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Voltammetric responses at modified electrodes and aggregation effects of two anticancer molecules: irinotecan and sunitinib†

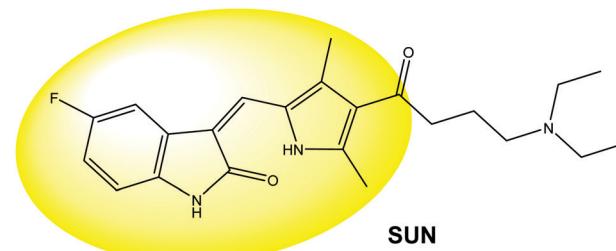
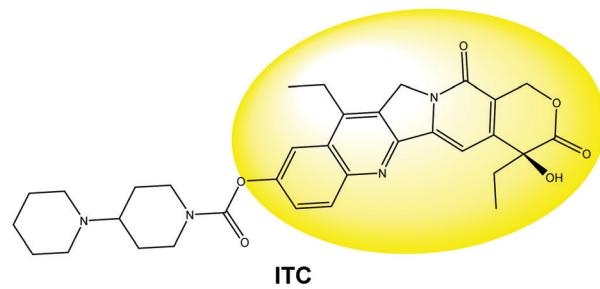
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The voltammetric responses of two anticancer molecules irinotecan (ITC) and sunitinib (SUN) were investigated at glassy carbon (GC) and indium tin oxide (ITO) electrodes bare or modified with carbon nanotubes (NTs), NT–Nafion composites (NafionNTs) and poly(3,4-ethylenedioxy-thiophene) (PEDOT). Bare GC electrodes revealed to be the most suitable for the detection of drugs at low concentration levels. Differential pulse voltammetry (DPV) analyses showed that the responses of the anticancer molecules are affected by the formation of aggregates. In particular, aggregation of SUN results in the formation of H-aggregates. DPV analysis extended to human plasma samples revealed the drugs in micromolar concentrations with good reproducibility.

Introduction

Irinotecan (ITC, Chart 1) is a semi-synthetic camptothecin derivative and is used as a single agent or in combination with other chemotherapeutics for the treatment of colorectal, ovarian and small cell lung cancers.¹ Sunitinib (SUN, Chart 1) is a novel oral multi-targeted tyrosine kinase inhibitor with antitumor and antiangiogenic activities.^{2,3} The exposure of these drugs to patients is characterized by considerable variation due to patient non-compliance (e.g. due to drug-related toxicity), drug interactions with other co-medication, variability in oral drug availability and many other factors. To solve the inter-individual variability in the pharmacokinetics of ITC and SUN, the evaluation of the drugs levels in plasma during infusion is necessary. Thus, the development of analytical techniques that can monitor the drugs in patients' plasma in real time and which are also rapid, simple and cost effective is a challenging issue for the development of personalized chemotherapy. The analytical approaches usually employed are based on liquid chromatography (LC) followed by ultra-violet visible spectroscopy (UV-vis)⁴ or tandem mass spectrometry

(MS/MS).^{5–7} However, the clinical application of LC is prevented by problems related to cost, portability of the instruments, which also requires highly specialized personnel, and long analysis time.⁸ In this scenario, electrochemical methods could be a challenging alternative because of their simplicity and low cost.⁹ Recent research works have reported on modified electrodes for electrochemical sensors and biosensors for bio-medical applications.^{10–15}



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Chart 1 Irinotecan (ITC) and sunitinib (SUN) un-protonated structures. Yellow color highlights the oxidative centers of the molecules.



In a previous work we reported the electrochemical behavior of ITC and SUN at bare glassy carbon (GC) electrodes.¹⁶ In the present work we aim to extend our studies to modified electrodes.

A reported work by Madrakian *et al.*¹⁷ employs carbon paste electrodes modified with a nanocomposite composed of polyacrylonitrile nanofibers and nickel-zinc ferrite nanoparticles for the electrochemical determination of SUN, here we focused our attention on oxides and carbon-based materials; we did not use metals, such as Pt and Au, because in the potential range of analysis their background signals may lead to wrong interpretation of the anticancer molecule responses.¹⁸

