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Exploring a new dinuclear Fe(\mathfrak{m}) complex for the fixation of atmospheric CO₂ and optical recognition of nano-molar levels of Zn²⁺ ions†

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A dinuclear Fe(III) complex (F1) of an imine derivative (L1) derived from 3-ethoxy-2-hydroxy-benzaldehyde and hydrazine, structurally characterised via single crystal X-ray studies, is employed for the catalytic

conversion of epoxides to cyclic carbonates utilizing carbon dioxide. In addition, F1 is employed for the

selective optical recognition of nano-molar levels of Zn^{2+} (42.23 nM) via a metal displacement approach.

The Job plot reveals interactions between F1 and Zn²⁺ at a 1 : 3 molar ratio with an association constant of 7.71×10^4 M⁻¹. Studies on the catecholase-like activity of F1 reveal a k_{cat} value of 4.42×10^3 h⁻¹.

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Introduction

Despite the high abundance of iron at the earth's surface, its bio-availability at physiological pH under aerobic conditions is low.^{1,2} However, the role of iron in bio-systems is highly significant. For example, it is involved in energy transfer and bio-catalysis,³ and acts as a cofactor of several enzymes.⁴ Iron deficiency may cause liver dysfunction and anaemia, whereas its presence at higher concentrations may lead to hemochromatosis and Parkinson's disease.⁵ It is well known that iron deficiency in pregnant women can adversely affect both the mother and infant, including an increased risk of sepsis, maternal mortality, perinatal mortality, and low birth weight.⁶

Similarly, Zn²⁺ plays a significant role in several biological processes,⁷ including in brain activity, gene transcription, immune functioning, *etc.*^{8,9} It is a cofactor of almost three hundred enzymes,¹⁰ including those involved in DNA repair and gene expression.¹¹ On the other hand, abnormal zinc metabolism leads to several health issues, including delayed sexual maturation, prostate cancer, type 2 diabetes mellitus (T2DM), Wilson's disease, amyotrophic lateral sclerosis (ALS), and agerelated macular degeneration (AMD).¹²⁻¹⁴ It is believed that Zn²⁺ homeostasis has some correspondence with the pathology

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† Electronic supplementary information (ESI) available: Contains information regarding materials; physical measurements; the general UV-vis and fluorescence measurement methods; Job experiments; binding constant, detection limit, and quantum yield measurements; copies of different spectra scans; CV data; crystal parameters; bond lengths and angles; catalytic performance data; *etc.* The CCDC no. of **F1** is 1864282. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/d0ra01698e

of Alzheimer's disease¹⁵ and other brutal neurological problems, such as cerebral ischemia⁴ and epilepsy.⁵

Currently, the catalytic activities of metal complexes have acquired significant attention. Scientists are attempting to lower carbon dioxide levels^{16–18} in the atmosphere through its conversion to cyclic carbonates using metal complexes as catalysts. The cyclo-addition of CO₂ and epoxides, forming cyclic carbonates, is a process with 100% atom economy^{19–21} that can lead to reduced CO₂ levels. Moreover, cyclic carbonates have enormous numbers of industrial applications, for example as non-toxic/polar/high boiling solvents, electrolytes in lithium ion batteries, reactive intermediates, *etc.*²² However, the low reactivity of CO₂ means that a catalyst^{23–25} is required for the cyclo-addition reactions to be feasible. Iron compounds,²⁶ having low intrinsic toxicity, can catalyse the combination of CO₂ and epoxides to form cyclic carbonates.^{27,28}

These facts motivated us to study the catalytic activity of a new iron complex for the fixation of CO_2 as cyclic carbonates using different epoxides.

Additionally, catechol oxidase of the oxidoreductase class of enzymes is very vital in biological systems.^{29–34} The use of transition metal complexes as bio-catalysts that can mimic catechol oxidase has become a pioneering field of research.

Therefore, the above discussion clearly demonstrates our motivation for investigating the optical recognition^{35–38} of the bio-relevant metal ion Zn(n) and the (bio)catalytic application of a new Fe(m) complex.

