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10901**Dinuclear Tb and Dy complexes supported by hybrid Schiff-base/calixarene ligands: synthesis, structures and magnetic properties†**Peter Hahn,^a Steve Ullmann,^a Jennifer Klose,^a Yan Peng,^b Annie K. Powell ^b and
Berthold Kersting ^{*a}

The synthesis of the new lanthanide complexes [HNEt₃][Dy₂(HL¹)(L¹)] (**5**), and [Ln₂(L²)₂] (Ln = Tb^{III} (**7**), Dy^{III} (**8**)) supported by the hybrid Schiff-base/calix[4]arene ligands H₄L¹ (25-[2-((2-methylphenol)imino)ethoxy]-26,27,28-trihydroxy-calix[4]arene) and H₃L² (25-[2-((2-methylpyridine)imino)ethoxy]-26,27,28-trihydroxy-calix[4]arene) are reported. Spectroscopic data (for **5**) and X-ray crystallographic analysis (for **7**·4MeCN, **8**·4MeCN) reveal the presence of dimeric structures, featuring doubly-bridged NO₄Ln (μ-O)₂LnO₄N (**5**) or N₂O₃Ln(μ-O)₂LnO₃N₂ cores (**7**, **8**) with seven-coordinated Ln³⁺ ions. The magnetic properties of polycrystalline samples of **5**, **7** and **8** were studied by variable temperature dc and ac magnetic susceptibility measurements. The χ''(T) vs. T plots show no maxima in zero field, but the maxima can be detected under a 3 kOe dc field. The relaxation times τ obey the Arrhenius law above 5 K. Anisotropy barriers of ~18 cm⁻¹ (26 K) for **5** and ~23 cm⁻¹ (33 K) for **8** were determined.

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Introduction

The synthesis and investigation of magnetic properties of novel types of lanthanide complexes are very important for the development of molecular based magnetic materials,^{1–3} and many different types of lanthanide complexes with remarkable magnetic properties have been reported in the last several years.^{4–7} In 2003, slow relaxation of the magnetization was observed in the mononuclear lanthanide complexes [Pc₂Tb][–] and [Pc₂Dy][–], where Pc = phthalocyaninato, and this discovery spurred the search for other lanthanide-based single-molecule magnets (SMMs).⁸ Considerable efforts have been dedicated to the understanding of the mechanisms in mono- and polynuclear 4f SMMs in order to inform and direct ongoing research. It has been established that variation of the ligand field symmetry and coordination geometry, alteration of the charge and type of the donor atoms, and the type and magnitude of exchange interactions (in polynuclear systems) can have a great effect on the single ion anisotropy of the coordinated Ln³⁺ ions and thus the dynamics of the magnetiza-

tion.⁹ Polydentate ligands have turned out to be ideal candidates for systematic investigations because they can be rationally designed to control the coordination environment of the lanthanide ions.¹⁰ Thus, phthalocyaninato,¹¹ β-diketonato,¹² calix[4]arenes,¹³ and various Schiff base ligands^{14,15} or mixtures thereof¹⁶ have been used extensively in synthesizing functional coordination compounds of the 4f elements. In addition, these complexes have the added advantage that they can be developed for surface deposition, and in turn for device applications.^{17,18}

Several dinuclear terbium^{19–24} and dysprosium complexes with SMM behaviour have been reported.^{16,25–32} Most are mixed-ligand complexes, with the coordination spheres often being partly completed by solvate molecules or the anions. We previously reported a hybrid Schiff-base/calixarene ligand H₄L¹ (Fig. 1) that supports homoleptic lanthanide complexes [HNEt₃][Ln₂(HL¹)(L¹)] without additional coligands.³³ Hybrid ligands of this sort are known to complex first-row transition metals readily, but their lanthanide chemistry, particularly their magnetic properties,^{13,34} remains largely unexplored.^{35–42} In order to investigate the effect of donor atom variations on the structures and properties of such lanthanide calix[4]arene complexes, we have now focused on the coordination properties of the ligand H₃L², bearing a pyridyl-aldiminato podand arm in place of a salicylaldehyde unit.

Herein, we report the synthesis, structures and magnetic properties of the corresponding Tb^{III} and Dy^{III} complexes **7** and **8** (Fig. 1). Their properties are compared with those of

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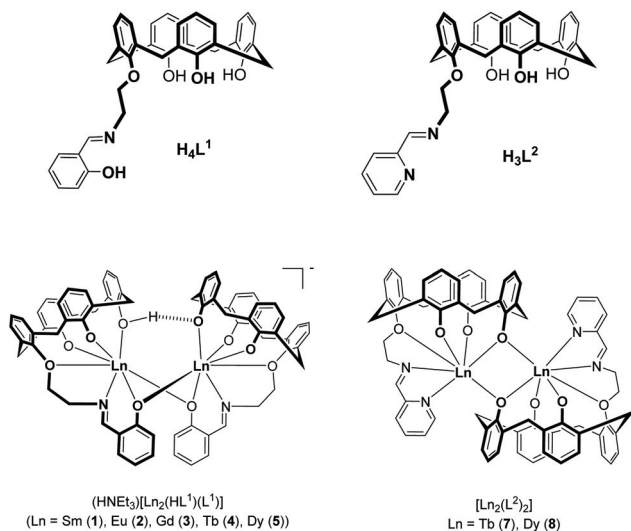


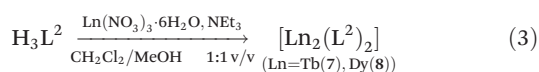
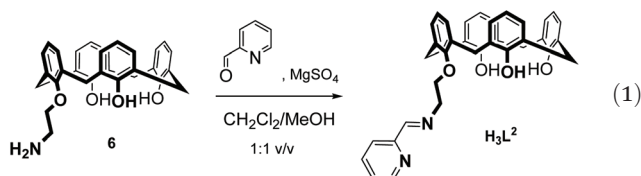
Fig. 1 Structures of hybrid Schiff-base/calix[4]arene ligands H_4L^1 and H_3L^2 and their lanthanide complexes **1–5**, **7**, and **8**. The complexes **1–4** have been reported previously.³³

complexes **4** and **5**. The contrasting responses to dynamic magnetic fields (in an applied DC field) are discussed in terms of the different coordination environments and geometries.

