


Cite this: *RSC Adv.*, 2020, 10, 2812

Probing the interaction effects of metal ions in $\text{Mn}_x\text{Fe}_{(3-x)}\text{O}_4$ on arsenite oxidation and adsorption

Linda Ouma ^{*ab} and Augustine Ofomaja^a

Wastewater treatment is still a global concern and materials capable of pollutant sequestration continue to be improved in a bid to ensure water reusability and curb water shortages. Some of the most promising materials so far are nanosized materials because of their unique properties and the ease of manipulation to improve their properties. In this work we investigated the effects of varying $\text{Fe}^{3+} : \text{Fe}^{2+}$ ratios in magnetite nanoparticles and the influence of manganese doping. Diffraction measurements indicated that the manganese introduced into the magnetite matrix displaced some Fe atoms resulting in the formation of a uniform phase matching the card data for magnetite with no additional manganese phases being formed. XPS confirmed the presence of manganese on the surface of the doped nanomaterials and that both As(III) and As(V) were bound on the adsorbent surface. The central composite design (CCD) of response surface methodology (RSM) was used to determine the effects the nanoparticle compositions had on As(III) adsorption and oxidation. A quadratic equation was used to model the experimental data with a correlation coefficient close to unity indicating that the model was a good fit for the data. The interaction between Fe^{3+} and Mn had a positive influence in the reduction of As(III) in solution while $\text{Fe}^{3+}/\text{Fe}^{2+}$ interactions had antagonistic effects and the Fe^{2+}/Mn interactions were found to be insignificant. Increasing the amounts of Fe^{3+} and manganese therefore resulted in the highest reduction in As(III) concentration.

Received 15th November 2019
Accepted 27th December 2019

DOI: 10.1039/c9ra09543h

rsc.li/rsc-advances

Introduction

Nano size adsorbents offer a high surface area for adsorption and are thus favorable over conventional adsorbent materials however, their recovery after water treatment may be time consuming due to their small sizes allowing for dispersion in the treated water.^{1,2} Magnetic nanoparticles (MNP) however provide a unique advantage in that they can easily be recovered from treated water by application of an external magnetic field allowing them to be separated without the use of expensive and energy intensive processes.^{3,4} Magnetite nanoparticles are superparamagnetic and consist of both ferric and ferrous ions which allow them to possess unique properties.^{5,6} They have been applied for the remediation of arsenic,^{4,7,8} lead,^{9,10} chromium,⁹⁻¹¹ cerium,¹² mercury,¹³ thorium¹⁴ and cadmium¹⁰ contaminated water. Nanoparticles can be modified to improve their efficiency by the introduction of other ions into their structure allowing them to have a wider range of applications.¹⁵ Manganese oxides, like iron oxides are naturally occurring and have high affinities for trace metals as compared to other adsorbents.¹⁶⁻¹⁸ Manganese doping on magnetite provides

higher magnetic susceptibility than pure magnetite due to the presence of five single unpaired electrons of Mn^{2+} .¹⁹ A material consisting of both oxides therefore yields magnetic particles with better sorption capacities for heavy metal pollutants.²⁰

In the preparation of manganese doped magnetite nanoparticles (Mn MNP), Warner *et al.*, (2012), confirmed that Mn presented no significant changes in the particle size of magnetite nanoparticles and XRD data revealed it was incorporated into the ferrite structure.²⁰ Upon annealing of mixed iron and manganese oxides, Lai *et al.*, (2004) observed that either Mn(III) occupied Fe sites or Fe(III) occupied Mn sites where iron and manganese oxides are a majority respectively. They were able to determine the concentration window where phase separation occurred and observed a structural progression from spinel to bixbyite and a decrease in superparamagnetism as Mn concentration increased.²¹ Warner and coworkers applied the prepared nanoparticles in the sorption of heavy metals from water with an increase in the analyte collection and retention as the doping levels increased.²⁰ Gibbons and Gagnon, (2011) observed that ferric containing water treatment residuals (WTRs) had the greatest amount of arsenic adsorbed on a molar basis and described arsenate removal by ferric ions as a surface complexation between ferric ions and arsenate ions.²² Later Ociński *et al.*, (2016) carried out adsorption of arsenites and arsenates onto iron and manganese containing WTRs, and showed that manganese acts not only as

^aBiosorption and Water Treatment Research Laboratory, Department of Chemistry, Vaal University of Technology, Vanderbijlpark, South Africa

^bDepartment of Science, Technology and Engineering, Kibabii University, Bungoma, Kenya. E-mail: oumachieng@gmail.com



an oxidant but also increases adsorption sites as manganese oxide is reduced to divalent manganese and increases the positive charge on the surface allowing for an increase in arsenate adsorption.²³

Optimization of the adsorbent synthesis parameters is therefore necessary for the efficient application of manganese doped magnetite nanoparticles (Mn MNP) adsorbent. Response surface methodology (RSM) using central composite experimental design is an appropriate technique to obtain the best conditions for the adsorbent synthesis.²⁴ Response surface methodology (RSM) is a multivariate statistical technique used in analytical optimization. It employs a collection of mathematical and statistical techniques to fit experimental data to a polynomial equation.²⁵ Central composite design (CCD) is a widely used RSM approach which is based on a second order polynomial design that is used to understand the interactive effects of variables on the studied responses.²⁶ The RSM enables determination of the optimum operating conditions in an effective manner and evaluates the effect of interaction of multivariable systems using statistical methods. This is more advantageous than a one variable at a time (OVAT) experimental design which is time consuming and does not cater for the interactive effects of variables.^{27,28}

This article discusses the optimization of one pot synthesis of Mn MNP using response surface methodology. The optimization was based on varying the amounts of ferric, ferrous and manganese ions in the adsorbent for optimal adsorption efficiency. This work was carried out with an aim to study the effect of interaction between the different constituents of the adsorbent material and their overall effect on arsenic adsorption and post adsorption speciation of arsenic species as could be influenced by the adsorbent composition. The responses analyzed were adsorption efficiency and the dominant arsenic species in the supernatant after the oxidation coupled adsorption process.

Materials and methods

Materials

Ammonium hydroxide (NH₄OH, 25%) was supplied by Labchem (South Africa). Hydrochloric acid (32%) and ferrous sulphate (FeSO₄·7H₂O > 98%) were supplied by Merck. Ferric chloride (FeCl₃·6H₂O > 99%), manganese sulphate (MnSO₄) and sodium arsenite (NaAsO₂ > 90%) were purchased from Sigma-Aldrich. All chemicals were used without any further purification. All syntheses were carried out under nitrogen atmosphere with vigorous stirring to ensure uniform dispersions.

Methods

Magnetite nanoparticles and manganese doped nanoparticles were synthesized in a one-pot reaction *via* co-precipitation method according to a previously reported procedure.²⁹ Briefly, deionized water was degassed and purged with nitrogen gas while stirring before the addition of pre-determined amounts ferric and ferrous salts. To the solution, ammonium hydroxide

was added and stirring continued till a black precipitate was formed. Manganese doped nanoparticles were synthesized by adding the requisite amount of manganese salt to the solution followed by precipitation with ammonium hydroxide. The prepared particles were washed with deionized water and ethanol and dried in a vacuum oven to prevent further oxidation. The adsorption of As(III) was carried out by dissolving sodium arsenite in deionized water to prepare a stock solution which was diluted further to prepare 100 mg L⁻¹ solutions for adsorption experiments. All adsorption experiments were carried out for 120 minutes at 298 K, adsorbent mass 5 g L⁻¹, adsorbate concentration of 100 mg L⁻¹ and agitation speed of 200 rpm.

