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Solvent-driven isomerization of *cis,cis*-muconic acid for the production of specialty and performance-advantaged cyclic biobased monomers†

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The quest for green plastics calls for new routes to aromatic monomers using biomass as a feedstock. Suitable feedstock molecules and conversion pathways have already been identified for several commodity aromatics through retrosynthetic analysis. However, this approach suffers from some limitations as it targets a single molecule at a time. A more impactful approach would be to target bioprivileged molecules that are intermediates to an array of commodity and specialty chemicals along with novel compounds. Muconic acid (MA) has recently been identified as a bioprivileged intermediate as it gives access to valuable aliphatic and cyclic diacid monomers including terephthalic acid (TPA), 1,4-cyclohexanedicarboxylic acid (CHDA), and novel monounsaturated 1,4-cyclohexanedicarboxylic acids (CH1DA, CH2DA). However, accessing these cyclic monomers from MA requires to first isomerize biologically-produced *cis,cis*-MA to Diels–Alder active *trans,trans*-MA. A major impediment in this isomerization is the irreversible ring closing of MA to produce lactones. Herein, we demonstrate a green solvent-mediated isomerization using dimethyl sulfoxide and water. The mechanistic understanding achieved here elucidates the role of low concentrations of water in reducing the acidity of the system, thereby preventing the formation of lactones and improving the selectivity to *trans,trans*-MA from less than 5% to over 85%. Finally, a Diels–Alder reaction with *trans,trans*-MA is demonstrated with ethylene. The monounsaturated cyclic diacid obtained through this reaction (CH1DA) can be converted in a single step into TPA and CHDA, or can be directly copolymerized with adipic acid and hexamethylenediamine to tailor the thermal and mechanical properties of conventional Nylon 6,6.

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Introduction

The polymer industry has become a key driver for the research and development of biorenewable chemicals due to end user appeal for green plastics.^{1–4} For instance, biobased terephthalic acid (TPA) is receiving significant attention due to its central role in the production of polyethylene terephthalate (PET), a commodity polyester broadly used in bottling and

packaging.^{5–17} Other cyclic targets include specialty chemical 1,4-cyclohexanedicarboxylic acid (CHDA), which is gaining attention for tuning properties in polyesters while enhancing their sustainability.^{18,19} The market size for these cyclic building blocks is currently valued at \$60 billion and is expected to further grow at an annual rate of 5%.^{20,21} Therefore, even a conservative 5% replacement by biobased drop-in chemicals represents tremendous opportunities for the emerging biorenewable chemical industry.

Technologies developed to access cyclic molecules from renewable feedstocks have largely used retrosynthesis strategies. In the case of biobased TPA production, research has primarily focused on renewable *p*-xylene as a drop-in replacement for catalytic oxidation in the AMOCO process.^{6–8,13,22,23} An alternate route towards these cyclic diacids is through muconic acid (MA), an intermediate produced from either sugar or lignin using metabolically engineered yeasts and bacteria.^{24–28} The latter approach is particularly attractive as

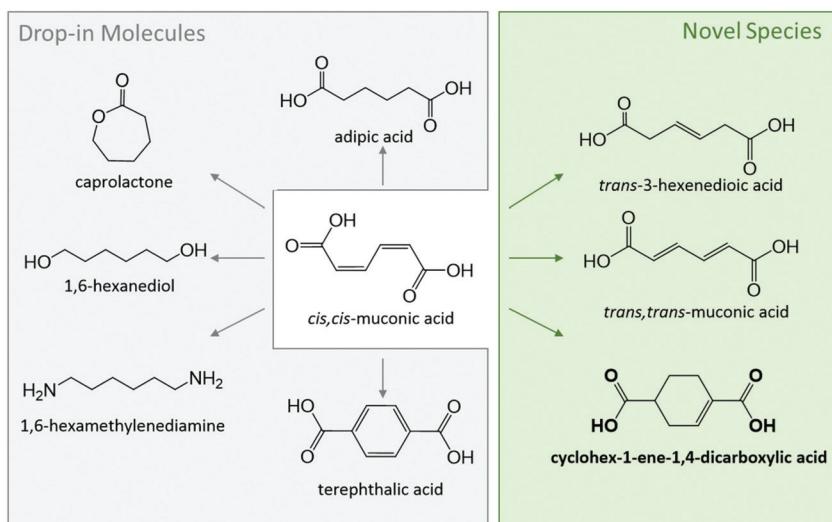
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Scheme 1 Bioprivileged molecule ccMA produced biologically from sugar or lignin can be diversified into various drop-in and novel species.

MA is a bioprivileged molecule with substantial potential for diversification to commodity and specialty chemicals, as well as novel molecules for enhanced end-use properties (Scheme 1).^{29–31} Previous work has already demonstrated the conversion of MA to an array of aliphatic commodity monomers including adipic acid and hexamethylenediamine,^{24,25,32–36} cyclic monomers such as ϵ -caprolactam,^{37,38} TPA and CHDA,^{5,19,39,40} and novel mono-unsaturated compounds such as 3-hexenedioic acid and 1,4-cyclohex-1/2-enedicarboxylic acid (CH1DA, CH2DA).^{19,41–44}

Although the downstream production of cyclic molecules has experienced some significant advances,^{39,45} the initial isomerization of biologically-produced *cis,cis*-muconic acid (ccMA) to Diels–Alder active *trans,trans*-muconic acid (*tt*MA) remains a major bottleneck.⁴⁶ Only a handful of MA isomerization technologies have been reported in the literature. Notable ones use catalysts such as Pd/C (methanol reflux) or iodine either under UV light or in solvents such as tetrahydrofuran, methanol, or acetonitrile.^{40,47–49} Various disadvantages, such as cost of noble metal catalysts, low selectivity, feed concentration limitations (e.g., 10 wt% MA for the iodine system), hinder them from reaching the commercial realm.⁵ The key obstacle in this seemingly simple isomerization is the spontaneous ring-closing of *cis,cis*- and *cis,trans*-MA (*ct*MA) to form the muconolactone (Mlac) and dilactone (Lac2). A previous report published by our group detailed the driving forces behind this lactonization when the isomerization is performed in water, and suggested several strategies to circumvent those driving forces.¹⁹

The present work investigates an organic solvent-mediated isomerization and demonstrates its clear potential to be scaled. It is shown here that the addition of water to the green aprotic solvent dimethyl sulfoxide (DMSO),^{50,51} enhances selectivity to *tt*MA ($S_{tt\text{MA}}$) by 13-fold at 20% conversion. A comprehensive mechanistic understanding of the role of water was achieved through a series of reactions with *ct*MA in DMSO/

water solvent systems. As a proof of concept, *tt*MA was reacted with ethylene to produce the novel unsaturated molecule 1,4-cyclohex-1-enedicarboxylic acid (CH1DA); finally, a novel polyamide was formed through copolymerization with adipic acid and hexamethylenediamine. Overall, this study demonstrates a green, solvent-driven, scalable isomerization of MA in order to synthesize renewable cyclic diacids.

