REVIEW ARTICLE
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Photo/electrocatalysis and photosensitization using metal nanoclusters for green energy and medical applications
Photo/electrocatalysis and photosensitization using metal nanoclusters for green energy and medical applications

Tokuhisa Kawawaki, Yuichi Negishi* and Hideya Kawasaki*

Owing to the rapidly increasing demand for sustainable technologies in fields such as energy, environmental science, and medicine, nanomaterial-based photo/electrocatalysis has received increasing attention. Recently, synthetic innovations have allowed the fabrication of atomically precise metal nanoclusters (NCs). These NCs show potential for green energy and medical applications. The present article primarily focuses on evaluation of the recent developments in the photo/electrocatalytic and photosensitizing characteristics of metal and alloy NCs. The review comprises two sections: (i) photo/electrocatalysis for green energy and (ii) photosensitization for biomedical therapy applications. Finally, the challenges associated with the use of metal NCs are presented on the basis of current developments.

1. Introduction

During the past few decades, photocatalysis and electrocatalysis have received significant attention as a result of the increasing demand for sustainable technologies in the fields of energy, environmental science, and medicine. Photo/electrocatalytic approaches rely on electronic excitation, and their performance depends on the ability to create electron (e^-)–hole (h^+) pairs that successively undergo chemical reactions with other compounds via oxidative (e.g., 2H_2O(l) → O_2(g) + 4H^+ (aq) + 4e^-) and reductive reactions (e.g., 2H^+ (aq) + 2e^- → H_2(g)). Many advanced nanomaterial-based photo/electrocatalysts have been synthesized and reported, and their advantages include large surface-to-volume effects, numerous catalytic active sites, quantum size effects and high stability. These catalysts are considered to be promising for energy and environmental applications, such as photo/electro water splitting to generate hydrogen (H_2) and conversion of carbon dioxide (CO_2). Furthermore, they are used in the fuel industry and in water treatment and disinfection, air purification, and self-cleaning surfaces. Usually, the term “photo/electrocatalysis” refers to photo/electrochemical reactions, which involve an electron transfer. Conversely, when energy transfer occurs in

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Received 15th September 2019
Accepted 17th October 2019
DOI: 10.1039/c9na00583h
rsc.li/nanoscale-advances

Cite this: Nanoscale Adv., 2020, 2, 17
a photochemical reaction, the process is often called “photosensitization”. Triplet photosensitizers are used not only for triplet energy transfer, but are also utilized in photodynamic therapy (PDT) for medical applications.7

One drawback of these nanomaterial-based photo/electrocatalysts is that they are generally poly-dispersed in size, morphology, and chemical composition. The heterogeneity of photo/electrocatalysts with multiple active sites often results in different catalytic activities and varying catalytic selectivity. Therefore, it is important to fabricate nanomaterial-based photo/electrocatalysts that can be controlled well at the atomic level. Recent synthetic innovations have allowed the fabrication of atomically precise metal nanoclusters (NCs).8–18 These metal NCs contain a few to hundreds of metal atoms that are protected by organic ligands (e.g., thiolate, phosphine, small molecules, synthetic polymers, and biomolecules). Particularly, thiolate-protected metal NCs and their alloy NCs can be prepared with atomic precision as $M_{n}(SR)_{m}$ (M = gold (Au), silver (Ag), copper (Cu), platinum (Pt), palladium (Pd), or other element; SR = thiolate ligand) in a wide range of sizes by varying the $n$ and $m$ values. As mentioned previously, such NCs have potential applications in the fields of energy, environmental science, medicine, and other.8–18 The $M_{n}(SR)_{m}$ NCs have discrete energy levels due to the quantum size effect, making them significantly different from the larger (>3 nm) plasmonic metal (Au or Ag) nanoparticles (NPs) with continuous energy levels. By adjusting $n$ and $m$, one can control the discrete electronic/ geometrical structure of $M_{n}(SR)_{m}$ NCs and thereby their photosensitizing and photo/electrocatalytic capability. Recently, several excellent reviews have discussed the syntheses of metal NCs,19 photoluminescence,20 and their biomedical21 and catalytic applications.22 However, to the best of our knowledge, there are no publications focusing on the photo/electrocatalytic and photosensitizing characteristics of metal NCs for applications in the green energy and medical fields. Therefore, a survey of recent progress, with a view to improving the photo/electrocatalytic and photosensitizing capability of metal NCs is needed.

In this review, we discuss recent achievements and evaluate the photo/electrocatalytic and photosensitizing characteristics of metal NCs and their alloy NCs. The review comprises two sections: (i) photo/electrocatalysis and photosensitization for green energy applications (specifically solar cell), and (ii) photosensitization for medical therapeutic applications (Fig. 1). We hope to provide a clear picture on how the photo/electrocatalytic and photosensitizing capability of metal NCs can improve their performance and promote the development of new green energy and medical applications.

2. Photo/electrocatalysis of metal NCs for green energy applications

In this section, we discuss recent progress of the photo/electrocatalytic (specifically reaction of water splitting and fuel cells) and photosensitizing applications (specifically solar cell) of metal NCs (Pd, Ag, Pt, or Au) and their alloy NCs. We summarized references used in this section in Table 1.

2.1. Electrocatalytic applications for water splitting

Water electrolysis is one of the cleanest energy-producing processes, generating $H_2$ using only water and electricity, which is regarded as the next-generation energy source or raw material for chemicals. Water electrolysis comprises two half-reactions, which are the $H_2$ evolution reaction (HER) and the $O_2$ evolution reaction (OER). When supplying electricity to metal electrodes, a reduction reaction occurs on the cathode and an oxidation reaction on the anode, and thereby water molecules are split into $H_2$ and $O_2$. However, the reaction does not proceed if the voltage supplied to the electrodes is higher than the redox potential of each reaction (HER: 0 V vs. SHE, OER: 1.23 V vs. SHE) due to the high activation energy. Therefore, investigations have been conducted to enhance the reactions using metal catalysts, such as precious metal NPs. An increased specific surface area of the active sites in the metal

![Fig. 1 Schematic of photo/electrocatalysis and photosensitization using metal NCs for green energy and medical applications. Reprinted with permission from ref. 51 and 75. Copyright American Chemical Society; ref. 127. Copyright 2015 Royal Society of Chemistry.](image-url)
The electrocatalytic reaction for fuel cells is closely related to the energy level (Fermi level) of the d-band center of each metal and the geometrical structure of the surface, which depends on the crystal plane and interatomic distances. These characteristics result in the tendency to form peaks in the activity volcano plots. The optimum metal types and their alloy NPs have been prepared based on the activity volcano plots obtained using various metal elements. Further reduction of the metal particle sizes induces the appearance of novel electronic and/or geometrical structures, which are different from the bulk metal. Therefore, recently, a number of studies on the electrocatalytic application of metal NCs composed of approximately several tens of atoms or less have been reported.

### 2.1.1. H₂ evolution reaction

HER, the half-reaction of water splitting, proceeds through the Volmer–Heyrovsky or Volmer–Tafel mechanism, whereby metal surface atoms form bonding orbitals with protons, resulting in the production of H₂.

