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Electrochemical performance of a self-assembled two-dimensional heterostructure of rGO/MoS₂/h-BN†

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We report the preparation and electrochemical performance evaluation of a two-dimensional (2D) self-assembled heterostructure of graphene oxide (rGO), molybdenum disulphide (MoS₂), and hexagonal boron nitride (h-BN). In the present study, the rGO–MoS₂–h-BN (GMH) multi-layered GMH heterostructure is fabricated *via an in situ* chemical route. Based on material analysis, the composite consists of bond conformations of C–B–C, Mo–S, C–N, B–N, and Mo–C, indicating the layered stacks of rGO/h-BN/MoS₂. In electrochemical analysis, the composite showed superior performance in the aqueous medium of cobalt sulphate (CoSO₄) over other samples. CV measurements, performed over the range 10 to 100 mV s⁻¹, showed a change in specific capacitance (C_{sp}) from 800 to 100 F g⁻¹. GMH showed almost no degradation up to 20 000 cycles @ 100 mV s⁻¹. The calculated C_{sp} , energy density (E_D), and power density (P_D) are discussed in light of Nyquist, Bode, and Ragone analysis. An equivalent circuit is simulated for the cell and its discrete electronic components are discussed. Due to its larger effective electron diffusion length > 1000 μm, broadly, the composite showed *battery-like* characteristics, as supported by radical paramagnetic resonance and transport response studies. The symmetric electrodes prepared in one step are facile to fabricate, easy to integrate and involve no pre or post-treatment. They possess superior *flat cell* character, are cost effective, and are favourable towards practicality at an industrial scale, as demonstrated on the laboratory bench. The details are presented.

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1. Introduction

Today's industrial market requires compact instruments with high energy storage devices, and the demand is ever-increasing for these power backup systems. The widespread usage of lithium-ion batteries is mainly due to advances in this technology. Rechargeable lithium-ion batteries have already acquired a *milestone* position as dominant energy sources; they power almost all forms of consumer electronics, including electric vehicles, strategic systems, and digital home appliances. This is due to their capability to provide high energy density (E_D). In spite of their extensive use, lithium-ion batteries possess several drawbacks, such as poor cycle life, low power performance, fire risk, carcinogenic waste, aging effects, and high cost.¹ In recent years, supercapacitors have been considered as alternatives to lithium-ion batteries due to their safety aspects, their capability to deliver ~10 times more power

density (P_D) than batteries, and their stability over millions of charge/discharge cycles.^{2,3} Further, current batteries are not sufficient to enable cars to drive distances comparable to those of combustion energy-driven vehicles. In order to commercially realise the technology of supercapacitors, enormous amounts of research endeavours are being performed to improve their technical specifications, such as E_D , P_D , and life cycle. A considerable amount of work has been carried out to fabricate electrode materials, select appropriate electrolyte environments, *etc.* Several two-dimensional (2D) materials, such as graphene, molybdenum disulphide, hexagonal boron nitride, transition metal dichalcogenides (TMDCs), g-C₃N₄, phosphorene layered double hydroxides and MXenes,^{4–8} are desirable for energy storage applications due to their outstanding structural-property relationships. In addition to graphene, transition metal dichalcogenides (TMD), post-graphene contenders, and transition metal oxides (TMO)/hydroxides (TMH) are considered as potential candidates. The structures of TMO and TMH in the form of films, flakes, platelets, petals, belts, *etc.*, significantly alter the inherent properties of supercapacitors, showing excellent E_D . This is a huge class, among which we have specifically chosen reduced graphene oxide (rGO), MoS₂ and h-BN in the form of a blend to investigate their electrochemical performance. Basically, the use of carbon compounds in energy storage devices has its own advantages in the form of superior

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performance, high surface area, low cost, affordability of large scale applications and facile preparation protocols. There are a number of reports on the use of nanocarbons such as rGO, single-walled carbon nanotubes (CNTs), CNTs/polymers, rGO-TMO, activated carbon, and multi-walled CNTs to achieve superior specific capacitance (C_{sp} , in $F g^{-1}$). Particularly, rGO showed $C_{sp} \sim 200$ @ 1200 cycles in the electrolyte KOH,^{9,10} >300 @ 1000 cycles in PVA,¹¹ and ~ 600 @ 1000 in H_2SO_4 .¹²

Similarly, the MoS_2 layers alone show the ability to store electrostatic charge by an electrode double layer capacitance (ELDC) mechanism. However, they suffer from inferior electrochemical performance in terms of poor cyclic life, inherently low charge transport, large volume changes during cycling, and restacking,^{13,14} yielding $C_{sp} \sim 100 F g^{-1}$ @ $1 mV s^{-1}$ (scan rate).¹⁵ In order to overcome this issue, MoS_2 has been mixed, wrapped or deposited with highly conductive/electroactive materials such as carbonaceous materials or conducting polymers using various top-down/bottom-up synthetic methods as well as combinatorial approaches.^{16–25}

Pristine h-BN, being an electrical insulator, is not suitable for this branch of application. However, due to its structural similarity with graphene and comparatively weak van der Waals forces, it offers strong ionic bonds and is expected to provide superior C_{sp} if combined with graphene and MoS_2 .²⁶

There are many reports of binary composites. Further, binary composites of rGO are reported to show excellent electrochemical properties. In the work carried out by Jiao *et al.*, a MoS_2 /rGO composite prepared *via* hydrothermal delivered excellent electrochemical properties with almost 95% capacitance retention.²⁷ In other work reported by Zhao *et al.*, MoS_2 /rGO showed superior electrochemical properties with a specific capacitance of more than $1200 F g^{-1}$.²⁸ A layer-controlled MoS_2 /graphene aerogel composite was also reported to show reversible capacity, with a specific capacitance of more than $1000 mA g^{-1}$ at a current density of $100 mA g^{-1}$.²⁹

The focus of our current work is to investigate the electrochemical parameters of a rGO/h-BN/ MoS_2 heterostructure. This self-assembled composite is prepared by facile addition of equal volume ratios. The blend showed successful formation of the composite 2D material. The electrochemical properties were studied to realize the effects of the blend of all the 2D materials; it showed a $C_{sp} > 650 F g^{-1}$ (current density) with a highest $E_D > 90+ Wh kg^{-1}$ and a P_D of $1700 W kg^{-1}$. The combination studied herein has not been reported to date for electrochemical applications. The details are presented.

2. Experimental section

2.1 Reagents

Natural graphite flakes were purchased from Sigma Aldrich. All other reagents are analytical grade chemicals.

