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Efficient NO₂ sensing performance of a low-cost nanostructured sensor derived from molybdenite concentrate†

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Accumulation of industrial solid waste necessitates the development of utilization processes and technologies to reduce their negative environmental impact. Herein, molybdenite concentrate from the mining-metallurgy industry is systematically characterized as a valuable starting material for the fabrication of an efficient and low-cost nanostructured gas sensor. Few-layer MoS₂ is obtained from molybdenite concentrate by liquid nitrogen exfoliation and deposited on different substrates by spin coating and drop casting. It is found that spin coating is advantageous over drop casting in fabricating a homogeneous and dense few-layer MoS₂ film. The charge-transfer-based sensing performance of the fabricated few-layer MoS₂ film is investigated upon exposure to NO₂ at different temperatures (50, 100, and 120 °C). At an optimized temperature of 120 °C, a faster recovery is achieved, and the fabricated device exhibits 28, 38, and 44% sensitivity to 10, 50, and 100 ppm NO₂, respectively, making it suitable for practical applications. Furthermore, the adsorption affinity of NO₂ to the predominant (002) crystallographic plane of MoS₂ is estimated from the distribution of field density and the calculated differential adsorption energies. According to the molecular modeling data, NO₂ in the Ar/NO₂ mixture has better interaction ($dE_{ad}/dN_{NO_2} = 4.77 \text{ kcal mol}^{-1}$) with the few-layer MoS₂ surface than individual NO₂ ($dE_{ad}/dN_{NO_2} = 2.57 \text{ kcal mol}^{-1}$), and van der Waals interaction ($\approx 14 \text{ kcal mol}^{-1}$) is the main adsorption force compared to the relatively weaker electrostatic interaction ($< 1 \text{ kcal mol}^{-1}$). This work demonstrates a straightforward approach not only for the conversion of molybdenite concentrate into an efficient and low-cost nanostructured gas sensor but also for the reduction of the negative impact of accumulated molybdenum concentrate on the environment and human health.

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1. Introduction

Rapid industrialization and urbanization have led to severe air pollution that has increasing negative impacts on human health and the environment alike. The United States Environmental Protection Agency (EPA) has even started an AirNow program that receives, manages and shares the data related to the real-time global air quality.¹ This program also forecasts the local air quality and supports actions to reduce

exposure to air pollution. Along with other hazardous air pollutants like sulfur dioxide (SO₂), carbon dioxide (CO₂), and nitrogen monoxide (NO), nitrogen dioxide (NO₂) is emitted into the atmosphere as a product of combustion processes (vehicles, industry, heating, shipping, etc.)² and explosives.³ NO₂ is the main source of tropospheric ozone and PM_{2.5} and forms nitrous acid (acid rain) upon reaction with water, which is harmful to ecosystems. Short- and long-term exposure to NO₂ can have adverse respiratory and cardiovascular health effects and even mortality.⁴ According to the current World Health Organization's guideline, the NO₂ annual mean value has been set to 40 micrograms per cubic meter in order to protect public health.⁵ It is a prerequisite to constantly monitor the air quality using highly sensitive and fast gas sensing systems to comply with laws and regulations.

In recent years, two-dimensional (2D) transition metal dichalcogenides, such as MoS₂, MoSe₂, WS₂, WSe₂, SnS₂, etc., have gained enormous interest in sensing applications with low-power consumption due to their exceptional electrical,

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chemical, and physical properties.⁶ Among transition metal dichalcogenides, MoS₂ has recently emerged as a potential candidate for NO₂ sensing application due to its atomic-scale thickness, high surface to volume ratio, tunable band gap, and high on/off switching ratio.

A one-dimensional MoS₂ nanowire network (NW) fabricated by using controlled turbulent vapor flow from the chemical transport reaction showed about 2-fold enhanced sensitivity for NO₂ at 60 °C with a low detection limit (4.6 ppb) and enhanced sensitivity and selectivity towards NO₂ because of a combination of abundant active edge sites and a large surface area and tuning of the potential barrier at the intersections of nanowires during adsorption/desorption.⁷ Vertically aligned MoS₂ fabricated by using the rapid sulfurization method of CVD process showed about 5-fold enhanced sensitivity to NO₂ compared to horizontally aligned MoS₂ due to the high density of the exposed edge sites.⁸ Hierarchical MoS₂ microsphere hollow structures synthesized by a hydrothermal method showed enhanced NO₂ detection owing to the improved active edge sites.⁹ Some researchers have used a UV light source to improve the sensing performance of a MoS₂ sensor at room temperature. For instance, a photoactivated mixed in-plane and edge-enriched p-type MoS₂ flake-based sensor demonstrated a fast response with a good sensitivity of ~10.36% for 10 ppm NO₂ without and with complete recovery at room temperature and under UV light irradiation, respectively.¹⁰ Under photo excitation, MoS₂ exhibited an enhanced sensitivity with ultrafast response time of ~29 s and excellent recovery to NO₂ (100 ppm) at room temperature due to charge perturbation on the surface of the sensing layer during the NO₂/MoS₂ interaction under light irradiation.¹¹ A high-performance gas sensor prepared using a hybrid of temperature-assisted sulfur vacancy within the edge-oriented vertically aligned MoS₂ and reduced graphene oxide exhibited a significantly enhanced response to NO₂ with fast response and complete recovery due to the controlled electrical/chemical sensitization level of MoS₂ through controllable vacancy and interface engineering.¹²

However, it is necessary to develop efficient and low-cost sensing materials based on MoS₂ to find practical applications in NO₂ sensing. One of the plausible ways to develop efficient and low-cost sensing materials is to utilize industrial solid wastes. In this work, we aim to characterize molybdenite concentrate from the mining-metallurgy industry and to fabricate an efficient and low-cost nanostructured sensor by utilizing molybdenite concentrate. Simple liquid nitrogen exfoliation is used to exfoliate few-layer MoS₂ and to remove any impurities from molybdenite concentrate. The gas sensing behavior of the fabricated sensor by spin coating of few-layer MoS₂ to the ppm concentration of NO₂ is investigated. Furthermore, the adsorption affinity of NO₂ to the predominant (002) crystallographic plane of molybdenite-2H is estimated from the distribution of field density and the calculated differential adsorption energies. This work demonstrates a straightforward approach not only for the conversion of molybdenite concentrate into an efficient and low-cost nanostructured gas sensor

but also for the reduction of the negative impact of accumulated molybdenum concentrate on the environment and human health.