In particular, we employed indium tin oxide (ITO), carbon nanotubes (NTs), NT-Nafion composites (NafionNTs) and the conducting polymer poly(3,4-ethylenedioxithiophene) (PEDOT). ITO has recently been used as a thin film electrode material in electrochemical biosensors owing to its good electrical conductivity and high transparency to optical wavelengths, rendering it a good candidate for either electrochemical and/or optical biosensors.¹⁹ On the other side, NTs have been widely employed in electrode modification because they exhibit good conductivity, high chemical stability, high active surface area and resistance to surface fouling.²⁰ Moreover, when they are dispersed in Nafion to form composites, NTs give rise to uniform and stable films which are easy to prepare with good reproducibility and have shown remarkable electrocatalytic properties.^{21–26} Finally, conducting polymers (such as polyanilines, polythiophenes, polypyrroles, *etc.*) were extensively studied as electrode modifiers because they were found to improve the sensitivity and selectivity of the analytical detection and even lower the detection limit.²⁷ In particular, poly(3,4-ethylenedioxithiophene) (PEDOT) has been the object of great attention owing to its high electrochemical and chemical stability, superior conductivity, good compatibility and adhesion ability with other materials.^{28,29}

In this work we employed cyclic voltammetry (CV) to study the electrochemical responses of ITC and SUN at GC and ITO electrodes bare or modified with NTs, NT-Nafion composites and PEDOT. Then we explored the drugs responses at low concentration levels by means of differential pulsed voltammetry (DPV) determinations. In particular, we focused on the influence of the molecular aggregations on their DPV responses. Finally we extended DPV analyses to human plasma samples to detect the anticancer molecules at micromolar levels.

Experimental

Materials

Irinotecan (ITC) hydrochloride, sunitinib (SUN) malate, carboxylate-functionalized multi-walled carbon nanotubes (NTs) (avg. diam. \times length: 9.5 nm \times 1.5 μm) with a carboxylate content of >8 mole percent and NafionTM solutions (5% solution in lower aliphatic alcohols + 10% water, 1200 equiv weight) were purchased from Sigma-Aldrich.

ITC and SUN stock solutions (10^{-3} M) were prepared by dissolving ITC hydrochloride in deionized water and SUN

malate in dimethyl sulfoxide (DMSO) and were stored in a cool and dark place. For SUN the used concentration is limited by the low solubility in water, which is sufficiently high at pH 7 (6×10^{-3} M) but decreases sharply with increased pH down to 1×10^{-4} M at pH 9. For this reason we have used a 1:3 v/v DMSO-water mixture to obtain stable solutions. The strength of all the employed solutions was adjusted to physiological levels by the addition of sodium chloride 0.1 M.

Blank human plasma of healthy subjects was obtained from the Tissue Biobank of the first Surgical Clinic of Padua Hospital (Italy). Human plasma solutions were routinely prepared by 1:10 dilution of human plasma in pH 7 buffer.

Apparatus and procedure

Electrochemistry was performed at room temperature in three electrode cells; the counter electrode was platinum and the reference electrode was SCE. The supporting electrolyte was 0.1 M NaCl. Cyclic voltammetry (CV) was performed at a scan rate 0.02 V s^{-1} . Differential pulse voltammetry (DPV) response was maximized at pulse amplitude $A = 50 \text{ mV}$, pulse duration $T = 50 \text{ ms}$ and scan rate $v = 0.02 \text{ V s}^{-1}$. Width of cyclic voltammograms ΔE_p is calculated as the difference between peak E_p and half-peak $E_p/2$ potentials, whereas for DPV it is the width at half-height.

Working electrodes were glassy carbon (GC, 0.2 cm^2) and indium tin oxide (ITO, 2 cm^2 , $20 \Omega \text{ sq}^{-1}$, purchased from KINTEC, Hong-Kong) bare or modified as follows:

ITO/NT

According to the literature,²⁰ 10 mg of NTs were dispersed in 10 ml of dimethylformamide (DMF) under ultrasonic stirring for 2 h prior to use. By means of a micro-syringe, 10 μL of the dispersed NTs were dropped on the electrode surface (2 cm^2 , *ca.* $5 \mu\text{g cm}^{-2}$) followed by heating the electrode in oven at 50°C for 15 min.

GC/NafionNT

According to the literature,^{22,30} 5 mg of NTs were dispersed in 10 ml of Nafion (wt. 0.1%) solution in methanol (MeOH) under ultrasonic stirring for about 30 min. By means of a micro-syringe, 10 μL of Nafion-dispersed NTs were dropped on the electrode surface (0.2 cm^2 , *ca.* $25 \mu\text{g cm}^{-2}$) and carefully dried in air.