2.2 Preparation of the rGO/ MoS_2 /h-BN multi-layered material: GMH-composite

Initially, graphene oxide (GO) was prepared from a natural graphite precursor by a modified Hummers route.³⁰ GO was

washed with DI water several times in order to obtain GO free from acidic moieties. Hexagonal boron nitride (h-BN) was prepared according to the protocols given in ref. 31. For this, around 3 g of boric acid (HBO_3) and approximately 9 g of urea were mixed in acetone, and the solution was stirred for approximately 15 m. Following this, the acetone was allowed to evaporate under natural conditions, and the obtained powder was heated for ~ 700 °C for 5 h under nitrogen atmosphere. In a round bottom flask, subsequently, GO (500 mg), h-BN (500 mg), ammonium molybdate ($(NH_4)_6Mo_7O_{24} \cdot 4H_2O$, (1.25 g)) and thiourea (CH_4-NH_2S , 2.31 g) were added to DI water.³² To the reaction mixture, hydrazine hydrate (10 ml) was added, and the entire system was maintained with heating at ~ 80 °C in an oil bath for about 4 h. After this, the system was allowed to cool to room temperature. During the synthesis of h-BN, the small amount of B_2O_3 formed may be reduced in the presence of hydrazine hydrate. Moreover, we prepared MoS_2 *in situ* during the reduction process of GO. The water was separated using vacuum filtration and the *black* residue was washed using hot DI water and ethanol, sequentially. The residue was sonicated further for about 4 h to exfoliate the interlayers. Finally, the product was assumed to be multi-layered rGO/ MoS_2 /h-BN and termed GMH-composite. It was further characterized for its structure–property relationships and by electrochemical application studies.

2.3 Characterization of GMH-composite

Electron spectroscopy for chemical analysis (ESCA) measurements were performed using an Omicron ESCA probe (Omicron Nanotechnology). X-ray powder diffraction studies were carried out using a Rigaku instrument with $Cu K_\alpha$ radiation (1.5406 \AA) over the 2θ range of $10\text{--}90^\circ$ @ a scanning rate of $2^\circ m^{-1}$. FTIR measurements were performed from 400 to $4000 cm^{-1}$ using a Bruker Tensor 37, and Raman measurements were performed over 200 to $3500 cm^{-1}$ using a LABRAM HR 800 (@ $\lambda \sim 533 nm$). Nanostructure imaging was carried out by field emission scanning microscopy (FSEM, S-4700), high resolution transmission electron microscopy (HRTEM, JEOL-2100F) and energy dispersive X-ray (EDAX) elemental mapping. ESR measurements were performed at 9.4 GHz (X-band) with a microwave input of $950 \mu W$.

2.4 Electrochemical measurements

The electrochemical performance of GMH was determined using cyclic voltammetry (CV) and galvanostatic charge-discharge (CD) studies. In order to understand the various impedance factors experienced by the electrode/electrolyte interface and the charge transfer process, electrochemical impedance spectroscopy measurements (EIS, Ivium Technology, Vertex) @ 0.1 to $10^6 Hz$ were performed. All studies were carried out using an Autolab workstation (PGSTAT 30, Eco Chemie) at room temperature. Electrode preparation was performed as per the protocols given in ref. 33. The composite of GMH and polyvinylidene fluoride (PVDF) was mixed with the solvent *N*-methyl-2-pyrrolidone (NMP) in a ratio of 9 : 1, and the slurry was dropped onto glassy carbon electrode using the drop



cast technique. The electrode was dried to perform CV, CD and EIS measurements. In the CV measurements, the three electrodes used were platinum wire as a counter electrode, Ag/AgCl as a reference electrode, and CNS as a working electrode. The CV curves were obtained in a potential window of 0.2 to 0.8 V @ scan rates of 10, 50, 100, 300, 500 and 1000 mV s⁻¹. In a similar fashion, two-electrode CV measurements were performed in addition to constant CD measurements @ 1.0 to 3.5 A g⁻¹. The electrolytes used were H₂SO₄ (1 M), Na₂SO₄ (1 M), and Co₂SO₄ (1 M). The cyclic stability of the cell was investigated up to 20 000 cycles @ 1 A g⁻¹.

3. Results and discussion

3.1 Structure–property relationship: pre-analysis of GMH-composite

Fig. 1(a) shows the recorded FTIR spectra of MoS₂, rGO, h-BN, and GMH-composite. For h-BN, a typically broad B–N peak appeared at ~1400 cm⁻¹, displaying in-plane B–N–B stretching (str.), whereas the peak at ~750 shows out-of-plane B–N–B stretching. For rGO, the peaks that appeared at 1300 to 1500 cm⁻¹ are assigned to C–C str. and C–H bending, whereas the peak at ~665 cm⁻¹ shows C–H out-of-plane vibrations. In MoS₂, the peak at ~615 is due to the presence of Mo–S str. In addition, the profile shows small peaks at ~950 and 850 cm⁻¹ due to Mo–O str., as shown in the inset to the corresponding zoomed region in plot (c). These peaks also shifted to lower wavelengths, supporting the formation of the composite. Their presence may be due to atmospheric handling of the material. However, the IR profile of GMH-composite shows the features of individual components, such as the peak at ~1600 due to C=C str. In the composites, the peaks associated with B–N, C–C, and C–H appear to overlap due to their similar vibration frequencies; hence, the bands appearing at 1300 to 1500 cm⁻¹ are attributed to their presence. A sharp peak observed at 900 cm⁻¹ is assigned to Mo–O, at 700 cm⁻¹ to B–N–B, and at ~630 cm⁻¹ to Mo–S stretching vibrations. Primarily, the prepared composite shows the vibration features of individual MoS₂, MoO₃, h-BN, and reduced graphene oxide 2D layers. Further, the plot (b) is the recorded XRD pattern of the composite obtained on calcination at 350 °C; it shows distinct diffraction peaks associated

