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Catalytic activation of peroxymonosulfate with manganese cobaltite nanoparticles for the degradation of organic dyes[†]

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In this work, we report the facile hydrothermal synthesis of manganese cobaltite nanoparticles ($MnCo_2O_{4.5}$ NPs) which can efficiently activate peroxymonosulfate (PMS) for the generation of sulfate free radicals ($SO_4^{\cdot-}$) and degradation of organic dyes. The synthesized $MnCo_2O_{4.5}$ NPs have a polyhedral morphology with cubic spinel structure, homogeneously distributed Mn, Co, and O elements, and an average size less than 50 nm. As demonstrated, $MnCo_2O_{4.5}$ NPs showed the highest catalytic activity among all tested catalysts (MnO_2 , CoO) and outperformed other spinel-based catalysts for Methylene Blue (MB) degradation. The MB degradation efficiency reached 100% after 25 min of reaction under initial conditions of 500 mg L⁻¹ Oxone, 20 mg L⁻¹ $MnCo_2O_{4.5}$, 20 mg L⁻¹ MB, unadjusted pH, and $T = 25$ °C. $MnCo_2O_{4.5}$ NPs showed a great catalytic activity in a wide pH range (3.5–11), catalyst dose (10–60 mg L⁻¹), Oxone concentration (300–1500 mg L⁻¹), MB concentration (5–40 mg L⁻¹), and temperature (25–55 °C). HCO_3^- , CO_3^{2-} and particularly Cl^- coexisting anions were found to inhibit the catalytic activity of $MnCo_2O_{4.5}$ NPs. Radical quenching experiments revealed that sulfate radicals are primarily responsible for MB degradation. A reaction sequence for the catalytic activation of PMS was proposed. The as-prepared $MnCo_2O_{4.5}$ NPs could be reused for at least three consecutive cycles with small deterioration in their performance due to low metal leaching. This study suggests a facile route for synthesizing $MnCo_2O_{4.5}$ NPs with high catalytic activity for PMS activation and efficient degradation of organic dyes.

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1. Introduction

Organic dyes are universally found in the wastewater of paint, printing and textile industries. Their high solubility, low biodegradability, and toxicity cause severe environmental pollution and threaten human health. To deal with these problems, many different methods have been exploited for the removal and degradation of organic dyes, and it is found that advanced oxidation processes (AOPs) are very promising due to their high efficiency. During traditional AOPs such as Fenton's

reagent, photo-Fenton, electro-Fenton, ozonation, and UV/TiO₂ treatment, the highly reactive hydroxyl radicals ($\cdot OH$, reduction potential $E^0 = 2.8$ V) are *in situ* generated and unselectively oxidize organic contaminants to smaller and less harmful molecules.¹ Nevertheless, classic AOPs have intrinsic limitations for practical applications such as sludge generation, strong pH-dependent efficiency, and cost ineffectiveness. Recently, sulfate free radicals ($SO_4^{\cdot-}$) have been attracting increasing attention due to its advantages such as stronger oxidizy ($E^0 = 2.5$ –3.1 V), longer half-life (30–40 μ s, vs. 10⁻³ μ s for $\cdot OH$), higher selectivity (to organic compounds), and the ability to work in a wide pH range.² Ultimately, these lead to the acceleration of the degradation of dissolved organic pollutants and improvement of degradation yield. $SO_4^{\cdot-}$ radicals are commonly generated from persulfate (PS) or peroxymonosulfate (PMS) through different activation methods using UV irradiation,³ heat,⁴ base,⁵ ultrasound,⁶ and catalyst.^{2,7} Differing from other methods, catalytic activation requires no high-energy input or sophisticated equipments and thus has great potential for practical applications. In addition, PMS has its own advantage as compared to PS for the generation of $SO_4^{\cdot-}$, thanks to its asymmetric molecular structure and large length of O–O bonds which greatly promote the activation process.²

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The catalytic activation of PMS is mostly based on transition metals such as Fe, Mn, Ni, Co, Cu and Ag in the forms of homogeneous or heterogeneous catalysts.⁸ Although Co, Cu and Ag are excellent PMS activators, their relatively high toxicity (especially Co) and low abundance have hindered their practical uses. Moreover, the utilization of these transitional metals ions as homogeneous activators has several drawbacks such as difficult recovery of activators and secondary pollution. Compared to homogeneous activation, heterogeneous activators show excellent PMS activation capability, chemical stability, low solubility (less metal leaching), easy separation and reuse. Therefore, the development of efficient heterogeneous catalysts for PMS activation becomes a priority in many studies. So far, Co(II)-based heterogeneous catalysts such as CoFe_2O_4 ,^{9,10} CoMn_2O_4 ,¹¹ $\text{Co}_x\text{Mn}_{3-x}\text{O}_4$,¹² CoNi_2O_4 , and CoCu_2O_4 have been extensively studied. Recently, Co(III)-based spinels with general formula MCo_2O_x (M is transition metals) have been paid great attention due to their excellent electrochemical properties.^{13,14} For PMS activation, these compounds exhibit several key advantages such as high catalytic activity and low Co leachability in acidic environment (CuCo_2O_4 (ref. 15)). $\text{Co}_x\text{Mn}_{3-x}\text{O}_4$ showed greater catalytic performance than that of the component oxides (Co_3O_4 and Mn_2O_3), which indicates the synergistic effects of Co and Mn species.¹² Therefore, exploring novel routes for the synthesis and catalytic performance of cobaltite in PMS activation is of great importance in environmental remediation.

In this work, manganese cobaltite nanoparticles ($\text{MnCo}_2\text{O}_{4.5}$ NPs) have been prepared by a combined hydrothermal treatment/annealing route and applied as a heterogeneous catalyst to activate PMS for the degradation of various synthetic dyes, with Methylene Blue (MB) selected as a model pollutant. Our main goals are to: (i) investigate the influence of initial reaction parameters (MB and Oxone concentration, catalyst dosage, pH, reaction temperature) and the presence of common co-existing anions on MB degradation; (ii) identify the major radical species, reaction pathways, and mechanism; (iii) assess the reusability of the catalyst. The crystal structure, morphology, chemical composition and oxidation states of the elements in $\text{MnCo}_2\text{O}_{4.5}$ NPs were investigated. The influence of various parameters on catalytic activity of $\text{MnCo}_2\text{O}_{4.5}$ NPs for PMS activation towards degradation of MB was studied. The mechanism for generating free radicals was investigated based on quenching experiments. We also evaluated the capability of $\text{MnCo}_2\text{O}_{4.5}$ NPs as heterogeneous catalyst for the treatment of textile wastewater and landfill leachate.

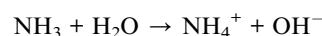
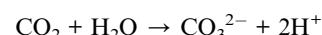
2. Experiments

2.1. Chemicals

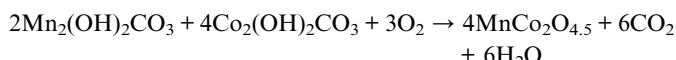
Manganese chloride tetrahydrate ($\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$), cobalt(II) chloride hexahydrate ($\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$), urea ($(\text{NH}_2)_2\text{CO}$), manganese(IV) oxide (MnO_2), cobalt(II) oxide (CoO), Oxone ($\text{KHSO}_5 \cdot 0.5\text{KHSO}_4 \cdot 0.5\text{K}_2\text{SO}_4$), ethanol ($\text{C}_2\text{H}_5\text{OH}$, 99.5%), Methylene Blue (MB), Rhodamine B (RhB), Orange G (OG), and Methyl Orange (MO) were purchased from Shanghai Aladdin Bio-Chem Technology Co., Ltd. All the reagents were used without further purification.

2.2. Synthesis of $\text{MnCo}_2\text{O}_{4.5}$

$\text{MnCo}_2\text{O}_{4.5}$ nanomaterials were synthesized by hydrothermal method using Mn^{2+} , Co^{2+} chloride salts and urea as precursors. Typically, $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ (5 mmol), $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ (10 mmol) and urea (0.9 g) were dissolved in 75 mL DI H_2O and magnetically stirred for 30 min, then transferred to a Teflon-lined autoclave and maintained at 120 °C for 6 h. In this condition, urea is hydrolyzed and decomposed into ammonia and carbon dioxide, which also undergo immediate hydrolysis to generate carbonate and hydroxyl anions. They combine with Mn^{2+} and Co^{2+} cations to form metal(II) hydroxycarbonates. The reactions involved in the hydrothermal treatment can be described by the following equations:



The resulting mixed metal hydroxycarbonates was washed with distilled water and ethanol several times, and dried at 60 °C for 12 h. Finally, the obtained pink powder was annealed in air at 400 °C for 2 h, during which the following reaction occurs:



The as-synthesized $\text{MnCo}_2\text{O}_{4.5}$ sample was obtained in the blackish and powdery form.