GC/PEDOT

According to the literature,³¹ the monomer 3,4-ethylenedioxithiophene (EDOT), 0.1 M in acetonitrile solvent with 0.1 M tetrabutylammonium triflate ($\text{Bu}_4\text{NCF}_3\text{SO}_3$), as the supporting electrolyte, was potentiostatically oxidized at $0.9 \text{ V vs. Ag/Ag}^+$ reference electrode (silver/0.1 M silver perchlorate in acetonitrile, 0.34 V vs. SCE). The deposition charge was 50 mC cm^{-2} , with the formation of *ca.* 100 nm thick layer on the electrode surface (which corresponds to a polymer amount of *ca.* $30 \mu\text{g cm}^{-2}$ for an electrode area of 0.2 cm^2 and for a reported PEDOT density value of 1.5 g cm^{-3} ^{32,33}).



The voltammetric apparatus was Metrohm Autolab 128N potentiostat/galvanostat.

UV-vis spectra were recorded on a PerkinElmer Lambda 35 spectrometer.

Results and discussion

Cyclic voltammetry

The analyses were performed in pH 7 and pH 9 buffers, because pH has a major impact on the stability of both molecules. In particular, it was found that ITC decomposes in pH < 3 and pH > 11 media.²⁰ SUN has an exocyclic alkanyl group and is capable of showing *Z*-*E* isomerism (Scheme 1) due to the presence of a double bond between 2-oxindol and the pyrrole ring.^{34,35}

The *Z* isomer is the principal clinical form of the drug and it is stabilized by internal hydrogen bonds between the 2-oxindol and the pyrrole ring. The literature data³⁴ show that the *Z*-isomer is not stable in acidic media, because a high concentration of H⁺ ions in the medium can interfere with the formation of internal hydrogen bonds, thus counteracting the thermodynamic stability of the *Z*-isomer and favoring the prevalence of the *E*-isomer.

ITO

In pH 9 buffer both molecules exhibited an ill-defined shoulder at *ca.* 1 V (Fig. 1 and Fig. S1, ESI[†]), with peak current values that are *ca.* 3 and 5 times lower, for ITO and SUN, respectively, than the ones at bare GC electrodes (Tables 1 and 2). The width of the cyclic voltammogram ΔE_p is *ca.* 200 mV for both drugs, a sign of a complex mechanism of electron exchange between ITO and the anticancer molecules. So we modified ITO electrodes with NTs (ITO/NT), in order to optimize the electrochemical response.

ITO/NT

In pH 9 buffer, ITC exhibited an oxidation peak at 0.67 V, *ca.* 100 mV less positive than at bare GC electrodes (Fig. 1), with a peak current value that is *ca.* 2 times lower (Table 1). ΔE_p is 70 mV a sign of an electron transfer process faster than that at

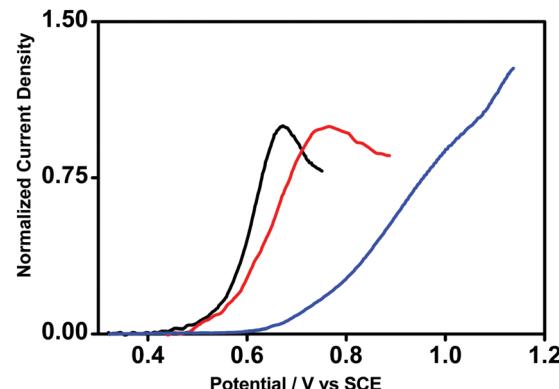
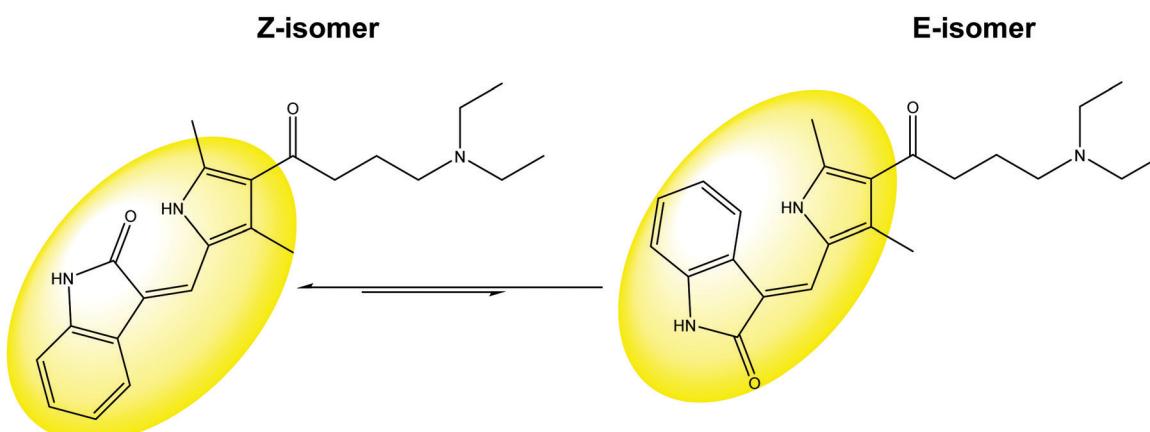