with rGO, h-BN and MoS₂ in the blend. The individual XRD patterns are provided in the ESI.† For MoS₂, namely, peaks 1–5 match with JCPDS card no. 37-1492. The values for 2θ are ~12° (1), 33° (2), 33° (3), 39° (4), and 42° (5), with corresponding Miller planes of (0 0 2), (1 0 0), (1 0 3), (1 0 5) and (1 1 0). The intensity of these lines is observed to be low due to exfoliation and intercalation within the layers of rGO and h-BN and is indicative of lowering of the number of layers.³⁴ Moreover, h-BN and rGO have very similar crystal structures, with a very small lattice constant mismatch of ~2%. Therefore, the main peaks have similar positions and intensities. For rGO, they are indicated by number 6. The broad peak shifted to sharp crystalline at 2θ = 23.34° is indicative of rGO, corresponding to a *d* spacing of 0.326 nm. During the formation of the heterostructure, the *d* spacing of rGO in the blend is reduced by almost half. The lattice structure is significantly different from that of the pristine graphite flakes; this shows substantial exfoliation of the conjugated 2D carbon layers. For h-BN, the XRD pattern matches JCPDS card number 73-2095, as shown in Fig. 7–10. The corresponding matching planes are (1 0 0), (1 0 1), (0 0 4), and (1 1 0), as shown by the numbers.³⁵ To summarize, the composite of GMH shows the presence of the multi-layered heterostructure h-BN/MoS₂/rGO, as desired for electrochemical applications. Plot (c) shows the Raman spectrum recorded at an excitation wavelength of 533 nm for the composite. The inset shows the magnified region of 240–450 cm⁻¹, in which 4 peaks were clearly identified for MoS₂. The D band at 1340 cm⁻¹ has a higher intensity than the G band at 1590 cm⁻¹, which indicates the presence of impurities induced due to MoS₂ and h-BN. In addition, 2D doublets were registered in the blend at ~2700 cm⁻¹. h-BN shows a characteristic peak at 1370 cm⁻¹, which is due to the E_{2g} phonon mode and is similar to the mode associated with the D band in graphene.³⁶ In the composite samples, their presence is not observed separately due to mutual overlap; however, this resulted in a mixed band at a somewhat lower wavelength of 1340 cm⁻¹. The position of the D band also shifted to a lower wavelength.³⁷ Characteristic bands of MoS₂ appeared at 280, 335, 370, and 418 cm⁻¹. The bands at 370 and 418 cm⁻¹ are attributed, respectively, to E_{2g}¹ and A_g¹ modes, while the major characteristic peaks at 280 and 335 cm⁻¹ belong to E_{1g} and longitudinal acoustic phonon

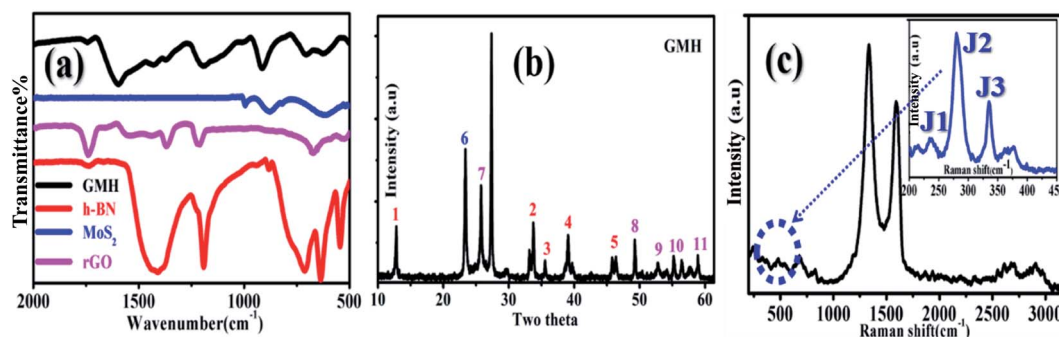


Fig. 1 (a) FTIR spectra of the individual components and GMH-composite, (b) XRD pattern of the composite (the coloured numbers show the characteristic peaks of the respective components), and (c) Raman spectra recorded at 533 nm for the composite (the zoomed portion of the dotted circle connects the typical J1, J2 and J3 peaks; the inset shows the characteristic Raman features for 1T-MoS₂).



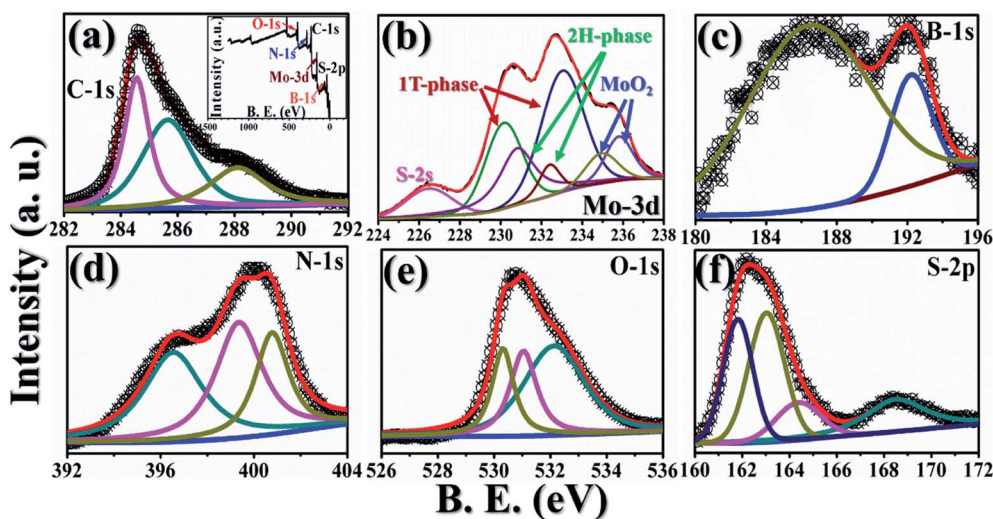


Fig. 2 Recorded XPS spectra for GMH-composite: (a) C-1s (inset: elemental survey scan), (b) Mo-3d (indicating 1T, 2H and oxide states), (c) B-1s, (d) N-1s, (e) O-1s, and (f) S-2p.

modes, respectively. These three peaks are associated with the 1T phase of MoS₂. This indicates that during composite formation, 1T phase is formed; this may improve the structural stability of MoS₂.^{38,39}

In order to understand the surface chemical composition and the valence state of the GMH blend, XPS analysis was carried out. Fig. 2 shows the corresponding spectra. The inset of plot (a) shows a survey spectrum indicating the coexistence of elements such as B, S, Mo, C, N, and O in GMH. Further, the composition of MoS₂ in GMH was confirmed by the high resolution of the MoS₂ spectrum, shown in plot (b). The Mo-3d peak can be deconvoluted to four peak fits at binding energies of 226.5, 230.4, 232.8, and 235.2 eV. The peak at 230.4 eV is assigned to Mo (+4) 3d_{5/2} and that at 232.8 eV is assigned to 3d_{3/2}, with a spin-orbit coupling separation of nearly 2.4 eV. The shift of these additional peaks suggests that they arise from 1T phase. This can be identified to be metallic 1T phase partially mixed with oxidized MoS₂, which is consistent with the Raman analysis.³⁹ The peak at 235.05 is assigned to the respective Mo (+6) 3d electronic shell. The emergence of this peak may be due to the exposure of MoS₂ to air, showing the presence of MoO₃, or may be due to incomplete reduction of Mo (+6).²⁷ In addition, the peak at a binding energy of 226.4 eV corresponds to S 2s of MoS₂ and also shows a resemblance to the spectral region of S-2p (plot (f)). For sulphur, the peaks at 161.8 and 163.04 are assigned to Mo-S bonding, with spin-orbit coupling of ~1.24 eV. Moreover, in (f), the spectrum shows that the binding energy of S-2p_{3/2} is 161.9, which corresponds to a charge state of S of -2 in MoS₂. Further, the peak at 164.3 corresponds to C-S bonding, whereas the peak at 168.5 corresponds to carbon-sulfonate bonding. Meanwhile, in plot (e), O-1s is deconvoluted to three components at 532.2, 531.04, and 530.3 eV; these broadly correspond to C-O and M-O bonding. In plot (a), the C-1s envelope is fitted to three peaks. Particularly, the peak at 284.5 eV appears to be broad and has a asymmetric tail towards the higher binding energy side, which shows a high