2.3. Characterization of $\text{MnCo}_2\text{O}_{4.5}$

The as-synthesized $\text{MnCo}_2\text{O}_{4.5}$ sample was characterized by X-ray diffraction (XRD; Siemens D5005 X-ray diffractometer; $\text{Cu K}\alpha$ radiation, $\lambda = 1.5418 \text{ \AA}$; voltage/current 30 kV/40 mA; scan rate 0.9°/min; scan step 0.03°; 2θ from 10 to 70°), scanning electron microscopy (SEM; Hitachi S4800 field-emission scanning electron microscope), scanning TEM (STEM; JEOL JEM-ARM200F microscope; high-angle annular dark-field (HAADF) detector; voltage 200 kV; spherical aberration corrector; nominal resolution 0.8 \AA), X-ray photoelectron spectroscopy (XPS; Shimadzu Kratos AXIS-ULTRA DLD; monochromated $\text{Al K}\alpha$ radiation). The average crystallite size of $\text{MnCo}_2\text{O}_{4.5}$ particles was calculated using Scherrer equation:¹⁶

$$d = k \times \lambda / (\beta \times \cos \theta) \quad (1)$$

where λ is the wavelength of $\text{Cu K}\alpha$ irradiation, $\lambda = 1.5418 \text{ \AA}$, β is the full width at maximal half (FWHM), θ is Bragg diffraction angle (peak position), and $k = 0.9$.

2.4. Catalysis studies

The catalytic activity of $\text{MnCo}_2\text{O}_{4.5}$ NPs for PMS activation was evaluated through the degradation of organic dyes in batch

experiments. We first prepared the stock suspension/solutions of $\text{MnCo}_2\text{O}_{4.5}$ NPs, MB, and Oxone (1000, 100, and 5000 mg L⁻¹, respectively) by dispersing pre-weighted amount of respective materials in deionized (DI) water using volumetric flasks. The calculated concentration of PMS (KHSO₅, $M = 152.2$) in the Oxone stock solution is 2475.76 mg L⁻¹. In a typical experiment, a known volume of catalyst suspension was introduced into a 5 mL volumetric flask containing a predetermined volume of MB solution and stirred for 30 min to reach adsorption equilibrium. Then, a known volume of Oxone solution was added to the above solution and shaken to commence the catalytic activation for dye degradation. Immediately after adding Oxone, the resulting solution was mixed well and transferred to a cuvette for UV-Vis measurement (*ca.* 4 mL), with the reference solution containing $\text{MnCo}_2\text{O}_{4.5}$ and Oxone of the same concentrations with those of the studied samples (without the presence of MB). The MB degradation by Oxone in the absence of $\text{MnCo}_2\text{O}_{4.5}$ catalyst was also performed using the same concentration of MB (20 mg L⁻¹) and Oxone (500 mg L⁻¹), employing a 500 mg L⁻¹ Oxone solution as a reference. When necessary, initial solution pH was adjusted by adding 10 mM NaOH or 10 mM HNO₃ solution. At various time intervals, the remaining concentration of MB was calculated according to UV-Vis data (Biochrom SP60, optical absorbance at 664 nm) and Beer-Lambert law. The degradation efficiency of MB (γ) was calculated using the formula:

$$\gamma (\%) = (C_o - C_t) \times 100/C_o \quad (2)$$

where C_o is the initial concentration and C_t is the real-time concentration of MB. The term $\gamma_{\Delta t}$ indicates the degradation efficiency of MB obtained after Δt min of treatment. The investigated experimental parameters were catalyst dose (10, 20, 30, 40, and 60 mg L⁻¹), Oxone concentration (100, 300, 500, 700, 1000, and 1500 mg L⁻¹), MB concentration (5, 10, 15, 20, 30, and 40 mg L⁻¹), temperature (25, 35, 45, and 55 °C), initial solution pH (3.5, 4.5, 5.5, 7.0, 8.0, 9.0, 10.0, 11.0, and 12.5), type of dye (RhB, MO, OG, and MB), reaction sequence, co-existing anions (Cl⁻, HCO₃⁻, CO₃²⁻, NO₃⁻), type of oxidant (PS, PMS, and H₂O₂). In the experiment to investigate the influence of temperature, the radical-generating reactions were stopped by quickly adding 0.2 mL of Na₂SO₃ 0.1 M and the catalyst NPs were removed by centrifugation. The influence of different coexisting anions on MB degradation was investigated by using their corresponding sodium salts at concentration of 0–20 mM. Quenching tests were performed by adding a predetermined amount of quencher into MB solution before adding Oxone. In this work, *tert*-butyl alcohol (TBA), ethanol (EtOH), and phenol have been employed as quenchers. Cycling tests were carried out by centrifugation, washed with double distilled water and EtOH till neutral, dried at 60 °C, then re-used. The concentration of possibly leached Mn and Co ionic species was quantified using an inductively coupled plasma mass spectrometry (ICP-MS, PerkinElmer NexION® 2000).

3. Results and discussion

3.1. Synthesis and characteristics of the as-prepared $\text{MnCo}_2\text{O}_{4.5}$ NPs

The catalyst was synthesized in two steps, as illustrated in Fig. 1a. In the first step, an aqueous solution containing Mn²⁺, Co²⁺, and urea is hydrothermally treated to form a pink precursor. In the second step, the precursor is annealed to get the final black solid. Its major XRD peaks (Fig. 1b) occurring at 18.6, 31.4, 36.7, 44.3, 59.1 and 64.6° are well indexed to (111), (220), (311), (422), (511) and (440) planes of $\text{MnCo}_2\text{O}_{4.5}$ (JCPDS #32-0297), correspondingly. This material possesses cubic spinel structure and space group $Fd\bar{3}m$ (227), with Mn ions occupying octahedral sites, and Co ions occupying both octahedral and tetragonal sites, as shown in Fig. 1c. In the FT-IR spectrum (Fig. 1d), the wide and weak bands observed at 3347.2 and 1633.3 cm⁻¹ are respectively associated with -OH stretching and bending vibrations of adsorbed water molecules. The sharp and strong peaks observed at 652.8 and 547.2 cm⁻¹ correspond to stretching vibrations of metal–oxygen bondings at tetrahedral and octahedral sites.¹⁷ The peak occurred at 1118.1 cm⁻¹ indicates the presence of C–O bondings, possibly due to the presence of adventitious carbon species. TG curve of the catalyst (Fig. 1e) displays two major events of the weight loss, which represent the dehydration and deoxygenation of the catalyst, respectively. The percentages of both weight losses were only 1.2 and 0.8%, indicating the high thermal stability of the material.

The morphology and structure of the catalyst were studied using SEM and STEM. As shown in Fig. 2a and b, the $\text{MnCo}_2\text{O}_{4.5}$ catalyst consists of polyhedral nanoparticles (NPs) with an average size less than 50 nm. These $\text{MnCo}_2\text{O}_{4.5}$ NPs are composed of Mn, Co, and O elements, as confirmed by STEM-HAADF image and EDS elemental mappings (Fig. 2b). These elements are homogeneously distributed throughout all NPs, indicating that the manganese cobaltite has been successfully formed. To identify the oxidation state of these elements, XPS analyses were performed. The presence of Mn, Co, and O in $\text{MnCo}_2\text{O}_{4.5}$ NPs is well confirmed in the XPS survey spectra (Fig. 2c), in which C peak is due to adventitious carbon contamination. The XPS core-level spectra of Mn 2p, Co 2p, and O 1s were calibrated using C 1s line as reference, background subtracted, deconvoluted and fitted by Gaussian method. The deconvoluted Mn 2p spectrum (Fig. 2d) showed four peaks, which can be ascribed to spin-orbit characteristics of Mn³⁺ (655.3 and 642.9 eV) and Mn²⁺ (653.4 and 641.0 eV),¹⁸ respectively. The spin-orbit splitting of Mn 2p was observed to be 11.7 eV. Fig. 2e shows that the XPS Co 2p spectrum is deconvoluted into Co²⁺ (796.2 and 780.4 eV), Co³⁺ (794.1 and 779.3 eV), and shake-up satellites peaks of Co (804.6 and 789 eV).¹⁹ The deconvoluted O 1s spectrum (Fig. 2f) showed O species in the metal–oxygen bonding (529.3 eV) and in surface defect sites (531 eV).²⁰ The presence of Mn²⁺, Mn³⁺, Co²⁺, Co³⁺, and O species on the surface of $\text{MnCo}_2\text{O}_{4.5}$ NPs clearly indicates that $\text{MnCo}_2\text{O}_{4.5}$ NPs have been successfully formed using our synthetic route.

