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# Effect of nickel ion doping in MnO<sub>2</sub>/reduced graphene oxide nanocomposites for lithium adsorption and recovery from aqueous media

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Novel and effective reduced graphene oxide–nickel (Ni) doped manganese oxide (RGO/Ni-MnO<sub>2</sub>) adsorbents were fabricated *via* a hydrothermal approach. The reduction of graphite to graphene oxide (GO), formation of  $\alpha$ -MnO<sub>2</sub>, and decoration of Ni-MnO<sub>2</sub> onto the surface of reduced graphene oxide (RGO) were independently carried out by a hydrothermal technique. The physical and morphological properties of the as-synthesized adsorbents were analyzed. Batch adsorption experiments were performed to identify the lithium uptake capacities of adsorbents. The optimized parameters for Li<sup>+</sup> adsorption investigated were pH = 12, dose loading = 0.1 g, Li<sup>+</sup> initial concentration = 50 mg L<sup>-1</sup>, in 10 h at 25 °C. It is noticeable that the highest adsorption of Li<sup>+</sup> at optimized parameters are in the following order: RGO/Ni3-MnO<sub>2</sub> (63 mg g<sup>-1</sup>) > RGO/Ni2-MnO<sub>2</sub> (56 mg g<sup>-1</sup>) > RGO/Ni1-MnO<sub>2</sub> (52 mg g<sup>-1</sup>). A Kinetic study revealed that the experimental data were best designated pseudo-second order for each adsorbent. Li<sup>+</sup> desorption experiments were performed using HCl as an extracting agent. Furthermore, all adsorbents exhibit efficient regeneration ability and to some extent satisfying selectivity for Li<sup>+</sup> recovery. Briefly, it can be concluded that among the fabricated adsorbents, the RGO/Ni3-MnO<sub>2</sub> exhibited the greatest potential for Li<sup>+</sup> uptake from aqueous solutions as compared to others.

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## 1. Introduction

Lithium is the most valuable metal in this century because of its wide range of applications in various areas including, energy storage Li batteries, nuclear fusion fuels, lubricants, alloys, refrigerants, drug delivery systems, and aerospace materials.<sup>1–6</sup> Therefore, the global demand for Li resources has been rising.<sup>7–8</sup> Li can be extracted from two main sources. The first is mineral alloys; castorite and spodumene alloys, while the second includes salt lakes, mines, sea-water, and geothermal liquids.<sup>9</sup> Among these resources, seawater is considered as an essential Li source to fulfil the global need.<sup>10</sup> Even though seawater possesses a very low Li concentration (0.17 mg L<sup>-1</sup>), the great area of the earth covered with seawater makes it an effective source for Li recovery.<sup>11</sup> Therefore, many scientists reported a number of methods for obtaining Li from seawater including chemical-precipitation,<sup>12</sup> nano-filtration, resin-ion exchange, membrane technology,<sup>13</sup> liquid extraction,<sup>14</sup> and adsorption.<sup>15,16</sup> Except for adsorption, the other techniques all possess some demerits, including high energy consumption, production of

hazardous waste, low effectiveness, and high costs. On the other hand, adsorption seems an efficient, inexpensive, and eco-friendly method.<sup>17</sup>

Currently, manganese oxide (Mn<sub>x</sub>O<sub>y</sub>) ion sieves have gained significant focus for metal adsorption (Pb, Hg, Cu, and Sr)<sup>18–20</sup> because of their favorable characteristics, such as eco-friendliness, low cost, and high stability with tunable morphology.<sup>21,22</sup> But fine size prohibited their utilization in various fields due to slow aqueous-solid partitioning. Thus, it becomes necessary to deposit Mn<sub>x</sub>O<sub>y</sub> on larger molecular components to enhance its experimental performance.<sup>23</sup> In this regard, graphene oxide (GO) has been gaining more attention because it has a greater specific surface area while exhibiting excellent adsorption activity along with identical electrical characteristics.<sup>24,25</sup> Consequently, it recognized that some distinctive challenging cations (for example, Sr<sup>2+</sup>, Co<sup>2+</sup>, Cu<sup>2+</sup>, Li<sup>+</sup>, and Hg<sup>2+</sup>) can possibly interfere with the adsorption activity of GO. It has been reported that doping of nickel contents improved MnO<sub>2</sub> applications, like increasing the rate of reduction reaction as a nanocatalyst, cycling of Li ion batteries.<sup>26–28</sup> Literature survey illustrated that, influence of Ni contents on MnO<sub>2</sub> with expanded surface area for Li recovery has not been studied yet.

Therefore, we fabricated Ni doped  $\alpha$ -MnO<sub>2</sub> nanorods above a reduced graphene oxide sheet by a facile hydrothermal approach and focused on the loading of Ni contents on

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adsorbents for Li adsorption and recovery performance in liquid media. In this research work, we also aimed to characterize the synthesized nanocomposites, to optimized various adsorption parameters (pH, adsorbent loading, contact time, and  $\text{Li}^+$  initial concentration), to analyze the adsorption data by using various models. Furthermore, the desorption, regeneration ability and influence of coexisting ions were also taken into consideration for selective ion adsorption.

## 2. Experimental

### 2.1. Materials

$\text{KMnO}_4$  and ethanol (99.9%) were purchased from OCI Company Ltd., South Korea. Nickel(II) acetate tetrahydrate ( $\text{Ni}(\text{CH}_3\text{COO})_2 \cdot 4\text{H}_2\text{O}$ ) and lithium atomic adsorption standard (AA) were obtained from Sigma Aldrich.  $\text{H}_2\text{O}_2$  (34.5% pure) was purchased from Samchun pure chemicals, Korea. Graphite powder, HCl, NaOH,  $\text{HNO}_3$  (60% purity), and  $\text{H}_2\text{SO}_4$  (98% purity) were bought from Daejung KOSDAQ, South Korea. All chemicals were of analytical grade and used as received without undergoing any further purification procedures. Distilled water (DW) was used throughout the experiments.

### 2.2. Synthetic protocol

RGO was prepared by a modified Hummer's method. At first, graphite powder (2 g) was added in  $\text{H}_2\text{SO}_4$  (120 mL) at  $0^\circ\text{C}$ . After that, 8.75 g of  $\text{KMnO}_4$  was added and under stirred for 4 h at  $30^\circ\text{C}$ . Then, mixture was highly diluted with DW and 35%  $\text{H}_2\text{O}_2$  at  $0^\circ\text{C}$  until bubbles formation stopped. The black precipitates were vacuum filtered, washed with DW.

$\alpha\text{-MnO}_2$  fabricated *via* hydrothermal technique. Firstly, 4.5 mL of 60%  $\text{HNO}_3$  was dropwise added in  $\text{KMnO}_4$  (0.043 M) 300 mL and allowed to stirrer for 4 h. Then solution was autoclaved for 15 h at  $125^\circ\text{C}$ . The formed product was vacuum filtered, rinsed with water, ethanol and dried.

RGO/Ni-MnO<sub>2</sub> adsorbents synthesized as follows, RGO powder (0.6 g) ultra-sonicated for 120 min with DW. Meanwhile, Ni ( $\text{CH}_3\text{COO}$ )<sub>2</sub> · 4H<sub>2</sub>O (0.01 g) and  $\alpha\text{-MnO}_2$  (0.2 g) in 60 mL DW stirred for 3 h individually with different nickel salt ratios 1 : 1, 1 : 2, 1 : 3 respectively.<sup>29</sup> After that, prepared solutions were separately mixed with sonicated RGO solution and autoclaved at  $120^\circ\text{C}$  for 5 h. The adsorbents were filtered and rinsed. A schematic diagram of the synthetic approach is shown in Fig. 1. The RGO/Ni-MnO<sub>2</sub> adsorbents prepared with different nickel salt ratios from 1–3 denoted as RGO/Ni1-MnO<sub>2</sub>, RGO/Ni2-MnO<sub>2</sub>, and RGO/Ni3-MnO<sub>2</sub>.