Fig. 1 Normalized single sweep cyclic voltammograms of ITC 10^{-4} M in pH 9 buffer + NaCl 0.1 M at: ITO (blue line), ITO/NT (black line) and GC (red line) electrodes. $v = 0.02$ V s $^{-1}$.

Table 1 Peak current i_p , peak potential E_p (vs. SCE) and peak width ΔE_p for 10^{-4} M ITC in 0.1 M NaCl pH 9 and 7 buffers at a scan rate of 0.02 V s $^{-1}$

Electrode	pH	$i_p/\mu\text{A cm}^{-2}$	E_p/V	$\Delta E_p/\text{mV}$
GC	9	15	0.77	120
GC/NafionNT	9	45	0.77	120
GC/PEDOT	9	70	0.70	60
ITO/NT	9	8.4	0.67	70
ITO	9	5	1.0	>200
GC	7	15	0.89	120
GC/NafionNT	7	45	0.89	120
GC/PEDOT	7	60	0.77	60

bare GC electrodes. SUN oxidation potentials are *ca.* 30 mV less positive than at bare GC electrodes (Fig. S1, ESI[†]) with comparable peak current values (Table 2). ΔE_p of the first oxidation process is 60 mV, like that at bare GC electrodes. In pH 7 buffer, oxidation peaks of SUN shifted to higher potential by *ca.* 120 mV, but the peak current and ΔE_p were the same (Table 2). As expected for an oxidation process involving one-proton loss per exchanged electron the peaks shift to higher potential by 60 mV per pH unit, *i.e.* passing from pH 9 to pH 7



Scheme 1 Structures of SUN *Z* and *E*-isomers. Yellow color highlights the oxidative center of the molecule.



Table 2 Peak current i_p , peak potential E_p and peak width ΔE_p for SUN in 1:3 v/v DMSO–water + 0.1 M NaCl pH 9 and pH 7 buffers (10^{-4} M and 4×10^{-5} M SUN concentration, respectively) at a scan rate of 0.02 V s^{-1}

Electrode	pH	$i_p/\mu\text{A cm}^{-2}$	E_p/V	$\Delta E_p/\text{mV}$
GC	9	15	0.55; 0.65	60
GC/NafionNT	9	55	0.52; 0.64	70
ITO/NT	9	15	0.53; 0.63	60
ITO	9	3	0.80	150
GC	7	15	0.67; 0.78	70
GC/NafionNT	7	55	0.65; 0.75	70
ITO/NT	7	15	0.65; 0.75	70
ITO	7	5	0.90	ca. 200

(see ref. 16 and references therein). However, the NTs deposited on the ITO surface were found to be nonuniform (they formed eye-visible aggregates) and not robust enough to tolerate a standard sticky-tape test (they could be scratched by a simple fingers' touch). So we dispersed NTs in Nafion solutions (GC/NafionNT), in order to obtain more uniform and mechanically stable films.^{22,30}

GC/NafionNT

In pH 7 buffer, ITC exhibited an oxidation peak at 0.89 V with peak current values that are *ca.* 3 times higher than those at bare GC electrodes (Fig. S2, ESI[†] and Table 1). ΔE_p is 120 mV similar to bare GC electrodes. As observed above for SUN, in pH 9 buffer the oxidation peak of ITC shifted cathodically by 120 mV by the 2-unit increase of pH.¹⁶ In both pH media, SUN exhibited two oxidation peaks separated by *ca.* 100 mV (Fig. 2a) with peak current values that, similar to ITC, are *ca.* 3 times higher than those at bare GC electrodes (Table 2). ΔE_p is 70 mV similar to bare GC electrodes (Table 2). The oxidation peaks shift from $E_p = 0.65 \text{ V}$ and 0.75 V at pH 7 to 0.53 V and 0.63 V at pH 9 as observed for ITC.

