


 Cite this: *RSC Adv.*, 2020, 10, 11435

Purification of metallurgical-grade silicon combining Sn–Si solvent refining with gas pressure filtration

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In this study, the purification of metallurgical-grade silicon using a combination of solvent refining and gas pressure filtration was investigated in the Sn–Si alloy. After the solvent refining process, the silicon was separated from the solvent by gas pressure filtration. The effects of pressure differentials (p), separation temperatures (T), and silicon contents in the alloy ($\omega_{(a)}$) on the separation efficiency were evaluated. The filtration result was improved with a higher pressure differential. The separation temperature had little effect on the separation efficiency, whereas a higher silicon content in the alloy led to a decrease of the separation efficiency. The final purification result after separation was examined, and a better separation contributed to the removal of impurity. The optimal result for separation was obtained at $p = 0.30$ MPa, $T = 250$ °C, and $\omega_{(a)} = 20$ wt%, and 93.6% of tin was separated into the filtrate, while almost all the silicon was recovered and formed the separated silicon with a silicon content of 80.0 wt%. At the same time, most impurities were eliminated and 94.9% of B was removed after refining the sample twice.

Received 3rd November 2019

Accepted 11th March 2020

DOI: 10.1039/c9ra09077k

rsc.li/rsc-advances

1. Introduction

The abundant and low-cost solar-grade silicon (SoG-Si) feedstock is essential for the widespread use of solar cells. The traditional method for SoG-Si production (the Siemens process) has less potential in reducing the cost. The high consumption of chemical energy is inevitable in the Siemens process, because the conversion of metallurgical-grade silicon (MG-Si) to gaseous compounds is followed by distillation, reduction, and deposition into high-purity silicon.¹ Therefore, metallurgical refining, a promising process for producing SoG-Si, has gradually become a topic of interest. It is based on a low-cost metallurgical purification technique, using MG-Si as the starting material and preventing silicon from becoming involved in a chemical reaction. The purification of MG-Si by metallurgical methods, such as solvent refining,^{2,3} slag refining,⁴ directional solidification,⁵ and so on,^{6,7} has been investigated.

MG-Si (purity, 90%~99.9%) contains more impurities than SoG-Si (purity, 99.999%~99.9999%), which are mainly metallic impurities, boron, and phosphorus. The removal of boron and phosphorus from MG-Si by metallurgical methods is relatively difficult compared with that of metallic impurities.^{2,5} Solvent refining is efficient for the removal of impurities from MG-Si, including boron and phosphorus. It relies on the impurities preferentially segregating into the liquid phase during the solidification process. Many metals, such as tin,^{8,9} aluminum,¹⁰ copper,¹¹ and iron,¹² have been proved to be suitable solvents

that could decrease the segregation coefficient of impurity contributing to its removal from the silicon. Among them, the Sn–Si system is highly efficient for the removal of boron. The segregation coefficient of boron in the Sn–Si melt can decrease to 0.038, far less than 0.8, which is the coefficient of boron in pure silicon.⁹ In addition, neither eutectic phase nor intermetallic compound is generated in the Sn–Si alloy, which means no loss of silicon during the refining process. Therefore, solvent refining in the Sn–Si system is promising for the purification of MG-Si.

Solvent refining has been proved as a potential method,^{9,13} but the separation of refined silicon from the solvent needs to be improved. Direct acid leaching is clearly undesirable due to the massive loss of tin and generation of waste acid.¹⁴ The electromagnetic force can agglomerate the silicon under an alternating magnetic field during the solidification process,^{15,16} but much tin still remains in the enriched silicon. Separation under centrifugal force is effective, but difficult to widely apply.¹⁷ A novel method of gas pressure filtration was investigated for the separation of silicon in the Al–Si solvent refining process,¹⁸ and was demonstrated to be a promising method with a high efficiency.

In this work, a method combining gas pressure filtration and Sn–Si solvent refining was studied for the purification of MG-Si and the separation of purified silicon from solvent. The influences of pressure differentials (p), separation temperatures (T), and silicon contents in the alloy ($\omega_{(a)}$) on the separation efficiency were investigated, and final purification results after separation process were examined.

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2. Experimental

2.1 Apparatus

Gas pressure filtration, a method of strengthening mass transfer under the driving force of a pressure differential, can separate a liquid phase from solid particles. During the filtration process, the liquid melt is forced through the filtration media by the applied pressure while the solid particles are intercepted. The filtration apparatus is shown in Fig. 1, and it was produced by Hefei Kejing Material Technology Co., Ltd in China. It mainly consisted of an induction heating furnace and a pressurization system. The temperature was controlled by a program controller with a type B thermocouple. A graphite crucible was used for placing the sample and enclosing spaces. A carbon fiber felt (CFF) with a thickness of 3 mm was supplied by Jing Long Te Tan Ltd in China and was used as the filtration media in this work. It has a real density of 1.82 g cm^{-3} , bulk density of 0.171 g cm^{-3} , and average fiber diameter of $17 \mu\text{m}$. The graphite flake had 20 uniformly distributed holes (2 mm in diameter), which were used to support the CFF and had little effect on the filtration process. The chamber in the apparatus was divided into two parts when the sample melted, and the upper part could be pressurized with a gasholder. The molten tin passed through the CFF by applying the pressure differential until the two parts of the chamber were connected.