Characterization

X-ray diffraction (XRD) spectra were recorded on a Bruker AXS D8 advanced diffractometer equipped with Cu K α (λ = 1.5418 Å) X-ray source. XRD was used to determine the crystallinity of the materials and confirm the formation of magnetite nanoparticles by comparing with reference data for synthetic magnetite. Particle sizes were calculated following Debye-Scherrer's equation to confirm the formation of nano-sized particles. X-ray photoelectron spectroscopy (XPS) studies were performed using XPS microprobe (PHI 5000 Scanning ESCA Microprobe ULVAC-PHI Inc). The concentration of arsenic species in solution was determined on a Thermo Fischer Scientific ICAP 7000 inductively coupled plasma-optical emission spectrometer (ICP-OES). As(III) concentration was determined after passing the solution through arsenic speciation cartridges (Metalsoft Centre, New Jersey, USA).^{30,31} The concentration of As(V) was determined as the difference between the concentrations of the raw and speciated solutions.

Experimental design and RSM optimization

Response surface methodology employing a central composite design (CCD) using Design Expert 11.0.5.0 software from Stat-Ease Inc (USA) was used to study the interactions and their effect on the responses. The factors studied were the amounts of ferric, ferrous and manganese salts used in the preparation of doped magnetite nanoparticles. Responses recorded were removal percentages of arsenite and speciation of arsenic species after adsorption. A combination of parameters (variables) that ensure effective As(III) removal and/or oxidation to As(V) are considered as most suitable due to the elimination of the most toxic arsenic species (As(III)). Each parameter investigated was varied as shown in Table 1. The total number of

Table 1 Variable levels for RSM experiment

Variables			Levels		
Actual	Alias	Units	−1	0	+1
Fe ³⁺	A	moles	0	0.012	0.024
Fe ²⁺	B	moles	0	0.008	0.016
Mn ²⁺	C	moles	0.0018	0.0036	0.0054



experiments was twenty consisting of eight factorial, six axial and six center points.²⁵

In each case, the data obtained was used to develop a mathematical model that best correlates the variables to the responses in the form of a quadratic polynomial equation (eqn (1)) where Y is the response, b_0 is the offset term, b_i is the linear effect, b_{ii} is the quadratic effect and b_{ij} is the interaction effect.

$$Y = b_0 + \sum_{i=1}^k b_i x_i + \sum_{i=1}^k b_{ii} x_i^2 + \sum_{i=1}^k \sum_{j>1}^k b_{ij} x_i x_j \quad (1)$$

Results and discussions

Characterization

The surface properties of magnetite nanoparticles and manganese doped nanoparticles were analyzed by Brunauer–Emmett–Teller (BET) surface area and isoelectric point determination. The surface area of the magnetite nanoparticles was improved while the pore sizes were reduced by the introduction of manganese (Table 2). The reduction in pore size was due to the incorporation of manganese in the magnetite lattice resulting in the formation of smaller pores while providing additional binding sites, resulting in increased adsorption capacities.³²

Iron atoms on magnetite surfaces coordinate with H_2O molecules which readily dissociate resulting in hydroxyl functionalized surfaces. These surface hydroxyl groups are amphoteric reacting with either acids or bases resulting in a near neutral pH_{PZC} (Table 2). MNP surfaces can either be negatively or positively charged depending on the pH of the solution. Below the isoelectric point, the surface is protonated leading to the formation of $\equiv Fe-OH_2^+$ resulting in a net positive charge. Above the isoelectric point, the surface hydroxyl groups are deprotonated forming $\equiv Fe-O^-$ surface groups.³³ The observed isoelectric point of pH 7.1 is in good agreement with reported values for magnetite particles reported between pH 6.5 and 6.8.^{33–36}

XRD

X-ray diffraction patterns were obtained to determine the effects of doping on magnetite nanoparticles crystal structure. The diffraction patterns for the synthesized magnetite nanoparticles and manganese doped nanoparticles matched card data for magnetite (JCPDS 19-0629) with no peaks corresponding to secondary phases or impurities being detected.³⁷ As doping levels increased, diffraction peaks were shifted to lower 2θ values since the radius of Mn^{2+} at 67 pm is larger than both Fe^{2+}

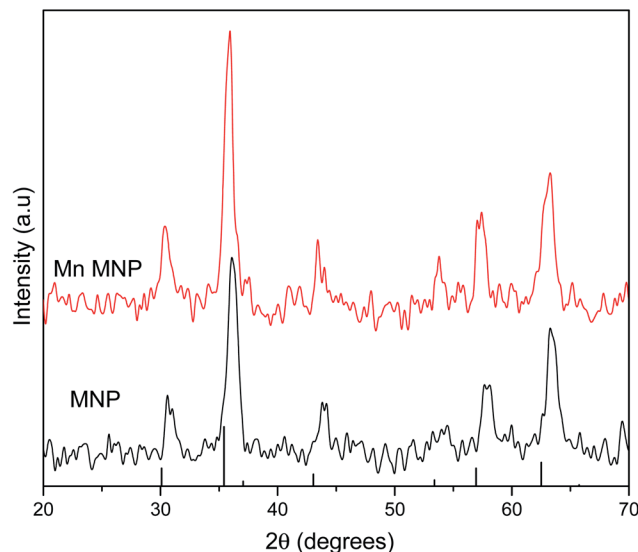


Fig. 1 XRD diffractograms of magnetite (MNP) and manganese doped magnetite (Mn MNP).

(61 pm) and Fe^{3+} (55 pm) therefore resulting in increased lattice distance as compared to magnetite.¹⁹ Bragg's law (eqn (2)) where n is an integer number of wavelengths, λ is the wavelength of the incident X-ray beam, d is the distance between atomic layers and θ is the angle of incidence or angle of scattering of the X-ray beam defines the relation between d and θ .

$$n\lambda = 2d \sin \theta \quad (2)$$

Since n and λ are constant for all the samples therefore as the distance between atomic layers increase, the angle of scattering decreases resulting in shifts to lower 2θ values as observed in the diffractograms in Fig. 1.