### 2.3. Characterization

The characterization of the fabricated adsorbent was carried out using following analytical techniques. The morphological and elemental properties were analyzed by scanning electron microscopy (FE-SEM; Model S-4300 SE, Hitachi Co., Ltd.), EDAX and field emission transmission electron microscopy (FE-TEM; JEM-2100F). X-ray diffraction (XRD) was used to characterize the crystal phases (D2 Phaser, Bruker Co., USA). The analysis of functional groups was performed by using Fourier transform

infrared spectroscopy (FTIR) using an FTIR vacuum-vertex/80 V spectrometer. Textural properties of adsorbents including specific surface area and pore volume were analyzed *via* a BELSORP-mini (BEL Co. Japan Inc.). The  $\text{Li}^+$  ion concentrations in aqueous solutions were measured *via* atomic adsorption spectrometry (AAS) using a Model PerkinElmer, A-Analyst 400, (USA).

### 2.4. Adsorption/desorption experiment

Batch experimental studies were conducted to analyze the adsorption ability of materials. Concisely, the  $\text{Li}^+$  solution (250 mL) allow to stir (220 rpm) at  $25^\circ\text{C}$ , and adsorbent-sorbate solution centrifuged to analyze suspension by ASS. The influence of pH, dose loading, initial  $\text{Li}^+$  concentration, and contact time on adsorbents were studied in following ranges; 4–12, 0.1–0.25 g, 10–50 ppm and 4–24 h. The  $\text{Li}^+$  loaded adsorbents were stored for recovery experiments. The adsorption capacity of  $\text{Li}^+$  was determined by using the following eqn (1),

$$q_e = C_o - C_e \times \frac{V}{M} \quad (1)$$

Where  $C_o$  and  $C_e$  are the concentration ( $\text{mg L}^{-1}$ ) of adsorbate before and after adsorption experiment.  $V$  is volume (L), and  $M$  is adsorbent mass (g). The kinetic studies were performed at specified time intervals, the quantity of  $\text{Li}^+$  uptake at equilibrium and different time ( $q_t$ ) ( $\text{mg g}^{-1}$ ) were obtained using eqn (2),

$$q_t = C_o - C_t \times \frac{V}{M} \quad (2)$$

In the above equations,  $C_t$  ( $\text{mg L}^{-1}$ ) is  $\text{Li}^+$  concentration at any time ( $t$ ).

For  $\text{Li}^+$  recovery, DW and HCl can be used as an eluting agent. In this analysis, 0.1 g of  $\text{Li}^+$  loaded adsorbents were immersed in 250 mL of eluting agent and magnetically stirred for 12 h at  $25^\circ\text{C}$ .<sup>30</sup> Desorption percentage can be calculated by the formula:

$$\text{Desorption}\% = \frac{\text{amount of desorbed metal ion by eluting agent}}{\text{primary adsorbed amount of ions above adsorbents}} \times 100 \quad (3)$$

In order to check the regeneration ability and stability of prepared adsorbents the adsorption-desorption experiment was repeated with three continuous cycles. All cycles were performed with following optimized experimental condition; pH (12), adsorbent dose (0.1 g),  $\text{Li}^+$  initial concentration (50 ppm) for 10 h at ambient temperature. The adsorbents were recovered with 0.5 M HCl for 12 h and also rinsed with distilled water to eliminate impurities.

## 3. Results and discussion

### 3.1. Morphological properties

Fig. 2 shows the HR-SEM and FE-TEM micrographs of prepared  $\alpha\text{-MnO}_2$ , Ni-MnO<sub>2</sub> and RGO based samples. Fig. 2(a) depicts the



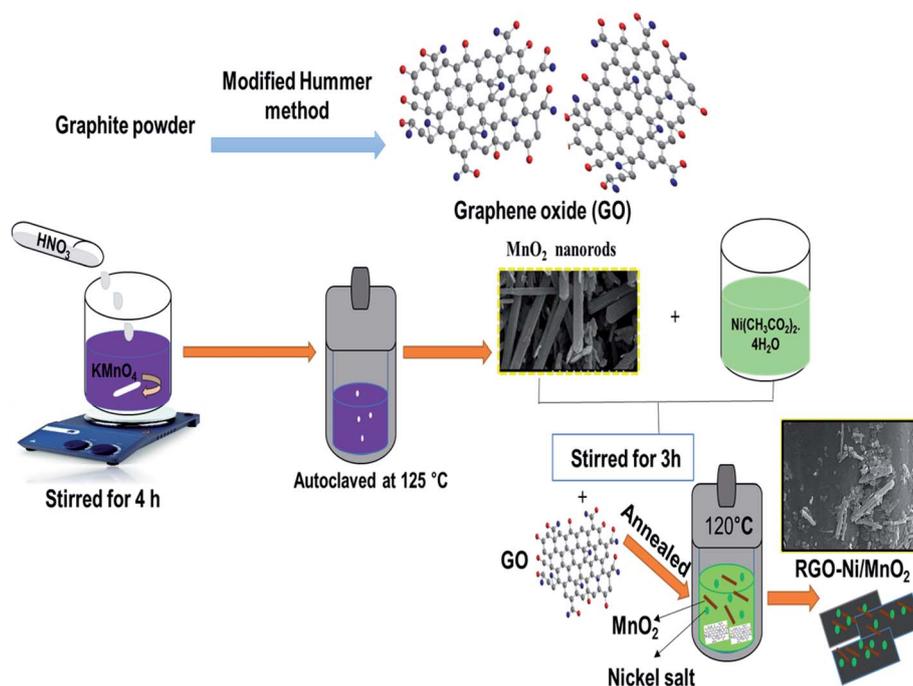


Fig. 1 General schematic layout of  $\alpha$ -MnO<sub>2</sub> nanorods and RGO/Ni-MnO<sub>2</sub> adsorbents fabrication by hydrothermal approach.

formation of rod like morphology of  $\alpha$ -MnO<sub>2</sub> nanoparticles. Fig. 2(b) shows that incorporation of nickel contents having no influence on its morphology, but it seems like coating of nanorods. The adsorbents in Fig. 2(c–g) indicates the random impregnation of Ni-MnO<sub>2</sub> nanorods above sheets of RGO at different resolutions.<sup>31</sup> Furthermore, Fig. 2(e) represents the interaction between the RGO and the Ni-MnO<sub>2</sub> nanorods at 500 nm. However, FE-TEM micrographs in Fig. 2(h–l) of as-synthesized samples show similar morphologies as the SEM within size ranges of 20, 100 and 200 nm. The crystalline nature of pure, Ni loaded, and RGO based as-synthesized adsorbents was studied through powder XRD and shown in Fig. 3(a), the diffraction peaks located at 12.24°, 17.45°, 28.27°, 36.89°, 42.31°, 49.51°, and 56.53° depict the clear formation of  $\alpha$ -MnO<sub>2</sub> nanorods. XRD pattern of nickel loaded adsorbents remain similar, no peak of nickel contents observed because it presents in very small amount but slight increase in peak intensity of nickel loaded  $\alpha$ -MnO<sub>2</sub> nanocomposites has been observed due to raise in size of nanorods and increase in crystallinity, gradually nickel doping assists the fabrication of  $\alpha$ -MnO<sub>2</sub> nanorods.<sup>32</sup> Another reason of increment in intensity consider that the tunnel spaces of  $\alpha$ -MnO<sub>2</sub> can allow to insert ions with radius lower than 1.5 Å, as the radius of nickel ion (Ni<sup>2+</sup>) is 0.72 Å that promotes the insertion of Ni<sup>2+</sup> ions within the tunnel spaces in order to enhance the stability of  $\alpha$ -MnO<sub>2</sub>.<sup>27</sup>

FTIR analysis scanned in the range of 4000 to 450 cm<sup>-1</sup> was used to confirm the presence of functional groups. The FTIR spectra in Fig. 3(b) exhibit broad peaks at 2100 cm<sup>-1</sup> and 1550 cm<sup>-1</sup>, suggesting the existence of C≡C and C=C bonds between carbons. However, the broad peak at 685 cm<sup>-1</sup> was attributed to Mn–O–Mn stretching,<sup>33</sup> while the peaks at

489 cm<sup>-1</sup> and 597 cm<sup>-1</sup> were attributed the existence of Mn–O–C linkages, which confirms the presence of interactions between the MnO<sub>2</sub> nanorods and RGO.<sup>34</sup> Fig. 4(b–d) elaborates the elemental analysis of fabricated adsorbents obtained by EDAX, which confirmed the doping of nickel contents in adsorbents.