Results and discussion

While *cc*MA readily isomerizes to *ct*MA at room temperature,^{19,52} further isomerization to Diels–Alder active *tt*MA is hindered by the intramolecular lactonization between unsaturated carbon–carbon and carboxylic acid moieties. We previously showed that the aprotic solvent DMSO impedes this lactonization and allows for high $S_{tt\text{MA}}$ (88%).¹⁹ To investigate this conversion on a more industrially relevant scale, a sample with a *cc*MA concentration on the order of 300 g L⁻¹ (~2 M concentration) was prepared in DMSO-d₆. Despite ¹H NMR spectra at such a high concentration being qualitative at best, heating a highly concentrated solution of *cc*MA at 100 °C overnight ($t_{\frac{1}{2}} \sim 2$ –4 hours) yielded primarily Mlac with no indication of the expected *tt*MA product. This result contrasted with our prior study,¹⁹ indicating an effect of reactant concentration on the isomerization reaction. Therefore, this system was investigated kinetically in order to develop optimization strategies for the isomerization to *tt*MA.

Effect of muconic acid concentration in dry DMSO-d₆

The experiments in this study used *ct*MA as the starting material, with concentrations ranging from 5 to 80 mM. Reactions were typically performed at 87 °C unless stated otherwise. The product selectivity at 20% and 50% conversion was determined as a function of initial *ct*MA concentration ($[ct\text{MA}]_0$) as shown in Table 1 and Fig. 1. Maximum $S_{tt\text{MA}}$



Table 1 Effect of initial ctMA concentration on observed rate constants at 87 °C and product selectivity at 20% and 50% conversion

[ctMA] ₀ (mM)	$k_{\text{obs}} \times 10^6$ ^a (s ⁻¹)	Half-life (days)	Selectivity (20% conversion)				Selectivity (50% conversion)			
			ccMA	ttMA	Mlac	Lac2	ccMA	ttMA	Mlac	Lac2
5.5	2.6	3.1	0.8%	19.0%	35.0%	37.3%	0.4%	14.6%	26.0%	31.0%
29.0	1.9	4.2	6.3%	12.0%	51.5%	25.0%	1.6%	6.4%	34.8%	34.2%
45.0	1.8	4.4	6.5%	6.3%	55.0%	26.0%	1.8%	4.0%	37.6%	35.6%
79.0	2.1	3.8	7.5%	3.0%	47.5%	15.0%	1.6%	3.0%	36.4%	32.4%

^a k_{obs} was determined at low conversion (ca. 20%) from eqn (1). Error in k_{obs} was determined by least squares method to be $\pm 1 \times 10^{-7}$ s⁻¹.

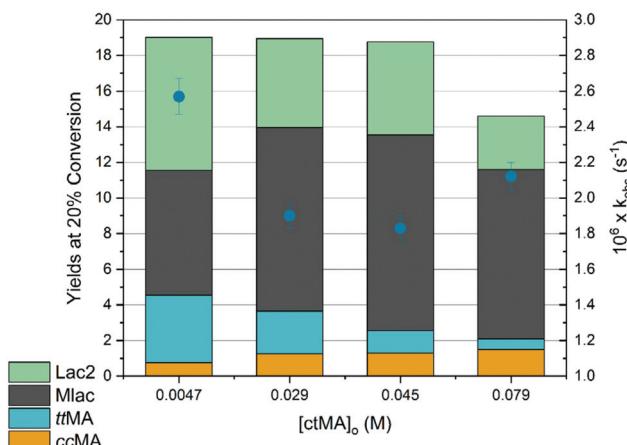
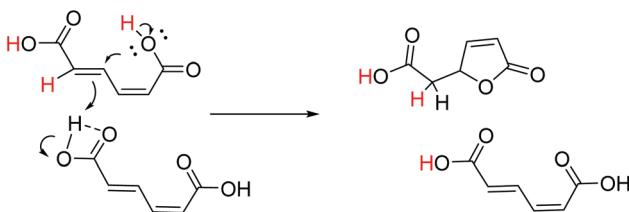


Fig. 1 Effect of [ctMA]₀ on yields at 20% conversion (left axis) and the observed rate constant (right axis). 5–80 mM ctMA solutions in DMSO- d_6 and heated to 87 °C. Kinetic traces and product distributions were obtained by ¹H NMR monitoring with DMSO₂ internal standard.

(19%) was observed at low conversion (ca. 20%) and the lowest [ctMA]₀. $S_{tt\text{MA}}$ gradually decreased from 19% to 3% as [ctMA]₀ increased from 5 mM to 79 mM at 20% conversion. Additionally, $S_{tt\text{MA}}$ at [ctMA]₀ of 5 mM decreased from 19% to roughly 15% as conversion increased from 20% to 50%. Several new signals consistent with MA degradation and formation of unidentified products were observed for all [ctMA]₀. This degradation was observed to occur more readily at elevated ctMA concentration. The significant drop in $S_{tt\text{MA}}$ with increasing [ctMA]₀ suggests an acid catalyzed pathway for lactonization analogous to the aqueous system.¹⁹ However, in this particular system, muconic acid is both the reactant and the catalyst as is shown in Scheme 2.



Scheme 2 Bimolecular pathway for muconic acid catalyzed lactonization scheme.

Kinetic traces obtained for the consumption of ctMA are consistent with a first order rate equation (eqn (1)) for all [ctMA]_t.

$$\ln([ctMA]_t/[ctMA]_0) = -k_{\text{obs}} \times t + \text{const.} \quad (1)$$

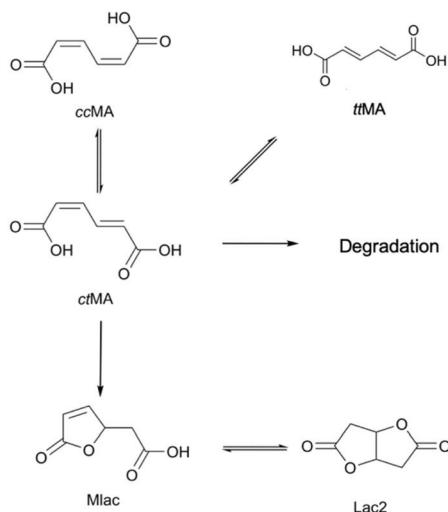
where [ctMA]_t = concentration of ctMA at time t , [ctMA]₀ = initial concentration of ctMA, k_{obs} = observed first order rate constant, t = time in seconds, and const. = integration constant of first order rate law.