Results and discussion

Synthesis and spectroscopic characterization

The new ligand H_3L^2 was obtained in 74% yield by a Schiff-base condensation between 25-(aminoethoxy)-26,27,28-trihydroxy-calix[4]arene **6** and pyridine-2-carbaldehyde as illustrated in eqn (1).⁴³ H_3L^2 is soluble in aprotic solvents such as CH_2Cl_2 , $CHCl_3$, DMSO, MeCN and insoluble in alcohols. H_3L^2 was characterized by IR, 1H and ^{13}C NMR spectroscopy and electrospray ionization mass spectrometry (ESI-MS). The NMR spectra show one set of signals. 2D NMR experiments were performed to correctly assign the chemical shifts of hydrogen and carbon atoms (ESI†). The data are in accordance with the formulation of H_3L^2 and the “cone”-conformation of the calix[4]arene unit in H_3L^2 .



The syntheses of the Sm^{III} , Gd^{III} , Eu^{III} and Tb^{III} complexes **1–4** have been reported previously.³³ The corresponding

Dy^{III} complex **5** was synthesized in an analogous fashion (eqn (2)). Similarly, reaction of H_3L^2 with terbium(III) nitrate hexahydrate or dysprosium(III) nitrate hexahydrate and NEt_3 as a base in a mixed $CH_2Cl_2/MeOH$ solution at room temperature, followed by extraction from MeCN affords $[Ln_2(L^2)_2] \cdot 4MeCN$ complexes (**7**·4MeCN, **8**·4MeCN) as yellow crystals in 85–87% yield (eqn (3)). The compounds are air stable, have little solubility in common organic solvents, and lose solvate molecules upon standing in air. All new complexes were characterized by elemental analysis, mass spectrometry, and IR spectroscopy, and additionally by X-ray crystallography (for **7**·4MeCN, **8**·4MeCN). Table 1 lists selected analytical data.

A negative ESI-MS spectrum of a solution of **5** (10^{-3} M) in MeCN displays a molecular ion peak at $m/z = 1461.25$ with the correct isotopic peak pattern for a dimeric $[Dy_2(HL^1)(L^1)]^-$ anion (ESI). A signal at $m/z = 730.12$ for a monomeric $[Dy(L^1)]^-$ species is also observed. Similar peaks were observed for complexes **1–4**.³³ Complexes **7** and **8** were investigated by ESI MS in MeCN (concentration $\sim 10^{-3}$ M) using the positive mode. Under these conditions, molecular ion peaks with the correct isotopic peak pattern for dimeric $[Ln_2(L^2)_2 + H^+]^+$, $[Ln_2(L^2)_2 + Na^+]^+$ or $[Ln_2(L^2)_2 + K^+]^+$ cations are detected (Fig. S7–S10†). Signals for monomeric $[Ln(L^2) + H^+]^+$ species are also observed, but all attempts to isolate monomeric complexes failed so far. The dimerisation is believed to be driven by the coordination requirements of the lanthanide ions. As in H_3L^1 , the hexadentate N_2O_4 ligand H_3L^2 cannot saturate the coordination sphere of the lanthanide ions alone and so dimerization *via* bridging phenolate atoms takes place to increase the coordination number. Peaks characteristic for host guest complexes with intracavity MeCN molecules could not be detected.

The FTIR spectrum of **5** is similar to those of **1–4**, with a strong band at 1635 cm^{-1} for the $\nu(C=N)$ stretching frequency. A broad band at 3439 cm^{-1} is attributed to the O–H stretching vibration. The $\nu(C-O)$ stretching frequency observed for H_4L^1 at 1338 cm^{-1} is shifted to 1327 cm^{-1} , indicative of the coordination of the pendant phenol ether moiety.⁴⁴

The most characteristic bands in the FTIR-spectrum of H_3L^2 are the bands at 1651 cm^{-1} for the imine group ($\nu(C=N)$) and a broad feature at 3250 cm^{-1} for the O–H stretches of the calix[4]arene moiety. The latter band is not seen in the IR spectra of the $[Ln_2(L^2)_2]$ complexes in agreement with the triply deprotonated form of H_3L^2 . The stretching vibration for the imine group is observed at 1651 cm^{-1} for H_3L^2 , 1652 cm^{-1} for **7** and 1655 cm^{-1} for **8**, respectively. The intensity of the $\nu(C=N)$ band decreases significantly upon coordination of the lanthanide ions.

X-ray crystallography

We have not been able to grow single-crystals of the Dy^{III}_2 complex **5**. However, on the basis of the spectroscopic data, this complex is believed to be isostructural with the Sm^{III}_2 (**1**) and Gd^{III}_2 complexes (**3**), both of which have been crystallographically characterized.³³

Single crystals of $[Tb_2(L^2)_2] \cdot 4MeCN$ (**7**·4MeCN) obtained by extraction from MeCN were found to be suitable for X-ray crys-



Table 1 Selected analytical data for the supporting ligands H_4L^1 , H_3L^2 and their lanthanide complexes **5**, **7** and **8**

Compound	ESI-MS ^a		IR ^b /cm ⁻¹			Ref.
			$\nu(\text{OH})$	$\nu(\text{C}\equiv\text{N})$	$\nu(\text{C}-\text{O})$	
H_4L^1	570.3 ^c	$[\text{H}_3\text{L}^1]^-$	3635 3500 3320	1635	1338	33
5 (Dy)	1461.25 ^c 730.12 ^c	$[\text{Dy}_2(\text{HL}^1)(\text{L}^1)]^-$ $[\text{Dy}(\text{L}^1)]^-$	3439	1635	1327	This work
H_3L^2	557.2 ^d	$[\text{H}_3\text{L}^2]^+$	3250	1651	1377	This work
7 (Tb)	1447.29 ^d 1463.24 ^d 713.16 ^d	$[\text{Tb}_2(\text{L}^2)_2 + \text{Na}^+]^+$ $[\text{Tb}_2(\text{L}^2)_2 + \text{K}^+]^+$ $[\text{Tb}(\text{L}^2) + \text{H}^+]^+$	—	1652	1331	This work
8 (Dy)	1433.22 ^d 717.1 ^d	$[\text{Dy}_2(\text{L}^2)_2 + \text{Na}^+]^+$ $[\text{Dy}(\text{L}^2) + \text{H}^+]^+$	—	1655	1333	This work