### 3.2. Textural and surface properties

The textural characteristics of the synthesized materials were analyzed *via* nitrogen gas adsorption/desorption isotherms at 77 K using a Belsorp Max-BEL apparatus. First, the samples were degassed for 6 h at 423 K by a vacuum system that permitted outgassing at a pressure of 10<sup>-5</sup> torr. Generally, the specific surface area (SSA) of adsorbents is an essential contributor to its adsorption performance and can be calculated by the Brunauer–Emmett–Teller (BET) equation.<sup>35</sup> The measured SSA of RGO,  $\alpha$ -MnO<sub>2</sub>, Ni-MnO<sub>2</sub>, and RGO/Ni-MnO<sub>2</sub> adsorbents were 184, 57, 71, 105, 111, and 112, m<sup>2</sup> g<sup>-1</sup>, as shown in Fig. 4(a) and Table 1, which indicates that RGO significantly increased the SSA of Ni-MnO<sub>2</sub> materials. Briefly, it can conclude that slight increase in specific surface area has been observed due to raise in nickel contents.<sup>36</sup> The synthesized samples were also characterized by XPS measurement for determining the chemical compositions. Fig. 5(a) illustrated the survey scan of XPS spectra obtained from  $\alpha$ -MnO<sub>2</sub> and RGO. From  $\alpha$ -MnO<sub>2</sub> spectrum two signals with intense intensities were observed attributed to Mn 2p, and O 1s which confirms the absence of others impurities. Fig. 5(b) shows the XPS survey spectrum of all synthesized RGO/Ni1-MnO<sub>2</sub> adsorbents, the clear signals shown at 852.81 eV, 642.23 eV, 529.91 eV and 283.72 eV related to Ni 2p, Mn 2p, O 1s and C 1s correspondingly. It depicts that the



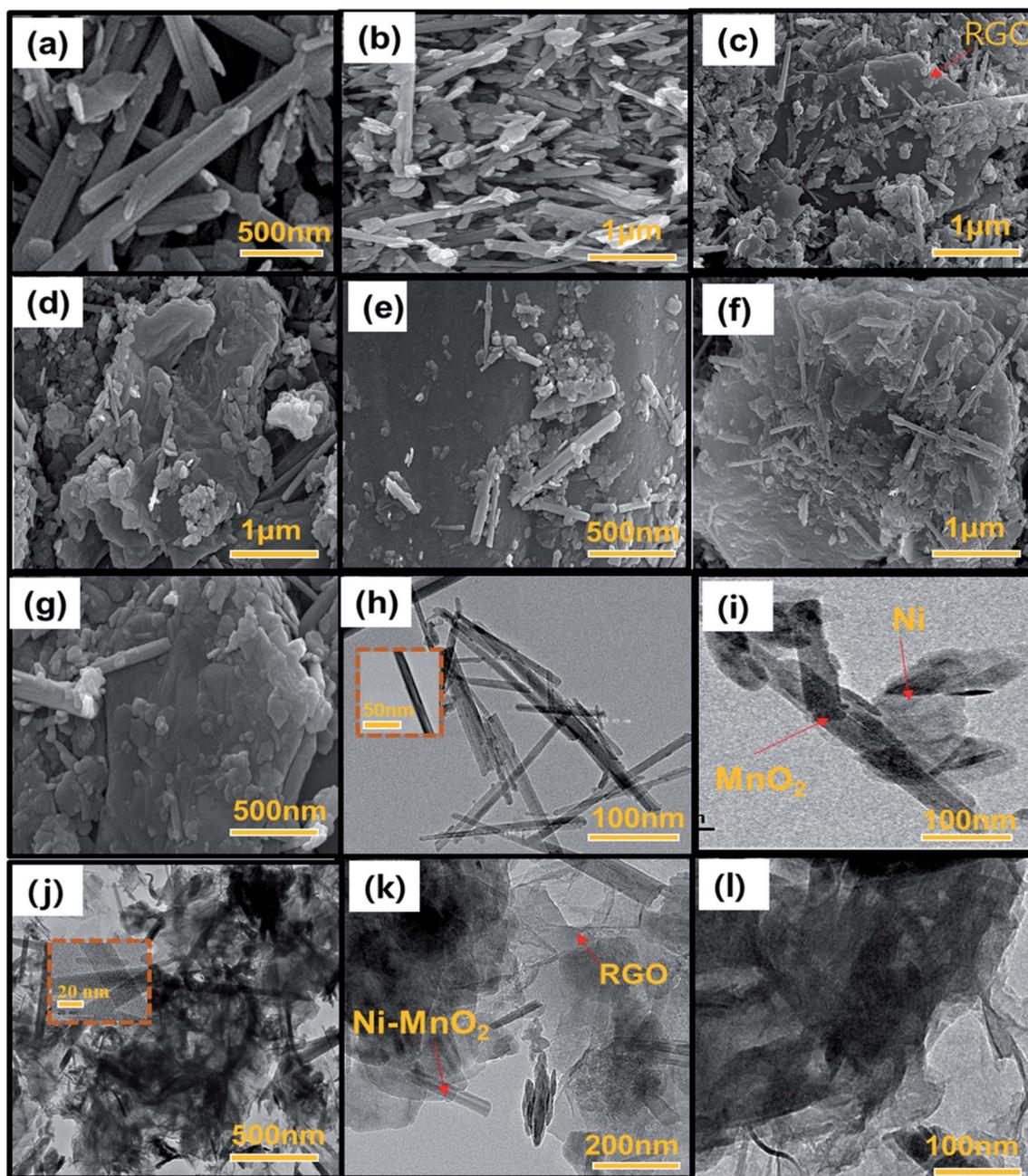


Fig. 2 HR-SEM images of (a)  $\alpha$ -MnO<sub>2</sub>, (b) Ni-MnO<sub>2</sub>, (c) RGO/Ni<sub>1</sub>-MnO<sub>2</sub>, (d and e) RGO/Ni<sub>2</sub>-MnO<sub>2</sub>, and (f and g) RGO/Ni<sub>3</sub>-MnO<sub>2</sub>. FE-TEM images of (h)  $\alpha$ -MnO<sub>2</sub> (i) Ni-MnO<sub>2</sub> and (j–l) RGO/Ni<sub>1</sub>-MnO<sub>2</sub> at different resolutions.

