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Ligand-centered reactivity of a pseudo-dearomatized phosphorus-nitrogen PN^3P^* rhodium complex towards molecular oxygen at room temperature†

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The ligand of an organometallic complex is typically considered as a spectator which can be modified to tune the steric and electronic properties of the coordination environment. We herein demonstrated a $\text{PN}^3\text{P}^*\text{Rh-CO}$ pincer system where one of the C–H bonds of the pseudo-dearomatized pyridine ring was oxidized by O_2 at room temperature to create an α,β -unsaturated carbonyl functionality on the ligand backbone. The resulting metal complex with the post-modified PN^3P ligand readily reacts with thiophenol and 4-methylaniline to afford the corresponding oxidative Michael addition products.

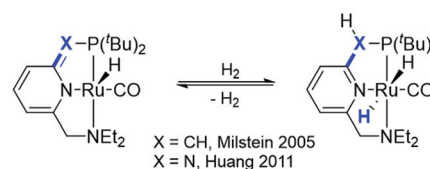
Introduction

Pincer ligand platforms have become popular due to the increased stability of the corresponding metal complexes arising from the tridentate binding mode. The tunability of stereo and electronic properties of the ligands allows broad applications in catalysis.¹ Early investigations on pincer complexes mainly focused on synthesis, coordination chemistry,² and substitution reactions on the metal center.³ It was not until 2005 when Milstein and co-workers demonstrated a new type of metal–ligand cooperation (MLC, Scheme 1a, X = CH) in a pyridine-based pincer ruthenium complex for the catalytic hydrogenation of alcohols to esters.⁴ The MLC proceeds *via* the dearomatization and rearomatization of the central pyridine ring of the pincer ligand backbone with the metal oxidation state unchanged. Such a bond activation mode has received much attention in the development of pincer catalysts.⁵ We have also recently demonstrated that the seemingly small change, by replacing the spacer arms from the CH_2 to NH group(s), has dramatically altered the thermodynamic and kinetic properties compared to those of the CH_2 analogs (Scheme 1a, X = N).^{1c,6} This new class of $\text{PN}^3(\text{P})$ metal complexes offered intriguing

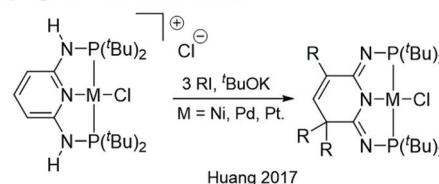
properties due to the distinct pseudo-dearomatized nature,⁷ greatly extending the potential for different reactivities in catalysis and bond activation.⁸ Accordingly, metal enhanced ligand centered reactivities were realized.⁹ Very interesting, such properties enabled the preparation of a second generation of PN^3P pincer complexes *via* a facile ligand post-modification to diversify the pincer platforms (Scheme 1b).¹⁰

Dioxygen is a triplet diradical ($^3\text{O}_2$) in its ground state and its reaction with a singlet closed shell organometallic complex is therefore spin-forbidden and quite unexpected. Only a few examples where a ligand is oxidized by O_2 could be found in the literature. For example, cobaltocene and other cobalt clusters can form peroxo-bridged dimers in the presence of O_2 .¹¹ A similar observation was made in $\text{Cp}(\text{Fe}^1)\text{arene}$ (Scheme 2a).¹²

(a) Metal-Ligand Cooperation of Pyridine-based Pincer Complexes



(b) Ligand Post-modification



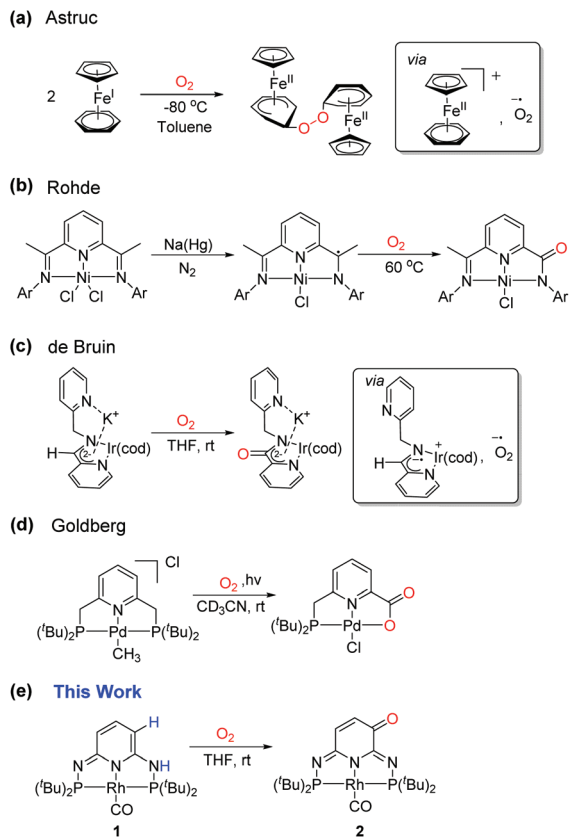
Scheme 1 Various reactivities in pincer metal complexes.

