

## RESEARCH ARTICLE

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# Aggregation enhances luminescence and photosensitization properties of a hexaiodo-BODIPY†

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Intramolecular rotations are known to interfere with the excited-state processes resulting in luminescence quenching. Efforts aimed at restricting intramolecular rotations have led to the development of highly luminescent systems. Here we report the luminescence properties of a BODIPY molecule (**I**<sub>6</sub>) containing six iodo-substituents. The luminescence of **I**<sub>6</sub> was highly quenched in the solution state whereas the conversion of this molecule into its nanoaggregates enhanced the emission intensity. Detailed photophysical studies and molecular simulations were employed to study the underlying mechanism which showed that the electron-withdrawing nature of iodine atoms and the peculiar structure of **I**<sub>6</sub> contributed towards the anomalous luminescence behaviour. Further, the luminescence and photosensitization properties of **I**<sub>6</sub> were exploited for cell imaging and photodynamic therapy. Thus, our simple but intelligent molecular design yielded a multi-faceted molecule having excellent photophysical properties for potential photobiological applications.

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## Introduction

Molecular motions are critical in determining luminescence efficiency of molecules. Intramolecular rotations about a C–C bond joining aromatic rings is a well-known non-radiative pathway in organic fluorophores. Any process that restricts the intramolecular rotations serves to hinder the non-radiative decays thus leading to enhanced emission. This principle served as the platform for the development of fluorophores exhibiting aggregation induced emission (AIE).<sup>1–5</sup> The signature of the AIE phenomenon is a weak emission in dilute solutions which intensifies upon aggregation or in the solid state. AIE is typically exhibited by molecules having propeller-type structures that allow intramolecular rotations in dilute solutions which facilitates non-radiative decays. Upon aggregation, the intramolecular rotations are restricted thereby hindering the non-radiative decay processes and eventually leading to an enhancement in emission. Because of the inverted emission characteristics, AIE luminogens were readily accepted in biology and materials science. These materials are ideal for biomedical applications because of the superior luminescence features and excellent stability in the aggregated state.<sup>6–11</sup>

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Boron-dipyrromethene (BODIPY) dyes are commonly used as fluorescent indicators, bio labelling agents, photosensitizers and as active material in optoelectronics owing to their exceptional optical and electronic properties.<sup>12–19</sup> BODIPYs are known for their sharp absorption and emission bands, high fluorescence quantum yields and small Stokes shifts in dilute solutions.<sup>20</sup> However, because of the hydrophobic core, non-functionalized BODIPYs undergo aggregation in the aqueous medium and their benchmark photophysical properties are compromised under these conditions.<sup>21</sup> The utility of simple BODIPY molecules are thus limited for biological applications.

Efforts are ongoing to develop new BODIPYs that retain their excellent photophysical properties in solution as well as in the aggregated/solid state.<sup>22</sup> A commonly adopted strategy in this direction is to attach bulky molecules on the main backbone to reduce the intramolecular rotations.<sup>23–32</sup> In another approach, BODIPYs are substituted with known moieties that exhibit the AIE phenomenon.<sup>33–39</sup> It has also been shown that molecules with twisted BODIPY cores show interesting luminescence properties.<sup>40–42</sup> Although a fair amount of success has been achieved in terms of luminescence properties, many of these strategies required multi-step synthetic routes which limits their widespread utility. Moreover, functional groups should be judiciously selected so as to improve cellular uptake and reduce toxicity if these molecules are to be employed for biological applications. Thus, design and synthesis of simple BODIPY dyes that retain their excellent photophysical properties

under a variety of conditions remains a challenging area of research.

Herein we report the synthesis and photophysical properties of a simple BODIPY molecule (**I**<sub>6</sub>) bearing six iodine atoms. Because of the peculiar structure, **I**<sub>6</sub> exhibited low emission in solution which could be turned-on by aggregation. Iodine atoms facilitated intersystem crossing thus allowing generation of singlet oxygen. Preliminary photobiological studies revealed that this molecule has the potential to be used as an efficient photosensitizer for photodynamic therapy (PDT).

## Results and discussion

The synthesis of the BODIPY molecule **I**<sub>6</sub> was achieved as shown in Fig. 1a. 1,3,5-Tribenzene carbonyl chloride on condensation with 2,4-dimethylpyrrole followed by complexation with  $\text{BF}_3\text{-Et}_2\text{O}$  yielded the BODIPY **B**<sub>3</sub> in moderate yields.<sup>43</sup> Iodination of **B**<sub>3</sub> with *N*-iodosuccinimide in dichloromethane gave **I**<sub>6</sub> in 88% yield. Both **B**<sub>3</sub> and **I**<sub>6</sub> were characterized by NMR and high resolution mass spectrometry (Fig. S1, ESI<sup>†</sup>).