Fig. 1 shows the observed rate constants (k_{obs}) obtained from eqn (1) plotted against [ctMA]₀. The trend for k_{obs} with increasing [ctMA]₀ is somewhat peculiar and initially unexpected as the isomerization of ccMA and ctMA to ttMA was shown to be unimolecular in an aqueous system.¹⁹ Therefore, an acid-catalyzed pathway that generates lactones would be expected to generate a linear plot with a y-intercept equal to the sum of all rate constants for unimolecular MA reactions and a slope corresponding to $k_{\text{acid-catalysed}}$ (M⁻¹ s⁻¹). The unimolecular reaction pathways would include isomerization, unimolecular lactonization, and unimolecular degradation (Scheme 3). However, k_{obs} was found to initially decrease with increasing [ctMA]₀ between 5 and 45 mM and then increase between 45 and 79 mM. The initial decrease at low [ctMA]₀ suggested an equilibrium with a non-reactive species like an unreactive MA complex. Though kinetically observable, the non-reactive complex was not identifiable spectroscopically. At higher [ctMA]₀ the acid catalyzed pathway shown in Scheme 2 appeared to be dominant. In addition to higher Mlac yields, elevated [ctMA]₀ also resulted in significantly higher degradation to unidentified byproducts. Clearly, in the absence of water, this system is rather complex and will not offer high $S_{tt\text{MA}}$ under industrially relevant conditions (i.e. high concentration). It was therefore decided to minimize efforts in this system and focus on approaches that take into consideration the effect of [ctMA]₀: (i) an acid-catalyzed lactonization pathway and (ii) the potential for reversible formation of a non-reactive complex/es. The former is more important as it clearly dominates near the solubility limit. It was therefore concluded that decreasing the apparent acidity should improve $S_{tt\text{MA}}$.

Effect of low water concentrations on muconic acid isomerization

The Dumesic group demonstrated that the addition of water to solid-acid catalyzed systems in DMSO and γ -valerolactone





Scheme 3 Unimolecular reaction pathways for ctMA.

(GVL) can significantly decrease the activity of the catalyst.⁵³ The decrease in catalyst activity could be attributed to a more widespread hydrogen bonding network near the active sites, which effectively decreases their pK_a . This rationale was extended to our reaction system by introducing water to minimize the acid-catalyzed lactonization of ctMA. The experiments in this section were carried out using 48 mM ctMA in dry DMSO-d₆ with varying concentrations of water ($[\text{H}_2\text{O}]$). ¹H NMR spectra of samples containing water (5 mM–744 mM) indicated that MA was in its fully protonated (ctMAH₂) form under all conditions.

At very low $[\text{H}_2\text{O}]$ the reactivity of ctMA was similar to that of the dry DMSO system at 20% conversion (compare 45 mM ctMA in Table 1 with 5 mM H₂O in Table 2). S_{ttMA} at 96 mM H₂O (H₂O : MA \sim 2 : 1) increased 13-fold to 81.5% (88% if conversion is considered as conv. = $[\text{ctMA}]_0 - ([\text{ctMA}]_t + [\text{ccMA}]_t)/[\text{ctMA}]_0$ due to the relatively rapid $\text{ccMA} \leftrightarrow \text{ctMA}$ equilibrium). For $[\text{H}_2\text{O}]$ of 96 mM, high S_{ttMA} (82%) was maintained up to 50%

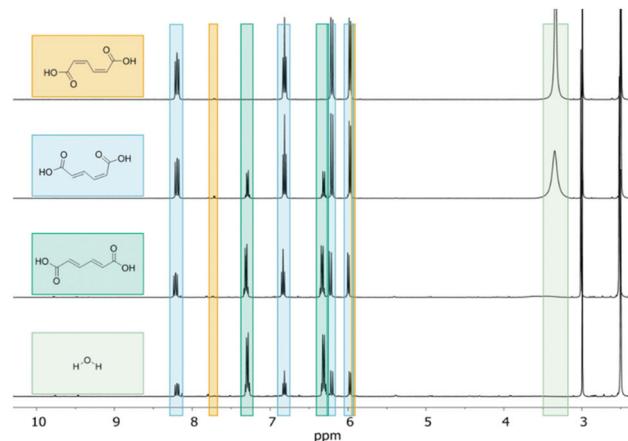


Fig. 2 ¹H NMR (600 MHz, DMSO-d₆) spectra of 48 mM ctMA + 96 mM H₂O heated to 87 °C at 0 (top), 18.5, 52, and 72.5% (bottom) conversion.

conversion. Additionally, throughout this reaction the ¹H NMR signal of H₂O (3.34 ppm) gradually decreased with increase in conversion as shown in Fig. 2, until it was not observable after 52% conversion. The majority of the samples showed an increased S_{ttMA} with decreasing H₂O signal, suggesting that an optimal ratio of H₂O : MA exists for a highly selective production of ttMA (\geq 75%). This optimal range appears to be broad, varying from a H₂O : MA ratio of 0.9 : 1 to 10 : 1.

Kinetically, the conversion of ctMA follows first order rate equations. As the water content was increased, there was an initial 20% drop in k_{obs} between 5 mM and 38 mM, followed by a gradual increase (70–120 mM H₂O), and a plateau from 120 mM to 744 mM. The initial decrease was expected under the assumption that an acid-catalyzed lactonization pathway exists in parallel to unimolecular isomerization, and that introduction of water to the system would decrease the apparent pK_a . The sigmoidal shape observed in Fig. 3 could be indicative of a ctMA-2H₂O complex that is roughly 20% more active for isomerization to ttMA than a ctMA-H₂O complex. The plateau achieved at high [H₂O] coupled with decreasing H₂O

Table 2 Effect of water concentration on muconic acid isomerization in DMSO^a

[H ₂ O] ₀ (mM)	$10^6 \times k_{\text{obs}}^b$ (s ⁻¹)	Half-life (days)	Selectivity at 20% conversion			Selectivity at 50% Conversion		
			ccMA	ttMA	Mlac	ccMA	ttMA	Mlac
5.0	1.8	4.5	6.5%	6.3%	55.0% ^c	1.8%	4.0%	37.6%
37.7	1.4	5.9	7.5%	70.0%	3.5%	1.8%	69.8%	2.8%
55.6	1.4	5.9	7.5%	75.0%	3.0%	1.8%	71.2%	2.8%
70.1	1.3	6.2	7.5%	77.5%	4.0%	1.9%	79.0%	2.7%
74.6	1.5	5.4	7.5%	76.0%	3.0%	1.8%	78.2%	2.6%
96.1	1.5	5.5	7.5%	81.5%	3.0%	1.8%	82.0%	2.6%
115.8	1.7	4.7	7.5%	80.0%	2.5%	1.8%	70.0%	2.4%
157.4	1.6	5.0	7.5%	80.0%	2.5%	1.8%	72.0%	2.2%
322.0	1.6	5.0	7.5%	75.0%	4.5%	2.1%	83.7%	2.5%
478.0	1.6	5.0	7.5%	75.0%	4.5%	2.4%	85.7%	2.4%
744.0	1.6	5.0	7.0%	70.0%	3.0%	2.4%	79.0%	2.4%

^a Selectivities at 20 and 50% conversion for 48 mM ctMA in DMSO-d₆ with varying concentrations of water at 90 °C. ^b k_{obs} acquired from fits to eqn (1) at 20% ctMA conversion. ^c 5 mM H₂O sample also contains Lac2.