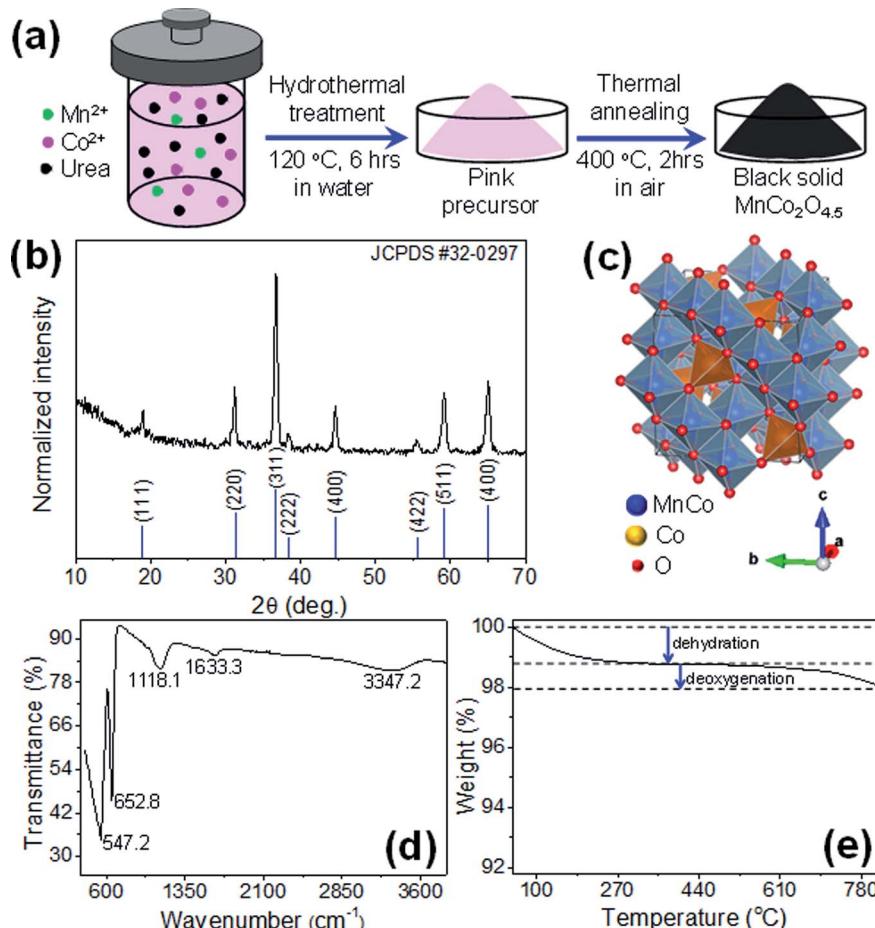


Fig. 1 (a) Schematic illustration of the preparation method for $\text{MnCo}_2\text{O}_{4.5}$ catalyst; (b) XRD patterns; (c) structural model; (d) FT-IR spectrum; and (e) TG curve of $\text{MnCo}_2\text{O}_{4.5}$ NPs.

3.2. Performance evaluation

3.2.1. Comparison of related catalysts. We firstly investigated MB degradation in the presence of various catalyst samples without pH adjustment, and the results are shown in Fig. 3a. When only $\text{MnCo}_2\text{O}_{4.5}$ was present, the degradation efficiency was very low ($\gamma_{20} = 5.2\%$), indicating that the degradation of MB by this catalyst itself was low in the absence of PMS. When PMS was used alone, the MB degradation proceeded more considerably, with γ_{20} reached 10.9%. The combination of PMS with commercially available MnO_2 and CoO bulky powders resulted in a γ_{20} of 25.8% and 25.3%, correspondingly. The homogeneous Mn^{2+} /PMS system gave a γ_{20} of only 22.9%, slightly lower than that of MnO_2 , indicating that Mn^{2+} cations are weak catalysts for PMS activation. The homogeneous Co^{2+} (at a very small concentration of 0.4 mg L^{-1}) displayed a high γ_{20} of 44.0%, as well documented in literatures.²⁷ When $\text{MnCo}_2\text{O}_{4.5}$ NPs were used, 96.7% of MB had been degraded after only 20 min. These comparisons show that $\text{MnCo}_2\text{O}_{4.5}$ catalyst performed better than its oxide components (commercial MnO_2 and CoO powder) and homogeneous Mn^{2+} . The excellent catalytic performance of $\text{MnCo}_2\text{O}_{4.5}$ NPs might originate from several aspects, namely (i) their small particle

size (as observed by TEM) and a high density of surface hydroxyl groups (as characterized by FTIR), and (ii) synergistic coupling of Mn–Co as previously observed and proposed for Mn–Fe,²¹ Cu–Fe,²² and Fe–Co.²³

3.2.2. Influence of catalyst dosage, initial Oxone and MB concentrations, and temperature. The influence of initial reaction parameters on MB degradation is shown in Fig. 4. As seen in Fig. 4a, the MB degradation proportionally increases with $\text{MnCo}_2\text{O}_{4.5}$ dosage. After 10 min of treatment, the MB degradation increased from 37.8 to 99.2% when $\text{MnCo}_2\text{O}_{4.5}$ dosage increased from 10 to 60 mg L^{-1} . This indicates that $\text{MnCo}_2\text{O}_{4.5}$ catalyst can be implemented at high dosages upon necessity without losing its activity, which is attributed to its excellent dispersibility and stability. Fig. 4b shows MB degradation at different concentrations of Oxone. When Oxone concentration was increased from 100 to 300 mg L^{-1} , the γ_{20} increased from 56.2% to 91.8% due to the increase of SO_4^{2-} concentration. At higher concentrations of Oxone (500 – 1500 mg L^{-1}), MB degradation was completed after 20 min. However, the MB degradation percentage increased very slowly with Oxone concentration, which might be caused by quenching reactions of redundant SO_4^{2-} radicals (eqn (3)–(5)). The generation of SO_5^{2-} radicals



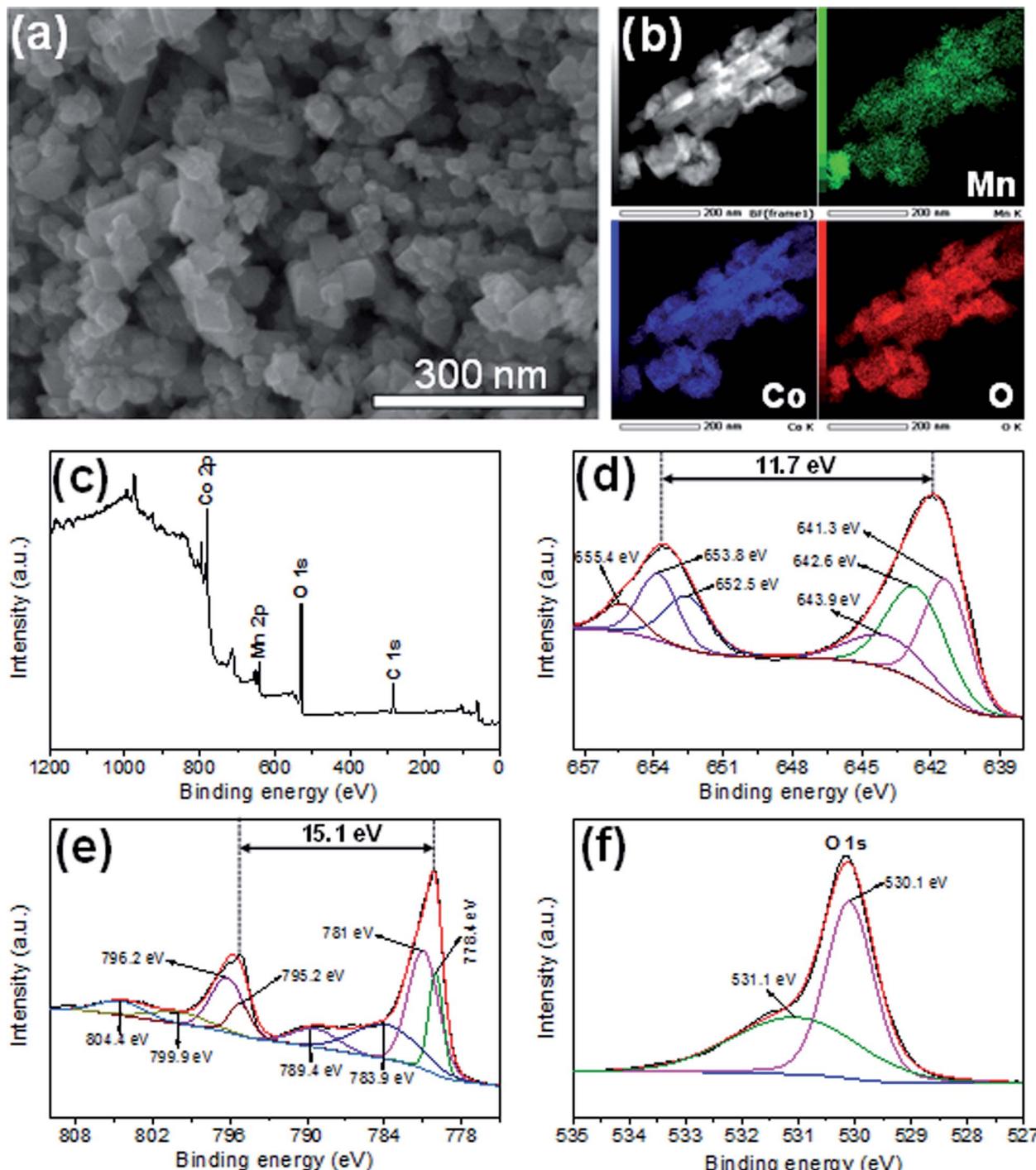
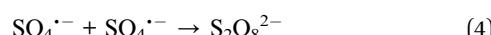
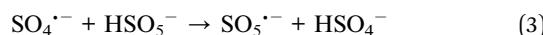


Fig. 2 (a) SEM image; (b) dark-field image and elemental mapping; XPS survey (c) and high-resolution spectra for Mn 2p (d); Co 2p (e); and O 1s (f).