2.2 Methods

The materials used in this work were MG-Si (99.6 wt% purity) and tin (99.99 wt% purity), and the impurity content in the raw materials was presented in a later section (Table 1). First, MG-Si was blended with the tin to form a mixture with varying weight percentages of silicon. Approximately 60 g of this mixture was placed in the graphite crucible of the apparatus. Then, the temperature of the apparatus was raised to $1500 \text{ }^\circ\text{C}$ and held for

1 h under argon atmosphere to melt the mixture. Third, the temperature was cooled to the separation temperature ($250\text{--}850 \text{ }^\circ\text{C}$) at $3 \text{ }^\circ\text{C min}^{-1}$ and held for 30 min contributing to the crystallization of silicon and segregation of impurity.^{9,10} Above mentioned was the solvent refining process. Next the separation process following, by applying a certain pressure of argon gas (0.05–0.3 MPa), the liquid tin was forced to flow through CFF and formed the filtrate, while the unfiltered part was intercepted by CFF and thought of as the separated silicon. Finally, the separated silicon was cleaned with hydrochloric acid ($70 \text{ }^\circ\text{C}$ for 2 h) after crushing ($100 \mu\text{m}$) for the removal of residual solvent, which obtained the refined silicon. The refined silicon was the product of this combined method for the MG-Si purification.

To evaluate the separation efficiency, two parameters were investigated, including the silicon content in the separated silicon ($\omega_{(s)}$) and the residual ratio of tin (η). The silicon content in the separated silicon ($\omega_{(s)}$) was measured, which obtained the weight of silicon in the tested specimen by dissolving tin with hydrochloric acid. The residual ratio of tin (η) was defined using eqn (1):

$$\eta = \frac{m_s \times (1 - \omega_{(s)})}{m_a \times (1 - \omega_{(a)})} \times 100\% \quad (1)$$

where m_s and m_a are the mass of separated silicon and Sn-Si alloy, respectively; $\omega_{(a)}$ denotes the mass fraction of silicon in the Sn-Si alloy.

As the assessment of impurity removal from silicon, the removal fraction of impurity element (R_i) was calculated by eqn (2):

$$R_i = \frac{C_{i,0} - C_{i,r}}{C_{i,0}} \times 100\% \quad (2)$$

where $C_{i,0}$ is the impurity content (element i) in MG-Si; $C_{i,r}$ denotes the impurity content (element i) in the refined silicon after refining process.

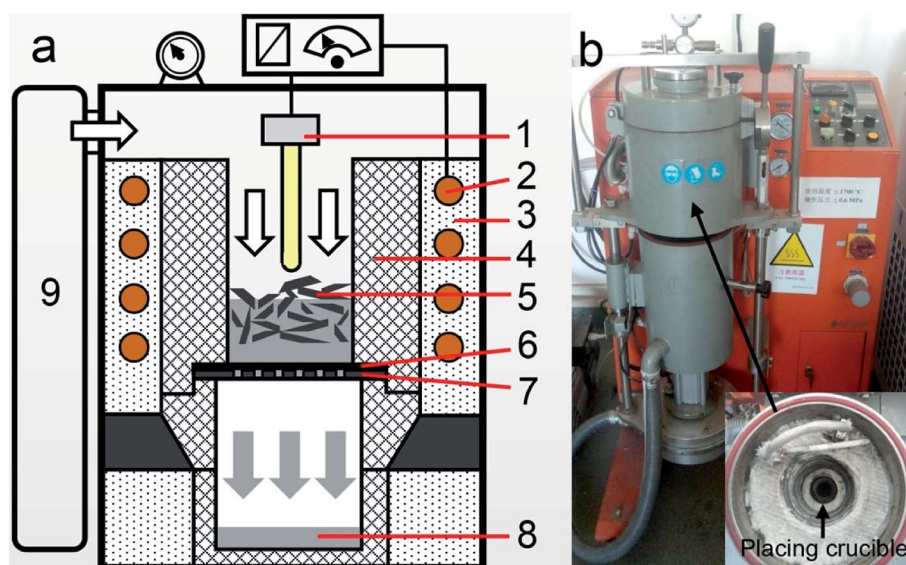


Fig. 1 (a) Schematic and (b) photograph of the pressurized filtering apparatus. ((1) Thermocouple, (2) induction coil, (3) thermal insulation materials, (4) graphite crucible, (5) separated silicon, (6) carbon fiber felt, (7) graphite flake, (8) filtrate, (9) gasholder).



Table 1 The impurity removal in the twice-refining process at $p = 0.30$ MPa, $T = 250$ °C, and $\omega_{(a)} = 20$ wt%

Element (i)	B	P	Al	Fe	Ti	V
MG-Si (C_{i_0})	31.5	54.3	521.4	2388.4	303.2	373.6
Sn	4.3	1.4	3.2	9.8	5.8	4.7
Refined silicon (C_{i_r})	6.2	17.8	76.2	1048.5	81.5	87.4
Removal fraction in the first refining (R_i)	80.3%	67.1%	85.4%	56.1%	73.1%	76.6%
Secondary refined silicon	1.6	12.9	18.3	198.2	9.8	9.1
Removal fraction in the second refining	94.9%	76.3%	96.5%	91.7%	96.8%	97.6%