concentration of sp² content in rGO. Meanwhile, other peaks at 285.6 and 288.15 eV with a difference of 1–2 eV show the presence of an sp³ fraction emerging from the C-O and C=O components, respectively. In plot (d), the N-1s spectrum is fitted for three components, in which the main broad peak at 396.5 is attributed to N-B bonding in h-BN; meanwhile, the other peak centred at ~400 is attributed to the presence of N-H bonds or amide groups,⁴⁰ and the peak at ~401 shows C-N bonds. In plot (c), the deconvoluted B-1s spectrum indicates the existence of sp²-like phase in B-N due to peaks appearing at ~193 and ~187 eV. These show the substitution of boron sites that exist in the rGO network of the GMH composite.^{41,42}

In the analysis of GMH-composite, broadly, the excess oxygen containing groups are significantly reduced, and only a small amount of oxygen is observed to reside in the 2D multilayers. From the rGO, h-BN and MoS₂ sheets, unwanted traces of hydroxide/epoxide, B₂O₃, and MoO₃, respectively, are removed due to the action of hydrazine hydrate as a reducing agent. The blend is found to be rich in C-B-C, Mo-S, C-N, B-N, Mo-C, *etc.*; this indicates that molecular bonding enabled the formation of composite layers.

3.2 Morphological analysis of GMH-composite

Electron microscopy imaging analysis was carried out in order to investigate the morphology of GMH-composite. Fig. 3 shows typical FESEM images of rGO, MoS₂, h-BN and GMH, and the EDX spectrum of GMH-composite. The FESEM image of rGO in Fig. 3(a) shows a smooth, homogeneous surface with typical wrinkles of rGO sheets that cause sheet folding. The corrugation and scrolling represent the intrinsic nature of graphene because the 2D membrane structure becomes thermodynamically stable *via* blending with the corrugation and scrolling.^{34,36} Fig. 3(b) shows coalesced plate-like MoS₂ particles. It can be seen that these microspheres are composed of stacked nanosheets assembled into uniform nanospheres. These nanosphere-like



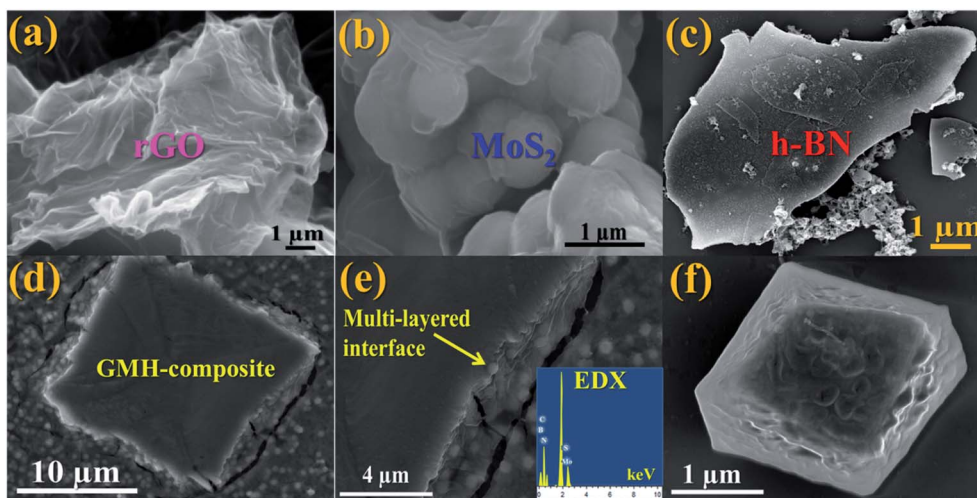


Fig. 3 Typical FESEM images. The upper panel shows individual 2D layers of (a) rGO, (b) MoS₂, and (c) h-BN. Lower panel: (d) GMH-composite, (e) sandwiched layers (inset: EDAX spectrum revealing the elemental composition), (f) a smaller block of the blend.

sheets are loosely intercrossed and contain void spaces between them. Part (c) shows an FESEM image of h-BN, indicating sharp edges and a flat, smooth surface. However, the FESEM images of GMH show an entirely different morphology. It is shown in part (d) that layered lump-like structures can be seen. Compact packing can be observed, and the layers of rGO appear to be broken down into small layers. For the GMH micrographs of cross section part (e) and part (f), it can be seen that the 2D planes are tightly stacked and the layered structure is well preserved, with h-BN and MoS₂ planes sandwiched in between the layers. The tightly packed layered structure plays a key role

in the efficient charge transfer as well as the storing between three nanosheets.

Fig. 4(a) shows a typical TEM image of rGO, indicating well-defined stacked graphitic planes, and the respective inset displays higher resolution imaging with lattice fringes. This indicates notable crystallization of the sheets. Image (b) shows two layers of MoS₂; it appears that due to stacking faults, the crystallinity is somewhat lost. In (c), crystalline h-BN can be seen, with some structural defects indicated by yellow double circles. These defects can be attributed to the presence of B₂O₃. They can be passivated by adding more hydrazine hydrate, as discussed previously. The corresponding inset shows the SAED

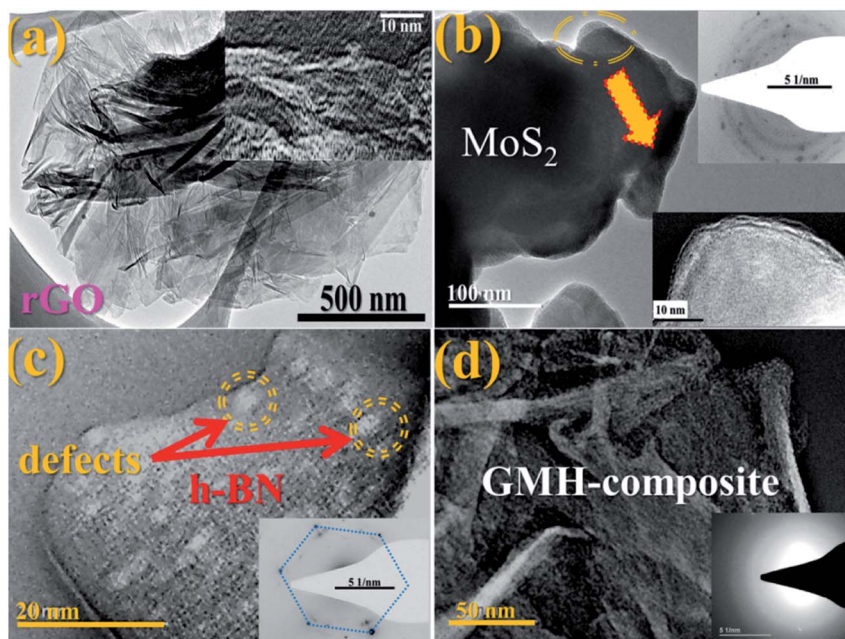


Fig. 4 Typical HRTEM images of (a) rGO (inset: higher resolution of layers), (b) MoS₂ (inset: higher resolution and SAED pattern), (c) h-BN arrows and double circles showing defects (inset corresponding SAED), and (d) GMH-composite with SAED.