Herein, a new dinulcear Fe(m) complex (F1), structurally characterized *via* single crystal X-ray diffraction analysis, is used for the selective optical recognition of nanomolar levels of Zn²⁺. F1 functions as an excellent catalyst for the fixation of atmospheric CO₂ as cyclic carbonates using epoxides. Moreover, the catecholase-like activity of F1 has been investigated. Finally, F1

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is also established as a good extractor of $\mbox{Zn}(\pi)$ from water to ethyl acetate.

Experimental

Synthesis

N,*N*[']-Bis-(2-hydroxo-3-ethoxy-benzylidene)-hydrazine (L1). 3-Ethoxy-2-hydroxy-benzaldehyde (166.17 mg, 1 mmol) was refluxed with a slight excess of hydrazine in methanol for 8 h to get L1 in 91.5% yield. Its molecular formula is $C_{18}H_{20}N_2O_4$ (MW = 328.36 g mol⁻¹). Anal. found (%): C, 65.65; H, 6.20; and N, 8.02; calcd (%): C, 65.84; H, 6.14; and N, 8.53. ESI-MS (*m*/*z*): [M]⁺, 327.75 (Fig. S1a, ESI[†]). ¹H NMR (CDCl₃, 400 MHz, *J*, Hz, δ ppm) (Fig. S1b, ESI[†]): 13.928 (1H, s) 8.234 (1H, s), 6.824–6.841 (1H, q) 6.783–6.761 (1H, q), 6.723–6.684 (1H, t, *J* = 7.8 Hz), 4.097–4.041 (2H, m), 3.304–3.280 (1H, q). FTIR (KBr, cm⁻¹): 3336, ν (O–H); 3196, ν (C–H, aromatic); 1684, ν (C=O, stretch); 1500, ν (C=N); 1068, ν (N–N) (Fig. S1c, ESI[†]).

Dinuclear Fe(III) complex (F1)

Methanol solution of FeCl₃·6H₂O (27 mg, 0.1 mmol, 5 mL) was added to DMF solution of L1 (50 mg, 1.5 mmol, 5 mL), and the mixture was stirred for 2 h. The mixture was kept for a week for the growth of black crystals, which were suitable for SC-XRD analysis (Fig. 1a). The yield was 64%. Its molecular formula is $C_{54}H_{54}Fe_2N_2O_{12}$ (MW = 1090.73 g mol⁻¹). Anal. found (%): C, 59.41; H, 5.13; and N, 7.10; calcd (%): C, 59.46; H, 4.99; and N, 7.70. ESI-MS (*m*/*z*): [M]⁺, 1113.08 (Fig. S2a, ESI⁺). FTIR (KBr, cm⁻¹): 3011, ν (C=H, aromatic); 1639, ν (C=O); 1451, ν (C=N); 1140, ν (N–N) (Fig. S2b, ESI⁺).

Zn(II) adduct (Z1)

Ethanol solution of **F1** was stirred with $Zn(OAc)_2$ and kept for 3 days to get the **Z1** adduct in 59% yield (Fig. 1b). Its molecular formula is $C_{20}H_{23}N_2O_5Zn$ (MW = 436.80 g mol⁻¹). Anal. found (%): C, 54.72; H, 5.40; and N, 6.04; calcd (%): C, 54.99; H, 5.31; and N, 6.41. ESI-MS (*m*/*z*): [M]⁺, 435.50 (Fig. S3a, ESI⁺). FTIR (KBr, cm⁻¹): 3169, ν (C–H, aromatic); 165, ν (C=O); 1481, ν (C=N); 1227, ν (N–N) (Fig. S3b, ESI⁺) (Scheme 1).



Scheme 1 The synthetic protocol for L1.

Results and discussion

The single crystal X-ray structure of F1 (ref. 39)

