

Cite this: *Dalton Trans.*, 2020, **49**,  
10574Received 1st July 2020,  
Accepted 13th July 2020

DOI: 10.1039/d0dt02331k

rsc.li/dalton

## The facile and additive-free synthesis of a cell-friendly iron(III)–glutathione complex†

Ziyu Gao,<sup>a,b</sup> Pablo Carames-Mendez,<sup>a</sup> Dong Xia,<sup>a</sup> Christopher M. Pask,<sup>a</sup> Patrick C. McGowan,<sup>a</sup> Paul A. Bingham,<sup>c</sup> Alex Scrimshire,<sup>c</sup> Giuseppe Tronci<sup>\*b,d</sup> and Paul D. Thornton<sup>id</sup> <sup>\*a</sup>

The straightforward creation of an unreported glutathione-stabilised iron(III) complex is disclosed. In contrast to previous reports, glutathione was shown to coordinate and stabilise iron directly under physiological conditions in the absence of additional sulfur containing molecules, such as sodium sulfide. The complex was extensively characterised; the molecular geometry was determined as two inequivalent octahedra, approximately 2/3 of which are slightly distorted towards more tetrahedral in character, with the remaining 1/3 more regularly octahedral. The dispersion of the iron(III)–glutathione complex in aqueous solution yielded particles of  $255 \pm 4$  nm in diameter that enhanced the growth and proliferation of L929 fibroblast cells over 7 days, and inhibited the activity of matrix metalloproteinase-13. Consequently, the unprecedented glutathione-stabilised iron(III) complex disclosed has potential use as a simple-to-prepare growth factor for inclusion within cell culture media, and is an excellent candidate as a therapeutic for the treatment of metalloproteinase-13-associated diseases.

### Introduction

Glutathione (GSH) is imperative for cellular defence against reactive oxygen species that cause apoptosis and tissue inflammation.<sup>1</sup> Free radicals and reactive oxygen species are deactivated by the sulfhydryl moiety that the tripeptide presents, undergoing conversion to stable and innocuous compounds. Numerous disease states are linked to GSH deficiency in cells, including chronic hypertension,<sup>2</sup> pulmonary fibrosis,<sup>3</sup> human immunodeficiency virus-related disease,<sup>4</sup> and respiratory distress syndrome.<sup>5</sup> GSH is administered as an oral supplement, however the half-life of GSH within the blood plasma ( $t = 1.6$  min) ensures that a significant proportion of the supplement is oxidised prior to reaching the gastrointestinal tract.<sup>6</sup> Consequently, alternative methods of GSH storage and

localised delivery, or reversible GSH modification, are required to elevate depleted tissue GSH levels.

Transition metal ions perform a number of cellular functions within eukaryotes, including photosynthesis and respiration.<sup>7</sup> Metalloproteins, which contain a metal ion cofactor, play a vital role in the storage and transport of enzymes and signal transduction proteins, and make-up a third of proteins within the human body.<sup>8</sup> Iron is the most abundant redox-active metal within the body owing to its presence within haemoglobin. The natural presence of haemoglobin *in vivo* provides great encouragement for the formation of iron-protein and iron-peptide complexes as materials that are suitable to be deployed within living organisms.

Complexes that contain redox-active metal centres may be exploited as therapeutics and diagnostic tools.<sup>9</sup> Notably, such organometallic materials have been proposed as contrast agents in which the variety of oxidation states that the redox-active metal can adopt is associated with distinct paramagnetic properties that enable their effective detection *in vivo*.<sup>10</sup> Iron has potential for use within a contrast agent composition due to its capability to form high-spin  $\text{Fe}^{3+}$  complexes. Wang *et al.* recently reported biochemically responsive  $\text{Fe}^{2+}$ -PyC3A complexes that, upon oxidation by reactive oxygen species, form  $\text{Fe}^{3+}$ -PyC3A that enable contrast enhancement within inflamed pancreatic tissue.<sup>11</sup> Within healthy tissue,  $\text{Fe}^{2+}$ -PyC3A is maintained and so signal enhancement does not occur.

GSH has been reported to form stable complexes with transition metals including mercury,<sup>12</sup> silver,<sup>13</sup> cadmium,<sup>14</sup> and

<sup>a</sup>School of Chemistry, University of Leeds, Leeds, UK.