2.3 Characterization

The separated silicon and filtrate were sectioned into two parts along the longitudinal center axes and then polished to a 1 μm finish for X-ray diffraction (XRD, SmartLab, Rigaku, Japan), optical microscopy (DM4M, Leica, Germany), scanning electron microscopy and energy dispersive X-ray spectroscopy (SEM-EDS, Phenom Prox, Phenom, Netherlands) analyses. The impurity content in the refined silicon was at a ppm level,^{10,17} while the silicon content in the filtrate was in the range of 0.05 wt% to 0.3 wt%. Therefore, the concentrations of impurity in refined silicon were analyzed by Inductively Coupled Plasma Mass Spectrometry (ICP-MS, ICP RQ, Thermo Fisher Scientific, USA), and the silicon content in the filtrate was analyzed by Inductively Coupled Plasma Optical Emission Spectrometry (ICP-OES, Optima 7000DV, PerkinElmer, USA). The two measurement methods were similar, except for the fact that different instruments were used. The sample was firstly dissolved into the solution by hydrofluoric acid and nitric acid, and the content of target element in the solution was then analyzed by ICP-MS/OES to obtain the impurity content in refined silicon or the silicon content in filtrate.

3. Results and discussion

3.1 Solvent refining process

In this work, the solvent refining was first completed. In the cooling process after the melting of tin and silicon, the silicon was firstly crystallized, while the impurities was not precipitated with silicon and preferentially segregated into the solvent.^{2,8} Therefore, the high-purity silicon was obtained, and the impurities came into the molten tin. The impurities in MG-Si and Sn-Si alloy after refining process are present in Fig. 2.

There were many impurities in the MG-Si, which were the bright white particles in Fig. 2a. They were mainly iron matrix impurities detected by energy-dispersive X-ray spectroscopy (EDS) point analysis (point#1 in Fig. 2a). According to the Sn-Si binary phase diagram,¹⁹ the vast majority of silicon precipitated into solid phase above 1200 °C, and the formed silicon grains exhibited a large size (black parts in Fig. 2b and c). These silicon grains contained almost no impurities, while the impurity elements segregated into the solvent, as shown in Fig. 2b and c. Most iron impurity existed in the binary silicide phase of FeSi_2 , detected by EDS point analysis (point#2 in Fig. 2b), and its particle size was obviously larger than that of other impurity. The Ti, V, and some Fe formed multicomponent phases

(point#3 in Fig. 2c) and surrounded by the tin. The mentioned above are basically consistent with other works.^{9,17}

The high-purity silicon grains were obtained after solvent refining, but they were still mixed with the solvent (tin). At the same time, the impurities distributed in the solvent also need to be removed. Since they are the main focus in this study, the separation of silicon from solvent by gas pressure filtration and final purification result in this combined method were investigated.

3.2 Separation efficiency of silicon by gas pressure filtration

3.2.1 Separation of silicon from solvent. The results of filtering separation at $p = 0.15$ MPa (pressure differential of 0.15 MPa), $T = 250$ °C (separation temperature of 250 °C), and $\omega_{(a)} = 20$ wt% (Sn-20% Si alloy) are shown in Fig. 3. The electromagnetic force agglomerates the silicon during the solidification process, and the silicon can be enriched in different positions under varied electromagnetic fields.^{20,21} In this work, the silicon grains were enriched at the bottom of the alloy (Fig. 3a). The morphology of separated silicon intercepted by CFF was distinctly different from that of the filtrate (Fig. 3b and c). The separated silicon was mainly composed of silicon grains (Fig. 3d), while tin was the main component of the filtrate (Fig. 3e). The silicon remained in the separated silicon, while most tin came into the filtrate, and refined silicon contained almost no tin after being cleaned with hydrochloric acid, as can be inferred from Fig. 3f.

The micro-morphologies of the separated silicon and filtrate observed by optical microscopy are presented in Fig. 4, referred to as areas 'A' and 'B' in Fig. 3, respectively. In the separated silicon, the silicon grains exhibited a size of a millimeter magnitude and a few residual tin remained in the voids (Fig. 4a). There were no large sized grains in the filtrate, and the black particle with a tiny size in the filtrate was the silicon (Fig. 4b). The silicon content in the separated silicon ($\omega_{(s)}$) and the residual ratio of tin (η) were 74.2 wt% and 9.1%, respectively, while the silicon content in the filtrate was only 0.08 wt%. It means that almost all the silicon was recovered and over 90% tin was separated, consistent with Fig. 3f. So it could be concluded that the separation of silicon from the solvent was achieved by the gas pressure filtration.

3.2.2 Effect of different conditions on separation. In this work, the silicon was separated under different conditions (*i.e.*, pressure differentials, separation temperatures, and alloy