It is desirable to confirm the structure of a new compound via single crystal X-ray diffraction (SC-XRD) analysis. Thus, in addition to the spectroscopic characterization of F1, its structure is authenticated from SC-XRD analysis (Fig. 1a). The crystallographic data and refinement parameters are presented in Table S1 (ESI).[†] Selected bond lengths and angles are listed in Table S2 (ESI).[†] F1 is triclinic, having the space group $P\overline{1}$ (CCDC 1864282), with a (Å): 14.8885(7); b (Å): 15.0825(6); c (Å): 15.1336(6); α (°): 104.859(3); β (°): 104.285(3); γ (°): 117.418(3); volume (Å³): 2639.0(2); and Z = 2, and adopts octahedral geometry around both Fe(m) centres. L1 exhibits cis geometry in the complex with selected bond lengths N1-C00P: 1.294(9) Å; and N2-C00U: 1.284(9) Å, indicating the presence of -CH=N double bonds. Bond angles of O004-Fe01-O007: 96.1(2)°; O004-Fe01-O003: 94.2(2)°; O003-Fe01-N3: 92.3(2)°; O004-Fe01-N00J: 83.9(2)°; O007-Fe01-N00J: 93.6(2)°; and N3-Fe01-N00J: 89.3(2)° clearly reveal the octahedral geometry around the Fe(m) centres. Symmetrical L1 contributes to each of the Fe(m) centres unbiasedly and co-ordinates via phenol O- and aldimine Ngroups, with Fe01-O003 and Fe01-N2 bond lengths of 1.920(4) Å and 2.197(6) Å, respectively.

Spectroscopic studies

The absorption characteristics of **F1** were studied using UV-vis spectroscopy. **F1** exhibits two significant absorption peaks at 296 nm and 358 nm, which can be assigned to π - π * and n- π * electronic transitions (Fig. 2a). Upon the gradual addition of





Fig. 1 (a) The SC-XRD structure of F1. (b) The ChemDraw structure of Z1.



Fig. 2 (a) Changes in the absorption spectra of F1 (20 μ M) upon the gradual addition of Zn²⁺ (0–2000 μ M) in HEPES-buffered EtOH/H₂O (4/1 v/v) media; pH: 7.4. (b) Absorption spectra of F1 (20 μ M) in the presence of different cations (2000 μ M), viz. Zn²⁺, K⁺, Ca²⁺, Mg²⁺, Hg²⁺, Pb²⁺, Cr³⁺, Mn²⁺, Fe³⁺, Co²⁺, Ni²⁺, Cu²⁺, Zn²⁺, Cd²⁺, and Ag⁺, in HEPES-buffered EtOH/H₂O (4/1 v/v) media; pH: 7.4.

 Zn^{2+} , the absorbance increases along with the slight red-shifting of the peaks from 296 nm and 358 nm to 310 nm and 380 nm, respectively (Fig. 2a). The selectivity of F1 towards Zn^{2+} is shown in Fig. 2b.

Also, the emission properties of F1 were monitored via fluorescence spectroscopy. In the presence of Zn²⁺ the emission spectrum of F1 is selectively perturbed, while other cations, viz. Al^{3+} , K^+ , Ca^{2+} , Mg^{2+} , Hg^{2+} , Pb^{2+} , Cr^{3+} , Fe^{3+} , Mn^{2+} , Co^{2+} , Ni^{2+} , Cu^{2+} , Zn^{2+} , Cd^{2+} and Ag^+ , have no effect (EtOH/H₂O: 4/1 v/v; 10 mM HEPES buffer; pH: 7.4; λ_{ex} : 332 nm; Fig. 3a). The weak emission of F1 at 437 nm gradually increases to a 19-fold maximum upon the gradual addition of Zn²⁺ (Fig. 3b). Fluorescence behavior is susceptible to pH, and the Zn²⁺-induced emission intensity of F1 becomes highest at pH 7 (Fig. S4, ESI[†]). Hence, fluorescence experiments have been performed at physiological pH: 7.4. Competitive experiments indicate that there is no interference from common cations (Fig. S5, ESI⁺). Moreover, a sigmoidal plot of the emission intensity vs. $[Zn^{2+}]$ has a linear region (Fig. S6, ESI[†]). The Zn²⁺ detection limit of F1 is 42.23 nM (Fig. S7, ESI[†]). The binding constant calculated via the Benesi-Hildebrand equation is $7.71 \times 10^4 \text{ M}^{-1}$ (Fig. S8, ESI[†]). The Job plot indicates a 1:3 (molar ratio) interaction between F1 and Zn²⁺ (Fig. S9, ESI[†]).

Sensing mechanism

To obtain a plausible Zn(II) recognition mechanism in the presence of F1, a proper interpretation and understanding of



Scheme 2 A plausible fluorescence sensing mechanism.