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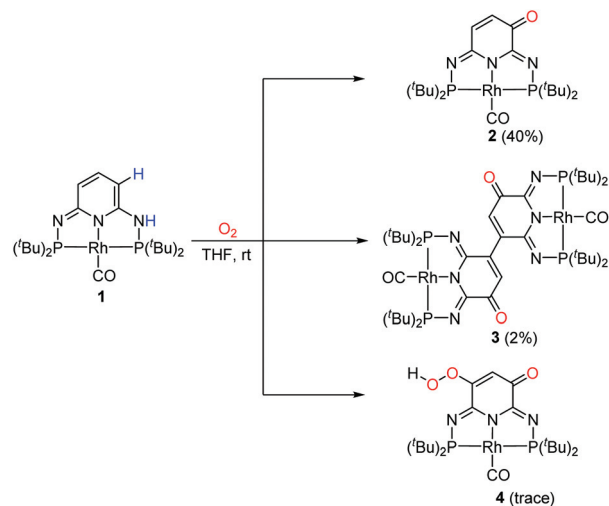
Scheme 2 Reactions between organometallic complex and O₂.

This reaction was proposed to proceed *via* a single-electron transfer (SET) from the Cp(Fe^I)arene complex to O₂ to create a superoxide radical anion, followed by the nucleophilic attack on the arene ring. The driving force for the SET was explained by the large difference of 1.0 V in redox potential between the Fe⁺/Fe²⁺ and O₂/O₂⁻ couples.¹³ In a bis(arylimino)pyridine nickel pincer complex, a radical could be chemically introduced on the ligand and could subsequently react with O₂ (Scheme 2b).¹⁴ The mono-oxygenation of a ligand backbone occurred at room temperature in a bis(picoyl)amine iridium complex *via* a ligand-centered radical intermediate as observed in the EPR spectroscopy (Scheme 2c).¹⁵ The PNP-ligated Pd-Me complex was found to react with O₂ leading to an oxidation of the CH₂ arm of the pincer ligand upon photolysis (Scheme 2d).¹⁶ Herein, we disclose yet another unique example of ligand reactivity towards O₂ in our PN³P*Rh-CO complex (**1**).^{8a} The subsequent reactions of the resulting α,β-unsaturated carbonyl functionality on the ligand backbone were also investigated.

Results and discussion

Reaction of **1** with O₂

The reaction with O₂ and **1** proceeded with a rapid color change from green to red-brown in 2 hours at room tempera-

Scheme 3 Reaction of **1** with O₂ at room temperature.

ture. **2** and **3** were isolated by flash chromatography as the major products in 40% and 2% yields, respectively, with the formation of a trace amount of **4** (Scheme 3). The ³¹P{¹H} NMR spectrum of **2** displays an AB system with two sets of doublet of doublet at δ 140.8 ppm (²J_{PP} = 252.5 Hz, ¹J_{RhP} = 127.3 Hz) and δ 139.5 ppm (²J_{PP} = 252.5 Hz, ¹J_{RhP} = 127.3 Hz), consistent with the non-equivalent phosphine groups. The ¹H NMR spectrum shows two different signals, a doublet at δ 6.91 ppm and multiplet at δ 6.10 ppm, in a 1 : 1 ratio, for two protons on the ligand backbone. The rhodium carbonyl carbon of **2** can be observed as a doublet of triplet at δ 199.80 ppm (¹J_{RhC} = 68.6 Hz, ²J_{PC} = 10.8 Hz) in the ¹³C NMR, while the carbonyl carbon on the ligand backbone appears at δ 176.46 ppm. Similarly, **3** also shows two sets of doublet of doublet at δ 141.4 ppm (²J_{PP} = 253.2 Hz, ¹J_{RhP} = 127.3 Hz) and δ 138.7 ppm (²J_{PP} = 253.2 Hz, ¹J_{RhP} = 127.3 Hz). The ¹H NMR spectrum exhibits a broad signal at δ 6.50 ppm, in agreement with only one proton remaining on the backbone.