Under acidic conditions, the reaction proceeds as follows:

**Volmer reaction:**

\[
\text{H}^+ + e^- \leftrightarrow \text{H}_{\text{ads}}
\]

**Heyrovsky reaction:**

\[
\text{H}_{\text{ads}} + \text{H}^+ + e^- \leftrightarrow \text{H}_2
\]

**Tafel reaction:**

\[
2\text{H}_{\text{ads}} \leftrightarrow \text{H}_2
\]

On the other hand, under alkaline conditions, the following reaction sequence occurs:

**Volmer reaction:**

\[
\text{H}_2\text{O} + e^- \leftrightarrow \text{H}_{\text{ads}} + \text{OH}^-
\]

**Heyrovsky reaction:**

\[
\text{H}_{\text{ads}} + \text{H}_2\text{O} + e^- \leftrightarrow \text{H}_2 + \text{OH}^-
\]

**Tafel reaction:**

\[
2\text{H}_{\text{ads}} \leftrightarrow \text{H}_2
\]

Here, the “ads” subscript denotes a reaction site on the catalyst surface. So far, noble metals, such as Pt, have been known for their high activity in HER, and attempts have been made to increase this activity further via miniaturization of the metals, and to predict their activity using density functional theory (DFT) calculations. For example, Sun et al. studied the HER characteristics by preparing single Pt atoms and Pt NCs supported on nitrogen-doped graphene nanosheets (NGNs) using the atomic layer deposition technique. The results revealed that the HER activity of Pt atoms on Pt NCs (ALD50Pt) was 10.1 A mg⁻¹, which is 37.4 times higher than the activity of the commercial carbon-supported Pt NP (Pt/C) catalyst (0.27 A mg⁻¹) (Fig. 2A and B). Furthermore, Chen et al. demonstrated that Pt NCs confined in a calixarene-based Ni₃₄ coordination cage (Fig. 2C) show much higher activity in HER than the commercial Pt/C catalyst (Fig. 2D and E). The same group also reported that the activated carbon-supported Pd₆₈/Pd₆₈/AC-V; the ligands were removed in this case) shows the current density 10.22 times higher than the commercial Pt/C catalyst at −0.423 V.

In contrast, bulk Au, which belongs to group 11, shows almost no HER activity. Reduction of the size to the cluster region, however, results in changes in the electronic structure, and thereby Au NCs exhibit HER activity. Based on the recent studies, Au NCs can now be synthesized using a very simple and easy approach. Moreover, the sizes can be controlled at atomic precision and their functions can be modulated via ligand exchange and heteroatom doping. In recent years, a number of studies have been conducted on the HER activity of Au NCs. For example, Teranishi and Sakamoto et al. studied the effect of the ligand structure on the HER activity using porphyrin-coordinated Au NCs (Fig. 3). In their study, the distance between the porphyrin-ring and the surface sulfur (S) directly bonded to Au NCs was controlled by varying the number of methylene groups between those moieties (Fig. 3A and C). They found that the HER activity increases with decreasing distance between the porphyrin-ring and the Au NC surface (Fig. 3G and H). In addition, it was demonstrated that tetrakis(5z,10z,15z,20z(2-acetylthiomethylphenyl)porphyrin (SC₄P)-protected Au NCs, with a 3.4 Å distance between the porphyrin-ring and the Au NC surface, show a high current density of 460% at 0.4 V vs. RHE, in comparison with Au NCs protected by ordinary phenylethanethiolate (PET). The study also demonstrates that the ligand structure strongly affects the HER activity and that the smaller Au₆(SR)₃₈ NCs exhibit the higher HER activity.
On the other hand, Lee and Jiang et al. investigated the effect of heteroatom doping of Au_{25}(SR)_{3m} NCs on the HER activity. They examined the HER activity of hexanethiolate (C6)-protected Au_{25} and PtAu_{24} NCs in a tetrahydrofuran (THF) electrolyte containing 1.0 M trifluoroacetic acid (TFA) (Fig. 4A–C). The obtained results demonstrated that PtAu_{24} NC has better catalytic activity with a higher starting potential (0.89 V) than Au_{25} NC (1.10 V). The same research group showed that the improved catalyst current and turnover frequency can be obtained with Pd as the central atom, instead of Pt (PtAu_{24} > PdAu_{24} > Au_{25}). The same tendency can be observed for the other-sized C6-protected Au_{38} NCs (Pt2Au_{36} > Pd2Au_{36} > Au_{38}). These results are in good agreement with those obtained from DFT calculations, which predicted that the HER activity significantly varies depending on the doping elements. Hence, the control of the electronic structure by alloying at the atomic level greatly affects the HER activity in the cluster size range.

In addition to these studies, Jin et al. investigated the HER activity of the composite comprising PET-protected Au_{25} NCs and molybdenum disulfide (MoS_{2}) nanosheet (MoS_{2} nanosheet carrying PET-protected Au_{25}; Au_{25}/MoS_{2}; Fig. 4D and E). The Au_{25}/MoS_{2} showed a current density 1.79 times higher than that of MoS_{2} at −0.4 V, indicating that the improvement of HER activity of MoS_{2} nanosheet was caused by loading Au_{25} NCs. X-ray photoelectron spectroscopy analysis showed that a negative shift of −0.4 eV in the Mo 3d orbital energy was induced by loading Au_{25} NCs. This implies that the partial charge transfer occurs from Au_{25} NCs to MoS_{2}, leading to the increased HER activity. The HER activity of the composite is, therefore, considerably improved by the electronic interaction between the materials. Regarding the HER activity of the composite comprising MoS_{2}, Zhu et al. also evaluated Au_{2}Pd_{6}/MoS_{2}.

2.1.2. O₂ evolution reaction. OER is a multi-step four-electron reaction dependent on the binding energy of the OER intermediate (O, OH, OO, etc.). Under acidic conditions, the reaction proceeds as follows:

\[
\text{H}_2\text{O} \rightarrow \text{OH}_{\text{ads}} + \text{H}^+ + e^- \\
\text{OH}_{\text{ads}} \rightarrow \text{O}_{\text{ads}} + \text{H}^+ + e^- \\
\text{O}_{\text{ads}} + \text{H}_2\text{O} \rightarrow \text{OOH}_{\text{ads}} + \text{H}^+ + e^- \\
\text{OOH}_{\text{ads}} \rightarrow 2\text{O}_{\text{ads}} + \text{H}^+ + e^- 
\]
On the other hand, under alkaline conditions, the following reaction sequence takes place:

\[
\begin{align*}
\text{OH}^- & \rightarrow \text{OH}_{\text{ads}} + e^- \\
\text{OH}_{\text{ads}} + \text{OH}^- & \rightarrow \text{O}_{\text{ads}} + \text{H}_2\text{O} + e^- \\
\text{O}_{\text{ads}} + \text{OH}^- & \rightarrow \text{OOH}_{\text{ads}} + e^- \\
\text{OOH}_{\text{ads}} + \text{OH}^- & \rightarrow 2\text{O}_{\text{ads}} + \text{H}_2\text{O} + e^- \\
2\text{O}_{\text{ads}} & \rightarrow \text{O}_2
\end{align*}
\]

According to this reaction pathway, the OER activity of the catalyst generally depends on the binding energy of the OER intermediates on the surface (O, OH, OOH, etc.). An appropriate binding energy of the oxygen species (neither too high nor too low) enables the catalyst suitable for OER. For this reason, the metal oxides, such as iridium oxide (IrO\textsubscript{3}), ruthenium oxide (RuO\textsubscript{2}), exhibit high activity. Therefore, extensive efforts have been made on their miniaturization, the theoretical prediction of the activity and the mechanism.\textsuperscript{15–38}

