

## PAPER

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## A rigid anionic Janus bis(NHC) – new opportunities in NHC chemistry†

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A phosphanido-type bridged bis(imidazolium) salt, readily prepared in two steps *via* reductive deselenization of a tricyclic 1,4-diphosphinine diselone, affords access to a novel anionic P-functional tricyclic bis(NHC) *via* deprotonation. The former also offers a P-functionalization/deprotonation sequence to access the first mixed P-substituted tricyclic bis(NHCs), as well as coordination of the phosphorus centers to rhodium(I) fragments.

## Introduction

Owing to the wide range of structural and property modifications of N-heterocyclic carbenes (NHCs), they have become potent ligands in organometallic chemistry and catalysis.<sup>1,2</sup> It has been reported that the incorporation of an anionic functionality confers higher stability to resulting NHC complexes compared to related neutral donor substituents.<sup>3</sup> Backbone-functionalized NHCs having anionic heteroatom substituents especially facilitate  $\pi$ -electron interactions with the heterocyclic ring, which can lead to electronic tuning of the donor properties of the carbene.<sup>4</sup> Only a small number of NHCs of type **I** possessing an anionic low-coordinate moiety such as enolate,<sup>5</sup> borate,<sup>6</sup> amido<sup>7</sup> and phosphanido,<sup>8</sup> have been reported (Fig. 1). Some of these enabled an additional and/or competing metal binding site, resulting in ligand polytopicity or ambidenticity in bimetallic coordination.<sup>9,10</sup> The anionic bis(NHC) **II** has been reported, obtained *via* reduction of a bis-(imidazol-2-thione-4-yl)phosphane using a large excess of potassium metal;<sup>11</sup> but **II** could neither be isolated nor structurally confirmed.

A new ambidentate Janus-type ligand combining a carbene and an anionic imidate centers within the same

heterocyclic framework was reported by Lavigne *et al.*, appears to be suitable for the directed construction of a variety of homo-and/or heteropolymetallic complexes.<sup>3c</sup> In 2017, a dianionic bis(malNHC)<sup>3d</sup> was reported as a bridging ligand to construct zwitterionic complexes by Tapu *et al.* Furthermore, it served as building block for the preparation of novel organometallic frameworks, not handy with neutral Janus-type bis(NHC)s. Due to their unique electronic properties, these zwitterionic NHC-metal species exhibit potentially valuable advantages such as enhanced catalytic activity and solubility relative to the classical cationic metal complexes of the neutral NHCs.

Recently, we established a new series of tricyclic rigid Janus bis(NHCs) **III**, tuned by  $\text{PR}_n$  moieties in different phosphorus oxidation states, and reported on their use in coinage metal(I) complex chemistry.<sup>12</sup> Thereafter, continuing efforts have been made to establish anionic low-coordinate P-linked bis(NHCs) in order to achieve tuneable electronic communication and redox activity.

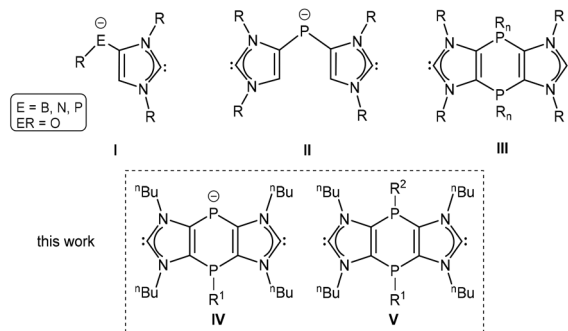


Fig. 1 Anionic mono and bis(NHCs) **I**, **II** and rigid Janus bis(NHCs) **III**–**V** ( $\text{R}, \text{R}^1, \text{R}^2$  = organic substituents).

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Herein, we report on the synthesis of a stable anionic Janus-type tricyclic bis(NHC) **IV** having a P-localized charge, and its use in main group and transition metal chemistry to access, *e.g.*, bis(NHCs) **V**, having a mixed P-substitution pattern.

## Results and discussion

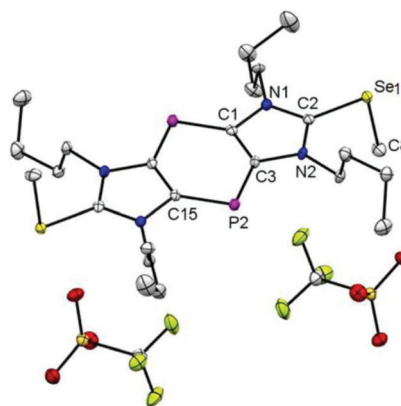
### Synthesis of anionic bis(NHC) **4** via reductive deselenization

Following our recent synthetic protocol,<sup>13</sup> the new tricyclic 1,4-diphosphinine diselone **1** (Scheme 1) was synthesized *via* mild reduction of the *P*-Cl substituted 1,4-dihydro-1,4-diphosphinine<sup>12</sup> precursor and, finally, isolated as a deep violet solid. According to TD-DFT calculations, the colour of this electron-delocalized compound could be attributed to a HOMO–LUMO transition (details in the ESI†). Diselone **1** has been treated with 2 equivalents of trifluoromethyl methylsulfonate (MeOTf) in dichloromethane to afford the doubly *Se*-methylated salt **2** which was isolated as a yellow solid and fully characterized, including single crystal X-ray diffraction analysis (Fig. 2).

The centrosymmetric molecular structure of **2** has a C2–Se1 bond distance of 1.896(4) Å that is slightly elongated compared to 1.8240(16) Å for C=Se in the precursor **1** (ESI†), but remains significantly shorter than the 1.952(4) Å Se–C8 bond to the methyl group.

Doubly *Se*-methylated salt **2** was subjected to reductive deselenization with NaBH<sub>4</sub> in the presence of [2.2.2]-cryptand in methanol to afford (somewhat surprisingly) bis(imidazolium) salt **3** (Scheme 1) with an anionic phosphorus centre. Salt **3**, obtained in pure form *via* extraction with dichloromethane followed by washing with diethylether, showed two resonance signals in a 1 : 1 ratio in the <sup>31</sup>P NMR spectrum at 20.1 ppm (POMe) and –67.3 ppm (anionic P), but no <sup>3</sup>J<sub>(P,P)</sub> coupling (Table 1). Further confirmation for **3** was obtained from NMR and MS experiments as well as elemental analysis. DFT calculations, performed on *N*-Me model compounds,<sup>14</sup> reveal that the aromatic character of the middle ring in **3'** is lower (NICS(0) = –4.9) than in **1'**<sup>14</sup> ((NICS(0) = –8.1) or **2'** (NICS(0) = –10.0); nevertheless some – apparently hyperconjugative – cyclic conjugation is still operative. In contrast, the outer rings retain high aromatic character (NICS(0) = –11.1).