At the outset, we were interested in studying the photophysical properties of **B**<sub>3</sub> and **I**<sub>6</sub> under a variety of conditions. As shown in Fig. 2 and Table 1, **B**<sub>3</sub> and **I**<sub>6</sub> showed absorption maxima at 502 and 520 nm, respectively in acetonitrile whereas the corresponding fluorescence maxima were observed at 525 and 550 nm. **B**<sub>3</sub> was observed to be negligibly fluorescent in solution ( $\phi_f = 0.0005$ <sup>43</sup>) whereas **I**<sub>6</sub> was observed to be moderately fluorescent. The fluorescence quantum yield and lifetime of **I**<sub>6</sub> were found to be 0.03 and 1.65 ns in acetonitrile (Fig. S2, ESI<sup>†</sup>), and the radiative ( $k_r$ ) and non-radiative ( $k_{nr}$ ) decay constants were calculated to be  $1.8 \times 10^8$  and  $1.8 \times 10^7 \text{ s}^{-1}$ , respectively. It is noteworthy that **I**<sub>6</sub> was observed to be fluorescent in the solid state as well. The solid state absorption

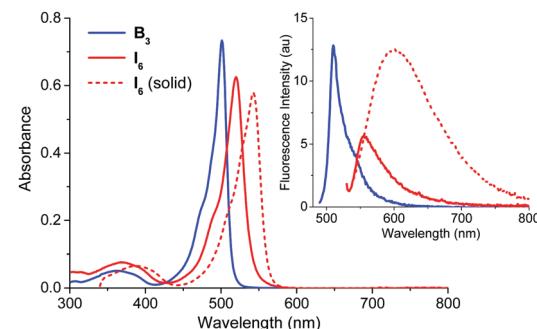


Fig. 2 Absorption and (inset) emission spectra of **B**<sub>3</sub> (6  $\mu\text{M}$ ) and **I**<sub>6</sub> (4.5  $\mu\text{M}$ ) in acetonitrile (solid lines) and that of **I**<sub>6</sub> in the solid state (dashed lines). The emission spectra were obtained by exciting the samples at the respective absorption maxima.

Table 1 Photophysical properties of various BODIPY molecules

	$\lambda_{\text{abs}}$ (nm)		$\lambda_{\text{em}}$ (nm)		$\phi_f$	$\phi_\Delta$	
	<i>a</i>	<i>b</i>	<i>a</i>	<i>b</i>		<i>a</i>	<i>b</i>
<b>B</b>	501 <sup>45</sup>	550 <sup>46</sup>	510 <sup>45</sup>	505 <sup>46</sup>	0.64 <sup>45</sup>	<i>c</i>	0.008 <sup>46</sup>
<b>B</b> <sub>2</sub>	502	<i>c</i>	<i>c</i>	<i>c</i>	0.02 <sup>47</sup>	<i>c</i>	0.071
<b>B</b> <sub>3</sub>	501	502	510	550	0.0005 <sup>43</sup>	0.013	<i>d</i>
<b>I</b> <sub>2</sub>	533 <sup>45</sup>	<i>c</i>	537 <sup>45</sup>	<i>c</i>	0.02 <sup>45</sup>	<i>c</i>	0.73 <sup>45</sup>
<b>I</b> <sub>4</sub>	540 <sup>44</sup>	<i>c</i>	<i>c</i>	<i>c</i>	<i>c</i>	<i>c</i>	0.68 <sup>44</sup>
<b>I</b> <sub>6</sub>	520	506	550	566	0.03	0.15	0.21
							0.82

<sup>a</sup> In solution. <sup>b</sup> As aggregates in aqueous medium. <sup>c</sup> Not reported.

<sup>d</sup> Not determined as photo-bleaching was observed.

spectrum of **I**<sub>6</sub> was slightly broader with a maximum at 542 nm and the emission maximum appeared at 604 nm. The fluorescence lifetime of **I**<sub>6</sub> was significantly longer ( $\tau = 4.71$  ns) in the solid state as compared to that in solution. On the other hand, **B**<sub>3</sub> was not observed to be emissive in the solid state. The bathochromically shifted absorption and emission peaks and longer fluorescence lifetime of **I**<sub>6</sub> in the solid state indicates aggregation of the molecules in the solid state.