prepared nanocomposites contained four main elements (Ni, Mn, O, and C). Furthermore, in the C 1s XPS spectrum of RGO/Ni<sub>3</sub>-MnO<sub>2</sub> (shown in Fig. 5(c)) three characteristic peaks obtained at 284.37, 286.04, and 287.99 eV attributed to C=C or C–C, epoxy or C–O and C=O or C–OOH successively.<sup>37</sup> The intensity at 284.37 eV is more intense than other signals which shows that C in the adsorbents exist in the form of C=C or C–C bond. Fig. 5(d) demonstrates the O 1s spectrum of RGO/Ni<sub>3</sub>-MnO<sub>2</sub> and deconvoluted into of two main peaks at 529.33 eV and 531.11 eV relates to Mn–O and Mn–OH functionalities.<sup>38</sup> Furthermore, as shown in Fig. 5(e) Ni 2p spectrum

is good-fitted with two spin orbit doublets at 855.63 eV, 873.19 eV and shake-up satellite at 861.23 and 878.23 eV respectively. The Ni 2p<sub>3/2</sub> peak shows three various chemical environments for Ni.<sup>39</sup> The weak peak at 852.09 eV indicates the metallic nickel,<sup>40</sup> while the intense peak at 855.6 eV attributed to the Ni<sup>2+</sup>–Mn or Ni<sup>2+</sup>–OH. Likewise, for Ni 2p<sub>1/2</sub> the peak at 873.1 eV attributed to Ni<sup>2+</sup>–Mn similar to previous literature.<sup>41</sup> As mentioned in Fig. 5(f) the Mn 2p spectrum, the two spin orbit doublets appeared at 642.2 eV and 653.7 eV corresponds to Mn 2p<sub>3/2</sub> and Mn 2p<sub>1/2</sub> respectively.<sup>42</sup> On the basis of Mn 2p an O 1s



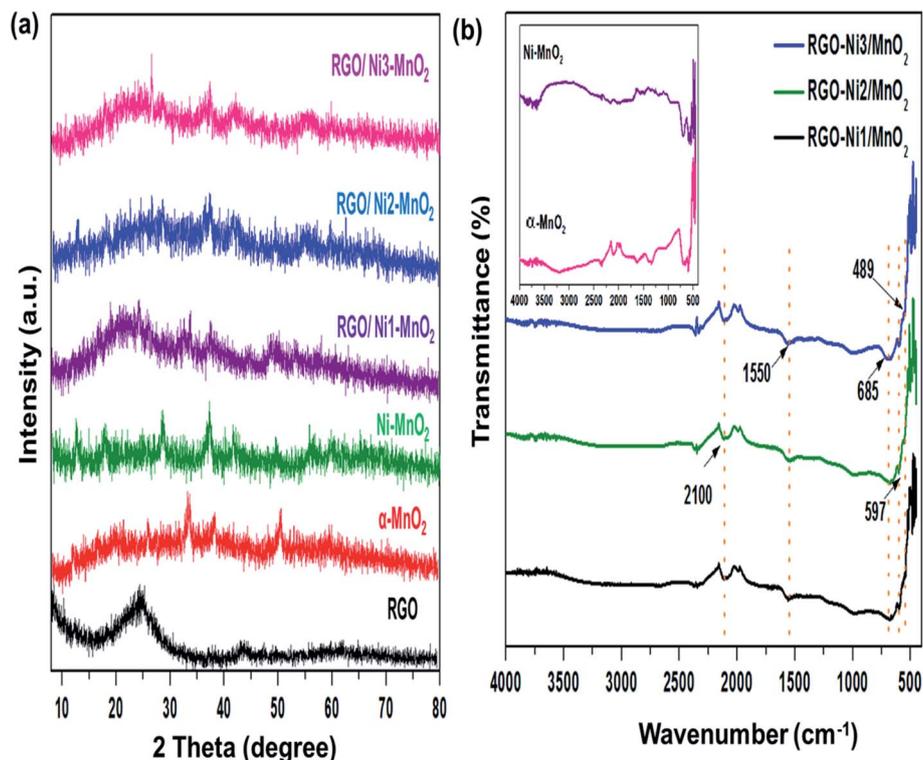


Fig. 3 (a) XRD patterns of RGO,  $\alpha$ -MnO<sub>2</sub>, Ni-MnO<sub>2</sub>, and RGO/Ni1-MnO<sub>2</sub>, RGO/Ni2-MnO<sub>2</sub>, and RGO/Ni3-MnO<sub>2</sub> and (b) FTIR spectra of  $\alpha$ -MnO<sub>2</sub>, Ni-MnO<sub>2</sub> and all adsorbents.

XPS spectrum, it can be resultant that the Mn within the nanocomposites existed in the form of MnO<sub>2</sub>.

### 3.3. Analysis of adsorption parameters and kinetics

The pH of solution influences the solubility of metals by charge variations above sorbent's surface. Fig. 6(a) indicates that the adsorption tendency declined in an acidic environment (pH = 4). When the solution's basicity was raised, a tremendous increase in Li<sup>+</sup> sorption rate was observed; the reason for this may be that under acidic conditions, the surface of adsorbents becomes highly protonated, leading to increased electrostatic repulsion between Li<sup>+</sup> and materials. This condition is also responsible for increased competition between Li<sup>+</sup> and H<sup>+</sup> ions for the active sites above the adsorbents. In contrast, under basic conditions, the concentration of OH<sup>-</sup> ions is increased and this makes the binding sites of the sorbents more negatively charged, which enhances the forces of attraction between metal ions and the negatively charged adsorbents binding sites.<sup>43</sup> The maximum adsorption  $q_e$  values recorded at pH 12 for RGO/Ni1-MnO<sub>2</sub>, RGO/Ni2-MnO<sub>2</sub>, and RGO/Ni3-MnO<sub>2</sub>, were 16.23 mg g<sup>-1</sup>, 18.73 mg g<sup>-1</sup>, and 22.23 mg g<sup>-1</sup> respectively. The influence of adsorbent loading was evaluated at optimized pH. Fig. 6(b) depicts that the Li<sup>+</sup> removal rate was rapidly decreased by raising dose from 0.1 to 0.25 g, because the active sites become blocked due to crowding and overlapping of adsorbent particles.<sup>44,45</sup> The adsorption of Li<sup>+</sup> proposed to be occurred by ion exchange mechanism between nickel and Li<sup>+</sup> above RGO surface as shown in Fig. 1. By optimizing dose loading

parameter, the  $q_e$  values raised to 19.35 mg g<sup>-1</sup>, 25.33 mg g<sup>-1</sup>, and 27.73 mg g<sup>-1</sup> for RGO/Ni1-MnO<sub>2</sub>, RGO/Ni2-MnO<sub>2</sub>, and RGO/Ni3-MnO<sub>2</sub> with minimum dose (0.1 g) of adsorbents.

The adsorption kinetics were studied as a function of contact time to determine the rate of adsorption. From Fig. 6(c), rapid adsorption was obtained from the initial stage with various equilibrium times. The products reached at its maximum equilibrium after 10 h with further raised in  $q_e$  values (46 mg g<sup>-1</sup>, 53 mg g<sup>-1</sup> and 58 mg g<sup>-1</sup>). It can be concluded that the Li<sup>+</sup> uptake rate was much higher in the first few hours and slowed down after attaining equilibrium, because initially, a greater number of binding sites were available, which enhanced the adsorption capacity.<sup>46,47</sup> Therefore, the pseudo-first order, pseudo-second order, and intraparticle diffusion models can applied to experimental values to understand the kinetics of the adsorption mechanism. The pseudo-first order and second order model equations are as follows,<sup>48,49</sup>

$$\ln(q_e - q_t) = \ln q_e - k_1 t \quad (4)$$

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \quad (5)$$

In the above equations,  $k_1$  and  $k_2$  represent the rate constants of 1<sup>st</sup> and 2<sup>nd</sup> order (g mg<sup>-1</sup> min<sup>-1</sup>), which are obtained from the slope, and  $q_e$  and  $q_t$  are the adsorbed concentration of Li<sup>+</sup> above the adsorbents at equilibrium and at a specific time ( $t$ ). A graph can be plotted of  $\ln(q_e - q_t)$  versus time ( $t$ ). However,