To access the first example of an anionic P-functional bis(NHC), deprotonation of the bis(imidazolium) salt **3** was performed in THF using two equivalents of KHMDS (Scheme 2).



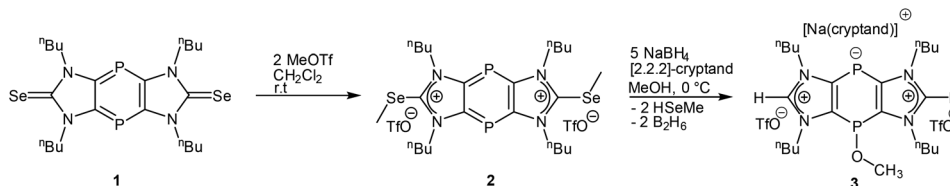
**Fig. 2** Molecular structure of compound **2** (ellipsoids at the 50% probability-substituted level); hydrogen atoms are omitted for clarity. Selected bond lengths [Å] and angles [°]: C2–Se1 1.896(4), Se1–C8 1.952(4), C2–N2 1.339(5), N2–C3 1.397(5), C3–P2 1.736(4), C1–C3 1.412(6); N1–C2–N2 108.8(3), C3–P2–C15 96.3(2).

**Table 1** <sup>31</sup>P{<sup>1</sup>H} and <sup>13</sup>C{<sup>1</sup>H} NMR data for **3**, **4**, **5**<sup>cis/trans</sup> (1 : 0.3) and **6**<sup>cis/trans</sup> (1 : 0.3)

	δ( <sup>31</sup> P)/ppm (CD <sub>2</sub> Cl <sub>2</sub> )	δ( <sup>13</sup> C)/ppm (CD <sub>2</sub> Cl <sub>2</sub> ) <sup>b</sup>
<b>3</b>	20.1 (s), –67.3 (s)	137.3 (d, <sup>3</sup> J <sub>P,C</sub> = 4.5 Hz)
<b>4</b> <sup>a</sup>	25.2 (s), –74.8 (s)	208.8 (br)
<b>5</b> <sup>cis/trans</sup>	δ = –71.58 (d, <sup>3</sup> J <sub>P,H</sub> = 5.2 Hz), –66.23 (d, <sup>3</sup> J <sub>P,H</sub> = 4.9 Hz), 39.57 (br), 43.7 (br)	142.4 (br), 143.37 (br)
<b>6</b> <sup>cis/trans</sup>	δ = –74.0 (d, <sup>2</sup> J <sub>P,H</sub> = 4.8 Hz), –68.6 (d, <sup>2</sup> J <sub>P,H</sub> = 3.7 Hz), 37.2 (d, <sup>3</sup> J <sub>P,H</sub> = 4.6 Hz), 41.3 (d, <sup>3</sup> J <sub>P,H</sub> = 3.8 Hz)	224.2, 223.8, (t, <sup>3</sup> J <sub>P,C</sub> = 2.7 Hz)

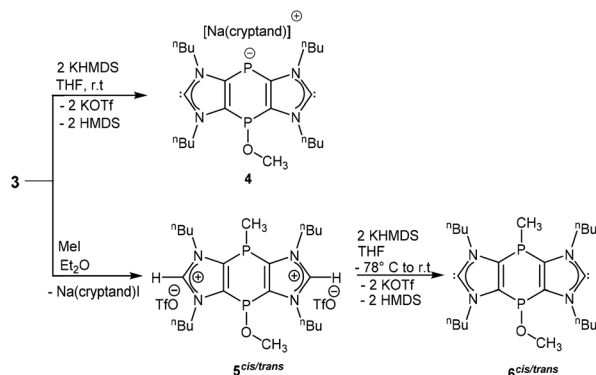
<sup>a</sup> In case of **4** and **6**<sup>cis/trans</sup> (1 : 0.4) THF-d<sub>8</sub>. <sup>b</sup> C<sup>2</sup> carbon.

After extraction with THF/diethyl ether (1 : 1.5), compound **4** was isolated as a dark orange solid which has two resonances in its <sup>31</sup>P NMR spectrum at 25.2 ppm (P-OMe) and –74.8 ppm (anionic P). The <sup>1</sup>H-NMR spectrum confirms the absence of the C<sup>2</sup>–H proton and the <sup>13</sup>C{<sup>1</sup>H} NMR spectrum a broad resonance at 208.8 ppm assigned to the C<sup>2</sup> atom of dicarbene **4**. The proposed constitution of **4** is also supported by HR-MS (negative ESI; exp. 449.2605 vs. calc. 449.2607). In order to establish the stability of the carbene **4**, an isodesmic reaction (see ESI†) yields 113.3 kcal mol<sup>–1</sup> stabilization for **4'**.<sup>14</sup> This is very similar to our earlier reported 111.1 kcal mol<sup>–1</sup> value<sup>12</sup> for **III** (R: Me, R': *cis*-NET<sub>2</sub>), indicating that the carbene character



**Scheme 1** Synthesis of doubly *Se*-methylated salt **2** and bis(imidazolium) salt **3** starting from tricyclic diselone **1**.





**Scheme 2** Synthesis of anionic bis(NHC) **4**, P<sup>III/III</sup> bis(imidazolium) salts **5<sup>cis/trans</sup>** (1 : 0.3) and mixed substituted bis(NHCs) **6<sup>cis/trans</sup>** (1 : 0.3).

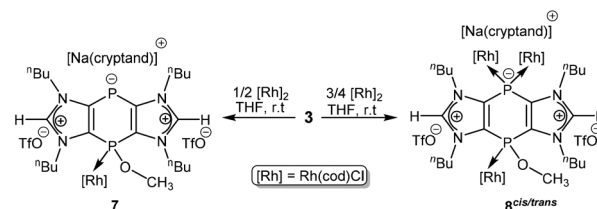
in the anionic tri-cycle is virtually unchanged. Compared to **3'**, the aromaticity of the middle ring is slightly higher (NICS(0) = −5.5), while that of the outer rings is lower (NICS(0) = −9.0), as is usual for NHCs if compared to imidazolium salts.<sup>15</sup>

#### Formation of mixed substituted P<sup>III/III</sup> bridged bis(NHCs) **6<sup>cis/trans</sup>**

We then targeted to use the P-anionic functionality to access bis(NHCs) with a mixed P-substitution pattern. Therefore, **3** was treated with MeI in diethyl ether at −80 °C which resulted in a clean formation of bis(imidazolium) salt **5<sup>cis/trans</sup>** with an isomeric ratio of 1 : 0.3 (signals not assigned) (Scheme 2). The <sup>31</sup>P{<sup>1</sup>H} NMR spectrum of the reaction mixture showed two sets of signals for two isomers at δ(P−CH<sub>3</sub>) = −71.6 (d, <sup>3</sup>J<sub>P,H</sub> = 5.0 Hz), −66.2 (d, <sup>3</sup>J<sub>P,H</sub> = 5.0 Hz) and δ(P−OCH<sub>3</sub>) = 39.6 (br), 43.7 (br) (ratio 1 : 0.4). The isolated mixture of **5<sup>cis/trans</sup>** (1 : 0.3) was subsequently deprotonated using two equiv. of KHMDS in THF to afford the mixed-substituted bis(NHCs) **6<sup>cis/trans</sup>** (ratio 1 : 0.3).