E-mail: p.d.thornton@leeds.ac.uk

<sup>b</sup>Biomaterials and Tissue Engineering Research Group, School of Dentistry,

St. James's University Hospital, University of Leeds, UK. E-mail: g.tronci@leeds.ac.uk

<sup>c</sup>Materials and Engineering Research Institute, Sheffield Hallam University, City Campus, Howard Street, Sheffield, UK

<sup>d</sup>Clothworkers' Centre for Textile Materials Innovation for Healthcare, School of Design, University of Leeds, UK

† Electronic supplementary information (ESI) available: Spectra corresponding to UV-Vis absorption spectroscopy, FTIR, mass spectrometry, XPS, X-ray powder diffraction, UV-Vis fluorescence spectroscopy, and TGA. Vibrating-sample magnetometer data, cell culture images and statistical analysis of cell viability. See DOI: 10.1039/d0dt02331k



zinc.<sup>15</sup> Consequently, it has been cited as a candidate for metal speciation in aqueous environments for the detection and remedy of heavy metal pollution.<sup>16</sup> However, the formation of hydrolytically-stable iron–GSH complexes is more limited, especially in a cluster-free configuration. Qi *et al.* reported that glutathione-complexed clusters [2Fe–2S] are stabilised by glutathione aggregates that form intermolecular salt bridges and hydrogen bonds, providing a stable binding pocket for the [2Fe–2S] cluster core.<sup>17,18</sup> However, despite these advances, the binding of glutathione to ferrous iron in the absence of iron–sulfur cluster formation has not been reported due to the rapid reduction of Fe<sup>3+</sup> in the presence of GSH.<sup>19,20</sup> In this case, a complex that contains GSH ligands, but lacks [Fe<sub>2</sub>S<sub>2</sub>] bridged dimers, is anticipated to form.

We report the formation of a Fe<sup>3+</sup>–GSH complex by combining equimolar quantities of GSH and FeCl<sub>3</sub>. Additives that promote iron–sulfur cluster formation, for instance sodium sulfide, were not included in the synthesis. Analysis reveals that reduced GSH forms a complex with Fe<sup>3+</sup>, yielding a stable complex suitable for extensive characterisation. Notably, the Fe<sup>3+</sup>–GSH complex possessed excellent cell compatibility, promoting the growth and proliferation of L929 fibroblast cells over seven days *in vitro*. Consequently, it is highly promising as a drug-free approach to enhance fibroblast growth *in vitro* within cell culture medium at low concentrations (<0.5 mg mL<sup>-1</sup>), and as a constituent to restore homeostasis in chronic tissue states.

## Results and discussion

To confirm complex formation upon GSH addition to FeCl<sub>3</sub> solution, UV-Vis optical absorption spectroscopy was used to track metal–chelate coordination for various molar ratios after 30 minutes, when the solution colour remained consistent. The specific absorbance peak at 456 nm was ascribed to Fe–S formation. However, following Fe<sup>3+</sup>–GSH complex formation, two additional absorption peaks were detected at positions of 452 nm and 550 nm (Fig. S1a†) in our system that did not present in the spectrum of FeCl<sub>3</sub> solution (Fig. S2†). These peaks are attributed to Fe<sup>3+</sup>–GSH coordination.<sup>21,22</sup> A regular decrease in absorbance intensity was observed in solutions containing a decreasing molar ratio of FeCl<sub>3</sub>:GSH (6 → 1), until no peak could be detected in solutions containing a molar ratio lower than 1 (Fig. S1†). This observation can be rationalised by the fact that when solutions with a FeCl<sub>3</sub>:GSH molar ratio of 1:1.2 or lower were analysed, Fe<sup>2+</sup> was present exclusively and so no Fe<sup>3+</sup>-associated absorbance could be recorded, illustrating that binding between Fe<sup>3+</sup> and GSH occurred only at increased FeCl<sub>3</sub>:GSH molar ratios (*i.e.* >1). This observation also reveals that the thiol-induced reduction of Fe<sup>3+</sup> to Fe<sup>2+</sup> is expected to occur prior to Fe chelation in solutions with increased GSH content, and that the chelating interaction detected by UV-Vis spectroscopy was exclusively between GSH and Fe<sup>3+</sup>, rather than between GSH and Fe<sup>2+</sup>, in solutions with decreased GSH content.

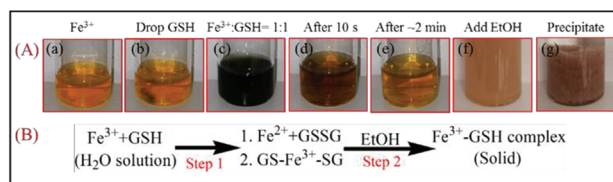


Fig. 1 Colour solution change that accompanies the transition of FeCl<sub>3</sub> solution to the precipitated Fe<sup>3+</sup>–GSH complex (A) and proposed mechanism of Fe<sup>3+</sup>–GSH complex formation (B).

Monitoring the colour of the Fe<sup>3+</sup>–GSH complex-yielding solution (FeCl<sub>3</sub>:GSH = 1:1) revealed that the starting yellow-like ferric chloride solution immediately shifted to a dark green colour for approximately three seconds upon addition of GSH (Fig. 1A). The solution then became lighter yellow, in comparison with the starting FeCl<sub>3</sub> solution, as shown in Fig. 1A. The quick formation of dark green solution is attributed to the reduction of Fe<sup>3+</sup> to Fe<sup>2+</sup>, and the oxidation of the GSH thiol, whereby either a disulfide bridge forms between two GSH molecules, a Fe–S bond forms, or a combination of both may form depending on the selected Fe<sup>3+</sup>:GSH molar ratio (Fig. 1B). As the pH of the solution containing both FeCl<sub>3</sub> and GSH is acidic (pH < 3), the disulfide bridge may be reduced and Fe<sup>2+</sup> undergo oxidation after approximately 3 seconds, when the light-yellow solution formed. Ethanol was then added to the yellow solution (Fig. 1A(f)), yielding a pale red precipitate (Fig. 1A(g)), which was dried at 37 °C and further characterised. The proposed mechanism for complex formation is provided in Fig. 1B. The capability of GSSG to induce iron oxidation, as shown in step 2, was confirmed by the GSSG-induced colour change of a FeCl<sub>2</sub> solution. The solution shifted to a pale red colour that is ascribed to the formation of a Fe<sup>3+</sup>–GSH complex (Fig. S3†).