pattern, which confirms the crystallinity of h-BN. Image (d) shows stacked compact layers of rGO, MoS₂ and h-BN, and the SAED pattern shows the amorphous nature of the blend. The supercapacitor performance of GMH-composite was investigated by cyclic voltammetry (CV) studies.

3.4 Electrochemical studies

In Fig. 5, plot (a) shows three electrode CV curves obtained for GMH-composite in different aqueous electrolytes at a scan rate of 10 mV s⁻¹. The values of the specific capacitance, C_{sp} , measured in F g⁻¹, were recorded to be 302.67, 564.18, and 745.18 for electrolytes H₂SO₄, Na₂SO₄, and CoSO₄, respectively.

These were calculated using the relation $C_{sp} = \frac{\int iVdV}{2 \mu m \times \Delta V}$. The magnitude of C_{sp} in H₂SO₄ electrolyte is low, and the shape of the CV curve is somewhat distorted. This may be due to the corrosive nature of acid towards the metallic species present in the composite. Acidic and basic electrolytes can act as current collectors and provide connexions. In order to overcome these technical drawbacks, they should be avoided.³⁹ Neutral electrolytes such as Na₂SO₄ and CoSO₄ showed symmetric current potential characteristics, indicating ideal capacitive behaviour. In general, the process of oxidation and reduction during cycling depends on the physical parameters of the cations, such as charge, charge density and size.⁴⁰ The highest value of C_{sp} was obtained for CoSO₄, which was used for further analysis. This is attributed to the smaller cationic radius of Co²⁺ ion (0.72 Å) than of Na⁺ (0.99 Å). Plot (b) presents profiles (three electrodes) at various scan rates from 10 to 100 mV s⁻¹ in CoSO₄ electrolyte. The CV curves exhibited nearly rectangular shapes, which is indicative of an ideal electrode double layer capacitor. From the device perspective, it is important to investigate two-electrode CV measurements. Therefore, we proceeded to do so; in (c), the variations in C_{sp} with changing scan rate can be

Table 1 Calculated C_{sp} , E_D and P_D values for GMH-composite for the two-electrode system

Current density (A g ⁻¹)	1.0	1.5	2.0	2.5	3.0	3.5
C_{sp} (F g ⁻¹)	683	482	401	296	224	214
E_D (W h kg ⁻¹)	92.3	65.1	54.2	40.0	30.0	28.9
P_D (W kg ⁻¹)	486	729	972	1213	1458	1701

observed to study the effective interaction between the ions and the electrode. Further, plot (IV) shows the galvanostatic charge-discharge curves at various current densities from 1 to 3.5 A g⁻¹. The increased scan rate and current density show a decrease in specific capacitance.

Table 1 shows the specific capacitance, C_{sp} , for the two-electrode technique with the corresponding power, P_D , and energy density, E_D . The values of C_{sp} for two electrodes are usually lower compared to those of the three-electrode configuration and were calculated using the formulae:

$$C_{sp} = 4 \times \frac{I\Delta t}{m\Delta V} \quad (1)$$

$$E_D = \frac{1}{2} \times \frac{1}{4} \times \frac{1}{1.36} \times C_{sp} \times \Delta V^2 \quad (2)$$

$$P_D = \frac{E}{\Delta t} \quad (3)$$

where I is the current, m is the specific mass, Δt is the time, and ΔV is the potential window.

At lower current density (lower scan rate), both the inner/outer active sites and pores of GMH can be sufficiently accessed by the electrolytic ions, leading to high C_{sp} values. For a high current density or scan rate, only the external surface of the electrode material contributes to the charge/discharge process. It appears that nanosheets significantly shorten the

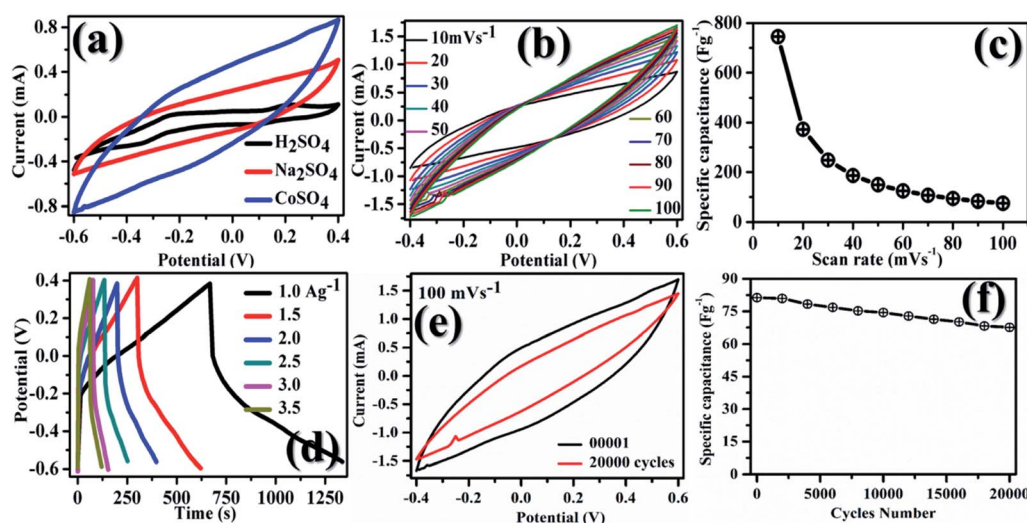


Fig. 5 CV measurements on GMH-composite for (a) different aqueous electrolytes (scan rate: 10 mV s⁻¹) with superior performance in CoSO₄, (b) variable scan rates between 10 and 100 mV s⁻¹, (c) change in C_{sp} (specific capacitance) with scan rate, (d) stability curves up to 20 000 cycles, (e) CV curves recorded after the first and the last cycle, (f) charge-discharge (CD) curves at variable current density (A g⁻¹). Plots (b–f) are CV measurements in CoSO₄ electrolyte.