2. Experimental

2.1. Preparation

Molybdenite concentrate (MC) is generally formed during the extraction process of copper and molybdenum from porphyry copper-molybdenum in Almalıy Mining-Metallurgical Complex, Uzbekistan and has been used as the starting material in this study. As shown in Fig. 1, after grinding for 30 min, MC was immersed in liquid nitrogen for 2 h,¹³ and the upper part of the suspension was separated and ground again for 1 h. Afterward, 45 mg of the exfoliated MC was dispersed in 3 mL of isopropyl alcohol under sonication for 2 h. The samples were prepared by spin coating of the MC-based suspension on a SiO₂/Si substrate at 1500 (MC1), 2500 (MC2) and 3500 (MC3) rpm for 30 s. The MC4 and MC5 samples were prepared by spin coating of the MC-based suspension on an n-type silicon substrate at 500 and 1500 rpm for 30 s, respectively. The MC6 sample was prepared by drop casting of 30 μL of the MC-based suspension on a metallic IDC-patterned SiO₂/Si substrate.

The conductivity of the fabricated sensing device is very important in examining its sensing response. Previously, it has been reported that few-layer MoS₂ has higher conductivity than its monolayer and multilayer counterparts.¹⁴ Therefore, we have used few-layer MoS₂ to fabricate a sensing device that shows a higher sensing response stability while maintaining a high surface to volume ratio. The fabricated sensing device was annealed at 300 °C for 1 h under a N₂ atmosphere to reduce the unintentional barrier at the metal/semiconductor interfaces.¹⁵ Au was used to create an Ohmic contact with few-layer MoS₂.^{16,17} The current-voltage characteristic curve of the fabricated sensing device at room temperature is shown Fig. S1 in the ESI,† indicating the Ohmic behavior of the fabricated sensing device.

2.2. Characterization

The X-ray diffraction (XRD) patterns were recorded on a MiniflexII (Rigaku) diffractometer using Cu Kα radiation ($\lambda = 0.15418$ nm) in the 2θ scan range from 10 to 80°. The morphology of the samples was examined using a JSM-7600F field-emission-type scanning electron microscope (JEOL). The energy-dispersive X-ray spectroscopy (EDX) element mapping images were obtained using a spectrometer attached to the SEM. The crystallinity and nanostructure of the samples were analyzed using a JEM-2100F HK high-resolution analytical scanning transmission electron microscope (JEOL). The ultraviolet-visible (UV-vis) diffuse reflectance spectrum of the samples was recorded on a UV-3600 UV-vis-NIR spectrophotometer (Shimadzu) equipped with an integrating sphere, and BaSO₄ was used as the reference. The Brunauer–Emmett–Teller (BET) specific surface area of the samples was determined from the N₂ adsorption–



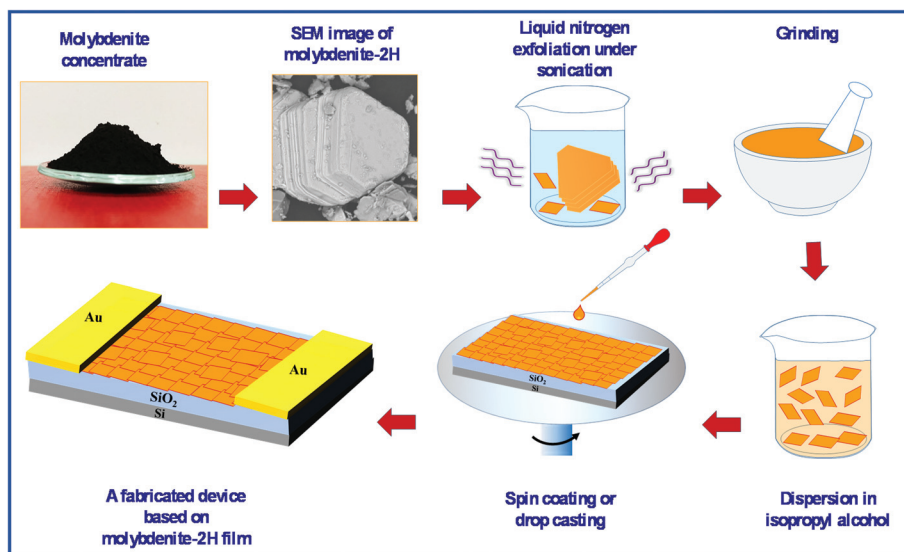


Fig. 1 Preparation procedure of a low-cost nanostructured sensor from molybdenite concentrate.

desorption isotherm measured using a BELSORP-mini instrument (BEL Japan, Inc.) at 77 K after degassing the sample at 120 °C for 12 h. The surface chemical states of elements were analyzed by X-ray photoelectron spectroscopy (JPS-9010MC, JEOL) using non-monochromated Mg K α radiation (1253.6 eV).

2.3. Gas-sensing measurement

Based on the characterization results, only the MC6 sample was selected for gas-sensing measurement. Its gas-sensing behavior was examined in a gas-sensing chamber using the desired concentrations (10, 50 and 100 ppm) of NO₂ after injecting a mixture of 5% NO₂ (99.99%, HPS Gases Limited)

and 95% Ar (99.99%, HPS Gases Limited). An external heating filament was used to increase the temperature of the device to 120 °C. The change in the current of the device at 3.0 V bias voltage was measured using a Keithley 4200-SCS semiconductor characterization system (Tektronix).