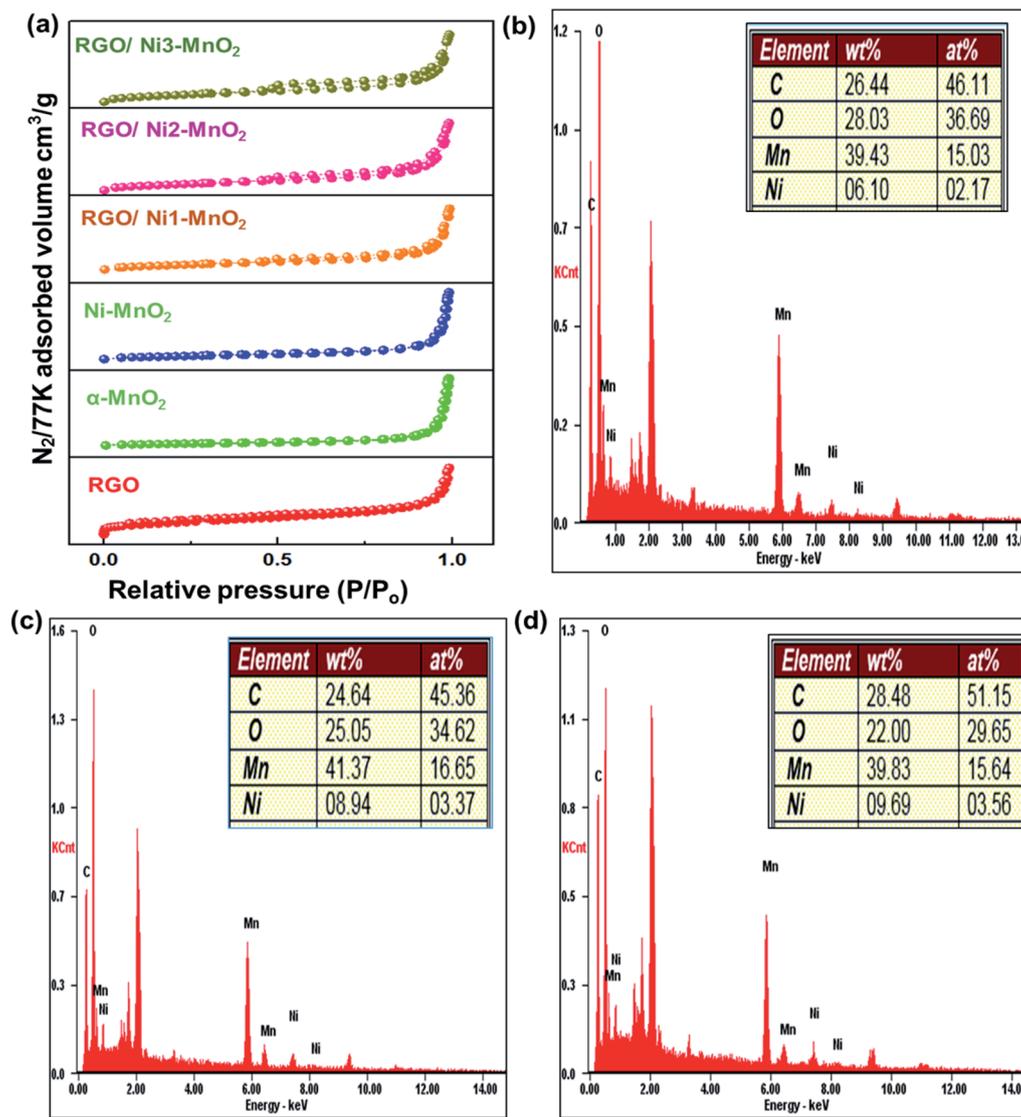


Fig. 4 (a) N<sub>2</sub> adsorption/desorption isotherms of fabricated RGO, nanoparticles and adsorbents, EDAX elemental analysis of (b) RGO/Ni<sub>1</sub>-MnO<sub>2</sub>, (c) RGO/Ni<sub>2</sub>-MnO<sub>2</sub>, and (d) RGO/Ni<sub>3</sub>-MnO<sub>2</sub>.

diffusion phenomena cannot be accurately modeled by these two equations. Therefore, the intra-particle diffusion model (IPD) was applied to understand the kinetics of the diffusion mechanism.<sup>50</sup>

Table 1 Textural properties of prepared nanoparticles and adsorbents

Specimens	Specific surface area (m <sup>2</sup> g <sup>-1</sup> )	Total pore volume (cm <sup>3</sup> g <sup>-1</sup> )	Mean pore diameter (nm)
α-MnO <sub>2</sub>	57	0.45	31.6
Ni-MnO <sub>2</sub>	71	0.46	26.0
RGO	184	0.43	9.3
RGO/Ni <sub>1</sub> -MnO <sub>2</sub>	105	0.30	11.5
RGO/Ni <sub>2</sub> -MnO <sub>2</sub>	112	0.31	11.3
RGO/Ni <sub>3</sub> -MnO <sub>2</sub>	115	0.26	9.3

$$q_t = K_{pi}t^{1/2} \quad (6)$$

Here,  $C_i$  is the intercept and  $K_{pi}$  is the intra-particle diffusion rate constant (mg g<sup>-1</sup> min<sup>-0.5</sup>). The experimental values were fitted to eqn (4)–(6) to yield the linear graphs depicted in Fig. 7 (a–c), and the calculated values of parameters are listed in Table 2. The correlation coefficient ( $R^2$ ) values of the pseudo second order kinetic model for all adsorbents were very close to 1, in contrast to those of the pseudo first order and intra-particle diffusion models. Besides this, the calculated  $q_t$  values of the pseudo second order models were found to resemble the experimental  $q_t$  values, which suggests that the pseudo-second order model precisely describes the kinetics of sorption process. The rate determining stage of adsorption was overcome by chemical sorption phenomena.<sup>51</sup>

The influence of different Li<sup>+</sup> initial concentration (10, 20, 30, 40, 50 ppm) on sorption tendency was studied, as