The iron valence of the Fe<sup>3+</sup>–GSH complex was evaluated by <sup>57</sup>Fe Mössbauer spectroscopy at room temperature, 293 K (Fig. 2). All iron in the complex was confirmed to be Fe<sup>3+</sup>, with no evidence of any magnetically ordered phases.<sup>17,18,23</sup> Approximately 60% of the spectral area is consistent with Fe<sup>3+</sup> with an octahedral structure that is slightly distorted towards tetrahedral in character, as evidenced by the relatively low  $\Delta$  (doublet 1, centre shift ( $\delta$ ) = 0.417 mm s<sup>-1</sup>, quadrupole splitting ( $\Delta$ ) = 0.650 mm s<sup>-1</sup>),<sup>17,18,23</sup> and approximately 40% of the spectral area is consistent with a more regular octahedral Fe<sup>3+</sup> structure (doublet 2,  $\delta$  = 0.447 mm s<sup>-1</sup>,  $\Delta$  = 0.987 mm s<sup>-1</sup>).<sup>17,18,23</sup> The <sup>1</sup>H-NMR spectrum of the Fe<sup>3+</sup>–GSH complex is displayed in Fig. 3, whereby broader peaks can be observed compared to the spectrum of GSH. Peak broadening is ascribed to the paramagnetism of ferric ions (Fe<sup>3+</sup>), which affects the magnetic field of the instrument. An obvious shift of the CH<sub>2</sub>-S (position c) protons from 2.87 ppm to 2.89 ppm and 3.19 ppm is evident, and is attributed to Fe–S binding, as the strong metal–chelate coordination affects the neighbouring carbon environment. FTIR spectroscopy also revealed differences between the spectra of GSH and the Fe<sup>3+</sup>–GSH complex.



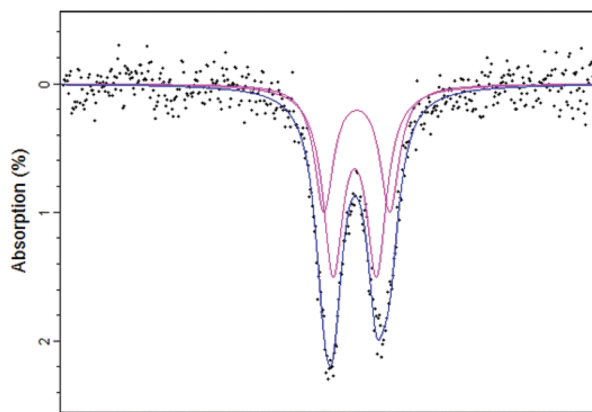


Fig. 2 Fitted Mössbauer spectrum for the  $\text{Fe}^{3+}$ -GSH complex.

A significant shift in the  $\text{C}=\text{O}$  peak position from  $1708\text{ cm}^{-1}$  (GSH) to  $1737\text{ cm}^{-1}$  ( $\text{Fe}^{3+}$ -GSH complex) suggests coordination between Fe and glutathione hydroxyl oxygen (Fig. S4†). The peak at  $2522\text{ cm}^{-1}$ , corresponding to a thiol vibration, disappeared from the spectrum of the complex, suggesting Fe-S binding. The peaks at  $3271\text{ cm}^{-1}$ ,  $1636\text{ cm}^{-1}$ , and  $1526\text{ cm}^{-1}$  correspond to the amine groups of GSH. Both  $^1\text{H}$  NMR and FTIR spectra revealed the formation of Fe-S and Fe-O bonds.

The  $\text{Fe}^{3+}$ -GSH complex was also found to display specific peaks in its CD spectrum (Fig. S5†). Compared with uncoordinated GSH in deionised water, the  $\text{Fe}^{3+}$ -GSH complex presented a broad positive peak from 750 to 460 nm and a negative peak from 460 to 302 nm, assigned to  $\text{Fe}^{3+}$  coordination.

In 60 mM HCl solution, no obvious structural change was observed for the GSH molecule, however, in the spectrum of the  $\text{Fe}^{3+}$ -GSH complex, an abrupt change was recorded, whereby the opposite trend was observed, at 750–422 nm and 422–360 nm, with respect to when the complex was incubated in  $\text{H}_2\text{O}$ . This change may be due to the presence of  $\text{H}^+$ ,<sup>24</sup> suggesting a reversible pH-induced peptide folding within the  $\text{Fe}^{3+}$ -GSH complex.