3. Results and discussion

Fig. 2a shows the XRD pattern of molybdenite concentrate. The major reflections in the XRD pattern are readily indexed to the hexagonal molybdenite-2H with the space group *P63/mmc* (194) (ICDD PDF# 06-0097), while the minor reflections

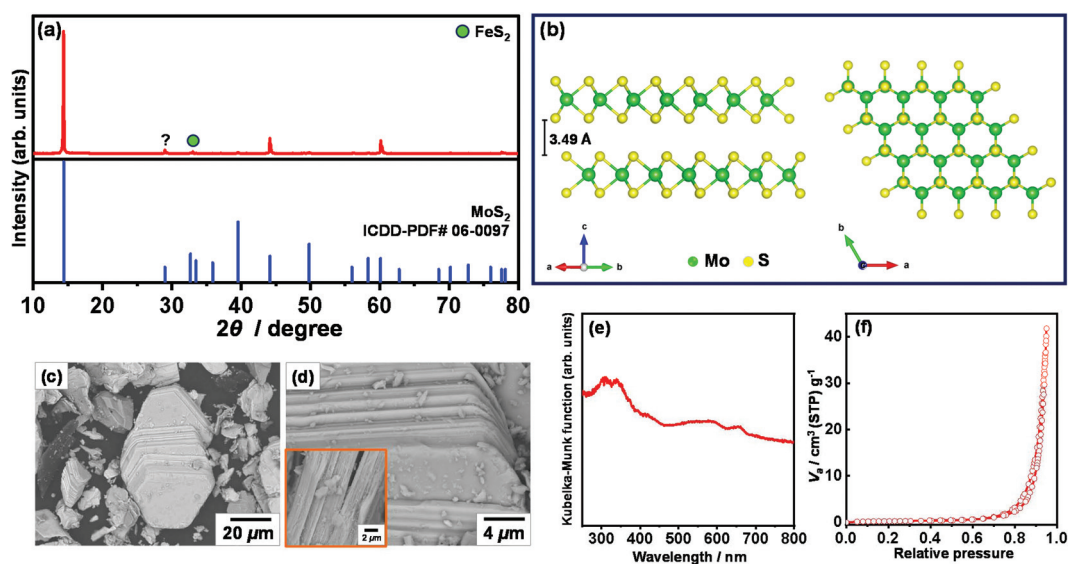


Fig. 2 (a) XRD pattern, (b) side and top views of the crystal structure of molybdenite (2H polytype), (c and d) BSE-SEM images, (e) UV-vis DRS spectrum, and (f) N₂ adsorption–desorption isotherm of molybdenite concentrate.



can be assigned to the triclinic FeS₂ with the space group P1 (ICDD PDF#71-1680) and an unidentified phase. The three main reflections observed at $2\theta = 14.39, 44.16, \text{ and } 60.11^\circ$ correspond to the (002), (006), and (008) atomic planes of crystalline molybdenite-2H, indicating a stacking layer along the *c* direction. These atomic planes are stacked in a laminar crystalline structure (common 2H polytype), as shown in Fig. 2b. These individual two-dimensional S–Mo–S crystalline layers are bonded by the weak van der Waals interactions and are capable of slipping over each other and generate low-energy and hydrophobic surfaces (basal planes). The in-plane S–Mo bonds are strong Mo–S covalent bonds, which create high-energy and hydrophilic sites known as edges when broken.^{18,19} Therefore, molybdenite particles exhibit anisotropic surface properties and are naturally hydrophobic due to the surface content of S atoms.^{19,20} Fig. 2c and d show the low- and high-magnification BSE-SEM images of molybdenite concentrate. Along with other irregular particles, truncated triangular particles, formed by the face-to-face stacking of the two-dimensional layers of molybdenite-2H, with a lateral size of <40 μm can be observed. Clearly, the surface of truncated triangular particles also contains some irregular particles of molybdenite-2H.

Fig. 2e shows the UV-vis diffuse reflectance spectrum of molybdenite concentrate. As shown, molybdenite concentrate exhibits a strong light absorption in the wavelength range of 200–800 nm. An absorption edge of molybdenite concentrate is observed at approximately 435 nm, corresponding to an optical indirect band-gap energy (E_g) of 2.86 eV. Clearly, the indirect band-gap energy of molybdenite concentrate is much lower than the reported value for a bulk MoS₂ crystal ($E_g = 1.29 \text{ eV}$)²¹ due to their chemical and structural differences. It is known that the band-gap energy of MoS₂ can easily be tuned by changing the number of layers²² and by applying a mechanical strain,²³ and a direct to indirect band-gap transition is also observed. Obviously, four characteristic peaks with lower intensity at about 290–300 nm and 550–685 nm are observed in the UV-vis diffuse reflectance spectrum of molybdenite concentrate. Presumably, the peaks observed at 290–300 nm may indicate a reduction in the 2H phase of MoS₂^{24,25} or the interband transition between the occupied and unoccupied orbitals,²⁶ while the peaks noted at 550–685 nm correspond to the excitonic interband transitions at the *K* point of the Brillouin zone.²⁷ It must be mentioned that the peaks representing the four excitonic electronic transitions are blue-shifted compared to those of stable dispersion of MoS₂ nanosheets.²⁶ As the specific surface area plays an important role in sensing, the specific surface area was estimated from the N₂ gas adsorption–desorption isotherm shown in Fig. 2f. According to the IUPAC classification,²⁸ the isotherm of molybdenite concentrate is type III, indicating a relatively weak adsorbent–adsorbate interactions on the surface of a nonporous or macroporous solid. The estimated specific surface area of molybdenite concentrate was 2.11 m² g^{−1}.

Various particles were observed in the BSE-SEM images of molybdenite concentrate in addition to the molybdenite-2H

particles. To identify these particles, energy-dispersive X-ray spectroscopy (EDX) analysis was performed. Fig. S2 in the ESI† shows the BSE-SEM images and the corresponding EDX spectra of different particles of molybdenite concentrate. Based on the EDX data shown in Fig. S2a and S2d,† the truncated cubic and plate-like particles can be assigned to FeS₂ and molybdenite-2H, respectively, and no impurities/dopants are detected, implying the high purity of both components. In Fig. S2b,† the irregular inhomogeneous particles accumulated between the truncated cubic crystals of FeS₂ and plate-like particles of molybdenite-2H contain iron, copper, magnesium, aluminum, sulfur, and oxygen elements. Interestingly, the sponge-like quasi-spherical particles contain iron, magnesium, aluminum, oxygen and a small amount of sulfur (Fig. S2c†). According to the EDX data, the chemical composition of molybdenite concentrate (wt%) is 24.54 Mo, 51.41 S, 18.53 O, 3.62 Fe, 0.87 Al, 0.57 Cu, and 0.46 Mg.