The lattice parameters were calculated to determine the effects of doping on the unit cell of magnetite. Magnetite crystals have a face-centered cubic pattern and the unit cell is characterised by a lattice parameter of 0.8396 nm (8.39 Å).^{38,39} The magnetite crystal is an O^{2-} face centered cubic lattice with Fe^{3+} occupying a $\frac{1}{2}$ of the tetrahedral interstices while $\frac{1}{8}$ of the octahedral interstices are occupied by a 1 : 1 mixture of Fe^{3+} and Fe^{2+} . The lattice parameter and cell volume values were calculated from diffraction data using Unit Cell Software.⁴⁰ The lattice parameter for the as-synthesized magnetite was 8.27 Å which is in good agreement with the value reported for nano-sized Fe_3O_4 .⁴¹ Manganese doping increased the lattice parameter due to the substitution of larger Mn into Fe lattices as observed by Liang *et al.*, (2014) that manganese substituted octahedral Fe in magnetite.^{42,43} The larger unit cell of $MnFe_2O_4$

Table 2 Surface properties for as-synthesized materials

	MNP	Mn MNP
Surface area ($m^2 g^{-1}$)	113.6	127.3
Pore volume ($cm^3 g^{-1}$)	0.6	0.4
Average pore diameter (nm)	20.0	10.9
pH_{PZC}	7.1	6.8

Table 3 Cell parameter values for synthesized samples

Sample	Lattice parameter a (Å)	Cell volume
Fe_3O_4	8.27	566.33
$Fe_{2.94}Mn_{0.06}O_4$	8.33	577.28



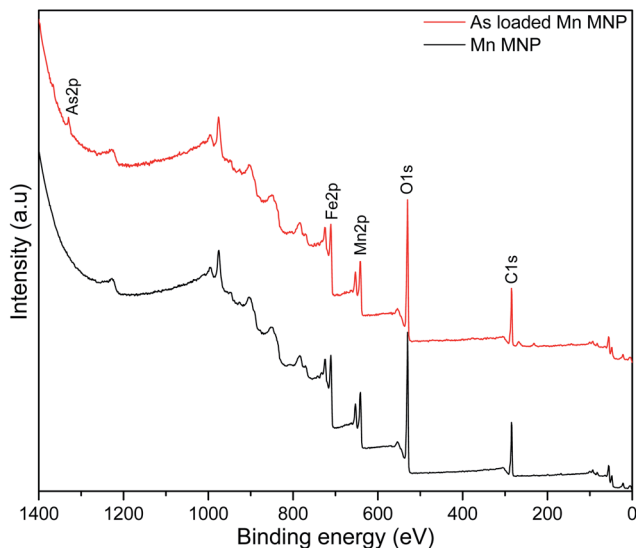


Fig. 2 XPS survey spectrum of Mn MNP and As loaded adsorbent.

(8.49–8.51 Å) therefore led to cell expansion and an increase in cell volume (Table 3).

XPS

X-ray photoelectron spectroscopic analysis was used to determine the surface composition of the nanoparticles and confirm manganese doping and arsenic adsorption. The survey

spectrum of the pristine nanoparticles (Fig. 2) indicates peaks at 710, 641, 530 and 284 eV corresponding to Fe, Mn, O and C respectively representing the predominant element on the adsorbent surface. After adsorption As2p peak at 1330 eV was introduced on the spectrum.⁴⁴

High resolution spectra were used to determine the states of the elements present on the adsorbent surface (Fig. 3). Fe2p peak (Fig. 3a) on the adsorbent were shifted to higher binding energies after adsorption as a result of arsenic complexation.²⁹ Mn2p peaks on the pristine and arsenic loaded adsorbent presented doublets due to spin coupling corresponding to Mn2p_{1/2} and Mn2p_{3/2} at ~653 eV and ~641 eV respectively (Fig. 3b). After arsenic adsorption a shift of the Mn2p_{1/2} (653.0–653.4 eV) and Mn 2p_{3/2} (641.1–641.6 eV) was due to the oxidation and binding of adsorbed arsenic.^{45,46} After adsorption the O1s peak (Fig. 3c) increased in intensity indicating an increase in oxygen atoms which was a likely result of the introduction of arsenic oxy-anions and water molecules.⁴⁷ Niu *et al.*, reported an increase in the O1s peak of titanate nanotubes after interaction with As(III) species while they observed no increase after interactions with As(V) species. The increase was attributed to strong interactions between As(III) and O atoms.⁴⁷

The As2p peak was deconvoluted to two peaks at 1329.4 and 1326.2 assigned to As³⁺ and As⁵⁺ respectively (Fig. 4). The presence of Fe and Mn on the adsorbent surface resulted in the oxidation of As³⁺ to As⁵⁺ resulting in the latter being more abundant on the surface after adsorption.

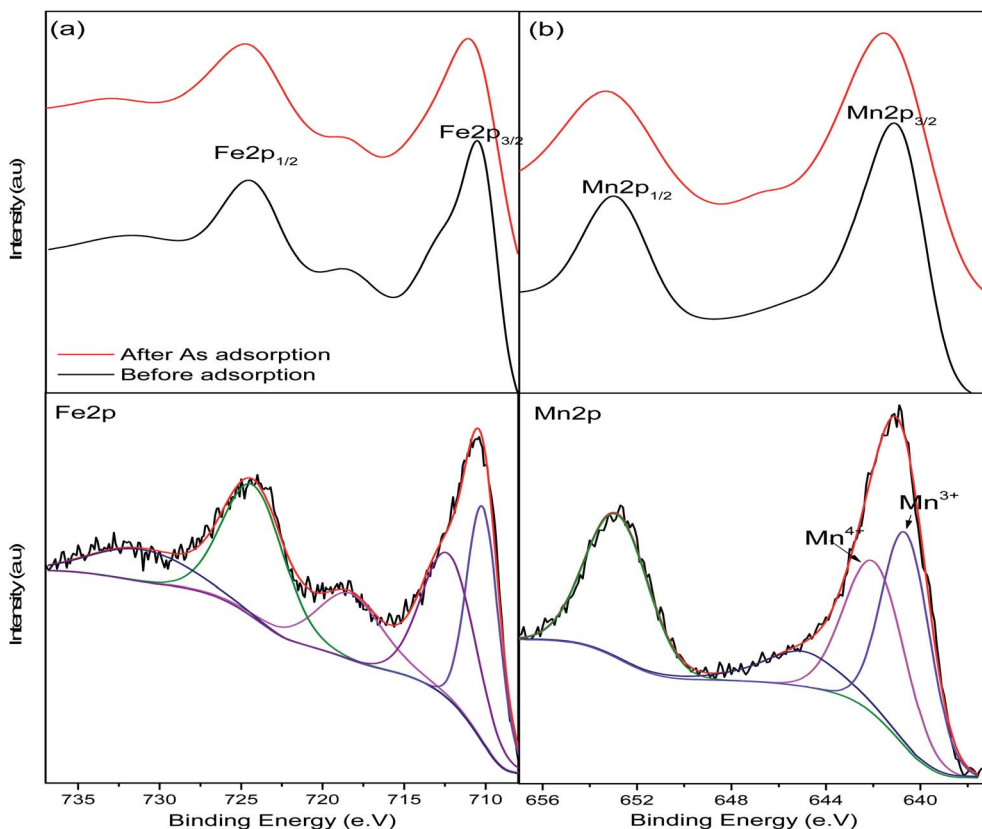


Fig. 3 XPS high-resolution spectra of the pristine and arsenic-loaded adsorbent showing deconvolutions of (a) Fe2p and (b) Mn2p peaks.

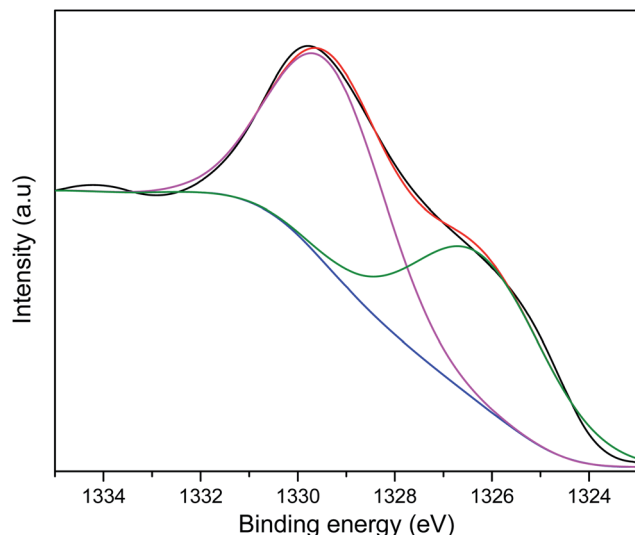


Fig. 4 High resolution XPS spectrum of the As 2p peak.