Clear evidence for the latter came from the <sup>1</sup>H NMR spectrum of this mixture, since the former C<sup>2</sup>−H signal (δ = 9.47 (t, <sup>3</sup>J<sub>P,H</sub> = 3.03 Hz), 9.56 (br)) of **5<sup>cis/trans</sup>** were absent. This was further supported by the <sup>13</sup>C{<sup>1</sup>H} NMR spectrum as the characteristic downfield signals for the C<sup>2</sup>-nuclei were found (Table 1), revealing the formation of the bis(NHCs) **6<sup>cis/trans</sup>** (ratio 1 : 0.3). **6<sup>trans</sup>** and **6<sup>cis</sup>** are computed to have high stability (the isodesmic reaction energies are 108.8 for **6<sup>trans</sup>**<sup>14</sup> and 108.6 kcal mol<sup>−1</sup> for **6<sup>cis</sup>**). The aromatic character of the middle ring is significantly lower (NICS(0) = −0.1), indicating that when both phosphorus centres are saturated the central ring loses its aromaticity, whilst that of the outer rings (NICS(0) = −9.7) remains high.

Initial coordination chemistry experiments were undertaken to explore the reactivity of tricyclic bis(imidazolium) salt **3**. Thus, reaction with a half equiv. of [Rh(cod)Cl]<sub>2</sub> affords exclusively the mono rhodium(i) complex **7** (Scheme 3). The coordination of **3** to the Rh(cod)Cl fragment is confirmed by the <sup>31</sup>P NMR spectrum as resonances of **7** appeared at high-field (−70.6 (s, anionic P) and 47.5 (d) ppm) having a rhodium–phosphorus coupling of <sup>1</sup>J<sub>Rh,P</sub> = 192.4 Hz, assigned



**Scheme 3** Synthesis of mononuclear rhodium complex **7** and trinuclear rhodium complexes **8<sup>cis/trans</sup>** (1 : 0.7).

to the neutral phosphorus atom. This coordination mode is surprising since the HOMO of **3'** is located at the dicoordinate (anionic) phosphorus atom, and the involved tricoordinate phosphorus, largely representing HOMO−1, is lower in energy by as much as 1.9 eV (see ESI†). Nevertheless, our calculations on the isomeric complexes of **7'** showed that the favoured coordination site is indeed the tricoordinate phosphorus, **7'** being more stable than the possible isomers (*cis* and *trans*) by more than 5 kcal mol<sup>−1</sup> (see ESI†). To examine how many Rh(i) fragments could be coordinated to **3**, it was then treated with 1.5 equiv. of the rhodium dimer which afforded selectively an isomeric mixture of trinuclear phosphanido complexes **8<sup>cis/trans</sup>** (ratio 0.9 : 1) (Scheme 3). Upon coordination, two sets of signals are present in the <sup>31</sup>P NMR spectrum at 65.0 (d, <sup>1</sup>J<sub>Rh,P</sub> = 195.6 Hz) (minor)/64.1 (dd, <sup>1</sup>J<sub>Rh,P</sub> = 200.1 Hz, <sup>3</sup>J<sub>P,P</sub> = 5.3 Hz) (major) and −120.3 (t br, <sup>1</sup>J<sub>Rh,P</sub> = 126.9 Hz) (major)/−123.4 (t, <sup>1</sup>J<sub>Rh,P</sub> = 126.9 Hz) (minor).

#### Cyclic voltametric studies supported by DFT calculations

The electrochemical properties of bis(NHCs) **4** and **6<sup>cis/trans</sup>** (1 : 0.3) were investigated by cyclic voltammetry (CV) in THF (0.2 M [*n*Bu<sub>4</sub>N][PF<sub>6</sub>]) at gold ceramic screen printed electrodes (Au CSPE) in an Ar-filled glove box. Voltammetric data were measured on solutions containing 1.0 mM analyte and representative CVs are presented in Fig. 3 (and in greater detail in the ESI†). The observed behaviour of **6<sup>cis/trans</sup>** (Fig. 3b) is reminiscent of dicarbenes **III** (R<sub>n</sub> = P<sup>III</sup>NEt<sub>2</sub>)<sup>12</sup> and we therefore start the analysis here.

There are a series of facile, chemically irreversible oxidations with anodic peak potentials  $E_p^{a1} = -0.29$  V,  $E_p^{a2} = -0.07$  V and  $E_p^{a3} = +0.65$ . The processes are nonetheless repeatable and stable to scanning first in anodic or cathodic directions, and continue right up to the anodic potential limit at around +1.0 V. This is consistent with oxidations involving the carbene σ(p)-centred HOMO and HOMO−1 as determined from B3LYP/6-31+G\*\*/M06-2X/6-31+G\* computations undertaken on the model structures **6<sup>cis/trans</sup>**<sup>15</sup> (Fig. 4). The more positive first oxidations  $E_p^{a1} = -0.29$  V for **6<sup>cis/trans</sup>** compared to **III** (R<sub>n</sub> = P<sup>III</sup>NEt<sub>2</sub>), for which  $E_p^{a1} = -0.61$  V (ref. 12) reflect the lower lying HOMO energies of −5.93/−5.92 eV of the former versus −5.78/−5.77 eV for the latter; as before, we are not able to identify separate CV processes for the two geometrical isomers that are known to co-exist in solution. The true electrochemical reductions for the two types of P<sup>III</sup> dicarbenes both occur at



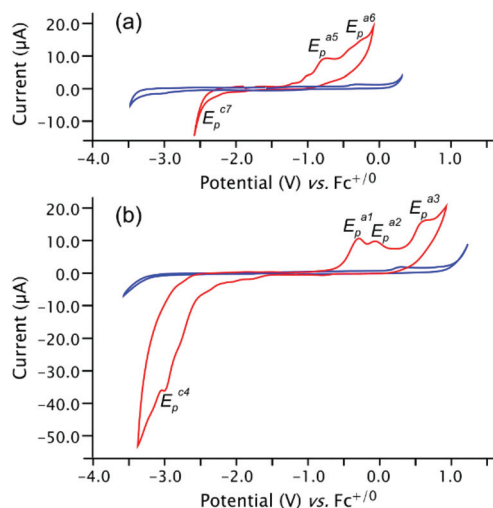


Fig. 3 CV diagrams (a) of **4**, and (b) of **6<sup>cis/trans</sup>** (scan rate: 200 mV s<sup>-1</sup>). In each case, the blue trace indicates the solvent/electrolyte background (0.2 M) [*n*Bu<sub>4</sub>N][PF<sub>6</sub>] in (THF), and the red that from a 1.0 mM solution of analyte.