The exfoliation in liquid nitrogen and dispersion in isopropanol under ultrasonication yielded a few-layer MoS₂ film. As shown in Fig. S3a in the ESI,† the presence of a strong diffraction peak at $2\theta = 14.3^\circ$ in the XRD pattern of the exfoliated/deposited sample confirms high crystallinity of few-layer MoS₂. Furthermore, Raman spectroscopy was used to estimate the thickness of the MoS₂ film. For a better understanding, the Raman spectrum of few-layer MoS₂ deposited on the sensing device is compared with that of bulk MoS₂. As shown in Fig. S3b,† the Raman spectrum of few-layer MoS₂ deposited on the sensing device shows red and blue shifts as compared to the Raman spectrum of bulk MoS₂. The two vibrational modes positioned at 383.30 and 407.28 cm^{−1} correspond to the in-plane (E_{12g}^1) and out-of-plane (A_{1g}) phonon modes. The difference between the two modes is $\sim 24 \text{ cm}^{-1}$ which ensures that the thickness of the deposited MoS₂ film in the sensing device is about 4 nm.^{29,30}

Molybdenite-2H collected after liquid nitrogen exfoliation was analyzed using a high-resolution analytical scanning transmission electron microscope to gain more insights into the crystal structure, and the results are shown in Fig. 3. The bright-field TEM image in Fig. 3a shows a molybdenite-2H particle with a lateral size of about 7 μm . Clearly, the main part of the molybdenite-2H particle had a darker color than its edge, suggesting that the molybdenite-2H particle is composed of several layers. No clear defects were seen in the lattice image (Fig. 3b), confirming the high crystallinity of exfoliated molybdenite-2H layers. The interplanar spacing in the lattice fringe image was found to be 0.27 nm, corresponding to the (100) atomic plane of hexagonal molybdenite-2H, which is in good agreement with the XRD data. A typical hexagonal lattice structure of molybdenite-2H was identified by indexing the SAED pattern shown in Fig. 3c. The well-ordered diffraction spots in the SAED pattern reveal the single-crystalline nature of the layered structure of molybdenite-2H. In Fig. 3d–f, the scanning transmission electron microscope (STEM) image and the corresponding elemental mapping images indicate the homogeneous distribution of Mo and S atoms (Mo:S ratio = 1.01 : 1.97) over the entire layered structure of molybdenite-2H.



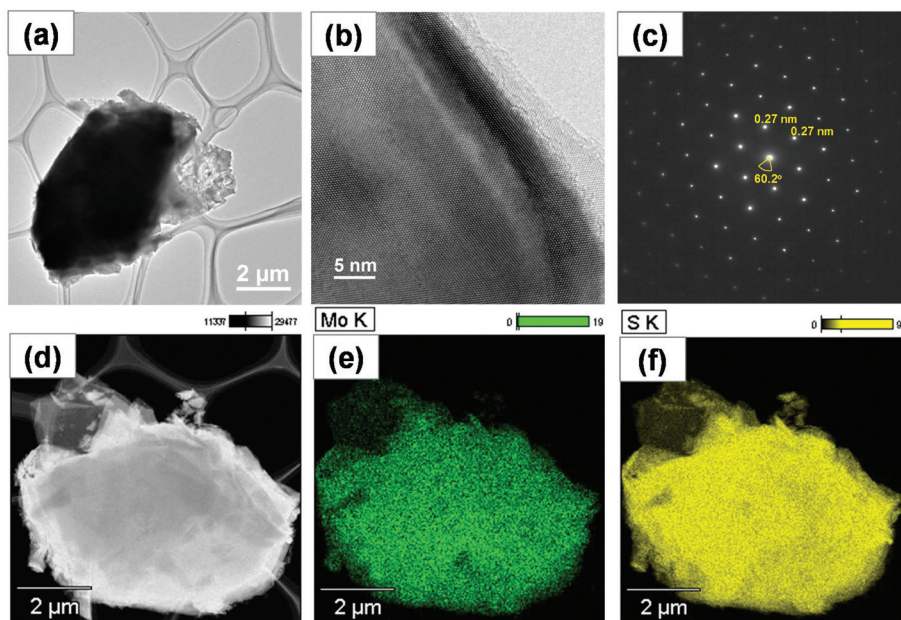


Fig. 3 TEM (a) and HRTEM (b) images, SAED pattern (c), STEM image (d) and the corresponding EDX elemental mapping images of Mo (e) and S (f) of exfoliated molybdenite.

The wide-scan XPS spectra of molybdenite concentrate before and after liquid nitrogen exfoliation are shown in Fig. 4. As shown in Fig. 4a, the XPS spectrum of molybdenite concentrate before exfoliation indicates the presence of Mo, S, O, Fe, Al, Cu, Mg and adventitious C elements. In contrast, the XPS spectrum of molybdenite concentrate after liquid nitrogen exfoliation (Fig. 4b) shows only Mo, S, and adventitious C and O elements, deducing that enriched molybdenite-2H can be produced by liquid nitrogen exfoliation. In Fig. 4c, the high-

resolution XPS spectrum of Mo 3d shows the predominant Mo 3d_{3/2} and Mo 3d_{5/2} peaks at 233.0 eV and 229.9 eV, respectively, which are ascribed to the Mo⁴⁺ components of the molybdenite-2H phase. In Fig. 4d, the high-resolution XPS spectrum of S 2p are deconvoluted into two characteristic S²⁻ components (S 2p_{1/2} and S 2p_{3/2}) of the molybdenite-2H phase at 162.1 eV and 163.31 eV, respectively.³¹