The absorption and emission properties of **B**<sub>3</sub> and **I**<sub>6</sub> are anomalous as compared to typical BODIPY dyes which are known for their sharp absorption and fluorescence spectra with high quantum yields that are typically independent of their environment. To gain insights into the luminescence behaviour of **I**<sub>6</sub>, we studied the effect of temperature and viscosity on the fluorescence of **I**<sub>6</sub>. When a solution of **I**<sub>6</sub> in acetonitrile was heated from 5 to 60 °C, the absorption spectrum remained unaffected whereas the emission intensity decreased regularly (Fig. 3a). At 60 °C, the fluorescence emission intensity was ~6 times lower as compared to that at 5 °C. The effect of viscosity on the absorption and emission spectra of **I**<sub>6</sub> was then studied. When a highly viscous solvent such as glycerol was added to a solution of **I**<sub>6</sub> in acetonitrile, we observed broadening of the absorption spectrum (Fig. 3b). Further, a new peak emerged at 463 nm and became prominent as the amount of glycerol was increased to 90%. Under similar conditions, the fluorescence intensity of **I**<sub>6</sub> was observed to increase ~17 times. However, we observed turbidity in the solution as the amount of glycerol was increased above 50% thereby indicating the aggregation of **I**<sub>6</sub>.

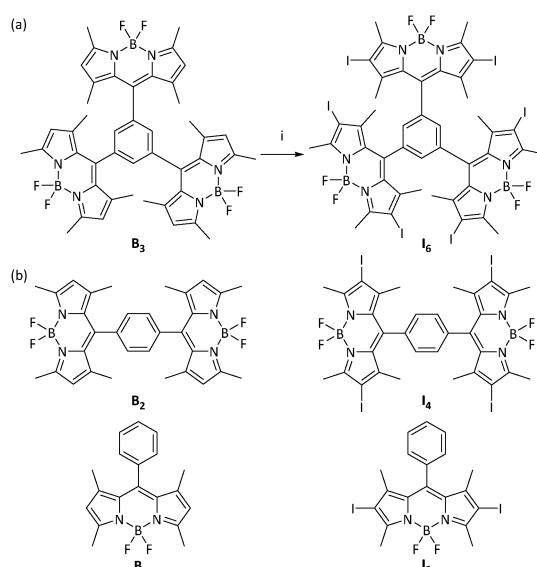


Fig. 1 (a) Synthesis of **I**<sub>6</sub>. Reactions conditions: (i) *N*-iodosuccinimide, dichloromethane, 25 °C, 4 h. (b) Chemical structures of related BODIPY molecules reported elsewhere.<sup>44,45</sup>

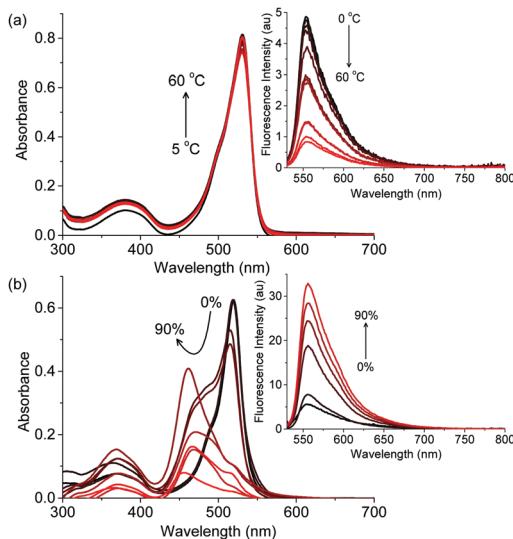


Fig. 3 Changes in the absorption and (inset) emission spectra of **I**<sub>6</sub> (4.5  $\mu$ M) as a function of varying (a) temperature and (b) amount of glycerol in acetonitrile.  $\lambda_{\text{ex}}$  is 520 nm.

As glycerol was promoting aggregation of **I**<sub>6</sub>, we were intrigued to investigate whether other solvents induced similar changes. In this context, the effect of solvent polarity was evaluated by recording the absorption and emission of **I**<sub>6</sub> in different solvents. We observed that in non-polar solvents **I**<sub>6</sub> showed a red-shift of  $\sim$ 20 nm in the absorption maximum as compared to that in polar solvents (Fig. S3 and S4, ESI<sup>†</sup>). A similar trend was observed in the emission behaviour as well wherein we observed a maximum at 545–550 nm in polar solvents whereas in non-polar solvents the emission maximum was red shifted by  $\sim$ 40 nm. These observations indicated the non-polar character of the **I**<sub>6</sub> molecule.