Moreover, research regarding cluster-sized particles is progressing. Chen and Gao reported that dodecanethiolate (C12)-protected Pd\textsubscript{6}/AC exhibits a current density 37 times higher than that of commercially available Pt/C at an applied voltage of 1.83 V (Fig. 5A and B).\textsuperscript{26} Although ligands were considered to be disadvantageous for electrode reactions, this study revealed that the higher electron density of C12-protected Pd\textsubscript{6}/AC enhances the desorption process of oxygen atoms or molecules and is favorable for OER. In addition, there are several reports describing the suitability of M\textsubscript{31}(SR)\textsubscript{31} NCs for OER. For example, Hussain and Joya \textit{et al.} reported that the OER activity of Ni\textsubscript{31}(PET)\textsubscript{31} NC is equal to that of a RuO\textsubscript{2} electrocatalyst.\textsuperscript{39}

As with HER activity, OER activity has also been studied for the composites including Au NCs. Jin \textit{et al.} investigated the OER activity of the composite comprising PET-protected Au\textsubscript{25} NCs and cobalt diselenide (CoSe\textsubscript{2}) nanosheet (CoSe\textsubscript{2}) nanosheet carrying PET-protected Au\textsubscript{25}; Au\textsubscript{31}/CoSe\textsubscript{2} and revealed that Au\textsubscript{25}/CoSe\textsubscript{2} shows a current density of 10 mA cm\textsuperscript{-2} with a small overvoltage of approximately 0.43 V (\textsim 0.52 V with CoSe\textsubscript{2} alone) (Fig. 6).\textsuperscript{40} The group also investigated the effect of Au NC size on OER activity using a series of Au\textsubscript{n}(SR)\textsubscript{m} NCs (Au\textsubscript{10}, Au\textsubscript{25}, Au\textsubscript{144}, and Au\textsubscript{333}). It was revealed that OER activity increases with increasing the Au NC size. Furthermore, Peng and Zeng demonstrated that Au NC/CoSe\textsubscript{2} is superior for increasing the OER activity compared to commercially available Ir/C.\textsuperscript{41}

### 2.2. Electrocatalytic application for fuel cells

The ultimate energy-conversion systems used for power generation are fuel cells utilizing materials such as H\textsubscript{2} and methanol obtained from natural energy sources. Such methods involve circulating energy systems that do not use any fossil fuels and only release water as the waste product.

Fuel cells are roughly classified into those using alcohol and those using H\textsubscript{2}. The latter involves the H\textsubscript{2} oxidation reaction (HOR) and O\textsubscript{2} reduction reaction (ORR), which are the reverse reactions of HER and OER, respectively. HOR is a one-electron reaction system where the catalysts displaying HER-activity can be useful. In contrast, ORR is a four-electron reaction system, and since it proceeds through a complex reaction pathway, its reactivity is different from that of OER.

![Fig. 4](image.png)

**Fig. 4** (A) Geometrical structures of C6-protected Au\textsubscript{25} and PtAu\textsubscript{24} NCs (golden, gold atoms of the core; olive, gold atoms of the shell; grey, sulfur). (B) The square-wave voltammogram. (C) HER polarization curves of C6-protected Au\textsubscript{25} and PtAu\textsubscript{24} NCs. (D) High angle annular dark-field SEM (HAADF-STEM) images. (E) HER polarization curves of the C6-protected Au\textsubscript{25}/MoS\textsubscript{2} composite. Panels (a)–(c) are reproduced with permission from ref. 29. Copyright 2017 Springer Nature. Panels (d) and (e) are reproduced with permission from ref. 33. Copyright 2017 Wiley-VCH.

![Fig. 5](image.png)

**Fig. 5** (A) OER polarization curves. (B) Mass OER activity of Pd\textsubscript{8}/AC, Pd\textsubscript{8}/AC-V and Pt/C at each overpotential. Panels (a) and (b) are reproduced with permission from ref. 26. Copyright 2017 Royal Society of Chemistry.
Fig. 6 (A) UV-vis and (B) MALDI-mass characterization of Au$_{25}$(PET)$_{18}$ NCs. Inset of (A): crystal structure of Au$_{25}$(PET)$_{18}$ NCs. (C) TEM image of CoSe$_2$ nanosheets. (D) and (E) HAADF-STEM images of Au$_{25}$/CoSe$_2$ composite with different magnifications. (F) OER polarization curves. (G) Mass OER activity of Au$_{25}$/CoSe$_2$ composite, CoSe$_2$, Au$_{25}$ supported on carbon black (CB) and commercial Pt/CB. Panels (a)–(g) are reproduced with permission from ref. 40. Copyright 2017 American Chemical Society.

Under acidic conditions, ORR proceeds as follows:

\[
\begin{align*}
O_2 + 4H^+ + 4e^- & \rightarrow H_2O \\
O_2 + 2H^+ + 2e^- & \rightarrow H_2O \\
H_2O_2 + 2H^+ + 2e^- & \rightarrow 2H_2O
\end{align*}
\]

On the other hand, under alkaline conditions, the following reaction sequence takes place:

\[
\begin{align*}
O_2 + H_2O + 4e^- & \rightarrow 4OH^- \\
O_2 + H_2O + 2e^- & \rightarrow OOH^- + OH^- \\
OOH^- + H_2O + 2e^- & \rightarrow 3OH^-
\end{align*}
\]

The first step in both of these reaction pathways is the scission of the O-O bond. Regarding fuel cell applications, the theoretical redox potential of the direct four-electron pathway is 1.23 V vs. SHE, which is greater than that of the indirect two-electron pathway (0.68 V vs. SHE). Therefore, the direct four-electron pathway is the preferred reaction system, owing to the higher energy-conversion efficiency. Since the ORR on the reduction side is slower than on the oxidation side, the former is the rate-limiting step in the fuel cell. Hence, understanding the factors controlling ORR and thereby the creation of high-performance ORR catalysts are essential for the development of efficient fuel cells.

Since Pt exhibits high ORR activity as well as high OER activity, the development of the miniaturization techniques, prediction of the activity by theoretical calculations, and the mechanism for Pt NCs have been studied extensively. For example, Yamamoto et al. examined the ORR activity of fine Pt NCs (Pt$_n$, NC, $n = 12–24$) using dendrimer-encapsulated atomically-controlled Pt$_n$ NCs (Fig. 7). The results demonstrated that Pt$_{19}$ NC exhibits the highest ORR activity (Fig. 7E). The edge site of Pt$_{19}$ has an ideal oxygen binding energy, which indicates that in addition to the electronic structure, the geometrical structure affects the ORR activity.

Moreover, the use of Au$_{n}$(SR)$_{m}$ NCs as catalysts for ORR has also attracted significant attention as an alternative to expensive Pt catalysts and to evaluate the catalytic reaction mechanism. Chen et al. reported that the ORR activity increases with a decrease of Au NC size (Au$_{15}$(PET)$_{18}$ > Au$_{16}$(PET)$_{24}$ > Au$_{14}$(PET)$_{30}$, Fig. 8A and B). This increase is suggested to be due to the enhancement of the oxygen adsorption caused by increasing the proportion of surface atoms and shifting the d-band center associated with a decrease of Au NC size. Chakraborty and Dass et al. synthesized a series of 4-tert-butylbenzenethiolate (TBBT)-protected Au NCs and analyzed the ORR activity of Au NCs supported on single-walled carbon nanotubes. The study on the reaction rate revealed that the ORR activity increases in the order of Au$_{36}$(TBBT)$_{24}$ > Au$_{133}$(TBBT)$_{32}$ > Au$_{279}$(TBBT)$_{84}$ > Au$_{32}$(TBBT)$_{20}$ (Fig. 8C and D). The same research group also used tert-butylthiolate (S-tBu) as a ligand and showed that the ORR activity increases in the order of Au$_{46}$(S-tBu)$_{29}$ = Au$_{133}$(S-tBu)$_{39}$ > Au$_{32}$(S-tBu)$_{14}$ > Au$_{32}$(S-tBu)$_{16}$ in the case of S-tBu-protected Au NCs.