Fig. 5 shows the XRD patterns of the molybdenite-2H films fabricated on SiO₂/Si (MC1–MC3), n-type silicon (MC4

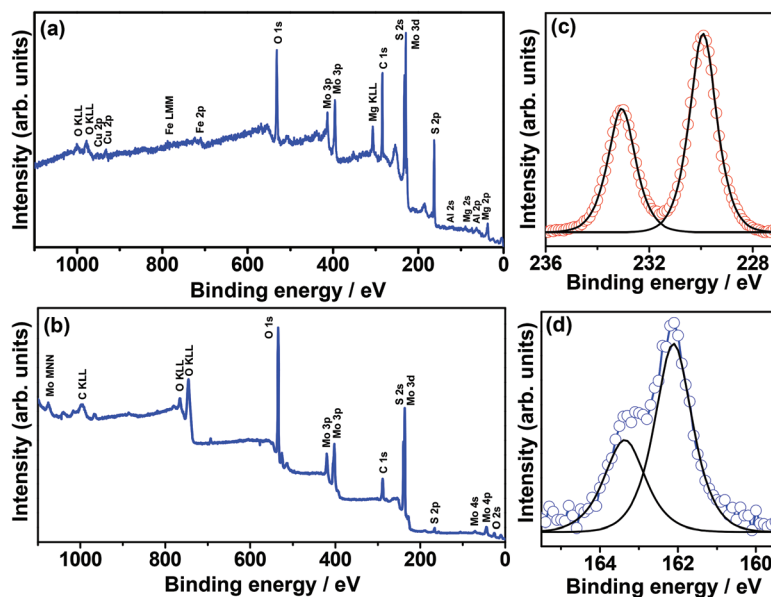


Fig. 4 Wide-scan XPS spectra of molybdenite concentrate before (a) and after (b) liquid nitrogen exfoliation and Mo 3d (c), and S 2p (d) XPS spectra of molybdenite concentrate after liquid nitrogen exfoliation.



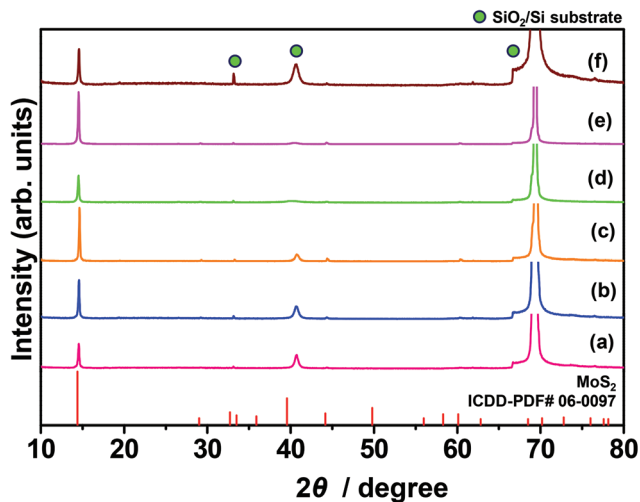


Fig. 5 XRD patterns of (a) MC1, (b) MC2, (c) MC3, (d) MC4, (e) MC5, and (f) MC6 samples.

and MC5) and metallic IDC-patterned SiO₂/Si (MC6) substrates. Upon increasing the speed of spin coating from 1500 rpm to 3500 rpm, the intensity of the 002 reflection in the XRD patterns of molybdenite-2H films fabricated on the SiO₂/Si substrate was increased presumably due to the formation of a thicker film. A similar trend was observed for molybdenite-2H films fabricated by spin coating on the n-type silicon substrate. That is, the intensity of the 002 reflection in the XRD pattern of the molybdenite-2H film fabricated by spin coating at 1500 rpm was higher than that of the molybdenite-2H film fabricated at 500 rpm. The molybdenite-2H film fabricated by drop casting on the metallic IDC-patterned SiO₂/Si substrate showed a slightly lower intensity of the 002 reflection than the other samples.

Fig. 6 shows the SEM images of molybdenite-2H films fabricated on the SiO₂/Si (MC1–MC3), n-type silicon (MC4 and MC5) and metallic IDC-patterned SiO₂/Si (MC6) substrates. It can be seen in Fig. 6a that the exfoliated molybdenite-2H sheets were deposited on the SiO₂/Si substrate by spin coating

at 1500 rpm, and the shape of the deposited molybdenite-2H sheets was visible. However, upon increasing the spin coating speed from 1500 rpm to 3500 rpm, the deposited molybdenite-2H films became denser and less rough, and the shape of the molybdenite-2H sheets was less visible (Fig. 6b and c). Similarly, the molybdenite-2H film fabricated on the n-type silicon substrate by spin coating at 500 rpm was the least dense and rougher, showing clearly the face-to-face deposited irregular molybdenite-2H sheets (Fig. 6d). In contrast, when the spin coating speed was raised to 1500 rpm, the molybdenite-2H film became denser (Fig. 6e), indicating a close relationship between the fabricated film density and the spin coating speed. Fig. 6f shows the SEM image of the molybdenite-2H film fabricated by drop casting on the metallic IDC-patterned SiO₂/Si substrate. It is clear that smaller and thinner molybdenite-2H sheets were homogeneously deposited, from which thicker molybdenite-2H sheets protruded. This indicates that spin coating is advantageous over drop casting in fabricating a homogeneous and dense molybdenite-2H film. Particularly, the MC3 sample shows a homogeneously deposited molybdenite-2H film. In Fig. S4a–c in the ESI,[†] the EDX spectrum and EDX elemental mapping images reveal that Mo and S are homogeneously distributed in the molybdenite-2H film, and no impurities are detected, implying high purity of the deposited film.