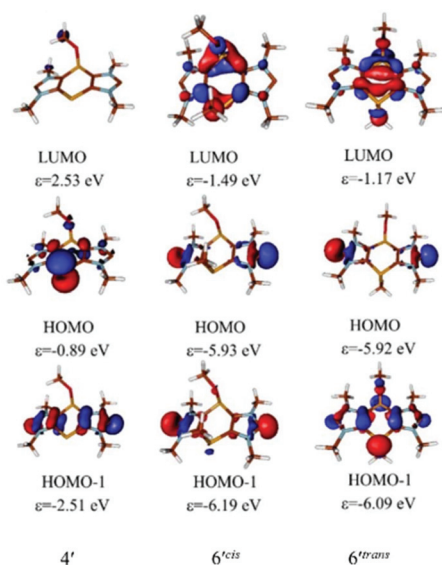
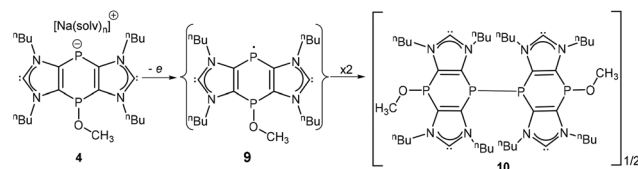


Fig. 4 FMO topologies and energies for the model (R = CH<sub>3</sub>) calculated structures of **4'**, **6<sup>cis</sup>** and **6<sup>trans</sup>** at the B3LYP/6-31+G\*/M06-2X/6-31+G\* level of theory (PCM solvent model).

very low potentials (less than -3.5 V) and probably cannot be measured accurately; the shoulder on the main reduction wave in the CVs (*e.g.*  $E_p^{c1}$ ) are likely from breakdown products of the IRR oxidations (see ESI† for further explanations).

In contrast to this established behaviour of the diphosphinine dicarbenes, the CVs measured on **4** are less well defined although the onset of oxidation is definitely lower in potential than in **6<sup>cis/trans</sup>** with  $E_p^{a5} = -0.74$  V (Fig. 3a). The computed HOMO of model system **4'** is very different (Fig. 4) and is essentially localized at the anionic P atom, and, as expected for an anion, is also at a much higher computed energy (-0.89



Scheme 4 One electron reduction of **4** to neutral **9** with rapid dimerization to **10**.

eV). Experimentally, the onset of oxidation does not have the expected well-separated peak for a first 1e oxidation of such a localized MO, followed by further processes after a considerable gap.

A plausible explanation for the observed behaviour is that 1e oxidation gives P-centred radical bis(NHC) **9** which rapidly dimerizes to give tetrakis NHC **10** (Scheme 4). Related tricyclic dithione P-radicals have been shown to be very short-lived furnishing structurally verified dimers with a P-P bond.<sup>16</sup> On this view, the remaining processes such as  $E_p^{a6} = -0.30$  V and  $E_p^{c1} = -2.5$  V recorded in the CVs of **4** (Fig. 3a) are best understood as redox processes of such a tetrakis NHC.

## Experimental section

### Experimental details and devices

All experiments were done under an argon atmosphere, using common Schlenk techniques and dry solvents. Tetrahydrofuran, *n*-pentane and diethyl ether were dried over sodium wire/benzophenone and further purified by subsequent distillation. The precursor for **1**, the 1,4-dichloro-1,4-dihydro-1,4-diphosphinine, was synthesized using standard protocols.<sup>12</sup> All NMR spectra were recorded on a Bruker AX-300 spectrometer (300.1 MHz for <sup>1</sup>H, 75.5 MHz for <sup>13</sup>C, and 121.5 MHz for <sup>31</sup>P) and pollux-500 (500.1 MHz for <sup>1</sup>H, 125.75 MHz for <sup>13</sup>C, and 500.0 MHz for <sup>31</sup>P) spectrometer. The <sup>1</sup>H and <sup>13</sup>C NMR spectra were referenced to the residual proton resonances and the <sup>13</sup>C NMR signals of the deuterated solvents and <sup>31</sup>P to 85% H<sub>3</sub>PO<sub>4</sub> as external standard, respectively. Elemental analyses were carried out on a Vario EL gas chromatograph. Melting points were determined in one-side melted off capillaries using a Büchi Type S or a Carl Roth Type MPM-2 apparatus, they are uncorrected Mass spectrometric data were collected on a Kratos MS 50 spectrometer using EI, 70 eV. IR spectra of all compounds were recorded on a Thermo IR spectrometer with an attenuated total reflection (ATR) attachment. The X-ray analyses were performed on a Bruker APEX-II CCD or a Bruker X8-KappaApexII type diffractometer at 100(2) K. The structures were solved by direct methods refined by full-matrix least-squares technique in anisotropic approximation for non-hydrogen atoms using SHELXS97 and SHELXL97<sup>17</sup> program packages. Crystallographic data for the structures reported in this paper have been deposited with the Cambridge Crystallographic Data Centre as supplementary publication no. CCDC 2032731 (**1**), 2032732 (**2**).†