The effect of pH on the first peak potential of SUN was investigated in the pH range of 5–10. SUN responses (Fig. S3, ESI[†]) showed that, similar to ITC,^{16,20} the peak potential shifts

to less positive values with increasing pH. The peak potential also shows a linear dependence on pH with a negative slope of -50.2 (Fig. S3, inset, ESI[†]):

$$E_p = 1018 - 50.2 \text{ pH} (r = 0.98107)$$

In pH 5 buffer SUN shows one oxidation peak at 0.76 V (Fig. S3, ESI[†]), and not two peaks separated by *ca.* 100 mV as in the pH range of 7–10.

A possible explanation may be found in the change in the structure of its oxidation centre (yellow highlighted in Scheme 1) caused by the *Z*–*E* isomerism of the molecule. In fact, in acidic media the high concentration of H^+ ions facilitates the rotation across the exocyclic group in the SUN structure and favours the switch from the *Z*-isomer, the principle clinical form, to the *E*-isomer (Scheme 1).³⁴

The active surface area of NTs in GC/NafionNT electrodes was estimated to be around $3 \text{ m}^2 \text{ g}^{-1}$ by CV, using $\text{K}_4\text{Fe}(\text{CN})_6$ 10^{-3} M (0.1 M NaCl). The estimated value is two orders of magnitude lower than the one reported for commercial NTs (*ca.* $400 \text{ m}^2 \text{ g}^{-1}$)³⁶ and clearly shows that the active area of NTs is limited by the dispersion in Nafion.

GC/PEDOT

In pH 7 buffer, ITC exhibited a peak at 0.77 V (Fig. 2b) with a peak current value that is *ca.* 4 times higher than that at bare GC electrodes (Table 1). ΔE_p is 60 mV a sign of an electron transfer process faster than the one at bare GC electrodes.

In pH 9 buffer, the oxidation peak of ITC shifted to 0.70 V, but the peak current and ΔE_p were the same (Table 1). Unfortunately, CV responses of SUN are ill-defined in both pH-media. We suppose that DMSO, added to obtain stable SUN solutions,¹⁶ may cause the swelling of the PEDOT film, thus leading to a difficult interpretation of the responses of the molecule.

The above reported results show that both drugs exhibited promising CV responses at ITO/NT and at GC/NafionNT electrodes. In fact, at ITO/NT electrodes the peak potential values of ITC and SUN are 100 mV and 30 mV, respectively, less positive

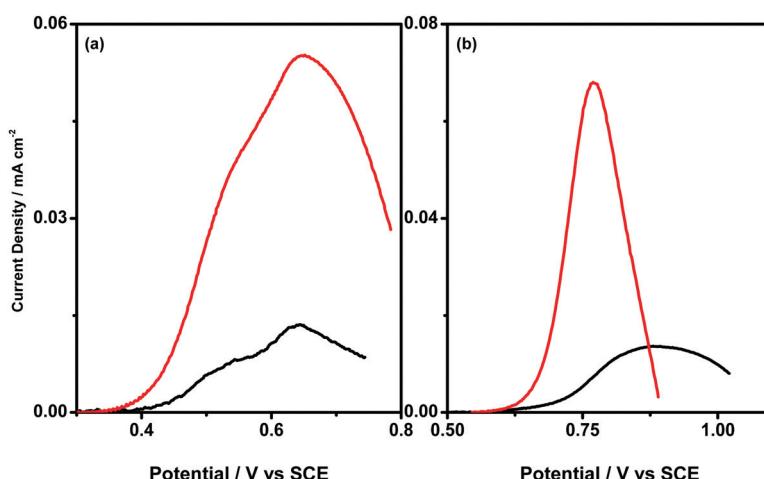


Fig. 2 Single sweep cyclic voltammograms of: (a) SUN 4×10^{-5} M in 1:3 v/v DMSO/water pH 9 buffer + 0.1 NaCl at GC (black line) and CG/NafionNT (red line) electrodes, and (b) ITC 10^{-4} M in pH 9 buffer + NaCl 0.1 M at GC (black line) and GC/PEDOT (red line) electrodes. $v = 0.02 \text{ V s}^{-1}$.



than those at bare GC electrodes and the electron transfer rate of ITC is faster (Tables 1 and 2). However, the NTs deposited on ITO appeared neither uniform nor sufficiently robust. So we dispersed the NTs in Nafion solutions to obtain more uniform and mechanically stable films.^{22,30} In fact, at GC/NafionNT electrodes the peak current values of both drugs are three times higher than the ones at bare GC electrodes (Tables 1 and 2), but the estimated active surface area of NTs is limited by the dispersion in Nafion. Moreover, the signal-to-noise ratio at GC/NafionNT electrodes, estimated from the faradaic and the capacitive currents, is higher than the one at bare GC electrodes. Thus, from these considerations, we selected the bare GC electrodes for the determination of the drugs in low concentration levels, as described in the following.