In recent years, MoS₂ has established itself as a promising gas sensing material for detecting various environmental pollutants. It offers a higher density of states, a higher value of mobility, and a lower value of interface scattering. Late *et al.*³² studied the sensing performance of single- and few-layer MoS₂ upon exposure to NO₂ gas molecules. After comparing the results from both types of devices, it was observed that few-layer MoS₂ is more suitable for sensing applications due to its excellent sensing response. The improvement in the sensing response of few-layer MoS₂ is due to its different electronic structures and more stable response over time. Hence, these appealing attributes of MoS₂ for gas sensing applications attract much research interest for developing highly sensitive, selective, and stable gas sensors. However, the incomplete

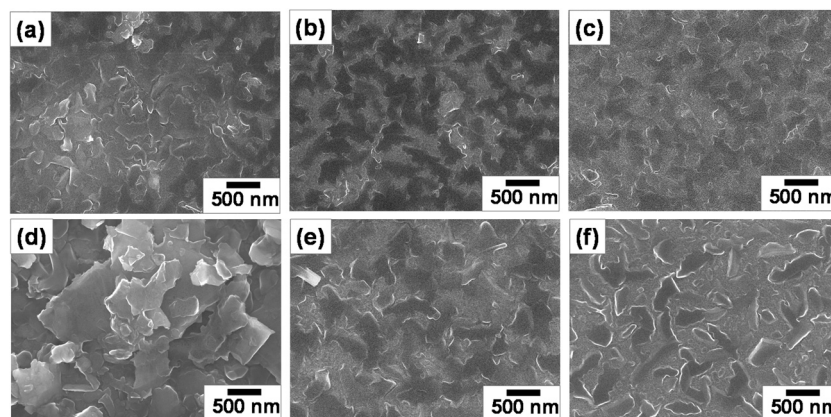


Fig. 6 SEM images of (a) MC1, (b) MC2, (c) MC3, (d) MC4, (e) MC5, and (f) MC6 samples.



recovery is a major bottleneck restricting the performance of MoS₂-based gas sensors at lower temperatures (<100 °C).³³ Therefore, a sufficient external stimulus is the stringent need for the complete recovery of practical gas sensors.

Among various available alternatives, thermal energy is one of the most widely used techniques to improve the sensing performance of the developed device. Therefore, we investigated the sensing performance of the device fabricated using the MC3 sample against different concentrations of NO₂ at different temperatures (50, 100, and 120 °C). The relative dynamic response of the device was calculated as $(\Delta R/R_0) \times 100\% = (R_g - R_0)/R_0 \times 100\%$, where R_g is the resistance of the MC3 sample in the presence of NO₂ and R_0 is the base resistance of the device. As an oxidizing gas, NO₂ traps electrons from the surface of the MC3 sample, resulting in the change of the charge density of the deposited film (Fig. 7a).³³ The variation in the carrier concentration leads to a large increase in the resistivity of the device, and hence, a high value of sensitivity was achieved. At 50 °C, 48% and 67% sensitivities were observed upon exposure to 50 and 100 ppm NO₂ gas molecules, respectively (Fig. 7b). However, the fabricated sensor was not fully recovered due to insufficient thermal energy. A strong binding energy between the MoS₂ film and the NO₂ molecules prevents the complete recovery of the fabricated device at 50 °C.³⁴ Therefore, to make the fabricated device fully recoverable, the temperature of the sensor was raised to 100 °C. At this temperature, 42% and 52% sensitivities were achieved for 50 and 100 ppm NO₂, respectively. Upon increasing the temperature of the device, the recovery was improved appreciably (Fig. 7c). Yet, a full recovery of the gas sensor was still not possible. Thus, the sensing behavior of the fabricated device was

further evaluated at a slightly higher temperature. As expected, a full recovery of the fabricated sensor was achieved at an optimized temperature of 120 °C (Fig. 8). The fabricated device exhibited 28, 38, and 44% sensitivities to 10, 50, and 100 ppm NO₂, respectively. Clearly, the sensitivity of the device increased with the increasing concentration of NO₂. It was noticed that at 120 °C, the sensitivity was slightly lowered as compared to the sensitivity at 50 and 100 °C. The decrement in the sensitivity is primarily due to a higher rate of desorption than adsorption at high temperatures. Besides acceleration of the desorption process, the reduced interaction between the MoS₂ surface and gas molecules at high temperature plays a

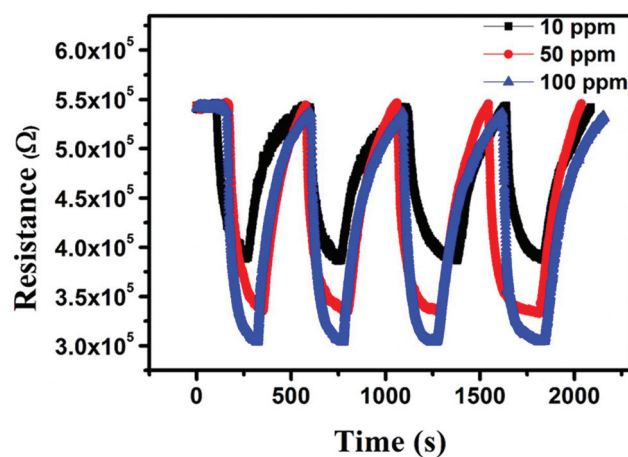


Fig. 8 Sensing performance of the MC6 sample against NO₂ gas with different concentrations at 120 °C.