## Synthesis of compounds

1. To a clear solution of 1,4-bis(diethylamino)-1,4-dihydro-1,4-diphosphinine<sup>12</sup> (2.5 g, 3.5 mmol) in dichloromethane, PCl<sub>3</sub> (0.61 mL, 6.9 mmol) was added and stirred for 4 hours at −40 °C. The reaction mixture was then warmed to room temperature and tris(*n*-butyl)phosphane (0.34 mL, 1.4 mmol) was added in a dropwise manner. After 10 minutes stirring, a colour change of the solution from orange to violet was observed. After concentrating the reaction mixture under reduced pressure, the residue was filtered *via* a silica® bed with diethyl ether and toluene mixture (1 : 1). It was then dried under reduced pressure (6 × 10<sup>−3</sup> mbar) and washed with *n*-pentane (3 × 10 mL) to get rid of the aminophosphane Et<sub>2</sub>NPCl<sub>2</sub>. Finally, the solution was concentrated *in vacuo* (6 × 10<sup>−3</sup> mbar) to get **1** as pure compound. Yield: 1.2 g (2.1 mmol) 61%; violet solid. M.p. 223 °C. <sup>1</sup>H NMR (300.1 MHz, C<sub>6</sub>D<sub>6</sub>, 25 °C): δ = 0.8 (t, 12H, <sup>3</sup>J<sub>H,H</sub> = 7.3 Hz, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.2–1.4 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.8–1.9 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 4.4–4.4 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me). <sup>13</sup>C{<sup>1</sup>H} NMR (75.5 MHz, C<sub>6</sub>D<sub>6</sub>, 25 °C): δ = 13.55 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 19.9 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 28.9 (t, <sup>3</sup>J<sub>P,C</sub> = 1.6 Hz, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 48.1 (t, <sup>3</sup>J<sub>P,C</sub> = 4.6 Hz, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 152.2 (t, <sup>1</sup>J<sub>P,C</sub> = 23.1 Hz, P–C of the middle ring), 168.32 (br, C<sup>2</sup>). <sup>31</sup>P{<sup>1</sup>H} NMR (121.5 MHz, C<sub>6</sub>D<sub>6</sub>, 25 °C): δ = 78.2 (s). <sup>77</sup>Se NMR (57.28 MHz, CDCl<sub>3</sub>): 179.4 (s). IR [cm<sup>−1</sup>]:  $\tilde{\nu}$  = 2998.1 (v), 2752.8 (m), 2654.0 (m), 1487.0 (s), 1289.1 (s), 1201.5 (s), 1098.2 (m), 964.0 (s). MS (EI, 70 eV): *m/z* (%) = 578.0 (100) [M]<sup>+</sup>, 498.1 (20) [M – Se]<sup>+</sup>. UV-vis (CH<sub>2</sub>Cl<sub>2</sub>): λ<sub>max</sub> [nm] (abs.): 296 (1.375), 383 (0.156), 554(0.139). EA [%]: theor./exp. C 45.84/45.64, H 6.30/6.46; N 9.72/9.63.

2. 2 equivalents of trifluoromethyl methylsulfonate (0.4 mL, 3.4 mmol) was added to a solution of **1** (1.0 g, 1.7 mmol) in dichloromethane, at room temperature. The reaction mixture was stirred for 2 hours resulting in a color change from violet to light yellow. After concentrating the reaction mixture under reduced pressure (6 × 10<sup>−3</sup> mbar), the residue was washed with *n*-pentane (2 × 5 mL) to get **2** as pure light yellow solid. Yield 1.4 g (1.5 mmol) 91%; light yellow solid. M.p. 96 °C. <sup>1</sup>H NMR (300.1 MHz, CDCl<sub>3</sub>, 25 °C): 1.04 (t, 12H, <sup>3</sup>J<sub>H,H</sub> = 7.4 Hz, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.52–1.62 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 2.05–2.15 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 2.8 (s, 6H, SeMe) 4.89–4.94 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me). <sup>13</sup>C{<sup>1</sup>H} NMR (75.5 MHz, CDCl<sub>3</sub>, 25 °C): δ = 11.72 (s, SeMe), 13.44 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 20.16 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 30.73 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 52.18 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 148.84 (t, <sup>3</sup>J<sub>P,C</sub> = 4.6 Hz, Se–C<sup>2</sup>), 154.94 (t, <sup>1</sup>J<sub>P,C</sub> = 26.1 Hz, P–C of the middle ring). <sup>31</sup>P NMR (121.5 MHz, CDCl<sub>3</sub>, 25 °C): δ = 119.95 (s). <sup>77</sup>Se NMR (57.28 MHz, CDCl<sub>3</sub>): 138.89 (s). IR [cm<sup>−1</sup>]:  $\tilde{\nu}$  = 3009.1 (v), 2992.8 (m), 2954.0 (m), 1623.0 (w), 1529.1 (w), 1461.5 (m), 1236.2 (m), 1201.3 (w), 1075.4 (m), 1032.2 (v), 974.0 (s). Pos. ESI-MS: [C<sub>25</sub>H<sub>42</sub>F<sub>3</sub>N<sub>4</sub>O<sub>3</sub>P<sub>2</sub>SSe]<sup>+</sup> calcd (found) 757.0728 (757.0750). neg. ESI-MS: TfO<sup>−</sup> theor./exp. 148.9(149.5). EA [%]: theor./exp. C 34.52/34.33, H 4.68/4.71; N 6.19/6.0, S 7.08/6.99.

3. To a solution of **2** (1.5 g, 1.6 mmol) in methanol, 5 equivalents of sodium tetrahydridoborate (0.3 g, 8.2 mmol)

and one equivalent of [2.2.2]-cryptand was added as solid at 0 °C. The reaction mixture turned to orange-red with strong odour due to a liberation of methylselane (HMeSe). The solution was then concentrated *in vacuo* (6 × 10<sup>−3</sup> mbar) after 30 minutes stirring. Extraction was done with dichloromethane followed by washing with diethyl ether (2 × 5 mL) to get **3** as pure orange red solid. Yield: 1.2 g (1.04 mmol) 65%; red orange solid, M.p. 142 °C <sup>1</sup>H NMR (300.1 MHz, CD<sub>2</sub>Cl<sub>2</sub>, 25 °C): δ = 1.0 (t, 12H, <sup>3</sup>J<sub>H,H</sub> = 7.3 Hz, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.3–1.5 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.9–2.1 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 2.4 (d, 3H, <sup>3</sup>J<sub>P,H</sub> = 7.3 Hz, O–Me) 2.6 (t, 12H, cryptand), 3.6 (t, 12H, cryptand), 3.7 (s, 12H, cryptand), 4.1–4.6 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 8.9 (t, 2H, <sup>4</sup>J<sub>P,H</sub> = 1.7 Hz, C<sup>2</sup>–H). <sup>13</sup>C{<sup>1</sup>H} NMR (75.5 MHz, CD<sub>2</sub>Cl<sub>2</sub>, 25 °C): δ = 13.2 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 19.6 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 30.2 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 47.8 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 67.5 (s, cryptand), 68.5 (s, cryptand), 120.9 (d, <sup>2</sup>J<sub>P,C</sub> = 7.3 Hz, O–CH<sub>3</sub>), 121.7 (q, <sup>1</sup>J<sub>P,F</sub> = 321.0 Hz, CF<sub>3</sub>), 137.2 (d, <sup>3</sup>J<sub>P,C</sub> = 4.5 Hz, H–C<sup>2</sup>), 155.5 (ddd, <sup>1/2</sup>J<sub>P,C</sub> = 47.0 Hz, C<sup>4/5</sup>). <sup>31</sup>P NMR (121.5 MHz, CD<sub>2</sub>Cl<sub>2</sub>, 25 °C): δ = 20.12 (P–OMe), −67.34 (anion P). IR [cm<sup>−1</sup>]:  $\tilde{\nu}$  = 2984 (v), 2921.8 (m), 2894.0 (m), 1542.0 (w), 498.1 (w), 1423.5 (m), 1246.2 (m), 1206.3 (w), 1012.4 (m), 968.4 (s). Pos. ESI-MS: *m/z* (%) = 451.3 (100) [M]<sup>+</sup>, 399.1 (97) [Na(C<sub>18</sub>N<sub>2</sub>H<sub>36</sub>O<sub>6</sub>)]<sup>+</sup>. HRMS: [C<sub>23</sub>H<sub>41</sub>N<sub>4</sub>OP<sub>2</sub>]<sup>+</sup> theor./exp. 451.2750 (451.2754). UV/vis (CH<sub>2</sub>Cl<sub>2</sub>): λ [nm] (abs.): 346 (0.791). EA [%]: theor./exp. C 44.94/45.12, H 6.75/6.64; N 7.31/6.74, S 5.58/5.68.