Surface element analysis confirmed the chemical composition and the coordinated binding energy of the  $\text{Fe}^{3+}$ -GSH powder. GSH powder was used as a control. The elements present in the  $\text{Fe}^{3+}$ -GSH complex are presented in Fig. S6,† which reveals that significant shifts were detected in  $\text{C}=\text{O}$ , C-O and S-related bonds (Fig. 4). No S-S bond was detected in this complex, which would be revealed by a peak at 166 eV, or greater.<sup>25</sup> The thiol group is protected from forming disulfide bridges with other GSH molecules by Fe chelation. In addition, no evidence for  $\text{Fe}^{2+}$  was found, which would be anticipated to appear at 708 eV,<sup>26</sup> rather than the pre-peak, which is normally observed in Fe related compounds.<sup>27</sup> This confirmed the exclusive existence of ferric iron, consistent with the results from Mössbauer spectroscopy.

X-ray powder diffraction analysis revealed that  $2\theta = 23.58^\circ$  with a full width at half maximum value of  $12.13^\circ$ , as shown in Fig. S7.† The  $d$  spacing of the complex was calculated to be 0.38 nm, and  $D_p = 0.70$  nm. The  $\text{Fe}^{3+}$ -GSH complex was confirmed to possess a 3-layer packing mode, which varied markedly to the crystal structure that was obtained for  $\text{FeCl}_3$  (Fig. S8†), confirming that a new stacking structure is formed for the  $\text{Fe}^{3+}$ -GSH complex.

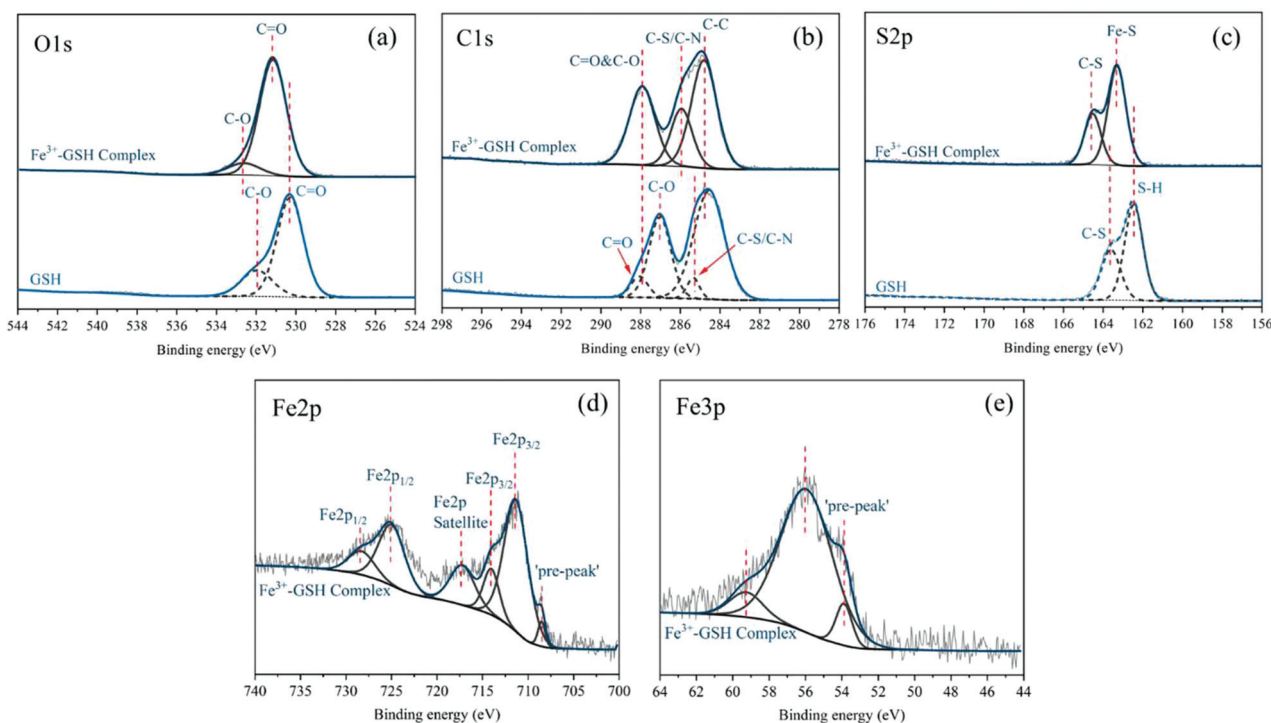


Fig. 3 X-ray photoelectron spectra of the  $\text{Fe}^{3+}$ -GSH complex and GSH, C 1s (a), S 2p (b), O 1s (c), Fe 2p (d) and Fe 3p (e) of  $\text{Fe}^{3+}$ -GSH.