RSM

The experimental variables and results for the measured responses are presented in Table 4. The adsorption efficiency of As(III) in aqueous solution ranged between 93% and 73% depending on the adsorbent composition (Table 4). The highest removal occurred in the absence of Fe(II) consistent with results reported by Tian *et al.*, (2017), which demonstrated that minimal precipitation of As(III) occurred in the presence of Fe(II) under both oxic and anoxic conditions.⁴⁸ The speciation of the arsenic species in the solution after adsorption was analyzed to give an indication of the environmental safety of treated

effluent. The treated effluent had different amounts of As(III) and As(V) species depending on the adsorbent composition. Manganese doped magnetite should be optimally synthesized to ensure high oxidation of the more toxic As(III) to As(V) and should be efficient in adsorbing both arsenic species. The variations in performance recorded in Table 4 was not consistent with the changes in a single variable level indicating that the performance was most likely affected by the interaction between variables.

Statistical analysis

The response data was analyzed using the design expert software and the models were analyzed using analysis of variance (ANOVA). The parameters used to determine the accuracy of the model were Fisher variation ratio (*F*-value), probability value (*p*-value), lack of fit, coefficient of determination (*R*²), adjusted coefficient of determination (*R*_{Adj}²) and adequate precision. The *F*-value and *p*-value determine the statistical significance of the model or coefficient term. The larger the *F*-value and the smaller the *p*-value the more statistically significant the model term. Adequate precision is a signal to-noise ratio, it compares the range of the predicted values to the average prediction error and ratios greater than 4 indicate adequate model discrimination.⁴⁹

Table 5 shows the summary of analysis of variance (ANOVA) for various polynomial models relating the variables to the responses. The *p*-values < 0.05 indicate that quadratic model adequately described the interaction between the adsorbent composition and As(III) removal through oxidation and adsorption. Quadratic model also described adequately the arsenic speciation in the treated wastewater. A low *p*-value for lack of fit in all cases confirms that the quadratic models were a good fit to the experimental data. All the variables and their

Table 4 CCD for As(III) adsorption on different compositions of iron oxides and manganese doped oxides

Run	Variables			Responses	
	A: Fe(III) (moles)	B: Fe(II) (moles)	C: Mn (moles)	As(III) removal (%)	As(V) conc. after adsorption (%)
1	1	0	0.0018	83	55
2	2	0.024	0.0018	88	67
3	5	0	0.0054	82	72
4	6	0.024	0.0054	93	58
5	11	0.012	0.0036	77	50
6	9	0	0.0036	80	81
7	10	0.024	0.0036	83	71
8	13	0.012	0.0018	73	78
9	14	0.012	0.0054	75	47
10	15	0.012	0.0036	73	53
11	16	0.012	0.0036	73	32
12	17	0.012	0.0036	73	40
13	18	0.012	0.0036	73	11
14	19	0.012	0.0036	73	21
15	20	0.012	0.0036	73	15
16	3	0	0.0018	89	17
17	4	0.024	0.0018	86	17
18	7	0	0.0054	88	17
19	8	0.024	0.0054	92	16
20	12	0.012	0.0036	79	16



Table 5 ANOVA for response models

Response	Polynomial model	Source	Sum of squares	DF	Mean square	F-value	p-value
As(III) removal	Quadratic	Model	919.20	9.00	102.13	1389.46	<0.0001
		A-Fe ³⁺	38.89	1.00	38.89	529.10	<0.0001
		B-Fe ²⁺	9.30	1.00	9.30	126.59	<0.0001
		C-Mn	11.49	1.00	11.49	156.34	<0.0001
		AB	27.60	1.00	27.60	375.44	<0.0001
		AC	20.24	1.00	20.24	275.37	<0.0001
		BC	0.23	1.00	0.23	3.16	0.1058
		A ²	198.72	1.00	198.72	2703.45	<0.0001
		B ²	63.97	1.00	63.97	870.32	<0.0001
		C ²	2.93	1.00	2.93	39.84	<0.0001
		Residual	0.74	10.00	0.07		
		Lack of fit	0.60	5.00	0.12	4.40	0.0650
% As(V) after adsorption	Quadratic	Model	11 304.40	9.00	1256.04	828.05	<0.0001
		A-Fe ³⁺	184.76	1.00	184.76	121.80	<0.0001
		B-Fe ²⁺	127.17	1.00	127.17	83.84	<0.0001
		C-Mn	145.03	1.00	145.03	95.61	<0.0001
		AB	314.56	1.00	314.56	207.37	<0.0001
		AC	190.14	1.00	190.14	125.35	<0.0001
		BC	11.89	1.00	11.89	7.84	0.0188
		A ²	2942.60	1.00	2942.60	1939.93	<0.0001
		B ²	963.24	1.00	963.24	635.02	<0.0001
		C ²	5.97	1.00	5.97	3.94	0.0753
		Residual	15.17	10.00	1.52		
		Lack of fit	13.73	5.00	2.75	9.56	0.0135

linear interactions were significant in the proposed models except the interaction between Fe²⁺ and Mn²⁺ in As(III) removal and the quadratic effect of Mn²⁺ in As(V) formation as indicated by *p*-values > 0.05.

Arsenic(III) removal

The As(III) in solution was removed through a two-step process involving oxidation of some amount to As(V) by the manganese-doped magnetite. Subsequently, the As(V) and remaining As(III) were removed by adsorption onto the manganese-doped magnetite. By applying regression analysis, the reduction in As(III) concentration through combined oxidation and adsorption was correlated to the parameters in terms of coded values as described by eqn (3):

$$\% \text{As(III)}_{\text{removal}} = 73.1 + 2A + B + C - 1.9AB + 1.6AC + 8.5A^2 + 4.8B^2 + C^2 \quad (3)$$

Positive terms in the equation imply a direct proportionality between factor and response while negative terms imply inverse proportionality.²⁶ From the equation, all variables and significant interactions between variables have direct proportionality effects on the model except the interaction between Fe³⁺ and Fe²⁺ (AB). Validation of the model was done by plotting the predicted values against the experimental values and *R*_{Adj}² value of 0.9985 was obtained (Fig. 5), implying that the model accurately described the effect of interaction of the amount of the iron species and manganese in the adsorbent on As(III) concentration reduction. The signal to noise ratio (adequate

precision) was determined to be > 100 indicating an adequate signal.