does not affect the degradation yield of MB since they are weak oxidants ($E^0 = 1.1$ V)²⁴



The γ value is also influenced by initial concentration of MB, as shown in Fig. 4c. Generally, the γ value decreases with increasing MB concentration, because a higher concentration of PMS is required to degrade MB completely. At small concentrations (5, 10, and 15 mg L⁻¹), high degradation percentage of 97.6, 95.1, and 97.6% was obtained after 8, 10, and 16 min of reaction, correspondingly. After 20 min, the γ of

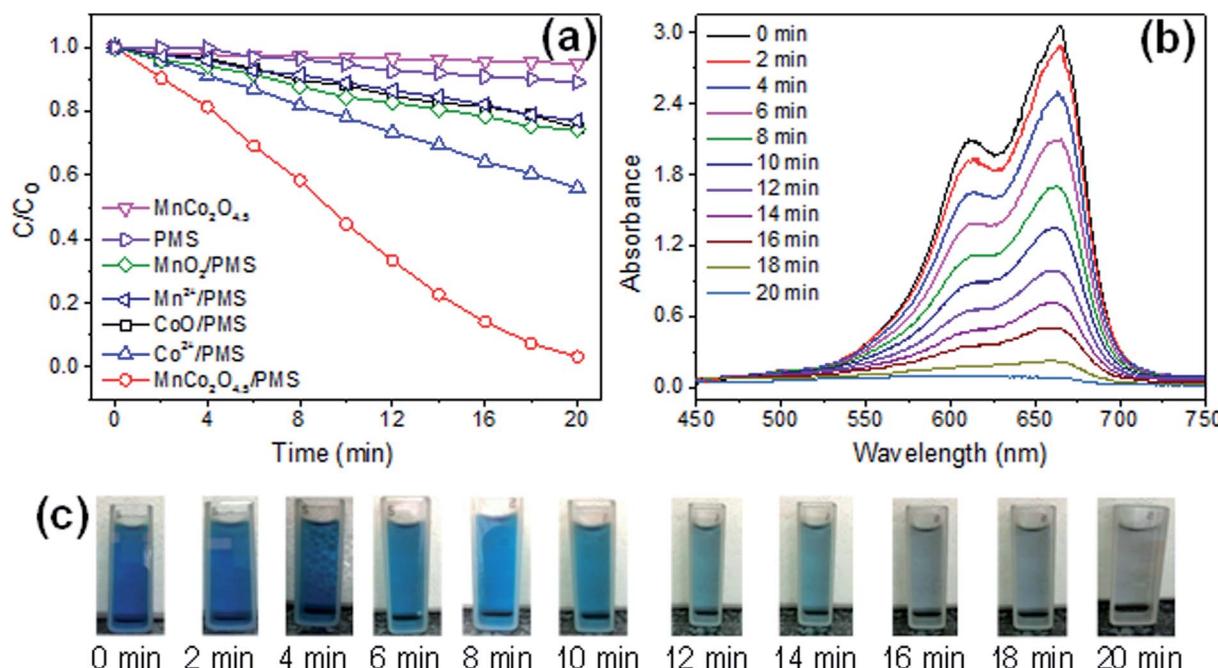


Fig. 3 (a) MB degradation in different comparison experiments. (b) UV-Vis spectra and (c) digital photos of \$\text{MnCo}_2\text{O}_{4.5}/\text{PMS}/\text{MB}\$ solution taken at different durations. Initial concentrations: \$[\text{MB}]_0 = 20 \text{ mg L}^{-1}\$, \$[\text{Oxone}]_0 = 500 \text{ mg L}^{-1}\$, \$[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}\$, \$[\text{MnO}_2]_0 = 20 \text{ mg L}^{-1}\$, \$[\text{Mn}^{2+}]_0 = 20 \text{ mg L}^{-1}\$, \$[\text{Co}^{2+}]_0 = 0.4 \text{ mg L}^{-1}\$, \$[\text{CoO}]_0 = 20 \text{ mg L}^{-1}\$, unadjusted pH, \$T = 25^\circ\text{C}\$.

MB solutions having initial concentration of 20, 30, and 40 mg L⁻¹ was 96.7, 85.2, and 69.2%, respectively. Herein, increasing MB concentration might limit the diffusion process to and from catalyst surface,²¹ therefore it required longer time for complete degradation of MB.²⁵ This result shows that the \$\text{MnCo}_2\text{O}_{4.5}/\text{PMS}\$ reaction system is capable of working in a vast range of dye concentration. It is well known that the reaction temperature has a great influence on the degradation of contaminants.⁷ The influence of reaction temperature on MB degradation was examined in the range of 25–55 °C. As seen in Fig. 4d, MB degradation was dramatically increased with increasing temperature. At 25, 35, 45, and 55 °C, the MB degradation reached 96.7% (20 min), 96.6% (14 min), 96.6% (8 min), and 91.5% (4 min), respectively. This result suggested that the generation of reactive radicals was promoted at high temperature.

3.2.3. Influence of initial solution pH. The pH value of natural and polluted water varies in a broad range, which might affect the generation of free radicals. Herein, we investigated the influence of initial solution pH (3.5–12.5) on MB degradation. We first examined the evolution of pH values of the reaction solution before and after catalysis tests. Without adjusting pH, mixing Oxone (pH ≈ 3.1–3.2) with MB solution (pH ≈ 7) and \$\text{MnCo}_2\text{O}_{4.5}\$ NPs resulted in a solution with a pH value of ca. 3.5. When the initial solution pH was adjusted to either acidic or neutral conditions (pH ≤ 7), the final solution was still acidic with pH < 4. The similar behavior was also observed in other catalyst systems such as \$\text{Co}_3\text{O}_4\$–\$\text{Bi}_2\text{O}_3\$/\$\text{PMS}\$, and mentioned as “buffering capacity”.^{26,27} The acidic pH of the post-reaction environment might be possibly caused by the decomposition of

PMS and the reaction of \$\text{SO}_4^{2-}\$ with water (\$\text{SO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{SO}_4^{2-} + \text{HO}^\cdot + \text{H}^\cdot\$) and MB which produces additional protons and small acidic molecules such as formic, acetic or oxalic acids.^{28,29} However, when the initial solution pH was adjusted to more alkaline conditions (pH > 7), the buffering capacity of \$\text{MnCo}_2\text{O}_{4.5}/\text{PMS}\$ system decreased, and at strongly alkaline conditions (pH ≥ 11), the \$\text{MnCo}_2\text{O}_{4.5}/\text{PMS}\$ system completely lost its buffering capacity. In order to elucidate the pH dependence of MB degradation in \$\text{MnCo}_2\text{O}_{4.5}/\text{PMS}\$ system, we investigated the surface charge of \$\text{MnCo}_2\text{O}_{4.5}\$ NPs as a function of solution pH. As shown in Fig. 5a, the point of zero charge (pH_{pzc}) of as-synthesized \$\text{MnCo}_2\text{O}_{4.5}\$ NPs was 7.22, indicating that \$\text{MnCo}_2\text{O}_{4.5}\$ NPs is positively charged at pH < 7.22, and negatively charged at pH > 7.22. It is also known that \$\text{H}_2\text{SO}_5\$ acid has \$pK_{a_1} < 0\$ (\$\text{H}_2\text{SO}_5 \leftrightarrow \text{HSO}_5^- + \text{H}^\cdot\$) and \$pK_{a_2} \approx 9.4\$ (\$\text{HSO}_5^- \leftrightarrow \text{SO}_5^{2-} + \text{H}^\cdot\$).² Therefore, PMS primarily exists as \$\text{SO}_5^{2-}\$ in a strong alkaline solution (pH > 9.4), as \$\text{HSO}_5^- \$ in acidic, neutral and mild alkaline solutions (pH < 9.4), and as both \$\text{HSO}_5^- \$ and \$\text{SO}_5^{2-}\$ at pH around 9.4.^{2,30–33}

The influence of the initial solution pH on MB degradation is presented in Fig. 5b. As observed, the \$\gamma_{20}\$ was highest in the range of pH from 4.5 to 7.0, where a nearly complete elimination of MB was achieved within ca. 14 min (96.4% at pH 4.5, 96.1% at pH 5.5, and 96.2% at pH 7.0). The high \$\gamma_{20}\$ obtained in this pH range could be attributed to strong Coulomb attraction force between positively charged \$\text{MnCo}_2\text{O}_{4.5}\$ NPs and anionic species generated by PMS. This result is consistent with earlier judgements that the maximal efficiency in PMS activation is achieved in mild acidic and neutral solutions.^{34–36} The \$\gamma_{20}\$ was declined in stronger acidic (pH = 3.5) or alkaline solutions (pH