Crystals suitable for X-ray analysis of **2** (Fig. 1), **3** (Fig. 2) and **4** (Fig. S14†) were obtained from their respective benzene solutions. **2** exhibits a distorted square-planar geometry around the tetracoordinated Rh(I) center with angles of P1-Rh1-P2 and N1-Rh1-C6 as 159.33 (2)° and 178.65(13)°, respectively. The C1-N3 (1.285 Å) and C5-N2 (1.291 Å) bond distances strongly indicate the double bond characters, in good agreement with the solution NMR data of **2**. **3** displays a dimeric structure resulting from the oxidative coupling of the backbone of **2** linked by formation of a C-C bond (C2-C23: 1.492 Å), with a distorted square-planar geometry around two Rh(I) centers (P1-Rh1-P2: 159.05(17), C43-Rh1-N1: 178.3(7), P4-Rh2-P3: 160.3(3) and C44-Rh2-N4: 178.2(7)). **4** shows a similar structure with the only significant difference being the presence of an -OOH moiety connecting to the C5 atom of the oxidized pyridine ring. The crystallography analysis also confirmed the tetracoordinated environment for the Rh(I) center (P2-Rh2-P3: 159.19(5)° and N2-Rh1-C6: 179.6(2)°).



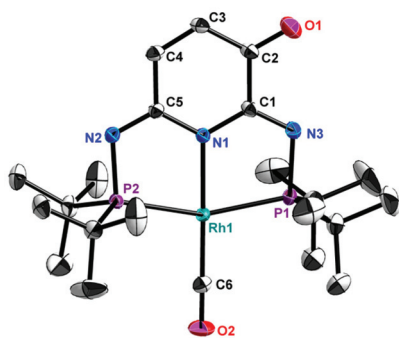


Fig. 1 Molecular structure of **2**. Thermal ellipsoids are shown at the 50% probability level. Selected bond lengths [Å] and bond angles [°]: C2–O1, 1.233(4); C3–C4, 1.358(4); C6–O2, 1.150(4); C1–N3, 1.285(3); C5–N2, 1.291(3); Rh1–N1, 2.059(2); Rh1–C6, 1.824(3); Rh1–P1, 2.2921(6); Rh1–P2, 2.2918(7); P2–Rh1–P1, 159.33(2); N1–Rh1–C6, 178.65(13); P2–Rh1–N1, 79.70(6); N1–Rh1–P1, 79.63(6); P2–Rh1–C6, 100.76(10); C6–Rh1–P1, 99.91(10).

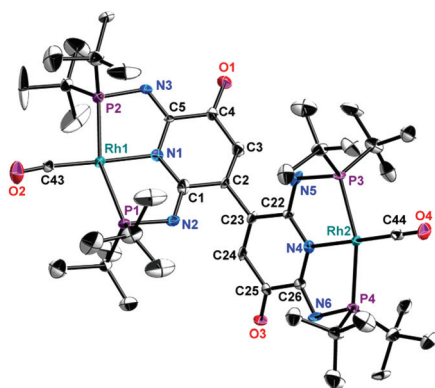
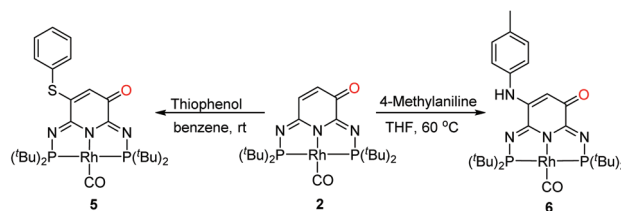