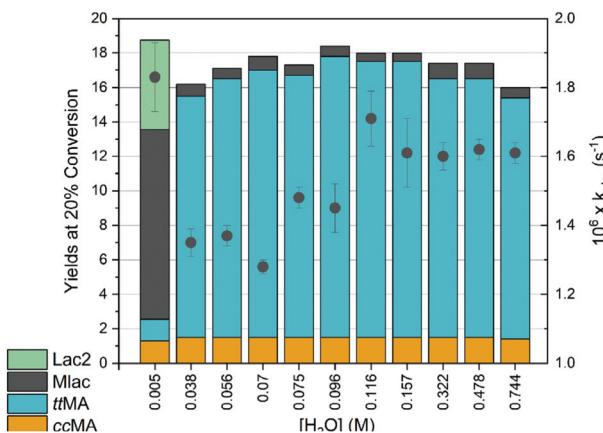


Fig. 3 Effect of water content on yields at 20% conversion (left axis) and the observed rate constant (right axis). Microliter amounts of water were added to 600 μ L of 48.0 mM ctMA solutions in DMSO-d₆ and heated to 87 °C. Kinetic traces and product distributions were obtained by ¹H NMR monitoring with DMSO₂ internal standard.

signal and steady decline in S_{ttMA} above 160 mM H₂O could also be indicative of H₂O being a potential reactant in the degradation pathway.

Probing the effect of muconic acid at low water concentration

To probe the effect of [ctMA]₀ on the kinetics and yields at low water concentrations, these experiments were carried out with D₂O : MA of (1.9–2.6) : 1 ([ctMA]₀ ranging from 0.049–0.16 M). Under these conditions the water should interact primarily with the carboxyl moieties, lowering their apparent acidity and minimizing the acid-catalyzed lactonization shown in Scheme 2. This assumption is supported by NOESY NMR experiments showing proximity between the carboxyl moieties and H₂O in the system (ESI 1†). However, it should be noted that the NOESY spectrum is not definitive proof for the absence of H₂O association with internal carbon atoms.

The selectivity to *cc*- and *tt*MA at 20% conversion remained relatively constant between 50 and 160 mM ctMA (*ca.* 6 and 75–80%, respectively), but selectivity to Mlac increased 6-fold (Fig. 4). Additionally, k_{obs} for ctMA consumption decreased with increasing [ctMA]₀, similar to that observed under dry DMSO. However, unlike the trend observed in Fig. 1, the rate constants did not increase again at higher [ctMA]₀. This observation supports the non-reactive complex hypothesis presented above.

Effect of high water concentrations on muconic acid isomerization

D₂O in concentrations ranging from 1.8 M to 16.0 M was added to a 50 mM ctMA solution resulting in a decrease in k_{obs} for ctMA consumption from 2.1×10^{-6} s⁻¹ to 1.0×10^{-6} s⁻¹, as shown in Fig. 5. *tt*MA yields at 20% conversion remained relatively constant ($15 \pm 1\%$) up to 8.5 M D₂O but were halved to 7.5% at 16 M D₂O. This was accompanied by a slight decrease in equilibrium concentrations of *cc*MA and a 10-fold increase in Mlac yields (Fig. 5). Mlac yields rose steadily from 0.1% to 0.5% as [D₂O] increased from 1.8 to 8.5 M, and then jumped to 5.5% at 16 M D₂O. Rate constants obtained by initial rates

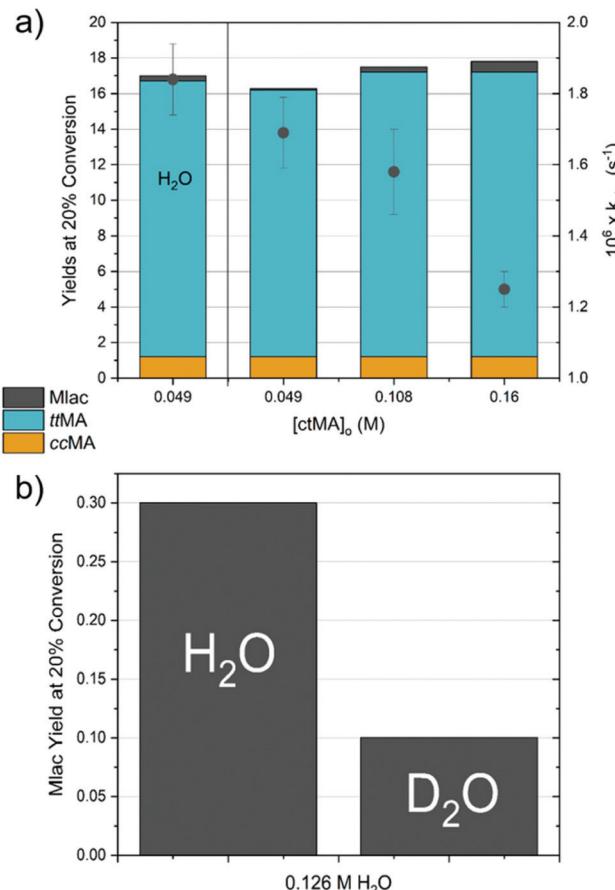


Fig. 4 (a) Effect of [ctMA]₀ with 2–2.6 equivalents of H₂O (column 1) and D₂O (columns 2–4) on yields at 20% conversion (left axis) and the observed rate constant (right axis) for ctMA solutions in DMSO-d₆ heated to 87 °C. Kinetic traces and product distributions were obtained by ¹H NMR monitoring with DMSO₂ internal standard. (b) Comparison of Mlac yield in H₂O and D₂O.