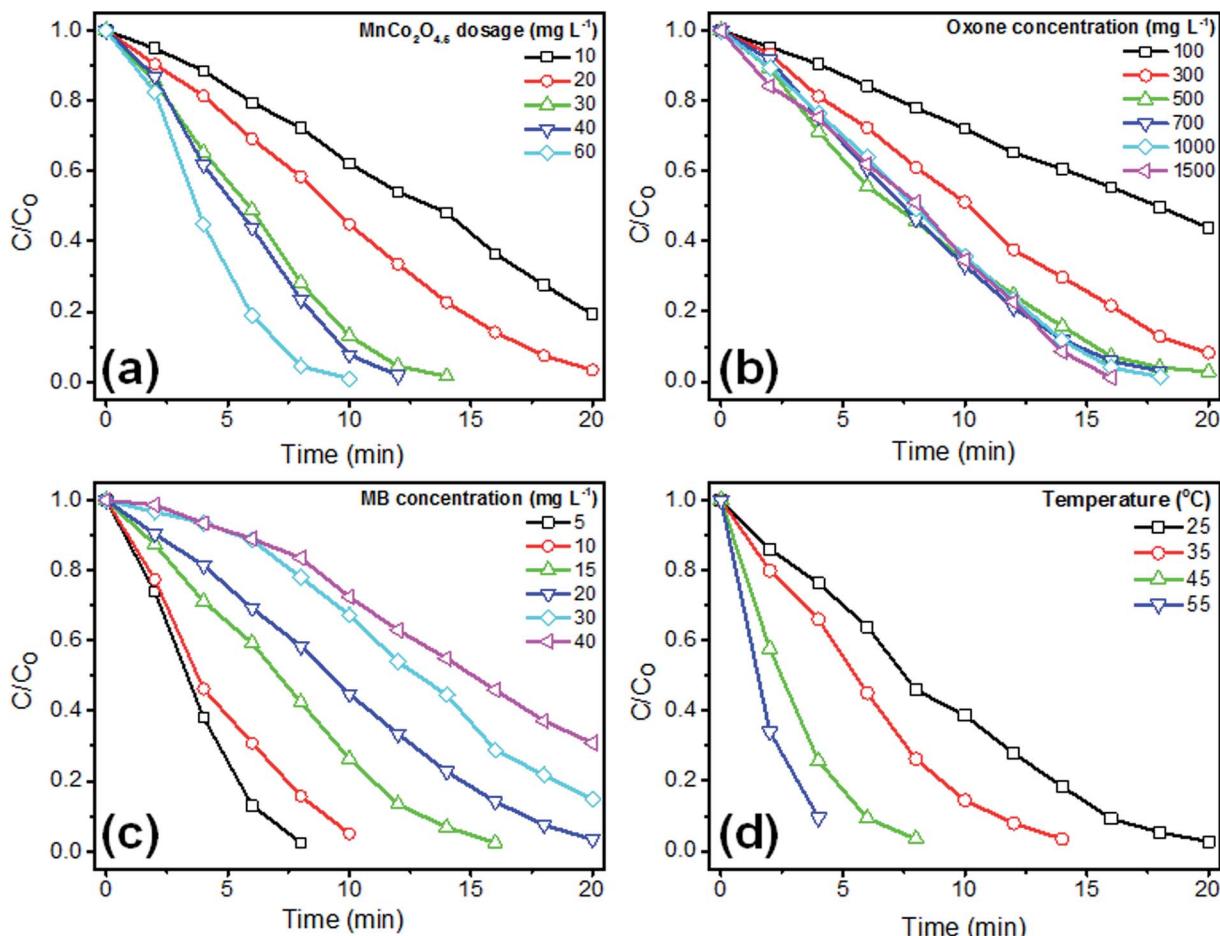


Fig. 4 Influence of initial parameters on MB degradation in $\text{MnCo}_2\text{O}_{4.5}$ /PMS system: (a) $\text{MnCo}_2\text{O}_{4.5}$ dosage; (b) Oxone concentration; (c) initial concentration of MB; and (d) temperature. Except for the examined parameter, other ones were fixed: $[\text{MB}]_0 = 20 \text{ mg L}^{-1}$; $[\text{Oxone}]_0 = 500 \text{ mg L}^{-1}$; $[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}$; unadjusted pH; $T = 25^\circ\text{C}$.

≥ 8), possibly due to several following factors: In acidic pH, the formation of H-bond between H^+ and the O-O groups of HSO_5^- becomes significant and thus hinders the interaction between HSO_5^- and positively charged catalyst surface.³⁷ In alkaline pH, $\text{MnCo}_2\text{O}_{4.5}$ catalyst is negatively charged and repulsive to both HSO_5^- and SO_5^{2-} . Moreover, SO_4^{2-} radicals may combine with hydroxyl anions in alkaline environment to form the weaker

oxidant, HO^\cdot ($\text{SO}_4^{2-} + \text{HO}^\cdot \rightarrow \text{SO}_4^{2-} + \text{HO}^\cdot$). The great efficiency of $\text{MnCo}_2\text{O}_{4.5}$ /PMS system at neutral and mild acidic pH suggested that this process can be employed for the treatment of ground and surface waters. This process also offers more advantages in comparison with Fenton processes, which are workable only in relatively limited pH ranges.³⁸

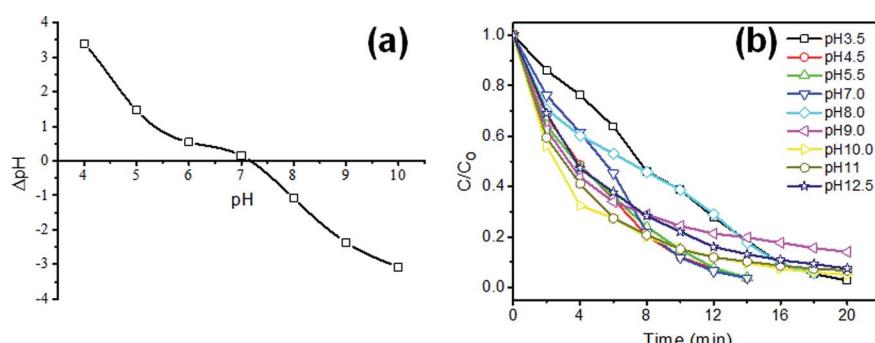


Fig. 5 (a) The isoelectric point of $\text{MnCo}_2\text{O}_{4.5}$ NPs at different pH. (b) Influence of initial solution pH on MB degradation in $\text{MnCo}_2\text{O}_{4.5}$ /PMS/MB system. Initial conditions: $[\text{MB}]_0 = 20 \text{ mg L}^{-1}$; $[\text{Oxone}]_0 = 500 \text{ mg L}^{-1}$; $[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}$; $T = 25^\circ\text{C}$.



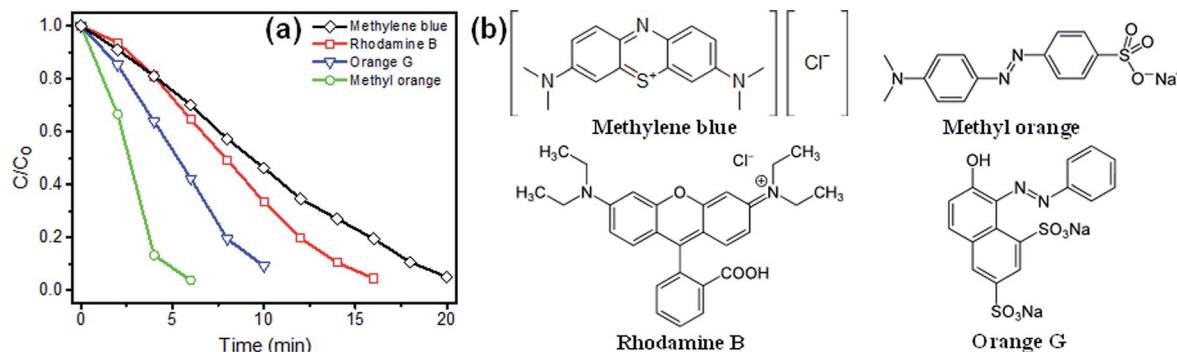


Fig. 6 (a) The degradation of different dyes by $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system. Initial conditions: $[\text{dye}]_0 = 20 \text{ mg L}^{-1}$; $[\text{Oxone}]_0 = 500 \text{ mg L}^{-1}$; $[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}$; $T = 25^\circ\text{C}$; unadjusted pH. (b) The molecular structure of MB, RhB, OG, and MO.

3.2.4. The degradation of different dyes. The versatility of $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system was examined for the degradation of various types of dyes. As shown in Fig. 6a, the degradation decreased following the order of MO > OG > RhB > MB. Under the same initial conditions ($\text{Oxone } 500 \text{ mg L}^{-1}$, $\text{MnCo}_2\text{O}_{4.5} 20 \text{ mg L}^{-1}$, $\text{MB } 20 \text{ mg L}^{-1}$, unadjusted pH, $T = 25^\circ\text{C}$), the degradation efficiencies of MB, RhB, OG, and MO were 96.7% (after 20 min), 95.6% (after 16 min), 91.0% (after 10 min), and 96.3% (after 6 min), respectively. The difference in degradation yield of these four dyes was due to their differences in molecular structure (Fig. 6b) and degradation mechanism.³⁹ The azo dyes (MO and OG) contain color centers ($-\text{N}=\text{N}-$), in which π -bondings are easily dissociated to form naphthalene ring.⁴⁰ The xanthene dye (RhB) is disrupted through the degradation of its conjugated structure and subsequent N-demethylation.⁴¹ The degradation of the thiazine dye (MB) first occurred on aromatic ring-bonded nitrogen atoms due to the presence of electron-donating groups ($-\text{N}(\text{CH}_3)_2$).⁴⁰ All of these organic dyes can be degraded in $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system, indicating that this catalyst can be employed in a wide range of organic pollutant environments.