^a Concentration of solutions were $\sim 1.0 \times 10^{-3}$ M. ^b MeCN solution. ^c ESI(−) mode. ^d ESI(+) mode.

tallographic analyses. The crystal structure determination revealed the presence of dimeric $[\text{Tb}_2(\text{L}^2)_2]\cdot 4\text{MeCN}$ complexes (Fig. 2). There are two heavily disordered MeCN solvent molecules, which were removed from the structure using the Squeeze routine implemented in Platon.⁴⁵ Two well-defined MeCN molecules are situated in the calix[4]arene cavities as in other calix[4]arene structures.⁴⁶

Complex **7** exhibits crystallographically imposed C_2 symmetry, with the C_2 axis passing through the centre of a central four-membered Tb_2O_2 ring. Two $[\text{Tb}^{\text{III}}(\text{L}^2)]$ units are joined by two phenolato bridges at a $\text{Tb}\cdots\text{Tb}$ distance of 3.8011(8) Å. This structure is strikingly different from that seen in **1–5**. In the latter, dimerization occurs *via* phenolato groups from the pendant salicylaldehyde units, while $[\text{Tb}^{\text{III}}(\text{L}^2)]$ monomers are

linked *via* phenolato groups from the capping calix[4]arene entities. In **1–5**, one Ln^{3+} ion adopts a capped octahedral geometry and the other is a capped trigonal prism as indicated by the SHAPE algorithm (ESI).⁴⁷ Conversely, in **7**, both Ln^{3+} ions are surrounded in a distorted capped trigonal prismatic fashion by five calix[4]arene O, one imine N and one pyridyl N atom. Also the Ln_2O_2 cores in the C_i symmetric complexes **1–5** are bent, while the Tb_2O_2 core in **7** is almost flat, with the TbO_2 planes intersecting at 12.5°. It should be noted that the dimerization mode of the calix[4]arene units in **7** is reminiscent of that seen in lanthanide complexes supported by unmodified calixarene ligands. The complexes $[\{\text{Eu}(4\text{-tert-butyl-calix}(4)\text{arene})(\text{DMF})_2\}]^{48}$ and $[\{\text{Nd}(\text{calix}[4]\text{H})(\text{thf})\}_2]^{49}$ represent two typical examples. There are other dimeric Ln_2 complexes containing phenolato bridges, but the coordination numbers tend to be higher than in the present case.^{50–52} Significant interactions between the complexes are not observed due to the steric bulk of the supporting ligand. Thus, the closest intermolecular $\text{Tb}\cdots\text{Tb}$ distance is at 9.145 Å.

Table 2 lists selected bond lengths and angles. The $\text{Tb}-\text{O}$ bond lengths vary considerably from 2.136(5) Å to 2.605(5) Å. The $\text{Tb}-\text{O}$ bonds involving the non-bridging phenolato groups (O2, O4) are significantly shorter than those involving the bridging phenolato groups (O3, O3') (mean value 2.33 Å), as in other seven-coordinate Tb^{III} calix[4]arene complexes.^{13,53} The

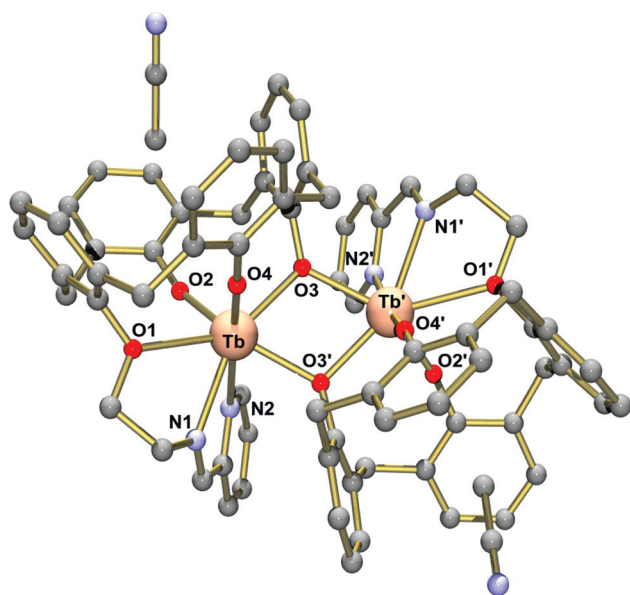


Fig. 2 Molecular structure of the neutral $[\text{Tb}_2(\text{L}^2)_2]$ complex in crystals of **7**·4MeCN (Ball and stick representation). The disordered MeCN solvate molecules are omitted for clarity. Symmetry code used to generate equivalent atoms $0.5 - x, -0.5 + y, -z$.

Table 2 Selected bond lengths/Å and angles/° for **7**·4MeCN and **8**·4MeCN

	$[\text{Tb}_2(\text{L}^2)_2]\cdot 4\text{MeCN}$	$[\text{Dy}_2(\text{L}^2)_2]\cdot 4\text{MeCN}$
M–O1	2.605(5)	2.608(6)
M–O2	2.136(5)	2.134(5)
M–O3	2.323(5)	2.311(5)
M–O3'	2.345(4)	2.343(5)
M–O4	2.109(5)	2.106(6)
M–N1	2.536(7)	2.509(7)
M–N2	2.691(7)	2.675(8)
M...M'	3.8011(8)	3.7819(9)
M–O3–M	109.03(18)	108.7(2)