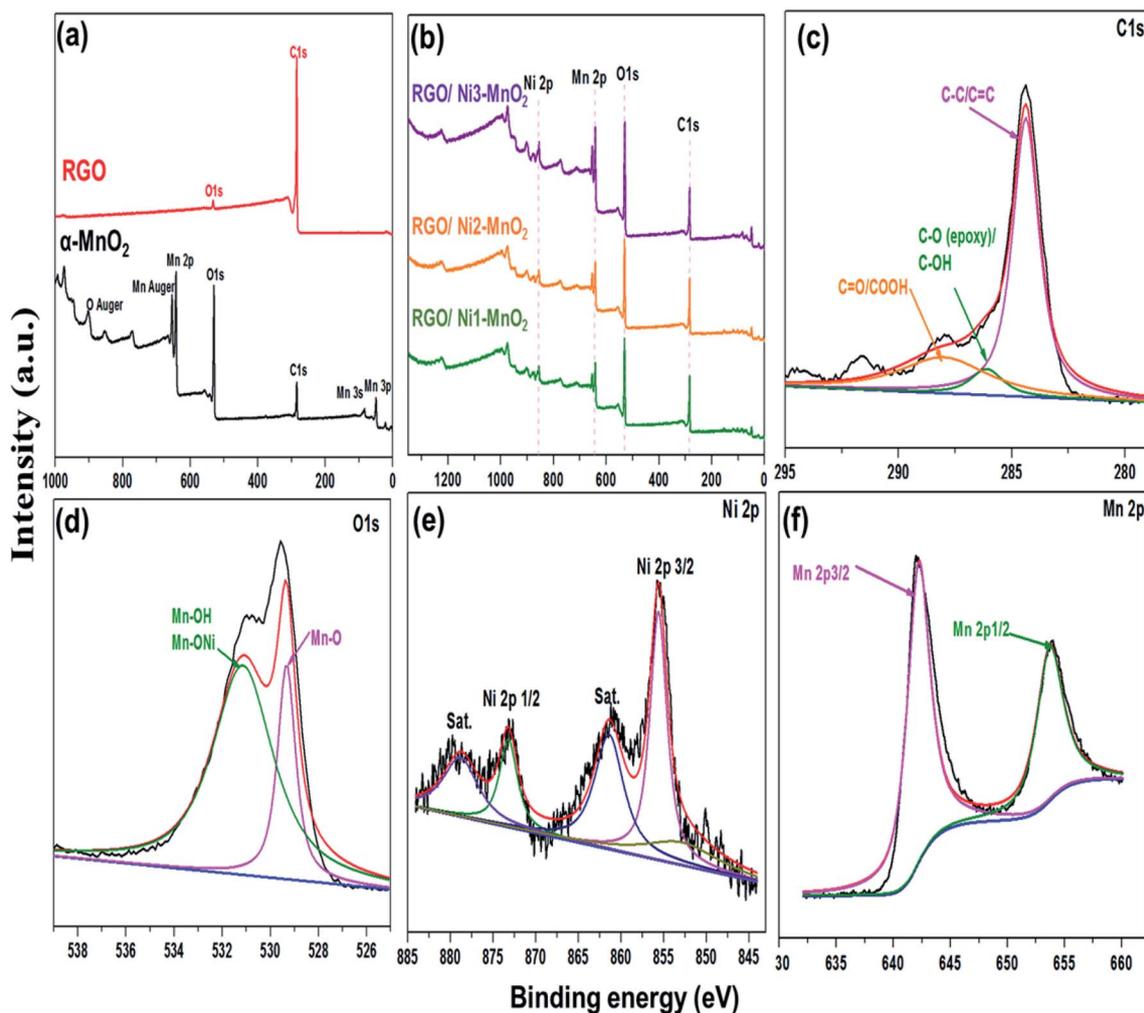


Fig. 5 Survey XPS spectrum of (a) RGO and  $\alpha$ -MnO<sub>2</sub>, (b) all RGO/Ni-MnO<sub>2</sub> adsorbents, XPS spectra of (c) C 1s, (d) O 1s, (e) Ni 2p and (f) Mn 2p of RGO/Ni3-MnO<sub>2</sub> adsorbent.

mentioned in Fig. 6(d). It is investigated that the adsorption amount possessed direct relationship with Li<sup>+</sup> initial concentration, reason for this considered that Li<sup>+</sup> initial amount may supply the driving force to overwhelm the hindrance of mass migration of Li<sup>+</sup> among aqueous and solid media.<sup>52</sup> Therefore, the maximum reported adsorption  $q_e$  were 52 mg g<sup>-1</sup>, 56 mg g<sup>-1</sup> and 63 mg g<sup>-1</sup> at highest initial concentration (50 ppm) for RGO/Ni1-MnO<sub>2</sub>, RGO/Ni2-MnO<sub>2</sub>, and RGO/Ni3-MnO<sub>2</sub> adsorbents respectively. Furthermore, in order to understand the nature of adsorption process (*e.g.* the amount of adsorption and related phenomena) the isothermal study was performed by applying two different adsorption models including Langmuir, Freundlich on experimental data.

At first, Langmuir<sup>53</sup> utilized this adsorption model to explain the adsorption behavior of gasses above the homogeneous material surface (usually crystalline base solids) that possess single kind of adsorption site. Different researchers have elaborated the Langmuir study to explain the sorption of solution molecules above the heterogeneous solid adsorbent,<sup>54</sup> while the Langmuir adsorption model equation is as follows,

$$\frac{C_e}{q_e} = \frac{C_e}{q_m} + \frac{a}{q_m}$$

where  $q_e$  represent the quantity of adsorbate adsorbed per unit amount of adsorbent in mg g<sup>-1</sup>,  $b$  is the Langmuir constant,  $a$  indicates the highest adsorption tendency of adsorbents (mg g<sup>-1</sup>) and  $C_e$  attributed to concentration of solution at equilibrium (mg L<sup>-1</sup>). The Freundlich adsorption model is valid to those adsorption natures that possess heterogenous solid surface. It provides information related to heterogeneous nature of surface material. The Freundlich adsorption model is described as follows,<sup>55</sup>

$$\log q_e = \log K_F + \frac{1}{n} \log C_e$$

Here,  $K_F$  indicated the adsorption capacity and it is Freundlich constant, and  $1/n$  represents the intensity of adsorption as well as heterogeneity of adsorbent solid sites, where  $n$  is empirical constant.

From the Fig. 6(e and f) it is clearly observed that the adsorption experimental data fitted well by both adsorption



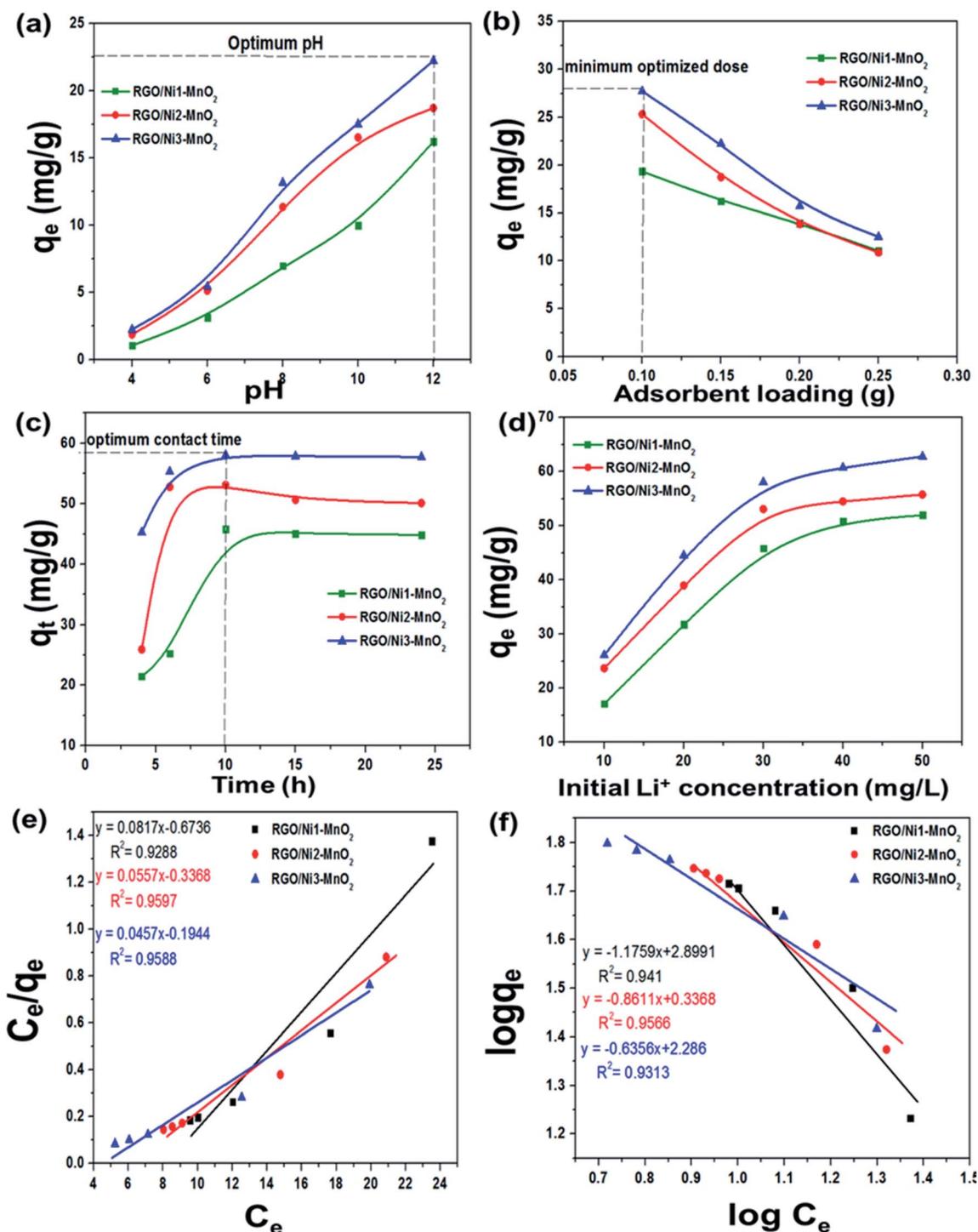


Fig. 6 (a) Effect of pH (conditions; 0.15 g/250 mL, 30 mg L<sup>-1</sup>, 5 h), (b) effect of adsorbent loading (conditions; pH = 12, 250 mL, 30 mg L<sup>-1</sup>, 5 h), (c) effect of contact time (conditions; pH = 12, 0.1 g/250 mL, 30 mg L<sup>-1</sup>), (d) effect of Li<sup>+</sup> initial concentration (conditions; pH = 12, 0.1 g/250 mL, 10 h) on Li<sup>+</sup> adsorption capacity by adsorbents at 30 °C, (e) Langmuir plot and (f) Freundlich plot.