Fig. 2 Molecular structure of **3**. Thermal ellipsoids are shown at the 50% probability level. Selected bond lengths [Å] and bond angles [°]: C4–O1, 1.23(2); C2–C3, 1.30(3); C43–O2, 1.147(19); C5–N3, 1.28(2); C1–N2, 1.28(2); Rh1–N1, 2.040(13); Rh1–C43, 1.834(16); Rh1–P1, 2.289(5); Rh1–P2, 2.292(5); C2–C23, 1.49(2); C25–O3, 1.23(2); C23–C24, 1.35(3); C44–O4, 1.17(2); C26–N6, 1.32(2); C22–N5, 1.28(2); Rh2–N4, 2.054(12); Rh2–C44, 1.82(2); Rh2–P3, 2.278(11); Rh2–P4, 2.276(5); N1–Rh1–C43, 178.3(7); P1–Rh1–P2, 159.05(17); N1–Rh1–P1, 79.8(4); P1–Rh1–C43, 101.5(6); C43–Rh1–P2, 99.4(6); P2–Rh1–N1, 79.3(4); N4–Rh2–C44, 178.2(7); P3–Rh2–P4, 160.3(3); N4–Rh2–P3, 80.8(5); P3–Rh2–C44, 98.5(6); C44–Rh2–P4, 100.7(6); P4–Rh2–N4, 79.9(4).

Oxidative Michael addition of **2**

The ligand backbone of **2** has an α,β -unsaturated carbonyl functionality which may react as a Michael acceptor. The reactions of thiophenol and 4-methylaniline were performed under air to produce **5** and **6** in 95% and 60% yields, respectively (Scheme 4). The $^{31}\text{P}\{^1\text{H}\}$ NMR spectra of **5** and **6** both show an AB pattern with two sets of doublet of doublet at δ 138.7 ppm ($^2J_{\text{PP}} = 252.7$ Hz, $^1J_{\text{RHP}} = 128.0$ Hz) and δ 137.0 ppm ($^2J_{\text{PP}} = 252.7$ Hz, $^1J_{\text{RHP}} = 128.0$ Hz), and δ 141.8 ppm ($^2J_{\text{PP}} = 254.2$ Hz, $^1J_{\text{RHP}} = 127.3$ Hz) and δ 135.6 ppm ($^2J_{\text{PP}} = 254.2$ Hz, $^1J_{\text{RHP}} = 127.3$ Hz), respectively. The ^1H NMR spectra exhibit a broad signal at δ 5.72 ppm and δ 6.38 ppm for **5** and **6**, respect-



Scheme 4 Oxidative Michael addition of **2**.

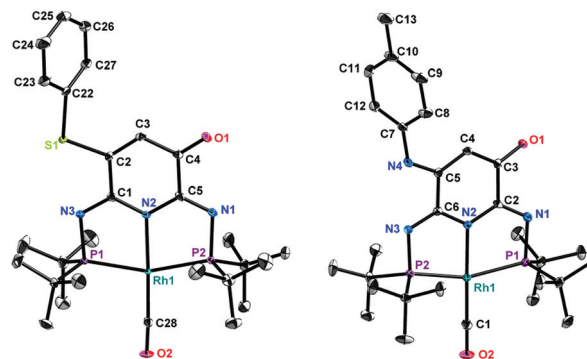


Fig. 3 Molecular structures of **5** (left) and **6** (right). Thermal ellipsoids are shown at the 50% probability level; hydrogen atoms are omitted for clarity. Selected bond lengths [Å] and bond angles [°]: for **5**: C4–O1, 1.220(3); C2–C3, 1.346(4); C2–S1, 1.753(3); C28–O2, 1.151(4); C1–N3, 1.295(4); C5–N1, 1.294(4); Rh1–N2, 2.061(2); Rh1–P1, 2.2849(7); Rh1–C28, 1.824(3); N2–Rh1, 2.2978(7); N2–Rh1–C28, 177.65(10); P1–Rh1–P2, 159.49(2); N2–Rh1–P1, 80.19(6); P1–Rh1–C28, 98.42(8); C28–Rh1–P2, 101.94(8); P2–Rh1–N2, 79.38(6). For **6**: C3–O1, 1.235(6); C4–C5, 1.362(7); C5–N4, 1.351(7); C1–O2, 1.156(7); C6–N3, 1.295(7); C2–N1, 1.291(8); Rh1–N2, 2.049(4); Rh1–C1, 1.810(6); Rh1–P2, 2.3054(15); Rh1–P1, 2.2894(15); N2–Rh1–C1, 179.0(3); P2–Rh1–P1, 158.56(6); N2–Rh1–P2, 79.85(13); P2–Rh1–C1, 100.85(19); C1–Rh1–P1, 99.78(19); P1–Rh1–N2, 79.46(13).

ively. These observations implied that **2** underwent an oxidative Michael addition. The molecular structures of **5** and **6** were finally confirmed by single crystal X-ray diffraction analysis (Fig. 3), revealing similar geometries to that of **4**.