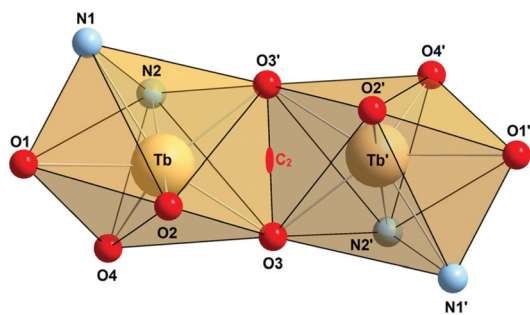


Fig. 3 Polyhedral representation of the dinuclear $N_2O_3Tb(\mu-O)_2TbO_3N_2$ core in $7 \cdot 4MeCN$ highlighting the mono-capped trigonal prismatic coordination geometry. The C_2 axis is indicated. Symmetry code used to generate equivalent atoms $0.5 - x, -0.5 + y, -z$.

phenol ether O atom (O1) forms the longest Tb–O bond at 2.61 Å. The Tb–N bonds are also quite long, with the Tb–N^{imine} bond length at 2.536(7) Å being somewhat shorter than the Tb–N^{pyridine} bond 2.691(7) Å. The Tb–O–Tb angles of 109.0° are normal for phenolato-bridged Ln_2 complexes (Fig. 3).¹⁰

The structure of the dysprosium complex $[Dy_2(L^2)_2] \cdot 4MeCN$ ($8 \cdot 4MeCN$) is isomorphous with $7 \cdot 4MeCN$ (Fig. S13†), having slightly shorter Dy–O and Dy–N distances (Table 2), in agreement with the smaller ion radius of Dy^{3+} .⁵⁴ The intramolecular Dy...Dy distance is at 3.7819(9) Å. Taken together, the two ligands behave similarly in the sense that they both support phenolato-bridged Ln_2 complexes, but the fashion in which this is individually achieved is strikingly different for the two cases.

Magnetic properties

The lanthanide complexes **5**, **7**, and **8** were further studied by variable temperature direct and alternating current (dc and ac) magnetic susceptibility measurements using a SQUID-Magnetometer (MPMS Quantum Design) in order to probe their magnetic properties. All samples were desolvated at 80 °C in vacuum to remove the enclathrated solvate and intracavity MeCN molecules. The magnetic properties of compounds **1–4** have been reported previously.³³

Plots of $\chi_M T$ versus T for **5**, **7**, and **8** are displayed in Fig. 4. At room temperature, the observed $\chi_M T$ values are 30.79 cm³ K mol^{−1}, 25.12 cm³ K mol^{−1} and 30.49 cm³ K mol^{−1} for complexes **5**, **7**, and **8**. The values for the Dy complexes **5** and **8** are slightly larger than the theoretical value of 28.34 cm³ K mol^{−1} for two non-interacting Dy^{3+} ions ($^6H_{15/2}$, $S = 5/2$, $L = 5$, $g = 4/3$, $\chi T = 14.17$ cm³ K mol^{−1}). In both cases, the $\chi_M T$ values decrease slightly upon decreasing the temperature reaching 21.54 cm³ K mol^{−1} (**5**) and 21.17 cm³ K mol^{−1} (**8**) at 4 K, respectively. This decrease is attributed to depopulation of the Stark sublevels and/or the magnetic anisotropy of the Dy^{3+} ions, although contributions from antiferromagnetic exchange interactions cannot be ruled out. In the Gd^{III} complex **3**, for example, a weak antiferromagnetic exchange interaction is present, $J = -0.065$ cm^{−1}.³³

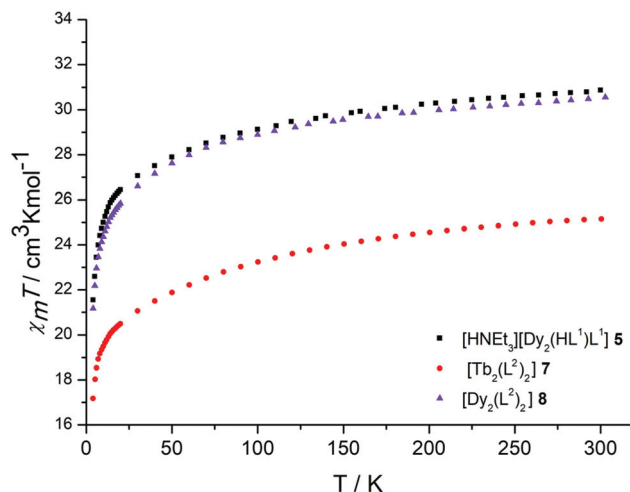


Fig. 4 Temperature dependence of the $\chi_M T$ product (at 500 Oe) for the dinuclear complexes **5**, **7** and **8**, χ_M being the molar susceptibility per dinuclear complex defined as M/H .

The observed $\chi_M T$ value of the Tb^{III} complex **7** is also slightly higher than the expected value of 23.60 cm³ K mol^{−1} for two non-interacting Tb^{3+} ions (7F_6 , $S = 3$, $L = 3$, $g = 3/2$, $\chi T = 11.82$ cm³ K mol^{−1}). As observed for **5** and **8**, the $\chi_M T$ values decrease with decreasing temperature reaching 17.17 cm³ K mol^{−1} at 4 K. A similar behaviour was observed for the Tb^{III} complex **4**. Tb^{3+} is known to exhibit significant magnetic anisotropy, and fitting of susceptibility data is therefore difficult.⁵⁵

The field dependence of the magnetization for complexes **5**, **7** and **8** was studied below 8 K in order to see whether magnetic anisotropy is present. Indeed, in all cases the magnetization values increase relatively rapidly at low fields and then linearly at higher field but without a clear saturation (Fig. S14–S16†). Thus, the magnetization values of 11.81 μ_B (**5**), 9.93 μ_B (**7**) or 11.92 μ_B for **8** (at 40 kOe and 2 K) are significantly lower than the theoretical value of 18 μ_B or 20 μ_B for dinuclear Tb^{III} and Dy^{III} complexes. Moreover, the iso-field lines are not superposed on each other as would be expected for an isotropic system with a well-defined ground state. These observations clearly reflect the presence of significant magnetic anisotropy and/or low lying excited states.^{56,57}