4. A solution of potassium hexamethyldisilazide (KHMDs) (0.7 g, 3.5 mmol) in 5 mL of THF was added dropwise to a solution of **3** (2 g, 1.7 mmol) in 10 mL of THF at room temperature. After 1 h, all volatiles were removed *in vacuo* (6 × 10<sup>−3</sup> mbar). The residue was washed (twice) with diethyl ether followed by extraction with mixture of THF and diethyl ether (1 : 1.5) to remove the potassium triflate. After concentrating the extracted solution, the product **4** was obtained as dark orange solid. Yield: 1.1 g (1.3 mmol) 76%; Dark orange. M. p. 207 °C. <sup>1</sup>H NMR (500.1 MHz, THF, 25 °C): δ = 0.9 (t, 12H, <sup>3</sup>J<sub>H,H</sub> = 6.5 Hz, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.2–1.3 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 1.8–1.9 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 2.1 (d, 3H, <sup>4</sup>J<sub>P,H</sub> = 6.3 Hz, O–Me) 2.5 (t, 12H, cryptand), 3.5 (t, 12H, cryptand), 3.6 (s, 12H, cryptand), 3.8–4.5 (m, 8H, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me). <sup>13</sup>C{<sup>1</sup>H} NMR (75.5 MHz, THF-d<sub>8</sub>, 25 °C): δ = 13.5 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 20.1 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 32.3 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 47.5 (s, NCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Me), 67.6 (s, cryptand), 70.4 (s, cryptand), 120.2 (d, <sup>2</sup>J<sub>P,C</sub> = 7.5 Hz, O–CH<sub>3</sub>), 154.2 (d, <sup>1/2</sup>J<sub>P,C</sub> = 43.9 Hz, C<sup>4/5</sup>), 215.5 (d, <sup>1</sup>J<sub>P,C</sub> = 2.7, C<sup>2</sup>). <sup>31</sup>P NMR (500.0 MHz, THF-d<sub>8</sub>, 25 °C): δ = 25.15 (P–OMe), −74.78 (anion P). IR [cm<sup>−1</sup>]:  $\tilde{\nu}$  = 3191.2 (w), 3045.0 (m), 2975.7 (m), 1623.4 (m), 1436.3 (m), 1375.3 (m), 1245.7 (s), 1184.8 (s), 1016.5 (s), 968.5 (s). Neg. ESI-MS: *m/z* (%) = 449.261 (15) [M]<sup>−</sup>, HRMS: [C<sub>23</sub>H<sub>39</sub>N<sub>4</sub>OP<sub>2</sub>]<sup>−</sup> theor./exp. 449.2605 (449.2608). UV/vis (THF): λ [nm] (abs.): 407 (0.517).

<sup>5</sup><sub>cis/trans</sub>. To a suspension of **3** (2.5 g, 2.1 mmol) in 50 mL of diethyl ether, methyl iodide (1.35 mL, 2.1 mmol) was added dropwise at −80 °C. The reaction mixture was stirred for 18 hours and warmed to room temperature. All volatiles were removed *in vacuo* (6 × 10<sup>−3</sup> mbar). Residue was extracted with



dichloromethane followed by washing (twice) with diethyl ether. The solvent was removed under vacuum ( $6 \times 10^{-3}$  mbar) which resulted in a pure colorless liquid. Yield: 1.2 g (1.6 mmol) 75%; colorless liquid. Ratio of two isomers 1:0.3. M.p. – (liquid at r.t.).  $^1\text{H}$  NMR (300.1 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 1.04, 1.06 (t, 12H,  $^3J_{\text{H,H}} = 7.4$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.41–1.54 (m, 8H,  $\text{PCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.67 (d,  $^2J_{\text{P,H}} = 5.2$  Hz,  $\text{P-Me}$ ), 1.87 (d,  $^2J_{\text{P,H}} = 6.8$  Hz,  $\text{O-Me}$ ); 1.92–2.11 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 4.28–4.57 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 9.47 (t, 2H,  $^3J_{\text{P,H}} = 3.03$  Hz,  $\text{C}^2\text{-H}$ ), 9.56 (brs,  $\text{C}^2\text{-H}$ ) 2nd isomer.  $^{13}\text{C}\{^1\text{H}\}$  NMR (75.5 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 13.08, 13.10 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$  of two isomers), 19.52 (br,  $\text{P-Me}$ ), 19.68 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 29.68 (br,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 31.78 (d,  $^3J_{\text{P,C}} = 2.6$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 32.23 (d,  $^3J_{\text{P,C}} = 2.4$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ) (2nd isomer), 49.13 (ddd,  $^3J_{\text{P,C}} = 9.1$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 49.96 (ddd,  $^3J_{\text{P,C}} = 8.2$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ) (2nd isomer), 122.84 (q,  $^1J_{\text{P,F}} = 319.5$  Hz,  $\text{CF}_3$ ), 131.54 (d,  $^2J_{\text{P,C}} = 9.5$  Hz,  $\text{O-CH}_3$ ), 135.38 (ddd,  $^{1/2}J_{\text{P,C}} = 3.7$  Hz,  $\text{P-C}$  of the middle ring), 135.72 (t,  $^{1/2}J_{\text{P,C}} = 3.0$  Hz,  $\text{P-C}$  of the middle ring) (2nd isomer), 142.38 (br,  $\text{H-C}^2$ ), 143.37 (br,  $\text{H-C}^2$ ) 2nd isomer.  $^{31}\text{P}$  NMR (500.0 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = –71.58 (d,  $^3J_{\text{P,H}} = 5.2$  Hz,  $\text{P-Me}$ ), –66.23 (d,  $^3J_{\text{P,H}} = 4.9$  Hz,  $\text{P-Me}$ ); 39.57 (br,  $\text{P-OMe}$ ) & 43.7 (br,  $\text{P-OMe}$ ). IR [ $\text{cm}^{-1}$ ]:  $\tilde{\nu}$  = 3204.7 (m), 3145 (m), 2975.5 (w), 2768.8 (m), 1534.3 (w), 1445.3 (s), 1317.7 (m), 1206.8 (m), 1046.9 (m), 1009.5 (s), 921.5 (s). Pos. ESI-MS:  $m/z$  (%) = 615.251 (54)  $[\text{M} - \text{TfO}]^+$ ; HRMS:  $[\text{C}_{25}\text{H}_{44}\text{F}_3\text{N}_4\text{O}_4\text{P}_2\text{S}]^+$  theor./exp. 615.2505 (615.2511). EA [%]: theor./exp. C 40.84/40.53, H 5.80/5.93, N 7.33/7.36.