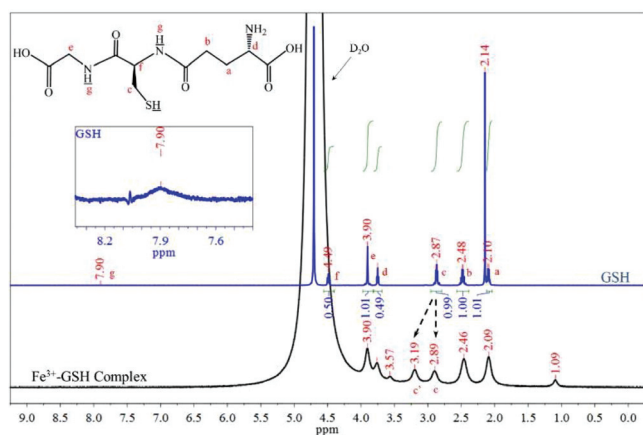


Fig. 4  $^1\text{H-NMR}$  spectra of  $\text{Fe}^{3+}$ -GSH complex (black) and GSH (blue) in  $\text{D}_2\text{O}$ .

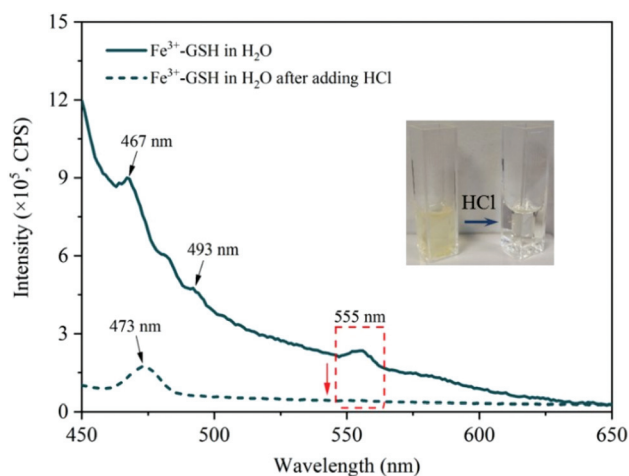


Fig. 5 UV-Vis fluorescence spectra of the  $\text{Fe}^{3+}$ -GSH complex in deionised water with a concentration of  $1.0 \text{ mg mL}^{-1}$ , initial status (solid) and after adding  $20 \text{ }\mu\text{L}$  HCl (dash).

UV-Vis fluorescence spectroscopy exposed an emission peak at  $555 \text{ nm}$  under excitation wavelength of  $408 \text{ nm}$  for the  $\text{Fe}^{3+}$ -GSH complex dispersed in deionised water (Fig. 5). This peak was also present in the spectrum corresponding to a mixture of  $\text{FeCl}_3$  and GSH (Fig. S9 $\dagger$ ). Upon the addition of HCl to the complex solution, the emission peak at  $555 \text{ nm}$  disappeared. A peak at  $473 \text{ nm}$  position formed, which lies between values of pure GSH ( $470 \text{ nm}$ ) and  $\text{FeCl}_3$  ( $493 \text{ nm}$ ) (Fig. S9 $\dagger$ ). The peak at  $555 \text{ nm}$  is considered to result from the strong binding between  $\text{Fe}^{3+}$  and GSH, yielding a rigid structure that provides a new emission peak. Its disappearance after adding HCl suggests that this structure is not stable in acidic solution, with the coordination between  $\text{Fe}^{3+}$  and GSH reduced, as supported by the accompanying solution colour change from a turbid pale yellow/creamy colour to a transparent colourless solution.

Thermogravimetric analysis (TGA) was carried out to further corroborate the chemical composition and composition of the complex. As observed in Fig. S10 $\dagger$ , the first mass loss of  $12.0\%$  was considered due to water evaporation; the next mass loss from  $88.0\%$  to  $11.5\%$ , which commenced at  $148.3 \text{ }^\circ\text{C}$ , is ascribed to the loss of GSH. After heating up to  $611.8 \text{ }^\circ\text{C}$ , the complex max remained constant at  $11.5\%$ , which was due to the presence of iron. Consequently, it was found that the complex contains  $87\%$  GSH and  $13\%$  Fe by mass. This is very close to the theoretical composition of  $84.6\%$  GSH and  $15.4\%$  Fe, based on a total complex mass of  $145.4 \text{ mg}$  that contains  $123 \text{ mg}$  GSH.