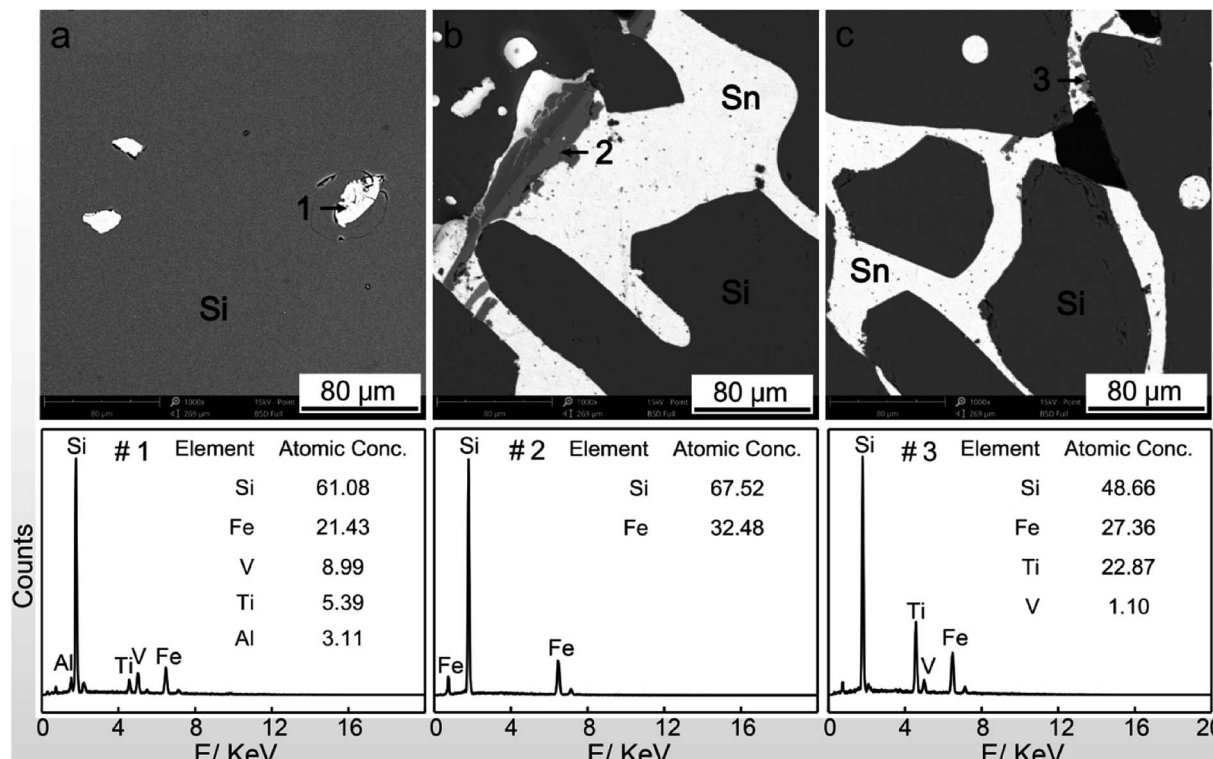


Fig. 2 Morphologies of impurities in (a) MG-Si and (b and c) Sn-Si alloy after refining process.

compositions), and the separation efficiencies are presented in Fig. 5.

The pressure differential is the key factor in this method, and it can make answers to the questions that whether efficient separation is achieved and how to meet its demand. The effect of pressure differentials on separation was evaluated at $T = 250\text{ }^{\circ}\text{C}$ and $\omega_{(a)} = 20\text{ wt}\%$, as shown in Fig. 5a. The value of $250\text{ }^{\circ}\text{C}$ was chosen as the separation temperature to ensure

sufficient superheat, while the melting point of tin is $232\text{ }^{\circ}\text{C}$. The separation results were better with higher pressure differentials. At $p = 0.05\text{ MPa}$, the result was undesirable, with $\omega_{(s)}$ of $51.4\text{ wt}\%$ and η of 22.3% . Additionally, the results were efficient when $p \geq 0.15\text{ MPa}$, with $\omega_{(s)}$ over $74.2\text{ wt}\%$ and η below 9.1% . At $p = 0.30\text{ MPa}$, the $\omega_{(s)}$ was $80.0\text{ wt}\%$ and η was 6.4% . The curve plateau of the separation results was reached at $p = 0.20\text{ MPa}$, and the improvement of separation efficiency was not

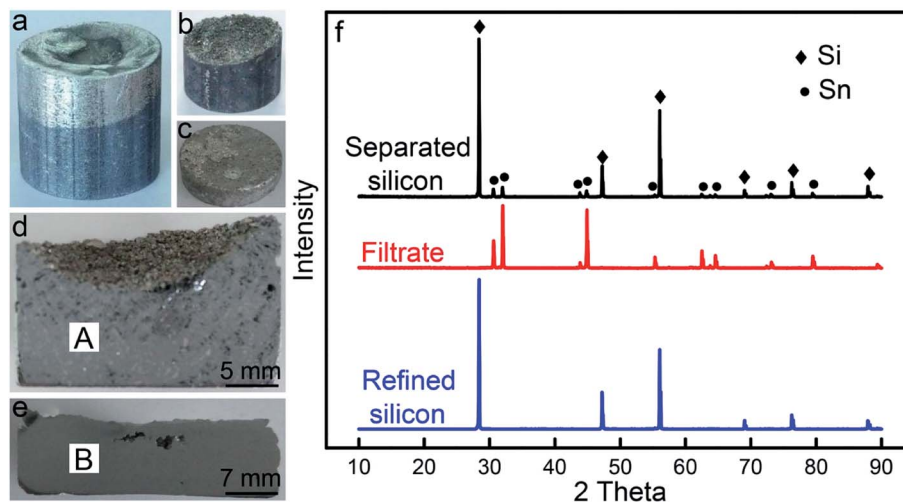


Fig. 3 Results of filtering separation at $p = 0.15\text{ MPa}$, $T = 250\text{ }^{\circ}\text{C}$, and $\omega_{(a)} = 20\text{ wt}\%$. Pictures of (a) Sn-Si alloy before filtration; (b) separated silicon and (c) filtrate; sections of (d) separated silicon and (e) filtrate; and (f) X-ray diffraction (XRD) patterns of separated silicon, filtrate, and refined silicon.



