 2, 5, 10, 20 A g^{-1}) in 6 M KOH aqueous solution display in Fig. 5a–d respectively. Basically, all these curves emerge in a nearly symmetric triangle shape along with a slight distortion, furthermore confirming the typical characteristic of EDLC. According to the formula,²⁸ the specific capacitance of S1 at 0.5 A g^{-1} is 263.6 F g^{-1} , higher than that of S2 (241.1 F g^{-1}), S3 (225.0 F g^{-1}) and S4 (217.7 F g^{-1}). Through the comparison of these four samples, it can be clearly seen that S1 shows the highest specific capacitance, indicating that the specific capacitance of NHCSs is negatively correlated with the thickness of carbon shell. The optimal specific capacitance of S1 may come from its highest specific surface area, the thinnest carbon shell and the largest pore volume pore size, which facilitates the formation of electric bilayers. In addition, nitrogen doped in the carbon structure can also create pseudo-capacitance and improve the surface wettability.

The specific capacitance values of all samples decrease with the increase of charge/discharge current densities (as shown in Fig. S4, ESI[†]), which is due to the satirical limitations for partial ions penetrating into micropores at higher current density, whereas smoothly diffusing at lower current density.⁴¹

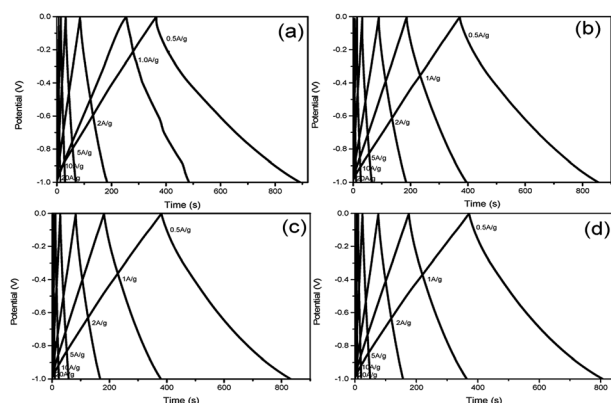


Fig. 5 Galvanostatic charge–discharge curves of NHCSs (a) S1, (b) S2, (c) S3 and (d) S4 at various current densities (0.5–20 A g^{-1}).

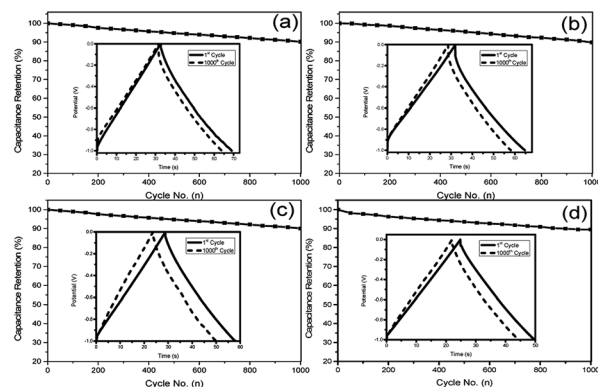


Fig. 6 Cycling stability measured for NHCSs (a) S1, (b) S2, (c) S3 and (d) S4 at 5 A g^{-1} over 1000 cycles in 6 M KOH electrolyte.

Obviously, the obtained S1 exhibits better electrochemical performance than others, which can be attributed to two reasons. Firstly, the thin carbon shell and the high S_{BET} can provide shorter ion transport distances and more electrochemically active sites for ion accommodation and fast ion migration. Secondly, the high amount of nitrogen doping and the tendency of graphitization can provide S1 with pseudo-capacitance and increase its conductivity. Both of the two factors generate a synergistic effect, which is advantageous for the high capacitance of S1. The cycling performance of supercapacitors is another crucial property for practical applications, and the cycling stability of NHCSs electrode was estimated by galvanostatic charge–discharge profiles for 1000 cycles at current density of 5 A g^{-1} . As depicted in Fig. 6a–d, the specific capacitances of those four samples all remain at about 90%, indicating that they all demonstrate good cycle performance. In addition, taking S1 with the best electrochemical performance among all samples, we also investigated the morphological changes of NHCSs after long time charging and discharging. Fig. S6[†] shows the SEM image of S1 after the long cycle test. The image shows that a small amount of carbon spherical morphology has been damaged, which could be ascribed to one of the reasons for its good chemical cycle stability.

4. Conclusions

In summary, using Cu_2O microspheres as recycle hard template, NHCSs with tunable shell thickness have been successfully synthesized through *in situ* polymerization of 3-aminophenol formaldehyde resin followed by high-temperature carbonization and templates etching. The prepared NHCSs combine the characteristics of special morphology, controllable shell thickness, unique porosity and nitrogen functional groups. Electrochemical tests showed that S1 with the thinnest carbon shell exhibited outstanding supercapacitor performance with good capacitance (263.6 F g^{-1} at 0.5 A g^{-1}), ultrahigh-rate performance (122.0 F g^{-1} at 20 A g^{-1}) and excellent long-term cycling stability (90.2% retention after 1000 cycles at 5 A g^{-1}) in 6 M KOH aqueous electrolyte. Based on the unique characteristics of the NHCSs, we believe that the well-developed



NHCSs with high electrochemical performance provide new opportunities for energy storage and energy conversion applications, such as supercapacitors and lithium–sulfur batteries.⁴²

Conflicts of interest

There are no conflicts to declare.

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