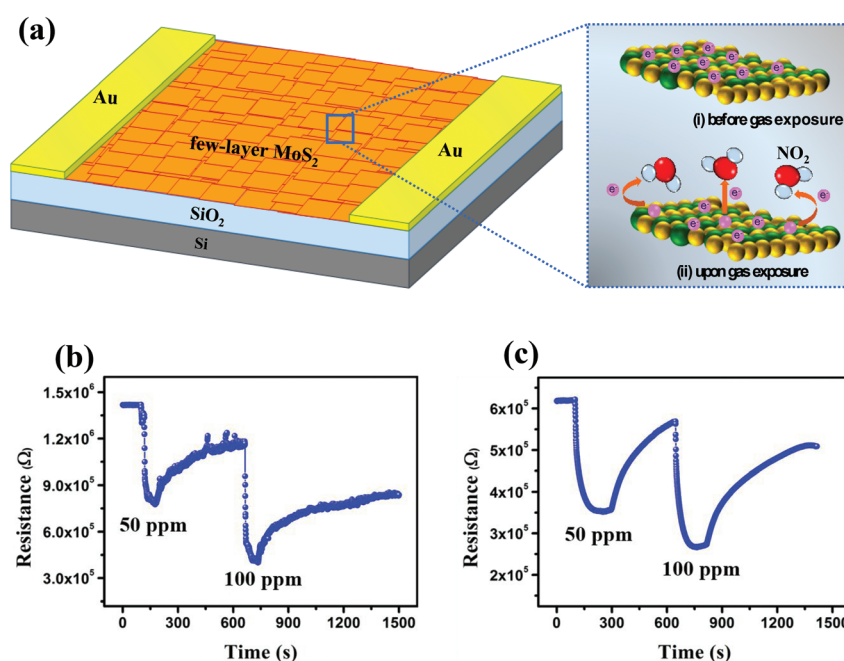


Fig. 7 (a) Schematic illustration of the sensing behavior of the fabricated device based on the charge-transfer mechanism, and the sensing performance of the MC6 sample against 50 and 100 ppm NO₂ at 50 °C (b) and 100 °C (c).



crucial role in determining the sensitivity of the device. In addition, the sensing performance of the as-received molybdenite concentrate without removing the impurities was also evaluated. As shown in Fig. S5,† the sensitivity decreased from a value of ~28% to ~21% in the presence of 10 ppm NO₂ at 120 °C. This decrement in the sensitivity is primarily due to the higher thickness of MoS₂ in molybdenite concentrate and lesser availability of favorable adsorption sites for NO₂. The performance of the fabricated sensing device using few-layer MoS₂ was compared with earlier reports. As shown in Table S1 in the ESI,† the sensing performance of the fabricated device is consistent with the previously reported values, indicating its good sensing response.

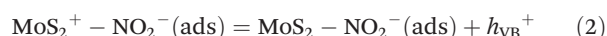
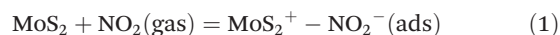
The adsorption of NO₂ onto the predominant (002) surface of molybdenite-2H was modeled by Forcite and Adsorption Locator modules in Accelrys Materials Studio using the experimental crystallographic structural data.³⁵ The adsorption affinities can be estimated from the distribution of field density (Fig. 9a) and the calculated differential adsorption energies (dE_{ad}/dN_{NO_2} and dE_{ad}/dN_{Ar}) which are the energies required for removing the sorbate of a particular component. According to the estimated differential adsorption energies, NO₂ can form comparatively similar substrate–adsorbate configurations at different surfaces with a slightly preferential adsorption on the (002) surface (Fig. 9a). As the sensing behavior of the molybdenite-2H film was tested using a 5% NO₂ and 95% Ar mixture, the dE_{ad}/dN_i values were estimated for the adsorption of the NO₂/Ar mixture onto different surfaces of molybdenite-2H:

NO₂: -4.77 kcal mol⁻¹ for (002) > -3.43 kcal mol⁻¹ for (006) > -3.35 kcal mol⁻¹ for (008)

Ar: -1.82 kcal mol⁻¹ for (008) > -1.15 kcal mol⁻¹ for (006) > -0.18 kcal mol⁻¹ for (002).

It is clear that NO₂ in the Ar/NO₂ mixture has better interaction with the surfaces of molybdenite-2H than individual NO₂ ($dE_{ad}/dN_{NO_2} = 2.57$ kcal mol⁻¹).

The layered crystal structure of molybdenite-2H shows closely packed double layers with S⁶-coordination of Mo atom (Fig. 9). The adsorption of different adsorbate molecules like H₂,³⁶ DNA,³⁷ CO,³⁸ thiophene,³⁹ and methanol⁴⁰ was studied by molecular modeling on the surface of molybdenite to gain more insights into the mechanisms of catalytic reaction^{36,38–40} and sensing.³⁷ It has been shown that molybdenite is a promising candidate for a new generation of gas sensors based on layered compounds. The chemical interaction mechanism for NO₂ sensing was proposed as follows:⁴¹



The calculation results of equilibrium adsorption of Ar/NO₂(gas) and Ar/NO₂⁻(ads) show negative values (exothermic process) of total adsorption energies at different temperatures (Table 1). Although the total adsorption energies and differential adsorption (desorption) energies dE_{ad}/dN_{NO_2} are negative, the differential adsorption energies $dE_{ad}/dN_{NO_2^-}$ are positive at different temperatures. This implies that higher temperatures should lead to the desorption of NO₂ and the shift of the adsorption–desorption equilibrium to adsorption of Ar/NO₂⁻(ads) on molybdenite-2H. The energy gain between the