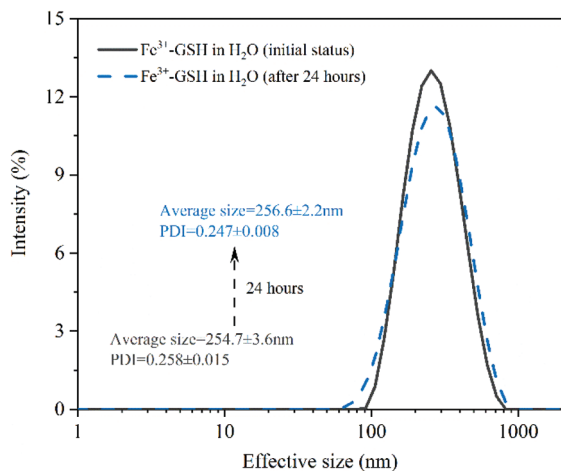
Although Fe was coordinated to water-soluble GSH, the solubility of the  $\text{Fe}^{3+}$ -GSH complex in water remained low due to the ionisation equilibrium of Fe itself. Iron ions tend to gain hydroxide ligands in neutral and basic water. Mass spectrometry analysis revealed peaks at  $615 \text{ m/z}$ , attributed to GSSG, and  $308 \text{ m/z}$ , attributed to GSH (Fig. S11 $\dagger$ ). In contrast to mass spectrometry, X-ray photoelectron spectroscopy (XPS) investigations revealed the absence of S-S bridges, suggesting that the formation of the oxidised product in the mass spectra is attributed to the electron source forming ionised species, which form GSSG in the instrument channel. Significant enhancement in peak intensity was observed after adding  $20 \text{ }\mu\text{L}$  HCl (final concentration of HCl was  $120 \text{ mM}$ ), from  $8.9 \times 10^5$  to  $13.2 \times 10^5$  peak intensity at  $615 \text{ m/z}$  and from  $0.8 \times 10^5$  to  $5.3 \times 10^5$  at  $308 \text{ m/z}$ . The peak at  $669 \text{ m/z}$  confirms the formation of the GS-Fe-SG complex, although its intensity surprisingly increased upon the addition of HCl. The enhancement of this peak following HCl addition, together with the change in solution colour from medium yellow to almost a colourless solution, suggested the ionisation of the  $\text{Fe}^{3+}$ -GSH complex had occurred, resulting in the oxidation of GSH to GSSG, and the partial formation of  $\text{Fe}^{2+}$  species.

DLS analysis revealed that the  $\text{Fe}^{3+}$ -GSH complex was stable in water, leading to nanoaggregates with an average effective size of  $255 \pm 4 \text{ nm}$  ( $\text{PDI} = 0.258 \pm 0.015$ ). Following 24-hour storage in water, the complex remained well dispersed with an average size of  $257 \pm 2 \text{ nm}$  ( $\text{PDI} = 0.247 \pm 0.008$ , Fig. 6). This characteristic may be exploited to generate bespoke  $\text{Fe}^{3+}$ -GSH carriers capable of the prolonged, or redox-mediated, release of GSH to, for example, chronic tissues that have downregulated levels of GSH.<sup>28</sup>

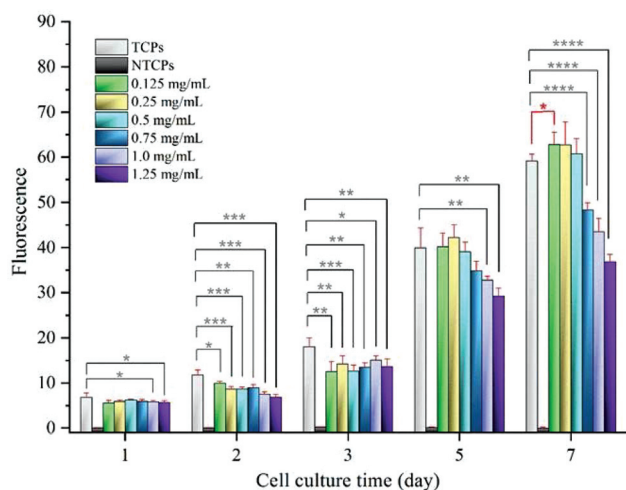
The paramagnetic property of  $\text{Fe}^{3+}$ -GSH was observed using a vibrating-sample magnetometer and increased in magnetic moment from  $-0.1474 \text{ emu g}^{-1}$  to  $0.1474 \text{ emu g}^{-1}$  under  $3 \text{ T}$  magnetic field, revealing the paramagnetism of the complex (Fig. S12 $\dagger$ ).

Following systematic characterisation of the complex, attention moved to its applicability within a biological environment. Cytotoxicity experiments were carried out with L929 mouse fibroblasts and confirmed high cellular tolerability of the  $\text{Fe}^{3+}$ -GSH complex. Growth and proliferation of L929 cells were measured during a 7-day culture (Fig. 7) with all the complex concentrations used in the cell culture medium, strongly supporting the excellent complex cytocompatibility (Table S1 $\dagger$ ).





**Fig. 6** The effective size of  $\text{Fe}^{3+}$ -GSH in  $\text{H}_2\text{O}$  at a concentration of  $0.5 \text{ mg mL}^{-1}$ , initial status (solid) and storage in water for 24 hours (dash).