Our next objective was to study the aggregation of **I**<sub>6</sub> in acetonitrile–water mixtures. **I**<sub>6</sub>, which showed an absorption peak at 520 nm in acetonitrile, exhibited significantly broadened and red-shifted absorption upon increasing water content. In 90% water–acetonitrile mixture, the long-wavelength absorption band was observed to redshift by 7 nm and a new sharp absorption peak appeared at 445 nm (Fig. 4a). The fluorescence intensity of **I**<sub>6</sub> was observed to increase with the addition of water and we noticed  $\sim$ 13-times enhancement in the fluorescence in 90% water–acetonitrile mixture. The fluorescence quantum yield of **I**<sub>6</sub> determined in 90% water–acetonitrile mixture was found to be  $\phi_f = 0.075$  which was 4-times higher than that in acetonitrile. Moreover, these changes were significant enough to result in visual changes in the fluorescence of **I**<sub>6</sub> and we observed a change in colour of **I**<sub>6</sub> from transparent pale green in acetonitrile to bluish-green in 90% water–acetonitrile mixture. Under similar experimental conditions, we also studied the effect of aggregation on the photophysical properties of **B**<sub>3</sub>. However, in contrast to the results obtained with **I**<sub>6</sub>, **B**<sub>3</sub> exhibited negligible changes in the absorption and emission spectra in acetonitrile–water mixtures thereby ruling out aggregation induced emission (Fig. S5, ESI<sup>†</sup>).

Since heavy atoms like bromine and iodine are known to enhance the intersystem crossing efficiency of organic dyes,<sup>48–51</sup>

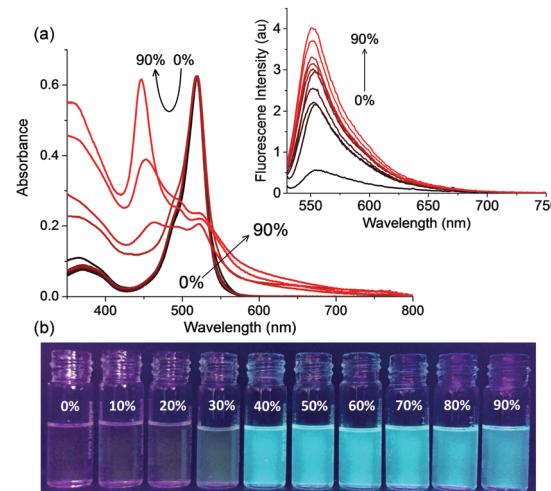


Fig. 4 (a) Absorption and (inset) emission spectra of **I**<sub>6</sub> (4.5  $\mu$ M) in acetonitrile–water mixtures.  $\lambda_{\text{ex}} = 520$  nm. (b) Visual fluorescence of **I**<sub>6</sub> upon excitation with a UV lamp ( $\lambda_{\text{ex}} = 365$  nm) in acetonitrile–water mixtures.

we inferred that triplet excited states will be formed in **I**<sub>6</sub>. The formation of triplet excited states can facilitate generation of reactive oxygen species thus making **I**<sub>6</sub> suitable as a photosensitizer for photodynamic therapy. To test this hypothesis, we studied the photosensitization properties of **I**<sub>6</sub> in acetonitrile using 1,3-diphenylisobenzofuran (DPBF) as a trapping agent for singlet oxygen. On photoirradiation of a mixture of **I**<sub>6</sub> and DPBF in acetonitrile, we observed a gradual decrease in the absorbance of DPBF at 420 nm whereas the absorbance of **I**<sub>6</sub> at 520 nm remained unperturbed (Fig. S6 and S7, ESI<sup>†</sup>). These results indicate that **I**<sub>6</sub> was photostable under our experimental conditions and that it sensitizes the generation of singlet oxygen. The presence of singlet oxygen was further confirmed by recording its characteristic emission at 1290 nm<sup>52</sup> in acetonitrile (Fig. S8, ESI<sup>†</sup>). The efficiency of singlet oxygen generation of **I**<sub>6</sub> was quantified by calculating the quantum yield using methylene blue as a reference standard and was obtained as  $\phi_A = 0.21$  (Fig. S9, ESI<sup>†</sup>).