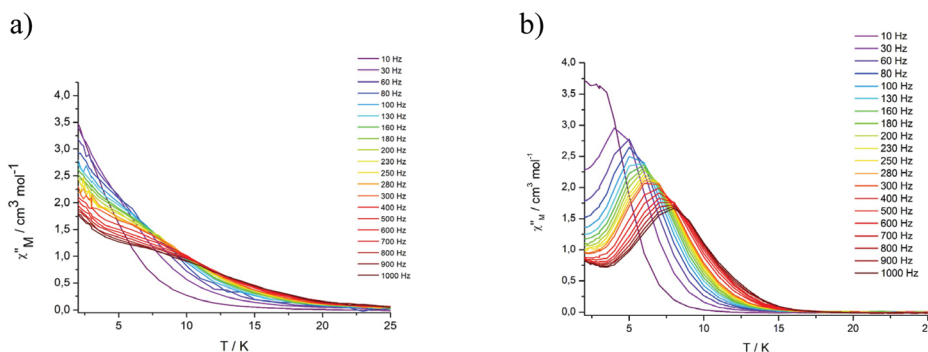
Variable temperature ac measurements were performed for **5**, **7**, and **8** in order to study their response to dynamic magnetic fields. Under a zero dc field, however, none of the complexes revealed frequency dependent in-phase (χ') or out-of-phase (χ'') signals, ruling out the presence of true SMM behaviour.⁴ This behaviour was also observed for the Tb^{III} complex **4** and is in striking contrast to many other phenolato-bridged structures, which are true SMMs (Table 3). Inspection of the data in Table 3 shows that the known phenolato-bridged Dy_2 and Tb_2 SMM systems exhibit coordination numbers 8 and 9 with a more pronounced axial symmetry, a factor which is known to enhance the anisotropy.⁵ Our complexes (*i.e.* the coordination polyhedra) lack axial symmetry and magnetic an-



Table 3 Structures and magnetic properties of selected Tb^{III} and Dy^{III} complexes exhibiting a phenolato-bridged Ln₂O₂ core structure

Complex	CN	Donor set	Coordination geometry	Magnetic parameters ^a U_{eff}/K (t_0/s)	Ref.
[Tb ₂ (HL ^a) ₄ (NO ₃) ₆]	9	O ₉	Distorted spherical capped square antiprism	Not determined	20
[NHET ₃] ₂ [Tb ₂ (μ-NO ₃) ₂ (NO ₃) ₂ (HL) ₂] ^b	9	O ₇ N ₂	Distorted muffin like	34 (1.1×10^{-8})	22
[NHET ₃] ₂ [Dy ₂ (μ-NO ₃) ₂ (NO ₃) ₂ (HL) ₂] ^b	9	O ₇ N ₂	Distorted muffin-like	80 (2.2×10^{-6})	22
[Dy(L ^c)(HL ^c)(phen)]	8	O ₆ N ₂	Triangular dodecahedron	160 (9×10^{-8})	25
[Dy ₂ (bfbpen) ₂ (H ₂ O) ₂](I) ₂	8	O ₄ N ₄	Square antiprism	16.9 (3.04×10^{-6})	26
[Dy ^{III}] ₂ (bcbpen) ₂ (H ₂ O) ₂ [(I) ₂ ·0.5H ₂ O]	8	O ₄ N ₄	Square antiprism	49.2 (6.79×10^{-7})	26
[Dy ₂ L ₂ (O ₂ CPh) ₂ ·2MeOH] ^d	8	O ₅ N ₃	Distorted triangular dodecahedron	47.5 (6.42×10^{-6})	31
[Dy ₂ L ₂ {(2-NO ₂)O ₂ CPh} ₂] ^d	8	O ₅ N ₃	Distorted triangular dodecahedron	67.6 (2.52×10^{-7})	31
[Dy ₂ L ₂ {(2-OH)O ₂ CPh} ₂] ^d	8	O ₅ N ₃	Distorted triangular dodecahedron	51.6 (7.50×10^{-6})	31
MeOH·MeCN ^d					
[Dy ₂ (hmi) ₂ (NO ₃) ₂ (MeOH) ₂]	8	O ₇ N ₁	Not determined	56 (3×10^{-7})	58
[Dy ₂ (api) ₂]	8	O ₄ N ₄	Square antiprism	25.8 (6.79×10^{-6})	59
[Dy(dbm) ₂ (L) ₂] ^e	8	O ₆ N ₂	Distorted dodecahedron	34.5 (1.54×10^{-6})	60
[Dy(dbm) ₂ (L) ₂] ^f	8	O ₆ N ₂	Distorted dodecahedron	67.6 (6.1×10^{-7})	9
[Dy ₂ (valdien) ₂ (NO ₃) ₂]	8	O ₅ N ₃	Square antiprism; dodecahedral	76 (6.04×10^{-7})	9
[HNET ₃] ₂ [Dy ₂ (HL ¹)(L ¹)] ⁵	7	O ₆ N ₁	Capped trigonal prism, capped octahedron	21.9 (4.73×10^{-5})	This work
[Dy ₂ (L ²) ₂] ⁸	7	O ₅ N ₂	Capped trigonal prism	32.9 (9.01×10^{-6})	This work

^a HL = 8-hydroxyquinoline. ^b H₃L = *N'*-(2-hydroxy-3-methoxy-5-nitrobenzylidene)-2-(hydroxyamino)propanehydrazide. ^c H₂L = 3,5-dichloro-4-hydroxy benzoic acid; phen: 1,10-phenanthroline; H₂bfbpen: *N,N'*-bis-(2-hydroxy-5-fluoro-benzyl)-*N,N'*-bis-(pyridin-2-ylmethyl)ethylenediamine; H₂bcbpen: *N,N'*-bis-(2-hydroxy-5-chloro-benzyl)-*N,N'*-bis-(pyridin-2-ylmethyl)ethylenediamine. ^d H₂L = *N*,*N*₃-bis(4-chlorosalicylaldehyde) diethylenetriamine; H₂hmi: (2-hydroxy-3-methoxyphenyl)methylene (isonicotino)hydrazine; H₃api: 2-(2-hydroxyphenyl)-1,3-bis[4-(2-hydroxyphenyl)-3-azabut-3-enyl]-1,3-imidazoline; dbm: 1,3-diphenyl-1,3-propanedione. ^e HL = 2-[[[(4-methoxy-phenyl)imino]methyl]-8-hydroxy-quinoline. ^f HL = 2-[[[(4-ethoxyphenyl)imino]methyl]-8-hydroxyquinoline; H₂valdien: *N*,*N*₃-bis(3-methoxysalicylidene) diethylenetriamine.