Table 2 Performance characteristics of reported supercapacitors compared to the current work

Electrode material	Electrolyte	C_{sp} (F g ⁻¹)	Cyclic stability	E_D (W h kg ⁻¹)	P_D (kW g ⁻¹)	Ref.
MoS ₂ -rGO	1 M Na ₂ SO ₄	388 @ 1.2 A g ⁻¹	100% up to 1000 cycles	—	—	44
h-BN/rGO	6 M KOH	145 @ 6 A g ⁻¹	89% up to 40 000 cycles	40	4200	26
Graphene/MnO ₂	Ethanol	315	95% up to 5000 cycles	13	110	45
<i>Aloe vera</i> -derived carbon	[EMIM][BF ₄]	244	83% up to 500 cycles	40	—	46
MoS ₂ /rGO	2 M KOH	220 @ 1 A g ⁻¹	91.8% up to 1000 cycles	—	—	47
Te/rGO	1 M H ₂ SO ₄	260 F g ⁻¹ @ 0.25 A g ⁻¹	~80% up to 10 000 cycles	8	250	48
GMH-composite	CoSO ₄	215 @ 3.5 A g ⁻¹	75% up to 20 000 cycles	30	1700	Current work

ion diffusion pathways during the CD process and help obtain the high specific capacitance.³⁶ Fig. 5(e and f), respectively, show the I - V curves after 1 and 20 000 cycles. After 20 000 cycles, a slight decrease in the specific capacitance of ~25% is observed by a reduction in the area under the curve. The good cyclic stability can be attributed to the excellent chemical stability, mechanical strength and flexibility of GMH-composite.⁴³

Table 2 provides a comparison of the performance of other previously reported supercapacitors with the current work. Typically, carbonaceous materials such as graphene, carbon nanotubes, and carbon aerogel have high electrical conductivities and surface areas. Developing TMD hybrids or composite materials with carbon-based materials provides a synergistic effect of the two materials; carbon offers conductive channels and enhanced interfacial contact, while TMDs contribute a short ion diffusion path and successively short electron transport path. GMH-composite shows comparatively good retention stability of 75% up to 20 000 cycles.

After cycling, GMH-composite was analysed with XPS, and it was observed that the chemical environment of the heterostructures is the same as that of pristine samples, as shown in Fig. 6.

It appears that interlayer electrostatic *cross-talk* plays an important role when one sheds light on the charging–discharging mechanism of this multi-layered blend composed of rGO, MoS₂ and h-BN. The inter-layer stacks of MoS₂ and h-BN in conjugation with the rGO layers may be mutually acting as spacer molecular layers that prevent restacking. Further, free radical electrons may be responsible for changing the oxidation state of intercalated Mo, B, N and S atoms, which contributes to the additional charge storage capacity of the blend. On the other hand, the intercalated MoS₂ provides additional channels for charge transport and increases the overall conductivity of the electrodes.

The re-hybridization of MoS₂ may also be possible with h-BN. This may allow the intercalated ions to move swiftly, providing active charging sites. The dominant electrical double layer capacitor (EDLC) behaviour of the composite GMH electrodes is a synergistic effect of the kinetics for non-faradaic proton adsorption at the interface of MoS₂/h-BN/rGO. Fig. 7 shows (a) the Nyquist plot and corresponding (b) Bode plot recorded for the GMH blend. Impedance analysis was employed to study the frequency characteristics of the electrode. The EIS was carried out in 2 M CoSO₄ solution over the frequency range of 0.01–10 000 Hz.

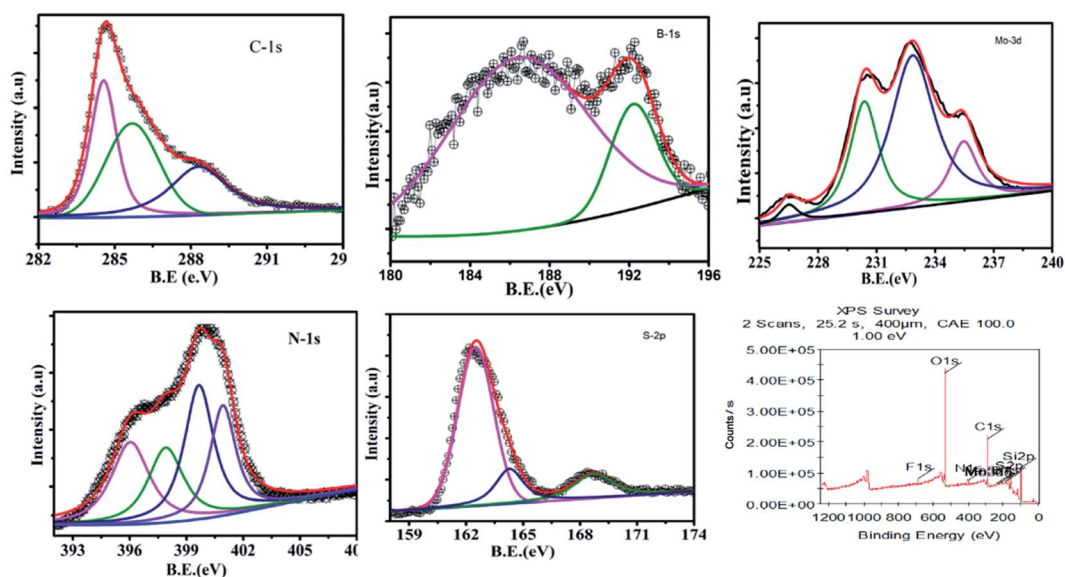


Fig. 6 Recorded XPS spectra for the GMH heterostructure after 20 000 cycles.



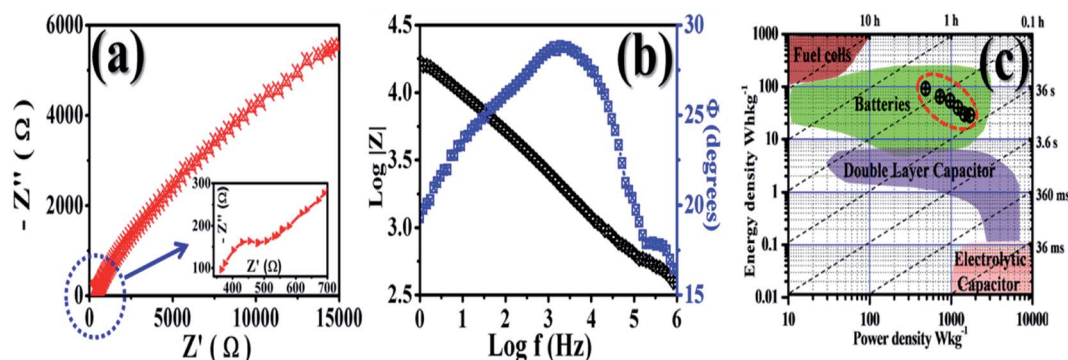


Fig. 7 (a) Nyquist (the arrow shows the inset), (b) Bode and (c) Ragone plots (the dotted circle indicates our results) for the GMH electrode.