Fig. 3 (a) Emission spectra of F1 (20 μ M) in the presence of different cations (2000 μ M), viz. Zn²⁺, K⁺, Ca²⁺, Mg²⁺, Hg²⁺, Pb²⁺, Cr³⁺, Mn²⁺, Fe³⁺, Co²⁺, Ni²⁺, Cu²⁺, Zn²⁺, Cd²⁺, and Ag⁺, in HEPES-buffered EtOH/H₂O (4/1 v/v) media; pH: 7.4; $\lambda_{ex} = 332$ nm. (b) Changes in the emission spectra of F1 (20 μ M) upon the gradual addition of Zn²⁺ (0–2000 μ M) in HEPES-buffered EtOH/H₂O (4/1 v/v) media; pH: 7.4; $\lambda_{ex} = 332$ nm.





Fig. 6 CV data from F1 and Z1

Table 1 Elucroscopes lifetime data of 71

Table I	luorescence	lifetime data			
Sample	B_1	B_2	t_1 (ns)	t_2 (ns)	χ^2
Z1	78.82	22.18	0.23	3.25	0.96

the spectroscopic properties of **F1** in the presence and absence of Zn(II) is very crucial. Thus, through interpreting the mass spectra, FTIR spectra and Job plot experiments, we can propose a "metal ion displacement protocol" as a plausible sensing mechanism (Scheme 2). The ν (N–N) band at 1140 cm⁻¹ for **F1** is shifted to 1227 cm⁻¹ in the case of **Z1**. The absence of the ν (O– H) band in the case of **Z1** also indicates the complexation of **L1** with Zn²⁺.

The selective recognition of an analyte by the probe using fluorescence spectroscopy is reflected by the fluorescence lifetime of the probe in the presence and absence of the analyte. Thus, Fig. 4 and Table 1 definitely suggest a significant interaction between **F1** and Zn^{2+} , leading to fluorescence enhancement. Non-fluorescent **F1** shows an average lifetime of 2.64 ns in the presence of Zn^{2+} .

It is very much expected that the two different systems possibly possess different morphologies, in terms of particle size, pattern, texture, *etc.* FESEM images indicate significant differences in morphological appearance between **F1** and **Z1** (Fig. 5). For **F1**, congested geometry is seen due to the large dinuclear Fe(m) complex, whereas changes detected in **Z1**, along



Fig. 5 FESEM images of F1 (a) and Z1 (b).

with a reduction in particle size, reveal the formation of a mononuclear complex.

Furthermore, the electrochemical behaviours of the two different metal complexes seem to be different in terms of reversibility, redox potential, and $E_{1/2}$. Fig. 6 indicates that the electrochemical properties of **Z1** are quite different from those of **F1**. The electrochemical behaviours of the **F1** and [**F1**–**Z**n²⁺] systems were monitored in anhydrous acetonitrile (ACN) media using tetra-butyl ammonium bromide (TBAB) as the supporting electrolyte at a scan rate of 100 mV s⁻¹, with [**F1**] = 1×10^{-4} M, [**F1**–**Z**n²⁺] = 1×10^{-4} M, and [TBAB] = 0.1 M at room temperature. The redox process associated with **F1** is quasi-reversible, having an $E_{1/2}$ value of +0.52 V, while the value is +0.31 V in the presence of Zn²⁺.

Kinetics of Fe(III) displacement from F1 by Zn²⁺

The proposed mechanism of Fe(III) replacement from F1 by Zn(II) is kinetically followed *via* the stopped-flow technique, where the change in absorbance of F1 at 380 nm is monitored as a function of the amount of added Zn^{2+} ions. The Zn^{2+} -assisted absorbance enhancement is attributed to the displacement of Fe³⁺ by Zn^{2+} within ~1.2 s. The rate constant of the pseudo first-order displacement reaction with respect to Zn^{2+} is 2.7315 × 10⁸ s⁻¹ (Fig. S10, ESI†).

Extraction of Zn²⁺ using F1

F1 is also employed for the enrichment of Zn^{2+} *via* solvent extraction from water to ethyl acetate. Three organic solvents, *viz.* EtOAc, CHCl₃ and CCl₄, are tested for the extraction of Zn^{2+} from aqueous solution. **F1** was equilibrated and partitioned at various concentrations (0.1×10^{-3} , 0.5×10^{-3} , 1.0×10^{-3} , 1.5



Scheme 3 The synthesis of cyclic carbonates from CO₂ and epoxides.