As **I**<sub>6</sub> exhibited enhanced emission in acetonitrile–water mixtures and was capable of generating singlet oxygen, we were interested in studying if nanoaggregates of **I**<sub>6</sub> could be synthesized for studying the photosensitization properties in the aqueous medium. These nanoaggregates would be dispersible in the aqueous medium and would show superior fluorescence and photosensitization properties thereby making it an ideal active ingredient for image-guided PDT. Nanoaggregates of **I**<sub>6</sub> (denoted hereafter as **I**<sub>6</sub>-NA) were synthesized by the re-precipitation method wherein a concentrated solution of **I**<sub>6</sub> in acetone was added to large excess of water and the nanoaggregates were collected by centrifugation (see the experimental section for the detailed procedure). **I**<sub>6</sub>-NA was dispersed in water and characterized by dynamic light scattering (DLS) and microscopy. As shown in Fig. S10 (ESI<sup>†</sup>), DLS analysis showed an average particle size of  $85.6 \pm 3.1$  nm for **I**<sub>6</sub>-NA. Similarly, high resolution transmission electron microscopy measurements (Fig. 5a) showed the presence of spherical particles and the

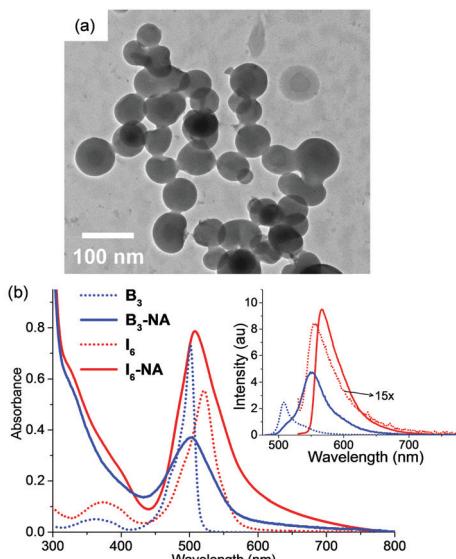


Fig. 5 (a) Transmission electron micrograph of **I<sub>6</sub>-NA** in water. (b) Absorption and (inset) fluorescence spectra of **B<sub>3</sub>** (6  $\mu$ M) and **I<sub>6</sub>** (4.5  $\mu$ M) in acetonitrile and **B<sub>3</sub>-NA** (10  $\mu$ g mL<sup>-1</sup>) and **I<sub>6</sub>-NA** (10  $\mu$ g mL<sup>-1</sup>) in water. The emission spectra were obtained by exciting the samples at the respective absorption maxima.

particle size corroborated well with that obtained from DLS experiments.

It is assumed that **I<sub>6</sub>** undergoes spontaneous self-assembly during the nanoaggregate formation the evidence for which was obtained from <sup>19</sup>F NMR spectroscopy. As shown in Fig. S11 (ESI<sup>†</sup>), **I<sub>6</sub>** in CDCl<sub>3</sub> showed a multiplet from  $\delta$  -145.94 to -146.28 ppm corresponding to the -BF<sub>2</sub> moieties. On the other hand, we observed a broad singlet at -119.28 ppm in the <sup>19</sup>F NMR spectrum of **I<sub>6</sub>-NA** in D<sub>2</sub>O. Although a comparison of the NMR spectra in organic and aqueous media are inappropriate, the peak broadening and the downfield shift indicates the aggregation of **I<sub>6</sub>** in **I<sub>6</sub>-NA**. The formation of **I<sub>6</sub>-NA** was also accompanied by changes in the absorption and emission spectra of **I<sub>6</sub>** (Fig. 5b). **I<sub>6</sub>-NA** showed a broad absorption peak that was blue shifted by 13 nm as compared to **I<sub>6</sub>** in acetonitrile whereas the emission maximum for **I<sub>6</sub>-NA** was observed at 568 nm. The fluorescence quantum yield and lifetime of **I<sub>6</sub>-NA** were observed to be 0.15 and 3.12 ns, respectively in water which were substantially higher as compared to that of **I<sub>6</sub>** in acetonitrile. Similarly, the radiative ( $k_r = 4.8 \times 10^8$  s<sup>-1</sup>) and non-radiative ( $k_{nr} = 4.8 \times 10^7$  s<sup>-1</sup>) decay constants were also observed to be higher in the aggregated state as compared to that in acetonitrile. Nanoaggregates of **B<sub>3</sub>** (denoted as **B<sub>3</sub>-NA**) were also synthesized by following a similar procedure. Fig. S10 (ESI<sup>†</sup>) and Fig. 5b shows the characterization data and the absorption and emission spectra of **B<sub>3</sub>-NA**. It is remarkable to note that **B<sub>3</sub>-NA** was observed to be negligibly fluorescent with a quantum yield of 0.013 and this observation was in stark contrast to the fluorescence properties of **I<sub>6</sub>-NA**.