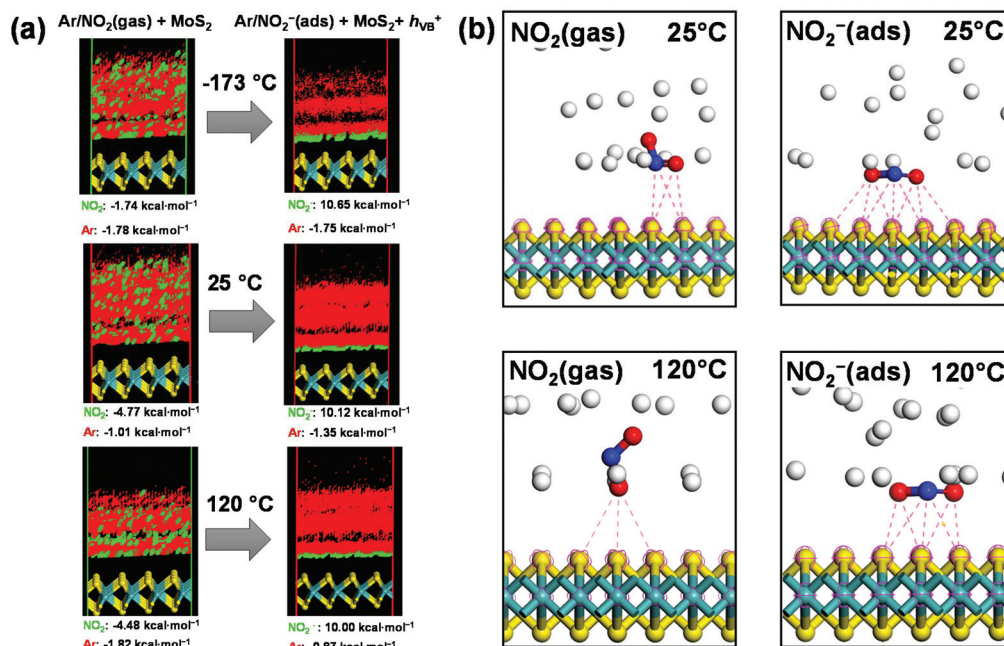


Fig. 9 (a) Field density distribution of a 5% NO₂ and 95% Ar mixture on the predominant (002) surface of molybdenite-2H. Atoms: sulphur, yellow; and molybdenum, light blue. Isosurface: NO₂, green; and Ar, red. (b) Close contacts (less than 3.7 Å) of NO₂/NO₂⁻ species on the predominant (002) surface of molybdenite-2H. Atoms: sulphur, yellow; molybdenum, light blue; nitrogen, dark blue; oxygen, red; and argon, grey.



Table 1 Energies of adsorption (desorption) of Ar/NO₂ and Ar/NO₂⁻ on the (002) surface of molybdenite-2H

Temperature, °C	Adsorption energy	Rigid adsorption energy	Deformation energy	dE _{ad} /dN (Ar)	dE _{ad} /dN (NO ₂)
MoS₂/NO₂(gas)					
-173 °C	-36.88	-36.88	0.00	-1.78	-1.74
25 °C	-33.63	-33.64	0.01	-1.01	-4.77
120 °C	-38.59	-38.59	0.00	-1.82	-4.48
MoS₂⁺/NO₂⁻(ads)					
-173 °C	-22.49	-22.49	0.00	-1.75	10.65
25 °C	-20.39	-20.40	0.01	-1.35	10.12
120 °C	-20.79	-20.80	0.00	-0.87	10.00

adsorption and desorption processes for NO₂/NO₂⁻ species can be estimated as:

$$\Delta = dE_{ad}/dN(\text{NO}_2) - dE_{ad}/dN(\text{NO}_2^-) \approx 14 \text{ kcal mol}^{-1} \quad (3)$$

In general, the dE_{ad}/dN_{NO₂} and dE_{ad}/dN_{NO₂⁻} values are similar at 25 °C and 120 °C. At 120 °C, the total adsorption of Ar/NO₂ (-38.59 kcal mol⁻¹) is greater than that at lower temperatures. However, the dE_{ad}/dN_{NO₂} value at 120 °C (-4.48 kcal mol⁻¹) is slightly decreased when compared to the dE_{ad}/dN_{NO₂} value at 25 °C (-4.77 kcal mol⁻¹). The highest desorption of Ar/NO₂⁻ is observed at 25 °C (-20.39 kcal mol⁻¹), while the desorption of NO₂⁻ at 25 °C (10.12 kcal mol⁻¹) is similar to that at 120 °C (10.00 kcal mol⁻¹). Thus, NO₂ adsorption is still high at 120 °C with high NO₂ desorption. Quantitatively, the optimal adsorption-desorption ratio of NO₂/NO₂⁻ and the energy gain (-14.99 kcal mol⁻¹) are observed at 25 °C although the largest gain in total adsorption-desorption is at 120 °C (17.80 kcal mol⁻¹).

The calculation results show that the adsorption of Ar and NO₂ molecules is impossible between the double S-Mo-S layers connected by weak van der Waals interactions. The negative and positive charges of the molybdenite-2H surface should have more influence on the electrostatic interactions between the molybdenite-2H surface and Ar/NO₂⁻ species. As a result, NO₂ is adsorbed in a vertical position when compared to NO₂⁻ which is adsorbed in a horizontal position correspondingly (Fig. 9b). It is consistent with the previous data⁴⁰ that the negatively charged DNA-molecules could not interact with the positively charged Mo atoms due to the stereochemical shield of the Mo atoms from the adsorption surface. Therefore, DNA could be adsorbed more effectively on the MoS₂ surface *via* van der Waals interactions than through electrostatic interactions between the phosphate and Mo atoms. As shown in Table 1, the interaction forces between NO₂⁻ and MoS₂ were analyzed, and the total interaction energy was found to be about -14 kcal mol⁻¹. The van der Waals force is similar to the total energy, while the contribution of electrostatic interaction force is less than -1.0 kcal mol⁻¹ in the case of NO₂⁻ adsorption. It indicates that the electrostatic interaction is relatively weak, and the main adsorption force is van der Waals interaction. This is in good agreement with our description for the molecular modeling data of close contacts between NO₂/NO₂⁻ and molybdenite-2H (Fig. 9b). The number of close contacts of NO₂ is less than the number of close con-

tacts of NO₂⁻, indicating better interaction of NO₂⁻ species with the positively charged surface of molybdenite-2H.

4. Conclusions

In summary, we have characterized molybdenite concentrate from the mining-metallurgy industry as a valuable starting material for the fabrication of an efficient and low-cost nanostructured gas sensor. After liquid nitrogen exfoliation, impurities were removed and the enriched molybdenite-2H was deposited on different substrates by spin coating and drop casting. It was found that spin coating was advantageous over drop casting in fabricating a homogeneous and dense molybdenite-2H film. The gas sensing behavior of the fabricated sensor was studied for different concentrations of NO₂ at different temperatures. It was found that the temperature strongly influences the key sensing parameters of the device. The sensing behavior of the fabricated device was explained based on the charge-transfer mechanism. The molecular modeling results showed that NO₂ in the Ar/NO₂ mixture had better interaction with the molybdenite-2H surface than individual NO₂, and the main adsorption force was the van der Waals interaction. The findings of this study reflect a straightforward approach for converting molybdenite concentrate into an efficient and low-cost nanostructured gas sensor.

Conflicts of interest

There are no conflicts of interest to declare.

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