analysis for Mlac formation (eqn (2)) were plotted as a function of [D₂O]. The results suggested a ternary reaction in which water promotes lactonization through a push–pull type interaction, where one water molecule donates a proton to the δ carbon of the *cis*-carboxylic acid while another accepts the proton from the *cis*-carboxylic acid (Scheme 4). This observation is consistent with the large negative entropy of activation ($\Delta S^\ddagger = -110$ J (mol K)⁻¹) obtained from simulations fitted to experimental data for the same reaction in an aqueous system.¹⁹ The rate constant for Mlac formation at 16 M D₂O obtained from eqn (2) is consistent with the aqueous simulations ($4 \pm 2 \times 10^{-7}$ s⁻¹ at 16 M D₂O and 3×10^{-7} s⁻¹ at 55 M).¹⁹ This of course is only consistent if the reaction reaches kinetic saturation near 16 M D₂O (*i.e.* D₂O : ctMA = 320 : 1). The elementary rate constant for the ternary reaction shown in Scheme 4 was calculated to be $2 \pm 1 \times 10^{-9}$ M⁻² s⁻¹.

$$\text{rate}_{\text{Mlac}} = [\text{Mlac}]_t/t \rightarrow k_{\text{Mlac}} = \text{rate}_{\text{Mlac}}/\left\{\frac{1}{2}([ctMA]_0 - [ctMA]_t)\right\} \quad (2)$$



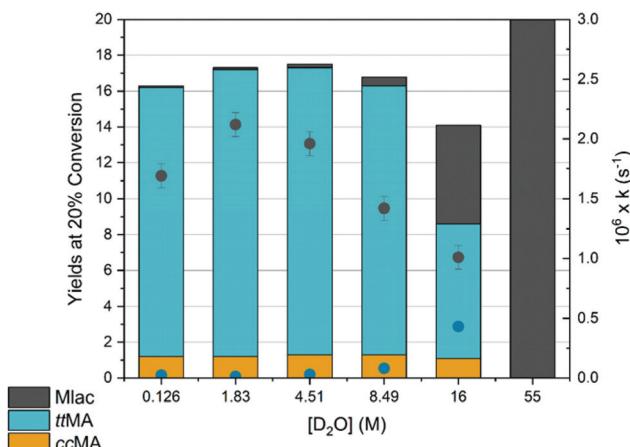
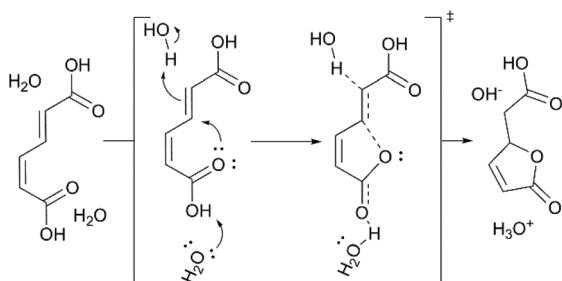


Fig. 5 Effect of high water content on yields at 20% conversion (left axis). The observed rate constant for ctMA conversion (black ●, right axis) was obtained from eqn (1) and the observed rate constant for Mlac formation (blue ●, right axis) was estimated from eqn (2). ca. 50 mM ctMA solutions in DMSO-d₆ + D₂O were heated to 87 °C. Kinetic traces and product distributions were obtained by ¹H NMR monitoring with DMSO₂ internal standard. The column at 55 M D₂O is from a previous experiment and the rate constant obtained is from simulations fit to experimental data in an aqueous system.¹⁹



Scheme 4 Ternary lactonization process with ctMAH₂.

where $[Mlac]_t$ = Mlac concentration at time t , t = time in seconds at which ~5% ctMA conversion was achieved, k_{Mlac} = rate constant for the formation of Mlac, $[ctMA]_0$ = initial ctMA concentration, and $[ctMA]_t$ = concentration of ctMA at time t .

In addition to the solvent driven lactonization, higher water content also showed a decrease in the rate constant for ttMA formation. The rate constant k_{ttMA} was obtained by the same method as k_{Mlac} shown in eqn (2). These rate constants were relatively unaffected up to 8.5 M D₂O but decreased by nearly a factor of 5 from 8.5 to 16 M D₂O. At 16 M D₂O k_{ttMA} was roughly equivalent to k_{Mlac} whereas at 8.5 M D₂O $k_{ttMA} = 84 \times k_{Mlac}$. The increase in k_{Mlac} with [D₂O] was expected for the reaction outlined in Scheme 4, however, the driving forces behind this significant decrease in k_{ttMA} were unclear. At 121 °C the ctMA to ttMA reaction was found to be reversible, *vide infra*. It is possible that water has an effect on the forward and/or reverse reaction(s). Therefore, the reactivity of 50 mM ttMA at 1.8 and 16 M D₂O was tested, which confirmed reversibility at 87 °C. Comparison of k_{obs} starting from both ctMA and ttMA yielded apparent equilibrium constants (K_{app}) that vary with [D₂O], $K_{app} = 7.2$ and 1.5 at 1.78 M and 16 M D₂O, respectively. The effect of water concentration on this equilibrium is exclusively a result of high water content slowing the forward reaction ($ct \rightarrow tt$) as the reverse reaction ($tt \rightarrow ct$) rate constant was the same at both 1.8 and 16 M D₂O ($k_{tt-ct} = 3 \pm 1 \times 10^{-7} \text{ s}^{-1}$).

The solvent driven lactonization at high water concentration in conjunction with the significant decrease in k_{ttMA} at 16 M D₂O (neat D₂O = 55 M) would help explain why ttMA is never observed in an aqueous system in the absence of catalyst like La³⁺.¹⁹

Effects of temperature & kinetic simulations

An investigation into the effects of temperature on the reaction system with ca. 2 equivalents of water was carried out. Product yields at 20% conversion and k_{obs} for ctMA conversion are shown in Fig. 6. S_{ttMA} was relatively high for all three temperatures; 88, 80, and 88% at 60, 87, and 121 °C, respectively (91, 88, and 93% considering the ccMA \leftrightarrow ctMA equilibrium). The reaction at 121 °C was fast enough ($t_{1/2} = 5.2$ weeks, 5.2 days, and 1.7 hours at 60, 87, and 121 °C, respectively) that we were able to achieve equilibrium between ctMA and ttMA. While this is reasonable, it was surprising due to previous observations in which aged aqueous solutions of ttMA and ttMA in DMSO with no added water at 200 °C do not appear to yield ctMA. To probe the temperature effects further, kinetic simu-