Next, we were interested in studying the photosensitization properties of the nanoaggregates in water. Similar to the experiments carried out in acetonitrile, the singlet oxygen generation

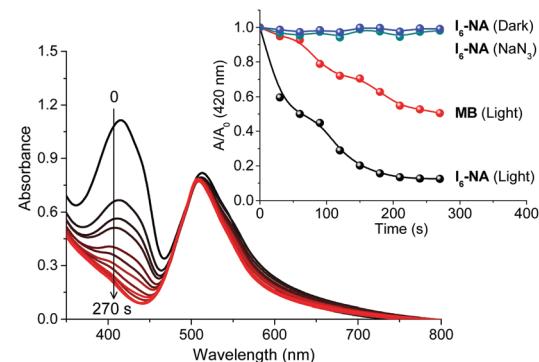


Fig. 6 Changes in the UV-Vis absorption spectrum of a mixture of 1,3-diphenylisobenzofuran (DPBF, 90  $\mu$ M) and **I<sub>6</sub>-NA** (10  $\mu$ g mL<sup>-1</sup>) as a function of irradiation time in 20% ethanol–water. Inset shows the relative decrease in the absorption of DPBF at 420 nm in the presence of **I<sub>6</sub>-NA** or methylene blue (MB) under different experimental conditions.

ability of **I<sub>6</sub>-NA** was studied using DPBF as a singlet oxygen trap. As shown in Fig. 6, we observed a gradual decrease in the absorbance of DPBF at 410 nm upon photoirradiating a mixture of **I<sub>6</sub>-NA** and DPBF. It is noteworthy that **I<sub>6</sub>-NA** was highly efficient in generating singlet oxygen as compared to **I<sub>6</sub>** and the quantum yield of singlet oxygen generation was found to be 0.82 (Fig. S9, ESI<sup>†</sup>). The formation of singlet oxygen was further confirmed by monitoring the emission of <sup>1</sup>O<sub>2</sub> at 1280 nm in water (Fig. S8, ESI<sup>†</sup>). Further, control experiments carried out in dark and in the presence of sodium azide conclusively proved the photosensitized generation of singlet oxygen by **I<sub>6</sub>-NA**.

Our results indicate that **I<sub>6</sub>-NA** generates singlet oxygen efficiently in the aqueous medium upon photoirradiation. The efficacy of **I<sub>6</sub>-NA** as photodynamic agents was tested against C6-glioblastoma cell line upon photoirradiation. When the cells were incubated with different concentrations of **I<sub>6</sub>-NA** such as 5, 10 and 20  $\mu$ g mL<sup>-1</sup> and photoirradiated for 10 minutes, we observed up to  $\sim$  60% cell death (Fig. 7a). On the other hand, cells incubated with similar concentrations of **I<sub>6</sub>-NA** and kept in dark conditions were observed to be almost 100% viable thereby confirming photosensitized generation of singlet oxygen only in light conditions. Since the nanoaggregates were fluorescent, their cellular localization was studied by incubating **I<sub>6</sub>-NA** with C6 cells followed by fixation and treatment with DAPI/Phalloidin-TRITC (nuclear/cytoskeletal labels) prior to imaging by confocal microscopy. As **I<sub>6</sub>-NA** exhibited an emission that tailed up to  $\sim$  750 nm, we could image the cells in the green and red channels. It was observed that **I<sub>6</sub>-NA** localized in the cytoplasm around the nuclear region when simultaneously scanned with 405/488/561 nm excitations. This was exemplified in the merged image of Fig. 7b wherein the **I<sub>6</sub>-NA** appears yellow around the blue region. Phalloidin-TRITC assisted in marking the boundary of the cells and its emission was clearly limited only to the cell surface as seen in control.