**Fig. 5** Frequency dependence of the out-of-phase (χ'') ac susceptibility from 2 to 25 K under an applied static field (3000 Oe) at indicated frequencies for Dy^{III}₂ complexes **5** (a) and **8** (b).

isotropy, and this observation is in good agreement with the reported trend.¹⁵

Variable temperature ac measurements were also performed in an applied DC field of 3 kOe to suppress quantum tunneling of the magnetization. Under these conditions, frequency dependent in-phase and out-of-phase signals could be observed, but only for the Dy complexes (Fig. 5, S17 and S18†).

Complex **5** displays only a broad shoulder between 6 and 10 K at 1000 Hz. For complex **8**, a well developed peak is clearly discernible with the maximum located at 8 K at 1000 Hz. In both cases, an enhancement of the peak intensity is indicated as the quantum tunneling is reduced with decreasing temperature. The relaxation time τ of complex **8** derived from the χ'' peaks follow the Arrhenius law (Fig. S22†) with $\tau_0 =$

9.01×10^{-6} s and the effective anisotropy barrier is determined to be $\Delta E/K_B = 32.9$ K. The corresponding values for the Dy₂ complex **5** are significantly smaller $\Delta E/K_B = 26.2$ K and $\tau_0 = 3.11 \times 10^{-5}$ s, attributable to the changes in the coordination environments, as detailed above. The values for **5** and **8** are otherwise in a typical range seen for other Dy₂ complexes displaying SMM behaviour in a dc field.^{21,26,61}

Conclusion

In summary, a new monofunctionalized hybrid Schiff-base calix[4]arene ligand H₃L², bearing a pyridyl-aldiminato podand arm has been synthesized and its ability to form dimeric com-



plexes $[\text{Ln}_2(\text{L}^2)_2]$ with $\text{Ln} = \text{Tb}^{3+}$ (7) and Dy^{3+} (8) has been demonstrated. The dinuclear complex $[\text{HNEt}_3][\text{Dy}_2(\text{HL}^1)(\text{L}^1)]$ (5), where a salicylaldimine podand arm is installed at the calix[4]arene moiety, has also been successfully isolated and characterized. Crystallographic studies show that the structures supported by H_4L^1 and H_3L^2 differ greatly. In 7 and 8, $[\text{Ln}(\text{L}^2)]$ monomers are linked *via* phenolato groups from the capping calix[4]arene, while in 5 dimerization occurs *via* phenolato groups from the pendant salicylaldimine units. Although both Dy compounds require applied static fields to allow for observation of their SMM, the impact of the change in the coordination environments on the dynamic magnetism is significant.

Experimental section

Materials and methods

Precursor compound 6 and the ligand H_4L^1 were prepared as previously described.³³ All other chemicals and solvents were obtained from commercial suppliers in reagent grade and were used without further purification. Melting points were determined in open-glass capillaries and are uncorrected. Mass spectra were obtained using the positive or negative electrospray ionization modus (ESI) on a Bruker Daltonics ESQUIRE 3000 Plus ITMS or Impact II UHR Qq-TOF instrument. ^1H and ^{13}C NMR spectra were recorded on a Bruker FOURIER 300 or a Bruker AVANCE DRX 400 spectrometer at 298 K. Chemical shifts refer to solvent signals. Mid (4000–400 cm^{-1}) infrared spectra at 2 cm^{-1} resolution were recorded on a Bruker TENSOR 27 (equipped with a MIRacle ZnSe ATR accessory from PIKE Technologies) FT-IR spectrometer. Elemental analyses were performed on a vario EL elemental analyzer (Elementar Analysensysteme GmbH, Hanau).

Synthesis of compounds

25-[2-((2-Methylpyridine)imino)ethoxy]-26,27,28-trihydroxy-calix[4]arene H_3L^2 . To a solution of calix[4]arene 6 (500 mg, 1.07 mmol, 1.00 eq.) in a mixture of CH_2Cl_2 (75 mL) and MeOH (75 mL) was added pyridine-2-carbaldehyde (111.90 μL , 1.18 mmol, 1.10 eq.) and an excess of MgSO_4 . The suspension was refluxed for 5 h, filtered and evaporated to dryness. The remaining solid was triturated with MeOH, filtered and washed with cold MeOH to give 0.44 g (74%) of pure H_3L^2 as a white solid. Yield: 74%. m.p. 209 °C. ^1H NMR (300 MHz, CD_2Cl_2 , see Fig. S1† for labeling scheme): δ 3.36–3.39 (d, 2 H, $^2J = 13.7$ Hz, $\text{Ar}-\text{CH}_{\text{eq}}\text{H}-\text{Ar}$, $\text{C}^{8/14}$); 3.45–3.48 (d, 2 H, $^2J = 13.0$ Hz, $\text{Ar}-\text{CH}_{\text{eq}}\text{H}-\text{Ar}$, $\text{C}^{2/20}$); 4.06–4.10 (d, 2 H, $^2J = 13.7$ Hz, $\text{Ar}-\text{CH}_{\text{ax}}\text{H}-\text{Ar}$, $\text{C}^{8/14}$); 4.36–4.49 (m, 6 H, $\text{O}-\text{CH}_2$, CH_2-N , $\text{Ar}-\text{CH}_{\text{ax}}\text{H}-\text{Ar}$, C^{29} , C^{30} , $\text{C}^{2/20}$); 6.61–6.66 (m, 3 H, *para* ArH, C^{11} , $\text{C}^{5/17}$); 6.88–6.90 (t, 1 H, $^3J = 7.6$ Hz, *para* ArH, C^{23}); 6.95–6.99 (m, 4 H, *meta* ArH, $\text{C}^{10/12}$, $\text{C}^{6/16}$); 7.04–7.10 (m, 4 H, *meta* ArH, $\text{C}^{4/18}$, $\text{C}^{22/24}$); 7.31–7.34 (m, 1 H, H_{py} , C^{35}); 7.69–7.73 (m, 1 H, H_{py} , C^{34}); 8.18–8.20 (d, 1 H, $^2J = 7.9$ Hz, H_{py} , C^{33}); 8.68–8.69 (d, 1 H, $^2J = 4.5$ Hz, H_{py} , C^{36}); 8.77 (s, 1H, $\text{N}=\text{CH}$, C^{31}). $^{13}\text{C}\{^1\text{H}\}$