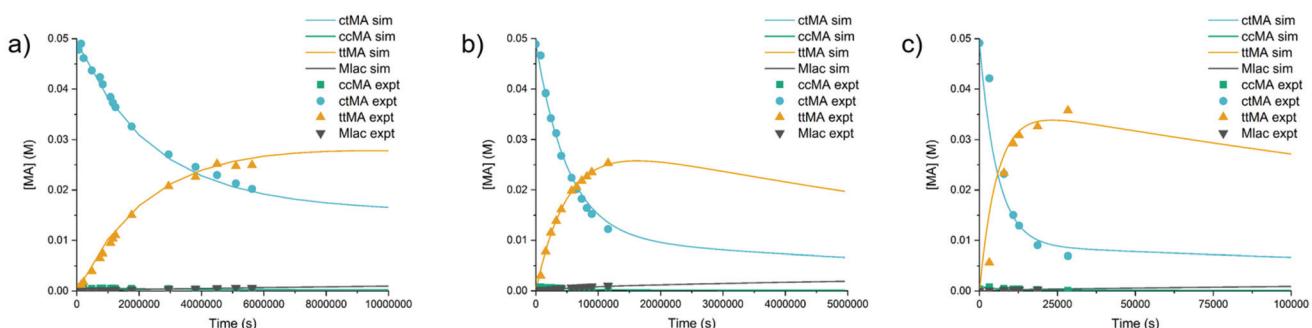


Fig. 6 Experimental (■) and simulated (—) traces of ctMA isomerization at 2 equivalents of D₂O at (a) 60 °C (b) 87 °C (c) 121 °C.

Table 3 Equilibrium constants, rate constants, and activation parameters for the reactions outlined in Scheme 4^a

Reaction	Temperature (°C)			Activation parameters ^d	
	60	87	121	ΔH^\ddagger kJ mol ⁻¹	ΔS^\ddagger J (mol K) ⁻¹
$K_{cc-ct_{b,c}}^b$	68	55	50	—	—
K_{ct-ct}	1.7	2.4	4.7	—	—
$k_{ct\rightarrow tt}$	2.5×10^{-7}	1.2×10^{-6}	1.4×10^{-4}	110 ± 31	-44 ± 85
$k_{tt\rightarrow ct}$	1.5×10^{-7}	5.0×10^{-7}	3.0×10^{-5}	92 ± 28	-103 ± 78
$k_{ct\rightarrow Mlac}$	4.0×10^{-9}	3.1×10^{-8}	6.0×10^{-6}	128 ± 31	-25 ± 86
$k_{ct\rightarrow deg}$	2.0×10^{-8}	3.5×10^{-7}	3.0×10^{-5}	128 ± 14	-11 ± 39

^a All reactions were carried out in DMSO-d₆ with 1.9–2.5 equivalents of D₂O relative to 50 mM ctMA. ^b Equilibrium constants expressed with the subscript ‘isomer1-isomer2’ are for reversible isomerization and should not be confused with equilibrium constants with the subscript ‘isomer1, isomer2’ mentioned above as they represent K for formation of non-reactive complexes of the different isomers. ^c $K_{ct\rightarrow tt}$ is an apparent equilibrium constant derived from kinetic information about the forward and reverse reactions obtained from fitting the simulation to experimental values and does not represent the thermodynamic equilibrium. ^d Activation parameters are calculated for the reactions outlined in Scheme 3 and are specific to the system at 50 mM ctMA with 2 equivalents of H₂O. They do not take into consideration the equilibrium formation of potential non-reactive complexes.

lations were run with KINSIM software in accordance with the simplified Scheme 4. Estimates for elementary rate constants were input based on the experimental values obtained from treatment of the data with either eqn (1) or (2). The rate constants were finely tuned until simulated kinetic traces matched experimental ones for all species, including the loss of signal due to degradation from ctMA. The equilibrium constants, rate constants, and activation parameters are shown in Table 3. While the simulations starting from ctMA fit well with the experimentally obtained kinetic traces, simulating the reverse reaction (e.g. starting from ttMA) required decreasing $k_{cc\rightarrow tt}$ by a factor of 10.

Perhaps non-reactive complexes dependent upon [MA] exist for each of the isomers (see next section on ccMA). The equilibria achieved for these unreactive complexes will likely have equilibrium constants that are dependent upon the specific isomer pair. That is to say $K_{cc,cc}$ may be significantly different than $K_{ct,ct}$, $K_{tt,tt}$, $K_{cc,ct}$, $K_{cc,tt}$, and $K_{ct,tt}$. The focus of this investigation, however, is the unprecedented catalyst-free isomerization of ctMA to ttMA. Therefore, a more thorough study of these potential non-reactive complexes was not carried out at this time.

As a high S_{ttMA} was achieved with 2 equivalents of D₂O at 121 °C (Fig. 6c), these conditions were selected to investigate the isomerization of ctMA under more industrially relevant conditions, namely at 70 and 300 g L⁻¹ (ca. 0.5 and 2.0 M). In contrast to the experiment in dry DMSO-d₆ that prompted this study (*vide supra*), the reaction with 2.0 M ctMA and 2 equivalents of D₂O produced minute amounts of Mlac. However, the reaction quickly reached an equilibrium between ctMA and ttMA favoring ctMA (nearly 75:1 at 121 °C). This [ctMA]₀ effect was not observed at 70 g L⁻¹ (~0.5 M), a concentration equivalent to the

Table 4 Rate constants for isomerization of ccMA^a

[ccMA] ₀ (mM)	[D ₂ O] (M)	$10^7 \times k_{obs}$ (s ⁻¹)
5	0.009	1.6
20	0.009	0.8
5	4.9	4.1
4	9.8	4.6
3	21.5	4.6

^a Isomerization was carried out at 22 ± 1 °C in DMSO-d₆ and monitored by ¹H NMR. Selectivity to ctMA was 100% and signal loss due to degradation was not observed.

highest titers reported to date for biologically-produced ccMA. S_{ttMA} reached 78% at 50% conversion, which is similar to the selectivity achieved with a 48 mM solution (Table 2).

Isomerization from *cis,cis*-muconic acid

As was the case with ctMA, ¹H NMR spectra obtained of solutions initially containing cc- and ttMA were consistent with the fully protonated form (MAH₂). These experiments are outlined in Table 4. Kinetic traces for the isomerization of ccMA to ctMA were fitted to 1st order rate equations with R^2 values ≥97%. Values for k_{obs} decreased with increasing [ccMA]₀ in the presence of 9 mM H₂O. This is consistent with the potential formation of a non-reactive MA complex and the lack of an acid catalyzed pathway for ccMA.¹⁹ Addition of D₂O to the system increased k_{obs} implicating water as a non-innocent solvent consistent with the aqueous $\Delta S^\ddagger = -88$ J (mol K)⁻¹ for cc- to ctMA isomerization of MAH₂.¹⁹ This effect appears to reach saturation at high H₂O:ccMA ratios. Unlike experiments in which ctMA was the initial isomer, no loss of signal was observed with reaction time, and selectivity to ctMA was 100%. This supports the above hypothesis that degradation of MA occurs from the ct-isomer.