Differential pulse voltammetry

For ITC, the DPV response (Fig. 3a) exhibits a single oxidation peak. The peak potential shifts to less positive values as the concentration is decreased, moving from 0.667 V to 0.621 and 0.615 V from 10^{-4} M to 10^{-5} and 10^{-6} M (at pH 9, see Table 3).

The peak width is 140 mV at 10^{-4} M concentration but decreases to *ca.* 80 mV in diluted solutions. The theoretical width of the DPV response is 90 mV for a fast electron transfer and *ca.* 150 mV, for a slow transfer.³⁷ Thus it appears that the transfer rate increases in the diluted solutions. The peak current is essentially linear in the 10^{-6} M to 10^{-4} M concentration range (Fig. S4, ESI[†]) and the corresponding regression equation is $y = 2.4 \times 10^5 + 0.8$, $R^2 = 0.993$. The oxidation peak in diluted solutions shifts from $E_p = 0.71$ V at pH 7 to $E_p = 0.61$ V at pH 9 as observed in cyclic voltammetry.

The DPV response of SUN in 1:3 v/v DMSO/water (Fig. 3b) exhibits a single oxidation peak at $E_p = 0.65$ V at pH 7 in 10^{-4} M solution. As summarized in Table 3, in more diluted solutions (10^{-5} M and 10^{-6} M) the response is split into a twin peak signal at $E_p = 0.60$ V and 0.70 V (at pH 7); at pH 9 the twin peaks are shown at $E_p = 0.49$ V and $E_p = 0.59$ V, *i.e.* with the cathodic shift observed in cyclic voltammetry. The peak width

Table 3 Peak current i_p , peak potential E_p and peak width ΔE_p for a typical DPV analysis of ITC and SUN in 0.1 M NaCl on the GC electrode at various concentrations, C

C/M	pH	$i_p/\mu\text{A cm}^{-2}$	E_p/V	$\Delta E_p/\text{mV}$
ITC				
10^{-4}	9	25	0.667	140
10^{-5}	9	4	0.621	80
10^{-6}	9	0.3	0.615	80
SUN				
10^{-4}	7	30	0.646	80
10^{-5}	7	4; 6	0.595; 0.682	80
10^{-6}	7	0.3; 0.6	0.585; 0.685	80

is *ca.* 80 mV at all concentrations indicating fast electron transfer in all cases. Also in the case of SUN the peak current is linear in the 10^{-6} M to 10^{-4} M concentration range (Fig. S4, ESI[†]), with a detection limit of *ca.* 10^{-7} M, the corresponding regression equation is $y = 2.9 \times 10^5 + 0.5$, $R^2 = 0.997$.

DPV analysis of SUN was also performed in 10^{-5} M and 10^{-6} M solutions at pH 7 obtained by dilution of the standard 10^{-4} M SUN solution into DMSO-free buffer. Lowering the DMSO content from 25% to 2.5% makes the single peak become a twin response peak (Fig. 4a). A possible explanation may be found in the formation of molecular aggregates as described in the following.

Aggregation effects

In contrast to organic solvents, aggregates (mainly dimers) are formed by ITC in water.³⁸ It is therefore suggested that aggregation causes the observed potential shifts and width changes in the ITC DPV signals with concentration, as when the oxidation charge-transfer rate is strongly decreased in the aggregated forms. The splitting of the DPV response of SUN with the decrease of concentration could be in principle attributed to analogous aggregation effects. The fact that the splitting is removed by increasing the DMSO content is in line with this suggestion (Fig. 4a). Also, the reported data on the aggregation of conjugated oligomers show the significant splitting of their oxidation responses upon aggregation.³⁹ Aggregation of ITC