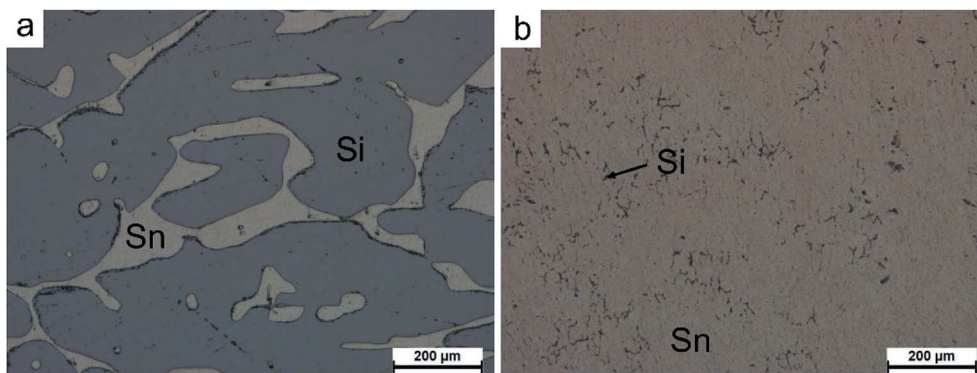


Fig. 4 Optical images of (a) separated silicon and (b) filtrate (areas 'A' and 'B' in Fig. 3, respectively).

obvious at $p \geq 0.20$. The efficient filtration results could be obtained without a high pressure differential, which is the advantage of this combined method.

Since optimal separation was achieved at $p = 0.30$ MPa, the effect of the separation temperature on filtration was investigated at $p = 0.30$ MPa and $\omega_{(a)} = 20$ wt%, as shown in Fig. 5b. The $\omega_{(s)}$ fluctuated in the range of 78.3 wt% to 80.9 wt% at $T = 250$ – 850 °C,

and η was around 6.5%. Varied separation temperatures had little effect on the separation efficiency. According to the phase diagram of the Sn–Si alloy,¹⁹ Sn–Si alloy melts mainly consist of molten tins and silicon grains from 232 °C to 1200 °C, and their amounts varied less with the temperature. Therefore, similar results were obtained, and 250 °C is a better choice because of the low superheat.