Isomerization from biobased *cis,cis*-muconate

The robustness of the isomerization process, specifically its tolerance to impurities, was investigated using a fermentation broth containing 70 g L⁻¹ (~0.5 M) of *cis,cis*-ammonium muconate (ESI†). The broth was first filtered over activated carbon to remove biogenic impurities and organic byproducts. ccMA was subsequently precipitated by acidification of the solution using 5 M sulfuric acid. The precipitate was recovered by filtration, rinsed with a minimum amount (10 ml) of 5 M sulfuric acid to prevent its redissolution, and dried overnight at room temperature. The biobased ccMA was then isomerized to ctMA using the procedure described previously. Further isomerization to ttMA was carried out in DMSO with 1.5 equivalents of H₂O. The ttMA yield at 50% conversion reached 34.8%, which is on par albeit lower than the yield obtained for reagent-grade ccMA (39.5%), see Fig. 7. This difference can be at least partly attributed to residual sulfuric acid that catalyzed the lactonization of ccMA during the isomerization step to ctMA, as indicated by ¹H NMR (ESI 2 and 3†). This issue could be easily addressed by improving the separation-purification process, which is beyond the scope of the present work.



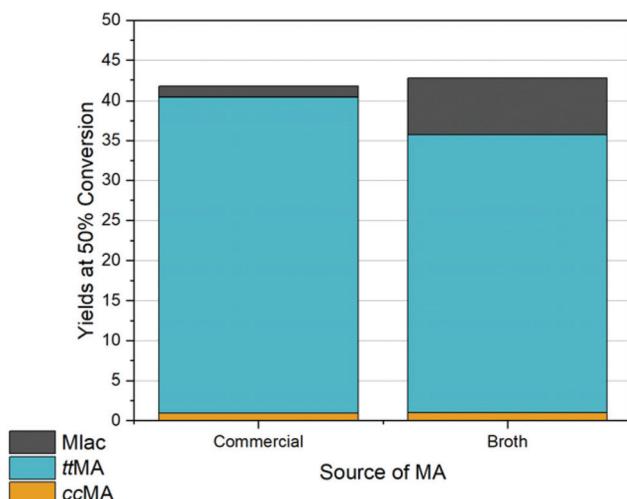


Fig. 7 Product distribution and yields obtained at 50% conversion when the reaction was carried out with reagent-grade ccMA and bio-based ccMA recovered from the fermentation broth. The reactions were performed with 48.0 mM ctMA solutions in DMSO- d_6 and 2 equivalents of H_2O and heated to 121 °C. Product distributions were obtained by 1H NMR monitoring with $DMSO_2$ internal standard.

Synthesis and copolymerization of the monounsaturated cyclic diacid CH1DA

Further conversion of Diels–Alder active *ttMA* to TPA and CHDA have been demonstrated in literature.^{19,54,55} Here, we show the preparation of novel monounsaturated cyclic diacids and their role on the physical and mechanical properties of Nylon-6,6 when copolymerized with adipic acid and hexamethylenediamine. *ttMA* underwent cycloaddition with ethylene in GVL to form cyclohex-1-ene-1,4-dicarboxylic acid (CH1DA) with >99% yield. Although cyclohex-2-ene-1,4-dicarboxylic acid (CH2DA) is the expected cycloadduct, CH2DA was not detected as it readily isomerizes in GVL likely due to the stabilization resulting from conjugation (Scheme 5).

The synthesized CH1DA was further copolymerized with adipic acid and hexamethylenediamine to demonstrate the potential of this new monomer for tuning the properties of conventional Nylon. To maintain an equimolar ratio between the diacid and diamine, 10 to 25 mol% of adipic acid were replaced by CH1DA (Fig. 8). This range was selected taking into consideration the economic feasibility of using novel chemicals. Gel permeation chromatography (GPC), differential

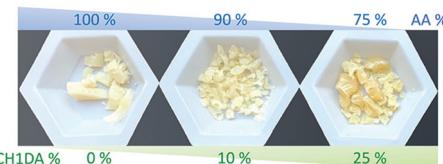


Fig. 8 Copolymerization of adipic acid (AA) and CH1DA with hexamethylenediamine.

Table 5 Properties of polyamides with various molar incorporation of CH1DA^a

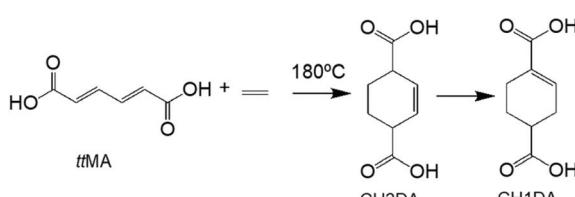
Sample	M_n [kDa]	PDI	T_{melt} [°C]	G' [GPa]	G'' [MPa]	T_g [°C]
PA66	16.5	2.5	261.0	1.32	16.52	73.3
CH1DA-10%	15.8	2.1	250.2	1.65	17.39	75.8
CH1DA-25%	12.4	2.3	232.6	1.65	10.99	80.3

^a M_n : number-average molecular weight with reference to poly(methyl methacrylate) standards; PDI: polydispersity index; G' , G'' : storage and loss moduli in the glassy plateau at 0 °C; T_g : glass transition temperature calculated from peak of $\tan(\delta)$.

scanning calorimetry (DSC), and dynamic mechanical analysis (DMA) were used to characterize these polymers using Nylon-6,6 as a reference (ESI 4–6†). The corresponding results are summarized in Table 5. The addition of the cyclic diacid into Nylon's aliphatic backbone altered its crystallinity due to the defects introduced by CH1DA sequences. This change reduces the polymer's melting temperature from 261 to 233 °C, which facilitates its processing through blow molding. Moreover, introducing a cyclic molecule in the polymer backbone increased the rigidity of polymer chains, thereby increasing the storage modulus and glass transition temperatures. The addition of CH1DA allows control over Nylon-6,6 thermal properties without sacrificing its mechanical properties, as shown in Table 5. Overall, CH1DA offers a renewable alternative to commercially available products such as INVISTA's Dytek diamines with the additional benefit of providing an unsaturated bond. This unsaturation allows for further functionalization to insert properties such as hydrophobicity and flame retardance directly into Nylon's backbone.²⁸

Conclusion

A novel solvent-driven, catalyst-free, isomerization of *ctMA* to *ttMA* was demonstrated and investigated mechanistically. The lactonization process that is dominant in aqueous media was overcome in solvent systems combining the polar aprotic solvent DMSO and water.¹⁹ In dry DMSO, MA generates a significant quantity of lactones through an acid-catalyzed pathway analogous to the aqueous system. This reaction is enhanced at higher [*ctMA*] due to MA acting as both the catalyst and reactant. The addition of 1 equivalent of water has the effect of reducing the apparent acidity of the system, thereby



Scheme 5 Reaction between *ttMA* and ethylene to produce monounsaturated cyclic diacid CH1DA.



reducing lactonization. Consequently, the S_{ttMA} was increased by 13-fold.