**6<sup>cis/trans</sup>**. A solution of potassium hexamethyldisilazide (KHMDs) (1.04 g, 5.2 mmol) in 5 mL of THF was added dropwise to a solution of **5<sup>cis/trans</sup>** (2.0 g, 2.6 mmol) in 10 mL of THF at room temperature. After 1 h, all volatiles were removed *in vacuo* ( $6 \times 10^{-3}$  mbar). Residue was extracted with diethyl ether to remove potassium triflate using filtering cannulation. After concentrating the extracted solution, the product **6<sup>cis/trans</sup>** was obtained as yellow liquid. Yield: 0.92 g (1.9 mmol) 76%; (ratio of two isomers 1:0.30). M.p. – (liquid at r.t.).  $^1\text{H}$  NMR (500.1 MHz,  $\text{THF-d}_8$ , 25 °C):  $\delta$  = 0.8, 1.1 (t, 12H,  $^3J_{\text{H,H}} = 7.1$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.2 (d,  $^2J_{\text{P,H}} = 5.3$  Hz,  $\text{P-Me}$ ), 1.2–1.4 (m, 8H,  $\text{PCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.9–2.1 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 2.7 (d,  $^2J_{\text{P,H}} = 7.2$  Hz,  $\text{O-Me}$ ), 3.9–4.2 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 4.3–4.5 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ; 2<sup>nd</sup> isomer).  $^{13}\text{C}\{^1\text{H}\}$  NMR (125.75 MHz,  $\text{THF-d}_8$ , 25 °C):  $\delta$  = 12.9, 12.8 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$  of two isomers), 19.9 (s,  $\text{P-Me}$ ), 23.7 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 25.7 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 32.8 (d,  $^3J_{\text{P,C}} = 2.1$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 32.9 (d,  $^3J_{\text{P,C}} = 2.2$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ; 2<sup>nd</sup> isomer), 48.5 (ddd,  $^3J_{\text{P,C}} = 9.6$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 49.6 (ddd,  $^3J_{\text{P,C}} = 8.3$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ; 2<sup>nd</sup> isomer), 118.5 (d,  $^2J_{\text{P,C}} = 9.2$  Hz,  $\text{O-CH}_3$ ), 131.4 (br,  $\text{P-C}$  of the middle ring), 132.2 (d,  $^{1/2}J_{\text{P,C}} = 2.5$  Hz,  $\text{P-C}$  of the middle ring; 2<sup>nd</sup> isomer), 223.4 ((t,  $^3J_{\text{P,C}} = 2.7$  Hz,  $\text{C}^2$ ), 224.2 (t,  $^3J_{\text{P,C}} = 2.7$  Hz,  $\text{C}^2$ ; 2<sup>nd</sup> isomer).  $^{31}\text{P}$  NMR (500 MHz,  $\text{THF-d}_8$ , 25 °C):  $\delta$  = –74.0 (d,  $^3J_{\text{P,H}} = 4.8$  Hz,  $\text{P-Me}$ ), –68.6 (d,  $^3J_{\text{P,H}} = 3.7$  Hz,  $\text{P-Me}$ ); 41.3 (d,  $^3J_{\text{P,H}} = 3.8$  Hz) & 37.2 (d,  $^3J_{\text{P,H}} = 4.6$  Hz). IR [ $\text{cm}^{-1}$ ]:  $\tilde{\nu}$  = 2992.2 (m), 2962.5 (m), 2842.2 (w), 1501.2 (w), 1472.4 (m), 1415.8 (s), 1367.3 (s), 1146.0 (s), 1052.1 (m), 986.5 (m). Pos.

ESI-MS:  $m/z$  (%) = 465.290 (31)  $[\text{M} + \text{H}]^+ [\text{C}_{24}\text{H}_{43}\text{ON}_4\text{P}_2]^{1+}$  theor./exp. 465.2907 (465.2909). UV/vis ( $\text{CH}_2\text{Cl}_2$ ):  $\lambda$  [nm] (abs.): 347 (0.124).

**7**. To a solution of **3** (2.0 g, 1.7 mmol) in dichloromethane,  $[\text{Rh}(\text{cod})\text{Cl}]_2$  (0.43 g, 0.87 mmol) was added as solid at ambient temperature. Reaction mixture was stirred for 6 hours, at which point volatiles were removed *in vacuo* ( $6 \times 10^{-3}$  mbar). Residue was washed (twice) with diethyl ether and subsequent drying *in vacuo* ( $6 \times 10^{-3}$  mbar) resulted in an orange solid. Yield: 2.1 g (1.5 mmol) 88%; Orange solid; M. p. 82 °C.  $^1\text{H}$  NMR (300.1 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 0.99 (t, 12H,  $^3J_{\text{H,H}} = 7.2$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.26–1.57 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.92 (m, 4H, cod), 1.99–2.21 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 2.36 (m, 4H, cod), 2.56 (d, 8H,  $^3J_{\text{P,H}} = 10.2$  Hz,  $\text{O-CH}_3$ ), 3.68 (m, 2H, cod), 4.06–4.21 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 5.37 (m, 2H, cod), 8.99 (brs,  $\text{C}^2\text{-H}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (125.75 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 13.46 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 19.95 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 28.52 (s, cod), 30.14 (br,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 33.27 (s, cod), 49.48 ((d,  $^3J_{\text{P,C}} = 7.2$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 69.6 (br, cod), 73.21 (d,  $^1J_{\text{Rh,C}} = 11.2$  Hz, cod), 122.13 (q,  $^1J_{\text{P,F}} = 322.8$  Hz,  $\text{CF}_3$ ), 108.68 (d,  $^2J_{\text{P,C}} = 11.1$  Hz,  $\text{O-CH}_3$ ), 155.96 (br,  $\text{P-C}$  of the middle ring), 156.35 (d,  $^{1/2}J_{\text{P,C}} = 44.2$  Hz,  $\text{P-C}$  of the middle ring), 137.96 (br,  $\text{H-C}^2$ ).  $^{31}\text{P}$  NMR (500.0 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = –70.56 (s), 47.45 (d,  $^1J_{\text{Rh,P}} = 188.2$  Hz). IR [ $\text{cm}^{-1}$ ]:  $\tilde{\nu}$  = 2975.2 (m), 2931.4 (m), 2840.1 (w), 1511.4 (s), 1480.4 (m), 1398.1 (m), 1247.8 (s), 1175.5 (m), 1129.8 (m), 1007.1 (s), 910.2 (s). Neg. ESI-MS:  $m/z$  (%) = 995.151 (29)  $[\text{M}]^+$ . HRMS:  $[\text{C}_{33}\text{H}_{53}\text{ClF}_3\text{N}_4\text{O}_7\text{P}_2\text{RhS}_2\text{F}_6]^{+}$  theor./exp. 995.1473 (995.1494). UV/vis ( $\text{CH}_2\text{Cl}_2$ ):  $\lambda$  [nm] (abs.): 299 (0.924). EA [%]: theor./exp. C 43.89/43.89, H 6.43/6.73, N 6.02/76.03, S 4.59/4.34.