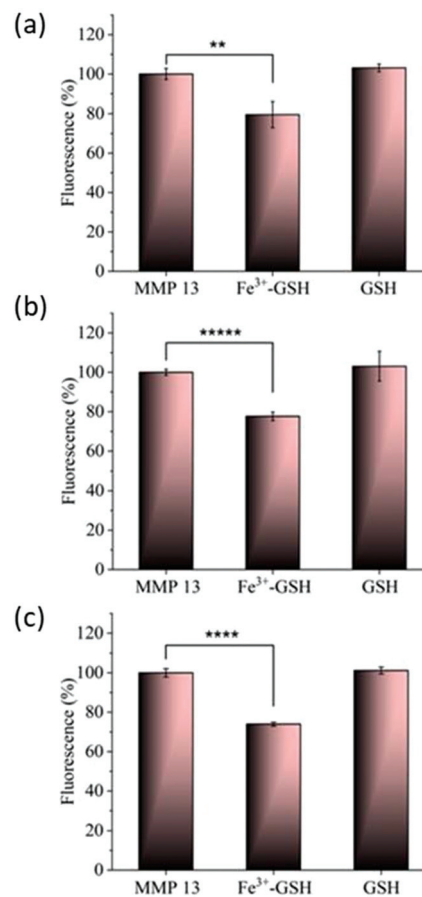


**Fig. 7** L929 cell viability without/with  $\text{Fe}^{3+}$ -GSH complex for 7 days culture.

Even compared with TCPs, only temporary inhibition was found in cell proliferation in the first 3 days. After 5-day culture, no significant difference could be observed in cell growth between cells cultured with either  $0.125 \text{ mg mL}^{-1}$ ,  $0.25 \text{ mg mL}^{-1}$ ,  $0.5 \text{ mg mL}^{-1}$  or  $0.75 \text{ mg mL}^{-1}$  of  $\text{Fe}^{3+}$ -GSH. After 7 days culture, L929 cells in  $0.125 \text{ mg mL}^{-1}$  presented significant higher viability and no significant difference was found at  $0.25 \text{ mg mL}^{-1}$  and  $0.5 \text{ mg mL}^{-1}$  conditions. The complex is described as being 'cell-friendly' due the positive effect that it has on cell growth when included in the assessed cell culture medium at  $0.125 \text{ mg mL}^{-1}$ , and due to it being non-cytotoxic when included in the assessed cell culture medium at concentrations  $\leq 0.5 \text{ mg mL}^{-1}$ . Conversely, when high  $\text{Fe}^{3+}$ -GSH content ( $>0.5 \text{ mg mL}^{-1}$ ) was supplemented to the cell culture medium, cells proved to grow slower with respect to cells cultured on tissue culture polystyrene (TCPs),

suggesting a significant GSH-induced impact on cellular activity at increased complex concentrations. In addition to possessing cellular tolerability, the  $\text{Fe}^{3+}$ -GSH complex proved also to be stable following 24-hour incubation in cell culture medium, suggesting long-lasting GSH storing capability (Fig. S13a-d<sup>†</sup>) and potential applicability as magnetic resonance contrast agent.

The effect of the complex on matrix metalloproteinase-13 (MMP-13) activity was then assessed to determine the effectiveness of the material as a viable therapeutic for ailments in which MMP-13 is overexpressed. Studies revealed that MMP-13 activity decreased to  $80 \pm 7\%$  after 12-hour incubation in the presence of the complex (Fig. 8). After 48 hours incubation, MMP-13 activity was downregulated to  $74 \pm 1\%$ . A significant difference was shown at all time points compared with the MMP-13 group, in which no complex was included. In addition, no significant difference was observed in the GSH group signifying the requirement of both organic and inorganic components for effective enzyme inhibition. This quick and lasting downregulation proved the ability of the  $\text{Fe}^{3+}$ -GSH complex to inhibit MMP-13 activity and suggested great potential for the complex to be applied to inflammation,



**Fig. 8** MMP-13 activity following incubation with  $\text{Fe}^{3+}$ -GSH or GSH for (a) 12 hours, (b) 24 hours, (c) 48 hours.



wound, and arthritis management by downregulating MMP activity.

## Conclusions

The formation of a  $\text{Fe}^{3+}$ -GSH complex without additional sulfur within its structure is reported and characterised for the first time. Aggregate formation of the complex in water yielded stable particulates with an initial average effective size of  $255 \pm 4$  nm. Such aggregates dispersed in aqueous solution and were found to be non-cytotoxic against L929 fibroblast cells. Fibroblast growth and proliferation was enhanced *versus* TCP after seven days of cell culture, suggesting that the material has great potential for use as a growth factor in cell culture medium. Additionally, the complex reduced the activity of MMP-13 over 48 hours, rendering it a highly promising candidate as a treatment for MMP-13-associated conditions.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The authors would like to thank Sarah Myers for support in the laboratory, and Dr Mark Howard for NMR training and guidance.