In addition to the NC size, the control of charge state of NCs is also a crucial factor for obtaining high activity. For example, the negatively charged Au$_{25}$(PET)$_{18}$ NC shows higher activity in the production of H$_2$O$_2$, which is a two-electron reduction reaction, than the neutral Au$_{25}$(PET)$_{18}$ NC (Fig. 8E and F). Since H$_2$O$_2$ is a useful raw material for a variety of chemical products, the development of a highly selective H$_2$O$_2$ production method is of importance. Likewise, the interaction with the substrate also requires consideration. Pt or Au NCs supported on various substrates, such as indium tin oxide (ITO) and reduced graphene oxide (rGO) sheets show outstanding electrocatalytic performance in ORR, which was found to be originated in the interaction between metal NC and the substrate.

2.3. Photocatalytic application for water splitting

Artificial photosynthesis, which has attracted significant attention in recent years, is an ultimate energy production technique involving conversion of light into chemical energy. In 1972, Honda and Fujishima et al. discovered that water can be split into H$_2$ and O$_2$ by irradiating the titanium dioxide (TiO$_2$) photocatalyst with ultraviolet light. Utilizing this method allows...
efficient generation of H₂ as an energy source using abundant water, without the production of carbon dioxide. In the photocatalytic water splitting reaction, the catalyst comprises a photocatalyst, which absorbs light, and a cocatalyst whose role is to reduce the overpotential of HER and OER. Moreover, the photocatalyst uses light energy instead of electric energy as the driving force for HER or OER. Therefore, the position of the conduction and valence bands and the width of the bandgap are essential factors when considering the photocatalyst.

Highly active metal NCs in electrochemical HER and OER are effective cocatalysts in photocatalytic systems. The photocatalytic reaction involves several key steps: (1) generation of the e⁻/h⁺ pair in the photocatalyst, (2) transfer of the e⁻ and h⁺ to the active sites, and (3) generation of H₂ and O₂ at the active sites. Here the cocatalyst greatly affects processes (2) and (3). The photodeposition and impregnation methods are typical approaches for loading cocatalysts on photocatalysts. They offer several advantages, such as the formation of fine metal particles directly on the photocatalyst. However, the disadvantages of these methods include large size dispersibility.

It is possible to load the controlled NC on photocatalyst and thereby to deepen the understanding on the effect of cocatalyst size on the photocatalytic activity when using an ultra-high vacuum apparatus. Heiz, Feldmann, and Jäckel et al. loaded a series of atomically precise Pt NCs (Pt₁₃, Pt₁₇, and Pt₁₉ NCs) on cadmium sulfide (CdS) nanorod utilizing an ultra-high vacuum apparatus (Fig. 9A and B).²⁸,²⁹ They investigated the water-splitting activity of the resulting photocatalysts (Pt NC/CdS) and consequently revealed that Pt₁₉/CdS exhibits the highest activity (Fig. 9C). In the NC region, the bandgap increases with a decrease of the NC size due to the quantum size effect. It was also reported that the control of the lowest unoccupied molecular orbital (LUMO) position of the cocatalyst by controlling the NC size is of great importance for the effective electron transfer from the photocatalyst to the cocatalysts. Furthermore, the group demonstrated that the LUMO position must be lower than the conduction band of the semiconductor photocatalysts and higher than the reduction potential of HER (Fig. 9D).

As well as investigation into electrochemical properties of HER and OER, research on photocatalytic water splitting is rapidly progressing also for Auₙ(SR)ₘ NCs. Negishi and Kudo et al. developed a photocatalytic system comprising atomically precise Au₁₃ NC and the BaLa₄Ti₄O₁₅ photocatalyst (Fig. 10).⁶⁰ In this study, the atomically precise Au₁₃ NC was loaded on the...
described that the ORR, which is a reverse reaction, can be suppressed by coating the Au25 NC with Cr2O3, and thereby the activity was improved by 19-fold compared to the non-coated Au25/BaLa4Ti4O15 (Fig. 10C–E).44 The cocatalysts with high HER and OER activities typically display good activity in HOR and ORR.41,63 Overall, as shown in this study, the control of the reverse reaction is an important factor for improving the activity.

The photocatalytic activity could also be modulated by heteroatom doping on the cocatalysts. Negishi and Yamazoe et al. demonstrated that Pd doping on Au25 NC reduces the water-splitting activity, whereas Pt doping on Au25 NC enhances it (Fig. 11A–F).46 Yang et al. also examined the effect of heteroatom doping on the photocatalytic activity of PtAg24-loaded graphitic carbon nitride (PtAg24/g-C3N4).66 They revealed that PtAg24/g-C3N4 shows higher activity for photocatalytic H2 production than Ag25/g-C3N4 (Fig. 11G and H). In addition to the above studies, there are several reports on water-splitting photocatalytic activity using composites including metal NCs, such as graphitic carbon nitride nanosheets carrying mercaptosuccinic acid-protected Ag9 NC,66 TiO2 carrying SG-protected Au NC, etc.67,68

2.4. Photosensitization for solar cell applications

Solar power generation systems, which directly convert light energy into electricity, continue to be important renewable energy sources and are the key components of the world’s technological advancement. A photoelectrode cell, commonly known as the solar cell, comprises a light absorption layer, an electrolyte, and two electrodes (working electrode and counter electrode) that independently take out e– and h+. The material used for the light absorption layer must effectively absorb light in the visible range, constituting approximately 50% of sunlight. In this respect, bulk gold reflects light due to a large number of free electrons. However, with the decreasing size of the NC, the number of electrons decreases, and discrete energy levels appear. Hence, a metal NC has a HOMO–LUMO gap similar to that of a molecule and exhibits light absorption in the visible to the near-infrared (NIR) region.

Tatsuma and Sakai previously demonstrated that metal NCs can be used for the light absorption layer in solar cells.69 They produced a photoelectrode in which a TiO2 electrode was coated with a series of SG-protected Au NC of various sizes (Au NC/TiO2). When hydroquinone was used as an electron donor, a photo-response was observed (Fig. 12A and B) and its photocurrent behavior was consistent with photocurrent action spectrum and absorption spectrum (Fig. 12C and D). These results indicate that an electron is excited and promoted from the HOMO to the LUMO by photoirradiation in Au NC and that the photoexcited electron of LUMO is transferred to the CB of TiO2. In this way, the Au NC functions like a dye in dye-sensitized solar cells (DSSCs). The same research group observed the generation of photovoltaic power even when the other element NCs (Ag, Pt and Pd NCs) were employed.70,71 In addition, various other metal and alloy cluster sensitized solar cells have been reported by a number of research groups.72–74

Regarding the solar cells, Kamat et al. achieved a high external quantum efficiency (IPCE) of 70% at 400–425 nm by
using a Co(bpy) redox pair (Co(bpy)₃(PF₆)₂/Co(bpy)₃(PF₆)₃) in SG-protected Au NC/TiO₂ system (Fig. 13). This value compares favorably even with the CdS-based quantum dot solar cells. The power conversion efficiency (PCE) was 2.36% at maximum, indicating that metal NCs can act as a high-performance sensitizer in sensitized solar cells. Bang, Kang, and Lee et al. achieved a PCE value as high as 3.8% when investigating SG-protected Au₁₈ NC/TiO₂ solar cells using the I²/Co/C₃I₃ redox pair (Fig. 14). They reported that Au₁₈ NC barely underwent exciton recombination and absorbed light in a wide wavelength range, resulting in high solar cell characteristics (Fig. 14A–C).