The response surface plot depicting the interactive effects of components of the adsorbent on As(III) removal is shown in Fig. 6. The characteristics of magnetite which affect its adsorption properties are affected by the amounts of the Fe²⁺ and Fe³⁺ precursors used. The chemical precipitation of

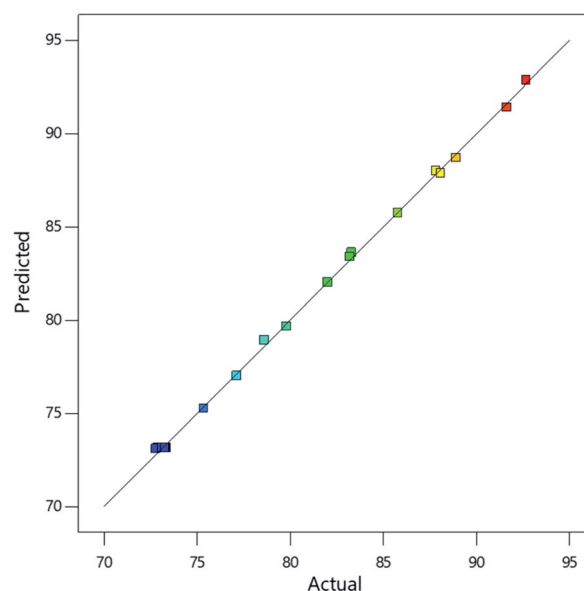
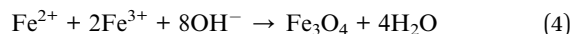


Fig. 5 Plot of model predictions against experimental findings for the percentage reduction of As(III) concentration.



dissolved iron salts to form magnetite in alkaline medium proceeds as illustrated in eqn (4).⁵⁰



Reaction conditions in this process significantly influence the dimensions and properties of the synthesized iron oxide nanoparticles. The ratio of $\text{Fe}^{3+}/\text{Fe}^{2+}$ affects the properties of the iron oxide products^{51,52} and subsequently the adsorptive properties of the nanoparticles tend to significantly vary with the synthesis conditions. For example, Roth *et al.*, (2015) observed an increase in the mean particle size of magnetite produced through co-precipitation when $\text{Fe}^{3+}/\text{Fe}^{2+}$ was decreased.⁵³ As explained by the classical nucleation theory where crystals in a co-precipitation reaction are formed through two distinct steps, nucleation from a supersaturated solution is followed by a slow growth of crystals. During the co-precipitation of magnetite from Fe^{3+} and Fe^{2+} , Fe^{3+} species form the primary nuclei formed from the supersaturated solution. Consequently, higher values of Fe^{3+} lead to an increase in the number of nuclei and therefore to more but smaller particles for an equal amount of iron salts.^{53,54} As a result, higher adsorption of $\text{As}(\text{III})$ is attained when Fe^{3+} concentration is highest (Fig. 6).

Also, in Fig. 6a the $\text{As}(\text{III})$ removal was highest at high $\text{Fe}(\text{III})$ dose of 0.02 M, at this condition the $\text{Fe}(\text{III})$ and $\text{Fe}(\text{II})$ concentrations were in the 2 : 1 stoichiometric proportion required to form magnetite as shown in eqn (4) since the amount of iron(II) was kept constant at a concentration of 0.01 M while OH^- concentration was always in excess at 0.14 M. The Mn^{2+} in solution then displaced some of the $\text{Fe}(\text{oct})^{2+}$ in magnetite to form manganese doped magnetite.^{43,55} Carvalho *et al.*, (2014) observed that manganese entered the magnetite lattice as Mn^{2+} and Mn^{3+} displacing Fe^{2+} and Fe^{3+} in octahedral and tetrahedral lattices respectively.⁵⁵ Comparably lower $\text{As}(\text{III})$ removal was obtained when $\text{Fe}(\text{III})$ dose was less than 0.02 M since low amount of magnetite was produced as $\text{Fe}(\text{III})$ was the limiting species.

In Fig. 6b, increasing the amount of $\text{Fe}(\text{II})$ from 0 to 0.02 M while the amount of $\text{Fe}(\text{III})$ was kept constant at 0.01 M did not lead to superior performance compared to that observed in Fig. 6a. This is because the concentration of $\text{Fe}(\text{III})$ was below the stoichiometric requirement compared to that of iron(II) which was in excess. Fig. 6 shows that manganese doped oxides of $\text{Fe}(\text{II})$ and $\text{Fe}(\text{III})$ had a nearly similar performance as manganese doped mixed oxide of $\text{Fe}(\text{II})$ and $\text{Fe}(\text{III})$ (manganese doped magnetite). However, manganese doped magnetite is advantageous due to its magnetic properties which allow for a simple post-adsorption separation.⁵⁶

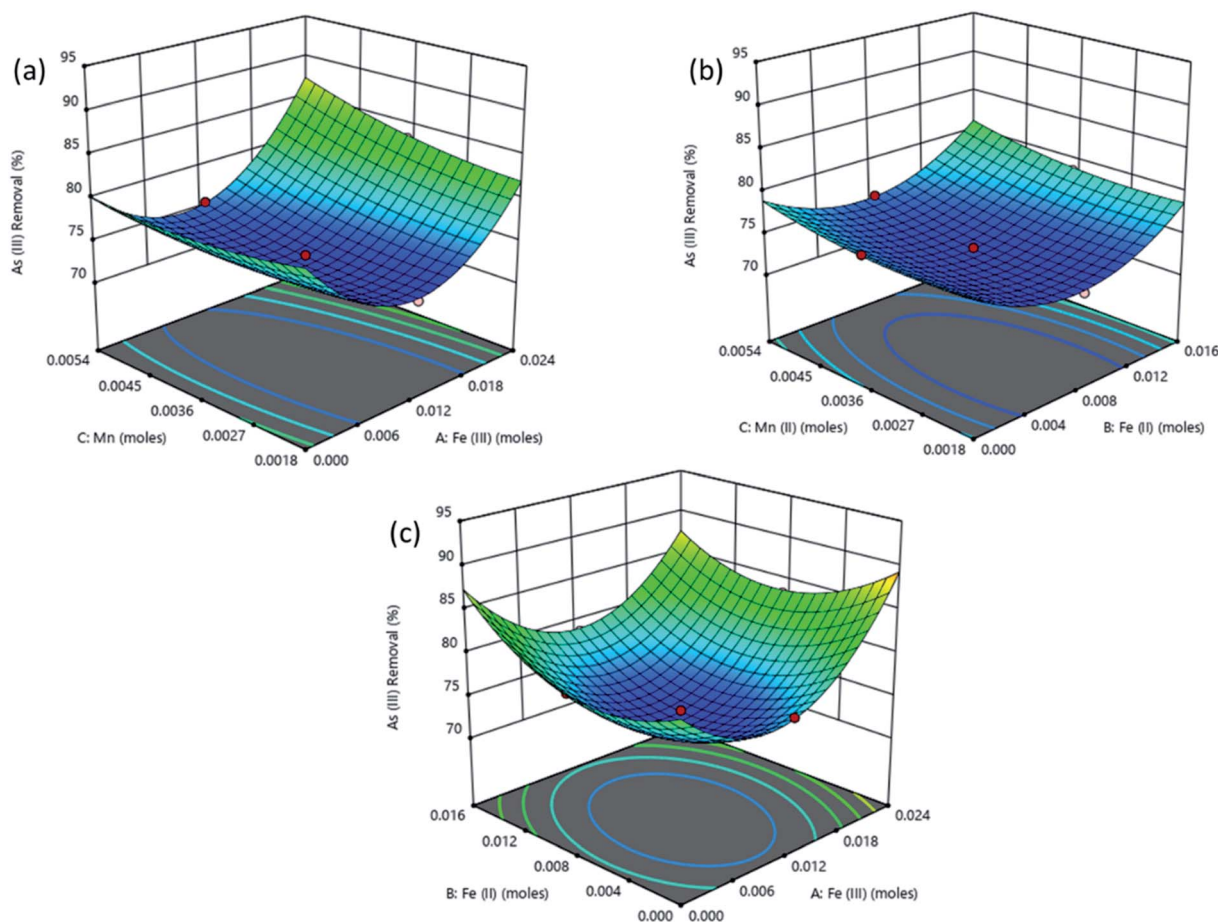


Fig. 6 Surface plots obtained from optimization using RSM for % removal of $\text{As}(\text{III})$ as a function of (a) $\text{Fe}(\text{III})$, (b) $\text{Fe}(\text{II})$ and (c) $\text{Mn}(\text{II})$ amount in the adsorbent.