models because the adsorption correlation coefficient values ( $R^2$ ) of both models are close to unity, which demonstrating that the adsorbent surface sites were heterogeneous. Briefly, it can have concluded that greater than single kind of active site involved in the adsorption process of Li<sup>+</sup> metal.

#### 3.4. Desorption and reusability analysis

The desorption percentage was found to be correlated with the ratio of Li<sup>+</sup> recovered to the amount of Li<sup>+</sup> adsorbed on the surface. The most efficient adsorbents for metal ions were those showing excellent adsorption as well as desorption behavior.<sup>56</sup>



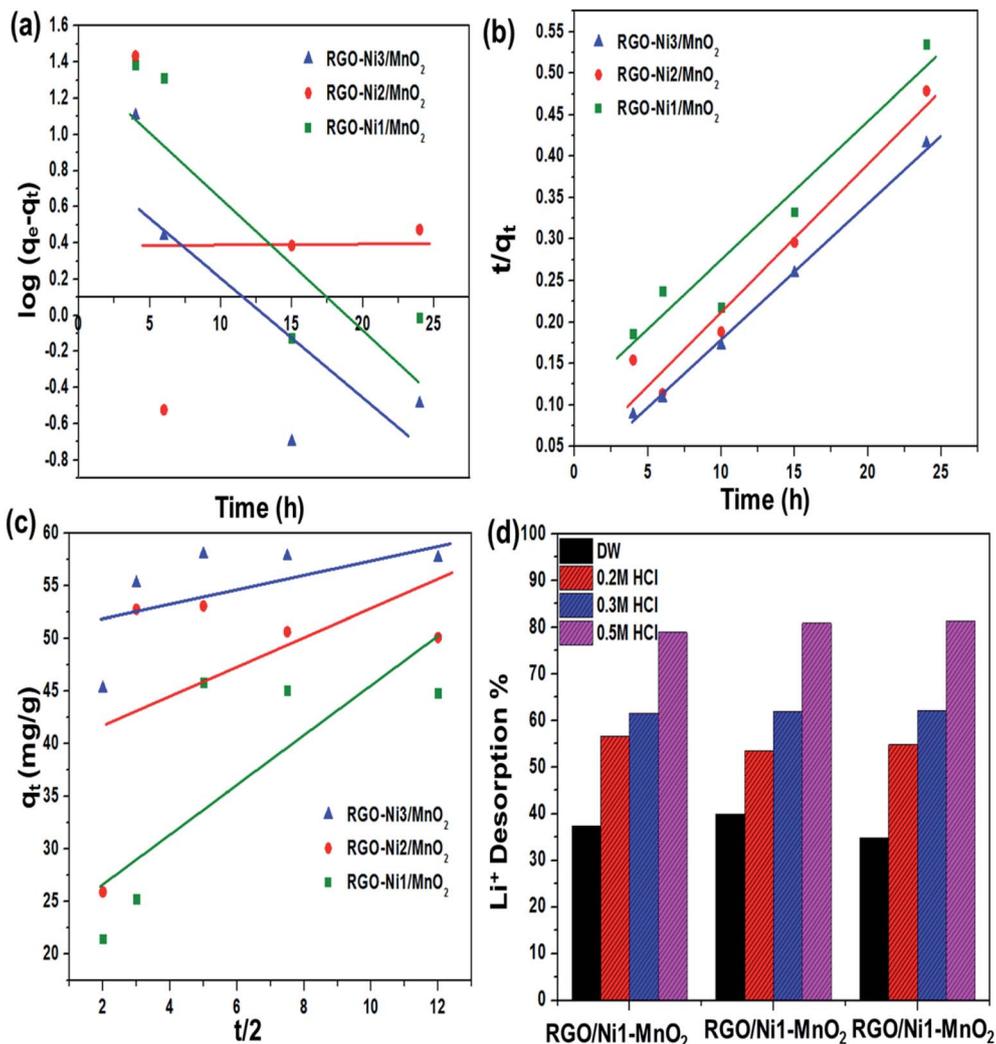


Fig. 7 Kinetics plots of  $\text{Li}^+$  adsorption (a) pseudo first order, (b) pseudo second order, (c) intra-particle diffusion model, (d) effect of distilled water (DW) and HCl concentrations on desorption capacity on adsorbents.

In this regard, desorption of  $\text{Li}^+$  ions from RGO/Ni- $\text{MnO}_2$  can be studied by selecting eluting reagents that are economical, effective, and non-damaging to the adsorbents. Based on a literature review, it was predicted that HCl should prove to be a more efficient eluting agent compared to  $\text{H}_2\text{SO}_4$  and  $\text{HNO}_3$ .<sup>57,58</sup> Therefore, DW and HCl were selected as an eluting reagent. The influence of DW and HCl solutions of various concentrations (0.2, 0.3, and 0.5 M) on desorption of  $\text{Li}^+$  from adsorbents are shown in Fig. 7(d). The results indicate that

about 39% of lithium was recovered if DW was used as the stripping agent. By increasing the concentration of acid, the metal desorption rate was raised; the maximum amount of  $\text{Li}^+$  recovered was 81% with a 0.5 M HCl solution. The metal desorbed because of the reversible phenomenon of cation exchange between the adsorbent surface and the acidic medium. In this regard, the dominating concentration of  $\text{H}^+$  ions in the as-synthesized adsorbents led to the extraction of the previously adsorbed  $\text{Li}^+$  ions.

Table 2 The constants of adsorption kinetics for  $\text{Li}^+$  uptakes by RGO/Ni- $\text{MnO}_2$  adsorbents

Samples	Pseudo 1 <sup>st</sup> order		Pseudo 2 <sup>nd</sup> order		$q_t$ (cal)	Intraparticle diffusion		
	$k_1$	$R^2$	$k_2$	$R^2$		$k_{pi}$	$R^2$	$C_i$
RGO/Ni1- $\text{MnO}_2$	0.1701	0.5965	$2.856 \times 10^{-3}$	0.9311	58.823	2.297	0.581	22.941
RGO/Ni2- $\text{MnO}_2$	0.0103	0.0025	$9.051 \times 10^{-3}$	0.9532	55.865	1.321	0.208	38.73
RGO/Ni3- $\text{MnO}_2$	0.1736	0.6898	$2.222 \times 10^{-3}$	0.9977	60.241	0.862	0.402	49.754



The regeneration ability of adsorbents always taking into intense consideration in order to increasing their practical applications. It was checked by consecutive adsorption-desorption cycles of all adsorbents as shown in Fig. 8(b). The outcomes showed that  $\text{Li}^+$  adsorption capacities were slightly declined (up-to 2%) until 4<sup>th</sup> adsorption-desorption cycle. The experimental  $q_e$  values showed that  $\text{Li}^+$  capture abilities from 1<sup>st</sup> to 4<sup>th</sup> cycles were as follows, 52–49 mg  $\text{g}^{-1}$  (RGO/Ni1-MnO<sub>2</sub>), 55–53 mg  $\text{g}^{-1}$  (RGO/Ni2-MnO<sub>2</sub>) and 62–60 mg  $\text{g}^{-1}$  (RGO/Ni3-MnO<sub>2</sub>) respectively. Overall, it is resultant that although with 2% decreased in  $\text{Li}^+$  adsorption ability after 4<sup>th</sup> round, the synthesized adsorbents still have tendency to capture  $\text{Li}^+$  ions from water and exhibits efficient stability. The adsorption mechanism was proposed to be ion exchange as mentioned in Fig. 8(c), because the ion exchange phenomena may be explained as the interchange of ions among the adsorbents and the surrounding

ions in aqueous solution.<sup>59</sup> The adsorption mechanism can be proposed to be ion exchange because the adsorption/desorption process is reversible in nature. As the adsorption process is reversible, that's why all prepared adsorbents can be reused after several cycles.