However, to improve the stability of the system, other ligand-protected metal NCs should be used. Pradeep et al. compared the photovoltaic performance of solar cells using various ligand-protected metal NCs, such as bovine serum albumin, TBBT, and 4-mercaptobenzoic acid. They reported that the Au₃₀ NC protected with bovine serum albumin has highest PCE (0.35%).

In the described systems, there is a correlation between the expansion of the light absorption wavelength range of the metal NCs and the decrease in opening photovoltage accompanying the bandgap decrease. To further improve the PCE, combination with other light-absorbing materials is required. DSSC using a squalene dye as the co-sensitizer was reported to provide a PCE of 4%, and a maximum PCE of 9.15% (ref. 78) by incorporating Au₃₈ clusters in organic thin-film solar cells.

3. Photosensitization of metal NCs for therapeutic applications

3.1. Photo-based therapeutic applications of water-soluble metal NCs

Water-soluble noble metal (such as Au and Ag) NCs are composed of a metal core with several to a few hundred atoms, capped by water-soluble ligands such as amino acids, peptides, proteins, deoxyribonucleic acid (DNA), and hydrophilic thiolate ligands with different functional groups (e.g., –COOH, –OH, and –NH₂). These water-soluble metal NCs exhibit unique physico-chemical properties including high photostability, low toxicity, and luminescence. In addition, the good penetration of such ultra-small metal NCs into the cells, the ability to target tumors through...
characteristics of (C) Au_26\textsuperscript{25} cells with I\textsuperscript{3} \textsuperscript{-}/C\textsuperscript{0} or Co\textsuperscript{2+}/Co\textsuperscript{3+} redox pair. Panels (a)–(d) are reproduced with permission from ref. 76. Copyright 2016 American Chemical Society.

proper surface modification, and their superior renal clearance make the water-soluble metal NCs promising materials for utilization in the field of biomedicine, e.g., bioimaging, sensing, drug delivery, and diagnostic and therapeutic applications.

Cancer is a class of common life-threatening diseases, understanding of which is incomplete, despite significant efforts and extensive research. Developed therapies against cancer include chemotherapy (chemo-T), radiotherapy (RT), surgery, immunotherapy, NPs-based therapy, photothermal therapy (PTT), and PDT.\textsuperscript{79} Among them, PDT is considered clinically reliable and non-invasive. In this method reactive oxygen species (ROS) generated by a photoexcited photosensitizer kill the tumor cells.

Traditional PDT utilizes organic photosensitizers with organic dye molecules such as porphyrin and its derivatives, chlorin e6, methylene blue, rose bengal, eosin Y, and indocyanine green. However, organic photosensitizers suffer from common drawbacks including low water-solubility, poor selectivity, toxicity, and photo-instability. Meanwhile, known inorganic photosensitizers include quantum dots, silicon nanocrystals, fullerene (C\textsubscript{60}), and metal NPs.\textsuperscript{80–83} More recently, Au\textsubscript{x}(SR)\textsubscript{m} NCs have been considered as promising photosensitizers for PDT. Their electronic structures can be optimized for efficient singlet oxygen (\textsuperscript{1}O\textsubscript{2}) production using atomically precise size control. Several fundamental properties of the Au\textsubscript{x}(SR)\textsubscript{m} NCs are also desirable for PDT, such as good biocompatibility in dark conditions, photosensitization under near-infrared light irradiation, good resistance to photobleaching, and target specificity via surface modification. The following section focuses on the photosensitizing ability of Au\textsubscript{x}(SR)\textsubscript{m} NCs toward PDT applications. A number of recent reviews on other biomedical applications of metal NCs have also been published.\textsuperscript{15–17}

3.2. Basic principle of photosensitization for PDT

Upon absorbing the appropriate light, the photosensitizer transitions from the ground singlet state (S\textsubscript{0}) to the excited singlet state (S\textsubscript{1}). Through non-radiative transition, it subsequently changes the multiplicity from S\textsubscript{1} to a triplet state (T\textsubscript{1}) via intersystem crossing (ISC). The photosensitizer in the T\textsubscript{1} state can then decay via two distinctive processes: Type I and Type II (Fig. 15). In Type I, the excited photosensitizer undergoes single-electron transfer with the substrate to produce a radical or a radical ion, such as hydroxyl radical (\textsuperscript{.}OH) and superoxide (O\textsubscript{2}\textsuperscript{-}). These highly reactive radical species can further react with biological substrates to modify their structure and/or function. In the Type II process, however, \textsuperscript{1}O\textsubscript{2} produced \textit{via} the energy transfer from the photosensitizer to oxygen, readily reacts with biological molecules (e.g., lipids, proteins, and DNA), which are the main components of cells and nuclear membranes. The majority of the effective photosensitizers are characterized by high quantum yields of their T\textsubscript{1} state, since a relatively longer living T\textsubscript{1} state allows more chance for energy and/or electron transfer. Thus, ROS such as O\textsubscript{2}\textsuperscript{-}, \textsuperscript{.}OH, and \textsuperscript{1}O\textsubscript{2} are key species in the PDT process.\textsuperscript{84,85}

The lifetime of \textsuperscript{1}O\textsubscript{2} in most biological environments is in the range of a few microseconds. Thus, \textsuperscript{1}O\textsubscript{2} produced from Type II photosensitization can affect biological substrates within a moderate distance from the photosensitizer itself (0.02–0.15 \textmu m). In contrast, \textsuperscript{.}OH produced in Type I photosensitization is highly reactive and interacts with biological molecules within a range of less than 5 nm, restricting its effects to the location where it is produced.\textsuperscript{84,85} Thus, \textsuperscript{1}O\textsubscript{2} formed from the Type II reaction is thought to be primarily responsible for the biological effect in PDT.

The direct confirmation of \textsuperscript{1}O\textsubscript{2} generation in photosensitization reactions comes from detecting phosphorescence at around 1270 nm resulting from the spontaneous decay of \textsuperscript{1}O\textsubscript{2} to its ground state. Other methods of \textsuperscript{1}O\textsubscript{2} detection are based on analyzing the spectral change of a probe molecule upon its interaction with \textsuperscript{1}O\textsubscript{2}, using electron spin resonance (ESR), UV-vis absorption, or fluorescence. The various \textsuperscript{1}O\textsubscript{2} detection techniques and their probes are summarized in Table 2.
1O2 scavengers can also be used to confirm the 1O2 generation. Azide ion and histidine are often utilized as quenchers for 1O2. They rapidly and preferentially interact with 1O2 and convert it to the ground state O2 before reacting with any other molecules in the system. Further confirmation of 1O2 can be obtained by replacing water with deuterated water, which extends the lifetime of 1O2 by as much as 15-fold.

3.3. Au NC-based photosensitizer for PDT

In previous studies, protein-stabilized Au NCs and Au NCs embedded in polymer matrix were reported to produce 1O2 upon photoexcitation at 330 and 532 nm (UV and visible light, respectively). Following these reports, Kawasaki and Jin et al. described the formation of 1O2 through direct sensitization of Au25(SR)18 (H-SR = phenylethanethiol or captopril) under visible/NIR light (532, 650, and 808 nm) irradiation. The generation of 1O2 by photoexcited Au25(SR)18 was confirmed by observation of the 1O2 emission at approximately 1270 nm, 1O2-selective probes, and the use of a 1O2 quencher. The larger HOMO–LUMO gap of Au25(SR)18 NCs (1.3 eV) than the energy of 1O2 (0.97 eV), the long lifetime of the electronic excitations, and the well-defined O2 adsorption sites are suggested to be the key factors promoting the energy transfer from the Au25(SR)18 NCs to molecular oxygen resulting in the formation of 1O2. Photodynamic activity of water-soluble Au25(Capt)18 NCs toward cancer cells under NIR light (808 nm) was also demonstrated. Moreover, Miyata and Miyaji et al. developed Au25(Capt)18 NCs exhibiting photodynamic activity toward oral bacteria.