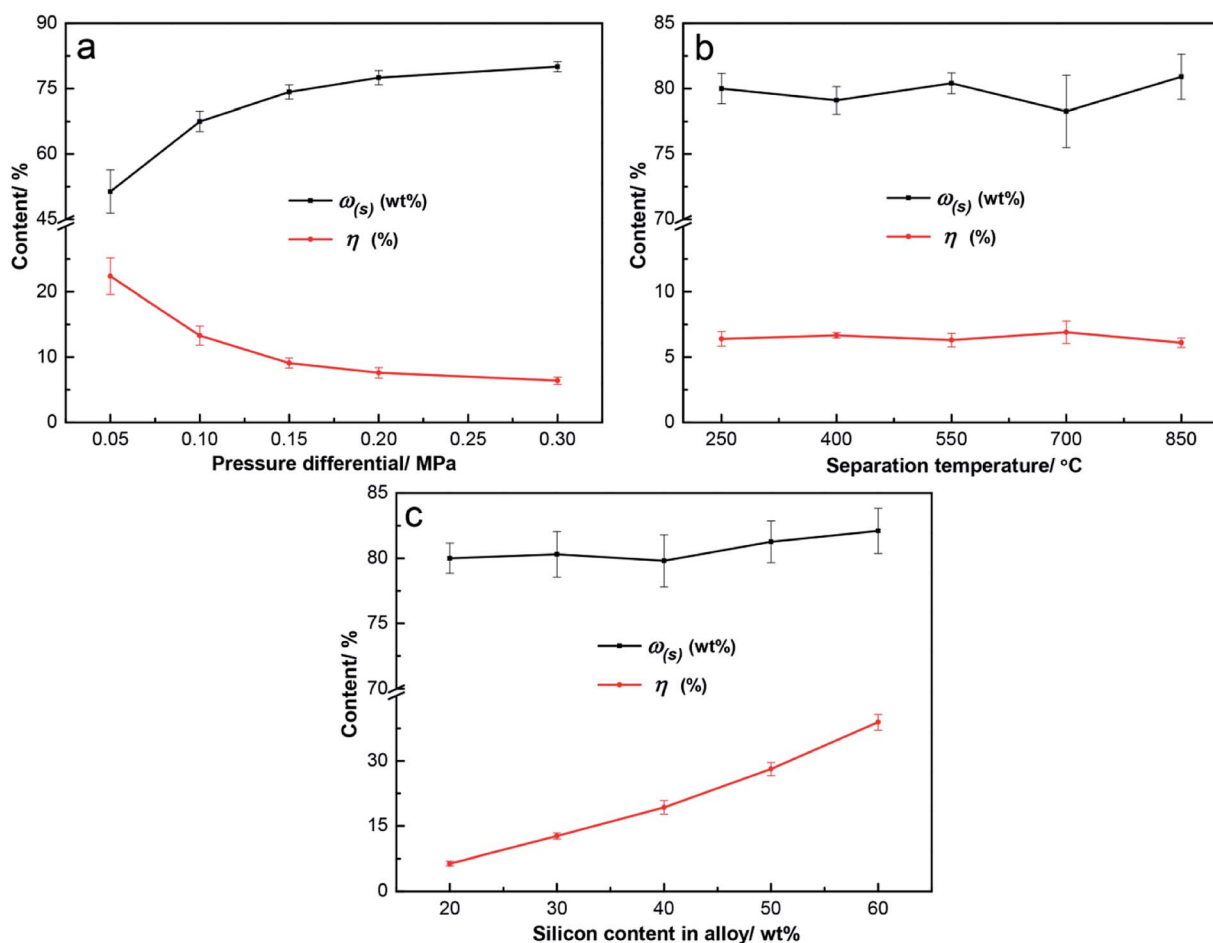


Fig. 5 Separation efficiencies with (a) pressure differential at $T = 250$ °C and $\omega_{(a)} = 20$ wt%, (b) separation temperature at $p = 0.30$ MPa and $\omega_{(a)} = 20$ wt%, and (c) alloy composition at $p = 0.30$ MPa and $T = 250$ °C.



Based on the above results, the effect of the alloy composition on separation was examined at $p = 0.30$ MPa and $T = 250$ °C, as shown in Fig. 5c. From $\omega_{(a)} = 20$ wt% to $\omega_{(a)} = 60$ wt%, the silicon content in separated silicon gradually increased from 80.0 wt% to 82.1 wt%, but η sharply increased from 6.4% to 38.9%. The separation efficiency obviously reduced with the higher silicon in the alloy, and the optimal result with $\omega_{(s)}$ of 80.0 wt% and η of 6.4% was obtained at $\omega_{(a)} = 20$ wt%. A higher amount of silicon in the alloy means more solid phases during the filtration process, which contributed to the decrease of the separation efficiency.

Optimal separation was obtained at $p = 0.30$ MPa, $T = 250$ °C, and $\omega_{(a)} = 20$ wt%; 93.6% of tin was separated into the filtrate; and the separated silicon was obtained with a silicon content of 80.0 wt%. In all of the above experiments, the silicon content in the filtrate was less than 0.3 wt%, as shown by ICP-OES. This means that the filtrate was basically the tin and almost all the silicon in the Sn-Si alloy remained in the separated silicon. This method is proved to be efficient for the separation of silicon.

3.2.3 Requirement of the filtration process. In this work, the separation of silicon from solvent was achieved by forcing the molten tin to pass through the CFF and intercepting the silicon grains upon the CFF. While the size of silicon grains had a millimeter magnitude and much larger than the average fiber diameter of CFF, the filtering process can be simplified as the molten tin flowing through the CFF forced by the pressure differential and the minimum pressure differential for filtration can be calculated. To filter the molten tin, the pressure differential must be larger than the filtration resistance (P_f) offered by the filtration media. On the basis of capillary law, an equation for calculation of the P_f of the infiltration of liquids into fibrous preforms was introduced:²²

$$P_f = -4\sigma \cos \theta \frac{V_f}{(1 - V_f)D} \quad (3)$$

where σ is the surface tension of the melt; θ denotes the wetting angle; V_f is the fiber volume fraction; and D denotes the average fiber diameter. At the separating temperature of 250 °C, substituting $\sigma = 0.529$ N m⁻¹,²³ $\theta = 150$ degree (tin wetting angle on graphite, approximately),²⁴ $V_f = 0.094$ (ratio of the bulk density of CFF to the real density of CFF), and $D = 1.7 \times 10^{-5}$ m, the P_f equals 0.011 MPa. Therefore, the molten tin could flow through the CFF forced by the pressure differential larger than 0.011 MPa, which was consistent with the experimental result. The initial filtering condition can be easily satisfied, and the efficient separation can be obtained at a low pressure differential (Fig. 5a). It could be concluded that the separation method has a great potential for production applications.

3.3 Purification efficiency in the combined method

3.3.1 Final purification result after separation. The purification process in this combined method was actually divided into two parts: one was the process of the impurity preferentially segregating into the solvent from silicon, and it was in the solvent refining process; the other was the process of the impurity distributed in solvent removing from the mixed melt

into the filtrate, and it was along with the separation process. The final purification result was examined after the separation under different conditions (Section 3.2.2). In previous works, the refined silicon was generally leached by aqua regia and hydrofluoric acid successively.^{2,17} This pickling method not only removed the residual solvent, but also dissolved the impurities segregated on the grain boundaries in silicon or attached to the surface of silicon during the refining process. In order to accurately evaluate the purification of the method combining solvent refining and gas pressure filtration, the refined silicon in this study was obtained by acid leaching with hydrochloric acid. After pickling, the refined silicon was basically free of the solvent (Fig. 3f).

The purification efficiency was studied after the separation results had been investigated. The concentrations of impurity in the refined silicon were analyzed by ICP-MS, and the removal fraction of impurity element (R_i) was calculated *via* eqn (2). Impurity removals in the combined method are presented in Fig. 6, while the impurity contents in MG-Si and tin are shown in Table 1.