Scheme 4 The synthesis of a cyclic carbonate from \mbox{CO}_2 and epichlorohydrin.

 \times 10⁻³, and 2.0 \times 10⁻³ M) with known concentrations of Zn²⁺ for 1 h at room temperature. After the removal of the organic layer, the concentration of Zn²⁺ in the aqueous phase is measured using the presented method. The highest extraction efficiency is observed for EtOAc (95.47%), followed by chloroform (90.11%), and carbon tetrachloride (87.35%). The effects of various solvents on the extraction efficiency are presented in Fig. S11 (ESI).†

F1 catalyzed cyclic carbonate synthesis

The importance of the fixation of atmospheric CO_2 in the form of cyclic carbonates has already been discussed above. Hence, attempts have been made to study the catalytic activity of the new iron complex (F1) for the fixation of CO_2 as cyclic carbonates using different epoxides (Scheme 3).

The reaction of epichlorohydrin with CO_2 at 70 °C and under atmospheric pressure under solvent-free conditions, leading to a cyclic carbonate, is studied as a model reaction (Scheme 4). The reaction conditions have been optimized with respect to diverse parameters, *viz.*, temperature, reaction time and TBAB concentration, for maximum yield (Table S3, ESI[†]).

Among different trials, 5 mol% loading of co-catalyst (TBAB) led to the maximum yield of cyclic carbonate (Table S3, entry 9, ESI†). Above 5 mol% TBAB, no further improvement in product yield is observed. Traces of the desired product formed without TBAB under identical reaction conditions. Interestingly, without the **F1** catalyst, TBAB alone fails to improve the yield of the product. A synergistic effect between **F1** and TBAB accelerated the product yield. It is noted that the reaction temperature plays a crucial role in determining the yield of the product (Table S3, ESI†). The highest yield is observed at 70 °C. Thus, the optimum parameters that lead to the highest yield of product are 5 mol% TBAB (co-catalyst), 70 °C (temp.), 12 h (reaction time), 15 mg of **F1** (catalyst load), and 1 atm. CO₂ (pressure) (Table S3, ESI†).

To verify the efficiency and performance of the **F1** catalyst, the fixation of CO_2 is tested with some substituted epoxides, like 1,2-epoxy-3-phenoxy propane epichlorohydrin and styrene oxide (Table 2). Of these, epichlorohydrin provided a better yield, probably due to the electron withdrawing effects of the substituent.

The catalytic efficiency of **F1**, the methodology, and the reaction conditions are highly competitive when compared with other reported iron catalysts (Table 3).

Plausible mechanism for epoxide formation

Based on experimental findings and the associated literature,⁴⁰⁻⁴⁴ a plausible mechanism for fixing CO₂ using an epoxide, leading to an organic cyclic carbonate, is portrayed in Scheme 5. At first, the active site of **F1** coordinates with the Odonor of the epoxide, followed by nucleophilic attack by Br⁻ (from TBAB) at the C-centre attached to the epoxide oxygen; this opens the ring to form an α -halo-alkoxide. Then, the O-donor attached to the Fe centre interacts with the C-centre of CO₂ to produce an alkyl cyclic carbonate along with the regeneration of the catalyst **F1**.

Catecholase-like activity

The catecholase-like activity of F1 (1×10^{-4} M) is monitored *via* time-dependent absorption spectroscopy through its reaction

Table 2 The synthesis of cyclic carbonates using the F1 catalyst ^a								
Entry	Epoxide	Product	Yield ^{b} (%)	Conversion ^c (%)	TON^d	TOF ^e		
1			87	90	32 142	2678		
2	CI	CI	97	99	35 357	2946		
3			91	94	33 571	2797		

^{*a*} Reaction conditions: epoxide (5 mmol), F1 catalyst (15 mg), TBAB (5 mol%), 70 °C, CO₂ (1 atm.), 12 h. ^{*b*} GC yield of cyclic carbonate. ^{*c*} Conversion is determined *via* GC. ^{*d*} Turnover number (TON) = (% conversion × mmol of substrate used)/(mmol of catalyst used). ^{*e*} Turnover frequency (TOF) = TON/time of reaction (h).