The following sections include discussion of several strategies to improve the 1O2 generation efficiency and PDT activity of Au NCs by using: (i) atomically precise size control, (ii) NIR activation for deep-tissue treatments, (iii) aggregation-induced emission (AIE), (iv) resonance energy transfer (RET), and (v) forming nanocomposites with other nanomaterials for multimodal synergistic therapy.

3.3.1. Atomically precise size control of Au NC (or Ag NC)-based photosensitizers. Au NCs with different sizes and ligands have been used to improve the 1O2 generation efficiency, depending on the size of NCs at the atomic level. For example, Au144(PET)24 NCs have been shown to be less effective compared to Au25(PET)18 since the optical gap of the former (~0.9 eV) is smaller than the energy of 1O2 (0.97 eV). Talavera et al. examined the excited state properties of tetraoctyl ammonium (TOA) capped Au10–12 (SG) NCs in toluene. The Au NCs exhibited a relatively long-lived triplet excited state, which originated from the ligand to the charge transfer state between the SG shell and the Au core. The quantum yields of 1O2 (~0.13) and the triplet excited state (~0.15) were evaluated by 1,3-diphenylisobenzofuran (DPBF, a 1O2-selective probe) and triplet–triplet absorption spectra. Ho-Wu et al. investigated the size effect of Au NCs on the 1O2-generation efficiency from Au25(SG)18, Au14(C6)60, and plasmonic Au NPs (~40 nm). The Au144 NCs had a much higher 1O2 generation rate (1O2/NC/min ~3.2) than Au25 NCs (~0.077). The following order for the 1O2 generation rate was reported: Au144 > Au25 > Au NPs (Fig. 18A).

Table 2 Singlet oxygen (1O2) detection methods

<table>
<thead>
<tr>
<th>Method</th>
<th>Probe</th>
<th>Detection</th>
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<tbody>
<tr>
<td>Phosphorescence</td>
<td>1O2</td>
<td>Phosphorescence of 1O2 at 1270 nm</td>
</tr>
<tr>
<td>UV-vis spectroscopy</td>
<td>9,10-Diphenylanthracene (DPA)</td>
<td>Decrease peak at 355 nm</td>
</tr>
<tr>
<td>UV-vis spectroscopy</td>
<td>9,10-Anthracenediy-bis(methylene) dimalonic acid (ABDA)</td>
<td>Decrease peak at 382 nm</td>
</tr>
<tr>
<td>UV-vis spectroscopy</td>
<td>9,10-Anthracenedipropionic acid (ADPA)</td>
<td>Decrease peak at 400 nm</td>
</tr>
<tr>
<td>Fluorescence</td>
<td>Anthracene-9,10-bisethanesulfonic acid (AES)</td>
<td>Decrease peak at 360, 378 and 400 nm</td>
</tr>
<tr>
<td>UV-vis spectroscopy</td>
<td>1,3-Diphenyl-isobenzofuran (DPBF)</td>
<td>Decrease peak at 410 nm</td>
</tr>
<tr>
<td>Fluorescence</td>
<td>p-Nitrosodimethylaniline imidazole (RNO method)</td>
<td>Decrease peak at 440 nm</td>
</tr>
<tr>
<td>Fluorescence</td>
<td>Single oxygen sensor Green (SOSG)</td>
<td>Increased fluorescence at 450 nm</td>
</tr>
<tr>
<td>ESR</td>
<td>Methotrexate (MTX)</td>
<td>Change in enhanced permeability and ESR spectrum</td>
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The superior $^1\text{O}_2$ generation with Au$_{144}$ was attributed to the combination of a high absorption cross section relative to the volume and a high triplet excited state population via core-localized plasmon resonance.

Compared to the Au NC-based photosensitizers, reports on Ag NC-based photosensitizers are limited due to their low photostability. Yu et al. examined bovine serum albumin (BSA)-protected Ag$_{13}$ NC as a photosensitizer. Due to a large population in the triplet state, the BSA-Ag$_{13}$ NCs showed high $^1\text{O}_2$ generation quantum efficiency (~1.26 using rose bengal as the standard). Owing to the good cellular uptake and high $^1\text{O}_2$ generation efficiency, the BSA-Ag$_{13}$ can effectively kill the MCF-7 breast cancer cells following uptake and white light treatment (Fig. 19B). Tominaga and Kawasaki et al. demonstrated that Ag$_7$(MBISA)$_6$ (MBISA = 2-mercapto-5-benzimidazolesulfonic acid sodium salt) NCs also generate $^1\text{O}_2$ with high efficiency under white light irradiation. The following order was suggested for the $^1\text{O}_2$ generation efficiency on different Ag NCs: Ag$_7$(MBISA)$_6 >$ BSA-Ag$_{13} >$ Ag$_{75}$(SG)$_{40} >$ Ag$_{35}$(SG)$_{18} >$ BSA-Ag$_8$ (not detected).

3.3.2. NIR activation of Au NC photosensitizers for deep-tissue treatments. NIR light (700–950 nm, also named the NIR-I window) has much deeper tissue penetration compared to visible light for cancer PDT. Ho-Wu et al. utilized two-photon excited Au$_{13}$SR$_{14}$ from a pulsed 800 nm laser to generate $^1\text{O}_2$ and trigger PDT. The two-photon NIR-triggered PDT was more effective in live cells compared to the one-photon excitation. More recently, the “NIR-II” window (1000–1700 nm) has also been attracting interest because of the decreased light scattering by tissues, lower background autofluorescence, and deeper tissue penetration. Vankayala et al. reported a mercaptopundecanoic acid (MUA)-protected Au NCs conjugated transactivator of transcription (TAT) peptide (peptide sequence: N-GRKKRRQRRR-C). The TAT peptide–Au NCs not only produced $^1\text{O}_2$ under NIR-II light (980 nm), but also acted as DNA nano-cargo to achieve high gene transfection efficiencies.
strategy for NCs of AuAg alloy to enhance the $^{1}\text{O}_2$ generation. Enhancement in both light emission and $^{1}\text{O}_2$ generation by reducing the singlet acceptor). Chromophore (the donor) to another chromophore (the donor chromophore and the absorption of the acceptor chromophore, (ii) small distance between the donor and the acceptor chromophore, and (iii) parallel orientations of the donor – MB nanocomposite exhibited good efficacy of the nanocomposite in animal experiments and tissue section analysis demonstrated the ability to inhibit tumor growth without damaging the major organs. The therapeutic efficacy of the nanocomposite in animal experiments and tissue section analysis demonstrated the ability to inhibit tumor growth without damaging the major organs.