Table 3 Circuit parameters generated after fitting to a simulated equivalent circuit

Parameter	Magnitude with units	Error (%)	Parameter	Magnitude with units	Error (%)
Z_s	0.30 kΩ	10.00	Q_s	0.10 milli stat C	07.00
Z_1	0.50 kΩ	10.50	Q_{S-GMH}	0.32 stat C	04.50
Z_2	1.50 kΩ	13.00	Q_G	0.01 stat C	01.75
Z_3	2.50 kΩ	23.00	Q_{eff}	03.18 milli stat C	16.00
Z_4	3.00 kΩ	06.70	L_1	14.61 mH	03.82
Z_5	3.50 kΩ	10.00	L_2	28.74 mH	04.50

The impedance spectrum of GMH in the high frequency region exhibits a semicircle with a low diameter and a straight line in the low frequency region. The intersection of the real axis in the high frequency region represents the solution resistance R_s and the related charge transfer frequency between the electrode material and the electrolyte. The value of R_s was found to be very low ($\sim 5 \Omega$), indicating the highly conducting nature of the electrode. The low solution resistance suggests that the GMH cell can be assembled in a series for high voltage applications with a minimum loss of energy. Further, the presence of a straight line in the high frequency region shows ideal supercapacitor behaviour. The inset shows that the enlarged plot consists of a broad, curved line instead of a semicircle, which is due to very low faradaic resistance and is indicative of pseudocapacitive behaviour.

In order to understand the frequency-dependent impedance characteristics of the GMH electrode, the logarithm of the absolute impedance $|Z|$ and the phase shift ϕ were studied. The almost linear nature is indicative of swift electron exchange at the solution/electrolyte interface. Two phase angles at 30° and 5° were found in the mid and low frequency regions, respectively, showing the pseudocapacitive behaviour of the blend electrode. This may be due to the diffusive characteristic of the electrode electrons.

Ragone plots were used to compare the performance of the current surfaces; they show the relationship between the power density, P_D , and the energy density, E_D . The calculated power densities are 486, 729, 1213, 1458, and 1701 $W kg^{-1}$ and the energy densities are 92.3, 65.1, 54.2, 30, 28.9 $W h kg^{-1}$, respectively, with various current densities @ 1, 1.5, 2, 2.5, 3,

and $3.5 A g^{-1}$. The values of C_s , P_D , and E_D are shown in Table 2 and are compared with those of reported forms of carbon. The material shows good performance in aqueous solution, where the pore accessibility limits the energy excitation at higher power. The Ragone plot shows the utility of GMH in the area of battery application. Therefore, the prepared composite material, GMH, is found to be suitable for application as an ideal electrode material for batteries. The parametric values obtained from the fitting are shown in Table 3.

In the Nyquist study, a semicircle pattern was obtained with a suppressed shape (shown in respective inset) above the real axis. Using the data values of real and imaginary $|Z|$, a circuit model was simulated and best fitted for discrete components such as impedance (R), inductance (L), and charge accumulated (Q). The simulated circuit was superimposed onto the GMH heterostructure and is schematically shown in Fig. 8. The impedance component, being a complex entity, can be represented in terms of a constant phase element due to the semi-circular pattern obtained in the Nyquist plot.

From the schematic, one can see that the equivalent circuit is composed of a series of impedances arising because of the GMH interface, its respective surfaces, the capacitive components (neglected during the simulation), the interface with the solution resistance, the mutual inductance arising from highly conducting graphene/MoS₂, and the effective charge accumulating at the grain boundaries. Notably, the sequence of the self-assembled heterostructure is immaterial. All this, together with the constant phasor element, leads to the equation:

$$Z_{eff} = Z^{-\alpha} \times L^{\beta} \times Q_{GMH}^{-1}$$



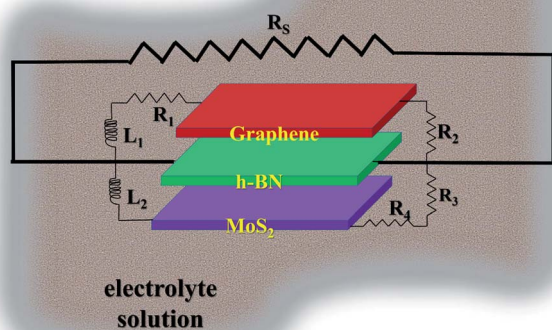


Fig. 8 Schematic of the GMH-blend electrode in electrolyte superimposed on the simulated equivalent circuit.

where the symbols have their usual meaning and the indices show the series/parallel components.

The parameters obtained for the discrete components by the circuit simulation are shown in Table 3. One can see that impedance, $Z^{-\alpha}$, systematically increases from the bulk solution to the self-assembled heterostructure electrodes, including the charge accumulation dynamics, Q_{GMH}^{-1} . Both mutual inductances, L^{β} , increased further by a factor of two in the cases of rGO/h-BN and h-BN/MoS₂; this indicates strong synergistic effects because of the combination of the two materials, namely MoS₂ and h-BN.

Further, electron spin resonance spectroscopy (ESR) revealed the dynamics of radical electrons in a solid system resonating para-magnetically by absorption of microwave energy. The purpose of ESR analysis is to estimate the diffusion dynamics of these radical electrons in the GMH blend *vis-a-vis* their individual counterparts.

As a first step towards computing the electron transport parameters, the magnitude of the line width determined by the

peak to peak distance, ΔH_{pp} , and the shapes of the ESR spectra were investigated. Fig. 9 shows the variation in intensity with applied magnetic field for rGO, MoS₂ and GMH-composite at room temperature. The ESR spectrum of h-BN was not obtained. In the cases of rGO and MoS₂, the first absorption derivative dY/dH as a function of the applied field shows an unsymmetrical nature, while for GMH-composite, symmetric and homogenous peaks were observed. Further, the principal parameter governing spintronics usability is the spin-lattice relaxation time, T_{sl} , which characterizes non-thermal spin states around the lattice. From the measured values of ΔH_{pp} and using $1/T_{\text{sl}} = ((28.0 \text{ GHz})/T \text{ (in K)}) \times \Delta H_{\text{pp}}$, the T_{sl} and electron diffusion length, D_{L} , were measured, with the Fermi velocity assumed to be $\sim 10^6 \text{ m s}^{-1}$ in all 2D materials.^{47–50} For MoS₂, the D_{L} was estimated to be $\sim 200 \mu\text{m}$, whereas that for rGO is $\sim 600 \mu\text{m}$. The effective D_{L} was found to be $\sim 1500 \mu\text{m}$ for GMH-composite. This indicates that on charging, the electron diffusion is quite high for GMH compared with its individual components. This also supports the conductivity measurements.