3.2.5. Influence of reaction sequence. The sequence of adding MB dye, $\text{MnCo}_2\text{O}_{4.5}$ catalyst, and Oxone into reaction flask (reaction sequence) might have an impact on the degradation yield. In this work, six possible reaction sequences were carefully examined (other initial conditions were same) and the obtained result is shown in Fig. S1.[†] As seen, the degradation obtained for these sequences were not significantly different. The MB- $\text{MnCo}_2\text{O}_{4.5}$ -PMS sequence was selected for further experiments because of its high degradation efficiency and minimal self-decomposition of PMS caused by sonication.

3.2.6. Influence of coexisting anions. Real wastewater sources are commonly encountered with different coexisting anions (water matrix species) such as chloride (Cl^-), nitrate (NO_3^-), bicarbonate (HCO_3^-) and carbonate (CO_3^{2-}) at ppm concentration, and their presence might affect the degradation process. The influence of these anions (in the concentration range of 0–20 mM) has been investigated and the results are shown in Fig. 7.

As shown in Fig. 7a, the inhibition of MB degradation could be observed at a concentration of Cl^- as low as 1 mM. The γ_{20}

dramatically decreased from 95.8 to 35.2% when Cl^- concentration was increased from 1 to 10 mM, but gradually decreased to 30.6 and 27.1 when Cl^- concentration was subsequently increased to 15 and 20 mM, respectively. These results show that Cl^- strongly inhibit $\text{SO}_4^{\cdot-}$ free radicals, and the inhibition mechanism can be explained as following: $\text{SO}_4^{\cdot-}/\text{SO}_4^{2-}$ redox couple has higher reduction potential ($E^\circ = 2.5 \text{ V}$) as compared to that of $\text{Cl}_2/2\text{Cl}^-$ ($E^\circ = 1.4 \text{ V}$) and HOCl/Cl^- ($E^\circ = 1.48 \text{ V}$). Therefore, $\text{SO}_4^{\cdot-}$ can oxidize Cl^- ions to form weaker oxidants such as Cl^\cdot , ClO^\cdot , and Cl_2^\cdot as described by Table S1.[†] The influence of bicarbonate on MB degradation is shown in Fig. 7b. At a very low concentration of HCO_3^- (1 mM), MB seems to degrade more quickly. If HCO_3^- was employed at higher concentrations, MB degradation was retarded, which obviously demonstrated the inhibition role of HCO_3^- towards PMS. It is not clear why the γ_{20} increased from 74.3 to 78.5% when HCO_3^- concentration increased from 5 to 15 mM. However, at the highest concentration of HCO_3^- examined in this work (20 mM), the γ_{20} reached only 58.6%. The involved reactions between HCO_3^- and PMS are also represented in Table S1.[†] The influence of CO_3^{2-} anions was relatively similar to that of HCO_3^- (Fig. 7c). In contrast, the presence of NO_3^- anions seemed not to affect MB degradation yield (Fig. S2[†]), indicating that free radicals were not inhibited by them. Fig. 7d compares the degradation of MB upon employing different coexisting anions at the same concentration (20 mM), and Fig. 7e visualizes the influence of different coexisting anions at various concentrations to the γ_{20} of MB. The γ_{20} was 96.7, 81.8, 58.6, and 27.1% for NO_3^- , CO_3^{2-} , HCO_3^- , and Cl^- , respectively. These results point out that the concentration of Cl^- , HCO_3^- , and CO_3^{2-} ions must be carefully controlled within allowed limitation before Oxone treatment for optimal performance.

3.2.7. Comparison of different oxidants (PMS, PS, and H_2O_2). In addition to PMS, PS and H_2O_2 are also employed as oxidants in AOPs.⁴² Herein we have compared the capability of $\text{MnCo}_2\text{O}_{4.5}$ NPs as a catalytic activator for these oxidants in MB degradation. As shown in Fig. 8, the γ_{20} of $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$, $\text{MnCo}_2\text{O}_{4.5}/\text{PS}$, and $\text{MnCo}_2\text{O}_{4.5}/\text{H}_2\text{O}_2$ systems was 96.7, 48.5, and 39.4%, correspondingly. The stronger oxidation power of the first two systems might be associated with the fact that O–O bonding energy in PMS and PS is smaller than that of H_2O_2 (Table



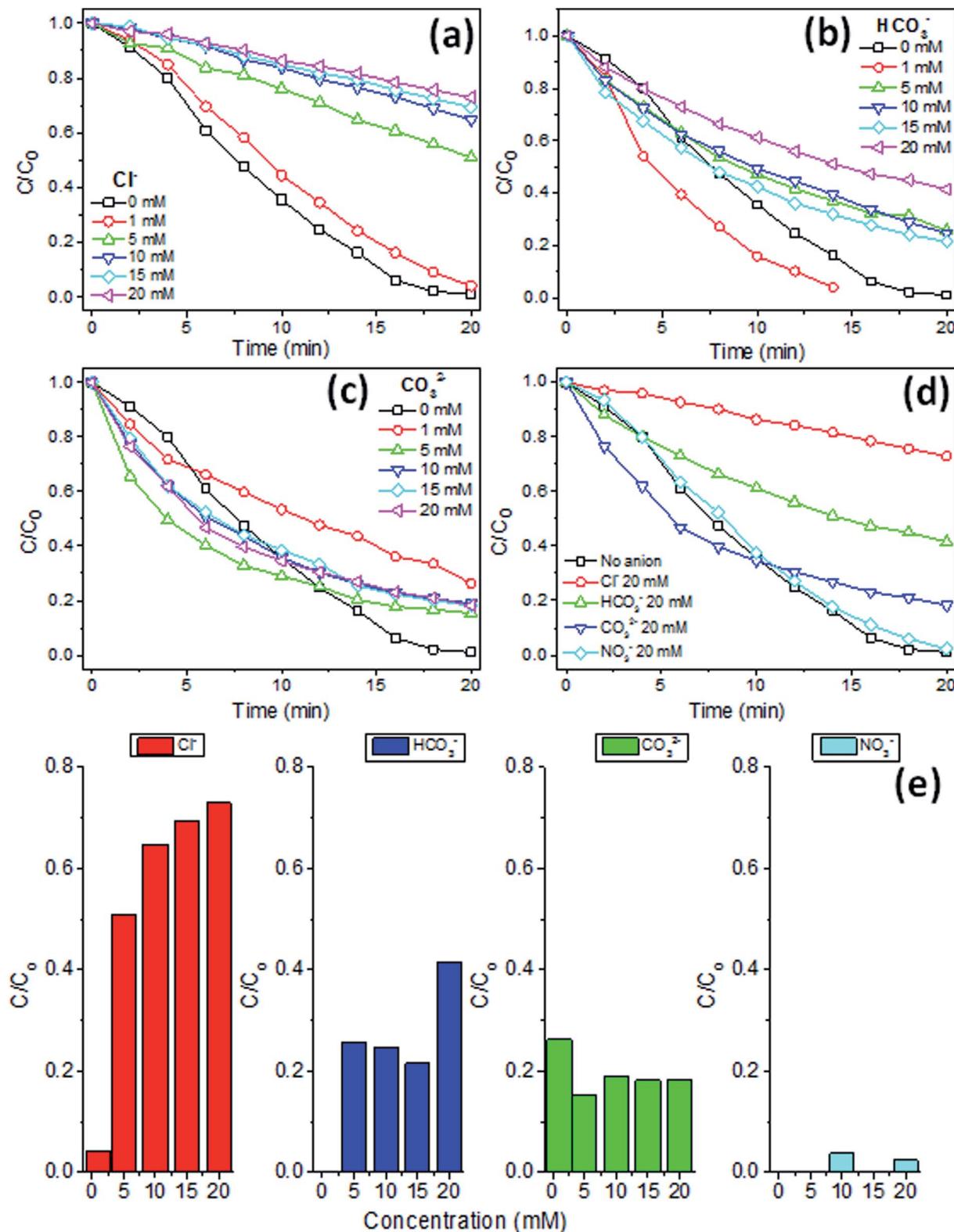


Fig. 7 Influence of different coexisting anions on MB degradation in $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system. (a) Cl^- ; (b) HCO_3^- ; (c) CO_3^{2-} ; (d) the degradation of MB in the presence of different coexisting anions at a concentration of 20 mM. (e) Comparison of γ_{20} in the presence of different coexisting anions at various concentrations. Initial conditions: $[\text{MB}]_0 = 20 \text{ mg L}^{-1}$; $[\text{Oxone}]_0 = 500 \text{ mg L}^{-1}$; $[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}$; $T = 25^\circ\text{C}$; unadjusted pH.