## Notes and references

- 1 A. T. Dharmaraja, *J. Med. Chem.*, 2017, **60**, 3221–3240.
- 2 S. Sagar, I. J. Kallo, N. Kaul, N. K. Ganguly and B. K. Sharma, *Mol. Cell. Biochem.*, 1992, **111**, 103–108.
- 3 K. M. Beeh, J. Beier, I. C. Haas, O. Kornmann, P. Micke and R. Buhl, *Eur. Respir. J.*, 2002, **19**, 1119–1123.
- 4 L. A. Herzenberg, S. C. De Rosa, J. G. Dubs, M. Roederer, M. T. Anderson, S. W. Ela, S. C. Deresinski and L. A. Herzenberg, *Proc. Natl. Acad. Sci. U. S. A.*, 1997, **94**, 1967–1972.
- 5 E. R. Pacht, A. P. Timerman, M. G. Lykens and J. Merola, *Chest*, 1991, **100**, 1397–1403.
- 6 A. Wendel and P. Cikryt, *FEBS Lett.*, 1980, **120**, 209–211.
- 7 I. Yruela, *Metallomics*, 2013, **5**, 1090–1109.
- 8 K. J. Waldron, J. C. Rutherford, D. Ford and N. J. Robinson, *Nature*, 2009, **460**, 823–830.
- 9 S. Betanzos-Lara, N. P. Chmel, M. T. Zimmerman, L. R. Barrón-Sosa, C. Garino, L. Salassa, A. Rodger, J. L. Brumaghim, I. Gracia-Moraa and N. Barba-Behrens, *Dalton Trans.*, 2015, **44**, 3673–3685.
- 10 Y. Miao, F. Xie, J. Cen, F. Zhou, X. Tao, J. Luo, G. Han, X. Kong, X. Yang, J. Sun and J. Ling, *ACS Macro Lett.*, 2018, **7**, 693–698.
- 11 H. Wang, V. C. Jordan, I. A. Ramsay, M. Sojoodi, B. C. Fuchs, K. K. Tanabe, P. Caravan and E. M. Gale, *J. Am. Chem. Soc.*, 2019, **141**, 5916–5925.
- 12 E. Bramanti, A. D'Ulivo, L. Lampugnani, G. Raspi and R. Zamboni, *J. Anal. At. Spectrom.*, 1999, **14**, 179–185.
- 13 Y. F. Ma, L. J. Wang, Y. L. Zhou and X. X. Zhang, *Nanoscale*, 2019, **11**, 5526–5534.
- 14 A. Fang, H. Chen, H. Li, M. Liu, Y. Zhang and S. Yao, *Biosens. Bioelectron.*, 2017, **87**, 545–551.
- 15 M. Gelinsky, R. Vogler and H. Vahrenkamp, *Inorg. Chim. Acta*, 2003, **344**, 230–238.
- 16 W. Chu, Y. Zhang, D. Li, C. J. Barrow, H. Wang and W. Yang, *Biosens. Bioelectron.*, 2015, **67**, 621–624.
- 17 W. Qi, J. Li, C. Y. Chain, G. A. Pasquevich, A. F. Pasquevich and J. A. Cowan, *J. Am. Chem. Soc.*, 2012, **134**, 10745–10748.
- 18 W. Qi, J. Li, C. Y. Chain, G. A. Pasquevich, A. F. Pasquevich and J. A. Cowan, *Chem. Commun.*, 2013, **49**, 6313–6315.
- 19 S. M. Hamed, J. Silver and M. Wilson, *Inorg. Chim. Acta*, 1983, **78**, 1–11.
- 20 R. Hider and X. Kong, *BioMetals*, 2011, **24**, 1179–1187.
- 21 S. Gervason, D. Larkem, A. B. Mansour, T. Botzanowski, C. S. Muller, L. Pecqueur, G. Le Pavec, A. Delaunay-Moisan, O. Brun, J. Agramunt, A. Grandas, M. Fontecave, V. Schunemann, S. Cianferani, C. Sizun, M. B. Toledano and B. D'Autreaux, *Nat. Commun.*, 2019, **10**, 3566.
- 22 C. Bonfio, L. Valer, S. Scintilla, S. Shah, D. J. Evans, L. Jin, J. W. Szostak, D. D. Sasselov, J. D. Sutherland and S. S. Mansy, *Nat. Chem.*, 2017, **9**, 1229–1234.
- 23 M. D. Dyar, D. G. Agresti, M. W. Schaefer, C. A. Grant and E. C. Sklute, *Annu. Rev. Earth Planet. Sci.*, 2006, **34**, 83–125.
- 24 L. Quintanar and L. Rivillas-Acevedo, *Methods Mol. Biol.*, 2013, **1008**, 267–297.
- 25 H. Mekaru, A. Yoshigoe, M. Nakamura, T. Doura and F. Tamanoi, *ACS Appl. Nano Mater.*, 2019, **2**, 479–488.
- 26 X. Wang, Q. Xiang, B. Liu, L. Wang, T. Luo, D. Chen and G. Shen, *Sci. Rep.*, 2013, **3**, 2007.
- 27 A. P. Grosvenor, B. A. Kobe, M. C. Biesinger and N. S. McIntyre, *Surf. Interface Anal.*, 2004, **36**, 1564–1574.
- 28 Z. Gao, B. Golland, G. Tronci and P. D. Thornton, *J. Mater. Chem. B*, 2019, **7**, 7494–7501.