BODIPYs are known to be highly emissive in solution with many molecules exhibiting fluorescence quantum yield close to unity.<sup>16,20,53-55</sup> However, **B<sub>3</sub>** is an exception and it was observed to be negligibly fluorescent. The peculiar structure of **B<sub>3</sub>** could

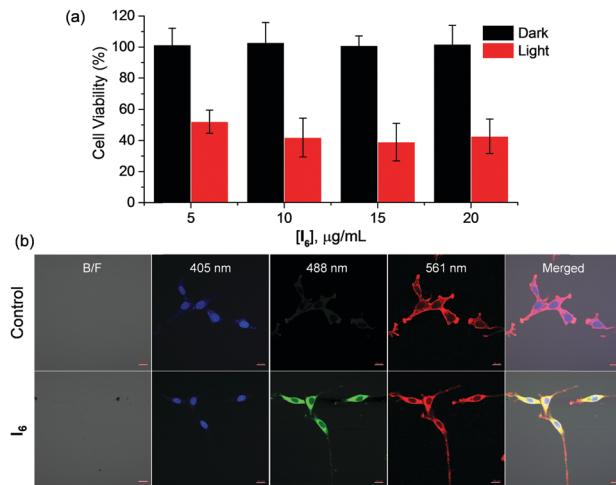


Fig. 7 (a) *In vitro* phototoxicity in C6 cells treated with **I**<sub>6</sub>-NA under dark and light conditions. (b) Confocal microscopic image of C6 cells treated with **I**<sub>6</sub>-NA (12.5  $\mu\text{g mL}^{-1}$ ) excited at different wavelengths. The cells were stained with nuclear staining dye Hoechst 33342 and phalloidin-TRITC. Scale bar is 10  $\mu\text{m}$ .

have contributed to its non-emissive character. Molecular simulations indicate that the BODIPY moieties are oriented orthogonally as compared to the central phenyl ring (Fig. 8). As compared to the previously reported monomeric and dimeric BODIPYs **B** and **B**<sub>2</sub> (Fig. 1b), BODIPY cores in **B**<sub>3</sub> are arranged in close proximity with an average distance of 7.649  $\text{\AA}$  between the BODIPY units. It is known that the presence of multiple chromophores in close proximity often leads to exciton coupling between dyes thereby resulting in significant changes in the absorption and emission properties.<sup>56–59</sup> Thus, the presence of three BODIPY units around the central phenyl ring plausibly promotes intramolecular interactions thereby resulting in quenching of fluorescence. The observation of negligible difference in the absorbance but a significant red-shift in the emission maximum of **B**<sub>3</sub> as compared to **B** and **B**<sub>2</sub> (Table 1) suggests excited state interactions. Further, intramolecular rotations about the C–C bond joining

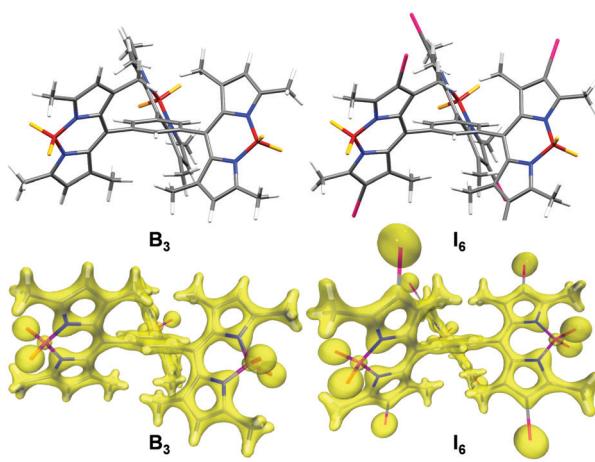


Fig. 8 Optimized geometry of **B**<sub>3</sub> and **I**<sub>6</sub> along with their electron cloud distribution.

the BODIPY unit to the phenyl ring contribute towards the quenching effect.