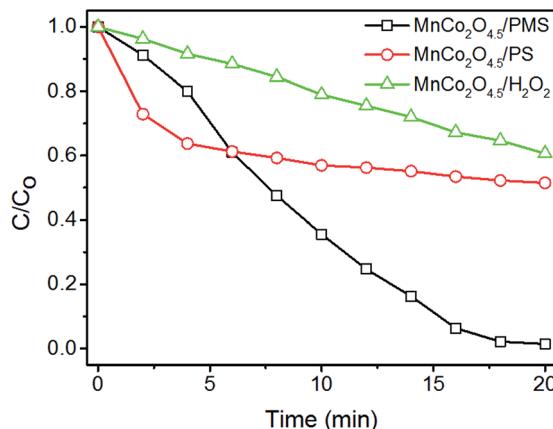


Fig. 8 Comparison of different oxidants on MB degradation using $\text{MnCo}_2\text{O}_{4.5}$ catalyst. Initial conditions: $[\text{MB}]_0 = 20 \text{ mg L}^{-1}$; $[\text{oxidant}]_0 = 500 \text{ mg L}^{-1}$; $[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}$; $T = 25^\circ\text{C}$; unadjusted pH.

$\text{S}2^\ddagger$), which makes the activation of PMS and PS easier in terms of energy. Although it is not clear why the degradation efficiency of PMS is higher than that of PS, one possible cause is the activation of PMS is more convenient due to its asymmetric structure.

3.2.8. Identification of contribution of different radicals based on quenching experiments. In order to explore activation mechanism, it is necessary to find the dominant species which cause the activation of PMS. The heterogeneous activation of PMS, particularly by transition metal oxides, is known to be strongly associated with the generation of $\text{SO}_4^{\cdot-}$, HO^{\cdot} and $\text{SO}_5^{\cdot-}$ free radicals.² MB degradation due to $\text{SO}_5^{\cdot-}$ can be neglected because the standard reduction potential of $\text{SO}_5^{\cdot-}$ at neutral solution is only 1.1 V ,¹⁵ much lower than that of $\text{SO}_4^{\cdot-}$ and HO^{\cdot} . In this work, the role of $\text{SO}_4^{\cdot-}$ and HO^{\cdot} free radicals in MB degradation has been determined through quenching experiments using EtOH, phenol, and TBA quenchers. The rate constant of EtOH with HO^{\cdot} ($1.2\text{--}2.8 \times 10^9 \text{ M}^{-1} \text{ s}^{-1}$) is $36\text{--}75$ times greater than that of $\text{SO}_4^{\cdot-}$ ($1.6\text{--}7.7 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$).⁴³ The rate constants of phenol with both HO^{\cdot} and $\text{SO}_4^{\cdot-}$ are very high (6.6×10^9 and $8.8 \times 10^9 \text{ M}^{-1} \text{ s}^{-1}$, correspondingly).⁴⁴ The rate constant of TBA with HO^{\cdot} ($3.8\text{--}7.6 \times 10^8 \text{ M}^{-1} \text{ s}^{-1}$) is 835 to 950 times greater than that of $\text{SO}_4^{\cdot-}$ ($4.0\text{--}9.1 \times 10^5 \text{ M}^{-1} \text{ s}^{-1}$).

Therefore, EtOH and phenol have been employed as scavenger for both $\text{SO}_4^{\cdot-}$ and HO^{\cdot} while TBA is scavenger for HO^{\cdot} . Fig. 9 shows the influence of different quenchers on MB degradation in a wide range of quencher concentration (0–200 mM). Without any quenchers (EtOH, phenol and TBA), 96.7% MB was eliminated after 20 min. In the presence of quenchers (EtOH and TBA) of increasing concentration (0–200 mM), the MB degradation yield decreased to 85.2% and 87.3% in 20 min, correspondingly. Whereas, with the increase of phenol concentration from 0 to 100 mM, MB degradation yield decreased to 18.0% in 20 min. These results point out that sulfate and hydroxyl radicals are generated from $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system, and sulfate radicals play a major role in MB degradation.

3.2.9. Cobalt leaching test. Metal leaching is a major concern in heterogeneous catalysts because metallic species such as Co^{2+} have been known as excellent PMS activator.^{2,45} Our own experiment also demonstrated the excellent catalytic activity of Co^{2+} (Fig. 3a) therefore the leaching of cobalt from corresponding heterogeneous catalysts must be carefully taken into account. Our ICP-MS analysis showed that the leaching concentration (percentage) of Co^{2+} and Mn^{2+} after the first run was 1.669 and 0.467 mg L^{-1} , respectively. The reported Co leaching concentration was 0.662 mg L^{-1} for $\text{CoFe}_2\text{O}_4/\text{TiO}_2$,⁹ 0.928 mg L^{-1} for CoFe_2O_4 ,⁹ 1.167 mg L^{-1} for SrCoO_3 ,²⁹ 2.45 mg L^{-1} for $\text{CoFe}_2\text{O}_4/\text{SBA-15}$,¹⁹ *ca.* 9.0 mg L^{-1} for Co-BTC⁴⁶ and 9.8 mg L^{-1} for $\text{PrBaCo}_2\text{O}_{5+\delta}$.⁴⁷ Although the measured leached Co concentration in our work is acceptable as compared to previous reports, future efforts should be made in order to meet environmental standard for surface water. A potential solution is to grow $\text{MnCo}_2\text{O}_{4.5}$ NPs on a suitable support such as carbon nanotube, which can simultaneously improve catalytic performance and suppress leaching ions.

3.2.10. Proposed mechanism of PMS activation on $\text{MnCo}_2\text{O}_{4.5}$ NPs. Based on the above results and previous researches, a mechanism for PMS activation on $\text{MnCo}_2\text{O}_{4.5}$ NPs is proposed as following: In the presence of $\text{MnCo}_2\text{O}_{4.5}$ NPs, PMS is activated to form $\text{SO}_5^{\cdot-}$ and $\text{SO}_4^{\cdot-}$ radicals (eqn (6) and (8)). Subsequently, $\cdot\text{OH}$ radicals are generated due to the reaction of $\text{SO}_5^{\cdot-}$ and hydroxylated metallic sites ($\text{M}^{3+}\text{--H}_3\text{O}^+$) (eqn (10) and (11)) or the reaction between $\text{SO}_4^{\cdot-}$ and H_2O or OH^- (eqn (12) and (13)). The two latter equations could be also considered as the quenching reactions of $\text{SO}_4^{\cdot-}$.⁴⁸ Mn^{2+} and

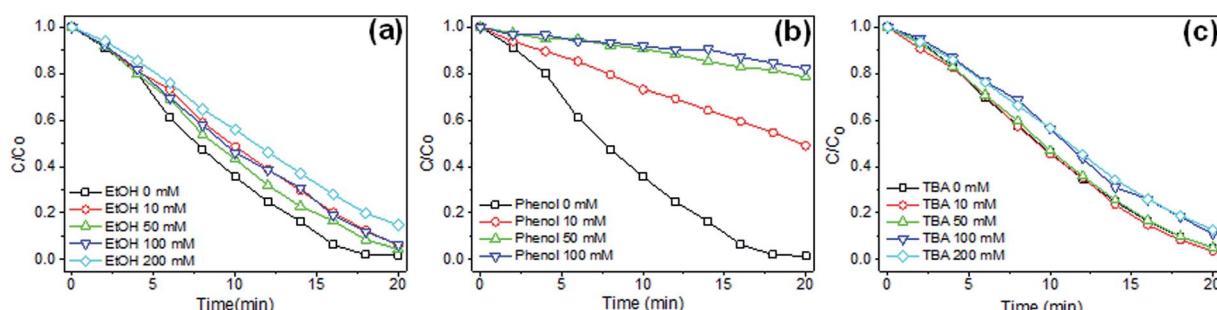
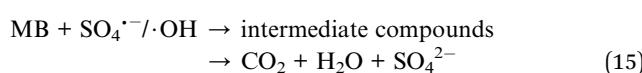
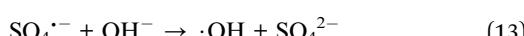
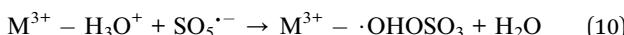
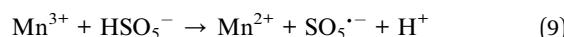
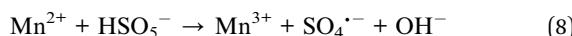
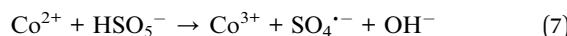
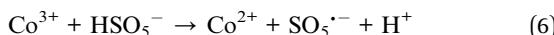


Fig. 9 Influence of different radical scavengers on MB degradation in $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system. (a) EtOH; (b) phenol; (c) TBA. Initial concentrations: $[\text{MB}]_0 = 20 \text{ mg L}^{-1}$, $[\text{Oxone}]_0 = 500 \text{ mg L}^{-1}$, $[\text{MnCo}_2\text{O}_{4.5}]_0 = 20 \text{ mg L}^{-1}$, unadjusted pH, $T = 25^\circ\text{C}$.



Co^{3+} can be regenerated as described by eqn (7) and (9). The generated radicals attack dye molecules, open the aromatic rings and form intermediate products (mostly oxalic and small-weight carboxylic acids) which are ultimately mineralized into CO_2 , H_2O and SO_4^{2-} (eqn (15)).