Table 3 Comparison of F1 with some reported iron catalysts

Reaction	Catalyst	Reaction conditions	Yield (%)	Ref.
Cyclic carbonate synthesis	Iron(m)aminobis(phenolate) complexes	Epichlorohydrin (5.0×10^{-2} mol), PPNCl (5.00×10^{-5} mol), CO ₂ (20 bar), 4 h, 100 $^{\circ}$ C	99	40
	Pyridine-bridged bispincer- type Fe(11) complex	Epichlorohydrin (10 mmol), TBAB (97.6 mg, 0.29 mmol), 25 °C, 0.5 MPa CO ₂ , 24 h.	91	41
	Fe(III) thio ether– triphenolate complex	Epichlorohydrin (7.0×10^{-2} mol), TBAB (7.0×10^{-5} mol), CO ₂ (2 MPa); 100 °C, 6 h.	78	42
	F1	Epichlorohydrin (5 mmol), TBAB (5 mol%), 70 °C, CO_2 (1 atm.), 12 h	97	Present study

with the model compound 3,5-DTBC (1 \times 10 $^{-2}$ M, 100 equiv.) in methanol under aerobic conditions at room temperature.

As saturation kinetics are observed at a high substrate concentration, the Michaelis–Menten equation and a Lineweaver–Burk plot (Fig. S12a, ESI†) are used to elucidate the kinetic parameters. The initial first order rate constant, $V(s^{-1})$, is calculated from the slope of the plot $\log[A_{\infty}/(A_{\infty\alpha} - A_t)]$ vs. time. Kinetic parameters like binding constant (K_M), maximum velocity (V_{max}), and rate constant (turnover number, k_{cat}) are determined from the linear plot of 1/V vs. 1/[S] (Lineweaver–Burk plot) as per the Michaelis–Menten enzymatic kinetics approach. From the Lineweaver–Burk plot, $1/V = \{K_M/V_{max}\} \times \{1/[S]\} + 1/V_{max}$, and the values for V_{max} , k_M , and k_{cat} are found to be 5.49 $\times 10^{-2}$ M s⁻¹, 1.23 $\times 10^{-4}$ M, and 4.42 $\times 10^{3}$ h⁻¹.



Scheme 5 The probable mechanistic pathway to organic cyclic carbonate formation.

Scheme S1 (ESI[†]) shows the well-known catalytic catechol oxidation protocol, supported by the ESI-MS(+) spectrum of a mixture of **F1** and 3,5-DTBC (1 : 50 molar ratio) in aqueous methanol (Fig. S12b, ESI[†]). The two predominant peaks at m/z values of 237.19 (calcd: 238.16) and 1135.13 (calcd: 1135.29) are attributed to $[3,5-DTBQ + H_2O]^+$ and $[F1 + 2H_2O + Li]^+$, respectively.

The electrochemical transformation of catechol to quinone has also been monitored in the presence of **F1** (Fig. S12c, ESI[†]). The electrochemical behaviours of the **F1** and [**F1**-catechol] systems have been investigated in anhydrous acetonitrile media in the presence of tetra-butyl ammonium bromide (TBAB) as a supporting electrolyte at a scan rate of 100 mV s⁻¹, where [**F1**] = 1×10^{-4} M, [**F1**-catechol] = 1×10^{-4} M, and [TBAB] = 0.1 M, at room temperature. Cyclic voltammograms (CV) of the **F1** and [**F1**-catechol] systems appear different, with associated $E_{1/2}$ values of +0.52 V and +0.47 V, respectively.

Conclusion

A new dinuclear Fe(m) complex (F1) of a symmetrical azine derivative, structurally characterized *via* single-crystal X-ray analysis, is employed as a catalyst for the fixation of atmospheric CO₂ using epoxides, leading to organic cyclic carbonates with 97% yield. In addition, F1 selectively recognizes Zn^{2+} at concentrations as low as 42.23 nM through TURN-ON fluorescence *via* a metal displacement protocol. Moreover, F1 efficiently (95.47%) extracts Zn^{2+} from aqueous media to ethyl acetate. The catecholase-like activity of F1 is demonstrated, providing a k_{cat} value of 4.42×10^3 h⁻¹.

Conflicts of interest

There are no conflicts to declare.

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