The removal of Fe, Ti, and V enhanced with the increasing pressure differential, while the removal of B, P, and Al exhibited little change, as shown in Fig. 6a. Upon increasing the pressure differential from 0.05 to 0.30 MPa, the removal fraction of B and P fluctuated in the range of 74.5% to 81.2% and 64.3% to 68.2%, respectively, and R_{Al} fluctuated in the range of 81.3% to 85.4%. The removal of Fe, Ti, and V was inefficient at $p = 0.05$ MPa, and only 30.8% of Fe, 44.3% of Ti, and 44.9% of V were removed during the refining process. Furthermore, their removal was improved at $p = 0.30$ MPa, at which point 56.1% of Fe, 73.1% of Ti, and 76.6% of V were eliminated in the refined silicon. It can be found that the plateaus in the curves of R_{Fe} , R_{Ti} and R_V were reached at 0.15 MPa or 0.2 MPa, and similar to that in the curve of separation result (Fig. 5a).

The purification efficiency with separation temperature fluctuated to a certain extent, as shown in Fig. 6b. The R_B and R_P fluctuated in the range of 78.2% to 82.3% and 65.4% to 70.1%, respectively, at $T = 250$ –850 °C. The removal of metallic impurity was also not changed notably with the increase of temperature, including all of the Al, Fe, Ti, and V. The R_{Al} and R_{Fe} fluctuated in the range of 81.5% to 86.5% and 49.4% to 58.2%, while R_{Ti} and R_V fluctuated in the range of 69.3% to 75.3% and 73.2% to 78.2%, respectively.

The removal fraction of Fe, Ti, and V significantly reduced with a higher silicon content in the alloy, while the removal of B, P, and Al fluctuated to a certain extent, as shown in Fig. 6c. From $\omega_{(a)} = 20$ wt% to $\omega_{(a)} = 60$ wt%, the R_B and R_P fluctuated in the range of 70.3% to 80.3% and 61.3% to 71.5%, respectively, and R_{Al} fluctuated in the range of 82.9% to 85.4%. More Fe, Ti, and V remained in the refined silicon when increasing the silicon content in the alloy: R_{Fe} reduced from 56.1% to 21.9%, while R_{Ti} and R_V decreased from 73.1% to 38.6% and 76.6% to 37.8%, respectively.

Compared with Fig. 5, the removal of Fe, Ti, and V is consistent with the separation efficiency of silicon under different conditions, while the removal of B, P, and Al was almost unchanged. The purification result obtained after the