3.2.11. Comparison with other spinel catalysts. Table 1 compares the performance of several spinel-based catalysts/PMS systems for MB degradation. Our $\text{MnCo}_2\text{O}_{4.5}$ /PMS system exhibited a high degradation efficiency of 96.7% after 20 min, and 100% after 25 min (data not shown here), while for other systems it usually requires 30–60 min to achieve 100% degradation of MB. In addition, the amount of $\text{MnCo}_2\text{O}_{4.5}$ catalyst employed in our experiments was 20 mg L^{-1} , much smaller than that of previous studies ($50\text{--}300 \text{ mg L}^{-1}$). This comparison showed that the performance of $\text{MnCo}_2\text{O}_{4.5}$ /PMS system is excellent and promising for practical applications in environmental remediation.

3.2.12. The reusability of as-prepared $\text{MnCo}_2\text{O}_{4.5}$ catalyst. The reusability and stability of $\text{MnCo}_2\text{O}_{4.5}$ catalyst are important aspects for practical applications, and were evaluated for

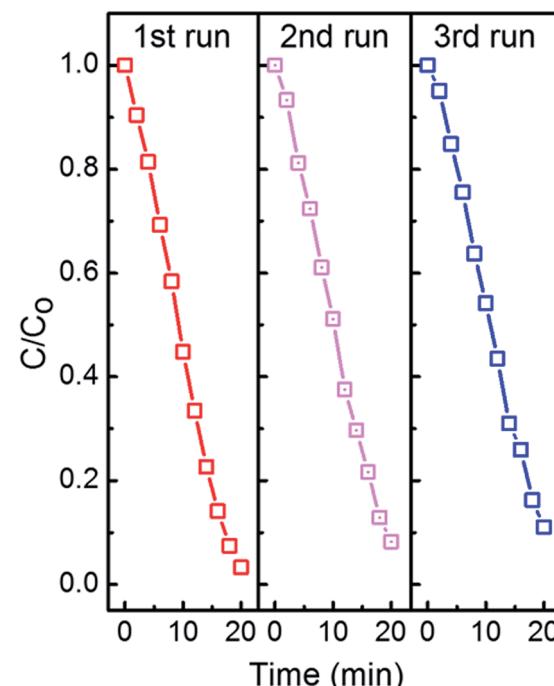


Fig. 10 MB degradation in three consecutive runs by $\text{MnCo}_2\text{O}_{4.5}$ /PMS system. Initial conditions: $[\text{MB}] = 20 \text{ mg L}^{-1}$, $[\text{MnCo}_2\text{O}_{4.5}] = 20 \text{ mg L}^{-1}$, $[\text{Oxone}] = 500 \text{ mg L}^{-1}$, $T = 25^\circ\text{C}$, unadjusted pH.

three successive runs. After each run, $\text{MnCo}_2\text{O}_{4.5}$ NPs were washed with ethanol and redispersed for the next catalysis reaction. As observed in Fig. 10, $\text{MnCo}_2\text{O}_{4.5}$ NPs sustained an excellent γ_{20} of 89.9% at the third run, indicating its very good reusability and high stability.

3.2.13. Evaluation of COD removal efficiency for real wastewater treatment application. Real wastewater such as leachate or textile discharges is a complex environment contains a diverse range of organic compound at high concentration, and COD is a parameter of major concern when treating civilian and industrial discharges. Here the COD removal efficiency of $\text{MnCo}_2\text{O}_{4.5}$ /PMS system was evaluated for three different sources: MB solution (20 mg L^{-1}), textile wastewater, and leachate. The textile wastewater was collected at Hopex Textile Co., Ltd. in Hai Duong Province, Vietnam. Random

Table 1 Comparison of various spinel-based catalysts as PMS activator for MB degradation

Catalyst	Synthesis method	Reaction conditions	Performance	Reference
$\text{Co}_3\text{O}_4\text{-Bi}_2\text{O}_3$	Co-precipitation	$[\text{MB}] = 7.5 \text{ mg L}^{-1}$; $[\text{catalyst}] = 50 \text{ mg L}^{-1}$; $[\text{PMS}] = 0.5 \text{ mM}$; $\text{pH} = 3.4$; $T = 25^\circ\text{C}$	98% (10 min) Co leaching: $43 \mu\text{g L}^{-1}$ (acidic pH) and $31 \mu\text{g L}^{-1}$ (pH 7.0)	49
CoMoO_4	Hydrothermal	$[\text{MB}] = 100 \text{ mg L}^{-1}$; $[\text{catalyst}] = 100 \text{ mg L}^{-1}$; $[\text{Oxone}] = 2 \text{ mM}$; unadjusted pH; $T = 25^\circ\text{C}$	100% (40 min)	50
$\text{Fe}_3\text{O}_4/\text{Mn}_3\text{O}_4/\text{rGO}$	Co-precipitation	$[\text{MB}] = 50 \text{ mg L}^{-1}$; $[\text{catalyst}] = 100 \text{ mg L}^{-1}$; $[\text{Oxone}] = 300 \text{ mg L}^{-1}$; $\text{pH} = 7.0$; $T = 25^\circ\text{C}$	98.8% (30 min)	21
$\text{Fe}_3\text{O}_4@\text{MnO}_2$	Co-precipitation	$[\text{MB}] = 30 \text{ mg L}^{-1}$; $[\text{catalyst}] = 300 \text{ mg L}^{-1}$; $[\text{PMS}] = 20 \text{ mM}$; $\text{pH} = 7.0$; $T = 25^\circ\text{C}$	99.8% (30 min)	51
$\text{CuFe}_2\text{O}_4/\text{AC}$	Co-precipitation/calcination	$[\text{MB}] = 20 \text{ mg L}^{-1}$; $[\text{catalyst}] = 200 \text{ mg L}^{-1}$; $[\text{PMS}]/[\text{dye}] = 200 \text{ mg L}^{-1}$; $\text{pH} = 5.0$; $T = 25^\circ\text{C}$	100% (60 min)	52
$\text{MnCo}_2\text{O}_{4.5}$	Hydrothermal/calcination	$[\text{MB}] = 20 \text{ mg L}^{-1}$; $[\text{MnCo}_2\text{O}_{4.5}] = 20 \text{ mg L}^{-1}$; $[\text{Oxone}] = 500 \text{ mg L}^{-1}$; $T = 30^\circ\text{C}$; $\text{pH} = 3.5$	96.7% (20 min); 100% (25 min)	This work



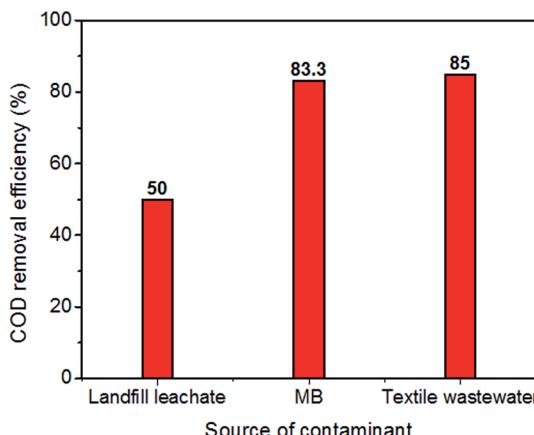


Fig. 11 COD removal efficiency of $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system for different sources of contaminant.

leachate samples were collected at Quang Hanh co-disposal landfill in Quang Ninh Province, Vietnam. The results are presented in Fig. 11. The input COD of MB solution, textile wastewater, and leachate was determined to be 144, 416, and 339.2 mg L^{-1} , respectively. After treatment with $\text{MnCo}_2\text{O}_{4.5}/\text{PMS}$ system for 20 min, the measured output COD was 22.4, 54.4, and 169.6 mg L^{-1} , correspondingly. The calculated COD removal efficiency for MB, textile wastewater, and leachate was 83.3, 86.9, and 50%, indicating dissolved organic matters have been significantly degraded in these systems. These results show that employing $\text{MnCo}_2\text{O}_{4.5}$ to activate PMS could be an alternative method to cope with real wastewaters.

4. Conclusions

In summary, well-crystallized, pure-phase spinel $\text{MnCo}_2\text{O}_{4.5}$ NPs are successfully synthesized *via* a new hydrothermal/calcination protocol. The synthesized $\text{MnCo}_2\text{O}_{4.5}$ NPs show the highest catalytic activity for PMS activation among all tested heterogeneous catalyst (MnO_2 , CoO) and outperform other spinel-based heterogeneous catalysts for dye degradation, which can be credited to its organic-free synthesis route and the synergy between Mn^{2+} and Co^{3+} . The catalyst dosage, initial Oxone and MB concentrations, reaction temperature, initial solution pH, and coexisting anions are found to influence the degradation of MB. The presence of HCO_3^- , CO_3^{2-} and particularly Cl^- ions reduces the MB degradation rate, while NO_3^- ions do not inhibit the generation sulfate radicals. A complex mechanism for the activation of PMS is proposed, whereas the generation of SO_4^{2-} , rather than $\cdot\text{OH}$, is primarily responsible for MB degradation. $\text{MnCo}_2\text{O}_{4.5}$ NPs can be separated, reused and displayed good catalytic performance after three consecutive cycles of use with small deterioration. This work thus suggests a promising route for the preparation of $\text{MnCo}_2\text{O}_{4.5}$ NPs with high catalytic activity for PMS activation for efficient degradation of organic dyes.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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