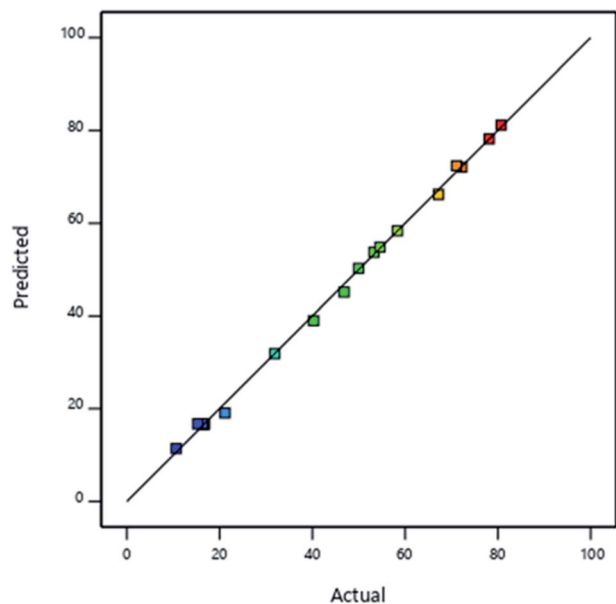


Fig. 7 Plot of model predictions against experimental findings for the concentration of As(v) in solution after adsorption.

Arsenic speciation after adsorption

Determination of the amount of As(v) in the solution after adsorption could give an indication of the potential environmental impact of the treated effluent as determined by the predominant As species. The models for predicting the amount of As(v) in solution is shown in eqn(5).

$$\% \text{As(v)}_{(\text{after adsorption})} = 16.7 + 4.3A + 3.6B + 3.8C - 6.3AB + 4.9AC + 1.2BC + 32.7A^2 + 18.7B^2 \quad (5)$$

Validation of the model is shown in Fig. 7 where R_{Adj}^2 value of 0.9994 implied that the model accurately described the effect of interaction of the iron and manganese species in the speciation of arsenic species after adsorption. The adequate precision was determined to be > 160 indicating that the signal was the adequate and the model can be used to navigate the design space.

Analysis of the As(v) concentration in the solution after adsorption could give an indication of the degree of oxidation of the more toxic As(III) to the less toxic As(v). It is also important to note that the analysis of the residual As(v) in the solution does not take into account the As(v) removed through adsorption.

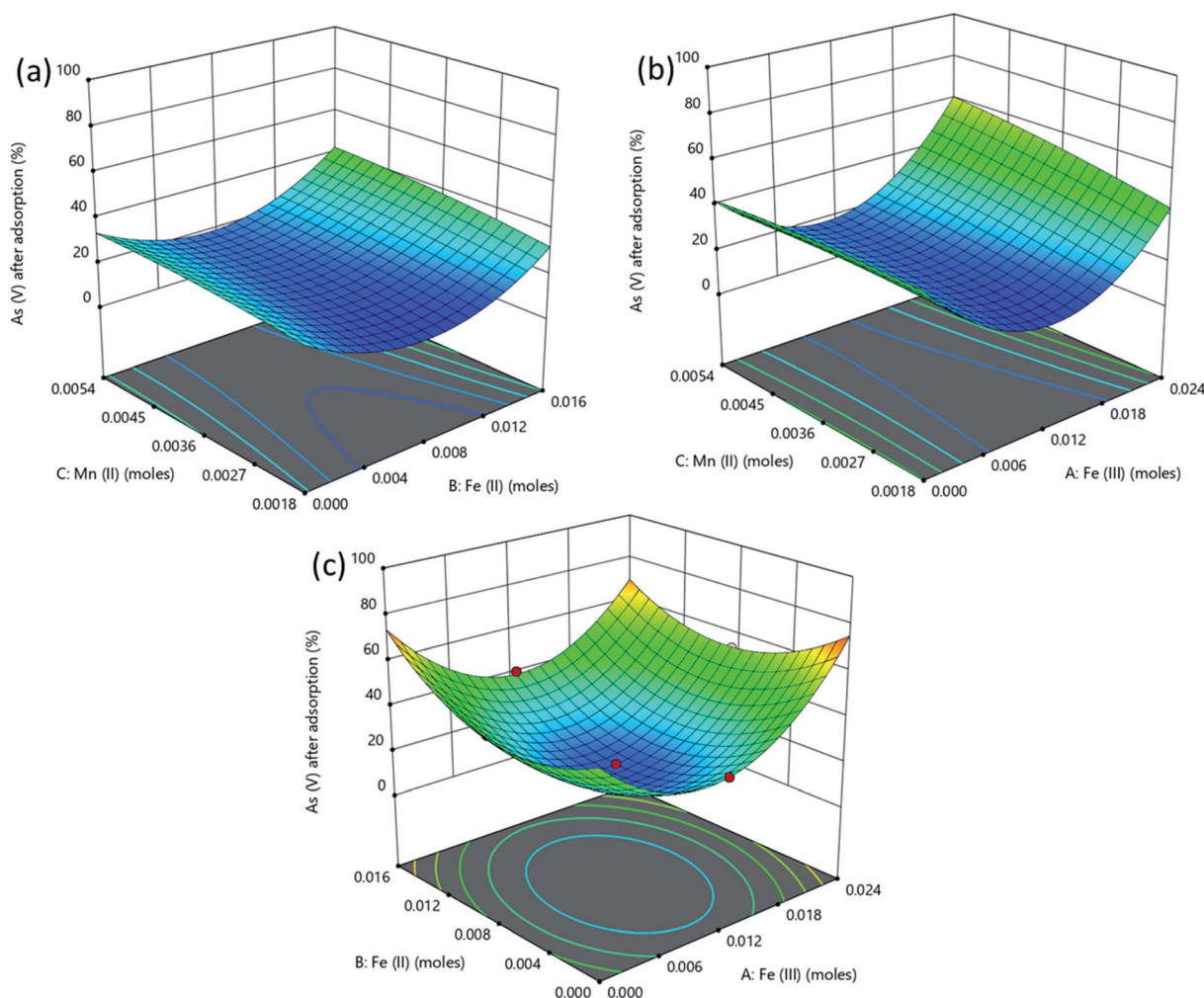


Fig. 8 Response surface plots for the solution concentration of As(III) and As(v) after adsorption.