NMR (100 MHz, CD_2Cl_2 , see Fig. S1† for labeling scheme): δ 31.85 ($\text{Ar}-\text{CH}_2-\text{Ar}$, $\text{C}^{2/20}$); 32.11 ($\text{Ar}-\text{CH}_2-\text{Ar}$, $\text{C}^{8/14}$); 61.13 (CH_2-N , C^{30}); 76.54 ($\text{O}-\text{CH}_2$, C^{29}); 121.06 (*para* C^{Ar} , $\text{C}^{5/17}$); 121.76 (C^{py} , C^{33}); 122.08 (*para* C^{Ar} , C^{11}); 125.44 (C^{py} , C^{35}); 126.33 (*para* C^{Ar} , C^{23}); 128.74 (*ortho* C^{Ar} , $\text{C}^{7/15}$); 128.83 (*ortho* C^{Ar} , $\text{C}^{3/19}$); 128.99 (*meta* C^{Ar} , $\text{C}^{4/18}$); 129.12 (*meta* C^{Ar} , $\text{C}^{6/16}$); 129.20 (*ortho* C^{Ar} , $\text{C}^{9/13}$); 129.25 (*meta* C^{Ar} , $\text{C}^{10/12}$); 129.85 (*meta* C^{Ar} , $\text{C}^{22/24}$); 134.87 (*ortho* C^{Ar} , $\text{C}^{1/21}$); 137.00 (C^{py} , C^{34}); 149.87 ($\text{C}^{\text{Ar}}-\text{OH}$, C^{27}); 149.91 (C^{py} , C^{36}); 151.66 ($\text{C}^{\text{Ar}}-\text{OH}$, $\text{C}^{26/28}$); 152.23 ($\text{C}^{\text{Ar}}-\text{OH}$, C^{25}); 155.28 (C^{py} , C^{32}); 165.86 ($\text{C}=\text{N}$, C^{31}). ATR-IR (ZnSe) $\nu/\text{cm}^{-1} = 3285$ (s), 2927 (m), 1651 (m), 1592 (m), 1465 (s), 1439 (s), 1375 (m), 1246 (s), 1193 (s), 1084 (w), 1044 (m), 921 (w), 822 (w), 786 (w), 768 (w), 751 (s). *m/z* (ESI+, MeCN): $\text{C}_{36}\text{H}_{32}\text{N}_2\text{O}_4$ (556.24) $[\text{M} + \text{H}]^+$ calcd: 557.24; found 557.2; $[\text{M} + \text{Na}]^+$ calcd: 579.23; found: 579.2. Elemental analysis for $\text{C}_{36}\text{H}_{32}\text{N}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (556.66 + 18.02) calc. C 75.24, N 4.87, H 5.96%; found. C 75.11, N 4.85, H 5.62%.

$[\text{HNEt}_3][\text{Dy}_2(\text{HL}^1)(\text{L}^1)]$ (5). To a solution of H_4L^1 (100 mg, 0.175 mmol, 1.00 eq.) in CH_2Cl_2 (5 mL) was added a solution of $\text{Dy}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ (88 mg, 0.192 mmol, 1.1 eq.) in MeOH (5 mL) followed by triethylamine (0.110 mL, 0.787 mmol, 4.5 eq.). The resulting mixture was stirred for 12 h, the CH_2Cl_2 was evaporated under vacuum to give a solid, which was filtered and washed with cold MeOH to give the title compound as a yellow powder. Yield: 79%. m.p. 243 °C. ATR-IR (ZnSe) $\nu/\text{cm}^{-1} = 3439$ (m), 3059 (w), 2912 (m), 1635 (s), 1591 (m), 1544 (m), 1461 (s), 1327 (s), 1301 (s), 1192 (m), 1155 (m), 1083 (m), 903 (m), 870 (w), 757 (s), 515 (w). $\text{C}_{80}\text{H}_{75}\text{N}_3\text{O}_{10}\text{Dy}_2$ (1555.396) $[\text{M}]^+$ calcd: 1461.29; found 1461.25; $[\text{M}]^-$ calcd 730.14; found 730.12. Elemental analysis for $\text{C}_{80}\text{H}_{75}\text{Dy}_2\text{N}_3\text{O}_{10} \cdot 5\text{H}_2\text{O}$ (1563.49 + 90.08) calc. C 58.11, N 2.54, H 5.18%; found. C 57.93, N 2.42, H 5.31%.

$[\text{Tb}_2(\text{L}^2)_2]$ (7). To a solution of H_3L^2 (100 mg, 0.180 mmol, 1.00 eq.) in CH_2Cl_2 (5 mL) was added a solution of Tb (NO_3)₃·6H₂O (90 mg, 0.198 mmol, 1.1 eq.) in MeOH (5 mL) followed by triethylamine (0.087 mL, 0.629 mmol, 3.5 eq.). The resulting mixture was stirred for 12 h, the CH_2Cl_2 was evaporated under vacuum to give a solid, which was filtered and washed with cold MeOH to give the title compound 7 as a yellow powder. Yield: 85%. m.p. 233 °C (decomp.). ATR-IR (ZnSe) $\nu/\text{cm}^{-1} = 3049$ (w), 2911 (w), 1652 (w), 1584 (m), 1457 (s), 1419 (m), 1329 (m), 1307 (m), 1244 (m), 1186 (m), 1082 (m), 1043 (w), 933 (w), 870 (w), 830 (w), 760 (s), 516 (w). *m/z* (ESI+, $\text{CH}_2\text{Cl}_2/\text{MeCN}$): $\text{C}_{72}\text{H}_{58}\text{N}_4\text{O}_8\text{Tb}_2$ (1425.13) $[\text{M} + \text{Na}]^+$ calcd: 1447.27; found 1447.29, $[\text{M} + \text{K}]^+$ calcd: 1463.24; found: 1463.27, $[\text{M} + \text{H}]^+$ calcd: 713.14; found 713.16. Elemental analysis for $\text{C}_{72}\text{H}_{58}\text{N}_4\text{O}_8\text{Tb}_2 \cdot \text{H}_2\text{O}$ (1425.13 + 18.02) calc. C 59.92; N 3.88; H 4.19%; found C 59.83; N 3.70; H 4.07%.