3.3.3. Au NC-based photosensitizers for hypoxic environment. Photosensitizers with Type II mechanism (i.e., $^{1}\text{O}_2$ generation) are the majority for PDT applications; however, their therapeutic efficacy is dependent on the oxygen level in the targeted microenvironment. Meanwhile, the native microenvironment of many tumors is low in oxygen (i.e., hypoxic), since the rapid tumor growth outpaces the oxygen supply. As a result,
the PDT efficiency is reduced in such hypoxic environments.\textsuperscript{119} Several approaches for solving this challenge have been proposed:\textsuperscript{120} (1) \textit{in situ} generation of O\textsubscript{2} \textsuperscript{-}, (2) O\textsubscript{2} independent PDT (Type I PDT), and (3) multimodal synergistic therapy. In the first approach (\textit{in situ} O\textsubscript{2} generation), H\textsubscript{2}O\textsubscript{2} is catalytically decomposed to form O\textsubscript{2} and water. In cancer cells, the levels of H\textsubscript{2}O\textsubscript{2} are usually larger than in healthy cells.\textsuperscript{122} With this in mind, Chen \textit{et al.} constructed a self-supply O\textsubscript{2} system for PDT against hypoxic tumors using the nanocomposite of AuNC@HSA/CAT.\textsuperscript{128} The AuNC@HSA/CAT could self-supply O\textsubscript{2} for PDT as a consequence of its intrinsic catalase-like activity to decompose H\textsubscript{2}O\textsubscript{2} and form O\textsubscript{2} (Fig. 20A). This resulted in greatly improved O\textsubscript{2} generation and increased PDT efficacy for hypoxic cancer cells (Fig. 20B and C). Liu \textit{et al.} also described self-supplying O\textsubscript{2} through catalase-like activity of PAMAM dendrimer-encapsulated Au NCs (AuNCs–NH\textsubscript{2}) for hypoxic tumor PDT.\textsuperscript{122} The AuNCs–NH\textsubscript{2} exhibited catalase-like activity in the physiological pH range (i.e., 4.8–7.4), hence decomposed H\textsubscript{2}O\textsubscript{2} to self-supply O\textsubscript{2}. The proposed mechanism was that in acidic solutions the protonated tertiary amine groups of dendrimers facilitate the pre-adsorption of OH on the metal surface, triggering the catalase-like reaction.

To overcome the hypoxia limitation in Type II PDT, an O\textsubscript{2} independent Type I PDT has recently been developed, as these photosensitizers produce HO\textsuperscript{-}.\textsuperscript{126,128} HO\textsuperscript{-} is a stronger oxidant than O\textsubscript{2} \textsuperscript{-} and could react with virtually any biological molecule, including DNA, proteins, lipids, and carbohydrates (Fig. 21).\textsuperscript{123} Yang \textit{et al.} fabricated Au\textsubscript{25} NCs anchored on black anatase TiO\textsubscript{2} \textsubscript{-} \textsubscript{x} nanotubes (Au\textsubscript{25}/B-TiO\textsubscript{2} \textsubscript{-} \textsubscript{x} NTs) for Type I PDT.\textsuperscript{124} Under NIR light, the Au\textsubscript{25}/B-TiO\textsubscript{2} \textsubscript{-} \textsubscript{x} NTs could achieve separation of the e\textsuperscript{-}–h\textsuperscript{+} pair, leading to the production of HO\textsuperscript{-} and O\textsubscript{2} \textsuperscript{-} radicals. \textit{In vivo} and \textit{in vitro} experiments demonstrated an obviously enhanced PDT effect of Au\textsubscript{25}/B-TiO\textsubscript{2} \textsubscript{-} \textsubscript{x} NTs. Cheng \textit{et al.} also fabricated titanium dioxide NP-Au NCs–graphene heterogeneous nanocomposites to induce a large production of HO\textsuperscript{-} and O\textsubscript{2} \textsuperscript{-} radicals, which could serve as Type I PDT for B16F1 melanoma cells.\textsuperscript{125}

3.3.6. Multimodal synergistic \textit{in vivo} therapy using Au NC-based photosensitizer. The efficacy of PDT using Au NC-based photosensitizers can be enhanced by combination with other therapeutic modalities, such as conjugates with organic photosensitizers, chemo-T, RT, PTT, and targeted imaging. In this context, we focused on Au NC-based multimodal synergistic therapies, particularly PTT,\textsuperscript{119,126–128} RT, chemo-T,\textsuperscript{129–131} and conjugates with organic photosensitizers.\textsuperscript{112,113}

He \textit{et al.} assembled Au\textsubscript{25}(Capt)\textsubscript{18} onto mesoporous silica-coated Nd\textsuperscript{3+}–sensitized up-conversion NPs (UCNPs@MS-Au\textsubscript{25}) to form a multifunctional platform for photoacoustic (PA) imaging and dual phototherapy (PTT/PDT).\textsuperscript{126} The UCNPs@MS-Au\textsubscript{25} exhibited considerable PTT effect, by combining its intrinsic PDT effect to further enhance \textit{in vivo} tumor inhibition under 808 nm laser.

Yang \textit{et al.} fabricated metal–organic frameworks (MOFs) of Fe\textsubscript{3}O\textsubscript{4}/ZIF-8-Au\textsubscript{25} (IZA) nanospheres.\textsuperscript{137} The IZA nanospheres not only exhibited PTT effects upon NIR light irradiation to effectively kill tumor cells, but also proved useful in targeting and magnetic resonance imaging (MRI). The Au\textsubscript{25}(Capt)\textsubscript{18} in IZA nanospheres produced highly reactive O\textsubscript{2} \textsuperscript{1} to cause PDT as well as PTT effects under NIR light irradiation. As a result, the IZA nanospheres exhibited synergistic therapeutic effect superior to any single therapy both \textit{in vitro} and \textit{in vivo} (Fig. 22A). Tumor-bearing mice treated only with PDT (Au\textsubscript{25}(Capt)\textsubscript{18} NCs under NIR) or PTT (IZ under NIR) displayed much less inhibition in tumor growth than those treated by the multimodal synergistic therapy with combined PDT, PTT, and MRI approaches (Fig. 22A).

Han \textit{et al.} described a BSA-stabilized multifunctional theragnostic nanoplatform, gadolinium oxide–Au NCs–indocyanine green (ICG) hybrid (Gd\textsubscript{2}O\textsubscript{3}–AuNCs–ICG). The Gd\textsubscript{2}O\textsubscript{3}–AuNCs–ICG nanocomposites demonstrated excellent \textit{in vivo} triple imaging capability for NIR fluorescence (NIRF), magnetic

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**Fig. 20** (A) Synthesis and characterization of AuNC@HSA/CAT NPs. (B) Oxygen generation in H\textsubscript{2}O\textsubscript{2} solutions (100 μM) incubated with AuNC@HSA or AuNC@HSA/CAT. (C) The generation of O\textsubscript{2} determined by the increased SOSG fluorescence from AuNC@HSA or AuNC@HSA/CAT, with or without addition of H\textsubscript{2}O\textsubscript{2}. Panels (a)–(c) are reproduced with permission from ref. 98. Copyright 2018 Springer Nature.