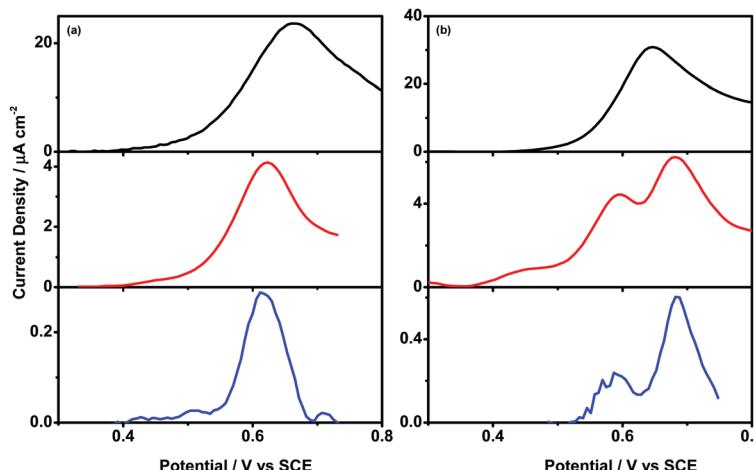


Fig. 3 Differential pulse voltammetry of: (a) ITC 10^{-4} M (black line), 10^{-5} M (red line) and 10^{-6} M (blue line) in pH 9 buffer + 0.1 NaCl; and (b) SUN 10^{-4} M (black line), 10^{-5} M (red line) and 10^{-6} M (blue line) in pH 7 buffer + 0.1 NaCl on the GC electrode.



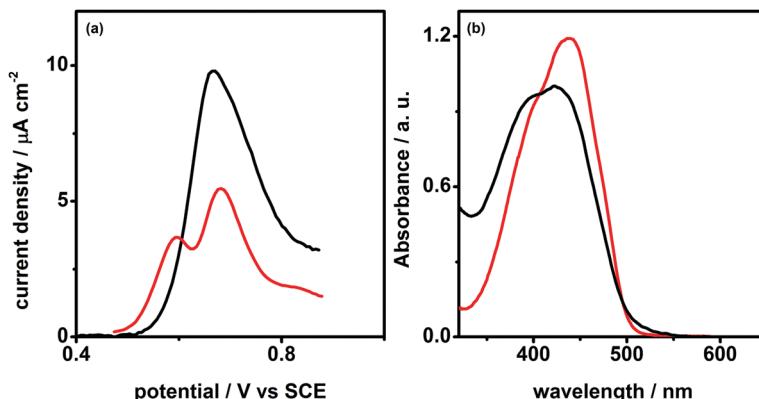


Fig. 4 (a) Differential pulse voltammetry of $\text{SUN } 10^{-5} \text{ M}$ in pH 7 buffer + 0.1 NaCl with 25% DMSO (black line) and 2.5% DMSO (red line) at the CG electrode. (b) UV-vis spectra of $\text{SUN } 10^{-5} \text{ M}$ in water (black line) and in 1:3 v/v DMSO/water mixture (red line).

was previously revealed by the UV-vis spectra of aqueous solution which exhibited an appreciable hypochromic effect when the concentration of the solute is increased from 10^{-5} M to 10^{-3} M , involving the two main absorption maxima at 355 and 368 nm. The absorption at the low-energy peak exhibited a conspicuous decrease of intensity when the concentration of the solute is increased, whereas that at the high energy side did not change.³⁸ We evidenced aggregation, also, for SUN by UV-vis spectroscopy. The spectrum of saturated water solution (*ca.* 10^{-5} M) was compared with the spectrum obtained in 1:3 v/v DMSO/water solution (Fig. 4b). The highest

energy band measured (*ca.* 400 nm) in water becomes a shoulder in DMSO/water solution where the organic solvent disaggregates the dimers. The picture resembles that of ITC with dilution.³⁸ In the case of SUN, the maximum peak is shifted hypsochromically from 440 to 423 nm suggesting that aggregation of the planar conjugated molecules results in the production of H-aggregate forms.⁴⁰

Human plasma

Blood plasma samples were obtained from healthy volunteers who had not received any treatment, including operation,