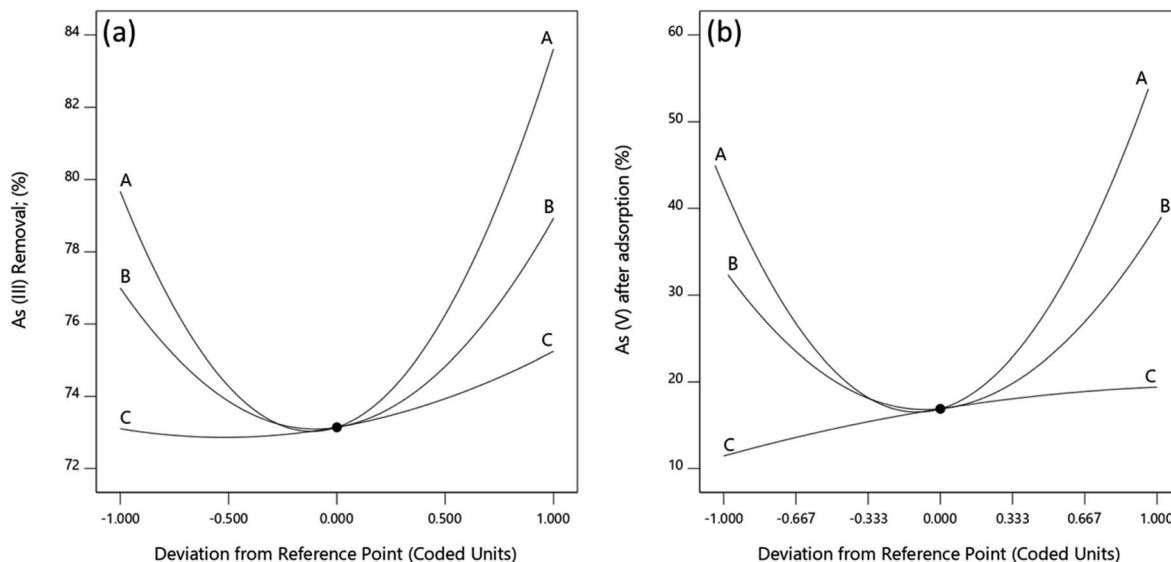


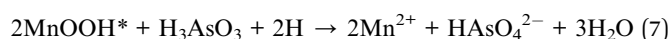
Fig. 9 Perturbation plots of As(III) removal and As(V) concentration in solution.

Fig. 8 shows that there was more As(III) remaining after adsorption in comparison to As(V) when the concentration of both the Fe ions were nearly equal at a constant manganese concentration. Under these conditions the Fe^{3+} concentration is insufficient while Fe^{2+} is in excess therefore formation of magnetite is limited. Subsequently, low amount of manganese doped magnetite is formed. This may cause poor oxidation of As(III) to As(V) and low adsorption due to low precipitation of As(III) by Fe(II) in magnetite.⁴⁸

Perturbation plots

The effects of Fe(III), Fe(II) and Mn concentrations on As(III) and total As removal were evaluated and the effect of each independent variable was determined by perturbation plots. Perturbation plots help to compare the effect of each independent variable at a particular point in the design space. Like in one factor at a time experiments, the response is plotted by varying only one factor over its design range while holding the other factors constant.⁵⁷ A relatively flat line for a given factor shows that a change in the factor has little influence on the performance, while a steep slope or curvature indicates that the response is very sensitive to the factor.

From Fig. 9, As(III) removal and As(V) concentration is least sensitive to change in Mn within the range studied while they are increasingly sensitive to change in Fe(II) and Fe(III) in that order. An increase in Mn concentration leads to a corresponding increase in As(III) removal and As(V) in solution. This is because manganese ions react with As(III) following eqn (6) and (7) to form Mn–O–As complexes releasing As(V) ions into the solution.⁵⁸



This results in a reduction in As(III) concentration and an increase in As(V) formed when Mn concentration is increased. The increase in As(III) removal with an increase in Mn concentration could also result from the formation of new adsorption sites produced during redox reaction of Mn(IV) and As(III) which bind both the formed As(V) and some of the As(III) in solution.⁵⁸

Conclusion

Manganese doped magnetite nanoparticles were synthesized and characterized using several techniques and it was confirmed that manganese atoms displaced iron atoms in the unit cell forming a uniform phase with increased cell volume. Manganese doping resulted in increased surface areas which enhanced pollutant removal due to the presence of manganese oxides on the adsorbent surface providing more binding sites. Statistical analysis revealed that ferric ions had the highest influence on arsenite reduction and arsenate formation while the interactions between ferrous and ferric ions had antagonistic effects on both processes. The interactions between ferric and manganese ions had a triple effect on the formation of arsenate as compared to arsenite reduction. These results confirmed that manganese aided in the oxidation of arsenite to arsenate and in the subsequent adsorption of both species onto the nanoparticle surface.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

The authors acknowledge financial support from the Vaal University of Technology.