The production of *tt*MA is a relatively slow process, particularly at high $[ctMA]$. This is believed to be due to the formation of a non-reactive MA complex that slow the kinetics. The system is also limited by the reversibility of the system in DMSO/H₂O mixtures, which was not previously observed in the aqueous system. Nevertheless, high yields of *tt*MA can be obtained at elevated temperatures in a relatively short period of time due to a shift in the *ct*MA to *tt*MA equilibrium constant favoring *tt*MA (nearly 5 : 1 at 121 °C). Although further increase in temperature would favor isomerization to *tt*MA, it needed to be balanced with the decomposition of DMSO at elevated temperatures.⁵⁰

Continued addition of water sheds light on the role of solvent in both isomerization and lactonization for the aqueous system. Previous work supported by simulations and kinetic measurements indicated that the isomerization of *cc*MA to *ct*MA and the lactonization of *ct*MA to *M*lac were unimolecular reactions. However, these reactions were both accompanied by relatively large entropies of activation (−88 and −110 J (mol K)^{−1}, respectively).¹⁹ While the involvement of water in the isomerization of *cc*MA is still not definitively shown, it is strongly supported by the increase in k_{obs} with water content and apparent saturation kinetics when D₂O was in large excess of MA. Lactonization of *ct*MA, on the other hand, is strongly supported by the dependence of k_{Mlac} on [D₂O]. Furthermore, this observed dependence on [D₂O] indicates that this transformation is a ternary reaction in which water acts as both a proton donor and a proton acceptor (Scheme 2). In addition to the enhanced lactonization at high [D₂O], $k_{ct \rightarrow tt}$ becomes slower while $k_{tt \rightarrow ct}$ remains unchanged; resulting in a shift in the apparent equilibrium constant to favor *ct*MA over *tt*MA. Perhaps this offers an explanation as to why *tt*MA has never been observed to form in aqueous solutions without the presence of a catalyst.

Under optimal conditions (2 equivalents of water, 121 °C), *ct*MA is isomerized to *tt*MA with 88% selectivity, with *M*lac and degradation products representing less than 3% of the mixture. The other species in solution are *ct*MA (in equilibrium with *tt*MA) and *cc*MA (in equilibrium with *ct*MA). In an industrial setting, *cc*MA and *ct*MA would be recovered and recycled, making it a scalable, green, and cost-efficient process.

This advancement in MA isomerization technology will not only allow for the development of biobased commodity chemicals such as TPA, but also for novel unsaturated cyclic molecules such as CH1DA. Incorporation of CH1DA in a polyamide backbone enabled tunable properties that could help with processability of polymers. Further modification of the alkene moiety in this diacid could additionally achieve targeted property enhancement in polyesters and polyamides.

Experimental

cis,cis-muconic acid (*cc*MA), *trans,trans*-muconic acid (*tt*MA), DMSO-d₆ (99.96% D), γ -valerolactone (GVL), tetrahydrofuran

(THF), hexamethylenediamine (HMDA), adipic acid (AA), tetramethylsilane and dimethyl sulfone (DMSO₂) were purchased from Sigma-Aldrich. Ethylene gas was purchased from Airgas at 99.5% purity. *cis,trans*-Muconic acid (*ct*MA) was prepared by methods previously described in the literature.⁴¹ Extra-dry DMSO-d₆ was refluxed over CaH₂ and stored over sieves.

For experiments in which strict water-free conditions were required, a mother solution containing *ct*MA and DMSO₂ (internal standard) with dried DMSO-d₆ was prepared in an inert atmosphere box. Experiments in which the effect of [MA] was investigated utilized 600 μ L of the mother solution in J. Young tubes containing additional solid *ct*MA. *ct*MA concentrations were determined by NMR vs. DMSO₂ internal standard in the mother solution. The tubes were sealed, placed in an Erlenmeyer flask with a thermometer, and heated in a laboratory oven. The experiments at 121 °C were carried out using an aluminum heating block equipped with a thermocouple. Samples were removed from the oven, cooled to room temperature, and analyzed by ¹H NMR throughout the duration of the experiment. Experiments that investigated the effect of water were prepared in DMSO-d₆ that had not been dried. Instead, aliquots of a mother solution containing *ct*MA and DMSO₂ were added to J. Young tubes and H₂O was added *via* syringe (<1 μ L–100 μ L). The water concentration was determined by ¹H NMR signal relative to DMSO₂ standard. High water concentration experiments (>1 M) utilized D₂O to minimize interference with ¹H NMR spectra. D₂O was added to the mother solution with an electronic micropipette (>100 μ L). Dilution of [*ct*MA] was adjusted by addition of solid *ct*MA and was again determined relative to [DMSO₂] internal standard.

Cyclohex-1-ene-1,4-dicarboxylic acid (CH1DA) was produced through a Diels–Alder cycloaddition of *tt*MA with 500 psig ethylene in γ -valerolactone at 180 °C for 24 hours in a 50 ml pressurized reactor (Parr 4590 Series). The products in the liquid phase were filtered using a cellulose filter, washed repeatedly with water, and dried overnight in a drying oven. The dried product was then dissolved in DMSO-d₆ and characterized using NMR using tetramethylsilane as internal standard.

To prepare the salt for polymerization, adipic acid and CH1DA were added in the molar ratio of $x:(1-x)$ with $0.75 \leq x \leq 1$ and dissolved in methanol and THF, respectively. The molar equivalent of HMDA was then added to the resulting mixtures and heated to 40 °C. The precipitated salt was filtered, washed and dried overnight in a fume hood. Polymerization was carried out in a Parr reactor charged with the salt and 60 wt% water and pressurized with 100 psig N₂. The mixture was heated to an internal temperature of 210 °C, held there for 80 minutes, followed by venting out the N₂ and water. The sample was then allowed to polymerize at an internal temperature of 270 °C at atmospheric pressure for 1.5 hours and cooled to room temperature. Molecular weight of copolymers was obtained through gel permeation chromatography (GPC) using EcoSEC GPC system. Polymer samples of around 5 mg were dissolved in 1,1,1,3,3,3-hexafluoroiso-



propanol and compared to poly(methyl methacrylate) standards. Samples were then annealed in an oven for 6 hours at 150 °C. DSC was carried out on the copolymer using TA Q100. Dynamic mechanical analysis was performed using a TA ARES-G2 rheometer.

¹H NMR spectra were collected with a Bruker AVIII600 spectrometer, and spectra were analyzed with MestReNova software. Data were plotted with OriginPro 9.1 software.

Conflicts of interest

There are no conflicts to declare.

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