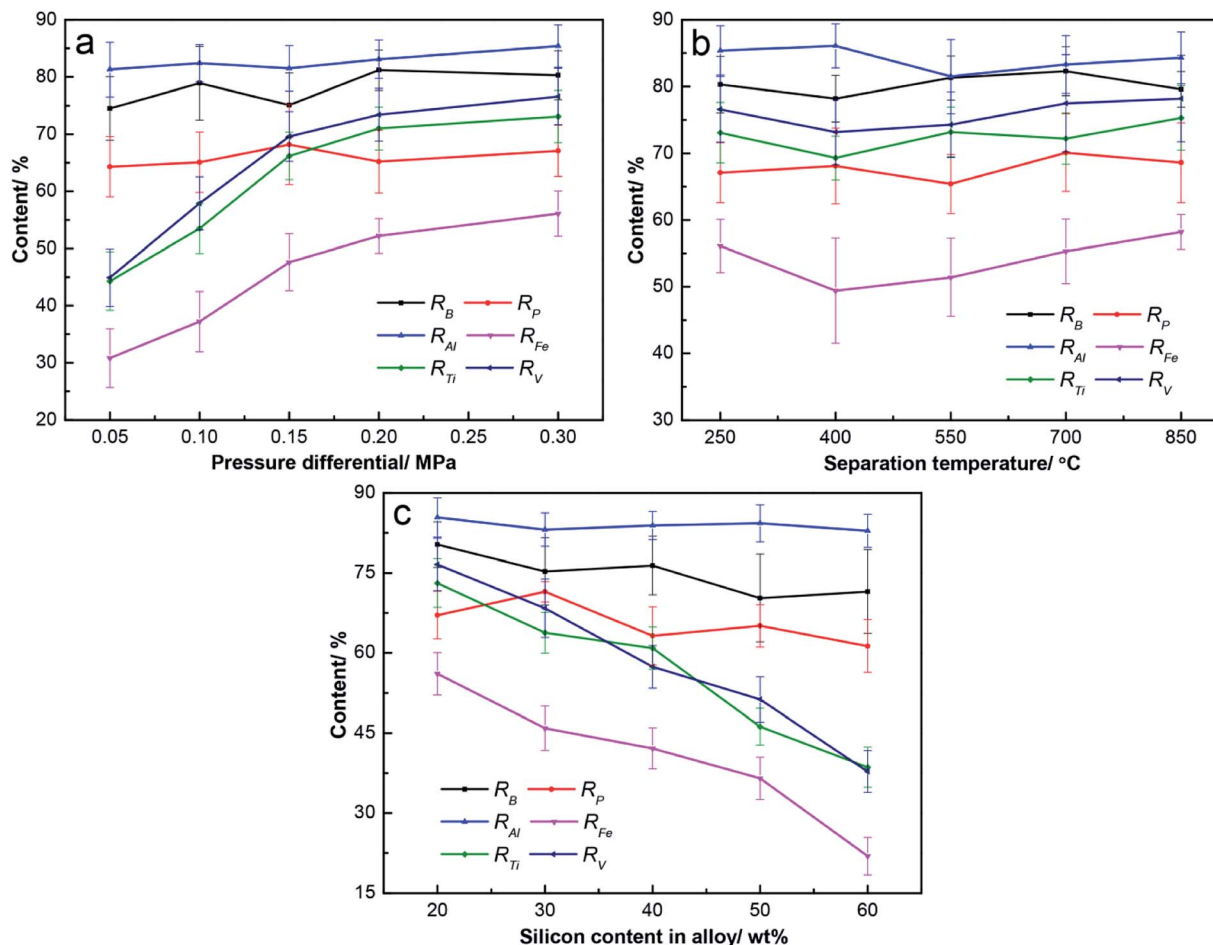


Fig. 6 Impurity removals with (a) pressure differential at $T = 250$ °C and $\omega_{(a)} = 20$ wt%, (b) separation temperature at $p = 0.30$ MPa and $\omega_{(a)} = 20$ wt%, and (c) alloy composition at $p = 0.30$ MPa and $T = 250$ °C.

optimal separation (at $p = 0.30$ MPa, $T = 250$ °C, and $\omega_{(a)} = 20$ wt%) was efficient, and the removal of B (80.3%) and P (67.1%) was comparable to the results presented in other literature, but the removal of metallic impurity (56.1% of Fe, 73.1% of Ti, and 76.6% of V) was less efficient, especially the iron removal.^{8,9,17}

In the combined method of this study, most impurities were in the company of molten tin and passed into the filtrate due to their tiny particle size. The poor separation efficiency led to more residues of molten tin and impurities in the separated silicon. Most notably, the particle size of $FeSi_2$ phase was obviously larger than that of other impurities (Fig. 2b), which made it difficult to separate and kept it in the separated silicon. All of the Ti, V, and $FeSi_2$ were difficult to dissolve using hydrochloric acid, and they remained in the refined silicon. Therefore, a better separation contributed to impurity removal.

3.3.2 Removal of impurity after refining sample twice. In order to obtain a better removal efficiency, a twice-refining process with the combined method was accomplished at $p = 0.30$ MPa, $T = 250$ °C, and $\omega_{(a)} = 20$ wt%, and the purification results are listed in Table 1. The second refining process was completed for the improvement of impurity removal, while the

refined silicon was used as the initial material instead of MG-Si and the same steps were repeated to obtain the secondary refined silicon. The removal fraction of impurity in the second refining was also calculated by eqn (2), using the impurity content in the secondary refined silicon instead of the impurity content in the refined silicon ($C_{i,r}$). It is emphasized that B was removed from 31.5 ppm to 1.6 ppm, with the removal fraction of 94.9%, and this is close to the requirement of SoG-Si. The removal fraction of P was 76.3%, and was limited by the phosphorus removal capacity of the Sn-Si system.⁹ After the twice-refining process, most impurities were eliminated, and more than 95% of Al, Ti, and V was removed. Although the removal fraction of iron was just 91.7% in this method, the iron impurity in the silicon can be easily removed by other methods.^{2,5} The solvent refining has been proven to be efficient, and it has potential for the application of MG-Si purification by combining gas pressure filtration.

3.4 Comparison with other methods

As a physical method, the method combining solvent refining and gas pressure filtration is believed to be superior to traditional method (the Siemens process) because of the simple



process and its low energy consumption.¹ In the case of similar effect for the impurity removal in solvent refining process, the subsequent separation process may become another key to its application. Compared with the direct acid leaching after refining, there was over 90% reduction in the acid consumption using gas pressure filtration, and most solvent can be recycled. As for the method of electromagnetic enrichment, its separation efficiency was investigated in the Al–Si system,²⁵ similar to the Sn–Si alloy. In the silicon enrichment part of Al–45% Si alloy, the silicon content was less than 70 wt%, measured by Yu.²⁵ Considering that the density of tin (7.28 g cm^{-3}) is clearly higher than that of aluminum (2.7 g cm^{-3}), gas pressure filtration is more efficient in the separation, which means that less acid is consumed and more solvent is recycled. In the centrifugal separation, Li obtained the separated silicon from Al–45% Si alloy, while its silicon content was close to 90 wt% (91.4 vol%) and most of the available silicon was kept in it.¹⁰ In this study, the comparable results were achieved in Sn–Si alloy by gas pressure filtration (80.0 wt% converting to 92.6 vol%). Since this method can be more easily scaled up, the combination of solvent refining and gas pressure filtration can serve an important role in the MG-Si purification.

4. Conclusions

In this study, the combination of solvent refining and gas pressure filtration was proved to be a promising method. After solvent refining process, efficient separation of silicon from solvent could be obtained at a low pressure differential. Optimal separation result was obtained at $p = 0.30 \text{ MPa}$, $T = 250 \text{ }^\circ\text{C}$, and $\omega_{(a)} = 20 \text{ wt\%}$, and 93.6% of tin was separated into the filtrate, while almost all the silicon was recovered and formed the separated silicon with a silicon content of 80.0 wt%. A better separation contributed to the removal of impurity. After a twice-refining process, most impurities were eliminated and 94.9% of B was removed.

Abbreviations

The following abbreviations are used in this manuscript:

SoG-Si	Solar-grade silicon
MG-Si	Metallurgical-grade silicon
CFF	Carbon fiber felt

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors acknowledge the financial supports from the National Natural Science Foundation of China (No. 51804030), the National Key Research and Development Program (2016YFB0601304), and the Fundamental Research Funds for the Central Universities (FRF-TP-19-010A2).

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