**Fig. 21** Identifying pathogen vulnerability using redox perturbation with different ROS. Compared to singlet oxygen and hydroxyl radicals, superoxide and peroxide radicals have much longer diffusion lengths and half-lives in the cellular environment (red circles, not to scale). This figure is reproduced with permission from ref. 123. Copyright 2019 Springer Nature.
resonance, and computed topography, as well as efficacy in combined PDT and PTT (Fig. 22B).\textsuperscript{136} Furthermore, Cui et al. constructed a nanocomposite of BSA-Au NCs–ICG to allow multimodal synergistic therapy, namely dual-modal NIRF/PA imaging, cancer treatment by synergistic action of PDT/PTT, and real-time therapeutic monitoring based on FRET.\textsuperscript{135} Multimodal synergistic therapy using this nanocomposite resulted in 95\% cancer cell death and complete tumor disappearance. Lv et al. produced a core/shell structured nanocomposite by conjugating Au\textsubscript{25}(SR)\textsubscript{18} NC (a PDT/PTT agent), the pH/temperature-responsive polymer P(NIPam-MAA), and the antitumor antibiotic doxorubicin (DOX) onto the surface of mesoporous silica-coated core–shell up-conversion NPs (UCNPs) (Fig. 22C).\textsuperscript{139} The controlled DOX release in the cancer cells was triggered by a high temperature from the PTT under NIR irradiation and low pH. The combined PDT, PTT, and pH/temperature-responsive chemo-T could significantly improve the therapeutic efficacy, as confirmed by both in vitro and in vivo assays. Moreover, the nanocomposite showed dual-modal imaging properties (computer tomography and up-conversion luminescence) when irradiated with 980 nm NIR light, demonstrating the potential for imaging-guided therapy.

In addition to their highly efficient \( ^1 \text{O}_2 \) generation, the physiological properties of Au NCs in biological environments must be further evaluated in terms of the toxicity, renal clearance, and stability for in vivo PDT applications.\textsuperscript{136-147} As for the renal clearance, NPs larger than 6 nm are eliminated from the blood stream by the reticuloendothelial system (RES; liver, 

spleen, etc.), while those smaller than 6 nm can be eliminated via the kidneys by passing through the glomerular capillary wall (GCW; i.e., the kidney filtration) (Fig. 23A).\textsuperscript{146} The longer circulation time of Au\textsubscript{10,11} NCs in the blood compared to Au\textsubscript{25} NCs allowed their accumulation in cancerous tissues through EPR (Fig. 23B).\textsuperscript{145} The unique renal clearance behavior of the sub-nm Au NCs was explained by the filtration mechanism of GCW like-size separation utilizing size exclusive chromatography. Compared to the neutral NCs, it is more difficult for the negatively charged NCs to cross the GCW, since the GCW is negatively charged. Zwitterionic or non-ionic surface ligands (e.g., SG or PEG) are desirable for the efficient urinary elimination of small NCs from the body.\textsuperscript{138,142,144,147}

4. Conclusions and perspectives

In summary, we have reviewed the recent advances in the fields of photo/electrocatalysis and photosensitization using metal NCs, and their related green energy and medical applications. Owing to the unique properties of metal NCs, including their attractive photo-/electro-functional properties, as well as recent
developments in their atomically precise synthesis, these functionalized moieties show excellent potential for the abovementioned applications.

Considering the use of metal NCs in photo/electrocatalysis and photosensitization in green chemistry, we summarized several strategies to improve conversion efficiency and selectivity of the following materials: (i) photocatalysts for water splitting and fuel cells, (ii) photocatalysts for water splitting, and (iii) solar cells. As described in this review, atomically precise control can contribute to the improvement of these materials. However, considering the practical applications, further developments are expected for simple and precise synthesis and isolation of metal NCs, as well as their loading. In addition, the following studies are envisaged for the improvement of these materials.

(1) Deeper understanding of the electronic structures of the loaded mono-metal and alloy NCs is essential for all of the applications. The electronic structures of the small NCs strongly depend on the number of the constituent atoms and the types of heteroatoms. Recent studies revealed the relationship between the electronic structures and these factors for the isolated NCs. However, the NCs loaded on the substrate seem to show different electronic structures as a result of the interaction between the NCs and the substrate, especially when the ligands are removed from the NC surface. It is expected that future studies would reveal the electronic structures of the loaded NCs for their application as active site on the substrate. In addition, the method for controlling the heteroatom position in the loaded alloy NCs is also expected to be developed in the future studies.

(2) For the improvement of photocatalysts, an effective method for separating e−h+ pairs should be established. The Fermi energy of the cocatalysts strongly affects the charge transfer of the excited electron. Therefore, the electronic structure of cocatalysts, which effectively facilitates the charge transfer from photocatalysts to cocatalysts must be designed. In addition, although most studies concentrate on the loading of cocatalysts only for the reduction reaction, the future studies should conduct the co-loading of the cocatalysts for both reduction and oxidation reactions. It is also important to deepen the understanding on the reaction intermediates, the mechanism, and the methods for suppression of the reverse reactions.

(3) Regarding the photosensitization application for solar cells, NCs, which are able to absorb the visible and infrared light to effectively use the sunlight, should be fabricated and utilized. The stability of NCs should also be improved for this application. In addition, improvements in the energy-conversion efficiency must be made by combining several systems.

Considering photosensitization of metal NCs for therapeutic applications, we summarized several strategies to improve the \(1\text{O}_2\) generation efficiency and PDT activity of metal NCs: (i) atomically precise size control, (ii) NIR activation for deep-tissue treatments, (iii) AIE, (iv) RET, and (v) forming nanocomposites with other nanomaterials for multimodal synergetic therapy. Despite these efforts, further research into the following aspects is necessary:

(4) By adjusting \(n\) and \(m\), the discrete electronic/geometrical structures of metals, and thereby their photosensitizing capability, can be controlled. Extensive studies on the syntheses of organo-soluble metal NCs with atomic precision have been described, while the synthesis of water-soluble metal NCs remains undeveloped. Therefore, further research into the synthesis of water-soluble metal NCs is required for therapeutic applications. Ligand exchange method using organo-soluble metal NCs may be useful for the synthesis of water-soluble metal NCs and alloy metal NCs would be valuable for the exploration of their photosensitizing capability.

(5) A photosensitizer is primarily required to possess the following properties:\textsuperscript{148} (1) constant composition, (2) simplicity of synthesis, (3) lack of toxicity in the dark, (4) target specificity, (5) having a triplet state energy higher than 0.97 eV (the energy of single state oxygen, \(1\text{O}_2\)), (6) photostability, (7) high triplet state quantum yield, (8) fast clearance from the body, and (9) minimal self-aggregation. Metal NCs satisfy the abovementioned requirements; however, their \(1\text{O}_2\) generation efficiencies are lower than those of organic photosensitizers. The high triplet state quantum yield and inhibition of energy dissipation through non-radiative pathways would be the key factors to improve the \(1\text{O}_2\) generation efficiency of the metal NCs. Since water-soluble thiolated metal NCs have multifunctionality such as imaging, drug delivery, and multiple sensitizing capability (photo-sensitizer, radiosensitizer, and sonosensitizer), it would be useful as a multipurpose ROS mediated nanomedicine\textsuperscript{20,149,150}

(6) Regarding the \textit{in vivo} PDT applications, the use of specific ligands for metal NCs increases the selectivity toward the target (e.g., cancerous tumors), an approach called “active targeting”. Targeting ligands on metal NCs are responsible for the selective interaction between the metal NCs and specific receptors on the surface of tumor cells, but not healthy cells, thus increasing the internalization of the NCs. Various types of surface modification of metal NCs, such as aptamer, specific peptides, glucose, DNA, and folic acid have been developed for efficient delivery of the metal NCs to the target (e.g., cancerous tumors) for \textit{in vivo} applications.\textsuperscript{123–126} Further studies involving active targeting of metal NCs would be important for \textit{in vivo} PDT applications.

Conflicts of interest

There are no conflicts to declare.

Authorship contributions

\textit{Photo/electrocatalysis of metal nanoclusters in green energy applications:} Tokuhisa Kawakaki and Yuichi Negishi. \textit{Photosensitization of metal nanoclusters for therapeutic applications:} Hideya Kawasaki. \textit{Introduction and conclusion:} equal contribution by all authors.

Notes and references