Plot (b) shows the IV plots for rGO, MoS₂ and GMH using two-probe conductivity measurements at room temperature. In the case of h-BN, no such curve was obtained. The calculated specific conductivity values of prepared pallets of rGO, MoS₂ and GMH are 45.5 ± 3.5 , $5.5 \pm 15 \times 10^6$, and $260 \pm 30 \text{ S m}^{-1}$ respectively. The specific conductivity of rGO is very high compared to those of GMH and MoS₂. The observed decrease is due to the presence of h-BN in the composite. Due to this, the magnitude of T_{sl} and the implication of D_{L} prove that the composite material is good for energy storage. Therefore, we further prepared a cell with the GMH blend, and its performance was tested.

3.5 GMH cell: performance

GMH-composite was implemented as an active electrode material for a laboratory scale demonstration of *blend-cell* action. The electrodes were developed as described in ref. 35; Fig. 10(a) shows the corresponding apparatus, starting

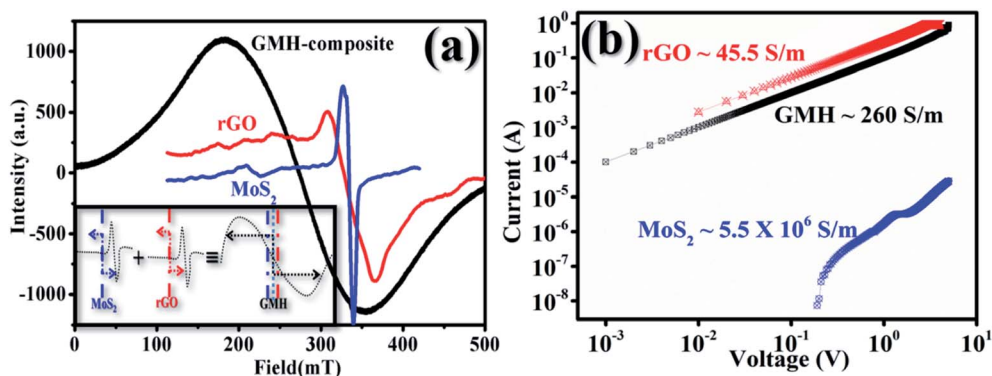


Fig. 9 (a) Recorded electron spin resonance profiles for rGO, MoS₂, and GMH-composite @ 300 K. The inset shows the scheme of electron diffusion for the individual 2D materials and their composite. The arrows schematically indicate the electron diffusion length, D_{L} , during charging in MoS₂, rGO, and GMH-composite, respectively, at 200, 600, and 1500 μm . (b) Recorded I - V profiles for rGO, MoS₂, and GMH-composite @ 300 K indicating the specific conductance. In both measurements, profiles for h-BN were not obtained.



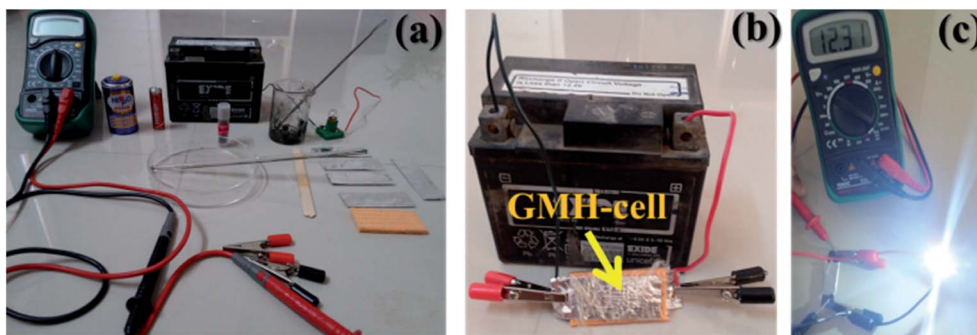


Fig. 10 Laboratory scale demonstration of GMH electrode demonstrating LED illumination. (a) Apparatus with the substrate, digital multimeter, 12.4 V commercial automobile battery, active electrode material, carbon powder, PVDF, etc., (b) charging of the blend-cell with a battery, (c) LED illumination after charging, indicating the voltage attained by the fully sealed fabricated blend electrode assembly.

materials, flat electrodes, etc. The assembly was packed with aluminium foil, as shown in image (b). Prior to charging, the voltage between the two electrodes was measured and was found to be negligibly small (~ 0.01 mV). The assembly was connected to *croc-clips* with conducting wires. The cell was charged for about 1.5 min with a nominal 12.50 V automobile battery. The voltage was instantaneously measured to be ~ 12.30 V after charging, as shown in image (c). The voltage decreased gradually with time, indicating recombination of developed charges across the 2D materials. The Helmholtz layer degradation time was found to be very small for such a primitive assembly. The charging measurements were performed repetitively. In one case, we connected the assembly to a commercially available LED equipped with current-carrying wires. The LED was found to be illuminated, as shown in image (c), for a period of approximately 5 min. These results show laboratory scale performance of the fabricated electrode using the GMH molecular blend. No such effect was observed for an rGO specimen prepared and tested in identical fashion.

Further ESCA analysis of GMH-composite after 20 000 cycles revealed that there was no change in the elemental composition of GMH (ESI Fig. S1†).

4. Conclusions

We have developed a porous self-assembled nano-composite medium of 2D materials comprised of GMH, rGO, MoS₂, and h-BN to use as an active material for evaluating supercapacitor performance. The symmetrical EDLC electrodes prepared using the GMH blend showed the highest C_{sp} of ~ 683 F g⁻¹ @ 1 A g⁻¹ (current density) with the highest energy density, E_D , of ~ 90 W h kg⁻¹; meanwhile, the power density, P_D , was ~ 1700 W kg⁻¹. The ESR studies underline that the diffusion length of the electrons is increased to a large extent, which effectively benefits the migration of electrolytic ions into the multi-layered structures. The presence of a small semicircle in the high frequency region shows favourable solution resistance and an effective charge transfer mechanism between the blend and electrolyte. At high frequency, the electrode shows behaviour of an ideal supercapacitor compounded with marginally low faradaic resistance. The phase angles of 30° and 5° in the mid and low

frequency regions, respectively, are indicative of hybrid behaviour as analysed by Bode. The Ragone analysis indicated electrode performance in the battery-like region. The present study suggests that this unique combination of GMH is an ideal material for high performance supercapacitor applications.

Conflicts of interest

There is no conflict to declare.

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