### 3.5. Influence of coexisting ions

Generally industrial wastewater has different types of metal ions. These coexisting ions may influence the adsorption of lithium ions *via* competing with targeted  $\text{Li}^+$  ions for sorption, because adsorbent may having tendency to adsorbed the coexisting metal ions within the wastewater.<sup>60</sup> Therefore, it is essential to identify the selectivity of the samples in the sorption process. For this purpose, 0.1 g of different adsorbents were added in 250 mL of 50 ppm solution (containing various ions;

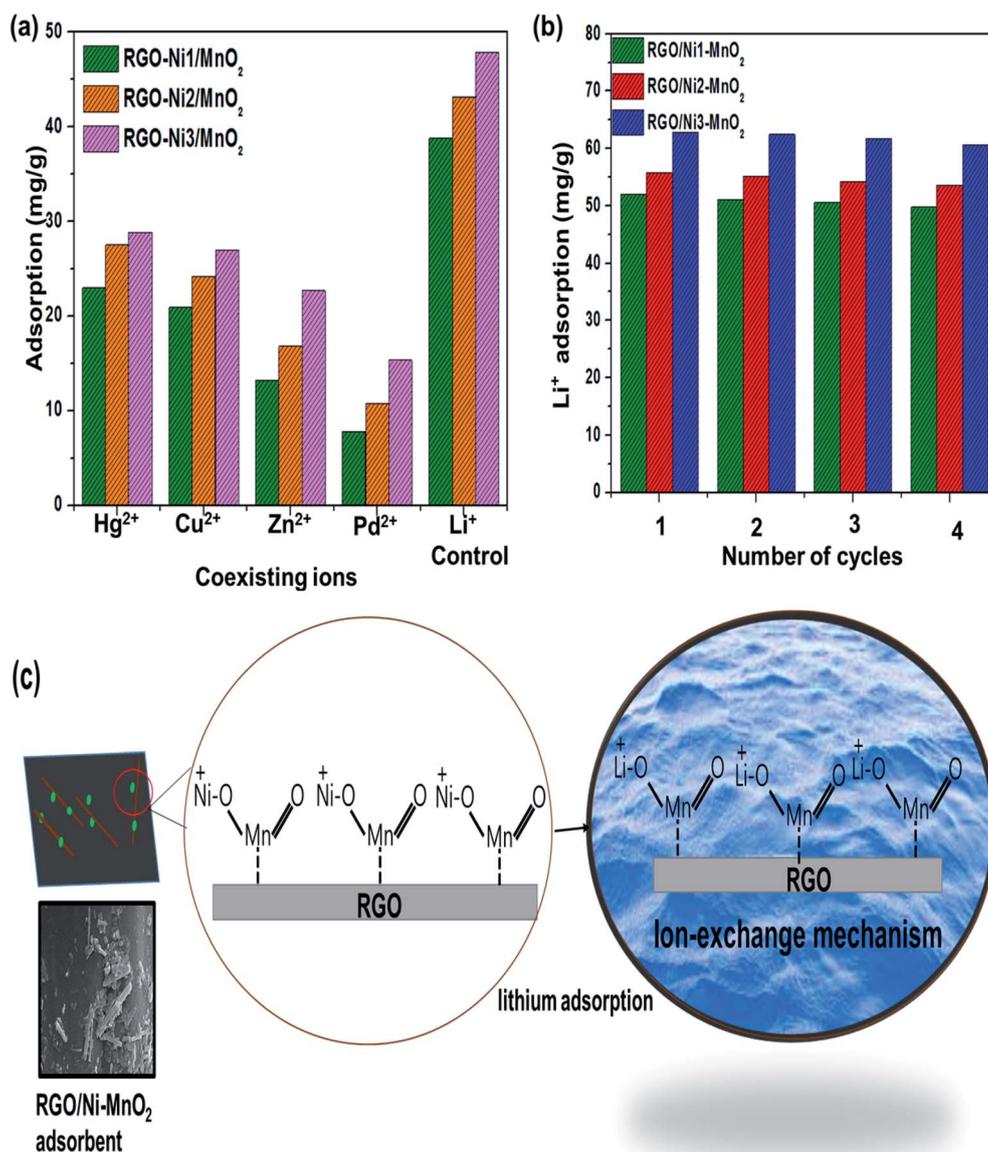


Fig. 8 (a) Effects of coexisting ions on the adsorption of adsorbents for  $\text{Li}^+$ ; (conditions; pH 12, 0.1 g, 250 mL, 50 ppm, 24 h), (b) reusability of adsorbents for 4 sequential cycles and (c) proposed ion exchange mechanism of lithium capture by adsorbents.



Hg<sup>2+</sup>, Cu<sup>2+</sup>, Zn<sup>2+</sup>, Pd<sup>2+</sup> and Li<sup>+</sup>) and allowed to stir for 24 h at room temperature. After centrifugation, the AAS analysis performed and result summarized in Fig. 8(a), which shows that the adsorption efficiency of adsorbents for Li<sup>+</sup> ions declined in the presence of coexisting ions with  $q_e$  values 38 mg g<sup>-1</sup>, 43 mg g<sup>-1</sup> and 47 mg g<sup>-1</sup> for RGO/Ni1-MnO<sub>2</sub>, RGO/Ni2-MnO<sub>2</sub> and RGO/Ni3-MnO<sub>2</sub> respectively. But this adsorption efficiency of adsorbents for Li<sup>+</sup> ions is still higher than the other coexisting ions in the aqueous medium.

## 4. Conclusion

In summary, RGO-Ni/MnO<sub>2</sub> adsorbents were fabricated using different mass ratios of nickel contents through a rapid, economical, and facile hydrothermal technique. The effect of adsorption parameters including pH, dose loading, contact time, and initial Li<sup>+</sup> concentrations were evaluated. Under optimized adsorption conditions, the reported highest equilibrium adsorption capacities ( $q_e$ ) of all adsorbents were 52 mg g<sup>-1</sup>, 56 mg g<sup>-1</sup> and 63 mg g<sup>-1</sup> in the following order: RGO/Ni3-MnO<sub>2</sub> > RGO/Ni2-MnO<sub>2</sub> > RGO/Ni1-MnO<sub>2</sub>. Among kinetic models, the pseudo-second order model provided the best fit results, as determined by its greater  $R^2$  values. The desorption study with eluting agents (DW and HCl) revealed that a high acid concentration (0.5 M) can recover about 81% of Li<sup>+</sup> ions compared to DW (39%). Furthermore, all adsorbents exhibit regeneration ability after four consecutive cycles and existence of interfering ions slightly affects the lithium adsorption. Conclusively, doping of nickel contents can consider to be very efficient strategy for improving Li<sup>+</sup> adsorption ability of adsorbents.

## Conflicts of interest

There are no conflicts to declare.

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