$[\text{Dy}_2(\text{L}^2)_2]$ (8). To a solution of H_3L^2 (100 mg, 0.180 mmol, 1.00 eq.) in CH_2Cl_2 (5 mL) was added a solution of Dy (NO_3)₃·6H₂O (90 mg, 0.198 mmol, 1.1 eq.) in MeOH (5 mL) followed by triethylamine (0.087 mL, 0.629 mmol, 3.5 eq.). The resulting mixture was stirred for 12 h, the CH_2Cl_2 was evaporated under vacuum to give a solid, which was filtered and washed with cold MeOH to give complex 8 as a green-yellow



powder. Yield: 87%. m.p. 236 °C (decomp.). ATR-IR (ZnSe) ν/cm^{-1} = 3055 (w), 2915 (w), 1655 (m), 1588 (m), 1458 (s), 1423 (m), 1327 (m), 1306 (m), 1245 (w), 1189 (m), 1084 (m), 1042 (w), 913 (w), 868 (w), 830 (w), 756 (s), 516 (w). m/z (ESI+, $\text{CH}_2\text{Cl}_2/\text{MeCN}$): $\text{C}_{72}\text{H}_{58}\text{N}_4\text{O}_8\text{Dy}_2$ (1432.28) $[2\text{M} + \text{H}]^+$ calcd: 1433.28; found 1433.22, $[2\text{M} + \text{Na}]^+$ calcd: 1455.27; found 1455.20, $[\text{M} + \text{H}]^+$ calcd: 717.14; found 717.1, $[\text{M} + \text{Na}]^+$ calcd: 740.13; found 740.1. Elemental analysis for $\text{C}_{72}\text{H}_{58}\text{N}_4\text{O}_8\text{Dy}_2 \cdot \text{H}_2\text{O}$ (1432.28 + 18.02) calc. C 59.63; N 3.86; H 4.17%; found C 59.72; N 3.96; H 3.99%.

X-ray crystallography

Single crystals of $[\text{Tb}_2(\text{L}^2)_2] \cdot 4\text{MeCN}$ (7·4MeCN) and $[\text{Dy}_2(\text{L}^2)_2] \cdot 4\text{MeCN}$ (8·4MeCN) were obtained from acetonitrile. Suitable specimens were selected and mounted on the tip of a glass needle using perfluoropolyether oil. The data sets were collected at 180(2) K using a STOE Stadivari diffractometer equipped with graphite monochromated Cu-K α radiation (λ = 1.54186 Å). The data were processed with the program XAREA.⁶² The structure was solved by direct methods⁶³ and refined by full-matrix least-squares techniques on the basis of all data against F^2 using SHELXL-2014/7.⁶⁴ PLATON was used to search for higher symmetry.⁴⁵ Two of the four MeCN solvate molecules in the structures of 7·4MeCN and 8·4MeCN were found to be heavily disordered and all attempts to model this disorder failed. The corresponding electron density was removed from the structures (and the corresponding F_0) with the Squeeze algorithm implemented in the Platon program suite. The SQUEEZE routine revealed total potential solvent area of 150 Å³ per unit cell for 7·4MeCN and 8·4MeCN corresponding to approximately two MeCN solvate molecules per formula unit. All non-hydrogen atoms were refined anisotropically. Graphics were produced with Ortep3 for Windows and PovRAY.

Crystal data for $[\text{Tb}_2(\text{L}^2)_2] \cdot 4\text{MeCN}$ (7·4MeCN). $\text{C}_{72}\text{H}_{58}\text{N}_4\text{O}_8\text{Tb}_2 \cdot 4\text{MeCN}$, M = 1589.31 g mol⁻¹, orthorhombic space group *Pccn*, a = 16.8317(4), b = 17.0263(5), c = 22.5381(6) Å, V = 6459.0(3) Å³, Z = 4 (the asymmetric units contains one half of the formula unit), ρ_{calc} = 1.55 g cm⁻³, μ = 11.132 mm⁻¹, 69 266 reflections collected, 6114 unique. Final R_1 [$F^2 > 2\sigma(F^2)$] = 0.0719, wR_2 (all data) = 0.1835.

Crystal data for $[\text{Dy}_2(\text{L}^2)_2] \cdot 4\text{MeCN}$ (8·4MeCN). $\text{C}_{72}\text{H}_{58}\text{Dy}_2\text{N}_4\text{O}_8 \cdot 4\text{MeCN}$, M = 1596.39 g mol⁻¹, orthorhombic space group *Pccn*, a = 16.8208(5), b = 16.9909(4), c = 22.5406(7) Å, V = 6442.1(3) Å³, Z = 4 (the asymmetric units contains one half of the formula unit), ρ_{calc} = 1.561 g cm⁻³, μ = 12.775 mm⁻¹, 69 509 reflections collected, 6070 unique. Final R_1 [$F^2 > 2\sigma(F^2)$] = 0.0713, wR_2 (all data) = 0.2067.

CCDC 2011007 (7) and 2011008 (8)[†] contain the supplementary crystallographic data for this paper.

SQUID magnetometry

Magnetic susceptibility measurements were carried out on a MPMS Quantum Design SQUID-magnetometer in the temperature range of 2–300 K. DC measurements were performed from 2–300 K in an applied external field of 0.5 T. AC measurements

were collected in a 3.0 Oe ac field at various frequencies (10–1000 Hz) with an applied 3 kOe dc field.

Conflicts of interest

There are no conflicts of interest to declare.

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