References

- 1 Y. C. Sharma, V. Srivastava, V. K. Singh, S. N. Kaul and C. H. Weng, *Environ. Technol.*, 2009, **30**, 583–609.
- 2 Y.-M. Hao, C. Man and Z.-B. Hu, *J. Hazard. Mater.*, 2010, **184**, 392–399.
- 3 Y. Haldorai, A. Rengaraj, T. Ryu, J. Shin, Y. S. Huh and Y.-K. Han, *Mater. Sci. Eng., B*, 2015, **195**, 20–29.
- 4 S. Lunge, S. Singh and A. Sinha, *J. Magn. Magn. Mater.*, 2014, **356**, 21–31.
- 5 A. Mohammadi and M. Barikani, *Mater. Charact.*, 2014, **90**, 88–93.
- 6 C. Su, *J. Hazard. Mater.*, 2017, **322**, 48–84.
- 7 L. Feng, M. Cao, X. Ma, Y. Zhu and C. Hu, *J. Hazard. Mater.*, 2012, **217–218**, 439–446.
- 8 C. Luo, Z. Tian, B. Yang, L. Zhang and S. Yan, *Chem. Eng. J.*, 2013, **234**, 256–265.
- 9 M. Kumari, C. U. Pittman and D. Mohan, *J. Colloid Interface Sci.*, 2015, **442**, 120–132.
- 10 X. Guo, B. Du, Q. Wei, J. Yang, L. Hu, L. Yan and W. Xu, *J. Hazard. Mater.*, 2014, **278**, 211–220.
- 11 S. Rajput, C. U. Pittman and D. Mohan, *J. Colloid Interface Sci.*, 2016, **468**, 334–346.
- 12 I. M. Ahmed, R. Gamal, A. A. Helal, S. A. Abo-El-Enain and A. A. Helal, *Part. Sci. Technol.*, 2016, **6351**, 1–10.
- 13 M. Hua, S. Zhang, B. Pan, W. Zhang, L. Lv and Q. Zhang, *J. Hazard. Mater.*, 2012, **211–212**, 317–331.
- 14 M. Karimi, S. A. Milani and H. Abolghashemi, *J. Nucl. Mater.*, 2016, **479**, 174–183.
- 15 T. M. Thi, N. T. H. Trang and N. T. Van Anh, *Appl. Surf. Sci.*, 2015, **340**, 166–172.
- 16 B. B. Johnson, *Environ. Sci. Technol.*, 1990, **24**, 112–118.
- 17 M. Zhang, Y. Cao, L. Wang, Y. Ma, X. Tu and Z. Zhang, *ACS Appl. Mater. Interfaces*, 2015, **7**, 4650–4658.
- 18 R. R. Pawar, Lalhmunsiam, H. C. Bajaj and S. M. Lee, *J. Ind. Eng. Chem.*, 2016, **34**, 213–223.
- 19 L. Yang, L. Ma, J. Xin, A. Li, C. Sun, R. Wei, B. W. Ren, Z. Chen, H. Lin and J. Gao, *Chem. Mater.*, 2017, **29**, 3038–3047.
- 20 C. L. Warner, W. Chouyyok, K. E. Mackie, D. Neiner, L. V. Saraf, T. C. Droubay, M. G. Warner and R. S. Addleman, *Langmuir*, 2012, **28**, 3931–3937.
- 21 J. Lai, K. V. P. M. Shafi, A. Ulman, K. Loos, N. Yang, M. Cui, T. Vogt, C. Estourne and D. C. Locke, *J. Phys. Chem. B*, 2004, **108**, 14876–14883.
- 22 M. K. Gibbons and G. A. Gagnon, *J. Hazard. Mater.*, 2011, **186**, 1916–1923.
- 23 D. Ociński, I. Jacukowicz-Sobala, P. Mazur, J. Raczky and E. Kociolek-Balawejder, *Chem. Eng. J.*, 2016, **294**, 210–221.
- 24 A. Nazari, M. R. Shishehbor and S. M. Poorhashemi, *J. Text. Inst.*, 2016, **107**, 1617–1631.
- 25 M. A. Bezerra, R. E. Santelli, E. P. Oliveira, L. S. Villar and L. A. Escaleira, *Talanta*, 2008, **76**, 965–977.
- 26 M. Sarkar and P. Majumdar, *Chem. Eng. J.*, 2011, **175**, 376–387.
- 27 Z.-B. Chen, M.-H. Cui, N.-Q. Ren, Z.-Q. Chen, H.-C. Wang and S.-K. Nie, *Bioresour. Technol.*, 2011, **102**, 8839–8847.
- 28 M. S. Podder and C. B. Majumder, *J. Environ. Chem. Eng.*, 2015, **3**, 1986–2001.
- 29 I. L. A. Ouma, E. B. Naidoo and A. E. Ofomaja, *J. Environ. Chem. Eng.*, 2018, **6**, 5409–5419.
- 30 X. Meng and W. Wang, in *Third International Conference on Arsenic Exposure and Health Effects*, Society of Environmental Geochemistry and Health, University of Colorado at Denver, 1998, pp. 1–6.
- 31 S. G. Meng, S. B. Bang and G. P. Korfiatis, *Water Res.*, 2000, **34**, 1255–1261.
- 32 Y. Liu, C. Luo, G. Cui and S. Yan, *RSC Adv.*, 2015, **5**, 54156–54164.
- 33 A. M. G. C. Dias, A. Hussain, A. S. Marcos and A. C. A. Roque, *Biotechnol. Adv.*, 2011, **29**, 142–155.
- 34 A. Z. M. Badruddoza, Z. B. Z. Shawon, W. J. D. Tay, K. Hidajat and M. S. Uddin, *Carbohydr. Polym.*, 2013, **91**, 322–332.
- 35 S. Aredes, B. Klein and M. Pawlik, *J. Cleaner Prod.*, 2012, **29–30**, 208–213.
- 36 S. S. Banerjee and D. H. Chen, *J. Hazard. Mater.*, 2007, **147**, 792–799.
- 37 C. Xiao, X. Liu, S. Mao, L. Zhang and J. Lu, *Appl. Surf. Sci.*, 2017, **394**, 378–385.
- 38 R. M. Cornell and U. Schwertmann, *The Iron Oxides*, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim, FRG, 2nd edn, 2003.
- 39 G. S. Parkinson, *Surf. Sci. Rep.*, 2016, **71**, 272–365.
- 40 T. J. B. Holland and S. A. T. Redfern, *J. Appl. Crystallogr.*, 1997, **30**, 84.
- 41 F. L. Deepak, M. Bañobre-López, E. Carbó-Argibay, M. F. Cerqueira, Y. Piñeiro-Redondo, J. Rivas, C. M. Thompson, S. Kamali, C. Rodríguez-Abreu, K. Kovnir and Y. V. Kolen'ko, *J. Phys. Chem. C*, 2015, **119**, 11947–11957.
- 42 D. Varshney and A. Yogi, *Mater. Chem. Phys.*, 2011, **128**, 489–494.
- 43 X. Liang, Z. He, G. Wei, P. Liu, Y. Zhong, W. Tan, P. Du, J. Zhu, H. He and J. Zhang, *J. Colloid Interface Sci.*, 2014, **426**, 181–189.
- 44 F. Zhang, X. Wang, J. Xionghui and L. Ma, *Environ. Pollut.*, 2016, **216**, 575–583.
- 45 K. Gupta, A. Maity and U. C. Ghosh, *J. Hazard. Mater.*, 2010, **184**, 832–842.
- 46 S. Ouvrard, P. de Donato, M. O. Simonnot, S. Begin, J. Ghanbaja, M. Alnot, Y. B. Duval, F. Lhote, O. Barres and M. Sardin, *Geochim. Cosmochim. Acta*, 2005, **69**, 2715–2724.
- 47 H. Y. Niu, J. M. Wang, Y. L. Shi, Y. Q. Cai and F. S. Wei, *Microporous Mesoporous Mater.*, 2009, **122**, 28–35.
- 48 Z. Tian, Y. Feng, Y. Guan, B. Shao, Y. Zhang and D. Wu, *Sci. Rep.*, 2017, **7**, 1–11.
- 49 P. Roy, U. Dey, S. Chatteraj and D. Mukhopadhyay, *Appl. Water Sci.*, 2017, **7**, 1307–1321.
- 50 M. Fang, V. Ström, R. T. Olsson, L. Belova and K. V. Rao, *Nanotechnology*, 2012, **23**, 145601.
- 51 D.-D. Herea, H. Chiriac and N. Lupu, *J. Nanopart. Res.*, 2011, **13**, 4357–4369.



- 52 G. A. Salazar-Alvarez, *Synthesis, Characterisation and Applications of Iron Oxide Nanoparticles*, K. T. H. Materialvetenskap, Stockholm, 2004.
- 53 H.-C. Roth, S. P. Schwaminger, M. Schindler, F. E. Wagner and S. Berensmeier, *J. Magn. Magn. Mater.*, 2015, **377**, 81–89.
- 54 S. J. Oh, D. C. Cook and H. E. Townsend, *Hyperfine Interact.*, 1998, **112**, 59–65.
- 55 H. W. P. Carvalho, P. Hammer, S. H. Pulcinelli, C. V. Santilli and E. F. Molina, *Mater. Sci. Eng., B*, 2014, **181**, 64–69.
- 56 A. S. Teja and P.-Y. Koh, *Prog. Cryst. Growth Charact. Mater.*, 2009, **55**, 22–45.
- 57 P. Roy, N. K. Mondal and K. Das, *J. Environ. Chem. Eng.*, 2014, **2**, 585–597.
- 58 Y. Yin, T. Zhou, H. Luo, J. Geng, W. Yu and Z. Jiang, *Colloids Surf., A*, 2019, **575**, 318–328.