Mechanistic insights for the reaction of **1** with O_2

In order to gain insight into the mechanism of the oxidation reaction between **1** and O_2 , multiple spectroscopic investigations were performed. From the similar reactions between metal complexes and O_2 shown in Scheme 2, a radical mechanism is highly plausible. Therefore, the reaction was investigated using the EPR spectroscopy. A toluene solution of **1** was exposed to O_2 and frozen in liquid nitrogen. The X-band EPR data at 10 K show signals originating from two paramagnetic species which grow together (Fig. 4A). One species has a g value of 2.06 and 1.99, showing typical characteristics of superoxide species,¹⁷ and the other has a sharp signal at $g = 2.004$, suggesting that the unpaired electron of these species is delocalized on the pyridine ligand. After being left at room temperature for 2 hours, the EPR signal gradually disappeared.



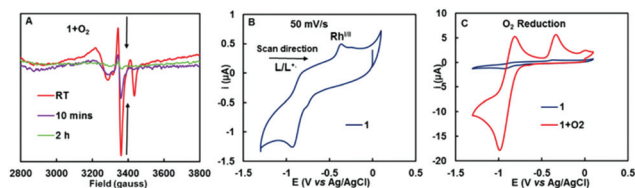


Fig. 4 (A) Detection of reaction intermediate in the reaction between **1** and O₂ with EPR spectroscopy. Conditions: *T* = 10 K, freq. = 9.3931 GHz, att. = 20 dB. (B) Cyclic voltammogram of **1** in THF (1.0 mM of complex **1**, scan rate: 50 mV s⁻¹, Glassy carbon electrode (GC, diameter 3 mm) as a working electrode, Pt-wire as a counter electrode, Ag/AgCl/1 M KCl as a reference electrode, 0.1 M ⁿBu₄NPF₆ as a supporting electrolyte. (C) Electrochemical investigation of **1** in presence of O₂, with clear indication of reduction current (enhancement of current) (scan rate: 50 mV s⁻¹).

With the presence of a ligand-centered radical intermediate established, cyclic voltammetry was performed to gain quantitative information about the charge transfer properties of **1**. Ideally, it should be possible to identify the possible redox couples of **1**, especially the ligand-centered redox couple **1**/**1**⁺. The hypothesis is to observe a ligand-centered redox couple that is more negative than that of the O₂/O₂⁻ couple, thereby explaining the driving force behind the electron transfer from the pincer ligand to O₂.

Electrochemical analysis of **1** showed a quasi-reversible wave in the cyclic voltammogram with *E*_{1/2} = -0.89 V (ΔE_p = 60 mV) vs. Ag/AgCl/1 M KCl in THF, corresponding to the ligand⁺/ligand redox couple (Fig. 4B). One broad redox couple was also observed at -0.56 V (*E*_{pc} = -0.77 V; *E*_{pa} = -0.35 V), corresponding to Rh^{II/I}. The CVs corresponding to two redox processes were obtained at different scan rates (Fig. S13A†) and showed a linear relationship, in agreement with a diffusion-controlled process (Fig. S13B and C†). The redox potential of O₂/O₂⁻ is -0.50 V vs. NHE (*i.e.* -0.70 V vs. Ag/AgCl),¹⁸ suggesting that the outer-sphere electron transfer (ET) from the pseudo-dearomatized pyridine ligand to O₂ should be exergonic by 4.3 kcal mol⁻¹. Thus, ET to O₂ is thermodynamically favorable, and a concerted process may be operative. Furthermore, electrochemical investigation of **1** in the presence of O₂, showed a clear indication of oxygen reduction current (enhancement of current; Fig. 4B). A loss of electrochemical reversibility of the ligand⁺/ligand redox couple upon addition of O₂ also implies the outer sphere electron transfer to O₂.

Conclusions

In summary, we have shown that the pseudo-dearomatized ligand backbone of PN³P*Rh-CO (**1**) underwent the C–H activation with O₂ to form **2** with a α,β -unsaturated carbonyl moiety at room temperature, showcasing an unprecedented ligand-centered reactivity with the central metal Rh(I) remaining untouched. **2** can function as a Michael acceptor to readily react with thiol and amino groups, such as thiophenol and

4-methylaniline, to generate the ligand post-modified **5** and **6**, respectively, demonstrating its potential for bio-conjugation. Further investigations on their bio-synthetic applications are ongoing and will be reported in due course.

Conflicts of interest

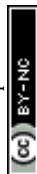
The authors declare that there are no conflicts of interest.

Acknowledgements

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