**8<sup>cis/trans</sup>**. To a solution of **3** (1.5 g, 1.3 mmol) in dichloromethane, 1.5 equivalent of  $[\text{Rh}(\text{cod})\text{Cl}]_2$  (0.97 g, 1.9 mmol) was added as solid at ambient temperature. Reaction mixture was stirred at ambient temperature for 12 hours. Solvent was then removed *in vacuo* ( $6 \times 10^{-3}$  mbar) and the residue washed (twice) with diethyl ether. Subsequent drying *in vacuo* ( $6 \times 10^{-3}$  mbar) resulted in a dark orange solid. Yield: 1.7 g (0.9 mmol) 69%; dark orange solid; (ratio of two isomers 1:0.7). M. p. 102 °C.  $^1\text{H}$  NMR (500.1 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 1.1–1.2 (t, 12H,  $^3J_{\text{H,H}} = 7.0$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 1.5–1.7 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 2.2–2.4 (m, 24H, cod), 2.5–2.6 (m, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 2.9, 3.0 (d, 3H,  $^3J_{\text{P,H}} = 12.4$  Hz,  $\text{O-CH}_3$ ), 3.9 (m, 6H, cod), 4.2 (br, 8H,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 5.2 (m, 6H, cod), 9.5 (brs,  $\text{C}^2\text{-H}$ ), 9.7 (brs,  $\text{C}^2\text{-H}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (125.75 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 13.5, 13.8 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ) two isomers, 19.9, 20.1 (s,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ) two isomers, 28.4 (s, cod), 31.3 (br,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 32.7 (s, cod), 49.9 (d,  $^3J_{\text{P,C}} = 6.3$  Hz,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 49.9 (br,  $\text{NCH}_2\text{CH}_2\text{CH}_2\text{Me}$ ), 72.0 (d,  $^1J_{\text{Rh,C}} = 13.0$  Hz, cod), 72.6 (d,  $^1J_{\text{Rh,C}} = 13.0$  Hz, cod), 73.21 (d,  $^1J_{\text{Rh,C}} = 13.1$  Hz, cod), 74.2 (d,  $^1J_{\text{Rh,C}} = 13.2$  Hz, cod), 74.6 (d,  $^1J_{\text{Rh,C}} = 13.0$  Hz, cod), 75.4 (d,  $^1J_{\text{Rh,C}} = 13.1$  Hz, cod), 113.4 (d,  $^2J_{\text{P,C}} = 9.5$  Hz,  $\text{O-CH}_3$ ), 121.1 (q,  $^1J_{\text{P,F}} = 332.7$  Hz,  $\text{CF}_3$ ), 131.9 (br,  $\text{P-C}$  of the middle ring), 133.3 (d,  $^{1/2}J_{\text{P,C}} = 45.2$  Hz,  $\text{P-C}$  of the middle ring), 141.9 (br,  $\text{H-C}^2$ ), 142.4 (br,  $\text{H-C}^2$ ) 2<sup>nd</sup> isomer.  $^{31}\text{P}$  NMR (500.0 MHz,  $\text{CD}_2\text{Cl}_2$ , 25 °C):  $\delta$  = 65.0 (d,  $^1J_{\text{Rh,P}} = 195.6$



Hz) (minor)/64.1 (dd,  $^1J_{\text{Rh,P}} = 200.1$  Hz,  $^3J_{\text{P,P}} = 5.3$  Hz) (major) and  $-120.3$  (t br,  $^1J_{\text{Rh,P}} = 126.9$  Hz) (major)/ $-123.4$  (t,  $^1J_{\text{Rh,P}} = 126.9$  Hz) (minor). IR [ $\text{cm}^{-1}$ ]:  $\tilde{\nu} = 2984.1$  (m), 2971.0 (m), 2861.8 (w), 1545.0 (s), 1491.7 (m), 1421.1 (m), 1327.8 (s), 1265.6 (m), 1129.8 (w), 1069.1 (s), 978.0 (s). Pos. ESI-MS:  $m/z$  (%) = 1153.205 (36)  $[\text{M} - \text{Cl} - 2\text{TfO}]^{+}$ . HRMS:  $[\text{C}_{47}\text{H}_{76}\text{Cl}_2\text{N}_4\text{OP}_2\text{Rh}_3]^{+}$  theor./exp. 1153.2031 (1153.2043). UV/vis ( $\text{CH}_2\text{Cl}_2$ ):  $\lambda$  [nm] (abs.): 386 (0.216). EA [%]: theor./exp. C 42.61/41.38, H 6.03/6.04, N 4.45/4.22, S 3.40/3.62.

## Conclusion

The first example of an anionic P-bridged tricyclic bis(NHC) was obtained *via* an unexpected reductive deselenization combined with a subsequent deprotonation of the bis(imidazolium) salt. Reaction of the latter with MeI as electrophile resulted finally in the first example of a neutral P-functional bis(NHCs) with a mixed substitution pattern. Initial studies of the coordination properties of the tricyclic bis(imidazolium) salt revealed a pre-ference of the neutral P-ligands site over the anionic phosphorus centre. Detailed electrochemical studies of the anionic bis(NHC) showed multiple, closely spaced oxidation processes owing, most probably, to the formation of a short lived P-centred radical that yields a new intermediate, a tetrakis(NHC), having a P–P bond linkage.

## Conflicts of interest

There are no conflicts of interest to declare.

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