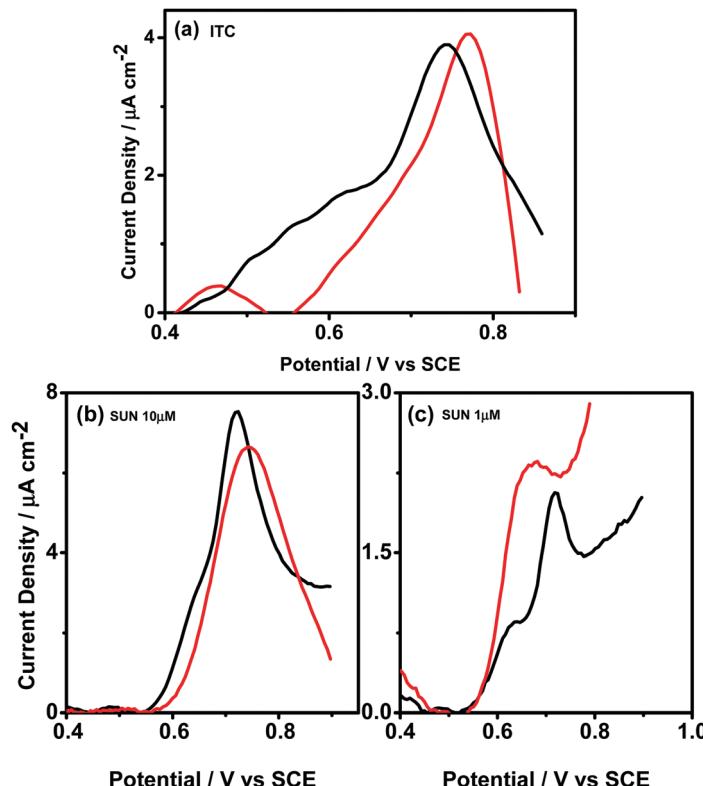


Fig. 5 Differential pulse voltammetry at GC electrodes of: (a) ITC, and (b) and (c) SUN in pH 7 buffer + 0.1 M NaCl before (black line, pure solutions) and after (red line, plasma samples) human plasma addition.

chemotherapy or radiation therapy. Plasma samples were routinely prepared by adding human plasma (1:10 dilution) to pH 7 buffers and then spiked with known amounts of anticancer drugs (5 μ M for ITC, 10 μ M and 5 μ M for SUN). DPV response of ITC in human plasma samples exhibits a single oxidation peak at 0.77 V, *i.e.* 30 mV more positive than that in pure solutions (Fig. 5a). In the case of SUN, the DPV twin peak signal observed in pure solutions (Fig. 5b and c) becomes a single oxidation response in human plasma samples. The peak current values of both drugs are comparable with those observed in pure solutions, for a given drug amount, showing that there are no significant interferences from plasma excipients.³⁴ The relative standard deviations of the peak current in plasma samples are 4% and 6% ($n = 3$) for ITC and SUN, respectively, demonstrating the good reproducibility of the measurements.

Conclusions

We have investigated the responses of the anticancer molecules irinotecan and sunitinib at modified electrodes in pH 7 and pH 9 media by CV determinations. We observed that the bare GC electrodes revealed to be more suitable for the detection of the drugs in low concentration levels. In fact at GC/NafionNT electrodes the peak current values of both drugs are three times higher than the ones at bare GC electrodes, but the estimated active surface area of NTs is limited by the dispersion in Nafion. On the other side, at ITO/NT electrodes the peak potential values of ITC and SUN are, respectively, 100 mV and 30 mV less positive than those at bare GC electrodes and the electron transfer rate of ITC is faster, but, for the employed electrode areas, the deposited NTs appeared neither uniform nor significantly robust. A possible alternative, which will be the object of further studies, is to deposit NT films on smaller-dimension electrodes, *i.e.* micro-electrodes.

DPV analyses showed that the drug responses are affected by molecular aggregate formation. In particular, the oxidation charge-transfer rate of ITC is strongly decreased in the aggregated forms, while the DPV response of SUN splits into a twin peak signal upon aggregate formation. Moreover, the aggregation of SUN results in the production of H-aggregate forms. DPV measurements were successfully extended to human plasma samples and revealed both drugs in micromolar concentration levels, with good reproducibility. To the best of our knowledge this is the first study on the voltammetric responses of two anticancer molecules at bare glassy carbon electrodes, also extended to human plasma samples.

Compliance with ethical standards

The study was conducted according to the principles expressed in the Declaration of Helsinki. Blank human plasma of healthy subjects was obtained from the Tissue Biobank of the first Surgical Clinic of Padua Hospital (Ethical committee of Centro Oncologico Regionale/I.O.V., Azienda Ospedaliera di Padova, approved protocol number: P.448).

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

With this work we want to remember Dr Gianni Zotti, who suddenly passed away on 15th of April 2020. Dr Zotti was an esteemed electrochemist of CNR, Padova. In 1984, he produced the first conjugated polymer films by the cathodic route, which opened the way to the electrochemistry of conducting polymers. His studies concerned the development of conjugated polymers and molecular materials with specific electronic properties for application in electronics, photonics, energy conversion, and sensors. Apart from his achievements in science, Gianni was an extraordinary human being with an amazing intellect, sense of humor, enthusiasm and a warm, generous personality. We will miss him deeply. We also would like to thank Dr Anna Berlin of CNR, Milan, for the helpful discussions. We are indebted to the Italian Government (MUR) for financial support.

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