The photophysical behaviour of **I**<sub>6</sub> in acetonitrile was on expected lines and was comparable to **I**<sub>2</sub> and **I**<sub>4</sub>. It was recently reported that **I**<sub>4</sub> having four iodine atoms fared poorly against **I**<sub>2</sub> having two iodines in terms of singlet oxygen generation.<sup>44</sup> This observation was rationalized on the basis of the competition between heavy atom effect and configuration of **I**<sub>4</sub>. Our results support this observation wherein **I**<sub>6</sub> exhibited a further lower singlet oxygen generation efficiency of 0.21 in acetonitrile as compared to 0.73 and 0.68 for **I**<sub>2</sub> and **I**<sub>4</sub>, respectively. However, we observed a dramatic change in the photophysical behaviour of **I**<sub>6</sub> upon aggregation: the fluorescence was turned-on and a substantial increase was observed in the generation of singlet oxygen. We propose a model for the aggregation of **I**<sub>6</sub> as depicted in Fig. 9 for explaining the unforeseen behaviour of the aggregates of **I**<sub>6</sub>. Because of the peculiar geometry consisting of the orthogonal BODIPY units, the chromophore (BODIPY) in the **I**<sub>6</sub> molecule cannot form aggregates in a conventional manner wherein a large number of the chromophores self-assemble through non-covalent interactions (Fig. 9a). Instead, the self-assembly is limited to the formation of H-type dimers with respect to the chromophore units. The observation of blue-shifted absorption in glycerol and in acetonitrile–water mixture exemplifies this proposition. However, the aggregation restricts the rotation of the BODIPY units and enhances the emission intensity and the efficiency of singlet oxygen generation.<sup>60</sup> It may also be noted that the increased absorption cross-section of **I**<sub>6</sub>-NA in water as compared to **I**<sub>6</sub> in acetonitrile also contributes to the enhanced  ${}^1\text{O}_2$  generation efficiency.

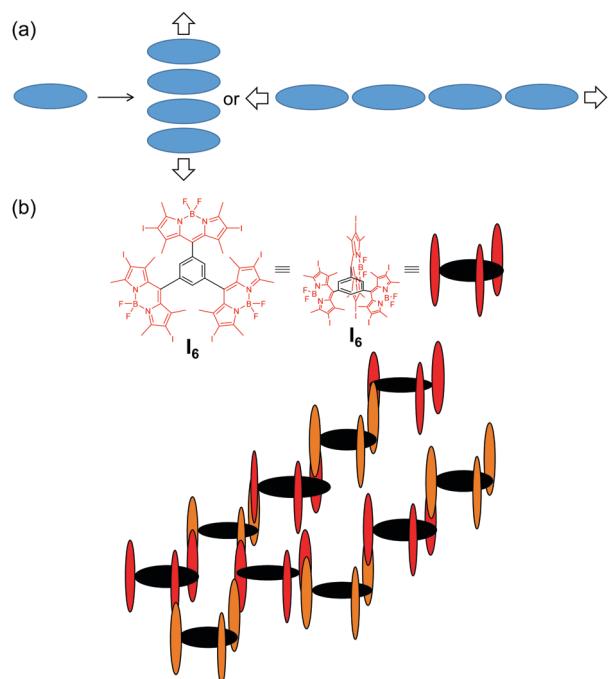


Fig. 9 Schematic representation of aggregation in (a) conventional planar molecules and (b) **I**<sub>6</sub>. The alternate molecules of **I**<sub>6</sub> are represented in different colours for better visualization.

It is intriguing to note that despite having similar structures, the aggregates of **B**<sub>3</sub> and **I**<sub>6</sub> exhibited extremely different fluorescence properties. The key to the remarkable photophysical properties of **I**<sub>6</sub>-NA is the presence of the iodo-substituents. It is proposed that the iodo-groups force the BODIPY units to stack in an off-set manner because of their steric bulkiness. Moreover, iodine atoms are known to form halogen bonds<sup>61</sup> and thus would promote intermolecular interactions. This in turn would reduce the intramolecular excitonic interactions between the individual BODIPY units in **I**<sub>6</sub> as compared to that in **B**<sub>3</sub> thereby resulting in an enhanced fluorescence and photosensitization efficiency. The progressive red-shift in the emission spectrum of **I**<sub>6</sub> from solution to the aggregated state in water and further to the solid state could also be rationalized on the basis of halogen bonding which would be stronger in the solid state as compared to that in **I**<sub>6</sub>-NA in water due to the competition from water molecules.

## Conclusions

We synthesised a hexaiodinated BODIPY molecule with interesting photophysical properties. The molecule showed low fluorescence in acetonitrile whereas the emission was enhanced significantly in the aggregated state. Similarly, the photosensitization properties were also observed to be more efficient in the aggregated state as compared to that in solution. Photobiological studies showed that this molecule was capable of functioning as a photosensitizer and imaging agent for photodynamic therapy. Molecular simulations indicated that the peculiar arrangement of the BODIPY units and the presence of iodine atoms as responsible for its emission characteristics.

## Conflicts of interest

There are no conflicts to declare.

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